New Series

LANDOLT-BÖRNSTEIN

Numerical Data and Functional Relationships in Science and Technology

GROUP IV

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Physical Chemistry

Electrochemistry

SUBVOLUME A

Electrochemical Thermodynamics and Kinetics

Springer

Landolt-Börnstein
Numerical Data and Functional Relationships in Science and Technology
New Series / Editor in Chief: W. Martienssen

Group IV: Physical Chemistry
Volume 9

Electrochemistry

Subvolume A

Electrochemical Thermodynamics and Kinetics

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Springer

ISSN 1615-2018 (Physical Chemistry)

ISBN 978-3-540-41038-6 Springer Berlin Heidelberg New York

Library of Congress Cataloging in Publication Data
Zahlenwerte und Funktionen aus Naturwissenschaften und Technik, Neue Serie
Editor in Chief: W. Martienssen
Vol. IV/9A: Editor: M.D. Lechner
At head of title: Landolt-Börnstein. Added t.p.: Numerical data and functional relationships in science and technology.
Tables chiefly in English.
Intended to supersede the Physikalisch-chemische Tabellen by H. Landolt and R. Börnstein of which the 6th ed. began publication in 1950 under title: Zahlenwerte und Funktionen aus Physik, Chemie, Astronomie, Geophysik und Technik.
Vols. published after v. 1 of group I have imprint: Berlin, New York, Springer-Verlag
Includes bibliographies.
1. Physics--Tables. 2. Chemistry--Tables. 3. Engineering--Tables.
I. Börnstein, R. (Richard), 1852-1913. II. Landolt, H. (Hans), 1831-1910.
III. Physikalisch-chemische Tabellen. IV. Title: Numerical data and functional relationships in science and technology.
QC61.23 502'.12 62-53136

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Cover layout: Erich Kirchner, Heidelberg
Typesetting: Author and Redaktion Landolt-Börnstein, Darmstadt
Printing and binding: AZ Druck, Kempten

SPIN: 1050 1024 63/3020 - 5 4 3 2 1 0 – Printed on acid-free paper

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Preface of the editor

Electrochemistry deals with the nature of electrolytes and with the processes at electrodes which are in contact with electrolytes. Generally speaking chemistry is electrochemistry as interactions in the atomic scale are of electrical nature.
Electrochemistry is the basis of many important and modern applications and scientific developments such as nanoscale machining (fabrication of miniature devices with three dimensional control in the nanometer scale), electrochemistry at the atomic scale, scanning tunneling microscopy, transformation of energy in biological cells, selective electrodes for the determination of ions, and new kinds of electrochemical cells, batteries and fuel cells.
All these developments need precise values of the electrochemical systems. Springer Verlag and the editor are very grateful to the author R. Holze who collects and presents all relevant data together with a precise description of the phenomenons in this volume. The data are divided in five chapters:
1 Electrode potentials (potentials between an electrode and the normal hydrogen electrode (NHE))
2 Cell voltages (electromotive force of a complete cell)
3 Electrode potentials of zero charge
4 Free enthalpy of adsorption
5 Exchange current densities
Springer-Verlag plans to publish a second volume on electrochemistry LB IV/9B which includes mainly electrical conductivities and equilibria of electrochemical systems.
The complete volume, including links to the original citations, is also available online: you can either navigate through the electronic version of this volume starting from the Landolt-Börnstein website www.landolt-boernstein.com; simply select and jump directly into the PDF data file of interest and then search for entries by the full text search engine; from this website you can also access the content via a substance index or a comprehensive index. Alternatively you can also search for relevant keywords via the full Springer search gate on www.springerlink.com leading to cross-publication hits. Constraints to the search will specify your results, particularly when combined with the navigation facility. You can find a brief description of the different entrances to Landolt-Börnstein online on the inner backcover page of this volume.
The editor wishes to express his thanks to the author R. Holze for this excellent volume which appears nearly fifty years after the corresponding volume of Landolt-Börnstein's 6th edition, "Elektrische Eigenschaften II (Elektrochemische Systeme)".
The encouraging and never ending support of the editor in chief W. Martienssen and of R. Poerschke, M. Müller, D. Rathgeber-Manns and H. Hämmer from Springer-Verlag is kindly acknowleged. The publisher and the editor are confident that this volume will increase the use of the "Landolt-Börnstein".

Osnabrück, May 2007 .....

Preface of the author

Data on the structure and the dynamics, the thermodynamics and the kinetics of phase boundaries in the presence of charged species and electric fields form the numerical basis of electrochemistry. They are of essential importance in chemistry, physics, materials science and numerous related fields of fundamental and applied science. This volume presents a collection of electrode potentials established when an electronically conducting material is brought into contact with an ionically conducting phase. Since single electrode potentials cannot be measured easily (for a thorough discussion see [90Tra]) the use of a reference electrode is a prerequisite for the measurement of electrode potentials. A list of reference electrode potentials is therefore included (see also [61Ive]). Sometimes the obtained voltage (i.e. the potential difference between the two electrode potentials) which should subsequently be converted into an electrode potential given the knowledge of the precise value of the electrode potential of one electrode (commonly called the reference electrode) contains unknown contributions from further local changes of potential in the system under investigation. In many cases the exact composition of the electrochemical system is hard to establish precisely. For these systems cell voltages are reported. Although recommendations provided elsewhere regarding the use of the terms voltage and potential remain inconclusive at best [96IUP] in this compilation the potential difference measured between two electrodes is always called voltage, the potential of an electrode given with respect to a reference electrode (preferably the standard hydrogen electrode) is called electrode potential. Certainly this might also be called a voltage because it is actually also a potential difference with one electrode potential set arbitrarily to zero. Nevertheless the daily use recommends assignment of voltage to cells and potential to electrodes. No distinction is mady anymore between reversible and irreversible electrode potentials [60LB]. The usage of the term "reversible" in the context of electrochemistry is a confused one. Frequently it is applied to electrode kinetics, i.e. fast electrode reactions showing particular features in e.g. cyclic voltammograms are called reversible. Closely related is the use as a term to characterize a reference electrode. A reversible electrode is an electrode which returns quickly to its thermodynamically correct rest potential after current has been drawn across the interface whereas an irreversible electrode shows a long time delay until the correct potential is established again (if at all) [61Ive]. Elsewhere reactions proceeding on the forward and backward reaction pathway according to exactly the same mechanism just in opposite directions are called reversible. Exactly (and thermodynamically) speaking the term applies only to a system being in its equilibrium state. Thus electrode potentials and cell voltages listed below are reversible because they are measured at zero current, i.e. in the equlibrium state. Because this condition is obvious such a remark would be redundant according to current standards. The term electromotive force (elektromotorische Kraft) has been defined and explained in classical terms [57Kor] as the voltage measured between the metallic leads (contacts) of a galvanic cell at infinitely small current. Already in a standard textbook published somewhat later [70Boc] this term has vanished completely, nevertheless it has been used frequently even in most recent publications listed and quoted below. It is nevertheless not used in this compilation anymore.

Because of the interaction between the electrode surface and the various species present in the ionic phase an excess of charged species may be present on the ionically conducting side of the phase boundary. Its extent depends upon the actual electrode potential. At a certain electrode potential this excess vanishes, the corresponding potential value is called potential of zero charge. Knowledge of this

value is important for the understanding of many electrochemical phenomena. A list of known values is therefore included.

Closely connected to this aspect is the process of adsorption at electrode surfaces, which is a competetive process like at any other solid/liquid surface where a solvent molecule is replaced by an adsorbate molecule. Compared to common adsorption processes at solid/liquid interfaces the strong electric field at electrochemical phase boundaries adds an influence of utmost importance. The strength of the interaction between the electrode and the adsorbed species can be expressed by the free enthalpy of adsorption. Employing a variety of methods many data have been reported, they are listed in this volume.

Presumably the most important kinetic parameter used in the description of the kinetics of an electrode is the exchange current density or the almost equivalent rate constant. It indicates the speed of the heterogeneous process of charging or discharging species at the phase boundary, i.e. the charge transfer process. Its value is influenced by numerous factors of the investigated system. For both applied and fundamental aspects of electrochemical research a list of reported values should be helpful. It concludes this volume.

The present volume presents the results of an attempt to collect all relevant data as complete as possible. In a few cases original sources were not accessible despite considerable efforts. To provide a critical review and a complete reevaluation of the reported data is not the purpose of this collection. Consequently all data are reported, only conversion into standard units and with respect to standard references were made when necessary and reasonable, obvious printing errors were corrected.

The author wishes to express his deep gratitude to all friends and colleagues who contributed to this volume by providing reprints, to all librarians who helped in retrieving even the most remote journals and documents. Without their support this book would have remained unwritten. Close cooperation and helpful assistance in planning and preparing this book by the staff of the Springer-Verlag are gratefully appreciated. The continuous (and almost neverending) encouragement provided by the editor M.D. Lechner was quintessential; without his support this collection would have remained unfinished. My wife's tolerance for many hours spent in front of computer terminals and scientific papers was the essential prerequisite, the book would have been impossible without this.

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References

57Kor Kortüm, G.: Lehrbuch der Elektrochemie, Weinheim: Verlag Chemie, 1957.

60LB Landolt-Börnstein, 6. Auflage, II. Band, Eigenschaften der Materie in ihren Aggregatzuständen, 7. Teil: Elektrische Eigenschaften II (Elektrochemische Systeme), Hellwege, K.-H., Hellwege, A.M., Schäfer, K., Lax, E. (eds.), Berlin: Springer-Verlag, 1960.

61Ive Ives, D.J.G., Janz, G.J.: Reference Electrodes, New York: Academic Press, 1961.

70Boc Bockris, J.O'M., Reddy, A.K.N.: Modern Electrochemistry, Vol. I, New York: Plenum Press, 1970.

90Tra Trasatti, S.: Electrochim. Acta 35 (1990) 269.

96IUP IUPAC (ed.): Größen, Einheiten und Symbole in der Physikalischen Chemie, Weinheim: VCH, 1996.

List of symbols and abbreviations [1])

Symbols	Definitions
A	electrode surface area
AN	acetonitrile
a	activity
aq	aqueous
Bu	butyl
C	capacitance
CE	counterelectrode
c	concentration
$c^{\#}$	standard unit concentration mol l^{-1}
D	diffusion coefficient
E	electrode potential
E_A	activation energy
E_H, E_{SHE}	electrode potential vs. standard hydrogen electrode
E_{NCE}	electrode potential vs. normal calomel electrode
E_{pzc}	electrode potential of zero charge
E_{SCE}	electrode potential vs. saturated calomel electrode
E_0	electrode potential at rest (i.e. $I = 0$ mA cm^{-2})
E_{00}	standard electrode potential
$E^{0'}$	electrode potential where the conditional rate constant k_c^{θ} is determined
ΔE	difference of electrode potentials, e.g. peak potentials
EDTA	ethylenediamine tetraacetate
Et	ethanol, ethyl
en	ethylenediamine
F	Faraday constant (96464 C mol^{-1})
F	fugacity; [**F**] = [**M**]
FeC	Ferrocene/ferrocinium (reference electrode)
f	activity coefficient (also in use: γ)
I	current
I_a	anodic current
I_c	cathodic current
j	current density
j_a	anodic current density
j_c	cathodic current density
j_0	exchange current density
j_{00}	standard exchange current density
k_{app}	rate constant at a selected electrode potential different from the rest potential
k^0, k_{sh}	rate constant of an electrode reaction
k_c^{θ}	conditional rate constant

Symbols	Definitions
M	molarity (number of moles dissolved per liter (dm^3) of solution)
Me	methyl
m	molality (number of moles dissolved per kilogram of solution)
N	normal concentration unit
NCE	normal calomel electrode (see also: E_{NCE})
n.a.	not available, information not provided in the original report
n	number of electrons transferred in the cell or electrode reaction
[n-Bu_4N][PF_6]	tetra-n-butylammonium hexafluorophosphate
Ph	phenyl
Q	charge
R_{ct}	charge transfer resistance
RE	reference electrode
RT	room temperature
SCE	standard calomel electrode
sat.	saturated (solution)
sol.	solid (state)
T	temperature (in units of °C or K)
TEAC	tetraethylammonium perchlorate
TEAH	tetraethylammonium hexafluorophosphate
TEAP	tetraethylammonium phosphate
U, U_T	cell voltage U at temperature T in K when not stated otherwise
U_0	cell voltage at rest (i.e. I = 0 mA cm^{-2}) and at standard temperature or some other defined reference temperature
v	rate (velocity) of reaction
x	mole fraction, atom fraction
z	number of electrons transferred in an electrochemical reaction step
α	charge transfer coefficient
η	overpotential
θ	coverage (relative units $0 < \theta < 1$)
υ	stochiometric coefficient
|	indicates a phase boundary between a solid and a liquid phase
‖	indicates a phase boundaries between two liquid phases established by e.g. a porous diaphragm
*	symbol indicating hydrate composition

[1]) For a compilation of symbols in electrochemistry see [74Par], symbols, abbreviations and acronyms used only infrequently are defined at their place of use or in an associated footnote.

List of abbreviations of methods

Abbreviation	Method
ACP	alternating current polarography
AE	acousto-electrochemical method
BAW	bulk acoustic wave method
C	chronocoulometry
CA	contact angle method
CC	corrosion current method
CCS	cyclic current step method
CD	circular dichroic spectroelectrochemical method
CF	controlled hydrodynamic convection
CH	chronopotentiometry
CM	chronocoulometric method
CO	CO charge displacement method
CP	current-potential curves
CPS	cyclic potential step method
CS	coulostatic method
CT	cathode overpotential method
CV	cyclic voltammetry
DCP	direct current polarography
DCVA	derivative cyclic voltabsorptometry
DE	desorption potential
DME	dropping mercury electrode
DR	double layer repulsion method
EB	film electrode bending method
EC	electrocapillary method
ED	electrode displacement method
EE	theory of electron emission
EL	ellipsometric method
ER	electroreduction of ions
F	friction method
FD	Faradaic distortion method
FL	flow measurement
FR	Faradaic rectification
G	gravimetric method
GCI	galvanostatically controlled current interruption
GD	galvanostatic double pulse method
GM	galvanostatic method
GS	galvanostatic single pulse method
H	hardness measurement
IA	ionic adsorption method
IM	immersion method
IP	impedance method
IR	infrared spectroscopic method

Abbreviation	Method
ISS	isoelectronic shift of potential
MP	mixed potential method
NS	null solution method
OA	organic adsorption method
OC	open-circuit method
OP	oscillographic polarography
OCS	oscillographically observed current interruption
P	photoemission method
PC	electrode capacitance measured with a pulse method
PE	piezoelectric method
POL	miscellaneous polarographic methods
PP	pulse polarography
PS	potential step method
R	surface stress measurement (ribbon extension method)
RDE	rotating disc electrode
RT	radiotracer technique
S	scratching or streaming electrode method
SFG	sum frequency generation
SO	salting out method
SS	stirred solution
T	tensammetry
TC	modified charge determination method
VE	vibrating electrode method
VIB	vibrational spectroscopic method
TLV	thin-layer voltammetry
TPF	turbulent pipe flow
WL	weight loss method

References

74Par Parsons, R.: Pure Appl. Chem. **37** (1974) 499.

1 Electrode potentials

1.1 Definition of electrode potential

When a piece of metal is brought into contact with a solution containing ions of the same metal the system thus created will change towards a chemical equilibrium. Basically this involves dissolution of the metal by means of oxidation of metal atoms resulting in solvated ions and electrons left in the metal or deposition of ions from the solution on the metal. The latter process will result in an excess of positive charges on the metal compared to the initial state, the former process leads to a negatively charged metal. Since consideration of the established equilibrium using the concepts of mixed phase thermodynamics includes charged species and immiscible phases with excess electric charges the simple chemical potential has to be corrected by a contribution of the electric charge. This can be deduced most easily by setting up the corresponding equations. The cell reaction under consideration is

$$\mathrm{M}_{\mathrm{sol}}^{z+} + z\,e \rightarrow \mathrm{M}_{\mathrm{s}} \tag{1.1}$$

In the state of equilibrium between both phases, i.e. the solution phase containing the $\mathrm{M}_{\mathrm{sol}}^{z+}$ species and the solid metal phase, the sum of the chemical potentials in both phases are equal. Since charged species are involved, the usual chemical potential μ_{i} has to be extended by a term representing the work necessary to bring one mol of charged species with a charge of $z_{\mathrm{i}}\,e$ into a phase where an electrostatic potential E is present

$$\overline{\mu}_{\mathrm{i}} = \mu_{\mathrm{i}} + N_{\mathrm{A}}\,e\,z_{\mathrm{i}}\,E = \mu_{\mathrm{i}} + z_{\mathrm{i}}\,F\,E \tag{1.2}$$

The equilibrium is expressed by

$$\overline{\mu}_{\mathrm{M}} - \left(\overline{\mu}_{\mathrm{M}_{\mathrm{sol}}^{z+}} + z\,\overline{\mu}_{\mathrm{e}^-,\mathrm{M}}\right) = 0 \tag{1.3}$$

with $\overline{\mu}_{\mathrm{M}_{\mathrm{sol}}^{z+}}$ being the electrochemical potential of the ionic species in the solution phase and $\mathrm{M}_{\mathrm{sol}}^{z+}$ the electrochemical potential of the respective species in the metal. Since atoms in the metal are uncharged, $\overline{\mu}_{\mathrm{M}} = \mu_{\mathrm{M}}$. Substituting the complete equations for the electrochemical potentials in eq. (1.3) gives

$$\overline{\mu}_{\mathrm{M}} - \left[\mu_{\mathrm{sol}}^{z+} + z_{\mathrm{i}}\,F\,E_{\mathrm{sol}} + z_{\mathrm{i}}\left(\mu_{\mathrm{e}^-,\mathrm{M}} + z_{\mathrm{i}}\,F\,E_{\mathrm{M}}\right)\right] = 0 \tag{1.3}$$

The difference of electric potentials across the interface ΔE is given by

$$\Delta E = \left[1/(z\,F)\right]\left(\mu_{\mathrm{M}_{\mathrm{sol}}^{z+}} + z\,\mu_{\mathrm{e}} - \mu_{\mathrm{M}}\right) \tag{1.4}$$

The chemical potential of the ions depends on their activity a, this results in

$$\Delta E = \left[1/(zF)\right]\left(\mu^0_{M^{z+}_{sol}} + z\mu_e - \mu_M\right) + \left[RT/(zF)\right]\ln a_{M^{z+}} \tag{1.5}$$

Assuming unit ion activity the potential difference is equal to the first term; this term is called 'standard potential difference' or more commonly 'standard electrode potential'. Eq. (1.5) thus simplifies to

$$E = E_0 + \left[RT/(zF)\right]\ln a_{M^{z+}} \tag{1.6}$$

This relationship is also called Nernst equation.

By now it is obvious that the actual value of an electrode potential will critically depend both on the nature and composition of the electronically conducting material and the ionically conducting phase. Although for the sake of simplicity the former has generally been called the electrode, the latter being named the electrolyte or electrolyte solution, W. Nernst has suggested to call the combination of both an electrode. Although this definition is considerably more precise than the general usage with the term electrode referring to the electron conducting phase only, its strict usage throughout this volume might result in confusion and complication, thus the term electrode will refer to the electron conductor mostly if not explicitly stated otherwise. The ionically conducting phase will be called electrolyte solution (or melt etc.) throughout. The term „electrolyte“ only is not used; this avoids the general confusion related to the fact, that the term electrolyte refers precisely speaking only to the component of the phase providing the ions (which does not include the solvent). Considerable confusion has been caused during the data compilation of the following lists as well as during the use of compilations prepared previously and published elsewhere by the sometimes poorly defined application of the term 'standard electrode potential'. As stated in textbooks (e.g. [72Moo]) a standard electrode potential is measured with an electrode, which contains all components in the electrode under consideration (or in the whole cell when cell voltages are considered) in their standard state. The actual measurement procedure is subject of textbooks (see e.g. [97Wed, 98Hol]). Besides temperature [1]) and pressure the activity of species involved in establishing the electrode potential is obviously crucial. As defined by IEC and by e.g. the German industrial norm DIN 50900 T2 (June 1975) the species have to be present at activity $a = 1$ [2]). In other textbooks every electrode potential measured versus a standard hydrogen electrode is considered a standard potential [96Atk]. This extends to reports on electrode potentials, wherein again a measurement versus this reference qualifies a measured electrode potential as a standard value (e.g. [74Bau]), even in early tabulations [60LB]. In the tables only electrode potentials of the first type are designated standard potentials, in the latter case the electrode potential values are quoted with the respective data on composition etc..

Various listings of calculated electrode potentials and cell voltages based on known or estimated thermodynamic and further chemical data have been compiled [52Lat, 64Bet, 82Ant]. Some compilations contain a large fraction of theoretically calculated potentials without providing experimental proofs [78Mil] or are limited to aqueous solutions only [85Bar]. Experimentally determined voltages and potentials involving the use of several cells and electrode combinations to yield a final value based on the mathematical combination of these numbers are nevertheless included.

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[1]) Although the standard temperature most frequently encountered is T_0 = 25 °C or 298.15 K there are deviating definitions like e.g. the German industrial standard DIN 1343 (November 1975) stating the normal temperature as T_n=273.15 K. In the following list the former standard is always assumed if not explicitly stated otherwise.

[2]) Activity a_i is defined as $a_i = (m_i\,\gamma_i)/m_0$ with m_i being the molality of ion i with activity coefficient γ_i and standard molality m_0. Thus activity is dimensionless.

1.2 Effect of temperature and pressure on cell voltages and electrode potentials

The temperature derivative of the free energy yields an expression for the entropy change occuring during the cell reaction

$$\Delta S = n F \left(\partial U / \partial T\right)_{\mathrm{p}} \tag{1.7}$$

which can be converted into

$$\Delta H = n F \left[T \left(\partial U / \partial T\right)_{\mathrm{p}} - U \right] \tag{1.8}$$

These relationships can be used to obtain thermodynamic data otherwise difficult to get. Vice versa they can be used to calculate the temperature coefficient of a cell voltage respectively an electrode potential based on known thermodynamic data.

The pressure dependency of the cell voltage (and correspondingly the electrode potential) can also be derived using standard thermodynamic equations

$$\left(\partial U / \partial p\right)_{\mathrm{T}} = -\Delta V / \left(n F\right) \tag{1.9}$$

With solid and liquid reactants participating in the electrochemical processes leading to the establishment of a cell voltage U and the respective electrode potentials E volume changes are generally negligible. If a gas is involved, considerable effects are to be expected. Integration of eq. (1.9) yields

$$U_{\mathrm{p}} = U_{1} - \frac{1}{n F} \int_{1}^{p} \Delta V \mathrm{d}p \tag{1.10}$$

Assuming perfect gas behavior a simplified equation can be used

$$U_{\mathrm{p}} = U_{1} + \frac{RT}{2F} \ln p \tag{1.11}$$

As reference electrode commonly a hydrogen electrode is used. In its standard state, i.e. with a proton activity $a = 1$ in the electrolyte solution and standard pressure of hydrogen in the gas phase, the potential of this electrode is set to $E_{00,\,\mathrm{SHE}} = 0$ V [3]) (see also [61Ive]). When comparing the corresponding potential scale with the energy scale commonly used in physics, a difficulty arises, because in the latter scale the energy of an electron being in vacuum without any interaction is considered as reference state. Both scales can be matched [78Noz, 80Noz]. The potential of the standard hydrogen electrode thus corresponds to – 4.5 eV on the vacuum scale (for an overview see e.g. [99Tra]). The relationship between the electrode potential, the metal electron work function, an electronic term, a dipolar term and a molecular term are thoroughly reviewed elsewhere [83Tra, 90Tra].

---

[3]) Sometimes the term normal hydrogen electrode (and respectively normal potential instead of standard potential) has been used referring to a hydrogen electrode with a platinized platinum electrode immersed in 1 **M** sulfuric acid irrespectively of the actual proton activity in this solution. With the latter electrode poorly defined diffusion (liquid junction) potentials will be caused, thus data obtained with this electrode are not included. The term normal hydrogen electrode should not be used either, because it implies a reference to the concentration unit normal which is not to be used anymore, see also below.

General remarks on the tables

An electrode potential is given a positive value (i.e. no sign at all) if it is the positive terminal in a cell, whose second electrode is the standard hydrogen electrode (SHE) [4]). This is the Gibbs-Stockholm electrode potential convention. The 'Oxidation Potential' introduced by Latimer has the opposite sign [52Lat]. The tables conform with the suggestions of IUPAC compiled elsewhere [74Par, 96IUP]. All electrode reactions (potential determining processes) are given as anodic (oxidation) reactions as suggested previously [64Bet]. Because of the delicate use of arrows in reaction equations only a single arrow is used tacitly assuming that the reaction proceeds at equilibrium in both directions at the same rate. Only if explicitly stated a difference is made and reported between absolute, international, average international and US-international Volt because of the very minor differences, which might become effective only in precision measurements. The respective relationships are 1 $V_{abs.}$ = 1.00034 $V_{int.}$ = 1.00034 $V_{av.int.}$ = 1.000330 $V_{US\text{-}int.}$ [61Ive, 96IUP]. The growing use of electrochemical methods including the measurement of electrode potentials assigned to well defined redox processes in organic and inorganic chemistry has resulted in an increasing demand for tabulations of the redox (or formal) potentials E_0 of these processes. Unfortunately in most experimental studies the choice of a redox electrode seems to be determined by the ease of availability instead of compatibility with data of other authors (for a typical example see [05Vat]). Consequently a considerable amount of data has been generated, wherein the communicated electrode potential E_0 does not refer to the standard hydrogen electrode neither in an aqueous system nor in the electrolyte system under investigation, but to one of the numerous other reference electrode systems. Any conversion into the normal hydrogen scale is prone to considerable errors, thus it has to be avoided [73Str]. As a help to the researcher in this field a tabulation of data has been attempted (Table 1.5) with values of E_0 stated with reference to a specified reference electrode. Only a brief separate table of reference electrodes and their respective potentials is provided. Various systems have been employed as reference electrodes, their choice depending on the actual application. As applications have developed and broadened, the use of references has changed, too. Thus this selection is personal and of temporary value only. The reader searching for any particular reference will nevertheless find the desired system in the tables arranged according to the potential-determining element/electrode reaction, for an older review see [61Ive] [5]).

[4]) The term normal hydrogen electrode (NHE) is frequently employed as a synonym. Since normal may be taken as a concentration unit (although use of this unit is discouraged, but nevertheless numerous results are given with this unit) its use in the designation of this reference electrode should be avoided.

[5]) No more recent complete review seems to be available.

References

52Lat Latimer, W.M.: The Oxidation States of the Elements and their Potentials in Aqueous Solutions, Englewood Cliffs: Prentice-Hall Inc., 1952.

60LB Landolt-Börnstein, 6. Auflage, II. Band, Eigenschaften der Materie in ihren Aggregatzuständen, 7. Teil: Elektrische Eigenschaften II (Elektrochemische Systeme), Hellwege, K.-H., Hellwege, A.M., Schäfer, K., Lax, E. (eds.), Berlin: Springer-Verlag, 1960.

61Ive Ives, D.J.G., Janz, G.J.: Reference Electrodes, New York: Academic Press, 1961.

64Bet de Bethune, A.J., Swendeman Loud, N.A.: Standard Aqueous Electrode Potentials and Temperature Coefficients at 25 °C, Skokie: C.A. Hampel, 1964.

72Moo Moore, W.J.: Physical Chemistry, London: Longman, 1972.

73Str Strehlow, H., Knoche, W., Schneider, H: Ber. Bunsenges. Phys. Chem. **77** (1973) 760.

74Bau Baucke, F.G.K.: Chem. Ing. Techn. **46** (1974) 71.

74Par Parsons, R.: Pure Appl. Chem. **37** (1974) 499.

78Mil Milazzo, G., Caroli, S., Sharma, V.K.: Tables of Standard Electrode Potentials, Chichester: John Wiley&Sons, 1978.

78Noz Nozik, A.J.: Annu. Rev. Phys. Chem. **29** (1978) 189.

80Noz Nozik, A.J.: Philos. Trans. R. Soc. (London) A **295** (1980) 453.

82Ant Antelman, M.S., Harris jr., F.J.: The Encyclopedia of Chemical Electrode Potentials, New York: Plenum Press, 1982.

83Tra Trasatti, S.: J. Electroanal. Chem. **150** (1983) 1.

85Bar Bard, A.J., Parsons, R., Jordan, J.: Standard Potentials in Aqueous Solutions, New York: Marcel Dekker, 1985.

90Tra Trasatti, S.: Electrochim. Acta **35** (1990) 269.

96Atk Atkins, P.W.: Physikalische Chemie, Weinheim: VCH, 1996.

96IUP IUPAC (ed.): Größen, Einheiten und Symbole in der Physikalischen Chemie, Weinheim: VCH, 1996.

97Wed Wedler, G.: Lehrbuch der Physikalischen Chemie, Weinheim: Wiley-VCH, 1997.

98Hol Holze, R.: Leitfaden der Elektrochemie, Stuttgart: Teubner, 1998.

99Tra Trasatti, S., Lust, E.: Modern Aspects of Electrochemistry, Vol. 33, White, R.E., Conway, B.E., Bockris, J.O'M., (eds.) New York: Kluwer Academic/Plenum Publisher, 1999, p. 1.

05Vat Vatsadze, S., Al-Anber, M., Thiel, W.R., Lang, H., Holze, R.: J. Solid State Electrochem. **9** (2005) 764.

Table 1.1. Metal electrode potentials in aequeous electrolyte systems. [1])

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Aluminum (Al)	$Al \rightarrow Al^{3+}+3e^-$		–1.66	28Lat, 25Mak
	$Al \rightarrow Al^{3+}+3e^-$		–1.662	72Moo
	$Al+Cl^- \rightarrow AlCl^{2+}+3e^-$		–1.802	42Bro1, 43Bro
Antimony (Sb)	$2Sb+3H_2O \rightarrow Sb_2O_3+6H^++6e^-$		0.245	61Ive
	$Sb \rightarrow Sb^{III}+3e^-$ E_{SHE} [V]=0.2552–(0.05893*pH)			41Hov
	$2Sb+3H_2O \rightarrow Sb_2O_3+6H^++6e^-$, acetate buffer		0.152	52Tou
	$Sb+H_2O \rightarrow SbO^++2H^++3e^-$, acetate buffer		0.212	52Tou
	$Sb+H_2O \rightarrow SbO^++2H^++3e^-$, $SbOClO_4$, $HClO_4$		0.2040	71Vas1
	$Sb+6H_2O \rightarrow Sb_4O_6+12H^++12e^-$, $SbOClO_4$, $HClO_4$		0.1504	71Vas1
	$Sb+4Cl^- \rightarrow SbCl_4^- +3e^-$		0.17	68Noz
	$Sb+4OH^- \rightarrow SbO_2^- +2H_2O+3e^-$		–0.67	23Gru, 26Lat
Arsenic (As)	$As+2H_2O \rightarrow HAsO_2+3H^++3e^-$		0.25	52Lat
	$4As+5H_2O+½O_2 \rightarrow As_4O_6+10H^++10e^-$		0.285	61Yag
	$H_3AsO_3+H_2O \rightarrow H_3AsO_4+2H^++2e^-$		0.559	52Lat, 31Lie
Barium (Ba)	$Ba \rightarrow Ba^{2+}+2e^-$		–2.92	52Lat, 28Dev, 34Lat
	$Ba \rightarrow Ba^{2+}+2e^-$, acidic		–2.906	72Moo
$Ba_{(Hg)}$	$Ba_{(Hg)} \rightarrow Ba^{2+}+Hg+2e^-$	283.15	–1.700	76Ard
$Ba_{(Hg)}$	$Ba_{(Hg)} \rightarrow Ba^{2+}+Hg+2e^-$	298.15	–1.717	76Ard
$Ba_{(Hg)}$	$Ba_{(Hg)} \rightarrow Ba^{2+}+Hg+2e^-$	313.15	–1.728	76Ard
$Ba_{(Hg)}$	$Ba_{(Hg)} \rightarrow Ba^{2+}+Hg+2e^-$	328.15	–1.741	76Ard
$Ba_{(Hg)}$	$Ba_{(Hg)} \rightarrow Ba^{2+}+Hg+2e^-$	343.15	–1.753	76Ard
	$BaCaO_3+Pb \rightarrow Ba^{2+}+PbCO_3+2e^-$, $BaCl_2$		–2.912	62Jak2
	$Ba+ Fe(CN)_6^{3-} \rightarrow BaFe(CN)_6^- +2e^-$		–2.997	52Gib
	$Ba+ Fe(CN)_6^{4-} \rightarrow BaFe(CN)_6^{2-} +2e^-$		–3.027	56Coh
	$Ba+ P_3O_9^{3-} \rightarrow BaP_3O_9^{3-} +2e^-$		–3.011	49Jon
	$Ba+ P_4O_{12}^{4-} \rightarrow BaP_4O_{12}^{2-} +2e^-$		–3.059	50Jon
	$Ba+ S_2O_3^{2-} \rightarrow BaS_2O_3+2e^-$		–2.979	51Den
	$Ba+formate^- \rightarrow Ba\ formate^++2e^-$, Na(formate), $Ba(IO_3)_2$ sat.		–2.930	52Col1
	$Ba+acetate^- \rightarrow Ba\ acetate^++2e^-$, Na(acetate), $Ba(IO_3)_2$ sat.		–2.924	52Col1
	$Ba+bromoacetate^- \rightarrow Ba\ bromoacetate^++2e^-$		–2.919	49Dav2
	$Ba+propionate^- \rightarrow Ba\ propionate^++2e^-$, Na propionate, $Ba(IO_3)_2$ sat.		–2.916	52Col1
	$Ba+n\text{-}butyrate^- \rightarrow Ba\ n\text{-}butyrate^++2e^-$, Na n-butyrate, $Ba(IO_3)_2$ sat.		–2.912	52Col1

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Ba	$Ba+glycolate^{-}\rightarrow Ba\ glycolate^{+}+2e^{-}$, Na glycolate, Ba glycolate, $Ba(IO_3)_2$ sat.		–2.942	52Col1, 54Dav1
	$Ba+lactate^{-}\rightarrow Ba\ lactate^{+}+2e^{-}$, Na lactate, $Ba(IO_3)_2$ sat.		–2.933	52Col1, 54Dav2
	$Ba+fumarate^{2-}\rightarrow Ba\ fumarate+2e^{-}$		–2.959	54Dav1
	$Ba+malate^{2-}\rightarrow Ba\ malate+2e^{-}$		–2.971	54Dav1, 51Bel
	$Ba+maleate^{2-}\rightarrow Ba\ maleate+2e^{-}$		–2.979	54Dav1
	$Ba+tartrate^{2-}\rightarrow Ba\ tartrate+2e^{-}$		–2.987	54Dav1
	$Ba+malonate^{2-}\rightarrow Ba\ malonate+2e^{-}$		–2.968	51Pea, 49Sto
	$Ba+iso\text{-}phthalate^{2-}\rightarrow Ba\ isophthalate+2e^{-}$		–2.958	54Dav1
	$Ba+pivalate^{-}\rightarrow Ba\ pivalate^{+}+2e^{-}$, Na pivalate, $Ba(IO_3)_2$ sat.		–2.914	52Col1
	$Ba+valerate^{-}\rightarrow Ba\ valerate^{+}+2e^{-}$, Na valerate, $Ba(IO_3)_2$ sat.		–2.918	52Col1
	$Ba+adipate^{2-}\rightarrow Ba\ adipate+2e^{-}$		–2.965	54Dav1, 51Pea
	$Ba+succinate^{2-}\rightarrow Ba\ succinate+2e^{-}$		–2.968	54Dav1, 51Pea
	$Ba+glutarate^{2-}\rightarrow Ba\ glutarate+2e^{-}$		–2.972	51Pea
	$Ba+alanine\rightarrow Ba\ alanine^{2+}+2e^{-}$, alanine, $Ba(IO_3)_2$ sat.		–2.935	40Top
	$Ba+glycine\rightarrow Ba\ glycine^{2+}+2e^{-}$, glycine, $Ba(IO_3)_2$ sat.		–2.934	51Mon3
	$Ba+8\text{-}hydroxyquinoline\text{-}5\text{-}sulfonate^{2-}\rightarrow$ Ba(8-hydroxyquinoline-5-sulfonate)$+2e^{-}$, NaOH, $BaCl_2$, 8-hydroxyquinoline-5-sulfonate^{2-}		–2.980	54Näs
Beryllium (Be)	$Be\rightarrow Be^{2+}+2e^{-}$		1.7	52Lat, 27Lat
Bismuth (Bi)	$Bi\rightarrow Bi^{3+}+3e^{-}$		–1.7	52Lat, 27Lat
	$Bi\rightarrow Bi^{3+}+3e^{-}$, HCl, BiOCl		0.200	47Cup
	$Bi+H_2O\rightarrow BiO^{-}+2H^{+}+e^{-}$		0.32	23Smi, 23Swi
	$Bi+Cl^{-}+H_2O\rightarrow BiOCl+H_2+e^{-}$, $HClO_4$, HCl		0.1697	69Vas
Cadmium (Cd)	$Cd\rightarrow Cd^{2+}+2e^{-}$	288.15	–0.401	79Lon1
	$Cd\rightarrow Cd^{2+}+2e^{-}$	293.15	–0.401	79Lon1
	$Cd\rightarrow Cd^{2+}+2e^{-}$	298.15	–0.402	79Lon1
	$Cd\rightarrow Cd^{2+}+2e^{-}$	303.15	–0.402	79Lon1
	$Cd\rightarrow Cd^{2+}+2e^{-}$	308.15	–0.402	79Lon1
	$Cd\rightarrow Cd^{2+}+2e^{-}$		–0.400	46Gap, 34Shr1, 36Har1, 33LaM, 23Lew
	$Cd\rightarrow Cd^{2+}+2e^{-}$		–0.4026	58Tre1
	$Cd\rightarrow Cd^{2+}+2e^{-}$	273.15	–0.4006	36Har1
	$Cd\rightarrow Cd^{2+}+2e^{-}$	278.15	–0.4010	36Har1
	$Cd\rightarrow Cd^{2+}+2e^{-}$	283.15	–0.4013	36Har1

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	$Cd \rightarrow Cd^{2+}+2e^-$	288.15	–0.4016	36Har1
	$Cd \rightarrow Cd^{2+}+2e^-$	293.15	–0.4018	36Har1
	$Cd \rightarrow Cd^{2+}+2e^-$	298.15	–0.4020	36Har1
	$Cd \rightarrow Cd^{2+}+2e^-$	303.15	–0.4021	36Har1
	$Cd \rightarrow Cd^{2+}+2e^-$	308.15	–0.40022	36Har1
	$Cd \rightarrow Cd^{2+}+2e^-$	313.15	–0.4026	36Har1
	$Cd \rightarrow Cd^{2+}+2e^-$, acidic		–0.4029	72Moo
Cd$_{(Hg)}$	$Cd_{(Hg)} \rightarrow Cd^{2+}+Hg+2e^-$	288.15	–0.379	79Lon1
Cd$_{(Hg)}$	$Cd_{(Hg)} \rightarrow Cd^{2+}+Hg+2e^-$	293.15	–0.379	79Lon1
Cd$_{(Hg)}$	$Cd_{(Hg)} \rightarrow Cd^{2+}+Hg+2e^-$	298.15	–0.379	79Lon1
Cd$_{(Hg)}$	$Cd_{(Hg)} \rightarrow Cd^{2+}+Hg+2e^-$	303.15	–0.380	79Lon1
Cd$_{(Hg)}$	$Cd_{(Hg)} \rightarrow Cd^{2+}+Hg+2e^-$	308.15	–0.380	79Lon1
	$Cd \rightarrow Cd^{2+}+2e^-$, $CdCl_2$, $Cd(ClO_4)_2$, $HClO_4$		–0.4030	58Tre1, 58Tre2, 66Bur
Cd$_{(Hg)}$	$Cd+Br^- \rightarrow CdBr^+ +2e^-$		–0.469	57Kiv
Cd$_{(Hg)}$	$Cd+2Br^- \rightarrow CdBr_2+2e^-$		–0.492	57Kiv
Cd$_{(Hg)}$	$Cd+3Br^- \rightarrow CdBr_3^- +2e^-$		–0.482	57Kiv
Cd$_{(Hg)}$	$Cd+4Br^- \rightarrow CdBr_4^- +2e^-$		–0.489	57Kiv
Cd$_{(Hg)}$	$Cd+Cl^- \rightarrow CdCl^+ +2e^-$, NaCl, $HClO_4$, $Cd(ClO_4)_2$		–0.463	53Van
Cd$_{(Hg)}$	$Cd+2Cl^- \rightarrow CdCl_2+2e^-$, NaCl, $HClO_4$, $Cd(ClO_4)_2$		–0.483	53Van
Cd$_{(Hg)}$	$Cd+3Cl^- \rightarrow CdCl_3^- +2e^-$, NaCl, $HClO_4$, $Cd(ClO_4)_2$		–0.468	53Van
Cd$_{(Hg)}$	$2Cd+Fe(CN)_6^{4-} \rightarrow Cd_2Fe(CN)_6+2e^-$, $Li_4Fe(CN)_6$		–0.6205	57Bas
	$Cd+S_2O_3^{2-} \rightarrow CdS_2O_3+2e^-$ $CdCl_2$, BaS_2O_3		–0.520	51Den
	$Cd+2S_2O_3^{2-} \rightarrow Cd(S_2O_3)_2^{2-} +2e^-$, $CdCl_2$, $Na_2S_2O_3$		–0.594	51Den
	$Cd+glycine \rightarrow Cd\ glycine^{2+}+2e^-$, glycine, $Cd(ClO_4)_2$		–0.547	55Eva
	$Cd+2glycine \rightarrow Cd(glycine)_2^{2+}+2e^-$, glycine, $Cd(ClO_4)_2$		–0.667	55Eva
	$Cd+glycylglycine \rightarrow Cd\ glycylglycine^{2+}+2e^-$, glycylglycine, $Cd(ClO_4)_2$		–0.502	55Eva
	$Cd+2glycylglycine \rightarrow Cd(glycylglycine)_2^{2+}+2e^-$, glycylglycine, $Cd(ClO_4)_2$		–0.578	55Eva
	$Cd+glycylglycylglycine \rightarrow Cd\ glycylglycylglycine^{2+}+2e^-$, glycylglycylglycine, $Cd(ClO_4)_2$		–0.501	55Eva
	$Cd+2glycylglycylglycine \rightarrow CD(glycylglycylglycine)_2^{2+}+2e^-$, glycylglycylglycine, $Cd(ClO_4)_2$		–0.577	55Eva
	$Cd+8\text{-hydroquinoline-5-sulfonate}^{2-} \rightarrow$ Cd(8-hydroquinoline-5-sulfonate)$+2e^-$, NaOH, $Cd(ClO_4)_2$, 8-hydroquinoline-5-sulfonic acid		–0.633	54Näs
	$Cd+2(8\text{-hydroxyquinoline-5-sulfonate}^{2-}) \rightarrow$ $Cd(8\text{-hydroxyquinoline-5-sulfonate})_2^{2-}+2e^-$, NaOH, $Cd(ClO_4)_2$, 8-hydroxyquinoline-5-sulfonic acid		–0.828	54Näs

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Cd	$Cd+oxalate^{2-} \rightarrow Cd\ oxalate+2e^-$, K_2 oxalate		–0.522	51Bar
	$Cd+2oxalate^{2-} \rightarrow Cd(oxalate)_2^{2-}+2e^-$, K_2 oxalate		–0.572	51Bar
Calcium (Ca)	$Ca \rightarrow Ca^{2+}+2e^-$		–2.84	52Lat, 34Lat
	$Ca \rightarrow Ca^{2+}+2e^-$, acidic		–2.866	72Moo
Ca$_{(Hg)}$	$Ca_{(Hg)} \rightarrow Ca^{2+}+Hg+2e^-$	283.15	–1.995	75Lon2
Ca$_{(Hg)}$	$Ca_{(Hg)} \rightarrow Ca^{2+}+Hg+2e^-$	298.15	–2.003	75Lon2
Ca$_{(Hg)}$	$Ca_{(Hg)} \rightarrow Ca^{2+}+Hg+2e^-$	313.15	–2.011	75Lon2
Ca$_{(Hg)}$	$Ca_{(Hg)} \rightarrow Ca^{2+}+Hg+2e^-$	328.15	–2.018	75Lon2
Ca$_{(Hg)}$	$Ca_{(Hg)} \rightarrow Ca^{2+}+Hg+2e^-$	343.15	–2.024	75Lon2
	$CaCaO_3+Pb \rightarrow Ca^{2+}+PbCO_3+2e^-$, $CaCl_2$		–2.868	62Jak2
	$Ca+HPO_4^{2-} \rightarrow CaHPO_4+2e^-$		–2.948	53Dav
	$Ca+H_2PO_4^- \rightarrow CaH_2PO_4^+ +2e^-$		–2.900	53Dav
	$Ca+P_3O_9^{3-} \rightarrow CaP_3O_9^- +2e^-$		–2.971	49Dav1, 49Jon
	$Ca+P_4O_{12}^{4-} \rightarrow CaP_4O_{12}^{2-} +2e^-$		–3.021	49Dav1, 50Jon
	$Ca+OH^- \rightarrow CaOH^+ +2e^-$		–2.906	53Bel
	$Ca+2OH^- \rightarrow Ca(OH)_2+2e^-$, basic		–3.02	72Moo
	$Ca+SO_4^{2-} \rightarrow CaSO_4+2e^-$		–2.936	53Bel
	$Ca+S_2O_3^{2-} \rightarrow CaS_2O_3+2e^-$		–2.926	49Dav2, 55Gim
	$Ca+formate^- \rightarrow Ca\ formate^+ +2e^-$		–2.892	52Col1
	$Ca+acetate^- \rightarrow Ca\ acetate^+ +2e^-$		–2.891	52Col1
	$Ca+trimethylacetate^- \rightarrow Ca\ trimethylacetate^+ +2e^-$		–2.884	50Jon
	$Ca+methoxyacetate^- \rightarrow Ca\ methoxyacetate^+ +2e^-$		–2.901	38Dav
	$Ca+bromoacetate^- \rightarrow Ca\ bromoacetate^+ +2e^-$, $Ca(IO_3)_2$ sat., Na bromoacetate		–2.884	52Col1
	$Ca+propionate^- \rightarrow Ca\ propionate^+ +2e^-$		–2.888	52Col1
	$Ca+malate^{2-} \rightarrow Ca\ malate+2e^-$		–2.942	40Top, 51Bel
	$Ca+maleate^{2-} \rightarrow Ca\ maleate+2e^-$		–2.940	40Top
	$Ca+malonate^{2-} \rightarrow Ca\ malonate+2e^-$		–3.608	35Win
	$Ca+n\text{-}valerate^- \rightarrow Ca\ n\text{-}valerate^+ +2e^-$		–2.877	52Col1
	$Ca+iso\text{-}valerate^- \rightarrow Ca\ iso\text{-}valerate^+ +2e^-$		–2.877	52Col1
	$Ca+fumarate^{2-} \rightarrow Ca\ fumarate+2e^-$		–2.927	40Top
	$Ca+n\text{-}butyrate^- \rightarrow Ca\ n\text{-}butyrate^+ +2e^-$, $Ca(IO_3)_2$ sat., Na n-butyrate		–2.884	52Col1
	$Ca+iso\text{-}butyrate^- \rightarrow Ca\ iso\text{-}butyrate^+ +2e^-$, $Ca(IO_3)_2$ sat., Na iso-butyrate		–2.883	52Col1
	$Ca+glycolate^- \rightarrow Ca\ glycolate^+ +2e^-$		–2.915	38Dav
	$Ca+tartrate^{2-} \rightarrow Ca\ tartrate+2e^-$		–2.951	40Top
	$Ca+adipate^{2-} \rightarrow Ca\ adipate+2e^-$		–3.468	40Top

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	Ca+glutamate^{2-}→Ca glutamate+2e$^-$, $Ca(IO_3)_2$ sat., Na_2 glutamate		–2.928	50Dav
	Ca+hippurate$^-$→Ca hippurate$^+$+2e$^-$, Na hippurate		–2.881	50Dav
	Ca+mandelate$^-$→Ca mandelate$^+$+2e$^-$		–2.911	38Dav
	Ca+salicylate^{2-}→Ca salicylate+2e$^-$,NaOH, $CaCl_2$, salicylic acid		–2.881	38Dav, 51Bel
	Ca+pyruvate$^-$→Ca pyruvate$^+$+2e$^-$		–2.900	50Dav
	Ca+isophthalate^{2-}→Ca isophthalate+2e$^-$		–2.927	40Top
	Ca+succinate^{2-}→Ca succinate+2e$^-$		–2.927	40Top
	Ca+glycylglycine→Ca glycylglycine^{2+}+2e$^-$		–2.904	50Dav
	Ca+alanylglycine→Ca alanylglycine^{2+}+2e$^-$		–2.887	50Dav
	Ca+leucylglycine→Ca leucylglycine^{2+}+2e$^-$, Na leucylglycine		–2.889	50Dav
	Ca+tyrosine→Ca tyrosine^{2+}+2e$^-$		–2.912	50Dav
	Ca+serine→Ca serine^{2+}+2e$^-$, $Ca(IO_3)_2$ sat., Na serine		–2.910	50Dav
	Ca+di-iodotyrosine→Ca di-iodotyrosine^{2+}+2e$^-$, $Ca(IO_3)_2$ sat., Na di-iodotyrosine		–2.913	50Dav, 35Win
	Ca+aminoacetate→Ca aminoacetate^{2+}+2e$^-$		–2.887	52Col1, 50Dav, 38Dav
	Ca+α-aminopropionate→Ca(α-aminopropionate)$^{2+}$+2e$^-$, Na(α-aminopropionate)		–2.904	50Dav
	Ca+8-hydroxyquinoline-5-sulfonate^{2-}→ Ca(8-hydroxyquinoline-5-sulfonate)+2e$^-$, NaOH, $CaCl_2$, 8-hydroxyquinoline-5-sulfonate^{2-}		–2.972	54Näs
Cerium	Ce+Br$^-$→CeBr^{2+}+e$^-$, $Ce(ClO_4)_3$,NaBr		–2.92	51May
(Ce)	Ce+ ClO_4^- → $CeClO_4^{2+}$ +3e$^-$, $Ce(ClO_4)_3$, Na ClO_4, $HClO_4$		–2.521	55Hei
	Ce+F$^-$→CeF^{2+}+e$^-$		–2.563	53Kur
	Ce+ PO_4^{3-} →$CePO_4$+3e$^-$, $CeCl_3$, NaH_2PO_4		–2.853	50May
	Ce+ $P_2O_7^{4-}$ →$Ce(P_2O_7)^-$+3e$^-$, $CeCl_3$, NaH_2PO_4		–2.825	50May
	Ce+ SO_3^{2-} → $CeSO_3^+$ +3e$^-$, $CeCl_3$, HCl, NaH_2SO_3		–2.641	50May
	Ce+ SO_4^{2-} → $CeSO_4^+$ +3e$^-$, $Ce_2(SO_4)_3$, HCl, NaH_2SO_3		–2.555	54Spe
	Ce+(2,4-dinitrophenate)$^-$→Ce(2,4-dinitrophenate)$^{2+}$+3e$^-$		–2.504	55Bab
	Ce+picrate$^-$→Ce picrate^{2+}+3e$^-$		–2.504	55Bab
	Ce^{3+}→Ce^{4+}+e$^-$, acidic		1.61	72Moo
Cesium (Cs)	Cs→Cs^++e$^-$		–2.92	26Mai, 39Ben
	Cs→Cs^++e$^-$, acidic		–2.923	72Moo
$Cs_{(Hg)}$	$Cs_{(Hg)}$→Cs^++Hg+e$^-$	283.15	–1.919	72Mus
$Cs_{(Hg)}$	$Cs_{(Hg)}$→Cs^++Hg+e$^-$	298.15	–1.950	72Mus
$Cs_{(Hg)}$	$Cs_{(Hg)}$→Cs^++Hg+e$^-$	313.15	–1.979	72Mus
$Cs_{(Hg)}$	$Cs_{(Hg)}$→Cs^++Hg+e$^-$	328.15	–2.006	72Mus
$Cs_{(Hg)}$	$Cs_{(Hg)}$→Cs^++Hg+e$^-$	343.15	–2.035	72Mus

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Chromium (Cr)	$Cr \rightarrow Cr^{2+}+2e^-$		–0.56	27Gru
	$Cr \rightarrow Cr^{3+}+3e^-$		–0.71	52Lat
	$Cr^{2+} \rightarrow Cr^{3+}+e^-$		–0.41	27Gru
	$2Cr^{3+}+7H_2O \rightarrow Cr_2O_7^{2-}+14H^++6e^-$, $Cr_2(SO_4)_3$, H_2CrO_4, H_2SO_4		1.232	46Dig
	$2Cr^{3+}+7H_2O \rightarrow Cr_2O_7^{2-}+14H^++6e^-$		1.36	52Lat, 25Neu
	$2Cr^{3+}+7H_2O \rightarrow Cr_2O_7^{2-}+14H^++6e^-$, acidic		1.33	72Moo
	$Cr(CN)_6^{4-} \rightarrow Cr(CN)_6^{3-}+e^-$		–1.28	43Hum
	[Cr(II)-trans-1,2-cyclohexandiaminetetraacetate]$^{2-}$ → [Cr(II)-trans-1,2-cyclohexandiaminetetraacetate]$^{2-}$, NaCl, acetate buffer		1.198	70Tan
Cobalt (Co)	$Co \rightarrow Co^{2+}+2e^-$		–0.283	32Hey, 34Har1, 10Sch2
	$Co \rightarrow Co^{2+}+2e^-$, $CoSO_4$		–0.28	66Gol
	$2Co+Fe(CN)_6^{4-} \rightarrow Co_2Fe(CN)_6+4e^-$		0.2691	66Mal2
	$Co+S_2O_3^{2-} \rightarrow CoS_2O_3+2e^-$, $CoCl_2$, BaS_2O_3		–0.337	51Den
	$Co^{2+} \rightarrow Co^{3+}+e^-$, H_2SO_4		1.83	58Ver, 63Che
	$Co^{2+} \rightarrow Co^{3+}+e^-$, acidic		1.808	72Moo
	$Co(CN)_6^{4-} \rightarrow Co(CN)_6^{3-}+e^-$		–0.83	26Lie
	$2Co(CO)_4^- \rightarrow [Co(CO)_4]_2+2e^-$, $NaCo(CO)_4$, $[Co(CO)_4]_2$	293	–0.40	53Hie
	$Co_2O_3+2OH^- \rightarrow 2CoO_2+H_2O+2e^-$, NaOH, Na_2CO_3, $Na_2B_4O_7$	291	0.58	50ElW2
	$Co(OH)_2+2OH^- \rightarrow CoO_2^-+2H_2O+e^-$, KOH, $Co(OH)_2$		0.22	67Bol
	$Co+malonate^{2-} \rightarrow Co\ malonate+2e^-$, $CoCl_2$, NaH malonate		–0.380	49Sto
	$Co+2(oxalate)^{2-} \rightarrow Co(oxalate)_2^{2-}+2e^-$, Co oxalate, K_2 oxalate		–0.474	51Bar
	$Co(oxalate)_3^{4-} \rightarrow Co(oxalate)_3^{3-}+e^-$, KCl		0.57	67Hin
	$Co\ oxalate_2(H_2O)_2^{2-} \rightarrow Co\ oxalate_2(H_2O)_2^-+e^-$, KCl		0.78	67Hin
	$Co+alanine \rightarrow Co\ alanine^{2+}+2e^-$, $CoCl_2$, alanine, NaOH		–0.419	51Mon4
	$Co+2alanine \rightarrow [Co\ alanine_2]^{2+}+2e^-$, $CoCl_2$, alanine, NaOH		–0.527	51Mon4
	$Co+glycine \rightarrow Co\ glycine^{2+}+2e^-$, $CoCl_2$, glycine, NaOH		–0.431	51Mon4
	$Co+2glycine \rightarrow [Co\ glycine_2]^{2+}+2e^-$, $CoCl_2$, glycine, NaOH		–0.550	51Mon4
	$Co+glycylglycine \rightarrow Co\ glycylglycine^{2+}+2e^-$, $CoCl_2$, glycylglycine, NaOH		–0.380	51Mon4
	$Co+2glycylglycine \rightarrow [Co\ glycylglycine_2]^{2+}+2e^-$, $CoCl_2$, glycylglycine, NaOH		–0.450	51Mon4
	$Co+glycylglycylglycine \rightarrow Co(glycylglycylglycine)^{2+}+2e^-$, $CoCl_2$, glycylglycylglycine, NaOH		–0.365	55Eva
	$Co+2glycylglycylglycine \rightarrow [Co\ glycylglycylglycine_2]^{2+}+2e^-$, $CoCl_2$, glycylglycylglycine, NaOH		–0.442	55Eva

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	$Co(phenanthroline)_3^{2+} \rightarrow Co(phenanthroline)_3^{3+} + e^-$,		0.419	72Lav
	$Co(phenanthroline)_3^{3+}$, $Co(phenanthroline)_3^{2+}$, HClCo+8-hydroxyquinoline-5-sulfonate$^{2-}\rightarrow$ Co 8-hydroxyquinoline-5-sulfonate+2e$^-$, 8-hydroxyquinoline-5-sulfonate^{2-}, NaOH, $Co(NO_3)_2$		–0.537	54Näs
	Co+2(8-hydroxyquinoline-5-(sulfonate)$^{2-}$)$\rightarrow$ [Co 8-hydroxyquinoline-5-sulfonate$_2$]$^{2-}$+2e$^-$, 8-hydroxyquinoline-5-sulfonate^{2-}, NaOH, $Co(NO_3)_2$		–0.746	54Näs
	Co propylenediaminetetraacetate$^{2-}\rightarrow$ Co propylenediaminetetraacetate$^-$+e$^-$, KNO_3, Co propylenediaminetetraacetate^{2-}, Co propylenediaminetetraacetate$^-$		0.36	65Tan2
	Co trimethylenediaminetetraacetate$^{2-}\rightarrow$ Co trimethylenediaminetetraacetate$^-$+e$^-$, KNO_3, Co trimethylenediaminetetraacetate^{2-}, Co trimethylenediaminetetraacetate$^-$		0.29	65Tan2
	[Co diethylenediamine$_2$]$^{2+}\rightarrow$[Co diethylenediamine$_2$]$^{3+}$+e$^-$, KCl, [Co diethylenediamine$_2$]$^{2+}$, [Co diethylenediamine$_2$]$^{3+}$, diethylenediamine		–0.232	71Bar3
	$Co(dipyridyl)_3^{2+} \rightarrow Co(dipyridyl)_3^{3+} + e^-$, $Co(NO_3)_2$, 2,2'-dipyridyl, monochloroacetic acid+ammonia buffer		0.31	61Vyd
	$Co(EDTA)^{2-} \rightarrow Co(EDTA)^- + e^-$, KNO_3, $Co(EDTA)^-$, $Co(EDTA)^{2-}$, acetate buffer		0.37	65Tan2
	[Co ethylenediamine$_3$]$^{2+}\rightarrow$[Co ethylenediamine$_3$]$^{3+}$+e$^-$, [Co ethylenediamine$_3$]Cl_2, [Co ethylenediamine$_3$]Cl_3, KCl		0.180	69Kim
	trans-[Co glycine$_3$]$^-\rightarrow$trans-[Co glycine$_3$] +e$^-$, KCl		0.20	67Hin
Copper (Cu)	$Cu \rightarrow Cu^{2+} + 2e^-$		0.337	52Lat, 59Par
	$Cu \rightarrow Cu^{2+} + 2e^-$		0.345	31Bur, 14Lew1, 41Mül, 42Mül, 27Nie, 38Qui
	$Cu \rightarrow Cu^{2+} + 2e^-$, acidic		0.337	72Moo
	$Cu \rightarrow Cu^+ + e^-$		0.52	26Fen
	$Cu^+ \rightarrow Cu^{2+} + e^-$		0.167	26Fen, 52Lat
	$Cu^+ \rightarrow Cu^{2+} + e^-$, acidic		0.153	72Moo
	$Cu \rightarrow Cu^{2+} + 2e^-$, $CuCl_2$		0.3352	74Gro
	$Cu \rightarrow Cu^{2+} + 2e^-$, $CuSO_4$, 0.01 **M**		0.268 [5])	61Hur2
	$Cu \rightarrow Cu^{2+} + 2e^-$, $CuSO_4$, 0.1 **M**		0.291 [5])	61Hur2
	Cu$\rightarrow$Cu2++2e–, $CuSO_4$, 1 **M**		0.311 [5])	61Hur2
	$Cu \rightarrow Cu^{2+} + 2e^-$, acidic		0.3419	64Bet

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Cu	$Cu+Br^- \rightarrow CuBr^+ +2e^-$		0.343	50Näs
	$Cu+2Br^- \rightarrow Cu(Br)_2^- +e^-$, 0.1 **M** KBr, 0.1 **M** KCl		0.170	02Böd
	$Cu+CN^- \rightarrow CuCN+e^-$		–0.639	50Vla
	$Cu+2CN^- \rightarrow Cu(CN)_2^- +e^-$		–1.170	50Vla
	$Cu+4CN^- \rightarrow Cu(CN)_4^{3-} +e^-$		–1.281	50Vla
	$2Cu+ Fe(CN)_6^{4-} \rightarrow Cu_2Fe(CN)_6+4e^-$		–0.130	02Böd
	$Cu+ 2IO_3^- \rightarrow Cu(IO_3)_2+2e^-$		–0.079	48Kee, 51Llo
	$Cu+ N_3^- \rightarrow CuN_3+e^-$		0.03	52Suz1
	$Cu+NH_3 \rightarrow Cu(NH_3)^{2+}+2e^-$		–0.223	44Näs
	$Cu+3NH_3 \rightarrow Cu(NH_3)_3^{2+} +2e^-$		0.025	54Llo
	$Cu+ NO_2^- \rightarrow Cu(NO)_2^+ +2e^-$		0.305	46Kos
	$Cu+OH^- \rightarrow CuO+H_2O+2e^-$		–0.262	65Kha
	$Cu+OH^- \rightarrow CuOH^+ +2e^-$		0.153	51Dav1
	$Cu+2OH^- \rightarrow Cu(OH)_2+2e^-$		–0.222	65Kha
	$2Cu+2OH^- \rightarrow Cu_2O+H_2O+2e^-$		–0.360	65Kha
	$3Cu+ 2PO_4^{3-} \rightarrow Cu_3(PO_4)_2+6e^-$		0.325	50Bat
	$Cu+ HPO_4^{2-} \rightarrow CuHPO_4+2e^-$		0.172	50Bat
	$Cu+ 2H_2PO_4^- \rightarrow Cu(H_2PO_4)_2+2e^-$		0.204	50Bat
	$Cu+4SCN^- \rightarrow Cu(SCN)_4^{3-} +e^-$		0.020	50Kor
	$Cu+ SO_4^{2-} \rightarrow CuSO_4+2e^-$		0.274	38Owe1, 49Näs
	$Cu+ 2S_2O_3^{2-} \rightarrow Cu(S_2O_3)_2^{2-} +2e^-$		–0.022	53Lev2
	$Cu+$1-methylallylalcohol$\rightarrow$Cu 1-methylallylalcohol$+2e^-$		0.247	49Kee
	$Cu+$2-methylallylalcohol$\rightarrow$Cu 2-methylallylalcohol$+2e^-$		0.296	49Kee
	$Cu+$formate$^- \rightarrow$Cu formate$^+ +2e^-$		0.2835	51Llo
	$Cu+$acetate$^- \rightarrow$Cu acetate$^+ +e^-$		0.276	51Llo
	$Cu+\beta$-hydroxypropionate$^- \rightarrow$Cu β-hydroxypropionate$^+ +2e^-$		0.255	51Llo
	$Cu+$n-hexanoate$^- \rightarrow$Cu n-hexanoate$^+ +e^-$		0.282	51Llo
	$Cu+$fumarate$^{2-} \rightarrow$Cu fumarate$+2e^-$		0.268	51Pea
	$Cu+$2dimethylmalonate$^- \rightarrow$Cu dimethylmalonate$_2+2e^-$		0.153	30Ril
	$Cu+$glycolate$^{2-} \rightarrow$Cu glycolate$+2e^-$		0.257	51Llo
	$Cu+$gluconate$^{2-} \rightarrow$Cu gluconate$+2e^-$		–0.023	55Pec
	$Cu+2$(gluconate)$^{2-} \rightarrow$Cu gluconate$_2^{2-} +2e^-$		–0.233	53Pec
	$Cu+$malonate$^{2-} \rightarrow$Cu malonate$+2e^-$		0.203	68Bon
	$Cu+$dimethylmalonate$\rightarrow$Cu dimethylmalonate$^{2+}+2e^-$		0.199	31Ive
	$Cu+$ethylmalonate$^{2-} \rightarrow$Cu ethylmalonate$+2e^-$		0.190	30Ril
	$Cu+$maleate$^{2-} \rightarrow$Cu maleate$+2e^-$		0.226	51Pea
	$Cu+$adipate$^{2-} \rightarrow$Cu adipate$+2e^-$		0.243	51Pea

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	Cu+phenylacetate$^-$→Cu phenylacetate$^+$+2e$^-$		0.284	51Llo
	Cu+pivalate$^-$→Cu pivalate$^+$+2e$^-$		0.281	51Llo
	Cu+iso-valerate$^-$→Cu iso-valerate$^+$+2e$^-$		0.284	51Llo
	Cu+n-valerate$^-$→Cu n-valerate$^+$+2e$^-$		0.276	51Llo
	Cu+6-methylpicolinate$^-$→Cu 6-methylpicolinate$^+$+2e$^-$		0.155	55Hol
	Cu+2(6-methylpicolinate$^-$)→Cu (6-methylpicolinate)$_2$+2e$^-$		–0.010	55Hol
	Cu+succinate^{2-}→Cu succinate+2e$^-$		0.243	51Pea
	Cu+salicylaldehyde-5-sulfonate$^-$→ Cu salicylaldehyde-5-sulfonate$^+$+2e$^-$		0.284	48Cal
	Cu+2salicylaldehyde-5-sulfonate$^-$→ Cu salicylaldehyde-5-sulfonate$_2$+2e$^-$		0.067	48Cal
	Cu+bromoacetate$^-$→Cu bromoacetate$^+$+2e$^-$		0.294	51Llo
	Cu+α-bromobutyrate$^-$→Cu α-bromobutyrate$^+$+2e$^-$		0.299	51Llo
	Cu+iso-bromobutyrate$^-$→Cu iso-bromobutyrate$^+$+2e$^-$		0.278	51Llo
	Cu+n-bromobutyrate$^-$→Cu n-bromobutyrate$^+$+2e$^-$		0.216	51Llo
	Cu+chloroacetate$^-$→Cu chloroacetate$^+$+2e$^-$		0.294	51Llo
	Cu+1,2-diaminopropane→Cu 1,2-diaminopropane^{2+}+2e$^-$		0.0294	54Bas
	Cu+2(1,2-diaminopropane)→Cu 1,2-diaminopropane$_2^{2+}$ +2e$^-$		–0.268	54Bas
	Cu+1,3-diaminopropane→Cu 1,3-diaminopropane^{2+}+2e$^-$		0.051	55Pou, 54Irv
	Cu+2(1,3-diaminopropane)→Cu 1,3-diaminopropane$_2^{2+}$ +2e$^-$		–0.163	55Pou, 54Irv
	Cu+2(ethylenediamine)→Cu ethylenediamine$_2^{+}$ +e$^-$		–0.119	48Bje
	Cu+EDTA^{4-}→Cu EDTA^{2-}+2e$^-$		–0.216	53Pec
	Cu+alanine→Cu alanine^{2+}+2e$^-$		0.092	48Kee, 51Mon1
	Cu+2alanine→Cu alanine$_2^{2+}$ +2e$^-$		–0.118	48Kee, 51Mon1
	Cu+glycine→Cu glycine^{2+}+2e$^-$		0.092	48Kee, 51Mon1
	Cu+2 glycine→Cu glycine$_2^{2+}$ +2e$^-$		–0.123	48Kee, 51Mon1
	Cu+glycylglycine→Cu glycylglycine^{2+}+2e$^-$		0.163	51Mon1
	Cu+2 glycylglycine→Cu glycylglycine$_2^{2+}$ +2e$^-$		–0.004	51Mon1
	Cu+glycylglycylglycine→Cu glycylglycylglycine^{2+}+2e$^-$		0.182	55Eva
	Cu+2 glycylglycylglycine→Cu glycylglycylglycine$_2^{2+}$ +2e$^-$		0.030	55Eva
	Cu+leucine→Cu leucine^{2+}+2e$^-$		0.109	49Mal1
	Cu+2leucine→Cu leucine$_2^{2+}$ +2e$^-$		–0.083	49Mal1
	Cu+valine→Cu valine^{2+}+2e$^-$		0.095	52Reb
	Cu+2valine→Cu valine$_2^{2+}$ +2e$^-$		–0.114	52Reb
	Cu+2histidine→Cu histidine$_2^{2+}$ +2e$^-$		0.023	50Mal
	Cu+norvaline→Cu norvaline^{2+}+2e$^-$		0.085	52Reb

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Cu	Cu+2norvaline→Cu norvaline$_2^{2+}$ +2e$^-$		–0.124	52Reb
	Cu+ornithine→Cu ornithine^{2+}+2e$^-$		0.138	52Reb
	Cu+2ornithine→Cu ornithine$_2^{2+}$ +2e$^-$		–0.028	52Reb
	Cu+8-hydroxyquinoline-5-sulfonate$^-$→ Cu 8-hydroxyquinoline-5-sulfonate$^+$+2e$^-$		–0.100	49Mal1, 54Näs
	Cu+2(8-hydroxyquinoline-5-sulfonate)$^-$→ Cu 8-hydroxyquinoline-5-sulfonate$_2$+2e$^-$		–0.308	49Mal1, 54Näs
Dysprosium (Dy)	Dy^{2+}→Dy^{3+}+e$^-$		1.7	72Ten
Erbium	Er^{2+}→Er^{3+}+e$^-$		1.6	72Ten
(Er)	Er+ SO_4^{2-} → $ErSO_4^+$ +3e$^-$		–2.368	54Spe
Europium (Eu)	Eu^{2+}→Eu^{3+}+e$^-$, Eu(ClO_4)$_3$, HCOOH, $NaClO_4$		–0.36	65Mac
Gadolinium	Gd+F$^-$→GdF^{2+}+3e$^-$		–2.482	53Kur
(Gd)	Gd+ SO_4^{2-} → $GdSO_4^+$ +3e$^-$, $Gd_2(SO_4)_3$		–2.470	54Spe
Gallium	Ga→Ga^{3+}+3e$^-$, $GaCl_3$, HCl, Ga(ClO_4)$_3$, $HClO_4$		–0.560	53Sal
(Ga)	Ga→Ga^{3+}+3e$^-$		–0.52	32Sch
	Ga+Cl$^-$→GaCl^{2+}+3e$^-$, $GaCl_3$, HCl		–0.548	54Kra
	Ga+2Cl$^-$→ $GaCl_2^+$ +3e$^-$, $GaCl_3$, HCl		–0.514	54Kra
	Ga+3Cl$^-$→ $GaCl_3$+3e$^-$, $GaCl_3$, HCl		–0.471	54Kra
	Ga+4Cl$^-$→ $GaCl_4^-$ +3e$^-$, $GaCl_3$, HCl		–0.447	54Kra
	Ga+F$^-$→GaF^{2+}+3e$^-$		–0.676	52Yat
	Ga+4OH$^-$→ $GaOH_4^-$ +3e$^-$, $KGaOH_4$, KOH		–1.326	67Van
Germanium	Ge→Ge^{2+}+2e$^-$, H_2SO_4		0.24	65Rei
(Ge)	Ge→Ge^{4+}+4e$^-$, H_2SO_4		0.124	65Rei
	Ge^{2+}→Ge^{4+}+2e$^-$, H_2SO_4		0.0	65Rei
	GeO(brown)+H_2O→GeO_2(white, hexag.)+2H$^+$+2e$^-$, HCl		–0.118	52Jol
Gold	Au→Au$^+$+e$^-$		1.7	52Lat
(Au)	Au→Au^{3+}+3e$^-$		1.42	52Lat
	Au$^+$→Au^{3+}+2e$^-$		1.2	29Abh
	Au+2Br$^-$→ $AuBr_2^-$ +e$^-$, HBr		0.959	63Eva
	Au+4Br$^-$→ $AuBr_4^-$ +3e$^-$, HBr		0.854	63Eva
	$AuBr_2^-$ +2Br$^-$→ $AuBr_4^-$ +2e$^-$, HBr		0.802	63Eva
	Au+Br$^-$+OH$^-$→AuBrOH$^-$+e$^-$, KBr		0.63	72Gad
	Au+2Cl$^-$→ $AuCl_2^-$ +e$^-$, HCl		1.1540	62Lin
	Au+2Cl$^-$→ $AuCl_2^-$ +e$^-$, HCl		1.1575	68Mur2
	Au+4Cl$^-$→ $AuCl_4^-$ +3e$^-$, HCl		1.002	62Lin
	$AuCl_2^-$ +2Cl$^-$→ $AuCl_4^-$ +2e$^-$, HCl		0.926	62Lin
	Au+Cl$^-$+OH$^-$→AuClH$^-$+e$^-$, NaCl		0.75	72Gad

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	$Au+2I^- \rightarrow AuI_2^- + e^-$		0.578	69Ere
	$2Au+3H_2O \rightarrow Au_2O_3+6H^++6e^-$, H_2SO_4		1.36	66Hoa
	$Au+2NH_3 \rightarrow Au(NH_3)_2^+ + e^-$, NH_4NO_3, $[Au(NH_3)_4](NO_3)_2$		0.563	74Ski
	$Au+4NH_3 \rightarrow Au(NH_3)_4^{3+} + 3e^-$, NH_4NO_3, $[Au(NH_3)_4](NO_3)_2$		0.325	74Ski
	$Au(NH_3)_2^+ + 2NH_3 \rightarrow Au(NH_3)_4^{3+} + 2e^-$, NH_4NO_3, $[Au(NH_3)_4](NO_3)_2$		0.206	74Ski
	$Au+2OH^- \rightarrow Au(OH)_2^- + e^-$		0.4	72Gad
	$Au+2SCN^- \rightarrow Au(SCN)_2^- + e^-$		0.661	71DeC
	$Au(SCN)_2^- + 2SCN^- \rightarrow Au(SCN)_4^- + 2e^-$		0.623	66Pou
	$Au+4SCN^- \rightarrow Au(SCN)_4^- + 3e^-$, HCl, KCl, NaSCN		0.636	66Pou
	$Au+S_2O_3^{2-} \rightarrow Au(S_2O_3)_2^{3-} + e^-$, KCl, $Na_3[Au(S_2O_3)_2]$		0.153	69Pou
Hafnium (Hf)	$Hf+H_2O \rightarrow HfO^{2+}+2H^++4e^-$		–1.7	52Lat
Holmium	$Ho^{2+} \rightarrow Ho^{3+}+e^-$		1.7	72Ten
(Ho)	$Ho+SO_4^{2-} \rightarrow HoSO_4^+ + 3e^-$		–2.391	54Spe
Indium	$In \rightarrow In^+ + e^-$	288.15	–0.340	64But2
(In)	$In \rightarrow In^+ + e^-$	298.15	–0.339	64But2
	$In \rightarrow In^+ + e^-$	308.15	–0.339	64But2
	$In \rightarrow In^+ + e^-$	318.15	–0.339	64But2
	$In \rightarrow In^+ + e^-$	333.15	–0.339	64But2
	$In \rightarrow In^{3+} + e^-$		–0.34	36Hat, 30Hak
$In_{(Hg)}$	$In_{(Hg)} \rightarrow In^+ + Hg + e^-$	298.15	–0.313	64But2
	$In \rightarrow In^+ + e^-$, $InClO_4$, $Ce(ClO_4)_3$, $HClO_4$		–0.14	53Hep
	$In \rightarrow In^{3+} + 3e^-$, $InCl_3$, HCl		–0.3382	63Cav
	$In^+ \rightarrow In^{2+} + e^-$, $InClO_4$, $Ce(ClO_4)_3$, $HClO_4$		–0.40	53Hep
	$In^{2+} \rightarrow In^{3+} + e^-$, $InClO_4$, $Ce(ClO_4)_3$, $HClO_4$		–0.49	53Hep
	$In+F^- \rightarrow InF^{2+} + 3e^-$, $InClO_4$, $Fe(ClO_4)_2$, $Fe(ClO_4)_3$, NaF, $HClO_4$		–0.429	54Hep
Iridium	$Ir \rightarrow Ir^{3+} + 3e^-$		1.0	52Lat
(Ir)	$IrBr_6^{3-} \rightarrow IrBr_6^{2-} + e^-$, NaBr		0.97	47Dwy2, 45Pti
	$IrCl_6^{3-} \rightarrow IrCl_6^{2-} + e^-$, K_3IrCl_6, HCl		0.8665	57Geo
	$Ir(H_2O)Cl_5^{2-} \rightarrow Ir(H_2O)Cl_5^- + e^-$, $(NH_4)_2[Ir(H_2O)Cl_5]$, HNO_3		1.0	65Cha
	trans-$[Ir(H_2O)_2Cl_4]^- \rightarrow$trans-$[Ir(H_2O)_2Cl_4] + e^-$, $HClO_4$, NaH_2PO_4		1.22	67ElA
	$Ir(H_2O)_3Cl_3 \rightarrow [Ir(H_2O)_3Cl_3]^+ + e^-$, $HClO_4$, HNO_3		1.30	67ElA
	$IrI_6^{3-} \rightarrow IrI_6^{2-} + e^-$		0.485	45Pti

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Iron (Fe)	$Fe \rightarrow Fe^{2+}+2e^-$		–0.441	23Lew, 32Ran, 24Ric
	$Fe \rightarrow Fe^{2+}+2e^-$, acidic		–0.4402	72Moo
	$Fe \rightarrow Fe^{3+}+3e^-$, acidic		–0.036	72Moo
	$Fe \rightarrow Fe^{2+}+2e^-$, $FeCl_2$, $FeSO_4$		–0.447	53Pat, 51Bon
	$Fe \rightarrow Fe^{2+}+2e^-$, $FeCl_2$		–0.447	60Hur
	$Fe^{2+} \rightarrow Fe^{3+}+2e^-$		0.771	60Mat, 37Sch, 45Bez, 72Moo
	$3Fe^{2+}+4H_2O \rightarrow Fe_3O_4+8H^++2e^-$		1.230	60Mat
	$3Fe^{2+}+3H_2O \rightarrow FeOH_3+3H^++e^-$		0.939	60Mat
	$Fe(CN)_6^{4-} \rightarrow Fe(CN)_6^{3-}+e^-$		0.358	67Han, 68Mur1
	$Fe(CN)_6^{4-} \rightarrow Fe(CN)_6^{3-}+e^-$		0.356	35Kol
	$Fe(CN)_6^{4-} \rightarrow Fe(CN)_6^{3-}+e^-$, acidic		0.69	72Moo
	$Fe(H_2O)(CN)_5^{3-} \rightarrow Fe(H_2O)(CN)_5^{2-}+e^-$, $K_3[Fe(H_2O)(CN)_5]$		0.412	66Mal1
	$Fe(NH_3)(CN)_5^{3-} \rightarrow Fe(NH_3)(CN)_5^{2-}+e^-$, $Na_3[Fe(NH_3)(CN)_5]$,$Na_2[Fe(NH_3)(CN)_5]$, H_2SO_4		–0.378	75Mal
	$Fe(H_2O)(NO)_5^{3-} \rightarrow Fe(H_2O)(NO)_5^{2-}+e^-$, $K_3[Fe(H_2O)(CN)_5]$		–0.075	69Mas
	$3Fe(CO)_4^{2-} \rightarrow [Fe(CO)_4]_3+6e^-$, Na_2, $Ba[Fe(CO)_4]$, $[Fe(CO)_4]_3$		–0.747	53Hie
	$3FeH(CO)_4^- \rightarrow [Fe(CO)_4]_3+3H^++6e^-$, Na_2, $Ba[Fe(CO)_4]$, $[Fe(CO)_4]_3$		–0.35	53Hie
	Fe alanine^{2+}→Fe alanine^{3+}+e^-, KCl	293	0.340	59Per
	Fe glycine^{2+}→Fe glycine^{3+}+e^-, KCl	293	0.380	59Per
	Fe glycylglycine^{2+}→Fe glycylglcine^{3+}+e^-, KCl	293	0.360	59Per
	Fe hydroxyproline^{2+}→Fe hydroxyproline^{3+}+e^-, KCl	293	0.445	59Per
	Fe leucine^{2+}→Fe leucine^{3+}+e^-, KCl	293	0.362	59Per
	Fe+2 mercaptoacetate$^-$→[Fe(mercaptoacetate)$_2$]$^{2-}$+2e^-, $FeCl_2$, mercaptoacetic acid		–0.768	53Leu
	Fe methionine^{2+}→Fe methionine^{3+}+e^-, KCl	293	0.400	59Per
	Fe ornithine^{2+}→Fe ornithine^{3+}+e^-, KCl	293	0.415	59Per
	Fe nitrilotriacetate$^-$→ Fe nitrilotriacetate+e^-, acetate or veronal buffer, KCl	293	0.3295	51Sch2
	Fe α-amino-n-butyric acid^{2+}→ Fe α-amino-n-butyric acid^{3+}+e^-, KCl	293	0.370	59Per
	Fe α-amino-iso-butyric acid^{2+}→ Fe α-amino-iso-butyric acid^{3+}+e^-, KCl	293	0.345	59Per
	Fe arginine^{2+}→Fe arginine^{3+}+e^-, KCl	293	0.420	59Per
	Fe asparagine^{2+}→Fe asparagine^{3+}+e^-, KCl	293	0.440	59Per
	Fe proline^{2+}→Fe proline^{3+}+e^-, KCl	293	0.395	59Per

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	Fe (β-phenylalanine)$^{2+}$→Fe (β -phenylalanine)$^{3+}$+e$^-$, KCl	293	0.415	59Per
	Fe sarcosine^{2+}→Fe sarcosine^{3+}+e$^-$, KCl	293	0.380	59Per
	Fe serine^{2+}→Fe serine^{3+}+e$^-$, KCl	293	0.405	59Per
	Fe threonine^{2+}→Fe threonine^{3+}+e$^-$, KCl	293	0.430	59Per
	Fe tryptophan^{2+}→Fe tryptophan^{3+}+e$^-$, KCl	293	0.415	59Per
	Fe valine^{2+}→Fe valine^{3+}+e$^-$, KCl	293	0.380	59Per
	Fe aspartic acid^{2+}→Fe aspartic acid^{3+}+e$^-$, KCl	293	0.330	59Per
	Fe glutamic acid^{2+}→Fe glutamic acid^{3+}+e$^-$, KCl	293	0.240	59Per
	Fe 2 cysteine^{2-}→Fe cysteine$_2^{2-}$ +e$^-$, $FeCl_2$, cysteine		–0.793	55Tan
	Fe+dipyridyl→Fe dipyridyl^{2+}+e$^-$, 2,2'-dipyridyl, $FeCl_2$		–0.571	49Kru
	[Fe(dipyridyl)$_3$]$^{2+}$→[Fe(dipyridyl)$_3$]$^{3+}$+e$^-$, HNO_3		1.120	59Geo
	[Fe(2,2',2"-terpyridyl)$_2$]$^{2+}$→ [Fe(2,2',2"-terpyridyl)$_2$]$^{3+}$+e$^-$, H_2SO_4		–0.927	54Dwy
	Fe(1,10-phenanthroline)$^{2+}$→ Fe(1,10-phenanthroline)$^{3+}$+e$^-$, H_2SO_4		1.06	52Bra
	Fe(1,10-phenanthroline)$^{2+}$→ Fe(1,10-phenanthroline)$^{3+}$+e$^-$, HNO_3,Fe(II,III)tris-o-orthophenanthroline	298.5	1.147	59Geo
	Fe(5-methyl-1,10-phenanthroline)$^{2+}$→ Fe(5-methyl-1,10-phenanthroline)$^{3+}$+e$^-$, H_2SO_4		1.02	52Bra
	Fe(5-phenyl-1,10-phenanthroline)$^{2+}$→ Fe(5-phenyl-1,10-phenanthroline)$^{3+}$+e$^-$, H_2SO_4		1.08	52Bra
	Fe(5-chloro-1,10-phenanthroline)$^{2+}$→ Fe(5-chloro-1,10-phenanthroline)$^{3+}$+e$^-$, H_2SO_4		1.12	52Bra
	Fe(5-nitro-1,10-phenanthroline)$^{2+}$→ Fe(5-nitro-1,10-phenanthroline)$^{3+}$+e$^-$, H_2SO_4		1.25	52Bra
	Fe(4,7-dihydroxyphenanthroline)$^{3-}$→ Fe(4,7-dihydroxyphenanthroline)$^{2-}$+e$^-$		–0.20	67Han
	Fe(4,7-dimethyl-1,10-phenanthroline)$^{2+}$→ Fe(4,7-dimethyl-1,10-phenanthroline)$^{3+}$+e$^-$, $LiClO_4$		0.936	66Far
	[Fe(dipyridyl)(CN)$_4$]$^{2-}$→[Fe(dipyridyl)(CN)$_4$]$^-$+e$^-$, HCl, Fe(II,III)dipyridyl(CN)$_4$		0.541	59Geo
	[Fe(dipyridyl)$_2$(CN)$_2$]→[Fe(dipyridyl)$_2$(CN)$_2$]$^+$+e$^-$, H_2SO_4		0.776	63Sch2
	[Fe(pyridine-2-aldoxime)$_3$]$^+$→[Fe(pyridine-2-aldoxime)$_3$]$^{2+}$+e$^-$, Fe(II,III)(pyridine-2-aldoxime)$_3$, NaOH, NaCl, NaH_2PO_4	293	0.348	68Han
	[Fe(pyridine-2,6-dialdoxime)$_2$]$^{2-}$			
	[Fe(pyridine-2,6-dialdoxime)$_2$]$^{2-}$→ [Fe(pyridine-2,6-dialdoxime)$_2$]$^-$+e$^-$		0.204	66Han
	Fe(EDTA)$^{2-}$→Fe(EDTA)$^-$+e$^-$, phosphate or acetate buffer		0.1172	51Sch1
	Ferrocene→Ferrocinium+e$^-$		0.400	78Lag
	Methylferrocene→Methylferricinium$^+$+e$^-$		0.082	62Gub
	Ethylferrocene→Ethylferricinium$^+$+e$^-$		0.081	62Gub
	1-Methyl-2-ethylferrocene→1-Methyl-2-ethylferricinium$^+$+e$^-$		0.098	65Nes
	1-Methyl-3-ethylferrocene→1-Methyl-3-ethylferricinium$^+$+e$^-$		0.097	65Nes
	Propylferrocene→Propylferricinium$^+$+e$^-$		0.079	62Gub

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Fe	iso-Propylferrocene→iso-Propylferricinium^{+}+e^{-}		0.082	62Gub
	Vinylferrocene→Vinylferricinium^{+}+e^{-}		0.000	62Gub
	Methoxyferrocene→Methoxyferricinium^{+}+e^{-}		0.114	62Gub
	Acetylferrocene→Acetylferricinium^{+}+e^{-}		–0.044	62Gub
	Phenylferrocene→Phenylferricinium^{+}+e^{-}		–0.004	62Gub
	Phenylmethoxyethylferrocene→ Phenylmethoxyphenylferricinium^{+}+e^{-}		–0.022	62Gub
	Aminoferrocene hydrochloride→ Aminoferrocene^{+}hydrochloride+e^{-}		–0.280	62Gub
	m-Aminophenylferrocene→m-Aminophenylferricinium^{+}+e^{-}		–0.056	62Gub
	p-Aminophenylferrocene→p-Aminophenylferricinium^{+}+e^{-}		–0.061	62Gub
	Trimethylferrocenylmethylammonium^{+}→ Trimethylferrocenylmethylammonium^{2+}+e^{-}		0.659	66L'vo
	m-Nitrophenylferrocene→m-Nitrophenylferricinium^{+}+e^{-}		–0.078	62Gub
	p-Nitrophenylferrocene→p-Nitrophenylferricinium^{+}+e^{-}		–0.110	62Gub
	m-Phenylazophenylferrocene→ m-Phenylazophenylferricinium^{+}+e^{-}		–0.049	62Gub
	Benzoxyferrocene→Benzoxyferricinium^{+}+e^{-}		–0.056	62Gub
	(2'-Bromoaminoethyl)ferrocene→ (2'-Bromoaminoethyl)ferricinium^{+}+e^{-}		–0.022	62Gub
	Butenylferrocene→Butenylferricinium^{+}+e^{-}		0.069	62Gub
	Carbacylferrocene→Carbacylferricinium^{+}+e^{-}		–0.204	62Gub
	1-Methyl-1'-carbacylferrocene→ 1-Methyl-1'-carbacylferricinium^{+}+e^{-}		–0.179	65Nes
	1-Methyl-2-carbacylferrocene→ 1-Methyl-2-carbacylferricinium^{+}+e^{-}		–0.157	65Nes
	1-Methyl-3-carbacylferrocene→ 1-Methyl-3-carbacylferricinium^{+}+e^{-}		–0.188	65Nes
	1-Ethyl-1'-carbacylferrocene→ 1-Ethyl-1'-carbacylferricinium^{+}+e^{-}		–0.277	65Nes
	1-Ethyl-2-carbacylferrocene→ 1-Ethyl-2-carbacylferricinium^{+}+e^{-}		–0.164	65Nes
	1-Ethyl-3-carbacylferrocene→ 1-Ethyl-3-carbacylferricinium^{+}+e^{-}		–0.188	65Nes
	1-Phenyl-1'-carbacylferrocene→ 1-Phenyl-1'-carbacylferricinium^{+}+e^{-}		–0.271	65Nes
	1-Phenyl-2-carbacylferrocene→ 1-Phenyl-2-carbacylferricinium^{+}+e^{-}		–0.262	65Nes
	1-Ethyl-2-hydroxyethylferrocene→ 1-Ethyl-2-hydroxyethylferricinium^{+}+e^{-}		–0.020	65Nes
	1'-Hydroxyethylferrocene→1'-Hydroxyethylferricinium^{+}+e^{-}		0.003	62Gub
	Hydroxymethylferrocene→Hydroxymethylferricinium^{+}+e^{-}		–0.029	62Gub
	p-Carbacylphenyl ferrocene→ p-Carbacylphenyl ferricinium^{+}+e^{-}		–0.111	62Gub
	Carbomethoxyferrocene→Carbomethoxyferricinium^{+}+e^{-}		–0.231	62Gub

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	Carboxyferrocene→Carboxyferricinium$^+$+e$^-$		–0.219	62Gub
	Carboxmethylferrocene→Carboxymethylferricinium$^+$+e$^-$		–0.008	62Gub
	Chloroferrocene→Chloroferricinium$^+$+e$^-$		–0.128	62Gub
	Chloromercuryferrocene→Chloromercuryferricinium$^+$+e$^-$		–0.002	62Gub
	Bromoferrocene→Bromoferricinium$^+$+e$^-$		–0.130	62Gub
	Iodoferrocene→Iodoferricinium$^+$+e$^-$		–0.104	62Gub
	(2'-Dimethylaminoethyl)ferrocene→ (2'-Dimetylaminoethyl)ferricinium$^+$+e$^-$		–0.016	62Gub
	(2-Dimethylaminoethyl)ferrocene→ (2-Dimetylaminoethyl)ferricinium$^+$+e$^-$		–0.165	62Gub
Lanthanum (La)	La→La^{3+}+3e$^-$		–2.4	52Lat
	La+ Co(CN)$_6^{3-}$ →LaCo(CN)$_6$+3e$^-$		–2.597	47Jam, 50Jam
	La+F$^-$→LaF^{2+}+3e$^-$		–2.592	64Rao
	La+ Fe(CN)$_6^{3-}$ →LaFe(CN)$_6$+3e$^-$		–2.594	48Dav
	La+ Fe(CN)$_6^{3-}$ →LaFe(CN)$_6$+3e$^-$, LaFe(CN)$_6$, glycine		–2.594	51Dun
	La+ Fe(CN)$_6^{3-}$ →LaFe(CN)$_6$+3e$^-$, LaFe(CN)$_6$		–2.594	53Ber2
	La+ P$_3$O$_9^{3-}$ →LaP$_3$O$_9$+3e$^-$		–2.636	49Mon3
	La+ P$_4$O$_{12}^{4-}$ →La(P$_4$O$_{12}$)$^-$+3e$^-$		–2.655	49Mon3
	La+ SO$_4^{2-}$ → LaSO$_4^+$ +3e$^-$, La$_2$(SO$_4$)$_3$		–2.596	50Jen, 52Jon, 54Spe
	La+adipate^{2-}→La adipate$^+$+3e$^-$, NaH adipate, LaCl$_3$		–2.604	51Pea
	La+fumarate^{2-}→La fumarate$^+$+3e$^-$, NaH fumarate, LaCl$_3$		–2.582	51Pea
	La+glutarate^{2-}→La glutarate$^+$+3e$^-$, NaH glutarate, LaCl$_3$		–2.598	51Pea
	La+malonate^{2-}→La malonate$^+$+3e$^-$, NaH malonate, LaCl$_3$		–2.620	51Pea
	La+phthalate^{2-}→La phthalate$^+$+3e$^-$, NaH phthalate, LaCl$_3$		–2.633	51Pea
	La+picrate$^-$→ La picrate^{2+}+3e$^-$		–2.543	55Bab
	La+2(picrate)$^-$→La picrate$_2^+$ +3e$^-$		–2.586	55Bab
	La+succinate^{2-}→La succinate$^+$+3e$^-$, NaH succinate, LaCl$_3$		–2.601	51Pea
Lead (Pb)	Pb→Pb^{2+}+2e$^-$		–0.126	61Ive, 38Har1, 22Ger, 30Ger, 30Ran, 38Lin, 39Har1, 72Moo
	Pb→Pb^{2+}+2e$^-$, PbCl$_2$		–0.1262	67Tou
	Pb→Pb^{2+}+2e$^-$, Pb(ClO$_4$)$_2$, NaClO$_4$		–0.1262	37Far
	Pb→Pb^{2+}+2e$^-$, Pb(ClO$_4$)$_2$, HClO$_4$		–0.1237	71Vas2
	Pb^{2+}→Pb^{4+}+2e$^-$, Pb(HCOO)$_2$, Pb(HCOO)$_4$, HClO$_4$		1.67	62Ber
	Pb+Br$^-$→PbBr$^+$+2e$^-$, PbBr$_2$		–0.174	55Nan

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Pb	$Pb+Br^- \rightarrow PbBr^+ + 2e^-$, KBr, $Pb(ClO_4)_2$		–0.174	55Big
	$Pb+CO_3^{2-} \rightarrow PbCO_3+2e^-$, basic		–0.509	72Moo
	$Pb+2Cl^- \rightarrow PbCl_2+2e^-$, KCl		–0.2675	64Alt
	$Pb+2OH^- \rightarrow PbO+H_2O+2e^-$, basic		–0.577	64Bet
	$Pb+3OH^- \rightarrow HPbO_2^- + H_2O+2e^-$, NaOH		–0.537	54Vle
	$Pb+SO_4^{2-} \rightarrow PbSO_4+2e^-$	273	–0.3335	61Ive
	$Pb+SO_4^{2-} \rightarrow PbSO_4+2e^-$	285.5	–0.3448	61Ive
	$Pb+SO_4^{2-} \rightarrow PbSO_4+2e^-$	298	–0.3563	61Ive
	$Pb+SO_4^{2-} \rightarrow PbSO_4+2e^-$	310.5	–0.3680	61Ive
	$Pb+SO_4^{2-} \rightarrow PbSO_4+2e^-$	323	–0.3802	61Ive
$Pb_{(Hg)}$	$Pb+SO_4^{2-} \rightarrow PbSO_4+2e^-$	273	–0.3281	61Ive
	$Pb+SO_4^{2-} \rightarrow PbSO_4+2e^-$	285.5	–0.3392	61Ive
	$Pb+SO_4^{2-} \rightarrow PbSO_4+2e^-$	298	–0.3505	61Ive
	$Pb+SO_4^{2-} \rightarrow PbSO_4+2e^-$	298	–0.351	35Har, 34Shr2
	$Pb+SO_4^{2-} \rightarrow PbSO_4+2e^-$	310.5	–0.3619	61Ive
	$Pb+SO_4^{2-} \rightarrow PbSO_4+2e^-$	323	–0.3738	61Ive
Lithium (Li)	$Li \rightarrow Li^+ + e^-$		–3.00	52Lat, 31Lar, 31Lew, 23Lew
	$Li \rightarrow Li^+ + e^-$		–3.0388	83Maz
	$Li \rightarrow Li^+ + e^-$, acidic		–3.045	72Moo
	$Li \rightarrow Li^+ + e^-$	283.15	–3.034	77Ron
	$Li \rightarrow Li^+ + e^-$	298.15	–3.042	77Ron
	$Li \rightarrow Li^+ + e^-$	313.15	–3.049	77Ron
	$Li+glycolate^{2-} \rightarrow Li(glycolate)^- + e^-$	288	–3.033	54Dav1
	$Li+lactate^{2-} \rightarrow Li(lactate)^- + e^-$	288	–3.052	54Dav2
$Li_{(Hg)}$	$Li_{(Hg)} \rightarrow Li^+ + Hg + e^-$, LiOH	288	–3.0401	68Hus
$Li_{(Hg)}$	$Li_{(Hg)} \rightarrow Li^+ + Hg + e^-$	283.15	–2.184	77Lon2
$Li_{(Hg)}$	$Li_{(Hg)} \rightarrow Li^+ + Hg + e^-$	298.15	–2.195	77Lon2
$Li_{(Hg)}$	$Li_{(Hg)} \rightarrow Li^+ + Hg + e^-$	313.15	–2.205	77Lon2
Magnesium (Mg)	$Mg \rightarrow Mg^{2+} + 2e^-$		–2.38	45Coa
	$Mg \rightarrow Mg^{2+} + 2e^-$, acidic		–2.363	72Moo
	$Cd+Mg(OH)_2 \rightarrow Cd(OH)_2+Mg^{2+}+2e^-$, cadmium electrode, $MgCl_2$		–2.372	61Jak1
$Mg_{(Hg)}$	$Mg_{(Hg)} \rightarrow Mg^{2+} + Hg + 2e^-$	298.15	–1.980	75Lon2
	$Mg+Fe(CN)_6^{4-} \rightarrow MgFe(CN)_6^{2-} + 2e^-$		–2.478	56Coh

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	$Mg + IO_3^- \rightarrow MgIO_3^+ + 2e^-$		–2.384	30Dav, 34Kil, 38Wis
	$Mg + OH^- \rightarrow MgOH^+ + 2e^-$		–2.440	48Sto
	$Mg + PO_4^{3-} \rightarrow MgPO_4^- + 2e^-$		–2.437	48Sto
	$Mg + P_4O_{12}^{4-} \rightarrow MgP_4O_{12}^{2-} + 2e^-$		–2.515	50Jon
	$Mg + SO_4^{2-} \rightarrow MgSO_4 + 2e^-$		–2.430	51Dun, 52Jon
	$Mg + lactate^{2-} \rightarrow Mg\ lactate + 2e^-$, Mg lactate, HCl		–2.404	54Dav2
	$Mg + malonate^{2-} \rightarrow Mg\ malonate + 2e^-$, Mg malonate, HCl		–2.447	52Eva
	$Mg + oxalate^{2-} \rightarrow Mg\ oxalate + 2e^-$, Mg oxalate, K_2 oxalate		–2.493	51Bar
	$Mg + glycolate^{2-} \rightarrow Mg\ glycolate + 2e^-$, Mg glycolate, HCl		–2.402	54Dav1
	$Mg + glycine \rightarrow Mg\ glycine^{2+} + 2e^-$, $MgCl_2$, NaOH, glycine		–2.465	51Mon2
	$Mg + glycylglycine \rightarrow Mg\ glycylglycine^{2+} + 2e^-$, $MgCl_2$, NaOH, glycylglycine		–2.394	51Mon2
	$Mg + (8\text{-hydroxyquinoline-5-sulfonate})^{2-} \rightarrow Mg(8\text{-hydroxyquinoline-5-sulfonate}) + 2e^-$, $MgCl_2$, NaOH, 8-hydroxyquinoline-5-sulfonate^{2-}		–2.465	54Näs
Manganese (Mn)	$Mn \rightarrow Mn^{2+} + 2e^-$		–1.05	23Cam1, 52Lat, 35San
	$Mn^{2+} \rightarrow Mn^{3+} + 1e^-$, $Mn(ClO_4)_2$, $HClO_4$		1.5415	65Cia
	$Mn^{2+} \rightarrow Mn^{4+} + 2e^-$	291	1.465	23Cam2
	$Mn^{2+} + 2H_2O \rightarrow MnO_2 + 4H^+ + 2e^-$		1.224	65Tan1
	$Mn^{2+} + 2H_2O \rightarrow MnO_2 + 4H^+ + 2e^-$, $Mn(ClO_4)_2$, $HClO_4$		1.224	47Wit, 55Max
	$Mn^{2+} + 2H_2O \rightarrow MnO_2 + 4H^+ + 2e^-$, acidic		1.23	72Moo
	$Mn^{2+} + 2H_2O \rightarrow MnO_2 + 4H^+ + 2e^-$, $MnCl_2$		1.224	62Cov
	$Mn^{2+} + 2H_2O \rightarrow MnO_2 + 4H^+ + 2e^-$		1.236	30Bro
	$Mn + P_3O_9^{3-} \rightarrow Mn(P_3O_9)^{3-} + 2e^-$		–1.291	49Jon
	$Mn + P_4O_{12}^{4-} \rightarrow Mn(P_4O_{12})^{2-} + 2e^-$		–1.355	50Jon
	$Mn + S_2O_3^{2-} \rightarrow MnS_2O_3 + 2e^-$, $MnCl_2$, BaS_2O_3		–1.243	51Den
	$Mn + S_2O_4^{2-} \rightarrow MnSO_4 + 2e^-$, $MnCl_2$, BaS_2O_3		–1.252	47Jam
	$2Mn(CO)_5^- \rightarrow [Mn(CO)_5]_2 + 2e^-$, $NaMn(CO)_5$		–0.68	58Hie
	$MnO_2 + 2H_2O \rightarrow MnO_4^- + 4H^+ + 3e^-$		1.679	65Tan1
	$MnO_4^{2-} \rightarrow MnO_4^- + 1e^-$, KOH, K_2MnO_4, $KMnO_4$		0.558	56Car2
	$MnO_4^{3-} \rightarrow MnO_4^{2-} + e^-$, NaOH,$NaClO_4$, $KMnO_4$		0.274	67Sch2
	$Mn + malonate^{2-} \rightarrow Mn(malonate) + 2e^-$, Mg(malonate), HCl		–1.282	51Den
	$Mn + alanine \rightarrow Mn(alanine)^{2+} + 2e^-$, $MnCl_2$, alanine, NaOH		–1.274	51Mon4
	$Mn + glycyne \rightarrow Mn(glycyne)^{2+} + 2e^-$, $MnCl_2$, glycyne, NaOH		–1.286	51Mon4

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Mn	Mn+glycylglycyne→Mn(glycylglycyne)$^{2+}$+2e$^-$, $MnCl_2$, glycylglycyne, NaOH		–1.248	51Mon4
	Mn+glycylglycylglycyne→Mn(glycylglycylglycyne)$^{2+}$+2e$^-$, $MnCl_2$,glycylglycylglycyne, NaOH		–1.227	55Eva
	[Mn trans-1,2-cyclohexanetetraacetate]$^{2-}$→ [Mn trans-1,2-cyclohexanetetraacetate]$^-$+e$^-$, CH_3COONa, $Mn(ClO_4)_2$, Na_2H_2trans-1,2-cyclohexanetetraacetate, K[Mn trans-1,2-cyclohexanetetraacetate]		0.814	67Ham3
	[Mn ethylenediaminetetraacetate]$^{2-}$→ [Mn ethylenediaminetetraacetate]$^-$+e$^-$, $Mn(ClO_4)_2$, Na_2H_2ethylenediaminetetraacetate, K[Mn ethylenediaminetetraacetate]		0.825	65Tan1
	[Mn ethylenediaminetetraacetate]$^{2-}$→ [Mn ethylenediaminetetraacetate]$^-$+e$^-$, acetate buffer, Na_2H_2ethylenediaminetetraacetate, K[Mn ethylenediaminetetraacetate]		0.825	67Ham3
	[Mn hydroxyethylenediaminetetraacetate]$^{2-}$→ [Mn hydroxyethylenediaminetetraacetate]$^-$ + e$^-$, $Mn(ClO_4)_2$, CH_3COONa, $NaClO_4Na_2H_2$hydroxyethylenediaminetetraacetate, K[Mn hydroxyethylenediaminetetraacetate]		0.782	67Ham3
	Mn+(8-hydroxyquinoline-5-sulfonate)$^{2-}$→ Mn(8–hydroxyquinoline-5-sulfonate)+2e$^-$, $MnCl_2$, NaOH, 8-hydroxyquinoline-5-sulfonate^{2-}		–1.390	54Näs
Mercury (Hg)	2Hg→ Hg_2^{2+} +2e$^-$		0.798	61Ive, 34Bra, 23Lew
	2Hg→ Hg_2^{2+} +2e$^-$, $Hg_2(NO_3)_2$		0.7973	53Bon
	2Hg→ Hg_2^{2+} +2e$^-$, $Hg_2(ClO_4)_2$		0.7973	50ElW1
	2Hg→ Hg_2^{2+} +2e$^-$, $Hg_2(ClO_4)_2$, $HClO_4$, $NaClO_4$		0.7889	71Pan
	2Hg→ Hg_2^{2+} +2e$^-$, Hg_2Cl_2		0.7973	63Mur
	Hg_2^{2+} →2Hg^{2+}+2e$^-$		0.905	31Pop
	Hg_2^{2+} →2Hg^{2+}+2e$^-$, acidic		0.92	72Moo
	Hg→Hg^{2+}+2e$^-$, acidic		0.854	72Moo
	2Hg+2Br$^-$→Hg_2Br_2+2e$^-$, HBr		0.13923	63Gup1
	2Hg+2Br$^-$→Hg_2Br_2+2e$^-$		0.13956	57Kor
			0.1397	40Lar
	2Hg+2Br$^-$→Hg_2Br_2+2e$^-$, KBr, KCl, KNO_3		0.13921	70Leu
	2Hg+2Br$^-$→Hg_2Br_2+2e$^-$, KBr E_0 [V]=0.13925–0.186 10^{-3}(T–25 °C)–0.32 10^{-5}(T–25 °C)^2		0.13921	40Dak
	2Hg+2Br$^-$→Hg_2Br_2+2e$^-$, HBr E_0 [V]=0.13925–0.186·10^{-3}(T–25°C)–0.32·10^{-5}(T–25°C)^2		0.13921	40Dak
	2Hg+2Br$^-$→Hg_2Br_2+2e$^-$, 1 **m**	278	0.14095	57Gup

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	$2Hg+2Br^- \rightarrow Hg_2Br_2+2e^-$, 1 **m**	283	0.14078	57Gup
	$2Hg+2Br^- \rightarrow Hg_2Br_2+2e^-$, 1 **m**	288	0.14041	57Gup
	$2Hg+2Br^- \rightarrow Hg_2Br_2+2e^-$, 1 **m**	293	0.13985	57Gup
	$2Hg+2Br^- \rightarrow Hg_2Br_2+2e^-$, 1 **m**	298	0.13917	57Gup
	$2Hg+2Br^- \rightarrow Hg_2Br_2+2e^-$, 1 **m**	303	0.13836	57Gup
	$2Hg+2Br^- \rightarrow Hg_2Br_2+2e^-$, 1 **m**	308	0.13726	57Gup
	$2Hg+2Br^- \rightarrow Hg_2Br_2+2e^-$, 1 **m**	313	0.13627	57Gup
	$2Hg+2Br^- \rightarrow Hg_2Br_2+2e^-$, 1 **m**	318	0.13503	57Gup
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, acidic		0.2679	64Bet
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, HCl		0.26808	53Pou, 67Cov, 63Gup2, 58Grz, 65Sch1, 68Sha
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$		0.26796	57Kor
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	298	0.26796	51Hil
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	278	0.27286	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	283	0.27189	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	288	0.27075	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	293	0.26943	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	298	0.26797	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	303	0.26639	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	308	0.26466	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	313	0.26278	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	318	0.26079	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	278	0.27283	53Pou
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	283	0.27187	53Pou
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	288	0.27078	53Pou
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	293	0.26949	53Pou
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	298	0.26804	53Pou
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	303	0.26642	53Pou
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	308	0.26466	53Pou
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	313	0.26273	53Pou
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	318	0.26063	53Pou
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	323	0.25839	53Pou
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	328	0.25598	53Pou
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	333	0.25347	53Pou
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	338	0.25088	53Pou
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	343	0.24767	53Pou
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	273	0.27406	58Grz
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	278	0.27321	58Grz
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	283	0.27218	58Grz
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	288	0.27099	58Grz
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	293	0.26962	58Grz

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Hg	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	298	0.26812	58Grz
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	303	0.26648	58Grz
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	308	0.26468	58Grz
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	313	0.26276	58Grz
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	318	0.26068	58Grz
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	323	0.25841	58Grz
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	328	0.25613	58Grz
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	333	0.25376	58Grz
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	278	0.27289	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	283	0.27191	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	288	0.27082	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	293	0.26955	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	298	0.26816	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	303	0.26659	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	308	0.26485	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	313	0.26303	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	318	0.26103	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, 1 **m**	298	0.26796	58Sch
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	298	0.2412	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	273	0.2602	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	283	0.2541	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	288	0.2509	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	293	0.2477	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	298	0.2444	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	303	0.2411	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	308	0.2377	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	311	0.2357	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	313	0.2343	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	323	0.2272	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	333	0.2199	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	343	0.2124	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	353	0.2047	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	363	0.1967	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. saturated in KCl	373	0.1885	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 3.5 **M** in KCl	313	0.2466	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 3.5 **M** in KCl	323	0.2428	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 3.5 **M** in KCl	333	0.2377	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 3.5 **M** in KCl	343	0.2331	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 3.5 **M** in KCl	353	0.2277	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 3.5 **M** in KCl	363	0.2237	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 1 **M** in KCl	273	0.2854	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 1 **M** in KCl	283	0.2839	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 1 **M** in KCl	293	0.2815	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 1 **M** in KCl	298	0.2801	61Ive [6])

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 1 **M** in KCl	303	0.2786	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 1 **M** in KCl	313	0.2753	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 1 **M** in KCl	323	0.2716	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 1 **M** in KCl	333	0.2673	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 1 **M** in KCl	343	0.2622	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 0.1 **M** in KCl	273	0.3338	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 0.1 **M** in KCl	283	0.3343	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 0.1 **M** in KCl	293	0.3340	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 0.1 **M** in KCl	298	0.3337	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 0.1 **M** in KCl	303	0.3332	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 0.1 **M** in KCl	313	0.3316	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 0.1 **M** in KCl	323	0.3296	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 0.1 **M** in KCl	333	0.3229	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 0.1 **M** in KCl	343	0.3236	61Ive [6])
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 1 **M** in KCl	298	0.2801	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 0.1 **M** in KCl	298	0.3337	61Ive
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. 1 **m** in KCl	298	0.2810	61Ive
	$Hg+F^- \rightarrow HgF^++2e^-$		0.876	55Pau
	$Hg+H_2O \rightarrow HgO+2H^++2e^-$, $Ba(OH)_2$ sat.	298	0.1462	35Sam
	$Hg+H_2O \rightarrow HgO+2H^++2e^-$, $Ca(OH)_2$ sat.	298	0.1932	35Sam
	$6Hg+2[Co(CN)_6^{3-}] \rightarrow (Hg_2)_3[Co(CN)_6]+6e^-$, $K_3[Co(CN)_6]$		–0.427	65Roc
	$2Hg+HPO_4^{2-} \rightarrow Hg_2HPO_4+2e^-$	298	0.638	61Ive
	$2Hg+HPO_4^{2-} \rightarrow Hg_2HPO_4+2e^-$, H_3PO_4		0.6359	50Lar
	$2Hg+H_2S \rightarrow Hg_2S+2H^++2e^-$		–0.661	61Ive
	$2Hg+2I^- \rightarrow Hg_2I_2+2e^-$	298	–0.0405	61Ive, 57Kor
	$2Hg+IO_3^- \rightarrow Hg_2(IO_3)_2+2e^-$, KIO_3		0.3942	43Tak
	$2Hg+IO_3^- \rightarrow Hg_2(IO_3)_2+2e^-$	298	0.3944	61Ive
	$Hg+N_3^- \rightarrow Hg_2(N_3)_2+2e^-$		–0.26	52Suz1
	$Hg+2OH^- \rightarrow HgO+H_2O+2e^-$, NaOH		0.0977	57Ham
	$Hg+4SCN^- \rightarrow Hg(SCN)_4^{2-}+2e^-$		0.267	56Tac
	$Hg+4SeCN^- \rightarrow Hg(SeCN)_4^{2-}+2e^-$		0.025	56Tac
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, acidic		0.6151	64Bet
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$		0.6141	57Kor
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4		0.6135	70Sha2
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4, c=1 **M**	273	0.63495	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4, c=1 **M**	278	0.63097	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4, c=1 **M**	283	0.62704	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4, c=1 **M**	288	0.62307	90Gal

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Hg	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4, c=1 **M**	293	0.61930	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4, c=1 **M**	298	0.61515	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4, c=1 **M**	303	0.61107	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4, c=1 **M**	308	0.60701	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4, c=1 **M**	313	0.60305	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4, c=1 **M**	318	0.59900	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4, c=1 **M**	323	0.59487	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4, c=1 **M**	328	0.59051	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4, c=1 **M**	333	0.58659	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$	273	0.63495	35Har
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$	278	0.63097	35Har
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$	283	0.62704	35Har
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$	288	0.62307	35Har
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$	293	0.61930	35Har
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$	298	0.61515	35Har
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$	303	0.61107	35Har
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$	308	0.60701	35Har
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$	313	0.60305	35Har
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$	318	0.59900	35Har
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$	323	0.59487	35Har
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$	328	0.59051	35Har
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$	333	0.58659	35Har
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, H_2SO_4		0.6125	65Cov
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	273	0.6718	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	278	0.6676	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	283	0.6635	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	288	0.6594	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	293	0.6553	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	298	0.6513	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	303	0.6473	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	308	0.6433	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	313	0.6392	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	318	0.6351	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	323	0.6309	90Gal

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	328	0.6266	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	333	0.6226	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	338	0.6177	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat	343	0.6133	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	348	0.6084	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	353	0.6034	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	358	0.5984	90Gal
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, K_2SO_4, sat.	363	0.5931	90Gal
	$2Hg+2S_2O_3^{2-} \rightarrow Hg(S_2O_3)_2^{2-}+2e^-$		0.028	63Mur
	$Hg+3S_2O_3^{2-} \rightarrow Hg(S_2O_3)_3^{4-}+2e^-$		–0.057	56Tac
	$Hg+4S_2O_3^{2-} \rightarrow Hg(S_2O_3)_4^{6-}+2e^-$		–0.100	56Tac
	$2Hg+2(formate)^- \rightarrow Hg_2$ $formate_2+2e^-$, Na formate		0.567	66Cha
	$2Hg+2(formate)^- \rightarrow Hg_2$ $formate_2+2e^-$, formic acid		0.566	74Ost
	$2Hg+2Ac^- \rightarrow Hg_2Ac_2+2e^-$, T = 298 K, 1 **M** acetic acid [7])		0.5109	61Ive
	$2Hg+2Ac^- \rightarrow Hg_2Ac_2+2e^-$, 1 **m** acetic acid	298	0.5117	61Ive
	$2Hg+2Ac^- \rightarrow Hg_2Ac_2+2e^-$, acetic acid		0.51163	63Lar
	$2Hg+2Ac^- \rightarrow Hg_2Ac_2+2e^-$, acetic acid, K acetate, Na acetate		0.51163	64Cov
	$2Hg+2Ac^- \rightarrow Hg_2Ac_2+2e^-$, acetic acid, KNO_3		0.51163	68Shu
	$2Hg+2Ac^- \rightarrow Hg_2Ac_2+2e^-$, acetic acid, Na acetate		0.5111	73Cho
	$2Hg+2(propionate)^- \rightarrow Hg_2$ $propionate_2+2e^-$, Na propionate		0.499	66Cha
	$2Hg+2(oxalate)^- \rightarrow Hg_2$ $oxalate_2+2e^-$, oxalic acid sat.		0.4158	48Fer
	$2Hg+Ox^- \rightarrow Hg_2Ox+2e^-$ [8])	298	0.4166	61Ive
	$2Hg+2(benzoate)^- \rightarrow Hg_2(benzoate)_2+2e^-$, benzoic acid		0.4263	67Ber
	$2Hg+2(picrate)^- \rightarrow Hg_2$ $picrate_2+2e^-$, picric acid		0.4924	71Cov
Molybdenum	$Mo \rightarrow Mo^{3+}+3e^-$		–0.2	52Lat
(Mo)	$MoO^{3+}+H_2O \rightarrow MoO_2^{2+}+2H^++2e^-$, molybdenum blue, HCl		0.482	53ElS
Neodymium	$Nd^{2+} \rightarrow Nd^{3+}+e^-$		1.8	72Ten
(Nd)	$Nd+SO_4^{2-} \rightarrow NdSO_4^+ +3e^-$, $Nd_2(SO_4)_3$		–2.504	54Spe
	$Nd+oxalate^{2-} \rightarrow Nd$ $oxalate^+ +3e^-$		–2.575	55Bab
	$Nd+2(oxalate)^{2-} \rightarrow Nd$ $oxalate_2^- +3e^-$		–2.661	55Bab
Neptunium (Np)	$NpO_2^+ \rightarrow NpO_2^{2+}+e^-$, $NpO_2(NO_3)_2$		1.130	54Spe
Nickel (Ni)	$Ni \rightarrow Ni^{2+}+2e^-$		–0.236	29Har1, 34Col
	$Ni \rightarrow Ni^{2+}+2e^-$, $NiSO_4$		–0.257	55Pal, 62Vag
	$Ni+2OH^- \rightarrow Ni(OH)_2+2e^-$, basic		–0.72	72Moo
	$Ni+P_3O_9^{3-} \rightarrow NiP_3O_9^- +2e^-$, $NiCl_2$, $NaPO_3$		–0.345	53Rin
	$Ni+S_2O_3^{2-} \rightarrow NiS_2O_3+2e^-$, $NiCl_2$, BaS_2O_3		–0.372	51Den

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Ni	$Ni(CN)_3^{2-} + CN^- \rightarrow Ni(CN)_4^{2-} + e^-$, $K_2Ni(CN)_3$ $K_2Ni(CN)_4$		–0.401	47Cag
	Ni+acetylacetone→[Ni acetylacetone]$^{2+}$+2e$^-$, acetylacetone		–0.425	49Mal2
	Ni+2acetylacetone→[Ni acetylacetone$_2$]$^{2+}$+2e$^-$, acetylacetone		–0.559	49Mal2
	Ni+malonate^{2-}→Ni malonate+2e$^-$, $NiCl_2$, NaH malonate		–0.368	49Sto
	Ni+2(oxalate)$^{2-}$→Ni oxalate$_2^{2-}$ +2e$^-$		–0.442	51Bar
	Ni+glycine→[Ni glycine]$^{2+}$+2e$^-$, $NiCl_2$, glycine, NaOH		–0.433	51Mon4
	Ni+2glycine→[Ni glycine$_2$]$^{2+}$+2e$^-$, $NiCl_2$, glycine, NaOH		–0.578	51Mon4
	Ni+glycylglycine→[Ni glyclglycine]$^{2+}$+2e$^-$, $NiCl_2$,glycylglycine, NaOH		–0.383	51Mon4
	Ni+2glycylglycine→[Ni glyclglycine$_2$]$^{2+}$+2e$^-$, $NiCl_2$,glycylglycine, NaOH		–0.483	51Mon4
	Ni+glycylglycylglycine→[Ni glyclglycylglycine]$^{2+}$+2e$^-$, $NiCl_2$,glycylglycylglycine, NaOH		–0.368	55Eva
	Ni+2glycylglycylglycine→[Ni glyclglycylglycine$_2$]$^{2+}$+2e$^-$, $NiCl_2$,glycylglycylglycine, NaOH		–0.459	55Eva
	Ni+8-hydroxyquinoline-5-sulfonate^{2-}→ Ni 8-hydroxyquinoline-5-sulfonate+2e$^-$, 8-hydroxyquinoline-5-sulfonate^{2-}, NaOH, $Ni(NO_3)_2$		–0.538	54Näs
	Ni+2(8-hydroxyquinoline-5-sulfonate^{2-})→ [Ni 8-hydroxyquinoline-5-sulfonate$_2$]$^{2-}$+2e$^-$, 8-hydroxyquinoline-5-sulfonate^{2-}, NaOH, $Ni(NO_3)_2$		–0.796	54Näs
Osmium	$OsBr_6^{3-} \rightarrow OsBr_6^{2-} + e^-$, HBr		0.452	47Dwy1
(Os)	$OsCl_6^{3-} \rightarrow OsCl_6^{2-} + e^-$, HCl		0.452	47Dwy1
	Os dipyridyl$_3^{2+}$ →Os dipyridyl$_3^{3+}$ +e$^-$, Os dipyridyl$_3(ClO_4)_2$, Os dipyridyl$_3(ClO_4)_3$, KCl, HCl, KNO_3		0.8821	52Bar
	Os dipyridyl$_3^{2+}$ →Os dipyridyl$_3^{3+}$ +e$^-$, Os dipyridyl$_3(ClO_4)_2$, Os dipyridyl$_3(ClO_4)_3$, KCl		0.8821	66Buc2
	Os dipyridyl$_2(CN)_2$→Os dipyridyl$_2(CN)_2^+$ +e$^-$, Os dipyridyl$_2(CN)_2$, H_2SO_4		0.792	63Sch2
	Os dipyridyl$_2$(2,4-pentanedione)$^+$→ Os dipyridyl$_2$(2,4-pentanedione)$^{2+}$+e$^-$, Os dipyridyl$_2$(2,4-pentanedione)$(ClO_4)_2$, Os dipyridyl$_2$(2,4-pentanedione)$(ClO_4)_3$, KCl		0.1530	66Buc2
	Os(dipyridyl)(pyridyl)$_4^{2+}$ →Os(dipyridyl)(pyridyl)$_4^{3+}$ +e$^-$, Os(dipyridyl)(pyridyl)$_4(ClO_4)_2$, Os(dipyridyl)(pyridyl)$_4(ClO_4)_3$		0.8052	52Bar
	Os(dipyridyl)$_2$(pyridyl)$_2^{2+}$ →Os(dipyridyl)$_2$(pyridyl)$_2^{3+}$ +e$^-$, Os(dipyridyl)$_2$(pyridyl)$_2(ClO_4)_2$, Os(dipyridyl)$_2$(pyridyl)$_2(ClO_4)_3$		0.8345	52Bar
	Os(dipyridyl)Br(pyridyl)$_3^+$ →Os(dipyridyl)Br(pyridyl)$_3^{2+}$ +e$^-$, Os(dipyridyl)Br(pyridyl)$_3(ClO_4)_2$, Os(dipyridyl)Br(pyridyl)$_3$Cl, KCl		0.4437	52Bar

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	$Os(dipyridyl)_2Br(pyridyl)^+ \rightarrow Os(dipyridyl)_2Br(pyridyl)^{2+} + e^-$, $Os(dipyridyl)_2Br(pyridyl)(ClO_4)_2$, $Os(dipyridyl)_2Br(pyridyl)Cl$, KCl		0.4871	52Bar
	$Os(dipyridyl)(pyridyl)_3Cl^+ \rightarrow Os(dipyridyl)(pyridyl)_3Cl^{2+} + e^-$, $Os(dipyridyl)Cl(pyridyl)_3(ClO_4)_2$, $Os(dipyridyl)Cl(pyridyl)_3Cl$, KCl		0.4255	52Bar
	$Os(dipyridyl)_2(pyridyl)Cl^+ \rightarrow Os(dipyridyl)_2(pyridyl)Cl^{+} + e^-$, $Os(dipyridyl)_2Cl(pyridyl)(ClO_4)_2$, $Os(dipyridyl)_2Cl(pyridyl)Cl$, KCl		0.4836	52Bar
	$Os(dipyridyl)_2Cl(pyridyl)^+ \rightarrow Os(dipyridyl)_2Cl(pyridyl)^{2+} + e^-$, $Os(dipyridyl)_2Cl(pyridyl)(ClO_4)_2$, $Os(dipyridyl)_2Cl(pyridyl)Cl$, KCl		0.4836	52Bar
	$Os(dipyridyl)I(pyridyl)_3^+ \rightarrow Os(dipyridyl)I(pyridyl)_3^{2+} + e$, $Os(dipyridyl)I(pyridyl)_3(ClO_4)_2$, $Os(dipyridyl)I(pyridyl)_3Cl$, KCl^-		0.4508	52Bar
	$Os(dipyridyl)_2(pyridyl)^+ \rightarrow Os(dipyridyl)_2(pyridyl)^{2+} + e^-$, $Os(dipyridyl)_2I(pyridyl)(ClO_4)_2$, $Os(dipyridyl)_2I(pyridyl)Cl$, KCl		0.4888	52Bar
	$Os\ tripyridyl_2^{2+} \rightarrow Os\ tripyridyl_2^{3+} + e^-$, $Os(tripyridyl)_2(ClO_4)_2$, $Os(tripyridyl)_2(ClO_4)_3$, H_2SO_4		0.9866	54Dwy
	$Os(dipyridyl)(tripyridyl)Br^+ \rightarrow Os(dipyridyl)(tripyridyl)Br^{2+} + e^-$, $Os(dipyridyl)Br(tripyridyl)_3(ClO_4)_2$, $Os(dipyridyl)Br(tripyridyl)_3Cl$, KCl		0.5675	52Bar
	$Os(dipyridyl)(tripyridyl)Cl^+ \rightarrow Os(dipyridyl)(tripyridyl)Cl^{2+} + e^-$, $Os(dipyridyl)Cl(tripyridyl)_3(ClO_4)_2$, $Os(dipyridyl)Cl(tripyridyl)_3Cl$, KCl		0.5628	52Bar
	$Os(dipyridyl)(tripyridyl)I^+ \rightarrow Os(dipyridyl)(tripyridyl)I^{2+} + e^-$, $Os(dipyridyl)I(tripyridyl)_3(ClO_4)_2$, $Os(dipyridyl)I(tripyridyl)_3Cl$, KCl		0.5666	52Bar
	$Os(tripyridyl)(dipyridyl)(pyridyl)^{2+} \rightarrow Os(tripyridyl)(dipyridyl)(pyridyl)^{3+} + e^-$, $Os(tripyridyl)(dipyridyl)(pyridyl)(ClO_4)_2$, $Os(tripyridyl)(dipyridyl)(pyridyl)(ClO_4)_3$, KCl		0.8713	52Bar
	$Os(tripyridyl)(dipyridyl)(4\text{-}ethylpyridyl)^{2+} \rightarrow Os(tripyridyl)(dipyridyl)(4\text{-}ethylpyridyl)^{3+} + e^-$, $Os(tripyridyl)(dipyridyl)(4\text{-}ethylpyridyl)(ClO_4)_2$, $Os(tripyridyl)(dipyridyl)(4\text{-}ethylpyridyl)(ClO_4)_3$, KCl		0.8587	52Bar
	$Os(tripyridyl)(dipyridyl)(3\text{-}methylpyridyl)^{2+} \rightarrow Os(tripyridyl)(dipyridyl)(3\text{-}methylpyridyl)^{3+} + e^-$, $Os(tripyridyl)(dipyridyl)(3\text{-}methylpyridyl)(ClO_4)_2$, $Os(tripyridyl)(dipyridyl)(3\text{-}methylpyridyl)(ClO_4)_3$, KCl		0.8672	52Bar
	$Os(tripyridyl)(dipyridyl)(4\text{-}methylpyridyl)^{2+} \rightarrow Os(tripyridyl)(dipyridyl)(4\text{-}methylpyridyl)^{3+} + e^-$, $Os(tripyridyl)(dipyridyl)(4\text{-}methylpyridyl)(ClO_4)_2$, $Os(tripyridyl)(dipyridyl)(4\text{-}methylpyridyl)(ClO_4)_3$, KCl		0.8506	52Bar

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Os	$Os(tripyridyl)(dipyridyl)(4\text{-}propylpyridyl)^{2+} \rightarrow Os(tripyridyl)(dipyridyl)(4\text{-}propylpyridyl)^{3+}+e^-$, $Os(tripyridyl)(dipyridyl)(4\text{-}propylpyridyl)(ClO_4)_2$, $Os(tripyridyl)(dipyridyl)(4\text{-}propylpyridyl)(ClO_4)_3$, KCl		0.8609	52Bar
	$Os(tripyridyl)(pyridyl)_3^{2+} \rightarrow Os(tripyridyl)(pyridyl)_3^{3+}+e^-$, $Os(tripyridyl)(pyridyl)_3(ClO_4)_3$, $Os(tripyridyl)(pyridyl)_3(ClO_4)_2$, KCl		0.7999	52Bar
Palladium (Pd)	$Pd \rightarrow Pd^+ + 2e^-$		0.951	63Gri
	$Pd \rightarrow Pd^+ + 2e^-$		0.83	32Chl, 38Owe2
	$Pd \rightarrow Pd^+ + 2e^-$, $Pd(ClO_4)_2$, $HClO_4$		0.951	67Iza
	$Pd+Cl^- \rightarrow PdCl^+ + e^-$		0.771	56Dro
	$Pd+2Cl^- \rightarrow PdCl_2 + 2e^-$		0.638	56Dro
	$Pd+3Cl^- \rightarrow PdCl_3^- + 2e^-$		0.567	56Dro
	$Pd+4Cl^- \rightarrow PdCl_4^{2-} + 2e^-$, $PdCl_4$		0.591	66Kra2
	$Pd+5Cl^- \rightarrow PdCl_5^{3-} + 2e^-$		0.553	56Dro
	$Pd+6Cl^- \rightarrow PdCl_6^{4-} + 2e^-$		0.615	56Dro
	$PdCl_4^{2-} + 2Cl^- \rightarrow PdCl_6^{4-} + 2e^-$		1.288	64Bet
	$[Pd\ ethylenediamine_2]^{2+}+2Br^- \rightarrow [Pd\ ethylenediamine_2\ Br_2]^{2+}+2e^-$, $[Pd\ ethylenediamine_2]^{2+}$, $[Pd\ ethylenediamine_2\ Br_2]^{2+}$		0.692	65Bab
	$[Pd\ ethylenediamine_2]^{2+}+2Cl^- \rightarrow [Pd\ ethylenediamine_2\ Cl_2]^{2+}+2e^-$, $[Pd\ ethylenediamine_2]^{2+}$, $[Pd\ ethylenediamine_2\ Cl_2]^{2+}$		1.13	65Bab
	$[Pd\ ethylenediamine_2]^{2+}+2I^- \rightarrow [Pd\ ethylenediamine_2\ I_2]^{2+}+2e^-$, $[Pd\ ethylenediamine_2]^{2+}$, $[Pd\ ethylenediamine_2\ I_2]^{2+}$		0.625	65Bab
Platinum (Pt)	$Pt \rightarrow Pt^{2+} + 2e^-$		1.28	52Lat
	$Pt+4Cl^- \rightarrow PtCl_4^{2-} + 2e^-$		0.755	68Gri
	$Pt+4I^- \rightarrow PtI_4^{2-} + 2e^-$		0.40	66Hub
	$PtCl_4^{2-} + 2Cl^- \rightarrow PtCl_6^{2-} + 2e^-$		0.72	30Smi
	$PtI_4^{2-} + 2I^- \rightarrow PtI_6^{2-} + 2e^-$		0.39	66Hub
	$Pt+2OH^- \rightarrow Pt(OH)_2 + 2e^-$		0.14	61Saw
	$PtCl(NCl_2)(NH_3)_4+2Cl^- \rightarrow [PtCl(NCl_2)(NH_3)_4]Cl_2+2e^-$, $PtCl(NCl_2)(NH_3)_4Cl_2$, H_2SO_4		1.28	69Kuk
	$[PtCl(NCl_2)(NH_3)_2]Cl^-+2Cl^- \rightarrow [PtCl_3(NCl_2)(NH_3)_2]Cl^-+2e^-$, $PtCl_3(NCl_2)(NH_3)2Cl^-$, H_2SO_4		0.95	69Kuk
	$PtCl(NO_2)(NH_3)_2+NO^{2-}+Cl^- \rightarrow Pt(Cl(NO_2))_2(NH_3)_2+2e^-$		0.722	66Che
	$Pt(NH_3)_4^{2+} + 2Br^- \rightarrow Pt(NH_3)_4Br_2^{2+} + 2e^-$		0.583	49Gri
	$Pt(NH_3)_4^{2+} + 2I^- \rightarrow Pt(NH_3)_4I_2^{2+} + 2e^-$		0.431	49Gri
	$cis\text{-}[Pt(NH_3)_2Br_2]+2Br^- \rightarrow cis\text{-}[Pt(NH_3)_2Br_4]+2e^-$, $cis\text{-}[Pt(NH_3)_2Br_2]$, $cis\text{-}[Pt(NH_3)_2Br_4]$		0.597	49Gri
	$trans\text{-}[Pt(NH_3)_2Br_2]+2Br^- \rightarrow trans\text{-}[Pt(NH_3)_2Br_4]+2e^-$, $trans\text{-}[Pt(NH_3)_2Br_2]$, $trans\text{-}[Pt(NH_3)_2Br_4]$		0.600	49Gri

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	cis-$[Pt(NH_3)_2I_2]+2I^- \rightarrow$ cis-$[Pt(NH_3)_2I_4]+2e^-$, cis-$[Pt(NH_3)_2I_2]$, cis-$[Pt(NH_3)_2I_4]$		0.390	49Gri
	trans-$[Pt(NH_3)_2I_2]+2I^- \rightarrow$ trans-$[Pt(NH_3)_2I_4]+2e^-$, trans-$[Pt(NH_3)_2I_2]$, trans-$[Pt(NH_3)_2I_4]$		0.383	49Gri
	$Pt((NO_2)(NH_3))_2+2Br^- \rightarrow Pt(Br_2((NO_2)(NH_3))_2+2e^-$, $Pt((NO_2)(NH_3))_2$, $Pt(Br_2((NO_2)(NH_3))_2$		0.782	66Che
	$Pt((NO_2)(NH_3))_2+2Br^- \rightarrow Pt(Br_2((NO_2)(NH_3))_2+2e^-$, $Pt((NO_2)(NH_3))_2$, $Pt(Br_2((NO_2)(NH_3))_2$		0.786	66Che
	$Pt(NO_2)(NH_3)_2Br+ NO_2^- +Br^- \rightarrow PtBr_2((NO_2)(NH_3))_2+2e^-$, $Pt((NO_2)(NH_3))_2Br$, $PtBr_2(NO_2)_2(NH_3)_2$		0.807	66Che
	$Pt((NO_2)(NH_3))_2+2Cl^- \rightarrow PtCl_2((NO_2)(NH_3))_2+2e^-$, $Pt((NO_2)(NH_3))_2$, $PtCl_2((NO_2)(NH_3))_2$		0.759	66Che
	$Pt(NO_2)_2(NH_3)_2+2Cl^- \rightarrow PtCl_2(NO_2)_2(NH_3)_2+2e^-$, $Pt(NO_2)_2(NH_3)_2$, $Pt(Cl_2(NO_2)_2(NH_3)_2$		0.763	66Che
	$PtCl(NO_2)(NH_3)_2+ NO_2^- +Cl^- \rightarrow PtCl_2(NO_2)_2(NH_3)_2+2e^-$, $PtCl(NO_2)_2(NH_3)_2$, $PtCl_2(NO_2)_2(NH_3)_2$		0.778	66Che
	Pt(methylamine)$_4^{2+}$ +2Cl$^- \rightarrow$ Pt(methylamine)$_4$ Cl$_2^{2+}$ +2e$^-$, Pt(methylamine)$_4^{2+}$, Pt(methylamine)$_4$ Cl$_2^{2+}$, HCl		0.6220	68Gri
	Pt(propylamine)$_4^{2+}$ +2Cl$^- \rightarrow$ Pt(propylamine)$_4$ Cl$_2^{2+}$ +2e$^-$, Pt(propylamine)$_4^{2+}$, Pt(propylamine)$_4$ Cl$_2^{2+}$, H_2SO_4		0.8327	68Gri
	Pt(butylamine)$_4^{2+}$ +2Cl$^- \rightarrow$ Pt(butylamine)$_4$Cl$_2^{2+}$ +2e$^-$, Pt(butylamine)$_4^{2+}$, Pt(butylamine)$_4$Cl$_2^{2+}$, H_2SO_4		0.8380	68Gri
	Pt(ethylamine)$_4^{2+}$ +2Cl$^- \rightarrow$ Pt(ethylamine)$_4$Cl$_2^{2+}$ +2e$^-$, Pt(ethylamine)$_4^{2+}$, Pt(ethylamine)$_4$Cl$_2^{2+}$, HCl		0.6919	68Gri
	Pt(ethylenediamine)$_2^{2+}$ +2Cl$^- \rightarrow$ Pt(ethylenediamine)$_2$Cl$_2^{2+}$ +2e$^-$, Pt(ethylenediamine)$_2^{2+}$, Pt(ethylenedamine)$_2$Cl$_2^{2+}$, HCl		0.5660	68Gri
Plutonium	$Pu^{3+} \rightarrow Pu^{4+}+e^-$, $HClO_4$		1.006	62Sch2
(Pu)	$Pu^{3+}+5acetate^- \rightarrow$ Pu acetate$_5^-$ $+e^-$		–0.33	62Sch2
Polonium (Po)	$Po \rightarrow Po^{3+}+3e^-$		0.56	35Hai, 29Sch
	$Po \rightarrow Po^{4+}+4e^-$		0.76	56Bag
	$Po+4Cl^- \rightarrow PoCl_4^{2-} +2e^-$		0.38	56Bag
	$PoCl_4^{2-} +2Cl^- \rightarrow PoCl_6^{2-} +2e^-$, HCl		0.72	56Bag
Potassium (K)	$K \rightarrow K^+ + e^-$		–2.922	34Arm, 31Lar, 52Lat, 12Lew
	$K \rightarrow K^+ + e^-$, acidic		–2.925	72Moo
	$K \rightarrow K^+ + e^-$	298.15	–2.928	77Gio
K$_{(Hg)}$	$K_{(Hg)} \rightarrow K^+ + Hg + e^-$		–2.931	68Dil

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
$K_{(Hg)}$	$K_{(Hg)} \rightarrow K^+ + Hg + e^-$	283.15	–1.950	77Gio
$K_{(Hg)}$	$K_{(Hg)} \rightarrow K^+ + Hg + e^-$	298.15	–1.975	77Gio
$K_{(Hg)}$	$K_{(Hg)} \rightarrow K^+ + Hg + e^-$	313.15	–1.998	77Gio
$K_{(Hg)}$	$K_{(Hg)} \rightarrow K^+ + Hg + e^-$	328.15	–2.020	77Gio
$K_{(Hg)}$	$K_{(Hg)} \rightarrow K^+ + Hg + e^-$	343.15	–2.038	77Gio
	$K + BrO_3^- \rightarrow KBrO_3 + e^-$		–2.895	44Jon1
	$K + ClO_3^- \rightarrow KClO_3 + e^-$		–2.928	31Ban
	$K + ClO_4^- \rightarrow KClO_4 + e^-$		–2.902	45Jon
	$K + Co(CN)_6^{3-} \rightarrow KCo(CN)_6^{2-} + e^-$		3.010	50Jam
	$K + Fe(CN)_6^{3-} \rightarrow KFe(CN)_6^{2-} + e^-$		3.008	50Jam, 37Har2
	$K + Fe(CN)_6^{4-} \rightarrow KFe(CN)_6^{3-} + e^-$		3.0667	36Jon, 49Jam
	$K + IO_3^- \rightarrow KIO_3 + e^-$		–2.917	42Kri, 48Mon
	$K + NO_3^- \rightarrow KNO_3 + e^-$		–2.917	31Ban, 27Dav
	$K + ReO_4^- \rightarrow KReO_4 + e^-$		–2.974	46Jon
	$K + SO_4^{2-} \rightarrow KSO_4^- + e^-$		–2.988	50Jen
	$K + S_2O_3^{2-} \rightarrow KS_2O_3^- + e^-$		–2.987	51Den
Praseodymium	$Pr^{2+} \rightarrow Pr^{3+} + e^-$		1.8	72Ten
(Pr)	$Pr + SO_4^{2-} \rightarrow PrSO_4^+ + 3e^-$, $Pr_2(SO_4)_3$		–2.534	54Spe, 55Gim
Rhenium (Re)	$ReO_3 + H_2O \rightarrow ReO_4^- + 2H^+ + e^-$, $HReO_4$, $NaReO_4$, NaOH		0.768	57Kin
Rhodium	$Rh \rightarrow Rh^{2+} + 2e^-$		0.6	52Lat
(Rh)	$Rh \rightarrow Rh^{3+} + 3e^-$, $Rh(ClO_4)_3$, $HClO_4$		0.758	71Amo
Rubidium (Rb)	$Rb \rightarrow Rb^+ + e^-$		–2.96	52Lat, 31Lar, 46Gap, 15Lew
$Rb_{(Hg)}$	$Rb_{(Hg)} \rightarrow Rb^+ + Hg + e^-$		–2.98	51Boc
$Rb_{(Hg)}$	$Rb_{(Hg)} \rightarrow Rb^+ + Hg + e^-$	283.15	–1.943	74Lon
$Rb_{(Hg)}$	$Rb_{(Hg)} \rightarrow Rb^+ + Hg + e^-$	298.15	–1.969	74Lon
$Rb_{(Hg)}$	$Rb_{(Hg)} \rightarrow Rb + Hg + e^-$	313.15	–1.996	74Lon
$Rb_{(Hg)}$	$Rb_{(Hg)} \rightarrow Rb^+ + Hg + e\, Ru(H_2O)(NH_3)_5^{3+-}$	328.15	–2.022	74Lon
$Rb_{(Hg)}$	$Rb_{(Hg)} \rightarrow Rb^+ + Hg + e^-$	343.15	–2.049	74Lon
Ruthenium	$Ru^{2+} \rightarrow Ru^{3+} + e^-$, HCl, $RuCl_2$, $RuCl_3$		0.2487	66Buc1
(Ru)	$Ru(CN)_6^{4-} \rightarrow Ru(CN)_6^{3-} + e^-$, $K_4Ru(CN)_6$, H_2SO_4, KCl		0.86	59Geo
	$Ru(NH_3)_6^{2+} \rightarrow Ru(NH_3)_6^{3+} + e^-$, $NaClO_4$		0.214	65And

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	$Ru(NH_3)_5^{2+} + H_2O \rightarrow Ru(H_2O)(NH_3)_5^{3+} + e^-$, $NaClO_4$		0.20	65And
	$RuO_4^- \rightarrow RuO_4 + e^-$, RuO_4^-, RuO_4		1.0	57Con2
	$RuO_4^- \rightarrow RuO_4 + e^-$, $KRuO_4$, RuO_4, $NaClO_4$		1.0	54Sil
	$RuO_4^{2-} \rightarrow RuO_4^- + e^-$, RuO_4^-, RuO_4		0.59	57Con2
	$RuO_4^{2-} \rightarrow RuO_4^- + e^-$, $KRuO_4$, RuO_4, $NaClO_4$		0.59	54Sil
	$Ru(dipyridyl)_3^{2+} \rightarrow Ru(dipyridyl)_3^{3+} + e^-$, Ru $dipyridyl_3Cl_2$		1.29	59Geo
	$Ru(dipyridyl)_3^{2+} \rightarrow Ru(dipyridyl)_3^{3+} + e^-$, H_2SO_4		1.29	49Bra
	$Ru(dipyridyl)_3^{2+} \rightarrow Ru(dipyridyl)_3^{3+} + e^-$, H_2SO_4		1.29	63Sch2
	$Ru(dipyridyl)_2(CN)_2 \rightarrow Ru(dipyridyl)_2(CN)_2^+ + e^-$, HNO_3, H_2SO_4		1.10	63Sch2
	$Ru(terpyridyl)_2^{2+} \rightarrow Ru(terpyridyl)_2^{3+} + e^-$, H_2SO_4		1.202	49Bra
Samarium (Sm)	$Sm + SO_4^{2-} \rightarrow SeSO_4^+ + 3e^-$, $Sm_2(SO_4)_3$		–2.487	54Spe
Scandium	$Sc + F^- \rightarrow ScF^{2+} + 3e^-$		–2.239	55Pau
(Sc)	$Sc + 2F^- \rightarrow ScF_2^+ + 3e^-$		–2.333	55Pau
	$Sc + 3F^- \rightarrow ScF_3 + 3e^-$		–2.428	55Pau
Selenium (Se)	$Se + 2e^- \rightarrow Se^{2-}$		–0.78	23Kra2, 23Lew
Silver (Ag)	$Ag \rightarrow Ag^+ + e^-$		0.799	61Ive, 23Lew
	$Ag \rightarrow Ag^+ + e^-$, acidic		0.7991	72Moo
	$Ag \rightarrow Ag^+ + e^-$, $AgClO_4$, $NaClO_4$		0.800	65Pap
	$Ag + Br^- \rightarrow AgBr + e^-$, HBr		0.07133	62Het, 60Tow
	$Ag + Br^- \rightarrow AgBr + e^-$, HBr		0.0714	70Sin
	$Ag + Br^- \rightarrow AgBr + e^-$		0.07131	57Kor
	$Ag + Br^- \rightarrow AgBr + e^-$		0.654	53Ber1
	$Ag + Br^- \rightarrow Ag^+ + e^- + Br^-$, acidic		0.0712	64Bet
	$Ag + 2Br^- \rightarrow AgBr_2^- + e^-$, $AgClO_4$, NaBr, $NaClO_4$		0.379	53Ber1
	$Ag + 3Br^- \rightarrow AgBr_3^{2-} + e^-$, $AgClO_4$, NaBr, $NaClO_4$		0.328	53Ber1
	$Ag + 4Br^- \rightarrow AgBr_4^{3-} + e^-$, $AgClO_4$, NaBr, $NaClO_4$		0.274	53Ber1
	$Ag + Br^- + 3Cl^- \rightarrow AgBrCl_3^{3-} + e^-$, AgBr sat., KCl, KBr		0.330	57Cha3
	$Ag + Cl^- \rightarrow Ag^+ + e^- + Cl^-$, acidic		0.2224	64Bet
	$Ag + Cl^- \rightarrow Ag^+ + e^- + Cl^-$, HCl		0.22230	67Sen1, 54Bat2, 60Gre, 62Lie
	$Ag + Cl^- \rightarrow Ag^+ + e^- + Cl^-$		0.22234	57Kor
			0.22233	71Kom1
	$Ag + Cl^- \rightarrow Ag^+ + e^- + Cl^-$, KCl, HCl		0.22238	72Kom

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Ag	$Ag+Cl^-+3Br^- \rightarrow AgClBr_3^{3-}+e^-$, AgBr sat., KCl, KBr		0.239	57Cha3
	$Ag+ClO_4^- \rightarrow AgClO_4+e^-$, Ag ClO_4		0.787	45Jon
	$3Ag+Co(CN)_6^{3-} \rightarrow Ag_3Co(CN)_6+3e^-$, $K_3Co(CN)_6$		0.298	65Roc
	$2Ag+CrO_4^{2-} \rightarrow Ag_2CrO_4+2e^-$, KNO_3, $AgNO_3$, K_2CrO_4		0.4470	55Pan
	$4Ag+Fe(CN)_6^{4-} \rightarrow Ag_4Fe(CN)_6+4e^-$, $K_4Fe(CN)_6$		0.1478	64Roc
	$Ag+I^- \rightarrow AgI+e^-$, HI		–0.15224	64Het
	$Ag+I^- \rightarrow AgI+e^-$ E_{00} [V]=–0.14590–1.696·10^{-4}T–2.934 10^{-6}T, 20<T<145 °C		–0.1510	37Can1, 37Can2 65Kor
	$Ag+IO_3^- \rightarrow AgIO_3+e^-$		0.763	55Ren
	$Ag+2IO_3^- \rightarrow Ag(IO_3)_2^- +e^-$		0.725	55Ren
	$Ag+MnO_4^- \rightarrow AgMnO_4+e^-$, $NaClO_4$, $AgMnO_4$		0.2149	75Das3
	$2Ag+2MoO_4^{2-} \rightarrow Ag_2MoO_4+2e^-$		0.4573	54Pan1
	$Ag+N_3^- \rightarrow AgN_3+e^-$		0.29	52Suz2
	$Ag+N_3^- \rightarrow AgN_3+e^-$		0.2919	57Kor
	$Ag+N_3^- \rightarrow AgN_3+e^-$	278	0.2959	38Tay
	$Ag+N_3^- \rightarrow AgN_3+e^-$	288	0.2942	38Tay
	$Ag+N_3^- \rightarrow AgN_3+e^-$	298	0.2919	38Tay
	$Ag+N_3^- \rightarrow AgN_3+e^-$	308	0.2889	38Tay
	$Ag+N_3^- \rightarrow AgN_3+e^-$	318	0.2854	38Tay
	$Ag+NH_3 \rightarrow AgNH_3^+ +e^-$		0.603	41Dor, 43Vos1, 44Kil
	$2Ag+2OH^- \rightarrow Ag_2O+H_2O+2e^-$, basic		0.342	64Bet
	$2Ag+2OH^- \rightarrow Ag_2O+H_2O+2e^-$, NaOH		0.342	62Gre
	$2Ag+2OH^- \rightarrow Ag_2O+H_2O+2e^-$, basic, sat.		0.342	57Ham
	$2Ag+2OH^- \rightarrow Ag_2O+H_2O+2e^-$, basic, sat.		0.338	61Dir
	$2Ag+2H_2O \rightarrow Ag_2O+2H^++2e^-$		1.1700	61Ive
	$2AgO+H_2O+2e^- \rightarrow Ag_2O+2OH^-$, basic, sat.		0.604	62Dir
	$3Ag+PO_4^{3-} \rightarrow Ag_3PO_4+3e^-$, H_3PO_4		0.3402	68Naq
	$2Ag+S^{2-} \rightarrow Ag_2S+2e^-$		–0.691	59Gol
	$2Ag+H_2S \rightarrow Ag_2S+2H^++2e^-$		–0.7125	61Ive
	$Ag+SCN^- \rightarrow AgSCN+e^-$, $HClO_4$, KSCN		0.08951	56Van
	$Ag+SCN^- \rightarrow AgSCN+e^-$, $HClO_4$, KSCN		0.08944	75Lal
	$Ag+SCN^- \rightarrow AgSCN+e^-$		0.0947	57Kor
	$Ag+2SCN^- \rightarrow Ag(SCN)_2^- +e^-$, AgSCN, KSCN		0.304	53Cav
	$Ag+3SCN^- \rightarrow Ag(SCN)_3^{2-} +e^-$, AgSCN, KSCN		0.231	53Cav

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	$Ag+4SCN^- \rightarrow Ag(SCN)_4^{3-}+e^-$, AgSCN, KSCN		0.214	53Cav
	$Ag+SeCN^- \rightarrow AgSeCN+e^-$, K_2CO_3, KSeCN		–0.0288	74Das
	$2Ag+SeO_3^{2-} \rightarrow Ag_2SeO_3+2e^-$, $NaHSeO_3$		0.3629	61Lin
	$2Ag+SeO_4^{2-} \rightarrow Ag_2SeO_4+2e^-$		0.8289	42Gel
	$2Ag+WO_4^{2-} \rightarrow Ag_2WO_4+2e^-$, Na_2WO_4,$NaNO_3$,$AgNO_3$		0.4660	54Pan2
	$Ag+formate^- \rightarrow Ag$ formate$+e^-$		0.502	66Cha
	$Ag+2(acetate)^- \rightarrow Ag$ acetate$_2^- +e^-$, acetic acid		0.421	52Mac, 42Mac, 47Mac, 51Dav2
	$2Ag+oxalate^{2-} \rightarrow Ag_2$ oxalate$+e^-$, Ag acetate, Na acetate		0.4647	48Fer
	$2Ag+tartrate^{2-} \rightarrow Ag_2$ tartrate$+2e^-$, tartaric acid		0.5610	46Ple1
	$Ag+picrate^- \rightarrow Ag$ picrate$+e^-$		0.17	68Mat
	$Ag+chloroacetate^- \rightarrow Ag$ chloroacetate$+e^-$, Ag chloroacetate		0.422	52Mac
	$Ag+2(chloroacetate)^- \rightarrow Ag_3$ citrate$+e^-$, citric acid		0.5046	66Mat
	$Ag+benzoate^- \rightarrow Ag$ benzoate$+e^-$, $AgClO_4$, $NaClO_4$		0.746	49Led
	$Ag+methylamine \rightarrow Ag(methylamine)^+ +e^-$, $CH_3NH_3NO_3$		0.614	50Bje
	$Ag+2methylamine \rightarrow Ag(methylamine)_2^+ +e^-$		0.399	55Fyf
	$Ag+2dimethylamine \rightarrow Ag(dimethylamine)_2^+ +e^-$		0.482	55Fyf
	$Ag+ethylamine \rightarrow Ag$ ethylamine$^+ +e^-$		0.600	48Bru
	$Ag+2ethylamine \rightarrow Ag(ethylamine)_2^+ +e^-$		0.368	55Fyf, 48Bru
	$Ag+ethylenediamine \rightarrow Ag$ ethylenediamine$^+ +e^-$		0.361	50Bje
	$Ag+2alanine \rightarrow Ag(alanine)_2^+ +e^-$, $AgNO_3$, alanine		0.375	51Mon2
	$Ag+2glycine \rightarrow Ag(glycine)_2^+ +e^-$, $AgNO_3$, glycine		0.392	51Mon2
	$Ag+2glycylglycine \rightarrow Ag(glycylglycine)_2^+ +e^-$, $AgNO_3$, glycylglycine		0.505	51Mon2
	$Ag+2nicotinamide \rightarrow Ag(nicotinamide)_2^+ +e^-$		0.609	55Mur
	$Ag+2isonicotinamide \rightarrow Ag(isonicotinamide)_2^+ +e^-$		0.699	55Mur
	$Ag+2methylnicotinamide \rightarrow Ag(methylnicotinamide)^+ +e^-$		0.625	55Mur
	$Ag+2methylisonicotinate \rightarrow Ag(methylisonicotinat)_2^+ +e^-$		0.655	55Mur
	$Ag+3thiourea \rightarrow Ag(thiourea)_3^+ +e^-$, thiourea		0.026	43Vos2
	$Ag+pyridine \rightarrow Ag$ pyridine$^+ +e^-$		0.682	55Mur, 61Sir
	$Ag+2pyridine \rightarrow Ag(pyridine)_2^+ +e^-$		0.549	55Mur, 61Sir
	$Ag(dipyridyl)_2^+ \rightarrow Ag(dipyridyl)_2^{2+} +e^-$, KNO_3		–0.758	54Scr
	$Ag+2(3\text{-}cyanopyridine) \rightarrow Ag$ 3-cyanopyridine$_2^+ +e^-$		0.629	55Mur
	$Ag+2(4\text{-}cyanopyridine) \rightarrow Ag$ 4-cyanopyridine$_2^+ +e^-$		0.618	55Mur

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Ag	Ag+(3-methoxypyridine)→Ag(3-methoxypyridine)$^+$+e$^-$		0.707	55Mur
	Ag+2(3-methoxypyridine)→Ag(3-methoxypyridine)$_2^+$+e$^-$		0.583	55Mur
	Ag+(4-methoxypyridine)→Ag(4-methoxypyridine)$^+$+e$^-$		0.665	55Mur
	Ag+2(4-methoxypyridine)→Ag(4-methoxypyridine)$_2^+$+e$^-$		0.538	55Mur
Sodium (Na)	Na→Na$^+$+e$^-$		–2.713	46Gap, 10Lew, 52Lat, 31Lar, 40Tay
	Na→Na$^+$+e$^-$, acidic		–2.714	72Moo
	Na→Na$^+$+e$^-$	298.15	–2.717	77Mus
	Na+ BrO_3^- →NaBrO_3+e$^-$	298	–2.692	55Mar1
	Na+OH$^-$→NaOH+e$^-$, 0.1 **M** NaOH		–2.670	33Har1
	Na+ $P_2O_7^{4-}$ → $NaP_2O_7^{3-}$ +e$^-$, 0.523...2.34·10^{-3} **M** $Na_4P_2O_7$, 0.256...1.17 10^{-3} **M** HCl		–2.849	49Mon1
	Na+ $HP_2O_7^{3-}$ → $NaHP_2O_7^{2-}$ +e$^-$	298	–3.610	49Dav1
	Na+ $P_3O_9^{3-}$ → $NaP_3O_9^{2-}$ +e$^-$	298	–2.779	55Dav
	Na+ $P_3O_{10}^{5-}$ → $NaP_3O_{10}^{4-}$ +e$^-$, 0.1941...2.749·10^{-3} **M** $Na_5P_3O_{10}$		–2.862	49Mon2
	Na+ SO_4^{2-} → $NaSO_4^-$ +e$^-$		–2.752	50Jen
	Na+ $S_2O_3^{2-}$ → $NaS_2O_3^-$ +e$^-$, 1.014·10^{-3} **M** $Na_2S_2O_3$, 0.01...0.02 **M** NaCl		–2.747	51Gim
	Na→Na$^+$+e$^-$	298	–2.71	46Ple2
Na$_{(Hg)}$	Na$_{(Hg)}$→Na$^+$+Hg+e$^-$	283.15	–1.941	71Mus
Na$_{(Hg)}$	Na$_{(Hg)}$→Na$^+$+Hg+e$^-$	298.15	–1.958	71Mus
Na$_{(Hg)}$	Na$_{(Hg)}$→Na$^+$+Hg+e$^-$	313.15	–1.976	71Mus
Na$_{(Hg)}$	Na$_{(Hg)}$→Na$^+$+Hg+e$^-$	328.15	–1.992	71Mus
Na$_{(Hg)}$	Na$_{(Hg)}$→Na$^+$+Hg+e$^-$	343.15	–2.008	71Mus
Strontium (Sr)	Sr→Sr^{2+}+2e$^-$		–2.89	52Lat, 28Dev, 34Lat
Sr$_{(Hg)}$	Sr$_{(Hg)}$→Sr^{2+}+Hg+2e$^-$	283.15	–1.889	75Lon1
Sr$_{(Hg)}$	Sr$_{(Hg)}$→Sr^{2+}+Hg+2e$^-$	298.15	–1.900	75Lon1
Sr$_{(Hg)}$	Sr$_{(Hg)}$→Sr^{2+}+Hg+2e$^-$	313.15	–1.911	75Lon1
Sr$_{(Hg)}$	Sr$_{(Hg)}$→Sr^{2+}+Hg+2e$^-$	328.15	–1.921	75Lon1
Sr$_{(Hg)}$	Sr$_{(Hg)}$→Sr^{2+}+Hg+2e$^-$	343.15	–1.929	75Lon1
	$SrCO_3$+Pb→Sr^{2+}+$PbCO_3$+2e$^-$, $SrCl_2$		–2.886	62Jak2, 53Har
	Sr+ $Fe(CN)_6^{3-}$ → $SrFe(CN)_6^-$ +2e$^-$		–2.972	52Gib
	Sr+ IO_3^- → $SrIO_3^+$ +2e$^-$		–2.931	52Col2
	Sr+OH$^-$→SrOH$^+$+2e$^-$		–2.913	52Col2, 53Har

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	$Sr+P_3O_9^{3-} \rightarrow SrP_3O_9^{-}+2e^-$		–2.987	52Mon
	$Sr+P_4O_{12}^{4-} \rightarrow SrP_4O_{12}^{2-}+2e^-$		–3.039	52Mon
	$Sr+S_2O_3^{2-} \rightarrow SrS_2O_3+2e^-$		–2.948	51Den
	$Sr+formate^- \rightarrow Sr\ formate^+ +2e^-$		–2.907	52Col1
	$Sr+bromoacetate^- \rightarrow Sr\ bromoacetate^+ +2e^-$		–2.896	52Col1
	$Sr+propionate^- \rightarrow Sr\ propionate^+ +2e^-$		–2.895	52Col1
	$Sr+lactate^- \rightarrow Sr\ lactate^+ +2e^-$		–2.917	52Col1, 54Dav2
	$Sr+pivalate^- \rightarrow Sr\ pivalate^+ +2e^-$		–2.892	52Col1
	$Sr+n\text{-}valerate^- \rightarrow Sr\ n\text{-}valerate^+ +2e^-$		–2.894	52Col1
	$Sr+alanine \rightarrow Sr\ alanine^{2+} +2e^-$		–2.909	52Col1
	$Sr+glycine \rightarrow Sr\ glycine^{2+} +2e^-$		–2.915	52Col1
	$Sr+8\text{-}hydroxyquinoline\text{-}5\text{-}sulfonate^{2-} \rightarrow$ Sr(8-hydroxyquinoline-5-sulfonate)$+2e^-$, NaOH, $SrCl_2$, 8-hydroxyquinoline-5-sulfonate^{2-}		–2.969	54Näs
Technetium (Tc)	$TcO_2+2H_2O \rightarrow TcO_4+4H^+ +3e^-$, $HTcO_4$		0.782	53Cob
	$TcO_2+4OH^- \rightarrow TcO_4^- +2H_2O+3e^-$, NaOH, $NaTcO_4$		–0.322	53Cob
Tellurium (Te)	$Te \rightarrow Te^{2+}+2e^-$		–0.92	23Kra1, 23Lew
	$Te \rightarrow Te^{4+}+4e^-$		0.56	33Get, 23Kra1
	$Te+3H_2O \rightarrow TeO_2 \cdot H_2O+4H^{++}+4e^-$, citric acid, phopshate buffer		0.593	59Ric
	$2H_2Te \rightarrow Te_2+4H^+ +4e^-$, HCl, $LaCl_3$		–0.50	62Awa
	$H_2Te_2 \rightarrow Te_2+2H^+ +2e^-$, HCl, $LaCl_3$		–0.365	62Awa
	$Te_2^{2-} \rightarrow 2Te+2e^-$, Na_2Te_2,NaOH		–0.790	71Kom2
	$Te_2^{2-} \rightarrow 2Te+2e^-$, $9 < pH < 13$		–0.845	64Pan
Terbium (Tb)	$Tb^{2+} \rightarrow Tb^{3+}+e^-$		1.6	72Ten
Thallium (Tl)	$Tl \rightarrow Tl^+ +e^-$		–0.3363	64Bet
	$Tl \rightarrow Tl^+ +e^-$		–0.335	61Ive
	$Tl \rightarrow Tl^+ +e^-$		–0.336	23Lew, 25Ger, 42Bro2, 34Cow
	$Tl \rightarrow Tl^+ +e^-$	288.15	–0.313	79Lon2
	$Tl \rightarrow Tl^+ +e^-$	293.15	–0.320	79Lon2
	$Tl \rightarrow Tl^+ +e^-$	298.15	–0.327	79Lon2
	$Tl \rightarrow Tl^+ +e^-$	303.15	–0.334	79Lon2
	$Tl \rightarrow Tl^+ +e^-$	308.15	–0.340	79Lon2
	$Tl^+ \rightarrow Tl^{3+}+2e^-$		1.25	36She
	$Tl^+ \rightarrow Tl^{3+}+2e^-$, HNO_3	273.15	1.193	36Noy
	$Tl^+ \rightarrow Tl^{3+}+2e^-$, HNO_3		1.2303	36Noy
$Tl_{(Hg)}$	$Tl_{(Hg)} \rightarrow Tl^+ +Hg+e^-$	283.15	–0.273	79Lon2

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
$Tl_{(Hg)}$	$Tl_{(Hg)} \rightarrow Tl^+ + Hg + e^-$	298.15	–0.293	79Lon2
$Tl_{(Hg)}$	$Tl_{(Hg)} \rightarrow Tl^+ + Hg + e^-$	313.15	–0.314	79Lon2
$Tl_{(Hg)}$	$Tl_{(Hg)} \rightarrow Tl^+ + Hg + e^-$	328.15	–0.333	79Lon2
	$Tl + Br^- \rightarrow TlBr + e^-$		–0.392	56Car1
	$Tl + 2Br^- \rightarrow TlBr_2^- + e^-$		–0.395	56Car1
	$Tl + Cl^- \rightarrow TlCl + e^-$, TlCl		–0.372	53Bel, 45Gar
	$Tl + Cl^- \rightarrow TlCl + e^-$, TlCl, HCl		–0.372	55Hu
	$Tl + 2Cl^- \rightarrow TlCl_2^- + e^-$, TlCl, HCl		–0.346	55Hu
	$2Tl + CrO_4^{2-} \rightarrow Tl_2CrO_4 + 2e^-$		–1.056	53Suz
	$Tl + F^- \rightarrow TlF + e^-$, TlF		–0.342	53Bel
	$Tl + IO_3^- \rightarrow TlIO_3 + e^-$, $TlIO_3$		–0.6	53Bel
	$Tl + N_3^- \rightarrow TlN_3 + e^-$		–0.554	52Suz3
	$Tl + 2N_3^- \rightarrow TlN_3^- + e^-$, TlN_3		–0.359	57Nai
	$Tl + NO_3^- \rightarrow TlNO_3 + e^-$, $TlNO_3$		–0.356	57Nai
	$Tl + OH^- \rightarrow TlOH + e^-$		–0.34	51Suz
	$Tl + OH^- \rightarrow TlOH + e^-$		–0.385	53Bel, 56Bel1, 56Bel2
	$Tl + SCN^- \rightarrow TlSCN + e^-$, TlSCN		–0.56 [9])	52Suz2, 56Gre
	$Tl + SCN^- \rightarrow TlSCN + e^-$, TlSCN		–0.384	53Bel
	$Tl + SO_4^{2-} \rightarrow TlSO_4^- + e^-$, TlSCN		–0.418	53Bel
$Tl_{(Hg)}$	$2Tl + H_2SO_4 \rightarrow Tl_2SO_4 + 2H^+ + 2e$		–0.4360	25Ger
Thorium (Th)	$Th \rightarrow Th^{4+} + 4e^-$		–2.1	52Lat
Thulium (Tm)	$Tm^{2+} \rightarrow Tm^{3+} + e^-$		2.0	72Ten
Tin	$Sn \rightarrow Sn^{2+} + 2e^-$		–0.135	23Lew
(Sn)	$Sn \rightarrow Sn^{2+} + 2e^-$, acidic		–0.136	72Moo
	$Sn \rightarrow Sn^{2+} + 2e^-$, $Sn(ClO_4)_2$, $HClO_4$		–1.375	57ElW1
	$Sn^{2+} \rightarrow Sn^{4+} + 2e^-$		0.154	34Hue
	$Sn^{2+} \rightarrow Sn^{4+} + 2e^-$		0.2	29Abh
	$Sn^{2+} \rightarrow Sn^{4+} + 2e^-$, acidic		0.15	72Moo
	$Sn + F^- \rightarrow SnF^+ + 2e^-$		–1.510	55Pau
	$SnH_4 \rightarrow Sn + 4H^+ + 4e^-$		–1.07	57DeZ
Titanium	$Ti \rightarrow Ti^{2+} + 2e^-$		–1.8	52Lat
(Ti)	$Ti^{2+} \rightarrow Ti^{3+} + e^-$		–0.37	24For
	$Ti^{3+} \rightarrow Ti^{4+} + e^-$		–0.04	29Abh
	$Ti^{3+} + H_2O \rightarrow Ti(OH)^{3+} + H^+ + e^-$, HCl or $HClO_4$		–0.055	58Tri
	$Ti^{3+} + H_2O \rightarrow TiO^{2+} + 2H^+ + e^-$		0.10	08Die
	$Ti^{3+} + 2Br^- + H_2O \rightarrow Ti(OH)(Br_2) + H^+ + e^-$, H+ KBr		0.235	62Kar

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
	$Ti(SCN)^{2+}+H_2O \rightarrow Ti(SCN)(OH)^{2+}+H^+ + e^-$, H+ KBr		–0.03	58Tri
Tungsten	$W+8OH^- \rightarrow WO_4^{2-}+4H_2O+6e^-$		–1.5	52Lat
(W)	$W+8OH^- \rightarrow WO_4^{2-}+4H_2O+6e^-$, Na_2WO_4, NaOH		–1.0069	57ElW2
	$W(CN)_8^{4-} \rightarrow W(CN)_8^{3-}+e^-$, K_4WCN_8, K_3WCN_8		0.457	55Baa
Uranium	$U+H_2O \rightarrow UO_2^{2-}+4H^+ + 6e^-$		–0.82	52Lat
(U)	$U_2^+ \rightarrow U_2^{2+}+e^-$		0.062	49Ker
	$U^{4+}+2H_2O \rightarrow UO_2^{2+}+4H^+ + 2e^-$		0.327	61Sob
	$UO_2^+ \rightarrow UO_2^{2+}+e^-$, UO_2NO_3, $NaClO_4$, $HClO_4$		0.062	49Ker
Vanadium	$V \rightarrow V^{2+}+2e^-$		–1.5	52Lat
(V)	$V^{2+} \rightarrow V^{3+}+e^-$, $V_2(SO_4)_3$, VSO_4		–0.255	44Jon2
	$V^{2+} \rightarrow V^{3+}+e^-$		–0.20	30Foe
	$V^{3+}+H_2O \rightarrow VO^{2+}+2H^+ + e^-$		0.314	30Foe
	$VO^{2+}+3H_2O \rightarrow V(OH)_4^+ + H^+ + e^-$		1.000	34Car, 33Cor
	$VO^{2+}+H_2O \rightarrow VO_2^+ + 2H^+ + e^-$, NH_4VO_3, $VOSO_4$		0.991	64Dzh
	$VO^{2+}+H_2O \rightarrow VO_2^+ + 2H^+ + e^-$, NH_4VO_3, $VOSO_4$		1.250	64Dzh
	$V(acetylacetone)_3^- \rightarrow V$ acetylacetone$+e^-$		1.0	65Sch2
Ytterbium	$Yb^{2+} \rightarrow Yb^{3+}+3e^-$, acidic, 1 **M**		–0.58	64Bet
(Yb)	$Yb+SO_4^{2-} \rightarrow YbSO_4^+ + 3e^-$		–2.339	54Spe
	$Yb+oxalate^{2-} \rightarrow Yb(oxalate)^+ + 3e^-$		–2.412	51Cro, 50Cro1
	$Yb+2(oxalate)^{2-} \rightarrow Yb\ oxalate_2^- + 3e^-$		–2.504	51Cro, 50Cro1
Yttrium	$Y+SO_4^{2-} \rightarrow YSO_4^+ + 3e^-$		–2.872	54Spe
(Y)				
Zinc	$Zn \rightarrow Zn^{2+}+2e^-$		–0.763	45Sto1, 40Rob, 34Shr1, 23Lew
(Zn)	$Zn \rightarrow Zn^{2+}+2e^-$, acidic		–0.7628	72Moo
	$Zn \rightarrow Zn^{2+}+2e^-$, $ZnCl_2$		–0.76194	76Lut
$Zn_{(Hg)}$	$Zn_{(Hg)} \rightarrow Zn^{2+}+Hg+2e$	283.15	–0.802	77Lon1
$Zn_{(Hg)}$	$Zn_{(Hg)} \rightarrow Zn^{2+}+Hg+2e^-$	293.15	–0.801	77Lon1
$Zn_{(Hg)}$	$Zn_{(Hg)} \rightarrow Zn^{2+}+Hg+2e^-$	298.15	–0.801	77Lon1
$Zn_{(Hg)}$	$Zn_{(Hg)} \rightarrow Zn^{2+}+Hg+2e^-$	303.15	–0.800	77Lon1
$Zn_{(Hg)}$	$Zn_{(Hg)} \rightarrow Zn^{2+}+Hg+2e^-$	313.15	–0.789	77Lon1
$Zn_{(Hg)}$	$Zn \rightarrow Zn^{2+}+2e^-$, acidic, 1 **M**		–0.7627	64Bet
	$Zn \rightarrow Zn^{2+}+2e^-$, acidic, $ZnBr_2$		–0.7618	45Sto2
	$Zn \rightarrow Zn^{2+}+2e^-$, acidic, $ZnCl_2$		–0.7618	57Car
	$Zn+F^- \rightarrow ZnF^+ + 2e^-$		–0.798	55Pau

Electrode [2])	Electrode reaction, electrolyte system [3]), concentration [4])	T [K]	E_{SHE} [V]	Ref.
Zn	$Zn+3OH^- \rightarrow Zn(OH)_3^- + 2e^-$, NaOH, $Zn(OH)_2$		–1.183	54Ful
	$Zn+4OH^- \rightarrow Zn(OH)_4^{2-} + 2e^-$, KOH		–1.214	54Dir
	$Zn+4OH^- \rightarrow Zn(OH)_4^{2-} + 2e^-$, KOH, $K_2Zn(OH)_4$, KOH		–1.205	71Bod
	$Zn+4OH^- \rightarrow Zn(OH)_4^{2-} + 2e^-$, basic		–1.215	72Moo
	$Zn+S_2O_3^{2-} \rightarrow ZnS_2O_3+2e^-$, $ZnCl_2$, BaS_2O_3		–0.831	51Den
	$Zn+malate^{2-} \rightarrow Zn\ malate+2e^-$, Zn malate		–0.859	54Ful
	$Zn+lactate^{2-} \rightarrow Zn\ lactate+2e^-$, Zn lactate		–0.826	54Dav2
	$Zn+mandelate^{2-} \rightarrow Zn\ mandelate+2e^-$, Zn mandelate		–0.828	51Bel
	$Zn+alanine \rightarrow Zn\ alanine^{2+}+2e^-$, alanine, NaOH, $ZnSO_4$		–0.915	51Mon4
	$Zn+2alanine \rightarrow Zn(alanine)_2^{2+} + 2e^-$, alanine, NaOH, $ZnSO_4$		–1.043	51Mon4
	$Zn+glycine \rightarrow Zn\ glycine^{2+}+2e^-$, glycine, NaOH, $ZnSO_4$		–0.925	51Mon4
	$Zn+2glycine \rightarrow Zn(glycine)_2^{2+} + 2e^-$, glycine, NaOH, $ZnSO_4$		–1.056	51Mon4
	$Zn+glycylglycine \rightarrow Zn\ glycylglycine^{2+}+2e^-$, glycylglycine, NaOH, $ZnSO_4$		–0.874	51Mon4
	$Zn+2glycylglycine \rightarrow Zn(glycylglycine)_2^{2+} + 2e^-$, glycylglycine, NaOH, $ZnSO_4$		–0.955	51Mon4
	$Zn+glycylglycylglycine \rightarrow Zn\ glycylglycylglycine^{2+}+2e^-$, glycylglycylglycine, NaOH, $ZnSO_4$		–0.859	51Mon4
	$Zn+2glycylglycylglycine \rightarrow Zn(glycylglycylglycine)_2^{2+} + 2e^-$, glycylglycylglycine, NaOH, $ZnSO_4$		–0.948	51Mon4
	$Zn+8$-hydroxyquinoline-5-sulfonate$^{2-} \rightarrow$ Zn(8-hydroxyquinoline-5-sulfonate)$+2e^-$, $Zn(NO_3)_2$, 8-hydroxyquinoline-5-sulfonic acid		–1.021	54Näs
	$Zn+2(8$-hydroxyquinoline-5-sulfonate$^{2-}) \rightarrow$ Zn(8-hydroxyquinoline-5-sulfonate)$_2^{2-} + 2e^-$, $Zn(NO_3)_2$, 8-hydroxyquinoline-5-sulfonic acid		–1.241	54Näs
Zirkonium (Zr)	$Zr+H_2O \rightarrow ZrO^{2+}+2H^++4e^-$		–1.5	52Lat
	$Zr+F^- \rightarrow ZrF^{3+}+4e^-$		–1.685	49Con

[1]) All data refer to room temperature if not stated otherwise.

[2]) Amalgam electrodes are denoted with a subscript (Hg), the metal content in the amalgam is stated when provided by the author; in case of standard electrodes the metal content may not be stated explicitly. For details of amalgam electrodes see [85Mus]. Amalgam electrodes are placed in the listings according to the same ordering criteria as non-amalgam ones. If an electronically conducting material different from the metal stated in column 1 is used (e.g. with a redox electrode system) this is not stated explicitly.

[3]) Experimental details are given only if relevant for the stated value of E_{SHE}, solvent is always water.

[4]) If no value is given unity activity has been used and thus a standard electrode potential is given.

[5]) Difference with respect to value calculated from accepted standard potential and activity coefficient is attributed to liquid junction potential.

[6]) Potential contains liquid junction potential.

[7]) Ac = acetate anion.

[8]) Ox = oxalic acid anion.

[9]) Mercury electrode used.

References

02Böd Bödlander, G., Starbeck, O.: Z. Anorg. Allg. Chem. **31** (1902) 458.
08Die Diethelm, B., Foerster, F.: Z. Phys. Chem. (Leipzig) **62** (1908) 129.
10Lew Lewis, G.N., Kraus, Ch.A.: J. Am. Chem. Soc. **32** (1910) 1459.
10Sch2 Schildbach, R.: Z. Elektrochem. **16** (1910) 967.
12Lew Lewis, G.N., Keyes, F.G.: J. Am. Chem. Soc. **34** (1912) 119.
14Lew1 Lewis, G.N., Lacy, W.N.: J. Am. Chem. Soc. **36** (1914) 804.
15Lew Lewis, G.N., Argo, W.L.: J. Am. Chem. Soc. **37** (1915) 1983.
22Ger Gerke, R.H.: J. Am. Chem. Soc. **44** (1922) 1684.
23Cam1 Campbell, A.N.: J. Chem. Soc. **123** (1923) 2323.
23Cam2 Campbell, A.N.: J. Chem. Soc. **123** (1923) 892.
23Gru Grube, G., Schweigardt, F.: Z. Elektrochem. **29** (1923) 257.
23Kra1 Krasanowski, I.A.: Z. Anorg. Allg. Chem. **128** (1923) 17.
23Kra2 Krasanowski, I.A.: Z. Anorg. Allg. Chem. **128** (1923) 3.
23Lew Lewis, G.N., Randall, M.: Thermodynamics and the Free Energy of Chemical Substances, New York: McGraw Hill, 1923.
23Smi Smith, D.F.: J. Am. Chem. Soc. **45** (1923) 360.
23Swi Swift, E.H.: J. Am. Chem. Soc. **45** (1923) 371.
24For Forbes, G.S., Hall, L.P.: J. Am. Chem. Soc. **46** (1924) 385.
24Ric Richards, Th.W., Richards, W.T.: J. Am. Chem. Soc. **46** (1924) 89.
25Ger Gerke, R.H.: Chem. Rev. **1** (1925) 377.
25Mak Makishima, S.: Z. Elektrochem. **31** (1925) 287.
25Neu Neumann, B., Richter, H.: Z. Elektrochem. **31** (1925) 287.
26Fen Fenwick, F.: J. Am. Chem. Soc. **48** (1926) 860.
26Lat Latimer, W.M.: J. Am. Chem. Soc. **48** (1926) 2868.
26Lie Lieder, H., Schachterle, P.: Z. Elektrochem. **32** (1926) 561.
26Mai Maier, Ch.G., Parks, G.S., Anderson, C.Fr.: J. Am. Chem. Soc. **48** (1926) 2564.
27Dav Davies, C.W.: Trans Faraday Soc. **23** (1927) 351.
27Gru Grube, G., Breitinger, G.: Z. Elektrochem. **33** (1927) 112.
27Lat Latimer, W.M.: J. Phys. Chem. **31** (1927) 1267.
27Nie Nielson, R.F., Brown, D.J.: J. Am. Chem. Soc. **49** (1927) 2423.
28Dev Devoto, G.: Z. Elektrochem. **34** (1928) 19.
28Lat Latimer, W.M., Greensfelder, B.S.: J. Am. Chem. Soc. **50** (1928) 2202.
29Abh Abhandlungen der Deutschen Bunsengesellschaft **10** (1929).
29Har1 Haring, M.M., van den Bosche, E.G.: J. Phys. Chem. **33** (1929) 161.
29Sch Scheidt, H.J.: Sitzungsber. Akad. Wiss. Wien, Math. Naturwiss. Kl. **138** (1929) 755.
30Bro Brown, D.J., Liebhafsky, H.A.: J. Am. Chem. Soc. **52** (1930) 2595.
30Dav Davies, C.W.: J. Chem. Soc. (1930) 2410.
30Foe Foerster, F., Böttcher, F.: Z. Phys. Chem. Abt. A **151** (1930) 321.
30Ger Gerding, H.: Z. Phys. Chem. Abt. A **151** (1930) 190.
30Hak Hakomori, S.I.: J. Am. Chem. Soc. **52** (1930) 2373.
30Ran Randall, M., Cann, J.Y.: J. Am. Chem. Soc. **52** (1930) 589.
30Ril Riley, H.: J. Chem. Soc. (1930) 1642.
30Smi Smith, E.R.: J. Res. Natl. Bur. Stand. **5** (1930) 735.
31Ban Banks, W.H., Righellato, E.C., Davies, C.W.: Trans. Faraday Soc. **27** (1931) 621.
31Bur Burian, R.: Z. Elektrochem. **37** (1931) 238.
31Ive Ives, D.J., Riley, H.L.: J. Chem. Soc. (1931) 1938.
31Lar Larsson, E., Adell, B.: Z. Phys. Chem. Abt. A **156** (1931) 352.
31Lew Lewis, G.N., Keyes, F.G.: J. Am. Chem. Soc. **53** (1931) 340.
31Lie Liebhafsky, H.A.: J. Phys. Chem. **35** (1931) 1648.
31Pop Popoff, S., Riddick, J.A., Wirth, V.J., Ough, L.D.: J. Am. Chem. Soc. **53** (1931) 1195.

32Chl Chloupek, J.B., Danes, V.Z.: Collect. Czech. Chem. Commun. **4** (1932) 124
32Hey Heymann, T., Jellinek, K.: Z. Phys. Chem. Abt. A **160** (1932) 34.
32Ran Randall, M., Frandsen, M.: J. Am. Chem. Soc. **54** (1932) 47.
32Sch Schwarz von Bergkampf, E.: Z. Elektrochem. **38** (1932) 847.
33Cor Coryell, C.D., Yost, D.M.: J. Am. Chem. Soc. **55** (1933) 1909.
33Get Getmann, F.H.: Trans. Am. Chem. Soc. **64** (1933) 201.
33Har1 Harned, H.S., Hamer, W.J.: J. Am. Chem. Soc. **55** (1933) 4496.
33LaM La Mer, V.K., Parks, W.G.: J. Am. Chem. Soc. **55** (1933) 4343.
34Arm Armbruster, M.H., Crenshaw, J.L.: J. Am. Chem. Soc. **56** (1934) 2525.
34Bra Bray, W.C., Hershey, A.V.: J. Am. Chem. Soc. **56** (1934) 1893.
34Car Carpenter, J.E.: J. Am. Chem. Soc. **56** (1934) 1874.
34Col Colombier, L.: C. R. Acad. Sci. (Paris) **199** (1934) 273; 408.
34Cow Cowperthwaite, J.A., La Mer, V.K., Barksdale, J.: J. Am. Chem. Soc. **56** (1934) 544.
34Har1 Haring, M.M., Westfall, B.B.: Trans. Am. Electrochem. Soc. **65** (1934) 235.
34Hue Huey, C.S., Tartar, H.V.: J. Am. Chem. Soc. **56** (1934) 2585.
34Kil Kilde, G.: Z. Anorg. Allg. Chem. **218** (1934) 113.
34Lat Latimer, W.M., Schutz, P.W., Hicks, J.F.G.: J. Chem. Phys. **2** (1934) 82.
34Shr1 Shrawder jr., J., Cowperthwaite, I.A., La Mer, V.K.: J. Am. Chem. Soc. **56** (1934) 2348.
34Shr2 Shrawder jr., J., Cowperthwaite, J.A.: J. Am. Chem. Soc. **56** (1934) 2340.
35Hai Haissinsky, M.: J. Chim. Phys. **32** (1935) 116.
35Har Harned, H.S., Hamer, W.J.: J. Am. Chem. Soc. **57** (1935) 27.
35Kol Kolthoff, I.M., Tomssicek, J.: J. Phys. Chem. **39** (1935) 945.
35Sam Samuelson, G.J., Brown, D.J.: J. Am. Chem. Soc. **57** (1935) 2711.
35San Sano, K.: Kinzoku No Kenkyu **12** (1935) 548.
35Win Winnek, P.S., Schmidt, C.L.A.: J. Gen. Physiol. **18** (1935) 889.
36Har1 Harned, H.S., Fitzgerald, M.: J. Am. Chem. Soc. **58** (1936) 2624.
36Hat Hattox, E.M., de Vries, T.: J. Am. Chem. Soc. **58** (1936) 2125.
36Jon Jones, G., Jelen, F.C.: J. Am. Chem. Soc. **58** (1936) 2561.
36Noy Noyes, A.A., Garner, C.S.: J. Am. Chem. Soc. **58** (1936) 1268.
36She Sherill, M.S., Haas jr., A.J.: J. Am. Chem. Soc. **58** (1936) 952.
37Can1 Cann, J.Y., Taylor, A.C.: J. Am. Chem. Soc. **59** (1937) 1484.
37Can2 Cann, J.Y., Taylor, A.C.: J. Am. Chem. Soc. **59** (1937) 1987.
37Far Farayo, K.: Acta Litt. Sci. Regiae Univ. Hung. Francisco-Josephinae Sect. Chem. Mineral. Phys. **5** (1937) 193.
37Har2 Hartley, G.S., Donaldson, G.W.: Trans. Faraday Soc. **33** (1937) 457.
37Sch Schumb, W.C., Sherrill, M.S., Sweetser, S.B.: J. Am. Chem. Soc. **59** (1937) 2360.
38Dav Davies, C.W.: J. Chem. Soc. (1938) 277.
38Har1 Haring, M.M., White, J.C.: Trans. Am. Chem. Soc. **73** (1938) 211.
38Lin Lingane, J.J.: J. Am. Chem. Soc. **60** (1938) 724.
38Owe1 Owen, B.B., Gurry, R.W.: J. Am. Chem. Soc. **60** (1938) 3074.
38Owe2 Owen, B.B., Brinkley, S.R.: J. Am. Chem. Soc. **60** (1938) 2233.
38Qui Quintin, M.: J. Chim. Phys. **35** (1938) 300.
38Tay Taylor, A.C., Nims, L.F.: J. Am. Chem. Soc. **60** (1938) 262.
38Wis Wise, C.W., Davies, C.W.: J. Chem. Soc. (1938) 273.
39Ben Bent, H.E., Forbes, G.S., Forziatti, A.F.: J. Am. Chem. Soc. **61** (1939) 709.
39Har1 Haring, M.M., Hatfield, M.R., Zapponi, P.P.: Trans. Am. Chem. Soc. **75** (1939) 167.
40Dak Dakin, T.W., Ewing, D.T.: J. Am. Chem. Soc. **62** (1940) 2280.
40Lar Larson, W.D.: J. Am. Chem. Soc. **62** (1940) 764.
40Rob Robinson, R.A., Stokes, R.H.: Trans. Faraday Soc. **36** (1940) 740.
40Tay Taylor, A.C.: J. Res. Natl. Bur. Stand. Sect. A **25** (1940) 731.
40Top Topp, N.E., Davies, C.W.: J. Chem. Soc. (1940) 87.
41Dor Dorr, P.D., Stockdale, R.M., Vosburg, W.C.: J. Am. Chem. Soc. **63** (1941) 2670.
41Hov Hovorka, F., Chapman, G.H.: J. Am. Chem. Soc. **63** (1941) 955.
41Mül Müller, F., Reuther, H.: Z. Elektrochem. **47** (1941) 640.

42Bro1 Brosset, C.: Dissertation, Stockholm, 1942.
42Bro2 Brouty, M.L.: C. R. Acad Sci. (Paris) **214** (1942) 258.
42Gel Gelbach, R.W., Brooks King, G.: J. Am. Chem. Soc. **64** (1942) 1054.
42Kri Krieger, K.A., Kilpatrick, M.: J. Am. Chem. Soc. **64** (1942) 7.
42Mac MacDougall, F.H., Allen, M.: J. Phys. Chem. **46** (1942) 730.
42Mül Müller, F., Reuther, H.: Z. Elektrochem. **48** (1942) 682.
43Bro Brosset, C., Orring, J.: Sven. Kem. Tidskr. **55** (1943) 101.
43Hum Hume, D.N., Kolthoff, I.M.: J. Am. Chem. Soc. **65** (1943) 1897.
43Tak Takács, I.: Mag. Chem. Foly. **49** (1943) 33; 100.
43Vos1 Vosburg, W.C., McClure, R.S.: J. Am. Chem. Soc. **65** (1943) 1060.
43Vos2 Vosburg, W.C., Cogswell, S.A.: J. Am. Chem. Soc. **65** (1943) 2412.
44Jon1 Jones, J.H.: J. Am. Chem. Soc. **66** (1944) 1115.
44Jon2 Jones, G., Colvin, J.H.: J. Am. Chem. Soc. **66** (1944) 1563.
44Kil Kilpi, S., Näsänen, R.: Suom. Kemistil. B **17** (1944) 9.
44Näs Näsänen, R.: Suom. Kemistil. B **17** (1944) 31.
45Bez Bezier, D.: Ann. Chim. (Paris) **20** (1945) 161.
45Coa Coates, G.E.: J. Chem. Soc. (1945) 478.
45Gar Garrett, A.G., Vellenga, S.J.: J. Am. Chem. Soc. **67** (1945) 225.
45Jon Jones, J.H.: J. Am. Chem. Soc. **67** (1945) 855.
45Pti Ptitsyn, B.V.: Zh. Neorg. Khim. **15** (1945) 277.
45Sto1 Stokes, J.M.: Trans. Faraday Soc. **41** (1945) 685.
45Sto2 Stokes, R.H., Stokes, J.M.: Trans. Faraday Soc. **41** (1945) 688.
46Dig Dighe, S.G.: J. Indian Chem. Soc. **23** (1946) 291.
46Gap Gapon, E.N.: J. Phys. Chem. USSR **20** (1946) 1209.
46Jon Jones, J.H.: J. Am. Chem. Soc. **68** (1946) 240.
46Kos Kossikoff, A., Sickmann, D.V.: J. Am. Chem. Soc. **68** (1946) 442.
46Ple1 Pleskov, V.A.: Acta Physicochim. URSS **21** (1946) 41; J. Phys. Chem. USSR **20** (1946) 153.
46Ple2 Pleskov, V.A.: Acta Physicochim. URSS **21** (1946) 235.
47Cag Caglioti, V., Sartori, G., Silvestroni, P.: Ric. Sci. **17** (1947) 624.
47Cup Cupr, V.: Publ. Fac. Sci. Univ. Masaryk **296** (1947).
47Dwy1 Dwyer, F.P., Mckenzie, H.A., Nyholm, R.S.: J. Proc. R. Soc. N. S. W. **80** (1947) 183.
47Dwy2 Dwyer, F.P., Mckenzie, H.A., Nyholm, R.S.: J. Proc. R. Soc. N. S. W. **80** (1947) 216.
47Jam James, J.C.: Thesis, London, 1947.
47Mac MacDougall, F.H., Peterson, S.: J. Phys. Chem. **51** (1947) 1346.
47Wit Witt Hutchison, A.: J. Am. Chem. Soc. **69** (1947) 3051.
48Bje Bjerrum, J., Nielson, E.J.: Acta Chem. Scand. **2** (1948) 297.
48Bru Bruehlman, R.J., Verhoak, F.H.: J. Am. Chem. Soc. **70** (1948) 1401.
48Cal Calvin, N., Melcier, N.C.: J. Am. Chem. Soc. **70** (1948) 3270.
48Dav Davies, C.W., James, J.C.: Proc. R. Soc. London, Ser. A **195** (1948) 116.
48Fer Ferrell jr., D.T., Blackburn, I., Vosburgh, W.C.: J. Am. Chem. Soc. **70** (1948) 3812.
48Kee Keefer, R.M.: J. Am. Chem. Soc. 7**0** (1948) 476.
48Mon Monk, C.B.: J. Am. Chem. Soc. **70** (1948) 3281.
48Sto Stock, D.I., Davies, C.W.: Trans. Faraday Soc. **44** (1948) 856.
49Bra Brandt, W.W., Smith, G.F.: Anal. Chem. **21** (1949) 1313.
49Con Connick, R.E., McVey, W.H.: J. Am. Chem. Soc. **71** (1949) 3182.
49Dav1 Davies, C.W., Monk, C.B.: J. Chem. Soc. (1949) 413.
49Dav2 Davies, C.W., Wyatt, P.A.H.: Trans. Faraday Soc. **45** (1949) 770.
49Gri Grinberg, A.A., Orlova, B.Z.: Zh. Prikl. Khim. (Leningrad) **22** (1949) 441.
49Jam James, J.C.: Trans. Faraday Soc. **45** (1949) 855.
49Jon Jones, H.W., Monk, C.B., Davies, C.W.: J. Chem. Soc. (1949) 2693.
49Kee Keefer, R.N., Andrews, L.J., Kepner, R.E.: J. Am. Chem. Soc. **71** (1949) 3906.
49Ker Kern, D.M.H., Orleman, E.F.: J. Am. Chem. Soc. **71** (1949) 2102.
49Kru Krumholz, P.: J. Am. Chem. Soc. **71** (1949) 3654.
49Led Leden, I.: Acta Chem. Scand. **3** (1949) 1318.

49Mal1 Maley, L.E., Mellor, D.P.: Aust. J. Sci. Res. Ser. A **2** (1949) 579.
49Mal2 Maley, L.E., Mellor, D.P.: Aust. J. Sci. Res. Ser. A **2** (1949) 92.
49Mon1 Monk, C.B.: J. Chem. Soc. (1949) 423.
49Mon2 Monk, C.B.: J. Chem. Soc. (1949) 427.
49Mon3 Monk, C.B.: J. Chem. Soc. (1949) 1317.
49Näs Näsänen, R.: Acta Chem. Scand. **3** (1949) 959.
49Sto Stock, D.I., Davies, C.W.: J. Chem. Soc. (1949) 1371.
50Bat Batashev, K.P., Nikitin, E.N.: Zh. Prikl. Khim. (Leningrad) **23** (1950) 263.
50Bje Bjerrum, J.: Chem. Rev. **46** (1950) 381.
50Cro1 Crouthamel, C.E., Martin, D.S.: J. Am. Chem. Soc. **72** (1950) 1382.
50Dav Davies, C.W., Waind, G.M.: J. Chem. Soc. (1950) 301.
50ElW1 El Wakkad, S.E.S., Salem, T.M.: J. Phys. Chem. **54** (1950) 1371.
50ElW2 El Wakkad, S.E.S., Hickling, A.: Trans. Faraday Soc. **46** (1950) 820.
50Jam James, J.C., Monk, C.B.: Trans. Faraday Soc. **46** (1950) 1041.
50Jen Jenkins, I.L., Monk, C.B.: J. Am. Chem. Soc. **72** (1950) 2695.
50Jon Jones, H.W., Monk, C.B.: J. Chem. Soc. (1950) 3475.
50Kor Korshunov, I.A., Nalyugina, N.I.: Zh. Obshch. Khim. **20** (1950) 1309.
50Lar Larson, W.D.: J. Phys. Chem. **54** (1950) 310.
50Mal Maley, I.E., Mellor, D.S.: Nature (London) **165** (1950) 453.
50May Mayer, S.W., Schwartz, S.D.: J. Am. Chem. Soc. **72** (1950) 5106.
50Näs Näsänen, R.: Acta Chem. Scand. **4** (1950) 816.
50Vla Vladimirova, M.G., Kakovskii, I.A.: Zh. Prikl. Khim. **23** (1950) 580.
51Bar Barney, J.E., Argersinger jr., W.J., Reynolds, G.A.: J. Am. Chem. Soc. **73** (1951) 3785.
51Bel Bell, R.P., Waind, G.M.: J. Chem. Soc. (1951) 2357.
51Boc Bockris, J., Haringshow, J.: Usp. Khim. **20** (1951) 216.
51Bon Bonhoeffer, K.F., Jena, W.: Z. Elektrochem. **55** (1951) 151.
51Cro Crouthamel, C.E., Martin, D.S.: J. Am. Chem. Soc. **73** (1951) 569.
51Dav1 Davies, C.W.: J. Chem. Soc. (1951) 1256.
51Dav2 Davies, C.W., Monk, C.B.: J. Chem. Soc. (1951) 2718.
51Den Denney, T.O., Monk, C.B.: Trans. Faraday Soc. **47** (1951) 992.
51Dun Dunsmore, H.S., James, J.C.: J. Chem. Soc. (1951) 2925.
51Gim Gimblett, F.G.R., Monk, C.B.: Trans. Faraday Soc. **47** (1951) 793.
51Hil Hills, G.J., Ives, D.J.G.: J. Chem. Soc. (1951) 305; 311; 318.
51Llo Lloyd, N., Wycherley, V., Monk, C.B.: J. Chem. Soc. (1951) 1786.
51May Mayer, E.W., Schwartz, S.D.: J. Am. Chem. Soc. **73** (1951) 222.
51Mon1 Monk, C.B.: Trans. Faraday Soc. **47** (1951) 285.
51Mon2 Monk, C.B.: Trans. Faraday Soc. **47** (1951) 292.
51Mon3 Monk, C.B.: Trans. Faraday Soc. **47** (1951) 1233.
51Mon4 Monk, C.B.: Trans. Faraday Soc. **47** (1951) 297.
51Pea Peacock, J.M., James, J.C.: J. Chem. Soc. (1951) 2233.
51Sch1 Schwarzenbach, G., Heller, J.: Helv. Chim. Acta **34** (1951) 576.
51Sch2 Schwarzenbach, G., Heller, J.: Helv. Chim. Acta **34** (1951) 1889.
51Suz Suzuki, S.: J. Chem. Soc. Jpn. **72** (1951) 265.
52Bar Barnes, G.T., Dwyer, F.P., Gyarfas, E.C.: Trans. Faraday Soc. **48** (1952) 269.
52Bra Brandt, W.W., Gullstrom, D.K.: J. Am. Chem. Soc. **74** (1952) 3532
52Col1 Colman-Porter, C.A., Monk, C.B.: J. Chem. Soc. (1952) 4363.
52Col2 Colman-Porter, C.A., Monk, C.B.: J. Chem. Soc. (1952) 1312.
52Eva Evans, J.I., Monk, C.B.: Trans. Faraday Soc. **48** (1952) 934.
52Gib Gibby, C.W., Monk, C.B.: Trans. Faraday Soc. **48** (1952) 632.
52Jol Jolly, W.L., Latimer, W.M.: J. Am. Chem. Soc. **74** (1952) 5751.
52Jon Jones, H.W., Monk, C.B.: Trans. Faraday Soc. **48** (1952) 929.
52Lat Latimer, W.M.: The Oxidation States of the Elements and their Potentials in Aqueous Solutions, Englewood Cliffs: Prentice-Hall Inc., 1952.
52Mac MacDougall, F.H., Topol, I.E.: J. Phys. Chem. **56** (1952) 1090.

52Mon Monk, C.B.: J. Chem. Soc. (1952) 1314.
52Reb Rebertus, R.L.: Dissertation, Urbana, 1952.
52Suz1 Suzuki, S.: J. Chem. Soc. Jpn. Pure Chem. Sect. **73** (1952) 278.
52Suz2 Suzuki, S.: J. Chem. Soc. Jpn. Pure Chem. Sect. **73** (1952) 153.
52Suz3 Suzuki, S.: J. Chem. Soc. Jpn. Pure Chem. Sect. **73** (1952) 150.
52Tou Tourky, A.R., Khairy, E.M.: J. Chem. Soc. (1952) 2626.
52Yat Yates, L.M., Dodgen, H.W.: 122nd Meeting of the American Chemical Society 1952, Abstr. Nr. 18P.
53Bel Bell, R.P., George, J.H.B.: Trans. Faraday Soc. **49** (1953) 619.
53Ber1 Berne, B., Leden, I.: Z. Naturforsch. A **8** (1953) 719.
53Ber2 Berg, D., Patterson jr., A.: J. Am. Chem. Soc. **75** (1953) 1484.
53Bon Bonner, O.D., Unietis, F.A.: J. Am. Chem. Soc. **75** (1953) 5111.
53Cav Cave, G.B.C., Hume, D.N.: J. Am. Chem. Soc. **75** (1953) 2893.
53Cob Cobble, J.W., Smith jr., W.T., Boyd, G.E.: J. Am. Chem. Soc. **75** (1953) 5777.
53Dav Davies, C.W., Hoyle, B.E.: J. Chem. Soc. (1953) 4134.
53ElS El Shamy, H.K., El Aggan, A.M.: J. Am. Chem. Soc. **75** (1953) 1187.
53Har Harned, H.S., Paxton, T.R.: J. Phys. Chem. **57** (1953) 531.
53Hep Hepler, L.G., Hugusti, Z.Z.: J. Am. Chem. Soc. **75** (1953) 5652.
53Hie Hieber, W., Hübel, W.: Z. Elektrochem. **57** (1953) 331.
53Kur Kury, J.W.: Thesis, University of California, Berkely, 1953.
53Leu Leussing, D.L., Kolthoff, I.M.: J. Am. Chem. Soc. **75** (1953) 3904.
53Lev2 Levin, A.I., Ukshe, E.A.: Sb. Statei Obshch. Khim. **2** (1953) 798.
53Pat Patrick, W.A., Thompson, W.E.: J. Am. Chem. Soc. **75** (1953) 1184.
53Pec Pecsok, R.: Anal. Chem. **25** (1953) 561.
53Pou Pouradier, J., Chateau, H.: C. R. Acad. Sci. (Paris) **237** (1953) 711.
53Rin Ringbom, A.: Solubility of Sulfides, Report to Analytical Section, IUPAC, July 1953.
53Sal Saltman, W.M., Nachtrieb, N.H.: J. Electrochem. Soc. **100** (1953) 126.
53Suz Suzuki, S.: J. Chem. Soc. Jpn. **74** (1953) 219.
53Van Vanderzee, C.E., Dawson, H.J.: J. Am. Chem. Soc. **75** (1953) 5659.
54Bas Basolo, F., Chen, Y.T.: J. Am. Chem. Soc. **76** (1954) 956.
54Bat2 Bates, R.G., Bowers, V.E.: J. Res. Natl. Bur. Stand. U.S. **53** (1954) 282.
54Dav1 Davies, P.B., Monk, C.B.: Trans. Faraday Soc. **50** (1954) 128.
54Dav2 Davies, P.B., Monk, C.B.: Trans. Faraday Soc. **50** (1954) 132.
54Dir Dirkse, T.P.: J. Electrochem. Soc. **101** (1954) 328.
54Dwy Dwyer, F.P., Gyarfas, E.C.: J. Am. Chem. Soc. **76** (1954) 6320.
54Ful Fulton, J.W., Swinehart, D.F.: J. Am. Chem. Soc. **76** (1954) 864.
54Hep Hepler, L.G., Hugusti jr., Z.Z., Kury, J.W.: J. Phys. Chem. **58** (1954) 26.
54Irv Irving, H.: J. Chem. Soc. (1954) 3494.
54Kra Kraus, K.A., Nelson, F., Smith, G.W.: J. Phys. Chem. **58** (1954) 11.
54Llo Lloyd, R.: Thesis, Temple University Philadelphia, 1954.
54Näs Näsänen, R., Uusitalo, E.: Acta Chem. Scand. **8** (1954) 112.
54Pan1 Pan, K.: J. Chin. Chem. Soc. (Taipei) Ser. II **1** (1954) 16.
54Pan2 Pan, K.: J. Chin. Chem. Soc. (Taipei) Ser. II **1** (1954) 26.
54Scr Scrocco, E., Marmani, G.: Atti Accad. Naz. Lincei Rend. Cl. Sci. Fis. Mat. Nat. **16** (1954) 637.
54Sil Silvermann, M.D., Levy, H.A.: J. Am. Chem. Soc. **76** (1954) 3319.
54Spe Spedding, F.H., Jaffe, S.: J. Am. Chem. Soc. **76** (1954) 882.
54Vle Vleck, A.A.: Chem. Listy **48** (1954) 1474.
55Baa Baadsgaard, H., Treadwell, W.D.: Helv. Chim. Acta **38** (1955) 1669.
55Bab Babtelsky, M., Kertes, S.: Bull. Soc. Chim. Fr. (1955) 328.
55Big Biggs, A.I., Parton, H.N., Robinson, R.A.: J. Am. Chem. Soc. **77** (1955) 5844.
55Dav Davis, J.A.: Thesis, Indiana University, 1955.
55Eva Evans, J.I., Monk, C.B.: Trans. Faraday Soc. **51** (1955) 1244.
55Fyf Fyfe, W.S.: J. Chem. Soc. (1955) 1347.
55Gim Gimblett, F.G.R., Monk, C.B.: Trans. Faraday Soc. **51** (1955) 793.

55Hei Heidt, L.J., Berestecky, J.: J. Am. Chem. Soc. **77** (1955) 2049.
55Hol Holmes, F., Crimmin, W.R.G.: J. Chem. Soc. (1955) 1175.
55Hu Hu, K.H., Scott, A.B.: J. Am. Chem. Soc. **77** (1955) 1380.
55Mar1 Martel, R.W., Kraus, C.A.: Proc. Natl. Acad. Sci. USA **41** (1955) 9.
55Max Maxwell, K.H., Thirsk, H.R.: J. Chem. Soc. (1955) 4054.
55Mur Murmann, R.K., Basolo, F.: J. Am. Chem. Soc. **77** (1955) 3484.
55Nan Nancollas, G.H.: J. Chem. Soc. (1955) 1458.
55Pal Palczewska, W.: Rocz. Chem. **29** (1955) 594.
55Pan Pan, K.: J. Chinese Chem. Soc. (Taipei) Ser. II **2** (1955) 15.
55Pau Paul, A.D.: Thesis, University of California, Berkely, 1955.
55Pec Pecsok, R., Juvet, R.S.: J. Am. Chem. Soc. **77** (1955) 20.
55Pou Poulsen, J., Bjerrum, J.: Acta Chem. Scand. **9** (1955) 1407.
55Ren Renier, J.J., Martin, D.S.: United States Atomic Energy Comission and Iowa State College, contribution 668-1955; see also: Renier, J.J., Martin, D.S.: J. Am. Chem. Soc. **78** (1956) 1833.
55Tan Tanacka, N., Kolthoff, I.M., Strieks, W.: J. Am. Chem. Soc. **77** (1955) 1996
56Bag Bagnall, K.W., Freeman, J.H.: J. Chem. Soc. (1956) 2770.
56Bel1 Bell, R.P., Panckhurst, M.H.: J. Chem. Soc. (1956) 2836.
56Bel2 Bell, R.P., Panckhurst, M.H.: Recl. Trav. Chim. Pays Bas **75** (1956) 725.
56Car1 Carpenter, L.G.: Thesis, Columbia University, 1956.
56Car2 Carrington, A., Symons, M.C.R.: J. Chem. Soc. (1956) 3373.
56Coh Cohen, S.R.: Thesis Cornell University, Ithaka, New York, 1956.
56Dro Droll, H.A.: Thesis, Pennsylvania University, 1956.
56Gre Grenier, J.N., Meites, L.: Anal. Chim. Acta **14** (1956) 482.
56Tac Tacopova, V.F.: Zh. Neorg. Khim. **1** (1956) 243.
56Van Vanderzee, C.E., Smith, W.E.: J. Am. Chem. Soc. **78** (1956) 721.
57Bas Basinski, A., Poczpko, S.: Rocz. Chem. **31** (1957) 449.
57Car Carraro, G., Stephens, H.L.: J. Electrochem. Soc. **104** (1957) 512.
57Cha3 Chateau, H., Hervier, B.: J. Chim. Phys. **54** (1957) 356.
57Con2 Connick, R.E., Hurley, C.R.: J. Phys. Chem. **61** (1957) 1018.
57DeZ De Zoubov, N., Deltombe, E.: Proc. Int. Comm. Electrochem. Thermodynamics and Kinetics, 1957, p. 240.
57ElW1 El Wakkad, S.E.S., Salem, T.M., Al Sayed, J.A.: J. Chem. Soc. (1957) 3770.
57ElW2 El Wakkad, S.E.S., Salem, T.M., Risk, H.A., Ebaid, J.G.: J. Chem. Soc. (1957) 3776.
57Geo George, P., Hanania, G.I., Irvine, D.H.: J. Chem. Soc. (1957) 3048.
57Gup Gupta, S.R., Hills, G.J., Ives, D.J.G.: Discuss. Faraday Soc. **24** (1957) 147.
57Ham Hamer, W.G., Craig, D.N.: J. Electrochem. Soc. **104** (1957) 206.
57Kin King, J.P., Cobble, J.W.: J. Am. Chem. Soc. **79** (1957) 1559.
57Kiv Kivalo, P., Ekari, P.: Suom. Kemistil. B **30** (1957) 116.
57Kor Kortüm, G.: Lehrbuch der Elektrochemie, Weinheim: Verlag Chemie, 1957.
57Nai Nair, V.S.K., Nancollas, G.H.: J. Chem. Soc. (1957) 318.
58Grz Grzybowski, A.K.: J. Phys. Chem. **62** (1958) 550
58Hie Hieber, W., Wagner, G.: Z. Naturforsch. B **13** (1958) 339.
58Sch Schwabe, K., Ziegenbalg, S.: Z. Elektrochem. **62** (1958) 172.
58Tre1 Treumann, W.B., Ferris, L.M.: J. Am. Chem. Soc. **80** (1958) 5048.
58Tre2 Treumann, W.B., Ferris, L.M.: J. Am. Chem. Soc. **80** (1958) 5050.
58Tri Tribalat, S., Delafosse, D.: Anal. Chim. Acta **19** (1958) 74.
58Ver Verbeek, F., Eechebaut, Z.: Bull. Soc. Chim. Belg. **67** (1958) 204
59Geo George, P., Hanania, G.I.H., Irvine, D.H.: J. Chem. Soc. (1959) 2548.
59Gol Golding, R.M.: J. Chem. Soc. (1959) 1838.
59Par Parsons, R.: Handbook of Electrochemical Constants, London: Butterworths, 1959.
59Per Perin, D.D.: J. Chem. Soc. (1959) 290.
59Ric Ricketts, J.A., Tresselt, L.W.: J. Am. Chem. Soc. **81** (1959) 2305.
60Gre Greeley, R.S., Smith, W.T., Stoughton, R.W., Lietzke, M.H.: J. Phys. Chem. **64** (1960) 652; 1445.

60Hur Hurlen, T.: Acta Chem. Scand. **14** (1960) 1533.
60Mat Mattoo, B.N.: Zh. Prikl. Khim. **33** (1960) 2015.
60Tow Towns, M.B., Greeley, Lietzke, N.H.: J. Phys. Chem. **64** (1960) 1861.
61Dir Dirkse, T.P.: J. Chem. Eng. Data **6** (1961) 538.
61Hur2 Hurlen, T.: Acta Chem. Scand. **15** (1961) 630.
61Ive Ives, D.J.G., Janz, G.J.: Reference Electrodes, New York: Academic Press, 1961.
61Jak1 Jakuszewski, B.J., Taneewska-Osinska, S.: Zesz. Nauk Uniw. Lodz. Ser. **2** (1961) 10.
61Lin Lin, C.L., Pan, K.: J. Chin. Chem. Soc. (Taipei) **8** (1961) 14.
61Saw Sawyer, D.T., Interrante, L.V.: J. Electroanal. Chem. **2** (1961) 310.
61Sir Sircair, S.C., Prasad, B.: J. Indian Chem. Soc. **38** (1961) 361.
61Sob Sobkovski, J., Minc, S.J.: J. Inorg. Nucl. Chem. **19** (1961) 101.
61Vyd Vydra, F., Pribil, R.: Talanta **8** (1961) 824.
61Yag Yagn, N.S., Bunazhnov, F.T., Sokolova, V.G.: Zap. Leningr. Gorn. Inst. im. G. V. Plekhanova **42** (1961) 35.
62Awa Awad, S.A.: J. Phys. Chem. **66** (1962) 890.
62Ber Berka, A., Dvorak, V., Nemec, I., Zyka, J.: J. Electroanal. Chem. **4** (1962) 150.
62Cov Covington, A.K., Cressey, T., Lever, B.G., Thirsk, H.R.: Trans. Faraday Soc. **58** (1962) 1975.
62Dir Dirkse, T.P.: J. Electrochem. Soc. **109** (1962) 173.
62Gre Gregor, V., Pitzner, K.S.: J. Am. Chem. Soc. **84** (1962) 2671.
62Gub Gubin, S.P., Perevalova, E.G.: Dokl. Akad. Nauk SSSR **143** (1962) 1351.
62Het Hetzer, H.B., Robinson, R.A., Bates, R.G.: J. Phys. Chem. **66** (1962) 1423.
62Jak2 Jakuszewski, B., Taniewska-Osinka, S.: Rocz. Chem. **36** (1962) 329.
62Kar Kartushinskaya, A.I., Stromberg, A.G.: Zh. Neorg. Khim. **7** (1962) 291.
62Lie Lietzke, N.H., Stoughton, R.W.: J. Chem. Educ. **39** (1962) 230.
62Lin Lingane, J.J.: J. Electroanal. Chem. **4** (1962) 332.
62Sch2 Schwabe, K., Nebel, D.: Z. Phys. Chem. (Leipzig) **220** (1962) 339.
62Vag Vagramyan, A.T., Uvarov, L.A.: Izv. Akad. Nauk SSSR, Otd. Khim. Nauk (1962) 1520.
63Cav Cavington, A.K., Hakeem, M.A., Wynne-Jones, W.F.K.: J. Chem. Soc. (1963) 4394.
63Che Chebrikova, Z.M., Belaya, Zh.V., Loshkarev, M.A.: Tr. Dnepropetr. Khim. Tekhnol. Inst. (1963) 55.
63Eva Evans, D.H., Lingane, J.J.: J. Electroanal. Chem. **6** (1963) 1.
63Gri Grinberg, A.A., Kiseleva, N.N., Gelfman, N.I.: Dokl. Akad. Nauk SSSR **153** (1963) 1327.
63Gup1 Gupta, S.R., Hills, G.J., Ives, D.J.G.: Trans. Faraday Soc. **59** (1963) 1886.
63Gup2 Gupta, S.R., Hills, G.J., Ives, D.J.G.: Trans. Faraday Soc. **59** (1963) 1874.
63Lar Larson, W.D.: J. Phys. Chem. **67** (1963) 937.
63Mur Murgulescu, J.G., Constantinescu, E.: Rev. Chim. Acad. Repub. Pop. Roum. **8** (1963) 5.
63Sch2 Schilt, A.A.: Anal. Chem. **35** (1963) 1599.
64Alt Altynov, V.I., Pritsyyn, B.V.: Zh. Neorg. Khim. **9** (1964) 2407.
64Bet de Bethune, A.J., Swendeman Loud, N.A.: Standard Aqueous Electrode Potentials and Temperature Coefficients at 25 °C, Skokie: C.A. Hampel, 1964.
64But2 Butler, J.N.: J. Phys. Chem. **68** (1964) 1828.
64Cov Covington, A.K., Talukdev, P.K., Thirsk, H.R.: Trans. Faraday Soc. **60** (1964) 412.
64Dzh Dzhabarov, F.Z., Gorbachev, S.V.: Zh. Neorg. Khim. **9** (1964) 2399.
64Het Hetzer, H.B., Robinson, R.A.: J. Phys. Chem. **68** (1964) 1929.
64Pan Panson, A.J.: J. Phys. Chem. **68** (1964) 1721.
64Rao Rao, J.P., Murthy, A.R.V.: J. Phys. Chem. **68** (1964) 1573.
64Roc Rock, P.A.: Inorg. Chem. **3** (1964) 1593.
65And Andicott, J.F., Taube, H.: Inorg. Chem. **4** (1965) 437.
65Bab Babaeva, A.V., Khananova, E.Ya.: Zh. Neorg. Khim. **10** (1965) 2579.
65Cha Chang, J.C., Garner, C.S.: Inorg. Chem. **4** (1965) 209.
65Cia Ciavatta, L., Grimaldi, M.: J. Inorg. Chem. **31** (1965) 3071
65Cov Covington, A.K., Dobson, J.V., Jones, L.W.: Trans. Faraday Soc. **61** (1965) 2050.
65Kha Khairy, E.M., Darwish, N.A.: Z. Phys. Chem. (Leipzig) **230** (1965) 327.

65Kor Kortüm, G., Häussermann, W.: Ber. Bunsenges. Phys. Chem. **69** (1965) 594.
65Mac Macero, D.J., Anderson, L.B., Malachesky, P.: J. Electroanal. Chem. **10** (1965) 76.
65Nes Nesmeyanov, A.N., Perevalova, E.G., Yureva, L.P., Gubin, S.P.: Izv. Akad. Nauk SSSR Ser. Khim. (1965) 909.
65Pap Papon, C., Jacq, J.: Bull. Soc. Chim. Fr. (1965) 13.
65Rei Reid Jr., W.E.: J. Phys. Chem. **69** (1965) 3168.
65Roc Rock, P.A.: Inorg. Chem. **4** (1965) 1667.
65Sch1 Schwabe, K., Ferse, E.: Ber. Bunsenges. Phys. Chem. **69** (1965) 383.
65Sch2 Schaefer, W.P.: Inorg. Chem. **4** (1965) 624.
65Tan1 Tanaka, N., Shirikashi, T.: Bull. Chem. Soc. Jpn. **38** (1965) 1515.
65Tan2 Tanaka, N., Ogino, H.: Bull. Chem. Soc. Jpn. **38** (1965) 1054.
66Buc1 Buckley, R.R., Mecer, E.E.: J. Phys. Chem. **70** (1966) 3103.
66Buc2 Buckingham, D.A., Dwyer, F.P., Sargeson, A.M.: Inorg. Chem. **5** (1966) 1243.
66Bur Burnett, J.L., Zirin, M.H.: J. Inorg. Nucl. Chem. **28** (1966) 902.
66Cha Chauchard, J., Gauthier, J.: Bull. Soc. Chim. Fr. (1966) 2635.
66Che Chernyaev, I.I., Muraveiskaya, G.S., Korablina, L.S.: Zh. Neorg. Khim. **11** (1966) 1339.
66Far Farba, F., Iwamoto, R.T.: Anal. Chem. **38** (1966) 143.
66Gol Goldberg, R.N., Riddell, R.G., Wing, M.R., Hopkins, H.P., Wuffs, C.A., Heple, L.G.: J. Phys. Chem. **70** (1966) 706.
66Han Hanania, G.I.H., Irvine, D.H., Michaelides, M.S., Shurayh, F.S.: Proc. 9th Int. Conf. Coord. Chem., St. Moritz-Bad, 1966, p. 222.
66Hoa Hoare, J.P.: Electrochim. Acta **11** (1966) 203.
66Hub Hubbard, A.T., Anson, F.C.: Anal. Chem. **38** (1966) 58.
66Kra2 Kravtsov, V.I., Zelensky, M.I.: Elektrokhimiya **2** (1966) 1138.
66Mal1 Malik, W.U.: Indian J. Chem. **4** (1966) 106.
66Mal2 Malik, W.U., Das, A.: Indian J. Chem. **4** (1966) 203.
66Mat Mathur, P.B., Naqvi, S.M.A.: Proc. Symp. Electrode Processes, 1966, p. 39.
66Pou Pouradier, J., Gadet, M.G.: J. Chim. Phys. **63** (1966) 1467.
67Ber Bertram, J., Bone, S.J.: Trans. Faraday Soc. **63** (1967) 415.
67Bol Boldin, R.V.: Elektrokhimiya **3** (1967) 1259.
67Cov Covington, A.K., Dobson, J.V., Jones, L.W.: Electrochim. Acta **12** (1967) 513.
67ElA El-Awady, A.A., Bounsall, E.J., Garner, C.S.: Inorg. Chem. **6** (1967) 79.
67Ham3 Hamm, R.E., Suwyn, M.A.: Inorg. Chem. **6** (1967) 139.
67Han Hanania, G.I.H., Irvine, D.H., Eaton, W.A., George, P.: J. Phys. Chem. **71** (1967) 2022.
67Hin Hin-Fat, L., Higginson, W.C.E.: J. Chem. Soc. A (1967) 298.
67Iza Izatt, R.M., Eatough, D., Christensen, J.J.: J. Chem. Soc. A (1967) 1301.
67Sch2 Schurig, H., Heusler, K.E.: Fresenius Z. Anal. Chem. **224** (1967) 45.
67Sen1 Sen, B., Johnson, D.A., Roy, R.N.: J. Phys. Chem. **71** (1967) 1523.
67Tou Tourchy, A.R., Abdul-Azim, A.A., Shalaby, L.A.: J. Chem. U.A.R. **10** (1967) 253.
67Van Van Vagramanian, A.T., Leshawa, T.I.: Z. Phys. Chem. **234** (1967) 57.
68Bon Bonnet, N., Mihalova, V.: Dokl. Bolg. Akad. Nauk **21** (1968) 1295.
68Dil Dill, A.J., Itzkovits, V.: J. Phys. Chem. **72** (1968) 4580
68Gri Grinberg, A.A., Dkhara, S.C., Gelfman, M.I.: Zh. Neorg. Khim. **13** (1968) 2199.
68Han Hanania, G.I.H., Irvine, D.H., Shurayh, F.R.: J. Phys. Chem. **72** (1968) 1355.
68Hus Huston, R., Butler, J.N.: J. Phys. Chem. **72** (1968) 4263.
68Mat Mathur, P.B., Naqvi, S.M.A.: Indian J. Chem. **6** (1968) 311.
68Mur1 Murray jr., R.C., Rock, P.A.: Electrochim. Acta **13** (1968) 969.
68Mur2 Murgulescu, I.G., Vartires, I.: Rev. Roum. Chim. **13** (1968) 1397.
68Naq Naqvi, S.M., Mathur, P.B.: Bull. Acad. Pol. Sci. Ser. Sci. Chim. **16** (9) (1968) 479.
68Noz Nozakzee, N.G., Kazakov, V.A., Vagramyan, A.T.: Elektrokhimiya **4** (1968) 1464.
68Sha Sharma, L., Sahu, G., Prasad, B.: J. Indian Chem. Soc. **45** (1968) 580.
68Shu Shu-Chen, L., Pan, K.: J. Chin. Chem. Soc. (Taipei) **15** (1968) 106.
69Ere Erenburg, A.M., Peshchevitskii, B.I.: Zh. Neorg. Khim. **14** (1969) 932.
69Kim Kim, J.J., Rock, P.A.: Inorg. Chem. **8** (1969) 563.

69Kuk Kukushkin, Yu.N., Drokina, Z.V.: Zh. Neorg. Khim. **14** (1969) 2452.
69Mas Mašek, J., Dempir, J.: Collect. Czech. Chem. Commun. **34** (1969) 727.
69Pou Pouradier, J., Gadet, M.G.: J. Chim. Phys. Phys. Chim. Biol. **66** (1969) 109.
69Vas Vasil'jev, V.P., Grechina, N.K.: Elektrokhimiya **5** (1969) 426.
70Leu Leuscke, W., Schwabe, K.: J. Electroanal. Chem. **25** (1970) 219.
70Sha2 Sharma, L., Prasad, B.: J. Indian Chem. Soc. **47** (1970) 379.
70Sin Sinha, H.K., Prasad, B.: J. Indian Chem. Soc. **47** (1970) 901.
70Tan Tanaka, N., Tomita, T., Yamada, A.: Bull. Chem. Soc. Jpn. **43** (1970) 2042.
71Amo Amosse, J., Ruband, M., Barbier, M.J.: C. R. Acad. Sci. (Paris), Ser. C **273** (1971) 1708.
71Bar3 Bartelt, H., Prügel, M.: Electrochim. Acta **16** (1971) 1815.
71Bod Boden, D.P., Wylie, R.B., Spera, V.J.: J. Electrochem. Soc. **118** (1971) 1298.
71Cov Covington, A.K., Srinivasan, K.V.: J. Chem. Thermodyn. **3** (1971) 795.
71DeC De Cugnac-Pailliotet, A., Pouradier, J.: C. R. Acad. Sci. (Paris), Ser. C **273** (1971) 1565.
71Kom1 Komar, N.P., Kaftanov, A.Z.: Zh. Anal. Khim. **26** (1971) 2086.
71Kom2 Komandenko, V.M.: Elektrokhimiya **7** (1971) 1255.
71Mus Mussini, T., Maina, T., Pagella, T.: J. Chem. Thermodyn. **3** (1971) 281.
71Pan Pan, K., Chang, J.-J., Hsin, S.-C.: J. Chin. Chem. Soc. (Taipei) **18** (1971) 1.
71Vas1 Vasil'ev, V.P., Shorokhova, V.I.: Elektrokhimiya **7** (1971) 185.
71Vas2 Vasil'ev, V.P., Glavin, S.R.: Elektrokhimiya **7** (1971) 1395.
72Gad Gadet, M.C., Pouradier, J.: C. R. Acad. Sci. (Paris), Ser. C **275** (1972) 1061.
72Kom Komar, N.P., Kaftanov, A.Z., Dunai, B.A.: Elektrokhimiya **8** (1972) 1177.
72Lav Lavrenova, L.G., Zegzhda, T.V., Shul'man, L.M.: Elektrokhimiya **8** (1972) 1153.
72Moo Moore, W.J.: Physical Chemistry, London: Longman, 1972.
72Mus Mussini, T., Longhi P., Riva, G.: J. Chem. Thermodyn. **4** (1972) 591.
72Ten Tendler, Y., Faraggi, M.: J. Chem. Phys. **57** (1972) 1358.
73Cho Choudhary, B.K., Prasad, B.: Indian J. Chem. **11** (1973) 931.
74Das Das, R.C., Sahu, G., Satyanarayana, D., Misra, S.N.: Electrochim. Acta **19** (1974) 887.
74Gro Gronlund, F., Noer, S.: J. Electrochem. Soc. **121** (1974) 25.
74Lon Longhi, P., Mussini, T., Osimani, C.: J. Chem. Thermodyn. **6** (1974) 227.
74Ost Ostannii, N.I., Zharkova, L.A., Erofeev, B.V., Kazakova, G.D.: Zh. Fiz. Khim. **48** (1974) 358.
74Ski Skibsted, L.H., Bjerrum, J.: Acta Chem. Scand. A **28** (1974) 764.
75Das3 Dash, U.N., Mohanty, J.: Thermochim. Acta **12** (1975) 189.
75Lal Lal, S.C., Prasad, B.: Indian J. Chem. **13** (1975) 372.
75Lon1 Longhi, P., Mussini, T., Vaghi, E.: J. Chem. Thermodyn. **7** (1975) 767.
75Lon2 Longhi, P., Mussini, T., Vaghi, E.: Chim. Ind. (Milan) **57** (1975) 169.
75Mal Malik, W.U., Bembi, R., Dwivedi, J.K.: Indian J. Chem. **13** (1975) 589.
76Ard Ardizzone, S., Longhi, P., Mussini, T., Rondinini, S.: J. Chem. Thermodyn. **8** (1976) 557.
76Lut Lutfullah, Dunsmore, H.S., Paterson, R.: J. Chem. Soc. Faraday Trans. 1 **72** (1976) 495.
77Gio Giordano, G.M., Longhi P., Mussini, T., Rondinini, S.: J. Chem. Thermodyn. **9** (1977) 997.
77Lon1 Longhi, P., Mussini, T., Rondinini, S., Vaghi, E: Chim. Ind. (Milan) **59** (1977) 249.
77Lon2 Longhi, P., Rondinini, S., Ardizzone, S., Mussini, T.: Ann. Chim. (Rome) **67** (1977) 177.
77Mus Mussini, T., Longhi, P.: Chim. Ind. (Milan) **59** (1977) 163.
77Ron Rondinini, S., Longhi, P., Mussini, T.: Ann. Chim. (Rome) **67** (1977) 305; 467.
78Lag Lagowski, J.J.: The Chemistry of Non-aqueous Solutions, Vol. 1, London: Academic Press, 1978.
79Lon1 Longhi, P., Mussini, T., Rondinini, S., Sala, B: J. Chem. Thermodyn. **11** (1979) 359.
79Lon2 Longhi, P., Mussini, T., Rondinini, S., Sala, B: J. Chem. Thermodyn. **11** (1979) 73.
83Maz Mazzarese, J., Popvych, O.: J. Electrochem. Soc. **130** (1983) 2032.
85Mus Mussini, T., Rondinini, S., Longhi, P.: Pure Appl. Chem. **57** (1985) 169.
90Gal Galster, H.: pH-Messung, Weinheim: VCH, 1990.

Table 1.2. Nonmetal electrode potentials in aqueous electrolyte systems. [1])

Electrode	Electrode reaction, electrode system [2])	T [K]	E_{SHE} [V]	Ref.
Boron (B)	$B+3H_2O \rightarrow H_3BO_3+3H^+ +3e^-$		–0.73	52Lat
Bromine (Br)	$2Br^- \rightarrow Br_2+2e^-$, HBr, Br_2		1.0873	66Mus
	$2Br^- \rightarrow Br_2+2e^-$		1.066	33Ish1, 17Lew1, 23Lew
	$2Br^- \rightarrow Br_2+2e^-$, acidic		1.0652	72Moo
	$2Br^- \rightarrow Br_2+2e^-$, KBr, KCl, Br_2, acetic acid		1.078	71Lav
	$Br^- + H_2O \rightarrow HBrO+H^+ +2e^-$		1.33	52Lat
Chlorine (Cl)	$Cl_2+2H_2O \rightarrow 2HClO+2H^+ +2e^-$, NaClO, HClO		1.611	59Fli
	$2Cl^- \rightarrow Cl_2+2e^-$, HCl		1.35827	69Cer
	$2Cl^- \rightarrow Cl_2+2e^-$		1.359	22Ger, 26Kam, 11Lew, 23Lew, 28Ran
	$2Cl^- \rightarrow Cl_2+2e^-$, acidic		1.3595	72Moo
	$Cl^- + H_2O \rightarrow HClO+H^+ +e^-$, NaClO, HClO		1.482	59Fli
	$Cl^- + 4H_2O \rightarrow ClO_4^- + 8H^+ + 8e^-$		1.35	52Lat
	$ClO_2^- \rightarrow ClO_2+e^-$, $NaClO_2$, $AgClO_2$, $Pb(ClO_2)_2$		0.954	62Nai
	$ClO_2^- \rightarrow ClO_2+e^-$, $NaClO_2$, ClO_2		0.936	59Tro
Deuterium (D)	$½D_2 \rightarrow D^+ +e^-$		–0.003	35Hai
	$½D_2 \rightarrow D^+ +e^-$, acidic		–0.0028	72Moo
Fluorine (F)	$2F^- \rightarrow F_2+2e^-$		2.85	26Lat
Hydrogen (H)	$H_2+2OH^- \rightarrow 2H_2O+2e^-$, basic		–0.82806	72Moo
	$I_2+2HCN \rightarrow 2ICN+2H^+ +2e^-$, $HClO_4$, NaCN, I_2		0.625	56Bow
Iodine (I)	$I^- + H_2O \rightarrow HIO+H^+ +2e^-$		0.536	11Bra
	$2I^- \rightarrow I_2+2e^-$		0.615	71Lav
	$2I^- \rightarrow I_2+2e^-$		0.536	23Lew, 29Ish, 37Wil
	$2I^- \rightarrow I_2+2e^-$, acidic		0.5355	72Moo
	$3I^- \rightarrow I_3^- +2e^-$		0.535	29Ish, 52Lat
	$I_2+6H_2O \rightarrow 2IO_3^- +12H^+ +10e^-$, $HClO_4$, HIO_3		1.1942	75Spi
	$I_2+6H_2O \rightarrow 2IO_3^- +12H^+ +10e^-$		1.19	29Abh
Nitrogen (N)	$NO+H_2O \rightarrow NO^+ +e^-$, NO, H_2SO_4		1.46	66Sch
	$NO+H_2O \rightarrow NO^+ +e^-$, NO, H_2SO_4		1.45	67Sch1, 68Mus

Electrode	Electrode reaction, electrode system [2])	T [K]	E_{SHE} [V]	Ref.
N	$NO+H_2O \rightarrow HNO_2+H^+ +e^-$, $NaNO_2$, H_2SO_4		0.983	67Sch1
	$2NO^+ +2H_2O \rightarrow N_2O_4+4H^+ +2e^-$, N_2O_4,H_2SO_4		0.77	68Mus
	$HNO_2+H_2O \rightarrow HNO_3+2H^+ +2e^-$, $NaNO_2$,HNO_3		0.945	71Ale
Lead dioxide	$PbO_2+4H^+ + SO_4^{2-} \rightarrow PbSO_4+2H_2O+2e^-$	273	1.67694	35Ham1
(PbO_2)	$PbO_2+4H^+ + SO_4^{2-} \rightarrow PbSO_4+2H_2O+2e^-$	278	1.67846	35Ham1
	$PbO_2+4H^+ + SO_4^{2-} \rightarrow PbSO_4+2H_2O+2e^-$	283	1.67998	35Ham1
	$PbO_2+4H^+ + SO_4^{2-} \rightarrow PbSO_4+2H_2O+2e^-$	288	1.68159	35Ham1
	$PbO_2+4H^+ + SO_4^{2-} \rightarrow PbSO_4+2H_2O+2e^-$	293	1.68322	35Ham1
	$PbO_2+4H^+ + SO_4^{2-} \rightarrow PbSO_4+2H_2O+2e^-$	298	1.68488	35Ham1
	$PbO_2+4H^+ + SO_4^{2-} \rightarrow PbSO_4+2H_2O+2e^-$	303	1.68671	35Ham1
	$PbO_2+4H^+ + SO_4^{2-} \rightarrow PbSO_4+2H_2O+2e^-$	308	1.68847	35Ham1
	$PbO_2+4H^+ + SO_4^{2-} \rightarrow PbSO_4+2H_2O+2e^-$	313	1.69036	35Ham1
	$PbO_2+4H^+ + SO_4^{2-} \rightarrow PbSO_4+2H_2O+2e^-$	318	1.69231	35Ham1
	$PbO_2+4H^+ + SO_4^{2-} \rightarrow PbSO_4+2H_2O+2e^-$	323	1.69436	35Ham1
	$PbO_2+4H^+ + SO_4^{2-} \rightarrow PbSO_4+2H_2O+2e^-$	328	1.69649	35Ham1
	$PbO_2+4H^+ + SO_4^{2-} \rightarrow PbSO_4+2H_2O+2e^-$	333	1.69861	35Ham1
Oxygen (O)	$2OH^- \rightarrow ½O_2+H_2O+2e^-$		0.400	10Brö, 06Lew, 14Lew2, 23Lew, 06Ner, 27Nie
	$3OH^- \rightarrow HO_2^- +H_2O+2e^-$, basic		0.878	72Moo
	$2H_2O \rightarrow O_2+4H^+ +4e^-$		1.229	66Hoa
	$H_2O_2 \rightarrow O_2+2H^+ +2e^-$, O_2, H_2O_2		0.695	54Ker
Phosphorus (P)	$H_2PO_2^- +3OH^- \rightarrow HPO_3^{2-} +2e^-$, basic		–1.565	72Moo
Quinone	hydroquinone→quinone+$2e^- +2H^+$	273	0.71798	54Bat1
	hydroquinone→quinone+$2e^- +2H^+$	278	0.71437	54Bat1
	hydroquinone→quinone+$2e^- +2H^+$	283	0.71073	54Bat1
	hydroquinone→quinone+$2e^- +2H^+$	288	0.70709	54Bat1
	hydroquinone→quinone+$2e^- +2H^+$	293	0.70343	54Bat1
	hydroquinone→quinone+$2e^- +2H^+$	298	0.69976	54Bat1
	hydroquinone→quinone+$2e^- +2H^+$	303	0.69607	54Bat1
	hydroquinone→quinone+$2e^- +2H^+$	308	0.69237	54Bat1
	hydroquinone→quinone+$2e^- +2H^+$	313	0.68865	54Bat1
	hydroquinone→quinone+$2e^- +2H^+$, E_0 [V]=0.69969–735.8·10^{-6}(T–273.15 K)–0.29·10^{-6}(T–273.15 K)			33Har2
	hydroquinone→quinone+$2e^- +2H^+$, pH=0		0.6990	22LaM

Electrode	Electrode reaction, electrode system [2])	T [K]	E_{SHE} [V]	Ref.
	Toluhydroquinone→Toluquinone+2e$^-$+2H$^+$, pH=0	298	0.6454	22LaM
	p-Xylohydroquinone→*p*–Xyloquinone+2e$^-$+2H$^+$, pH=0		0.5900	22LaM
	p-Thymohydroquinone→*p*-Thymoquinone+2e$^-$+2H$^+$, pH=0		0.5875	22LaM
	Dimethoxyhydroquinone→Dimethoxyquinone+2e$^-$+2H$^+$, pH=0		0.5139	22LaM
	1,4-Naphthohydroquinone→1,4-Naphthoquinone+2e$^-$+2H$^+$, pH=0		0.4698	22LaM
	Bromohydroquinone→Bromoquinone+2e$^-$+2H$^+$, pH=0		0.7151	22LaM
	Chlorohydroquinone→Chloroquinone+2e$^-$+2H$^+$, pH=0		0.7125	22LaM
	hydroquinone→quinone+2e$^-$+2H$^+$, 0.01 **N** HCl, 0.09 **N** KCl	291	0.5750	61Ive
	hydroquinone→quinone+2e$^-$+2H$^+$, 0.01 **N** HCl, 0.09 **N** KCl	298	0.5831	61Ive
	hydroquinone→quinone+2e$^-$+2H$^+$, 0.01 **N** HCl	291	0.5787	61Ive
	hydroquinone→quinone+2e$^-$+2H$^+$, 0.01 **N** HCl	298	0.5868	61Ive
	hydroquinone→quinone+2e$^-$+2H$^+$, 0.1 **N** HCl	291	0.6345	61Ive
	hydroquinone→quinone+2e$^-$+2H$^+$, 0.1 **N** HCl	298	0.6414	61Ive
Selenium (Se)	Se+3H$_2$O→H$_2$SeO$_3$+4H$^+$+4e$^-$, H$_2$SeO$_3$		–0.74	66Osm
Sulfur (S)	S^{2-}→S+2e^{-2}		–0.51	35Kim, 20Noy, 23Lew
	Ag$_2$S→S+2Ag$^+$+2e$^-$, H$_2$S, HCl		–1.99	51Goa1
	HgS→S+Hg^{2+}+2e$^-$, H$_2$S, HCl		–2.00	51Goa2
	PbS→S+Pb^{2+}+2e$^-$		–1.296	56Kiv
	Ag$^+$+4S+6H$_2$O→ Ag(S$_2$O$_3$)$_2^{3-}$ +12H$^+$+8e$^-$		0.365	57Cha1, 57Cha2
	Ba^{2+}+2S+3H$_2$O→BaS$_2$O$_3$+6H$^+$+4e$^-$		0.432	49Dav2
	Ba^{2+}+2S+3H$_2$O→BaS$_2$O$_3$+6H$^+$+4e$^-$, BaS$_2$O$_3$		0.394	51Den
	Ca^{2+}+2S+3H$_2$O→CaS$_2$O$_3$+6H$^+$+4e$^-$, CaCl$_2$, BaS$_2$O$_3$		0.436	51Den, 49Dav1, 55Gim
	Cd^{2+}+2S+3H$_2$O→CdS$_2$O$_3$+6H$^+$+4e$^-$, CdCl$_2$, Na$_2$S$_2$O$_3$, BaS$_2$O$_3$		0.407	51Den, 49Dav1, 55Gim
	Co^{2+}+2S+3H$_2$O→CoS$_2$O$_3$+6H$^+$+4e$^-$, CoCl$_2$, BaS$_2$O$_3$		0.435	51Den
	Hg^{2+}+4S+6H$_2$O→ Hg(S$_2$O$_3$)$_2^{2-}$ +12H$^+$+8e$^-$, HgNO$_3$, Na$_2$S$_2$O$_3$		0.245	54Tor
	Hg^{2+}+6S+9H$_2$O→ Hg(S$_2$O$_3$)$_2^{4-}$ +18H$^+$+12e$^-$, HgNO$_3$, Na$_2$S$_2$O$_3$		0.306	54Tor
	Hg^{2+}+8S+12H$_2$O→ Hg(S$_2$O$_3$)$_4^{6-}$ +24H$^+$+16e$^-$, HgNO$_3$, Na$_2$S$_2$O$_3$		0.341	54Tor
	K$^+$+2S+3H$_2$O→ KS$_2$O$_3^-$ +6H$^+$+4e$^-$, KCl, BaS$_2$O$_3$		0.452	51Den
	Mg^{2+}+2S+3H$_2$O→MgS$_2$O$_3$+6H$^+$+4e$^-$, MgCl$_2$, BaS$_2$O$_3$		0.438	51Den
	Mg^{2+}+2S+3H$_2$O→MgS$_2$O$_3$+6H$^+$+4e$^-$, MgCl$_2$, Na$_2$S$_2$O$_3$		0.438	55Gim
	Mn^{2+}+2S+3H$_2$O→MnS$_2$O$_3$+6H$^+$+4e$^-$, MnCl$_2$, BaS$_2$O$_3$		0.438	51Den
	Na$^+$+2S+3H$_2$O→ NaS$_2$O$_3^-$ +6H$^+$+4e$^-$, MnCl$_2$, BaS$_2$O$_3$		0.455	51Den
	Ni^{2+}+2S+3H$_2$O→NiS$_2$O$_3$+6H$^+$+4e$^-$, NiCl$_2$, BaS$_2$O$_3$		0.434	51Den

Electrode	Electrode reaction, electrode system [2])	T [K]	E_{SHE} [V]	Ref.
S	$Sr^{2+}+2S+3H_2O \rightarrow SrS_2O_3+6H^++4e^-$, $NiCl_2$, BaS_2O_3		0.435	51Den
	$Zn^{2+}+2S+3H_2O \rightarrow ZnS_2O_3+6H^++4e^-$, $NiCl_2$, BaS_2O_3		0.430	51Den
	$Zn^{2+}+2S+3H_2O \rightarrow ZnS_2O_3+6H^++4e^-$, $NiCl_2$, BaS_2O_3		0.430	51Den
	$Zn^{2+}+2S+3H_2O \rightarrow ZnS_2O_3+6H^++4e^-$, $NiCl_2$, $Na_2S_2O_3$		0.430	51Goa1, 51Goa2
	$Ce^{3+}+S+4H_2O \rightarrow CeSO_4^+ +8H^++6e^-$, $Ce_2(SO_4)_3$		0.322	54Spe
	$Er^{3+}+S+4H_2O \rightarrow ErSO_4^+ +8H^++6e^-$, $Er_2(SO_4)_3$		0.322	54Spe
	$Gd^{3+}+S+4H_2O \rightarrow GdSO_4^+ +8H^++6e^-$, $Gd_2(SO_4)_3$		0.321	54Spe
	$Hg_2^{2+}+S+4H_2O \rightarrow Hg_2SO_4+8H^++6e^-$, Hg_2SO_4		0.297	57Bro
	$Ho^{3+}+S+4H_2O \rightarrow HoSO_4^+ +8H^++6e^-$, $Ho_2(SO_4)_3$		0.322	54Spe
	$K^++S+4H_2O \rightarrow KSO_4^- +8H^++6e^-$, K_2SO_4		0.348	50Jen
	$La^{3+}+S+4H_2O \rightarrow LaSO_4^+ +8H^++6e^-$, $La_2(SO_4)_3$		0.322	50Jen
	$La^{3+}+S+4H_2O \rightarrow LaSO_4^+ +8H^++6e^-$, $La_2(SO_4)_3$		0.322	47Jam
	$Mg^{2+}+S+4H_2O \rightarrow MgSO_4+8H^++6e^-$, $MgSO_4$		0.338	51Dun
	$Mg^{2+}+S+4H_2O \rightarrow MgSO_4+8H^++6e^-$, HCl, $MgSO_4$		0.338	52Jon
	$Mn^{2+}+S+4H_2O \rightarrow MnSO_4+8H^++6e^-$		0.346	47Jam
	$Na^++S+4H_2O \rightarrow NaSO_4^- +8H^++6e^-$, Na_2SO_4		0.350	57Dav
	$Nd^{3+}+S+4H_2O \rightarrow NdSO_4^+ +8H^++6e^-$, $Nd_2(SO_4)_3$		0.321	54Spe
	$Pr^{3+}+S+4H_2O \rightarrow PrSO_4^+ +8H^++6e^-$, $Pr_2(SO_4)_3$		0.322	54Spe
	$Sm^{3+}+S+4H_2O \rightarrow SmSO_4^+ +8H^++6e^-$, $Sm_2(SO_4)_3$		0.321	54Spe
	$UO_2^{2+}+S+4H_2O \rightarrow UO_2SO_4+8H^++6e^-$, UO_2SO_4, $HClO_4$		0.328	57Dav
	$Y^{3+}+S+4H_2O \rightarrow YSO_4^+ +8H^++6e^-$, $Y_2(SO_4)_3$		0.323	54Spe
	$Yb^{3+}+S+4H_2O \rightarrow YbSO_4^+ +8H^++6e^-$, $Yb_2(SO_4)_3$		0.322	54Spe
	$S^{2-} \rightarrow S+2e^-$, H_2S		–0.47627	57Mar
	$2S^{2-} \rightarrow S_2^{2-} +2e^-$, H_2S		–0.5242	57Mar
	$S_2^{2-} \rightarrow 2S+2e^-$, H_2S		–0.42836	57Mar
	$3S_2^{2-} \rightarrow 2S_3^{2-} +2e^-$, H_2S		–0.5056	57Mar
	$S_3^{2-} \rightarrow 3S+2e^-$, H_2S		–0.38973	57Mar
	$4S_3^{2-} \rightarrow 3S_4^{2-} +2e^-$, H_2S		–0.4778	57Mar
	$S_4^{2-} \rightarrow 4S+2e^-$, H_2S		–0.36039	57Mar
	$5S_4^{2-} \rightarrow 4S_5^{2-} +2e^-$, H_2S		–0.4407	57Mar
	$S_5^{2-} \rightarrow 5S+2e^-$, H_2S		–0.34030	57Mar
	$2S_2O_3^{2-} \rightarrow S_4O_6^{2-} +2e^-$, $K_2S_2O_3$, $K_2S_4O_6$		0.169	76Cob
	$2S_2O_3^{2-} \rightarrow S_4O_6^{2-} +2e^-$, acidic		0.17	72Moo
	$SO_3^{2-} +2OH^- \rightarrow SO_4^{2-} +H_2O+2e^-$, basic		–0.93	72Moo
	$2SO_4^{2-} \rightarrow S_2O_8^{2-} +2e^-$, acidic		2.01	72Moo

[1]) First electrodes based on inorganic materials (arranged according to Hill's system) are listed, followed by those based on organic materials. All data refer to room temperature if not stated otherwise.

[2]) Experimental details are given only if relevant for the stated value of E_{SHE}.

References

06Lew Lewis, G.N.: J. Am. Chem. Soc. **28** (1906) 158.
06Ner Nernst, W., Wartenberg, H.v.: Z. Phys. Chem. **56** (1906) 534.
10Brö Brönsted, J.N.: Z. Phys. Chem. **68** (1910) 703.
11Bra Bray, W.C., Connolly, E.L.: J. Am. Chem. Soc. **33** (1911) 1485.
11Lew Lewis, G.N., Rupert, F.F.: J. Am. Chem. Soc. **33** (1911) 299.
14Lew2 Lewis, G.N., Randall, M.: J. Am. Chem. Soc. **36** (1914) 2468.
17Lew1 Lewis, G.N., Storch, H.: J. Am. Chem. Soc. **39** (1917) 2544.
20Noy Noyes, A.A., Freed, E.S.: J. Am. Chem. Soc. **42** (1920) 476.
22Ger Gerke, R.H.: J. Am. Chem. Soc. **44** (1922) 1684.
22LaM La Mer, V.K., Baker, L.E.: J. Am. Chem. Soc. **44** (1922) 1954.
23Lew Lewis, G.N., Randall, M.: Thermodynamics and the Free Energy of Chemical Substances, New York: McGraw Hill, 1923.
26Kam Kameyama, N., Yamamoto, H., Skumpei, O.: J. Soc. Chem. Ind. Jpn. **29** (1926) 679.
26Lat Latimer, W.M.: J. Am. Chem. Soc. **48** (1926) 2868.
27Nie Nielson, R.F., Brown, D.J.: J. Am. Chem. Soc. **49** (1927) 2423.
28Ran Randall, M., Young, L.E.: J. Am. Chem. Soc. **50** (1928) 989.
29Abh Abhandlungen der Deutschen Bunsengesellschaft **10** (1929).
29Ish Ishikawa, F., Shibata, E.: Sci. Rep. Tohoku Univ. Ser. 1 **18** (1929) 109.
33Har2 Harned, H.S., Wright, D.D.: J. Am. Chem. Soc. **55** (1933) 4849.
33Ish1 Ishikawa, F., Kimura, G., Murooka, T.: Sci. Rep. Tohoku Imp. Univ. Ser. 1 **22** (1933) 131.
35Hai Haissinsky, M.: J. Chim. Phys. **32** (1935) 116.
35Ham1 Hamer, W.J.: J. Am. Chem. Soc. **57** (1935) 9.
35Kim Kimura, G.: Bull. Inst. Phys. Chem. Res. Tokyo **14** (1935) 94.
37Wil Willard, H.H., Ning Kang Tang: J. Am. Chem. Soc. **59** (1937) 1188.
47Jam James, J.C.: Thesis, London, 1947.
49Dav1 Davies, C.W., Monk, C.B.: J. Chem. Soc. (1949) 413.
49Dav2 Davies, C.W., Wyatt, P.A.H.: Trans. Faraday Soc. **45** (1949) 770.
50Jen Jenkins, I.L., Monk, C.B.: J. Am. Chem. Soc. **72** (1950) 2695.
51Den Denney, T.O., Monk, C.B.: Trans. Faraday Soc. **47** (1951) 992.
51Dun Dunsmore, H.S., James, J.C.: J. Chem. Soc. (1951) 2925.
51Goa1 Goates, J.R., Cole, A.G., Cray, E.L., Faux, N.D.: J. Am. Chem. Soc. **73** (1951) 707.
51Goa2 Goates, J.R., Cole, A.G., Cray, E.L.: J. Am. Chem. Soc. **73** (1951) 3596.
52Jon Jones, H.W., Monk, C.B.: Trans. Faraday Soc. **48** (1952) 929.
52Lat Latimer, W.M.: The Oxidation States of the Elements and their Potentials in Aqueous Solutions, Englewood Cliffs: Prentice-Hall Inc., 1952.
54Bat1 Bates, R.G.: Electrometric pH Determinations, New York: Wiley, 1954.
54Ker Kern, D.M.H.: J. Am. Chem. Soc. **76** (1954) 4208.
54Spe Spedding, F.H., Jaffe, S.: J. Am. Chem. Soc. **76** (1954) 882.
54Tor Toropova, V.F.: Zh. Obshch. Khim. **24** (1954) 423.
55Gim Gimblett, F.G.R., Monk, C.B.: Trans. Faraday Soc. **51** (1955) 793.
56Bow Bowersox, D.F., Butler, E.A., Swift, E.H.: Anal. Chem. **28** (1956) 221.
56Kiv Kivalo, P., Ringbom, A.: Suom. Kemistil. B **29** (1956) 109.
57Bro Brown, S.A., Land, J.E.: J. Am. Chem. Soc. **79** (1957) 3015.
57Cha1 Chateau, H., Hervier, B., Pouradier, J.: J. Phys. Chem. **61** (1957) 250.
57Cha2 Chateau, H., Hervier, B., Pouradier, J.: J. Chim. Phys. **54** (1957) 246.
57Dav Davies, E.W., Monk, C.B.: Trans. Faraday Soc. **53** (1957) 442.
57Mar Maronny, G., Valensi, G.: Proc. Int. Comm. Electrochem. Thermodynamics and Kinetics, 1957, p. 155.
59Fli Flis, I.E., Mishchenko, K.P., Troitskaya, N.V.: Zh. Fiz. Khim. **33** (1959) 1744.
59Tro Trotiskaya, N.V., Mishchenko, K.P., Flis, I.E.: Zh. Fiz. Khim. **33** (1959) 1614.

61Ive Ives, D.J.G., Janz, G.J.: Reference Electrodes, New York: Academic Press, 1961.
62Nai Naito, T.: Kogyo Kagaku Zasshi **65** (1962) 749.
66Hoa Hoare, J.P.: Electrochim. Acta **11** (1966) 203.
66Mus Mussine, T., Faita, G.: Ric. Sci. **36** (1966) 175.
66Osm Osman-Zade, Sh.D., Vagramyan, A.T.: Elektrokhimiya **2** (1966) 85.
66Sch Schmid, G., Neumann, U.: Ber. Bunsenges. Phys. Chem. **70** (1966) 1165.
67Sch1 Schmid, G., Neumann, U.: Z. Phys. Chem. (Frankfurt/Main) **54** (1967) 150.
68Mus Mussini, T.: Chim. Ind. (Milan) **50** (1968) 783.
69Cer Cerquetti, A., Longhi, P., Mussini, T., Natta, G.: J. Electroanal. Chem. **20** (1969) 411.
71Ale Alekseeva, N.I., Zymmer, Ya.D., Nikol'skii, V.A.: Elektrokhimiya **7** (1971) 873.
71Lav Lavrenova, L.G., Zegzhda, T.V., Shul'man, V.M.: Elektrokhimiya **7** (1971) 83.
72Moo Moore, W.J.: Physical Chemistry, London: Longman, 1972.
75Spi Spitz, R.D., Liefbafsky, H.A.: J. Electrochem. Soc. **122** (1975) 363.
76Cob Cobble, J.W., Stephens, H.P., McKinnon, I.R., Westrum jr., E.F.: Inorg. Chem. **11** (1976) 1669.

Table 1.3. Metal electrode potentials in nonaqueous [1]) electrolyte systems [2]).

Electrode	Electrode reaction, electrolyte solution [3])	T [K]	E_{SHE} [4]) [V]	Ref.
Aluminum (Al)	$Al \rightarrow Al^{3+}+3e^-$, methanol		–0.632	64Jak2
Arsenic (As)	$As+3Br^- \rightarrow AsBr_3+3e^-$, toluene$AsBr_3$,$AlBr_3$		0.590	49Gal
Barium (Ba)	$BaCaO_3+Pb \rightarrow Ba^{2+}+PbCO_3+2e^-$, methanol, $BaCl_2$		–2.943	62Jak2
Beryllium (Be)	$Be \rightarrow Be^{2+}+2e^-$, molten $BeCl_2$, T=773 K		–2.144	56Ham
Bismuth (Bi)	$Be \rightarrow Be^{2+}+2e^-$, methanol, LiCl		–0.513	64Jak2
Cerium (Ce)	$Ce^{3+} \rightarrow Ce^{4+}+e^-$, acetonitrile; activity: 1.3			64Rao
Cesium (Cs)	$Cs \rightarrow Cs^+ + e^-$, propylene carbonate, $CsClO_4$		–3.012	72LHe
Cadmium	$Cd+2Cl^- \rightarrow CdCl_2+2e^-$, liquid ammonia, NH_4Cl, 1 **M**	298	–0.362	58Fis
(Cd)	$Cd \rightarrow Cd^{2+}+2e^-$, liquid ammonia, $CdCl_2$, $Pb(NO_3)_2$, NH_4SCN	239	–0.189	66Ros
Cd_{Hg}	$Cd+SO_4^{2-} \rightarrow CdSO_4+2e^-$, hydrazine, H_2SO_4		–0.246	55Fur
Cd_{Hg}	$Cd \rightarrow Cd^{2+}+2e^-$, methanol, $CdCl_2$		–0.410	65Hef
Cd_{Hg}	$Cd \rightarrow Cd^{2+}+2e^-$, ethanol, $Cd(NO_3)_2$, NH_4NO_3		–0.340	54Tur1
Cd_{Hg}	$Cd+3Br^- \rightarrow CdBr_3^- +2e^-$, ethanol, $Cd(NO_3)_2$, NH_4Br		–0.779	56Tur
$Cd_{(Hg)}$	$Cd+2SCN^- \rightarrow Cd(SCN)_2+2e^-$, ethanol, $Cd(NO_3)_2$, NH_4Br		–0.740	56Tur
Cd_{Hg}	$Cd \rightarrow Cd^{2+}+2e^-$, formic acid, $CdCl_2$		–0.75	46Ple1
	$Cd \rightarrow Cd^{2+}$, formamide, $CdCl_2$, KCl		–0.414	68Bro3, 67DeR
	$Cd \rightarrow Cd^{2+}+2e^-$, formamide, $CdCl_2$		–0.412	71And
	$Cd \rightarrow Cd^{2+}$, anhydrous formamide, 1 **M**	298	–0.408	54Pav
	$Cd+2Cl^- \rightarrow CdCl_2+2e^-$, anhydrous formamide, 1 **M**	298	–0.617	54Pav
	$Cd+2Cl^- \rightarrow CdCl_2+2e^-$, KCl saturated in anhydrous formamide, 1 **M**	298	–0.608	54Pav
	$Cd \rightarrow Cd^{2+}+2e^-$, acetonitrile, CdI_2		–0.47	48Ple
	$Cd \rightarrow Cd^{2+}+2e^-$, molten $CdCl_2$	773	–1.403	56Ham
	$Cd \rightarrow Cd^{2+}+2e^-$, molten $CdCl_2$	1073	–1.193	61Ive
	$Cd \rightarrow Cd^{2+}+2e^-$, molten $CdCl_2$	1273	–1.002	56Ham
Calcium	$Ca \rightarrow Ca^{2+}+2e^-$, ammonia	239	–2.185	66Ros
(Ca)	$CaCaO_3+Pb \rightarrow Ca^{2+}+PbCO_3+2e^-$, methanol, $CaCl_2$		–2.929	62Jak2
$(Ca)_{Hg}$	$Ca \rightarrow Ca^{2+}+2e^-$, formic acid, Ca formate		–3.20	46Ple1
	$Ca \rightarrow Ca^{2+}+2e^-$, propylene carbonate		–2.65	72LHe
	$Ca \rightarrow Ca^{2+}+2e^-$, acetonitrile		–2.75	48Ple
	$Ca \rightarrow Ca^{2+}+2e^-$, molten $CaCl_2$	773	–3.534	56Ham
	$Ca \rightarrow Ca^{2+}+2e^-$, molten $CaCl_2$	1273	–3.208	56Ham

Electrode	Electrode reaction, electrolyte solution [3])	T [K]	E_{SHE} [4]) [V]	Ref.
Cesium (Cs)	$Cs \rightarrow Cs^+ + e^-$, ammonia, $CsNO_3$	228	–3.93	46Ple2, 46Ple3
	$Cs \rightarrow Cs^+ + e^-$, formic acid, 0.01 **M** CsHCOO		–3.44	46Ple1
	$Cs \rightarrow Cs^+ + e^-$, acetonitrile, 0.01 **M** CsCl		–3.16	48Ple
Chromium (Cr)	$Cr \rightarrow Cr^{2+} + 2e^-$, molten $CrCl_2$	773	–1.537	56Ham
	$Cr \rightarrow Cr^{2+} + 2e^-$, molten $CrCl_2$	1073	–1.352	61Ive
	$Cr \rightarrow Cr^{2+} + 2e^-$, molten $CrCl_2$	1273	–1.262	56Ham
	$Cr \rightarrow Cr^{3+} + 3e^-$, molten $CrCl_3$	773	–1.376	56Ham
	$Cr \rightarrow Cr^{3+} + 3e^-$, molten $CrCl_3$	1073	–1.113	61Ive
	$Cr \rightarrow Cr^{3+} + 3e^-$, molten $CrCl_3$	1273	–1.006	56Ham
Cobalt (Co)	$Co \rightarrow Co^{2+} + 2e^-$, methanol,$CoCl_2$		–0.233	61Jak2
	$Co \rightarrow Co^{2+} + 2e^-$, molten $CoCl_2$	773	–1.140	56Ham
	$Co \rightarrow Co^{2+} + 2e^-$, molten $CoCl_2$	1073	–0.977	61Ive
	$Co \rightarrow Co^{2+} + 2e^-$, molten $CoCl_2$	1273	–0.900	56Ham
Copper (Cu)	$Cu \rightarrow Cu^+ + e^-$, methanol, $Cu(ClO_4)_2$	298	0.64	70Des
(Cu)$_{Hg}$	$Cu \rightarrow Cu^{2+} + 2e^-$, methanol, $Cu(ClO_4)_2$	298	0.513	70Des
(Cu)$_{Hg}$	$Cu^+ \rightarrow Cu^{2+} + e^-$, methanol, $Cu(ClO_4)_2$	298	0.4	70Des
(Cu)$_{Hg}$	$Cu \rightarrow Cu^{2+}$, anhydrous formamide, 1 **M**	298	0.279	54Pav
(Cu)$_{Hg}$	$Cu \rightarrow Cu^+$, acetonitrile	298	–0.344	71Sen
	$Cu^+ \rightarrow Cu^{2+} + e^-$, acetonitrile, $Cu(ClO_4)_2$	298	0.133	71Sen
	$Cu \rightarrow Cu^+ + e^-$, molten CuCl	773	–1.024	56Ham
	$Cu \rightarrow Cu^+ + e^-$, molten CuCl	1073	–0.970	61Ive
	$Cu \rightarrow Cu^+ + e^-$, molten CuCl	1273	–0.943	56Ham
	$Cu \rightarrow Cu^{2+} + 2e^-$, molten $CuCl_2$	1273	–0.528	56Ham
Iron (Fe)	Ferrocene$\rightarrow$Ferrocinium$+e^-$, methanol		0.410	78Lag
	Ferrocene$\rightarrow$Ferrocinium$+e^-$, 1,2-dimethoxyethane		0.07	71Gue
	Ferrocene$\rightarrow$Ferrocinium$+e^-$, tetrahydrofuran, Bu_4NClO_4		0.12	71Gue
	Ferrocene$\rightarrow$Ferrocinium$+e^-$, formamide		0.539	78Lag
	Ferrocene$\rightarrow$Ferrocinium$+e^-$, acetonitrile, ferrocene, ferricinium picrate, $(C_2H_5)_4NClO_4$		0.19	65Kol
	Ferrocene$\rightarrow$Ferrocinium$+e^-$, pyridine, ferrocene, ferricinium picrate, $(C_2H_5)_4NClO_4$		0.628	68Muk
	Fe(1,10-phenanthroline)$^{2+}\rightarrow$ Fe(1,10-phenanthroline)$^{3+}+e^-$, acetonitrile, $NaClO_4$		0.805	65Kol
	Fe(1,10-phenanthroline)$^{2+}\rightarrow$ Fe(1,10-phenanthroline)$^{3+}+e^-$, acetonitrile, $(C_2H_5)_4NClO_4$		0.805	66Kra3
	$Fe \rightarrow Fe^{2+} + 2e^-$, molten $FeCl_2$	773	–1.267	56Ham
	$Fe \rightarrow Fe^{2+} + 2e^-$, molten $FeCl_2$	1073	–1.118	61Ive
	$Fe \rightarrow Fe^{2+} + 2e^-$, molten $FeCl_2$	1273	–1.050	56Ham
Lead (Pb)	$Pb \rightarrow Pb^{2+} + 2e^-$, ethanol, NH_4NO_3		–0.146	54Tur2
	$Pb \rightarrow Pb^{2+} + 2e^-$, formic acid, $Pb(HCOO)_2$		–0.72	46Ple1
	$Pb \rightarrow Pb^{2+} + 2e^-$, anhydrous formamide, 1 **M**	298	–0.193	54Pav
	$Pb \rightarrow Pb^{2+}$, acetamide, $Pb(NO_3)_2$		–0.193	58Fis

Electrode	Electrode reaction, electrolyte solution [3])	T [K]	E_{SHE} [4]) [V]	Ref.
Pb	$Pb \rightarrow Pb^{2+}$, acetonitrile, $Pb(ClO_4)_2$		–0.12	48Ple
	$Pb+2Cl^- \rightarrow PbCl_2+2e^-$, methanol, KCl		–0.2873	50Car
	$Pb+2Cl^- \rightarrow PbCl_2+2e^-$, *N*, *N*-dimethylformamide, HCl		–0.3368	74Bat2
	$Pb \rightarrow Pb^{2+}+2e^-$, molten $PbCl_2$	773	–1.2710	56Ham
	$Pb \rightarrow Pb^{2+}+2e^-$, molten $PbCl_2$	1073	–1.112	61Ive
	$Pb \rightarrow Pb^{2+}+2e^-$, molten $PbCl_2$	1273	–1.039	56Ham
Lithium (Li)	$Li \rightarrow Li^+ + e^-$, methanol/water 30 wt%		–3.020	83Maz
	$Li \rightarrow Li^+ + e^-$, methanol/water 50.3 wt%		–3.001	83Maz
	$Li \rightarrow Li^+ + e^-$, methanol/water 69.3 wt%		–2.994	83Maz
	$Li \rightarrow Li^+ + e^-$, methanol/water 90.0 wt%		–3.006	83Maz
	$Li \rightarrow Li^+ + e^-$, methanol		–3.095	83Maz
	$Li \rightarrow Li^+ + e^-$, methanol	298	–3.095	59Jak
	$Li \rightarrow Li^+ + e^-$, ethanol/water 10 wt%		–3.025	83Maz
	$Li \rightarrow Li^+ + e^-$, ethanol/water 20.1 wt%		–3.014	83Maz
	$Li \rightarrow Li^+ + e^-$, ethanol/water 30.4 wt%		–3.002	83Maz
	$Li \rightarrow Li^+ + e^-$, ethanol/water 40.3 wt%		–2.991	83Maz
	$Li \rightarrow Li^+ + e^-$, ethanol/water 49.9 wt%		–2.983	83Maz
	$Li \rightarrow Li^+ + e^-$, ethanol/water 60.3 wt%		–2.975	83Maz
	$Li \rightarrow Li^+ + e^-$, ethanol/water 70.3 wt%		–2.961	83Maz
	$Li \rightarrow Li^+ + e^-$, ethanol/water 80.4 wt%		–2.952	83Maz
	$Li \rightarrow Li^+ + e^-$, ethanol/water 90.4 wt%		–2.942	83Maz
	$Li \rightarrow Li^+ + e^-$, ethanol		–3.016	83Maz
	$Li \rightarrow Li^+ + e^-$, ethanol		–3.022	53Bje
$Li_{(Hg)}$	$Li \rightarrow Li^+ + e^-$, ethanol, LiBr	298	–3.097	72Kun
$Li_{(Hg)}$	$Li \rightarrow Li^+ + e^-$, formic acid	298	–3.48	46Ple1
	$Li \rightarrow Li^+ + e^-$, propylene carbonate	298	–2.8874	70Bod
	$Li \rightarrow Li^+ + e^-$, acetonitrile	298	–3.23	48Ple
	$Li \rightarrow Li^+ + e^-$, dimethylsulfoxide	298	–2.4212	67Smy
	$Li \rightarrow Li^+ + e^-$, molten LiCl	773	–3.646	56Ham
	$Li \rightarrow Li^+ + e^-$, molten LiCl	1273	–3.352	56Ham
Magnesium (Mg)	$Cd+Mg(OH)_2 \rightarrow Cd(OH)_2+Mg^{2+}+2e^-$, methanol, cadmium electrode		–2.478	61Jak1
	$Mg \rightarrow Mg^{2+}+2e^-$, molten $MgCl_2$	773	–2.680	56Ham
	$Mg \rightarrow Mg^{2+}+2e^-$, molten $MgCl_2$	1073	–2.460	61Ive
	$Mg \rightarrow Mg^{2+}+2e^-$, molten $MgCl_2$	1273	–2.346	56Ham
Manganese (Mn)	$Mn \rightarrow Mn^{2+}+2e^-$, molten $MnCl_2$	773	–1.967	56Ham
	$Mn \rightarrow Mn^{2+}+2e^-$, molten $MnCl_2$	1073	–1.807	61Ive
	$Mn \rightarrow Mn^{2+}+2e^-$, molten $MnCl_2$	1273	–1.725	56Ham
Mercury (Hg)	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, methanol(20.22 wt%)/water, 1 **m**	298	0.2545	58Sch
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, methanol(43.12 wt%)/water, 1 **m**	298	0.2515	58Sch
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, methanol(68.33 wt%)/water, 1 **m**	298	0.2173	58Sch
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, methanol(99.29 wt%)/water, 1 **m**	298	0.1027	58Sch
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, dioxane(20 wt%)/water, 1 **m**	298	0.2501	59Sch2
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, dioxane(40 wt%)/water, 1 **m**	298	0.2104	59Sch2

Electrode	Electrode reaction, electrolyte solution [3])	T [K]	E_{SHE} [4]) [V]	Ref.
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, dioxane(70 wt%)/water, 1 **m**	298	0.1126	59Sch2
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, dioxane(82 wt%)/water, 1 **m**	298	–0.0014	59Sch2
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, ethylene glycol(19.25 wt%)/water, 1 **m**	298	0.2570	59Sch3
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, ethylene glycol(50 wt%)/water, 1 **m**	298	0.2364	59Sch3
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, ethylene glycol(77.91 wt%)/water, 1 **m**	298	0.20125	59Sch3
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, anhydrous acetic acid, 1 **m**	298	0.181	31Hut
	$2Hg+SO_4^{2-} \rightarrow Hg_2SO_4+2e^-$, methanol, H_2SO_4		0.5392	46Kan
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, acetic acid, KCl		0.27	56Tut2
	$2Hg \rightarrow Hg_2^{2+}+2e^-$, formic acid, HgHCOO		0.180	46Ple1
	$2Hg+2Ac^- \rightarrow Hg_2Ac_2+2e^-$, acetic acid, Na acetate		–0.1599	65Jak2
	$2Hg \rightarrow Hg_2^{2+}+2e^-$, formamide, KCl		0.2451	67DeR
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, formamide, HCl		0.2453	71Nay
	$2Hg \rightarrow Hg_2^{2+}+2e^-$, acetonitrile, $HgBr_2$		0.25	48Ple
	$Hg_2^{2+} \rightarrow 2Hg^{2+}+2e^-$, acetonitrile, $Hg_2(ClO_4)_2$, $HClO_4$		0.614	71Nay
	$2Hg \rightarrow Hg_2^{2+}+2e^-$, acetonitrile, $Hg_2(ClO_4)_2$, $HClO_4$		0.460	73Coe
	$2Hg \rightarrow Hg_2^{2+}+2e^-$, *N*-methylpyrrolidone, $Hg_2(ClO_4)_2$, $HClO_4$		0.069	70Bré
	$Hg_2^{2+} \rightarrow 2Hg^{2+}+2e^-$, *N*-methylpyrrolidone, $HClO_4$		0.061	70Bré
	$Hg \rightarrow Hg^{2+}+2e^-$, *N*-methylpyrrolidone, $Hg_2(ClO_4)_2$, $HClO_4$		0.065	70Leu
	$2Hg+2\ benzoate^- \rightarrow Hg_2\ benzoate_2+2e^-$, formamide, benzoic acid		0.2535	75Das5
	$2Hg \rightarrow Hg_2^{2+}+2e^-$, molten Hg_2Cl_2	773	–0.597	56Ham
Nickel	$Ni \rightarrow Ni^{2+}+2e^-$, LiCl		0.088	64Jak2
(Ni)	$Ni \rightarrow Ni^{2+}+2e^-$, molten $NiCl_2$	773	–1.070	56Ham
	$Ni \rightarrow Ni^{2+}+2e^-$, molten $NiCl_2$	1073	–0.875	61Ive
	$Ni \rightarrow Ni^{2+}+2e^-$, molten $NiCl_2$	1273	–0.763	56Ham
Palladium	$Pd \rightarrow Pd^{2+}+2e^-$, molten $PdCl_2$	773	–0.457	56Ham
(Pd)	$Pd \rightarrow Pd^{2+}+2e^-$, molten $PdCl_2$	1073	–0.331	61Ive
Potassium	$K \rightarrow K^+ + e^-$, methanol/water 28.9 wt%		–2.888	83Maz
(K)	$K \rightarrow K^+ + e^-$, methanol/water 49.7 wt%		–2.861	83Maz
	$K \rightarrow K^+ + e^-$, methanol/water 70.1 wt%		–2.833	83Maz
	$K \rightarrow K^+ + e^-$, methanol/water 90.6 wt%		–2.836	83Maz
	$K \rightarrow K^+ + e^-$, methanol		–2.928	83Maz
$\mathbf{K}_{(Hg)}$	$K \rightarrow K^+ + e^-$, methanol, 0.01...0.1 **M** KCl		–2.921	58Brä
$\mathbf{K}_{(Hg)}$	$K \rightarrow K^+ + e^-$, ethanol, 0.6266...3.1390 **M** KCl		–2.843	68Dil
$\mathbf{K}_{(Hg)}$	$K \rightarrow K^+ + e^-$, formic acid, 0.01 **M** KCOOH		–3.36	46Ple1
	$K \rightarrow K^+ + e^-$, propylene carbonate, KI or $KClO_4$		–2.9504	70Bod
$\mathbf{K}_{(Hg)}$	$K \rightarrow K^+ + e^-$, ethylene glycol, KBr	298	–2.998	72Kun
	$K \rightarrow K^+ + e^-$, acetonitrile, 0.01 **M** KI		–2.91	48Ple
	$K \rightarrow K^+ + e^-$, anhydrous formamide, 1 **M**	291	–2.868	54Pav
	$K \rightarrow K^+ + e^-$, anhydrous formamide, 1 **M**	298	–2.872	54Pav
	$K \rightarrow K^+ + e^-$, ethane bromide, 0.19 **M** KCl, 0.24 **M** $AlBr_3$		–2.516	46Mez
	$K \rightarrow K^+ + e^-$, nitrobenzene, KCl		–2.804	46Mez

Electrode	Electrode reaction, electrolyte solution [3])	T [K]	E_{SHE} [4]) [V]	Ref.
K	$K \rightarrow K^+ + e^-$, molten KCl	773	–3.755	56Ham
	$K \rightarrow K^+ + e^-$, molten KCl	1273	–3.155	56Ham
Rubidium$_{(Hg)}$ (Rb)	$Rb \rightarrow Rb^+ + e^-$, formic acid, 0.01 **M** RbHCOO		–3.45	46Ple1
	$Rb \rightarrow Rb + e^-$, propylene carbonate, $RbClO_4$		–2.982	72LHe
	$Rb \rightarrow Rb^+ + e^-$, anhydrous formamide, 1 **M**	291	–2.851	54Pav
	$Rb \rightarrow Rb^+ + e^-$, anhydrous formamide, 1 **M**	298	–2.855	54Pav
	$Rb \rightarrow Rb^+ + e^-$, acetonitrile		–3.17	48Ple
Silver (Ag)	$Ag \rightarrow Ag^+ + e^-$, liquid ammonia, $AgNO_3$, NH_4NO_3	243	0.49	74Bro2
	$Ag + Br^- \rightarrow AgBr + e^-$, D_2O, DBr		0.06014	74Lie
	$Ag \rightarrow Ag^+ + e^-$, *N*,*N*-dimethylformamide, $AgNO_3$, KNO_3, $NaNO_3$		0.812	63Cha
	$Ag \rightarrow Ag^+ + e^-$, acetonitrile, $AgClO_4$	295	0.240	68Ben
	$Ag + Br^- \rightarrow AgBr + e^-$, methanol, 1 **m**	298	–0.1328	42Kan
	$Ag + Br^- \rightarrow AgBr + e^-$, methanol, 1 **M**	308	–0.0815	56Muk
	$Ag + Br^- \rightarrow AgBr + e^-$, methanol		–0.139	67Alf
	$Ag + Br^- \rightarrow AgBr + e^-$, ethanol		–0.1939	61Nun
	$Ag + Br^- \rightarrow AgBr + e^-$, ethylene glycol, HBr		–0.1007	70Kun1
	$Ag + Br^- \rightarrow AgBr + e^-$, propylene glycol, NaBr		–0.1633	70Kun2
	$Ag + Br^- \rightarrow AgBr + e^-$, formamide, HBr		0.0967	68Mar1
	$Ag + Br^- \rightarrow AgBr + e^-$, formamide, KBr, CH_3COOH, CH_3COOK		0.0965	75Das1
	$Ag + BrO_3^- \rightarrow AgBrO_3 + e^-$, formamide, $NaClO_4$, $AgBrO_3$		0.4997	75Das2
	$Ag + Cl^- \rightarrow AgCl + e^-$, methanol(10 wt%)/water, 1 **m**	298	0.2155	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, methanol(20 wt%)/water, 1 **m**	298	0.2094	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, methanol(40 wt%)/water, 1 **m**	298	0.1968	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, methanol(60 wt%)/water, 1 **m**	298	0.1818	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, methanol(80 wt%)/water, 1 **m**	298	0.1492	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, methanol(90 wt%)/water, 1 **m**	298	0.1135	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, methanol(94.2 wt%)/water, 1 **m**	298	0.0841	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, methanol, 1 **m**	298	–0.0099	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, methanol		–0.0103	74Kun
	$Ag + Cl^- \rightarrow Ag^+ + e^- + Cl^-$, methanol, HCl		–0.0221	53Mar
	$Ag + Cl^- \rightarrow AgCl + e^-$, ethanol(10 wt%)/water, 1 **m**	298	0.2144	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, ethanol(20 wt%)/water, 1 **m**	298	0.2073	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, ethanol(30 wt%)/water, 1 **m**	298	0.2003	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, ethanol(40 wt%)/water, 1 **m**	298	0.1945	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, ethanol(50 wt%)/water, 1 **m**	298	0.1859	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, ethanol(71.9 wt%)/water, 1 **m**	298	0.1554	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, ethanol(88.5 wt%)/water, 1 **m**	298	0.1053	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, ethanol(98 wt%)/water, 1 **m**	298	0.0215	61Ive
	$Ag + Cl^- \rightarrow AgCl + e^-$, ethanol, 1 **m**	298	–0.0813	61Ive
	$Ag + Cl^- \rightarrow Ag^+ + e^- + Cl^-$, ethanol, HCl		–0.0848	61Téz
	$Ag + Cl^- \rightarrow Ag^+ + e^- + Cl^-$, n-propanol, HCl		–0.1135	61Téz
	$Ag + Cl^- \rightarrow Ag^+ + e^- + Cl^-$, isopropanol, HCl		–0.1345	61Téz
	$Ag + Cl^- \rightarrow AgCl^-$, 1-propanol(10 wt%)/water, 1 **m**	298	0.2141	61Ive

Electrode	Electrode reaction, electrolyte solution [3])	T [K]	E_{SHE} [4]) [V]	Ref.
	$Ag+Cl^- \rightarrow AgCl+e^-$, 1-propanol(20 wt%)/water, 1 **m**	298	0.2066	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, 2-propanol(5 wt%)/water, 1 **m**	298	0.2180	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, 2-propanol(10 wt%)/water, 1 **m**	298	0.2138	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, 2-propanol(20 wt%)/water, 1 **m**	298	0.2066	61Ive
	$Ag+Cl^- \rightarrow Ag^+ + e^- + Cl^-$, *n*-butanol, HCl		–0.1430	60Ryz
	$Ag+Cl^- \rightarrow AgCl+e^-$, ethylene glycol(5 wt%)/water, 1 **m**	298	0.2190	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, ethylene glycol(10 wt%)/water, 1 **m**	298	0.2161	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, ethylene glycol(15 wt%)/water, 1 **m**	298	0.2133	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, ethylene glycol(20 wt%)/water, 1 **m**	298	0.2101	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, ethylene glycol(30 wt%)/water, 1 **m**	298	0.2036	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, ethylene glycol(40 wt%)/water, 1 **m**	298	0.1972	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, ethylene glycol(60 wt%)/water, 1 **m**	298	0.1807	61Ive
	$Ag+Cl^- \rightarrow Ag^+ + e^- + Cl^-$, ethylene glycol, HCl		0.0235	67Sen2
	$Ag+Cl^- \rightarrow AgCl+e^-$, ethylene glycol, HCl		0.0233	75Ale2
	$Ag+Cl^- \rightarrow AgCl+e^-$, propylene glycol(10 wt%)/water, 1 **m**	298	0.2150	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, propylene glycol(20 wt%)/water, 1 **m**	298	0.2077	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, propylene glycol		–0.0323	70Kun1
	$Ag+Cl^- \rightarrow AgCl+e^-$, 2,3-butylene glycol(10 wt%)/water, 1 **m**	298	0.2144	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, 2,3-butylene glycol(20 wt%)/water, 1 **m**	298	0.2067	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, glycerol(4.9 wt%)/water, 1 **m**	298	0.2196	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, glycerol(10 wt%)/water, 1 **m**	298	0.2165	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, glycerol(21.2 wt%)/water, 1 **m**	298	0.2084	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, glycerol(30 wt%)/water, 1 **m**	298	0.2022	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, glycerol(50 wt%)/water, 1 **m**	298	0.1940	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, glycerol, HCl		0.02077	72Roy
	$Ag+Cl^- \rightarrow AgCl+e^-$, acetone(5 wt%)/water, 1 **m**	298	0.2190	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, acetone(10 wt%)/water, 1 **m**	298	0.2156	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, acetone(20 wt%)/water, 1 **m**	298	0.2079	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, acetone(40 wt%)/water, 1 **m**	298	0.1859	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, acetone(50 wt%)/water, 1 **m**	298	0.158	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, acetone(90 wt%)/water, 1 **m**	298	–0.034	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, acetone, 1 **m**	298	–0.53	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, acetone, HCl		–0.54	54Eve
	$Ag+Cl^- \rightarrow AgCl+e^-$, methyl ethyl ketone(10 wt%)/water, 1 **m**	298	0.2153	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, methyl ethyl ketone(20 wt%)/water, 1 **m**	298	0.2078	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, dioxane(20 wt%)/water, 1 **m**	298	0.2031	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, dioxane(45 wt%)/water, 1 **m**	298	0.1635	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, dioxane(70 wt%)/water, 1 **m**	298	0.0639	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, dioxane(82 wt%)/water, 1 **m**	298	–0.0614	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, D-glucose(5 wt%)/water, 1 **m**	298	0.2186	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, D-glucose(10 wt%)/water, 1 **m**	298	0.2142	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, D-glucose(20 wt%)/water, 1 **m**	298	0.2045	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, D-glucose(30 wt%)/water, 1 **m**	298	0.1935	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, D-glucose(50 wt%)/water, 1 **m**	298	0.1634	61Ive

Electrode	Electrode reaction, electrolyte solution [3])	T [K]	E_{SHE} [4]) [V]	Ref.
Ag	$Ag+Cl^- \rightarrow AgCl+e^-$, D-fructose(5 wt%)/water, 1 **m**	298	0.2190	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, D-fructose(10 wt%)/water, 1 **m**	298	0.2150	61Ive
	$Ag+Cl^- \rightarrow AgCl+e^-$, HCl in formic acid, 1 **m**	298	–0.1302	57Muk
	$Ag+Cl^- \rightarrow AgCl+e^-$, HCl in formic acid, 1 **M**	298	–0.1199	57Muk
	$Ag+Cl^- \rightarrow AgCl+e^-$, anhydrous acetic acid, 1 **m**	298	–0.6241	57Muk
	$Ag+Cl^- \rightarrow AgCl+e^-$, anhydrous acetic acid, 1 **M**	298	–0.6180	57Muk
	$Ag+Cl^- \rightarrow AgCl+e^-$, acetic acid, $AgNO_3$,	295	0.87	56Tut1
	$Ag+Cl^- \rightarrow AgCl+e^-$, acetic acid, HCl, $LiClO_4$		–0.610	75Nev
	$Ag+Cl^- \rightarrow AgCl+e^-$, trifluoro acetic acid		0.31	72Pet3
	$Ag+Cl^- \rightarrow AgCl+e^-$, anhydrous formamide, 1 **m**	298	0.204	58Man
	$Ag+Cl^- \rightarrow AgCl+e^-$, formamide, KCl		0.1979	68Bro2, 67Aga
	$Ag+Cl^- \rightarrow AgCl+e^-$, *N*-methylformamide, HCl		0.208	70Ber
	$Ag+Cl^- \rightarrow AgCl+e^-$, *N*-methylacetamide, HCl		0.21187	65Daw
	$Ag+Cl^- \rightarrow AgCl+e^-$, *N,N*-dimethylacetamide, HCl		0.1500	68Scr
	$Ag+Cl^- \rightarrow AgCl+e^-$, ethylenediamine, AgCl		0.757	65Muk
	$Ag+Cl^- \rightarrow AgCl+e^-$, pyridine, AgCl		0.551	69Muk
	$Ag+I^- \rightarrow AgI+e^-$, methanol, 1 **M**	308	–0.2530	56Muk
	$Ag+I^- \rightarrow AgI+e^-$, methanol, NaI, succinic acid, Na succinate		–0.3176	53Ber1
	$Ag+I^- \rightarrow AgI+e^-$, ethylene glycol, NaI, Na cetate, acetic acid		–0.2928	70Kun2
	$Ag+I^- \rightarrow AgI+e^-$, propylene glycol, NaI, Na cetate, acetic acid		–0.3468	70Kun2
	$Ag+I^- \rightarrow AgI+e^-$, formamide, CH_3COOH, CH_3COOK, KI		–0.0996	75Das1
	$Ag+ MnO_4^- \rightarrow AgMnO_4+e^-$, formamide, $NaClO_4$, $AgMnO_4$		0.2055	75Das4
	$Ag+acetate^- \rightarrow Ag\ acetate+e^-$, formamide, acetic acid		0.4266	73Nay
	$Ag \rightarrow Ag^+ + e^-$, molten AgCl	773	–0.896	56Ham
	$Ag \rightarrow Ag^+ + e^-$, molten AgCl	1073	–0.826	61Ive
	$Ag \rightarrow Ag^+ + e^-$, molten AgCl	1273	–0.763	56Ham
Sodium (Na)	$Na \rightarrow Na^+ + e^-$, methanol/water 69.8 wt%		–2.627	83Maz
	$Na \rightarrow Na^+ + e^-$, methanol		–2.726	83Maz
	$Na \rightarrow Na^+ + e^-$, methanol		2.72	59Jak
	$Na \rightarrow Na^+ + e^-$, ethanol/water 10 wt%		–2.701	83Maz
	$Na \rightarrow Na^+ + e^-$, ethanol/water 20.4 wt%		–2.688	83Maz
	$Na \rightarrow Na^+ + e^-$, ethanol/water 30.8 wt%		–2.672	83Maz
	$Na \rightarrow Na^+ + e^-$, ethanol/water 40.6 wt%		–2.656	83Maz
	$Na \rightarrow Na^+ + e^-$, ethanol/water 50.5 wt%		–2.639	83Maz
	$Na \rightarrow Na^+ + e^-$, ethanol/water 60.6 wt%		–2.626	83Maz
	$Na \rightarrow Na^+ + e^-$, ethanol/water 69.9 wt%		–2.607	83Maz
	$Na \rightarrow Na^+ + e^-$, ethanol/water 80.0 wt%		–2.597	83Maz
	$Na \rightarrow Na^+ + e^-$, ethanol/water 90.5 wt%		–2.557	83Maz
	$Na \rightarrow Na^+ + e^-$, ethanol		–2.643	83Maz
$Na_{(Hg)}$	$Na \rightarrow Na^+ + e^-$, ethanol, NaBr,	298	–2.787	72Kun
$Na_{(Hg)}$	$Na \rightarrow Na^+ + e^-$, formic acid		–3.42	46Ple1
	$Na \rightarrow Na^+ + e^-$, propylene carbonate, $NaClO_4$		–2.653	72LHe
	$Na \rightarrow Na^+ + e^-$, acetonitrile, 0.01 **M** NaI		–2.87	46Mez

Electrode	Electrode reaction, electrolyte solution [3])	T [K]	E_{SHE} [4]) [V]	Ref.
	$Na \rightarrow Na^+ + e^-$, nitrobenzene, 0.19 **M** NaCl, 0.24 **M** $AlBr_3$		–3.42	46Ple1
	$Na \rightarrow Na^+ + e^-$, ethane bromide, 0.19 **M** NaCl, 0.24 **M** $AlBr_3$		–2.45	47Mez
	$Na \rightarrow Na^+ + e^-$, molten NaCl	773	–3.519	56Ham
	$Na \rightarrow Na^+ + e^-$, molten NaCl	1273	–3.019	56Ham
Strontium	$SrCO_3 + Pb \rightarrow Sr^{2+} + PbCO_3 + 2e^-$, methanol, $SrCl_2$		–2.938	62Jak2
(Sr)	$Sr \rightarrow Sr^{2+} + 2e^-$, molten $SrCl_2$	773	–3.684	56Ham
	$Sr \rightarrow Sr^{2+} + 2e^-$, molten $SrCl_2$	1273	–3.333	56Ham
Thallium	$Tl + Cl^- \rightarrow TlCl + e^-$, liquid ammonia, 1 **M**	298	0.1	58Fis
(Tl)	$Tl + Cl^- \rightarrow TlCl + e^-$, *n*-hexanol, HCl		–0.973	74Con2
	$Tl \rightarrow Tl^+ + e^-$, propylene carbonate		–0.3817	70Bod, 73Bod
	$Tl \rightarrow Tl^{2+} + 2e^-$, anhydrous formamide, 1 **M**	298	–0.344	54Pav
	$Tl + Cl^- \rightarrow TlCl + e^-$, *N*,*N*-dimethylformamide, HCl		–0.5994	74Bat2
$Tl_{(Hg)}$	$Tl \rightarrow Tl^+ + e^-$, acetonitrile, $TlCl_4$, Et_4NClO_4		–0.410	67Coe
	$Tl \rightarrow Tl^+ + e^-$, molten TlCl	773	–1.606	56Ham
	$Tl \rightarrow Tl^+ + e^-$, molten TlCl	1073	–1.473	61Ive
	$Tl \rightarrow Tl^+ + e^-$, molten TlCl	1273	–1.470	56Ham
Tin	$Sn \rightarrow Sn^{2+} + 2e^-$, methanol, $SnCl_2$		–0.532	64Jak2
(Sn)	$Sn \rightarrow Sn^{2+} + 2e^-$, molten $SnCl_2$	773	–1.320	56Ham
	$Sn \rightarrow Sn^{2+} + 2e^-$, molten $SnCl_2$	1073	–1.259	61Ive
Titanium	$Ti^{3+} \rightarrow Ti^{4+} + 2e^-$, DMF, HCl		0.000	69Hla
(Ti)	$Ti^{3+} \rightarrow Ti^{4+} + 2e^-$, pyridine, HCl		–0.190	69Hla
Vanadium (V)	$V \rightarrow V^{2+} + 2e^-$, molten VCl_2	1073	–1.566	61Ive
Zinc	$Zn + 2Cl^- \rightarrow ZnCl_2 + 2e^-$, liquid ammonia, NH_4Cl, 1 **M**	298	–0.72	58Fis
(Zn)	$Zn \rightarrow Zn^{2+} + 2e^-$, liquid ammonia, $ZnCl_2$, $Pb(NO_3)_2$, NH_4SCN	239	–0.505	66Ros
	$Zn + \rightarrow ZnSO_4 + 2e^-$, anhydrous hydrazine, 1 **M**	273	–0.7100	37Uli
	$Zn \rightarrow Zn^{2+} + 2e^-$, formic acid, $ZnCl_2$		–1.05	46Ple1
	$Zn \rightarrow Zn^{2+} + 2e^-$, acetonitrile, $ZnCl_2$		–0.74	48Ple
	$Zn \rightarrow Zn^{2+} + 2e^-$, anhydrous formamide, $ZnCl_2$, 1 **M**	298	–0.757	54Pav
	$Zn + 2Cl^- \rightarrow ZnCl_2 + 2e^-$, pyridine, HCl		–0.788	68Muk
	$Zn \rightarrow Zn^{2+} + 2e^-$, molten $ZnCl_2$	773	–1.603	56Ham
	$Zn \rightarrow Zn^{2+} + 2e^-$, molten $ZnCl_2$	1073	–1.476	61Ive

[1]) Mixed solvent systems containing water are included below even in case water is only a minor constituent, molten salts are listed after liquid solvent-based systems.

[2]) All data refer to room temperature if not stated otherwise.

[3]) Details are given only if relevant for the stated value of E_{SHE}.

[4]) All potentials refer to a standard hydrogen electrode in the used solvent. Exemptions are stated explicitly.

References

31Hut Hutchinson, A.W., Chandlee, G.C.: J. Am. Chem. Soc. **53** (1931) 2881.
37Uli Ulich, H., Biostoch, K.: Z. Phys. Chem. Abt. A **178** (1937) 306.
42Kan Kanning, E.W., Campbell, A.W.: J. Am. Chem. Soc. **64** (1942) 1672.
46Kan Kanning, E.W., Bowman, M.G.: J. Am. Chem. Soc. **68** (1946) 2042.
46Mez Mezhennyi, Ya.F.: J. Phys. Chem. USSR **20** (1946) 493.
46Ple1 Pleskov, V.A.: Acta Physicochim. URSS **21** (1946) 41; J. Phys. Chem. USSR **20** (1946) 153.
46Ple2 Pleskov, V.A.: Acta Physicochim. URSS **21** (1946) 235.
46Ple3 Pleskov, V.A.: J. Phys. Chem. USSR **20** (1946) 151.
47Mez Mezhennyi, Ya.F.: J. Phys. Chem. USSR **21** (1947) 839.
48Ple Pleskov, V.A.: Acta Physicochim. URSS **22** (1948) 351.
49Gal Galinker, V.S.: Zh. Obshch. Khim. **19** (1949) 2048.
50Car Carlson, C.L., Malmberg, E.W., Brovon, D.I.: J. Phys. Chem. **54** (1950) 322.
53Ber1 Berne, B., Leden, I.: Z. Naturforsch. A **8** (1953) 719.
53Bje Bjerrum, J., Jörgensen, C.K.: Acta Chem. Scand. **7** (1953) 951
53Mar Marshall, H.P., Grünwald, E.: J. Chem. Phys. **21** (1953) 2143.
54Eve Everett, D.M., Rasmussen, S.E.: J. Chem. Soc. (1954) 2812.
54Pav Pavlopoulos, T., Strehlow, H.: Z. Phys. Chem. (Frankfurt/Main) **2** (1954) 89.
54Tur1 Turyan, Ya.I.: Zh. Fiz. Khim. **28** (1954) 2152.
54Tur2 Turyan, Ya.I.: Zh. Fiz. Khim. **28** (1954) 2129.
55Fur Furlani, C.: Ann. Chim. (Paris) **45** (1955) 264.
56Ham Hamer, W.H., Malmberg, M.S., Rubin, B.: J. Electrochem. Soc. **103** (1956) 8.
56Muk Mukherjee, L.M.: J. Phys. Chem. **60** (1956) 974.
56Tur Turyan, Ya.I.: Zh. Anal. Khim. **11** (1956) 71.
56Tut1 Tutundžik, P.S., Putanov, P.: Glas. Hem. Drust. (Beograd) **21** (1956) 19.
56Tut2 Tutundžik, P.S., Putanov, P.: Glas. Hem. Drust. (Beograd) **21** (1956) 257.
57Muk Mukherjee, L.M.: J. Am. Chem. Soc. **79** (1957) 4040.
58Brä Bräuer, K., Strehlow, H.: Z. Phys. Chem. N. F. **17** (1958) 346.
58Fis Fischer, L., Winkler, G., Jander, G.: Z. Elektrochem. **62** (1958) 1.
58Man Mandel, M., Decroly, P.: Nature (London) **182** (1958) 794.
58Sch Schwabe, K., Ziegenbalg, S.: Z. Elektrochem. **62** (1958) 172.
59Jak Jakuszewski, B., Taniewska-Osinka, S.: Acta Chim. Soc. Sci. Lodz. **4** (1959) 17.
59Sch2 Schwabe, K., Schwenk, W.: Z. Elektrochem. **63** (1959) 441.
59Sch3 Schwabe, K., Schwenk, W.: Z. Elektrochem. **63** (1959) 445.
60Ryz Ryzhov, E.M., Sukhotin, A.N.: Zh. Fiz. Khim. **34** (1960) 1402.
61Ive Ives, D.J.G., Janz, G.J.: Reference Electrodes, New York: Academic Press, 1961.
61Jak1 Jakuszewski, B.J., Taneewska-Osinska, S.: Zesz. Nauk Uniw. Lodz. Ser. **2** (1961) 10.
61Jak2 Jakuszewski, B., Taniewska-Osinska, S.: Acta Chim. Soc. Sci. Lodz. **7** (1961) 37.
61Nun Nunez, L.J., Day, M.C.: J. Phys. Chem. **65** (1961) 164.
61Téz Tézé, A., Schaal, R.: C. R. Acad. Sci. (Paris) **252** (1961) 3995.
62Jak2 Jakuszewski, B., Taniewska-Osinka, S.: Rocz. Chem. **36** (1962) 329.
63Cha Chateau, M., Moncet, M.C.: C. R. Acad. Sci. (Paris) **256** (1963) 1504.
64Jak2 Jakuszewski, B., Kozlowski, Z.: Rocz. Chem. **38** (1964) 93.
64Rao Rao, J.P., Murthy, A.R.V.: J. Phys. Chem. **68** (1964) 1573.
65Daw Dawson, R.L., Sheridan, R.C., Eckstrom, H.C.: J. Phys. Chem. **69** (1965) 1335.
65Hef Hefley, J.D., Amis, E.S.: J. Electrochem. Soc. **112** (1965) 336.
65Jak2 Jakuszewski, B., Jedraejwska, J.B.: Rocz. Chem. **39** (1965) 907.
65Kol Kolthoff, I.M., Thomas, F.G.: J. Phys. Chem. **69** (1965) 3049.
65Muk Mukherjee, L.M., Bruckenstein, S., Badawai, F.A.K.: J. Phys. Chem. **69** (1965) 2544.
66Kra3 Kratochvil, B., Krock, J.: J. Phys. Chem. **70** (1966) 944.
66Ros Rosztoczy, F.E., Tobias, C.W.: Electrochim. Acta **11** (1966) 857.

67Aga Agarwal, R.K., Nayak, R.: J. Phys. Chem. **71** (1967) 2062.

67Alf Alfenaar, M., De Ligny, C.L., Remijnse, A.G.: Recl. Trav. Chim. Pays Bas **86** (1967) 555.

67Coe Coetzee, T.F., Campion, J.J.: J. Am. Chem. Soc. **89** (1967) 2513.

67DeR De Rossi, M., Pecci, G., Scrosati, B.: Ric. Sci. **37** (1967) 342.

67Sen2 Sen, B., Kundu, K., Das, M.N.: J. Phys. Chem. **71** (1967) 3665.

67Smy Smyrl, W.H.: Diss. Abstr. B **28** (1967) 637.

68Ben Benoit, R.L.: Inorg. Nucl. Chem. Lett. **4** (1968) 723.

68Bro2 Broadbank, R.W.C., Dhabanandana, S., Morcom, K.W., Muju, B.L.: Trans. Faraday Soc. **64** (1968) 3311.

68Bro3 Broadbank, R.W.C., Muju, B.L., Morcom, K.W.: Trans. Faraday Soc. **64** (1968) 3318.

68Dil Dill, A.J., Itzkovits, V.: J. Phys. Chem. **72** (1968) 4580

68Mar1 Marcom, K.W., Muju, B.L.: Nature (London) **217** (1968) 1046.

68Muk Mukerjee, L.M., Kelly, J.J., Baranetzky, W., Sica, J.: J. Phys. Chem. **72** (1968) 3410.

68Scr Scrosati, B., Pecci, G., Pistoia, G.: J. Electrochem. Soc. **115** (1968) 506.

69Hla Hladky, Z., Vrestal, J.: Collect. Czech. Chem. Commun. **34** (1969) 984.

69Muk Mukherjee, L.M., Kelly, J.J., Richards, M., Lukacs jr., J.M.: J. Phys. Chem. **73** (1969) 580.

70Ber Berardelli, M.J., Pecci, G., Scrosati, B.: J. Electrochem. Soc. **117** (1970) 781.

70Bod Boden, D.P.: Diss. Abstr. B **31** (1970) 3223.

70Bré Bréant, M., Buisson, C.: J. Electroanal. Chem. **24** (1970) 145.

70Des Desmarquet, J.P., Trinh-Dinh, C., Bolck, O.: J. Electroanal. Chem. **27** (1970) 101.

70Kun1 Kundu, K.K., Chattopadhyay, P.K., Jana, D., Das, M.N.: J. Chem. Eng. Data **15** (1970) 209.

70Kun2 Kundu, K.K., Chattopadhyay, P.K., Jana, D., Das, M.N.: J. Phys. Chem. **74** (1970) 2625.

70Leu Leuscke, W., Schwabe, K.: J. Electroanal. Chem. **25** (1970) 219.

71And Andrews, A.W., Armitage, D.A., Broadbank, R.W.C., Morcom, K.W.: Trans. Faraday Soc. **67** (1971) 128.

71Gue Guérin, G.D., Caillet, A.: C. R. Acad. Sci. (Paris), Ser. C **273** (1971) 235.

71Nay Nayak, B., Saku, D.K.: Electrochim. Acta **16** (1971) 1757.

71Sen Senne, J.K., Kratochvil, B.: Anal. Chem. **43** (1971) 79.

72Kun Kundu, K.K., Rakshit, A.K., Das, M.N.: Electrochim. Acta **17** (1972) 1921.

72LHe L'Her, M., Courtot-Coupez, J.: Bull. Soc. Chim. Fr. (1972) 3645.

72Pet3 Petit, G., Besssiere, J.: J. Electroanal. Chem. **34** (1972) 489.

72Roy Roy, R.N., Vernon, W., Gibbons, J.J., Bothwell, A.L.M.: J. Electroanal. Chem. **34** (1972) 101.

73Bod Boden, D.P., Mukherjee, L.M.: Electrochim. Acta **18** (1973) 781.

73Coe Coetzee, J.F., Campion, J.J., Liberman, D.R.: Anal. Chem. **45** (1973) 343.

73Nay Nayak, B., Dash, U.N.: J. Electroanal. Chem. **41** (1973) 323.

74Bat2 Bates, F.L., Nee, Y.T.: J. Electrochem. Soc. **121** (1974) 79.

74Bro2 Brown, O.R., Thornton, S.A.: J. Chem. Soc. Faraday Trans. 1 **70** (1974) 1009.

74Con2 Contreras-Ortega, C.H., Rock, P.A.: J. Electrochem. Soc. **121** (1974) 1048.

74Kun Kundu, K.K., Jana, D., Das, M.N.: J. Chem. Eng. Data **19** (1974) 329.

74Lie Lietzke, M.H., Lemmonds, T.J.: J. Inorg. Nucl. Chem. **36** (1974) 2299.

75Ale2 Aleksandrov, V.V., Bezpalyi, B.N.: Zh. Fiz. Khim. **49** (1975) 1314.

75Das1 Dash, U.N., Nayak, B.: Aust. J. Chem. **28** (1975) 1657.

75Das2 Dash, U.N., Nayak, B.: Thermochim. Acta **11** (1975) 17.

75Das4 Dash, U.N.: Thermochim. Acta **12** (1975) 197.

75Das5 Dash, U.N.: Aust. J. Chem. **28** (1975) 1653.

75Nev Nevskaya, Yu.A., Korotkova, P.I., Sumarova, T.N.: Elektrokhimiya **11** (1975) 1660.

78Lag Lagowski, J.J.: The Chemistry of Non-aqueous Solutions, Vol. 1, London: Academic Press, 1978.

83Maz Mazzarese, J., Popvych, O.: J. Electrochem. Soc. **130** (1983) 2032.

Table 1.4. Nonmetal electrode potentials in nonaqueous [1]) electrolyte systems.

Electrode	Electrode reaction, electrolyte solution [2])	T [K] [3])	E_{SHE} [V] [4])	Ref.
Boron (B)	$B_{10}H_{14}^{2-} \rightarrow B_{10}H_{14}+2e^-$, acetonitrile, Et_4NClO_4, Bu_4NClO_4		–0.539 [5])	68Cha
Bromine	$2Br^- \rightarrow Br_2+2e^-$		0.8617	59Jak
(Br)	$3Br^- \rightarrow Br_3^- +2e^-$, acetonitrile, $(C_2H_5)_4NClO_4$	293	0.22	68Mar2
	$2Br_3^- \rightarrow 3Br_2+2e^-$, acetonitrile, $(C_2H_5)_4NClO_4$	293	0.83	68Mar2
	$2Br_3^- \rightarrow 3Br_2+2e^-$, nitromethane, $(C_2H_5)_4NClO_4$	293	1.24	68Mar2
	$3Br^- \rightarrow Br_3^- +2e^-$, nitromethane, $(C_2H_5)_4NClO_4$	293	0.63	68Mar2
Chlorine	$2Cl^- \rightarrow Cl_2+2e^-$, methanol		1.1272	59Jak
(Cl)	$2Cl^- \rightarrow Cl_2+2e^-$, acetonitrile		0.60	68Mar2
Fluorine (F)	$HF \rightarrow F+H^+ +e^-$, HF+1.5 wt%$H_2O$		2.85	68Wat
Iodine	$2I^- \rightarrow I_2+2e^-$, methanol		0.357	59Jak
(I)	$3I^- \rightarrow I_3^- +2e^-$, acetonitrile, $LiClO_4$		–0.124	61Des
	$3I^- \rightarrow I_3^- +2e^-$, nitromethane, $(C_2H_5)_4NClO_4$	293	0.28	68Mar2
Oxygen (O) [6])				
Tetra-chlor-benzo-quinone	tetrachlorbenzoquinone+$2e^-$+$2H^+ \rightarrow$tetrachlorohydroquinone, 1 m, anhydrous acetic acid		0.680	33Hes

[1]) Mixed solvent systems containing water are included below even in case water is only a minor constituent.

[2]) Details are given only if relevant for the stated value of E_{SHE}.

[3]) All data refer to standard (room) temperature T = 298 K if not stated otherwise.

[4]) All potentials refer to a standard hydrogen electrode in the used solvent. Exemptions are stated explicitly.

[5]) Potential reported versus a saturated calomel electrode, converted assuming E_{SCE}= –0.241 V vs. SHE.

[6]) An extensive review on oxygen reference electrodes with solid electroytes is available [00Mal].

References

33Hes Heston, B.O., Hall, N.F.: J. Am. Chem. Soc. **55** (1933) 4729; **56** (1934) 1462.

59Jak Jakuszewski, B., Taniewska-Osinka, S.: Acta Chim. Soc. Sci. Lodz. **4** (1959) 17.

61Des Desbarres, J.: Bull. Soc. Chim. Fr. (1961) 502.

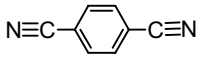
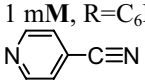
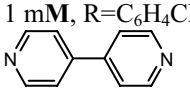
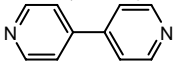
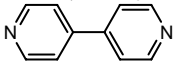
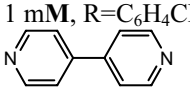
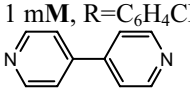
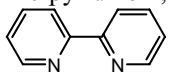
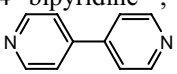
68Cha Chambers, J.Q., Norman, A.D., Bickell, M.R., Cadle, S.H.: J. Am. Chem. Soc. **90** (1968) 6056.

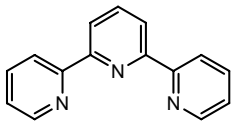
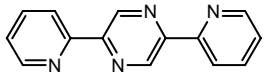
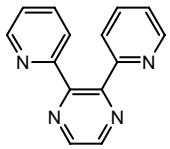
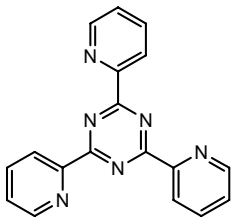
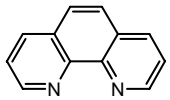
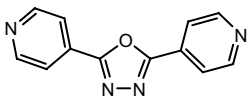
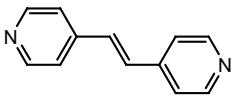
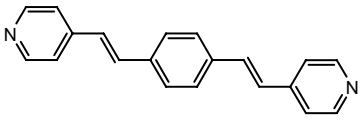
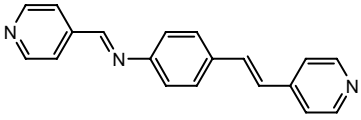
68Mar2 Marchon, J.C.: C. R. Acad. Sci. (Paris), Ser. C **267** (1968) 1123.

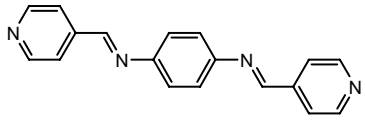
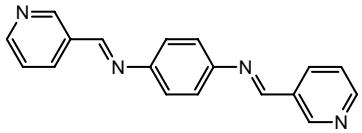
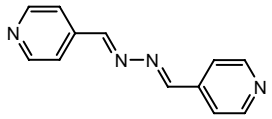
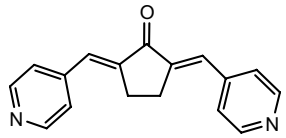
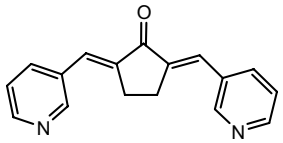
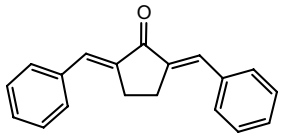
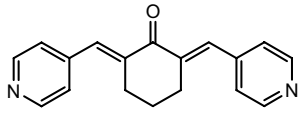
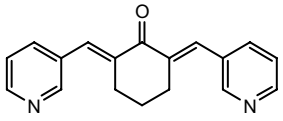
68Wat Watanabe, N., Chang, T., Nakanishi, K.: Denki Kagaku **36** (1968) 600.

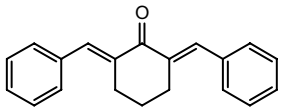
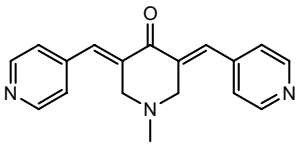
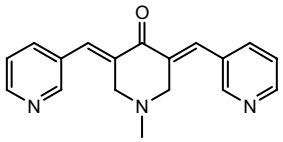
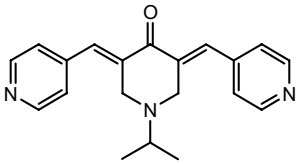
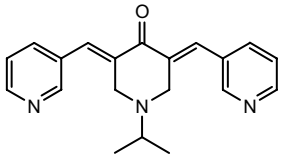
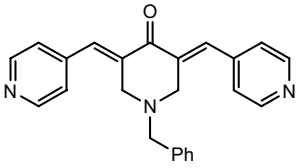
00Mal Mallika, C., Sreedharan, O.M., Subasri, R.: J. Eur. Ceram. Soc. **20** (2000) 2297.

Table 1.5. Formal potentials of organic redox systems.

Electrode	Electrode reaction [1]), electrolyte system [2])	RE	E_0 [V]	Ref.
Platinum (Pt)	C_6H_4-1,4-$(CN)_2^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 1 m**M** 	FeC	–2.16	03Les
	[(R_3P)Ag-4-NC_5H_4-CN-Ag(PR_3)]$^{2+/1+}$, 0.1 **M** [*n*-Bu_4N][PF_6], 1 m**M**, R=$C_6H_4CH_2NMe_2$-2, 4-NC_5H_4-$CN^{0/-}$ 	FeC	–2.46	79Rus
	[(R_3P)Ag-bipy-Ag(PR_3)]$^{2+/1+}$, 0.1 **M** [*n*-Bu_4N][PF_6], 1 m**M**, R=$C_6H_4CH_2NMe_2$-2, bipy$^{0/-}$ 	FeC	–2.57	03Les
	[(R_3P)Cu-bipy-Cu(PR_3)]$^{2+/1+}$, 0.1 **M** [*n*-Bu_4N][PF_6], 1 m**M**, R=$C_6H_4CH_2NMe_2$-2, Cu(I)/Cu(II) 	FeC	–0.44	03Les
	[(R_3P)Cu-bipy-Cu(PR_3)]$^{2+/1+}$, 0.1 **M** [*n*-Bu_4N][PF_6], 1 m**M**, R=$C_6H_4CH_2NMe_2$-2, bipy$^{0/-}$ 	FeC	–2.57	03Les
	[(R_3P)Ag-bipy-Cu(PR_3)]$^{2+/1+}$, 0.1 **M** [*n*-Bu_4N][PF_6], 1 m**M**, R=$C_6H_4CH_2NMe_2$-2, Cu(I)/Cu(II) 	FeC	–0.41	03Les
	[(R_3P)Ag-bipy-Cu(PR_3)]$^{2+/1+}$, 0.1 **M** [*n*-Bu_4N][PF_6], 1 m**M**, R=$C_6H_4CH_2NMe_2$-2, bipy$^{0/-}$ 	FeC	–2.57	03Les
	[(R_3P)Ag-4-NC_5H_4-CN-Ag(PR_3)]$^{2+/1+}$, 0.1 **M** [*n*-Bu_4N][PF_6], 1 m**M**, R=$C_6H_4CH_2NMe_2$-2, Cu(I)/Cu(II)	FeC	0.07	79Rus
	[(R_3P)Ag-4-NC_5H_4-CN-Ag(PR_3)]$^{2+/1+}$, 0.1 **M** [*n*-Bu_4N][PF_6], 1 m**M**, R=$C_6H_4CH_2NMe_2$-2, Cu(I)/Cu(II)	FeC	–0.38	03Les
	[(R_3P)Ag-4-NC_5H_4-CN-Ag(PR_3)]$^{2+/1+}$, 0.1 **M** [*n*-Bu_4N][PF_6], 1 m**M**, R=$C_6H_4CH_2NMe_2$-2, 4-NC_5H_4-$CN^{0/-}$	FeC	–2.46	79Rus
	[(R_3P)Cu-NC-C_6H_4-CN-Cu(PR_3)]$^{2+/1+}$, 0.1 **M** [*n*-Bu_4N][PF_6], 1 m**M**, R=$C_6H_4CH_2NMe_2$-2, Cu(I)/Cu(II)	FeC	–0.43	03Les
	[(R_3P)Cu-NC-C_6H_4-CN-Cu(PR_3)]$^{2+/1+}$, 0.1 **M** [*n*-Bu_4N][PF_6], 1 m**M**, R=$C_6H_4CH_2NMe_2$-2, C_6H_4-1,4-$(CN)_2^{0/-}$	FeC	–2.16	03Les
	2,2'-bipyridine$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–2.76	05Vat
	4,4'-bipyridine$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–2.51	05Vat

Electrode	Electrode reaction [1]), electrolyte system [2])	RE	E_0 [V]	Ref.
Pt	[2,2';6,6']terpyridine$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M**	FeC	–2.71	05Vat
				
	2,5-dipyridin-2-yl-pyrazine$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M**	FeC	–2.15	05Vat
				
	2,3-dipyridin-2-yl-pyrazine$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M**	FeC	–2.51	05Vat
				
	1,3,5-tris(2-pyridyl)triazine$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M**	FeC	–2.03	05Vat
				
	[1,10]phenanthroline$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M**	FeC	–2.66	05Vat
				
	2,5-bis(4-pyridinyl)[1,2,4]oxadiazol$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M**	FeC	–2.17	05Vat
				
	1,2-bis(4-pyridinyl)ethylene$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M**	FeC	–2.27	05Vat
				
	1,4-bis(2-(4-pyridinyl)ethenyl)benzene$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M**	FeC	–2.21	05Vat
				
	4-[2-(4-pyridinyl)ethenyl]-*N*-(4-pyridinylmethylene)-*N*-benzylamine$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M**	FeC	–2.06	05Vat
				

Electrode	Electrode reaction [1]), electrolyte system [2])	RE	E_0 [V]	Ref.
	N,*N*'-bis(4-pyridinylmethylene)1,4-benzenediamine$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–2.00	05Vat
	N,*N*'-bis(3-pyridinylmethylene)1,4-benzenediamine$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–2.18	05Vat
	4-pyridinaldazine$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–1.92	05Vat
	2,5-bis(4-pyridinylmethylene)cyclopentaone$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–1.61	05Vat
	2,5-bis(3-pyridinylmethylene)cyclopentanone$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–1.78	05Vat
	2,5-bis(phenylmethylene)cyclopentanone$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–1.92	05Vat
	2,6-bis(4-pyridinylmethylene)cyclohexaone$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–1.75	05Vat
	2,6-bis(3-pyridinylmethylene)cyclohexanone$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–1.89	05Vat

Electrode	Electrode reaction [1]), electrolyte system [2])	RE	E_0 [V]	Ref.
Pt	2,6-bis(phenylmethylene)cyclohexanone$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–2.07	05Vat
	1-methyl-3,5-bis(4-pyridinylmethylene)piperidin-4-one$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–1.71	05Vat
	1-methyl-3,5-bis(3-pyridinylmethylene)piperidin-4-one$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–1.83	05Vat
	1-isopropyl-3,5-bis(4-pyridinylmethylene)piperdin-4-one$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–1.85	05Vat
	1-isopropyl-3,5-bis(3-pyridinylmethylene)piperdin-4-one$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–1.93	05Vat
	1-benzyl-3,5-bis(4-pyridinylmethylene)piperidin-4-one$^{0/-}$, THF, 0.1 **M** [*n*-Bu_4N][PF_6], 5 m**M** 	FeC	–1.71	05Vat

[1]) In order to avoid cumbersome notation and excessive length of entries caused by repetition of long formulas assigned to the oxidized and the reduced state of the species under investigation the name of the potential-determining-compound is given once only. The superscript denotes the involved states of oxidation. The initial state is given first, the final state second as encountered in the experimental method applied.

[2]) Details are given only if relevant for the stated value of E_{SHE}.

References

79Rus Rusina, A., Vlcek, A.A., Zalis, S.: Z. Chem. **19** (1979) 27.

03Les Leschke, M., Lang, H., Holze, R.: J. Solid State Electrochem. **7** (2003) 518.

05Vat Vatsadze, S., Al-Anber, M., Thiel, W.R., Lang, H., Holze, R.: J. Solid State Electrochem. **9** (2005) 764.

Table 1.6. Potentials of common reference electrodes in aqueous solution.

Electrode	Electrode reaction, electrolyte solution [1])	T [K, °C]	E_{SHE} [2]) [V]	Ref.
Copper (Cu)	$Cu \rightarrow Cu^{2+}+2e^-$, sol. sat. $CuSO_4$		0.320	61Ive
Lead (Pb)	$Pb+SO_4^{2-} \rightarrow PbSO_4+2e^-$, sol. sat. K_2SO_4		–0.276	57Kor
$Pb_{(Hg)}$	$Pb+SO_4^{2-} \rightarrow PbSO_4+Hg+2e^-$		–0.351	35Har, 34Shr2
Mercury (Hg)	$2Hg+2I^- \rightarrow Hg_2I_2+2e^-$		–0.0405	57Kor
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. sat. Hg_2Cl_2 and in KCl		0.2412	61Ive,
	E [$V_{int.}$]=0.2412–6.61·10^{-5}(T–25°C)–1.75·10^{-6}(T–25°C)^2 –9·10^{-10}(T–25°C)			54Cha
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. sat. Hg_2Cl_2 and in KCl	0 °C	0.2568 *)	54Cha
		10 °C	0.2507 *)	54Cha
		20 °C	0.2444 *)	54Cha
		25 °C	0.2412 *)	54Cha
		30 °C	0.2378 *)	54Cha
		40 °C	0.2307 *)	54Cha
		50 °C	0.2233 *)	54Cha
		60 °C	0.2154 *)	54Cha
		70 °C	0.2071 *)	54Cha
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. sat. Hg_2Cl_2 and in NaCl		0.2360	01Bar
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. sat. Hg_2Cl_2, sol. 0.1 **M** in KCl		0.3337	01Bar
	E [$V_{int.}$]=0.3337–8.75·10^{-5}(T–25°C)–3·10^{-6}(T–25°C)^2			54Cha
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. sat. Hg_2Cl_2, sol. 0.1 **m** in KCl		0.3339	54Cha
	E [$V_{int.}$]=0.3339–7.75·10^{-5}(T–25°C)–3·10^{-6}(T–25°C)^2			54Cha
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. sat. Hg_2Cl_2, sol. 1 **M** in KCl		0.2801	61Ive,
	E [$V_{int.}$]=0.2801–2.75·10^{-5}(T–25 °C)–2.5·10^{-6}(T–25 °C)^2 –4·10^{-9}(T–25 °C)			54Cha
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$, sol. sat. Hg_2Cl_2, sol. 1 **m** in KCl		0.2809	54Cha
	E [$V_{int.}$]=0.2809–2.65·10^{-5}(T–25 °C)–2.63·10^{-6}(T–25 °C)^2			54Cha
	$2Hg+2Cl^- \rightarrow Hg_2Cl_2+2e^-$		0.268	43Mül, 28Ran
	$Hg+2OH^- \rightarrow HgO+H_2O+2e^-$, sol. sat. HgO, sol. 0.1 **M** NaOH		0.926	01Bar
	$Hg+2OH^- \rightarrow HgO+H_2O+2e^-$		0.098	26Fri, 34Kob
	$Hg+SO_4^{2-} \rightarrow HgSO_4+2e^-$, sol. 0.5 **M** K_2SO_4		0.680	01Bar
	$Hg+SO_4^{2-} \rightarrow HgSO_4+2e^-$, sol. sat. K_2SO_4		0.640	01Bar
Silver (Ag)	$Ag+Br^- \rightarrow AgBr+e^-$		0.071	36Har2, 37Har1, 43Har, 35Kes, 36Owe
	$Ag+Cl^- \rightarrow AgCl+e^-$		0.2224	32Har1, 43Har, 33Pre

Electrode	Electrode reaction, electrolyte solution [1])	T [K, °C]	E_{SHE} [2]) [V]	Ref.
Ag	$Ag+Cl^- \rightarrow AgCl+e^-$, sat. KCl		0.197	01Bar
	$Ag+Cl^- \rightarrow AgCl+e^-$, sol. 3 **M** KCl		0.207	01Bar
	$Ag+Cl^- \rightarrow AgCl+e^-$, sol. 0.1 **M** KCl		0.289	01Bar
	$Ag+I^- \rightarrow AgI+e^-$		–0.152	35Owe
	$Ag+I^- \rightarrow AgI+e^-$		–0.15225	57Kor
Thallium$_{(Hg)}$	$Tl+Cl^- \rightarrow TlCl+e^-$, TlCl, sol. 3.5 **M** KCl, HCl [3])	298 K [4])	–0.5706	74Bau
(Tl)$_{(Hg)}$	$Tl+Cl^- \rightarrow TlCl+e^-$, TlCl, sol. 3.5 **M** KCl, oxalate	298 K	–0.5703	74Bau
	$Tl+Cl^- \rightarrow TlCl+e^-$, TlCl, sol. 3.5 **M** KCl, phosphate	298 K	–0.5713	74Bau
	$Tl+Cl^- \rightarrow TlCl+e^-$, TlCl, sol. 3.5 **M** KCl, borax	298 K	–0.5709	74Bau
	$Tl+Cl^- \rightarrow TlCl+e^-$, TlCl, sat. KCl, HCl	298 K	–0.5734	74Bau
	$Tl+Cl^- \rightarrow TlCl+e^-$, TlCl, sat. KCl, phosphate	298 K	–0.5766	74Bau
	$Tl+Cl^- \rightarrow TlCl+e^-$, TlCl, sat. KCl, borax	298 K	–0.5795	74Bau
Quin-hydrone	hydroquinone→quinone$+2e^-+2H^+$		0.69969	33Har2

[1]) Details are given only if relevant for the stated value of E_{SHE}.

[2]) All potentials refer to a standard hydrogen electrode in the used solvent. Exceptions are stated explicitly.

[3]) This electrode is frequently called 'Thalamid™-electrode'.

[4]) For a detailed listing of temperature-dependent potential values see [74Bau].

[*]) In [$V_{int.}$].

References

26Fri Fried, F.: Z. Phys. Chem. **123** (1926) 406.

28Ran Randall, M., Young, L.E.: J. Am. Chem. Soc. **50** (1928) 989.

32Har1 Harned, H.S., Ehlers, R.W.: J. Am. Chem. Soc. **54** (1932) 1350.

33Har2 Harned, H.S., Wright, D.D.: J. Am. Chem. Soc. **55** (1933) 4849.

33Pre Prentiss, S.S., Scatchard, G.: Chem. Rev. **13** (1933) 139.

34Kob Kobayashi, Y., Wang, H.Y.: J. Sci. Hiroshima Univ. Ser. A Math. Phys. Chem. **5** (1934) 392.

34Shr2 Shrawder jr., J., Cowperthwaite, J.A.: J. Am. Chem. Soc. **56** (1934) 2340.

35Har Harned, H.S., Hamer, W.J.: J. Am. Chem. Soc. **57** (1935) 27.

35Kes Keston, A.S.: J. Am. Chem. Soc. **57** (1935) 1526.

35Owe Owen, B.B.: J. Am. Chem. Soc. **57** (1935) 1926.

36Har2 Harned, H.S., Keston, A.S., Donelson, J.G.: J. Am. Chem. Soc. **58** (1936) 989.

36Owe Owen, B.B., Foering, L.: J. Am. Chem. Soc. **58** (1936) 1575.

37Har1 Harned, H.S., Donelson, J.G.: J. Am. Chem. Soc. **59** (1937) 1280.

43Har Harned, H.S., Owen, B.B.: The Physical Chemistry of Electrolytic Solutions, New York: Reinhold, 1943.

43Mül Müller, F., Reuther, H.: Z. Elektrochem. **49** (1943) 497.

54Cha Chateau, H.: J. Chim. Phys. **51** (1954) 590.

57Kor Kortüm, G.: Lehrbuch der Elektrochemie, Weinheim: Verlag Chemie, 1957.

61Ive Ives, D.J.G., Janz, G.J.: Reference Electrodes, New York: Academic Press, 1961.

74Bau Baucke, F.G.K.: Chem. Ing. Techn. **46** (1974) 71.

01Bar Bard, A.J., Faulkner, L.R.: Electrochemical Methods, 2nd ed., New York: Wiley, 2001.

2 Cell voltages

2.1 Definition of cell voltage

Between the terminals of an electrochemical cell composed of two electrodes (i.e. two half cells) a voltage being equivalent to the electrode potential difference between both terminals appears. This voltage measured in the absence of any electrical current flowing across the participating interfaces has also been called electromotive force or - particularly confusing - cell potential. Subsequently - as in the previous section - the term potential is reserved for use with an electrode (a half cell) having of course in mind, that by definition this electrode potential value is numerically equivalent to the cell voltage of a cell composed of this electrode and the standard hydrogen electrode, whereas voltage refers to a complete cell. The measured cell voltage is composed of the electrode potentials and further contributions. As discussed in detail above an electrode potential is actually the potential difference between the bulk of two phases. Beyond these two obvious contributions of the electrodes to the cell voltage further potential steps or differences e.g. at the interface between the electrolyte solutions established by a porous frit or a membrane may be present. At first glance a listing of cell voltages might seem to be superfluous, because it might contain only differences between electrode potentials already listed in the previous chapter. As already mentioned additional contributions might be involved, also certain useful electrode compositions might have not been studied as single electrodes, but only in complete cells. In addition the growing number of sensor systems based on the measurement of the voltage of an electrochemical cell has stimulated interest in cell voltages. The use of solid state electrochemical cells has been reviewed [89Jac], the applicability of CaF_2 as a solid electrolyte has been discussed extensively [89Aza3].

The influence of interfacial potentials (diffusion or liquid junction potentials) established at the boundary between two different electrolyte solutions (based on e.g. aqueous and nonaqueous solvents) has been investigated frequently, for a thorough overview see Jakuszewski and Woszczak [68Jak2]. Concerning the usage of absolute and international Volt see preceding chapter.

General remarks on the tables

Cells are arranged according to the first chemical element of the electrode material first mentioned in the entry (left hand side according to the Stockholm convention, negative pole in common terms). When an element has been used as amalgam also the amalgam entries follow those obtained with the pure element. Within every group the potential-determining constituent of the other electrode serves as next order criterion. A final ordering criterion is the composition of the electrolyte solution, with the organic solvents arranged according to Hill's system [00Hil] with the ordering of the element symbol and trivial names kept in the usual manner. Metal contacts mentioned for completeness and other construction elements are not taken into account (although they are mentioned especially when explicitly stated in the original report) in this arrangement, although they may be mentioned when provided in the original report. In case of redox electrodes, where both potential-determining species are dissolved, the metal (or any other electron conducting material) is used instead. The reader searching for a cell containing such redox electrode may start looking at frequently used electrode materials (platinum, mercury, carbon) first. Concentration of the electrolyte solution constituents are given in the units as originally reported. When no concentration is given, unity values are assumed, thus the cell voltage refers to a standard cell voltage (although this term does not seem to be properly defined) with two standard electrodes combined.

Vertical lines indicate phase boundaries inside the cell. Double bars indicate inserted porous separators in order to avoid intermixing of liquid phases. In case of states composed of several solid phases serving as electronic conductors and/or as electrolytes with no well defined boundaries or other reasons reported in the original publications, where sometimes no phase boundaries are explicitly stated, this is carefully reflected in the tables below. The sequence of electrodes follows the Stockholm convention, i.e. the electrode forming the negative terminal of a cell (the minus pole of the cell) is always given first (left hand). Frequently data referring to variable concentrations of an electrolyte solution constituent are given. Sometimes the cell voltage varies as a function of the said concentration resulting in an inversion of polarity. Strictly adhering to the stated convention would result in a separation of this list from subsequent entries of a second part found at the entry of the respective element of the initially right-hand side electrode. This would cause considerable searching for the reader. To avoid this the cell voltages, where the negative terminal is on the right-hand side of the description of the cell, are marked by a minus sign.

References

00Hil Hill, E.A.: J. Am. Chem. Soc. **22** (1900) 478.

68Jak2 Jakuszewski, B., Kozlowski, Z.: Soc. Sci. Lodz. Acta Chim. **13** (1968) 21.

89Aza3 Azad, A.M., Sreedharan, O.M.: J. Appl. Electrochem. **19** (1989) 475.

89Jac Jacob, K.T., Swaminathan, K., Sreedharan, O.M.: Solid State Ionics **34** (1989) 167.

Table 2.1. Cell voltages with aqueous electrolyte systems. [1])

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Barium (Ba)			
Ba_xHg | $BaCl_2$, 0.01 **M** ‖ $BaCl_2$, 0.005 **M** | Ba_xHg		0.00700	32Mas2
Ba_xHg | $BaCl_2$, 0.005 **M** ‖ $BaCl_2$, 0.0025 **M** | Ba_xHg		0.00725	32Mas2
Ba_xHg | $BaCl_2$, 0.0025 **M** ‖ $BaCl_2$, 0.00125 **M** | Ba_xHg		0.00731	32Mas2
Ba_xHg | $BaCl_2$, 0.00125 **M** ‖ $BaCl_2$, 0.000625 **M** | Ba_xHg		0.00728	32Mas2
Ba_xHg | $BaCl_2$, 0.0500 **m**, Hg_2Cl_2 | Hg	273 K	2.0072	34Tip
	298 K	2.0453	34Tip
	308 K	2.0595	34Tip
Ba_xHg | $BaCl_2$, 0.1159 **m**, Hg_2Cl_2 | Hg	273 K	1.9832	34Tip
	298 K	2.0183	34Tip
	308 K	2.0317	34Tip
Ba_xHg | $BaCl_2$, 0.2293 **m**, Hg_2Cl_2 | Hg	273 K	1.9631	34Tip
	308 K	2.0090	34Tip
Ba_xHg | $BaCl_2$, 0.2373 **m**, Hg_2Cl_2 | Hg	298 K	1.9953	34Tip
Ba_xHg | $BaCl_2$, 0.4546 **m**, Hg_2Cl_2 | Hg	273 K	1.9433	34Tip
	298 K	1.9738	34Tip
Ba_xHg | $BaCl_2$, 0.4786 **m**, Hg_2Cl_2 | Hg	308 K	1.9838	34Tip
Ba_xHg | $BaCl_2$, 0.8807 **m**, Hg_2Cl_2 | Hg	273 K	1.9199	34Tip
	298 K	1.9489	34Tip
	308 K	1.9598	34Tip
Ba_xHg | $BaCl_2$, 1.4015 **m**, Hg_2Cl_2 | Hg	273 K	1.9003	34Tip
Ba_xHg | $BaCl_2$, 1.5271 **m**, Hg_2Cl_2 | Hg	298 K	1.9243	34Tip
Ba_xHg | $BaCl_2$, 1.5713 **m**, Hg_2Cl_2 | Hg	308 K	1.9333	34Tip
Ba_xHg | $BaCl_2$, 1.7783 **m**, Hg_2Cl_2 | Hg	298 K	1.9166	34Tip
Beryllium (Be)			
Be | $Be(ClO_4)_2$, 0.005 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.804	30Bir
Be | $Be(ClO_4)_2$, 0.05 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.825	30Bir
Be | $BeSO_4$, 0.05 **M**, KCl, 0.0 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.820	27Bod
Be | $BeSO_4$, 0.1 **M**, KCl, 0.5 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		1.029	27Bod
Bromine (Br)			
Pt, Br_2, 0.00385 **M** | HBr, 0.3 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.0433	32Wil
Pt, Br_2, 0.0068 **M** | HBr, 0.6 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.0764	32Wil
Pt, Br_2, 0.01275 **M** | HBr, 1.0 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.1020	32Wil
Pt, Br_2, 0.0233 **M** | HBr, 2.0 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.1433	32Wil
Pt, Br_2, 0.0446 **M** | HBr, 3.7 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.2013	32Wil
Pt, Br_2, 0.0594 **M** | HBr, 5.0 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.2397	32Wil
Pt, Br_2, 0.0684 **M** | HBr, 6.0 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.2681	32Wil
Pt, Br_2, 0.0733 **M** | HBr, 7.0 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.2949	32Wil
Pt, Br_2, 0.0743 **M** | HBr, 8.5 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.3340	32Wil
Pt, Br_2, 0.00385 **M** | HBr, 0.3 **M** ‖ KCl, 3.0 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.0264	32Wil
Pt, Br_2, 0.0068 **M** | HBr, 0.6 **M** ‖ KCl, 3.0 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.0475	32Wil
Pt, Br_2, 0.01275 **M** | HBr, 1.0 **M** ‖ KCl, 3.0 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.0648	32Wil

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, Br_2, 0.0233 **M** | HBr, 2.0 **M** ‖ KCl, 3.0 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.0989	32Wil
Pt, Br_2, 0.0446 **M** | HBr, 3.7 **M** ‖ KCl, 3.0 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.1437	32Wil
Pt, Br_2, 0.0594 **M** | HBr, 5.0 **M** ‖ KCl, 3.0 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.1745	32Wil
Pt, Br_2, 0.0684 **M** | HBr, 6.0 **M** ‖ KCl, 3.0 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.2002	32Wil
Pt, Br_2, 0.0733 **M** | HBr, 7.0 **M** ‖ KCl, 3.0 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.2236	32Wil
Pt, Br_2, 0.0743 **M** | HBr, 8.5 **M** ‖ KCl, 3.0 **M** ‖ HBr, 0.1 **M** | Br_2, 0.001575 **M**, Pt	18 °C	0.2588	32Wil
Cadmium (Cd)			
Cd | $CdSO_4$, 0.5 **m** | $Cd_{sat.}Hg$	0 °C	0.0554	34Par
	10 °C	0.0536	34Par
	20 °C	0.0515	34Par
	30 °C	0.0495	34Par
Cd | CdI_2, 0.0039 **m**, AgI,sol. | Ag		0.4648	28Get1
Cd | CdI_2, 0.0098 **m**, AgI,sol. | Ag		0.4425	28Get1
Cd | CdI_2, 0.0197 **m**, AgI,sol. | Ag		0.4325	28Get1
Cd | CdI_2, 0.0389 **m**, AgI,sol. | Ag		0.4250	28Get1
Cd | CdI_2, 0.0969 **m**, AgI,sol. | Ag		0.4102	28Get1
Cd | CdI_2, 0.2021 **m**, AgI,sol. | Ag		0.3993	28Get1
Cd | CdI_2, 0.4883 **m**, AgI,sol. | Ag		0.3869	28Get1
Cd | CdI_2, 1.0274 **m**, AgI,sol. | Ag		0.3740	28Get1
Cd | $CdBr_2$, 0.0012 **m**, Hg_2Br_2 | Hg		0.7624	28Get2
Cd | $CdBr_2$, 0.0029 **m**, Hg_2Br_2 | Hg		0.7507	28Get2
Cd | $CdBr_2$, 0.0048 **m**, Hg_2Br_2 | Hg		0.7368	28Get2
Cd | $CdBr_2$, 0.0117 **m**, Hg_2Br_2 | Hg		0.7202	28Get2
Cd | $CdBr_2$, 0.0293 **m**, Hg_2Br_2 | Hg		0.6993	28Get2
Cd | $CdBr_2$, 0.0476 **m**, Hg_2Br_2 | Hg		0.6915	28Get2
Cd | $CdBr_2$, 0.0732 **m**, Hg_2Br_2 | Hg		0.6805	28Get2
Cd | $CdBr_2$, 0.1172 **m**, Hg_2Br_2 | Hg		0.6733	28Get2
Cd | $CdBr_2$, 0.1872 **m**, Hg_2Br_2 | Hg		0.6640	28Get2
Cd | $CdBr_2$, 0.2929 **m**, Hg_2Br_2 | Hg		0.6551	28Get2
Cd | $CdBr_2$, 0.4756 **m**, Hg_2Br_2 | Hg		0.6452	28Get2
Cd | $CdBr_2$, 3.1594 **m**, Hg_2Br_2 | Hg		0.6051	28Get2
Cd | $CdCl_2$, 0.00086 **m**, Hg_2Cl_2 | Hg		0.9155	28Get2
Cd | $CdCl_2$, 0.0031 **m**, Hg_2Cl_2 | Hg		0.8667	28Get2
Cd | $CdCl_2$, 0.0062 **m**, Hg_2Cl_2 | Hg		0.8727	28Get2
Cd | $CdCl_2$, 0.0086 **m**, Hg_2Cl_2 | Hg		0.8485	28Get2
Cd | $CdCl_2$, 0.0164 **m**, Hg_2Cl_2 | Hg		0.8325	28Get2
Cd | $CdCl_2$, 0.0311 **m**, Hg_2Cl_2 | Hg		0.8198	28Get2
Cd | $CdCl_2$, 0.0622 **m**, Hg_2Cl_2 | Hg		0.8063	28Get2
Cd | $CdCl_2$, 0.0819 **m**, Hg_2Cl_2 | Hg		0.8014	28Get2
Cd | $CdCl_2$, 0.1556 **m**, Hg_2Cl_2 | Hg		0.7866	28Get2
Cd | $CdCl_2$, 0.3113 **m**, Hg_2Cl_2 | Hg		0.7752	28Get2
Cd | $CdCl_2$, 5.035 **m**, Hg_2Cl_2 | Hg		0.7311	28Get2

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Cd | $CdCl_2$, 5.909 **m**, Hg_2Cl_2 | Hg		0.7286	28Get2
Cd | $CdCl_2$, 0.00966 **m**, Hg_2Cl_2 | Hg		0.8465	31Get2
Cd | $CdCl_2$, 0.04805 **m**, Hg_2Cl_2 | Hg		0.8117	31Get2
Cd | $CdCl_2$, 0.09986 **m**, Hg_2Cl_2 | Hg		0.7954	31Get2
Cd | $CdCl_2$, 0.20027 **m**, Hg_2Cl_2 | Hg		0.7821	31Get2
Cd | $CdCl_2$, 0.28766 **m**, Hg_2Cl_2 | Hg		0.7763	31Get2
Cd | $CdCl_2$, 0.49084 **m**, Hg_2Cl_2 | Hg		0.7663	31Get2
Cd | $CdCl_2$, 0.99472 **m**, Hg_2Cl_2 | Hg		0.7555	31Get2
Cd | $CdSO_4$, 0.0014 **m**, Hg_2SO_4 | Hg		1.1438	28Get2
Cd | $CdSO_4$, 0.0055 **m**, Hg_2SO_4 | Hg		1.1484	28Get2
Cd | $CdSO_4$, 0.0139 **m**, Hg_2SO_4 | Hg		1.1399	28Get2
Cd | $CdSO_4$, 0.0554 **m**, Hg_2SO_4 | Hg		1.1285	28Get2
Cd | $CdSO_4$, 0.1328 **m**, Hg_2SO_4 | Hg		1.1178	28Get2
Cd | $CdSO_4$, 0.1386 **m**, Hg_2SO_4 | Hg		1.1153	28Get2
Cd | $CdSO_4$, 0.2217 **m**, Hg_2SO_4 | Hg		1.1130	28Get2
Cd | $CdSO_4$, 0.5542 **m**, Hg_2SO_4 | Hg		1.1034	28Get2
Cd | $CdSO_4$, 1.386 **m**, Hg_2SO_4 | Hg		1.0926	28Get2
Cd | $CdSO_4$, 2.125 **m**, Hg_2SO_4 | Hg		1.0810	28Get2
Cd | $CdSO_4$, 2.656 **m**, Hg_2SO_4 | Hg		1.0768	28Get2
Cd | $CdSO_4$, 3.698 **m**, Hg_2SO_4 | Hg		1.0656	28Get2
Cd | $CdSO_4$*8/3 H_2O,sol., H_2SO_4, 0.078 **M** ‖ H_2SO_4, 0.078 **M**, Hg_2SO_4 | Hg	15.67 °C	1.0184	18Oba
	20 °C	1.0182	18Oba
Cd | $CdSO_4$*8/3 H_2O,sol., H_2SO_4, 0.195 **M** ‖ H_2SO_4, 0.195 **M**, Hg_2SO_4 | Hg	20 °C	1.0181	18Oba
Cd | $CdSO_4$*8/3 H_2O,sol., H_2SO_4, 0.756 **M** ‖ H_2SO_4, 0.756 **M**, Hg_2SO_4 | Hg	15.67 °C	1.0178	18Oba
Cd | $CdSO_4$*8/3 H_2O,sol., H_2SO_4, 1.151 **M** ‖ H_2SO_4, 1.151 **M**, Hg_2SO_4 | Hg	20 °C	1.0172	18Oba
Cd | $CdSO_4$*8/3 H_2O,sol., H_2SO_4, 0.021 **M** ‖ H_2SO_4, 0.021 **M**, Hg_2SO_4 | Hg		1.0180	26Vos
Cd | $CdSO_4$*8/3 H_2O,sol., H_2SO_4, 0.078 **M** ‖ H_2SO_4, 0.078 **M**, Hg_2SO_4 | Hg		1.0179	26Vos
Cd | $K_2Cd(CN)_4$, 9.13 m**M**, KCN, 3.56 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	14 °C	0.952	32Bri
Cd | $K_2Cd(CN)_4$, 9.13 m**M**, KCN, 16.93 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	14 °C	1.013	32Bri
Cd | $K_2Cd(CN)_4$, 7.44 m**M**, KCN, 15.83 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	14 °C	0.982	32Bri
Cd | $K_2Cd(CN)_4$, 2.92 m**M**, KCN, 26.55 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	14 °C	1.022	32Bri
Cd | CdI_2, 0.0030 **m**, PbI_2,sol. | Pb		0.2542	28Get1
Cd | CdI_2, 0.0040 **m**, PbI_2,sol. | Pb		0.2487	28Get1
Cd | CdI_2, 0.0050 **m**, PbI_2,sol. | Pb		0.2463	28Get1
Cd | CdI_2, 0.0099 **m**, PbI_2,sol. | Pb		0.2332	28Get1
Cd | CdI_2, 0.0196 **m**, PbI_2,sol. | Pb		0.2232	28Get1
Cd | CdI_2, 0.0389 **m**, PbI_2,sol. | Pb		0.2142	28Get1
Cd | CdI_2, 0.0968 **m**, PbI_2,sol. | Pb		0.2017	28Get1
Cd | CdI_2, 0.2030 **m**, PbI_2,sol. | Pb		0.1923	28Get1
Cd | CdI_2, 0.4897 **m**, PbI_2,sol. | Pb		0.1807	28Get1
Cd | CdI_2, 1.0274 **m**, PbI_2,sol. | Pb		0.1674	28Get1
Cd | CdI_2, 1.2967 **m**, PbI_2,sol. | Pb		0.1678	28Get1

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Cd | $Cd_{11wt\%}Hg$ | $CdCl_2$, AgCl | Ag	25 °C	0.57390	36Har1
(U_T [V]=0.57390–394·10^{-6}(T–25 °C)–3.63·10^{-6}(T–25 °C)^2)			
Cd | $Cd_{11wt\%}Hg$ | $CdCl_2$, 0.0005 **m**, AgCl | Ag	25 °C	0.8539	36Har1
(U_T [V]=0.8539+560·10^{-6}(T–25 °C)–3.48·10^{-6}(T–25 °C)^2)			
Cd | $Cd_{11wt\%}Hg$ | $CdCl_2$, 0.001 **m**, AgCl | Ag	25 °C	0.82997	36Har1
(U_T [V]=0.82997+494·10^{-6}(T–25 °C)–3.40·10^{-6}(T–25 °C)^2)			
Cd | $Cd_{11wt\%}Hg$ | $CdCl_2$, 0.005 **m**, AgCl | Ag	25 °C	0.77851	36Har1
(U_T [V]=0.77851+356·10^{-6}(T–25 °C)–3.13·10^{-6}(T–25 °C)^2)			
Cd | $Cd_{11wt\%}Hg$ | $CdCl_2$, 0.01 **m**, AgCl | Ag	25 °C	0.75846	36Har1
(U_T [V]=0.75846+305·10^{-6}(T–25 °C)–2.93·10^{-6}(T–25 °C)^2)			
Cd | $Cd_{11wt\%}Hg$ | $CdCl_2$, 0.05 **m**, AgCl | Ag	25 °C	0.71739	36Har1
(U_T [V]=0.71739+185·10^{-6}(T–25 °C)–2.05·10^{-6}(T–25 °C)^2)			
Cd | $Cd_{11wt\%}Hg$ | $CdCl_2$, 0.1 **m**, AgCl | Ag	25 °C	0.70175	36Har1
(U_T [V]=0.70175+136·10^{-6}(T–25 °C)–1.68·10^{-6}(T–25 °C)^2)			
Cd | $Cd_{11wt\%}Hg$ | $CdCl_2$, 0.5 **m**, AgCl | Ag	25 °C	0.66864	36Har1
(U_T [V]=0.66864+48·10^{-6}(T–25 °C)–0.80·10^{-6}(T–25 °C)^2)			
Cd | $Cd_{11wt\%}Hg$ | $CdCl_2$, 1.0 **m**, AgCl | Ag	25 °C	0.65640	36Har1
(U_T [V]=0.65640+12·10^{-6}(T–25 °C)–0.44·10^{-6}(T–25 °C)^2)			
$Cd_{sat.}Hg$ | $CdBr_2$*4 H_2O,sat., AgBr | Ag		0.4883	32Spe1
$Cd_{sat.}Hg$ | $CdBr_2$,sat., AgBr | Ag	35.6 °C	0.3776	32Spe1
	40 °C	0.4776	32Spe1
$Cd_{sat.}Hg$ | $CdBr_2$, 0.0100 **m**, AgBr | Ag		0.5932	29Luc1
$Cd_{sat.}Hg$ | $CdBr_2$, 0.01993 **m**, AgBr | Ag		0.5756	29Luc1
$Cd_{sat.}Hg$ | $CdBr_2$, 0.05629 **m**, AgBr | Ag		0.5515	29Luc1
$Cd_{sat.}Hg$ | $CdBr_2$, 0.100 **m**, AgBr | Ag		0.5395	29Luc1
$Cd_{sat.}Hg$ | $CdBr_2$, 0.3026 **m**, AgBr | Ag		0.5264	29Luc1
$Cd_{sat.}Hg$ | $CdBr_2$, 0.5444 **m**, AgBr | Ag		0.5101	29Luc1
$Cd_{sat.}Hg$ | $CdBr_2$, 0.9996 **m**, AgBr | Ag		0.5005	29Luc1
$Cd_{sat.}Hg$ | $CdBr_2$, 1.896 **m**, AgBr | Ag		0.4888	29Luc1
$Cd_{sat.}Hg$ | $CdBr_2$, 3.082 **m**, AgBr | Ag		0.4777	29Luc1
$Cd_{sat.}Hg$ | $CdCl_2$, 0.0101 **m**, AgCl | Ag		0.7406	29Luc1
$Cd_{sat.}Hg$ | $CdCl_2$, 0.05056 **m**, AgCl | Ag		0.6995	29Luc1
$Cd_{sat.}Hg$ | $CdCl_2$, 0.1000 **m**, AgCl | Ag		0.6843	29Luc1
$Cd_{sat.}Hg$ | $CdCl_2$, 0.5101 **m**, AgCl | Ag		0.6534	29Luc1
$Cd_{sat.}Hg$ | $CdCl_2$, 0.9999 **m**, AgCl | Ag		0.6424	29Luc1
$Cd_{sat.}Hg$ | $CdCl_2$, 2.045 **m**, AgCl | Ag		0.6317	29Luc1
$Cd_{sat.}Hg$ | $CdCl_2$, 6.146 **m**, AgCl | Ag		0.6105	29Luc1
Cd_{Hg} | $CdCl_2$ | AgCl | Ag		0.5744	58Tre1
Cd_{Hg} | $CdCl_2$, 4.0972 m**m**, Na_3citrate, 5.2310 m**m**, AgCl | Ag	24.96 °C	0.8293	58Tre2
Cd_{Hg} | $CdCl_2$, 4.9487 m**m**, Na_3citrate, 7.6584 m**m**, AgCl | Ag	24.96 °C	0.82803	58Tre2
Cd_{Hg} | $CdCl_2$, 5.0353 m**m**, Na_3citrate, 6.3399 m**m**, AgCl | Ag	24.96 °C	0.82809	58Tre2
Cd_{Hg} | $CdCl_2$, 5.6126 m**m**, Na_3citrate, 7.1657 m**m**, AgCl | Ag	24.96 °C	0.81870	58Tre2
Cd_{Hg} | $CdCl_2$, 6.5715 m**m**, Na_3citrate, 7.7265 m**m**, AgCl | Ag	24.96 °C	0.81021	58Tre2
Cd_{Hg} | $CdCl_2$, 7.2661 m**m**, Na_3citrate, 8.8274 m**m**, AgCl | Ag	24.96 °C	0.80792	58Tre2
Cd_{Hg} | $CdCl_2$, 7.8425 m**m**, Na_3citrate, 10.012 m**m**, AgCl | Ag	24.96 °C	0.80722	58Tre2
Cd_{Hg} | $CdCl_2$, 8.4232 m**m**, Na_3citrate, 10.551 m**m**, AgCl | Ag	24.96 °C	0.80476	58Tre2

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
$Cd_{10wt\%}Hg \mid CdBr_2*4\ H_2O, CdBr_2,sat., Hg_2Br_2 \mid Hg$	20 °C	0.5581	33Ish4
	25 °C	0.5561	33Ish4
	30 °C	0.5539	33Ish4
	35.5 °C	0.5510	33Ish4
	36 °C	0.5507	33Ish4
$Cd_{10wt\%}Hg \mid CdBr_2,sol., CdBr_2,sat., Hg_2Br_2 \mid Hg$	36.5 °C	0.5509	33Ish4
	37.5 °C	0.5512	33Ish4
	38 °C	0.5513	33Ish4
$Cd_{10wt\%}Hg \mid CdBr_2$, 52.90wt%, $Hg_2Br_2 \mid Hg$	25 °C	0.5563	33Ish4
	30 °C	0.5580	33Ish4
	35 °C	0.5597	33Ish4
$Cd_{10wt\%}Hg \mid CdBr_2$, 46.34wt%, $Hg_2Br_2 \mid Hg$	25 °C	0.5633	33Ish4
$Cd_{10wt\%}Hg \mid CdBr_2$, 48.15wt%, $Hg_2Br_2 \mid Hg$	25 °C	0.5614	33Ish4
$Cd_{10wt\%}Hg \mid CdBr_2$, 50.82wt%, $Hg_2Br_2 \mid Hg$	25 °C	0.5586	33Ish4
$Cd_{10wt\%}Hg \mid CdBr_2$, 52.14wt%, $Hg_2Br_2 \mid Hg$	25 °C	0.5571	33Ish4
$Cd_{9wt\%}Hg \mid CdI_2,sol., CdI_2,sat., Hg_2I_2 \mid Hg$	25 °C	0.4171	28Yos
	30 °C	0.4189	28Yos
	35 °C	0.4207	28Yos
$Cd_{sat.}Hg \mid CdI_2$, 0.001025 **m** $\Vert$ CdI_2, 0.001025 **m**, $Hg_2I_2 \mid Hg$	25 °C	0.5695	37Bat
$Cd_{sat.}Hg \mid CdI_2$, 0.0050 **m** $\Vert$ CdI_2, 0.0050 **m**, $Hg_2I_2 \mid Hg$	25 °C	0.5252	37Bat
	30 °C	0.5296	37Bat
	35 °C	0.5341	37Bat
$Cd_{sat.}Hg \mid CdI_2$, 0.0100 **m** $\Vert$ CdI_2, 0.0100 **m**, $Hg_2I_2 \mid Hg$	25 °C	0.5083	37Bat
	30 °C	0.5123	37Bat
	35 °C	0.5165	37Bat
$Cd_{sat.}Hg \mid CdI_2$, 0.02015 **m** $\Vert$ CdI_2, 0.02015 **m**, $Hg_2I_2 \mid Hg$	25 °C	0.4932	37Bat
	30 °C	0.4969	37Bat
	35 °C	0.5006	37Bat
$Cd_{sat.}Hg \mid CdI_2$, 0.0500 **m** $\Vert$ CdI_2, 0.0500 **m**, $Hg_2I_2 \mid Hg$	25 °C	0.4779	37Bat
	30 °C	0.4811	37Bat
	35 °C	0.4846	37Bat
$Cd_{sat.}Hg \mid CdI_2$, 0.100 **m** $\Vert$ CdI_2, 0.100 **m**, $Hg_2I_2 \mid Hg$	25 °C	0.4679	37Bat
	30 °C	0.4709	37Bat
	35 °C	0.4739	37Bat
$Cd_{sat.}Hg \mid CdI_2$, 0.250 **m** $\Vert$ CdI_2, 0.250 **m**, $Hg_2I_2 \mid Hg$	25 °C	0.4557	37Bat
	30 °C	0.4584	37Bat
	35 °C	0.4611	37Bat
$Cd_{sat.}Hg \mid CdI_2$, 0.4843 **m** $\Vert$ CdI_2, 0.4843 **m**, $Hg_2I_2 \mid Hg$	25 °C	0.4467	37Bat
	30 °C	0.4493	37Bat
	35 °C	0.4519	37Bat
$Cd_{sat.}Hg \mid CdI_2$, 1.000 **m** $\Vert$ CdI_2, 1.000 **m**, $Hg_2I_2 \mid Hg$	25 °C	0.4351	37Bat
	30 °C	0.4375	37Bat
	35 °C	0.4400	37Bat
$Cd_{sat.}Hg \mid CdI_2$, 1.983 **m** $\Vert$ CdI_2, 1.983 **m**, $Hg_2I_2 \mid Hg$	25 °C	0.4209	37Bat
	30 °C	0.4233	37Bat
	35 °C	0.4258	37Bat
$Cd_{sat.}Hg \mid CdI_2$, 2.366 **m** $\Vert$ CdI_2, 2.366 **m**, $Hg_2I_2 \mid Hg$	25 °C	0.4170	37Bat

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
$Cd_{sat.}Hg \mid CdI_2$, 2.404 **m** $\Vert CdI_2$, 2.404 **m**, $Hg_2I_2 \mid Hg$	30 °C	0.4188	37Bat
$Cd_{sat.}Hg \mid CdI_2$, 2.453 **m** $\Vert CdI_2$, 2.453 **m**, $Hg_2I_2 \mid Hg$	35 °C	0.4206	37Bat
$Cd_{sat.}Hg \mid CdCl_2$*2.5 H_2O,sat. $\Vert CdCl_2$*2.5 H_2O,sat., Hg_2Cl_2,sol. $\mid Hg$	17 °C	0.6713	30Pri
	20 °C	0.6711	30Pri
	25 °C	0.6706	30Pri
	30 °C	0.6701	30Pri
$Cd_{sat.}Hg \mid CdCl_2$*H_2O,sat. $\Vert CdCl_2$*1H_2O,sat., Hg_2Cl_2,sol. $\mid Hg$	35 °C	0.6699	30Pri
	40 °C	0.6714	30Pri
$Cd_{10wt\%}Hg \mid Cd(OH)_2$,sol., NaOH, 1.0%, HgO,sol. $\mid Hg$		0.8598	32Ish2
$Cd_{10wt\%}Hg \mid Cd(OH)_2$,sol., NaOH, 33%, HgO,sol. $\mid Hg$		0.8502	32Ish2
$Cd_{8wt\%}Hg \mid CdSO_4 \mid Cd_{0.74wt\%}Hg$	0 °C	0.0159	31LaM
	17 °C	0.02389	31LaM
	25 °C	0.02759	31LaM
$Cd_{8wt\%}Hg \mid CdSO_4 \mid Cd_{2.12wt\%}Hg$	25 °C	0.01395	31LaM
$Cd_{8wt\%}Hg \mid CdSO_4 \mid Cd_{2.77wt\%}Hg$	0 °C	0.00023	31LaM
$Cd_{8wt\%}Hg \mid CdSO_4 \mid Cd_{4.76wt\%}Hg$	25 °C	0.00331	31LaM
$Cd_{8wt\%}Hg \mid CdSO_4 \mid Cd_{5.87wt\%}Hg$	25 °C	0.00040	31LaM
$Cd_{10...13wt\%}Hg \mid CdSO_4 8{\cdot}8/3H_2O_{sat.}$, $H_2O \mid Hg_2SO_{4,sat.}$, $H_2O \mid Hg$ (U_T [V]=U_{293K}–[4.06·10^{-5}(T–293 K)–0.95·10^{-6}(T–293 K) +1·10^{-8}(T–293 K)]) (Weston-Cell)	293 K	1.0183	55Sti [3])
$Cd_{sat.}Hg \mid CdCl_2$*2.5 H_2O,sat. $\Vert CdCl_2$*2.5 H_2O,sat., $PbCl_2$,sol. $\mid Pb_{6wt\%}Hg$	17 °C	0.1425	30Pri
	20 °C	0.1419	30Pri
	25 °C	0.1407	30Pri
	30 °C	0.1395	30Pri
$Cd_{sat.}Hg \mid CdCl_2$*H_2O,sat. $\Vert CdCl_2$*1H_2O,sat., $PbCl_2$,sol. $\mid Pb_{6wt\%}Hg$	35 °C	0.1385	30Pri
	40 °C	0.1392	30Pri
$Cd_{sat.}Hg \mid CdSO_4$, 0.0005 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$	0 °C	0.2004	33LaM
	20 °C	0.2028	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.001 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$	0 °C	0.1869	33LaM
	20 °C	0.1887	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.002 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$	0 °C	0.1739	33LaM
	20 °C	0.1749	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.005 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$	0 °C	0.1580	33LaM
	20 °C	0.1580	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.01 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$	0 °C	0.1465	33LaM
	20 °C	0.1459	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.02 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$	0 °C	0.1350	33LaM
	20 °C	0.1340	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.000458 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$		0.2024	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.000727 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$		0.1943	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.001077 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$		0.1861	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.001993 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$		0.1738	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.005618 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$		0.1546	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.008351 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$		0.1475	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.01096 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$		0.1428	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.02237 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$		0.1310	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.0529 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$		0.1190	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.1268 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$		0.1086	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 0.3496 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$		0.0963	33LaM

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
$Cd_{sat.}Hg \mid CdSO_4$, 1.027 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$		0.0826	33LaM
$Cd_{sat.}Hg \mid CdSO_4$, 3.698 **m**, $PbSO_4$,sol. $\mid Pb_{sat.}Hg$		0.0464	33LaM
Calcium (Ca)			
$Ca_{0.0256wt\%}Hg \mid CaCl_2$, 0.0099 **M** $\mid$ KCl, 3.0 **M** $\mid$ KCl, 1.0 **M**, $Hg_2Cl_2 \mid Hg$	294 K	2.2649	26Dru1
$Ca_{0.0256wt\%}Hg \mid CaCl_2$, 0.0197 **M** $\mid$ KCl, 3.0 **M** $\mid$ KCl, 1.0 **M**, $Hg_2Cl_2 \mid Hg$	294 K	2.2564	26Dru1
$Ca_{0.0256wt\%}Hg \mid CaCl_2$, 0.0495 **M** $\mid$ KCl, 3.0 **M** $\mid$ KCl, 1.0 **M**, $Hg_2Cl_2 \mid Hg$	294 K	2.2459	26Dru1
$Ca_xHg \mid CaCl_2$, 0.00125 **M** $\Vert$ $CaCl_2$, 0.000625 **M** $\mid Ca_xHg$	303 K	0.00664	32Mas1
$Ca_xHg \mid CaCl_2$, 0.0025 **M** $\Vert$ $CaCl_2$, 0.00125 **M** $\mid Ca_xHg$	303 K	0.00675	32Mas1
$Ca_xHg \mid CaCl_2$, 0.005 **M** $\Vert$ $CaCl_2$, 0.0025 **M** $\mid Ca_xHg$	303 K	0.00659	32Mas1
$Ca_xHg \mid CaCl_2$, 0.01 **M** $\Vert$ $CaCl_2$, 0.005 **M** $\mid Ca_xHg$	303 K	0.00615	32Mas1
$Ca_{0.012wt\%}Hg \mid Ca(OH)_2$, 0.0022 **m** $\mid$ KCl,sat. $\mid$ KCl,sat., $Hg_2Cl_2 \mid Hg$		2.2266	29Fos
$Ca_{0.012wt\%}Hg \mid Ca(OH)_2$, 0.0044 **m** $\mid$ KCl,sat. $\mid$ KCl,sat., $Hg_2Cl_2 \mid Hg$		2.2208	29Fos
$Ca_{0.012wt\%}Hg \mid Ca(OH)_2$, 0.0110 **m** $\mid$ KCl,sat. $\mid$ KCl,sat., $Hg_2Cl_2 \mid Hg$		2.2135	29Fos
$Ca_{0.012wt\%}Hg \mid Ca(OH)_2$, 0.0153 **m** $\mid$ KCl,sat. $\mid$ KCl,sat., $Hg_2Cl_2 \mid Hg$		2.2113	29Fos
$Ca_{0.012wt\%}Hg \mid Ca(OH)_2$, 0.0219 **m** $\mid$ KCl,sat. $\mid$ KCl,sat., $Hg_2Cl_2 \mid Hg$		2.2053	29Fos
$Ca_{0.012wt\%.}Hg \mid CaSO_4$, 0.0038 **m** $\mid$ KCl,sat. $\mid$ KCl,sat., $Hg_2Cl_2 \mid Hg$	23.2 °C	2.1959	29Fos
$Ca_{0.012wt\%}Hg \mid CaSO_4$, 0.0077 **m** $\mid$ KCl,sat. $\mid$ KCl,sat., $Hg_2Cl_2 \mid Hg$	23.2 °C	2.1906	29Fos
$Ca_{0.012wt\%}Hg \mid CaSO_4$, 0.0153 **m** $\mid$ KCl,sat. $\mid$ KCl,sat., $Hg_2Cl_2 \mid Hg$	23.2 °C	2.1853	29Fos
Chlorine (Cl)			
Pt, $Cl_2 \mid$ HCl, 0.01 **m**, NaCl, 0.0 **m** $\Vert$ KCl, 1.0 **m**, HCl, 0.1 **m** $\mid Cl_2$, Pt		–0.0680	25Wil
Pt, $Cl_2 \mid$ HCl, 0.01 **m**, NaCl, 4.5 **m** $\Vert$ KCl, 1.0 **m**, HCl,0.1 **m** $\mid Cl_2$, Pt		–0.0668	25Wil
Pt, $Cl_2 \mid$ HCl, 0.1 **m**, NaCl, 2.0 **m** $\Vert$ KCl, 1.0 **m**, HCl, 0.1 **m** $\mid Cl_2$, Pt		–0.0380	25Wil
Pt, $Cl_2 \mid$ HCl, 0.1 **m**, NaCl, 4.0 **m** $\Vert$ KCl, 1.0 **m**, HCl, 0.1 **m** $\mid Cl_2$, Pt		–0.0721	25Wil
Pt, $Cl_2 \mid$ HCl, 1.0 **m**, NaCl, 0.0 **m** $\Vert$ KCl, 1.0 **m**, HCl, 0.1 **m** $\mid Cl_2$, Pt		0.0685	25Wil
Pt, $Cl_2 \mid$ HCl, 1.0 **m**, NaCl, 2.0 **m** $\Vert$ KCl, 1.0 **m**, HCl, 0.1 **m** $\mid Cl_2$, Pt		0.0943	25Wil
Pt, $Cl_2 \mid$ HCl, 1.0 **m**, NaCl, 3.5 **m** $\Vert$ KCl, 1.0 **m**, HCl, 0.1 **m** $\mid Cl_2$, Pt		0.1068	25Wil
Pt, $Cl_2 \mid$ HCl, 2.0 **m**, NaCl, 0.0 **m** $\Vert$ KCl, 1.0 **m**, HCl, 0.1 **m** $\mid Cl_2$, Pt		0.1276	25Wil
Pt, $Cl_2 \mid$ HCl, 2.0 **m**, NaCl, 2.0 **m** $\Vert$ KCl, 1.0 **m**, HCl, 0.1 **m** $\mid Cl_2$, Pt		0.1458	25Wil
Pt, $Cl_2 \mid$ HCl, 4.0 **m**, NaCl, 0.0 **m** $\Vert$ KCl, 1.0 **m**, HCl, 0.1 **m** $\mid Cl_2$, Pt		0.1820	25Wil
Pt, $Cl_2 \mid$ HCl, 5.0 **m**, NaCl, 0.0 **m** $\Vert$ KCl, 1.0 **m**, HCl, 0.1 **m** $\mid Cl_2$, Pt		0.1978	25Wil
Pt, $Cl_2 \mid$ HCl, 5.0 **m**, NaCl, 1.25 **m** $\Vert$ KCl, 1.0 **m**, HCl, 0.1 **m** $\mid Cl_2$, Pt		0.2126	25Wil
Chromium (Cr)			
Cr $\mid Cr_2(SO_4)_3 \mid$ KCl,sat. $\Vert$ KCl, 0.1 **M**, $Hg_2Cl_2 \mid$ Hg	18 °C	0.867	26Gru
Cr $\mid CrSO_4$, 0.426 **M**, H_2SO_4, 0.0 **M** $\Vert$ KCl, 1.0 **M**, $Hg_2Cl_2 \mid$ Hg	18 °C	0.879	27Gru
Cr $\mid CrSO_4$, 0.168 **M**, H_2SO_4, 0.0 **M** $\Vert$ KCl, 1.0 **M**, $Hg_2Cl_2 \mid$ Hg	18 °C	0.884	27Gru
Cr $\mid CrSO_4$, 0.034 **M**, H_2SO_4, 0.001 **M** $\Vert$ KCl, 1.0 **M**, $Hg_2Cl_2 \mid$ Hg	18 °C	0.910	27Gru
Cobalt (Co)			
Co $\mid CoCl_2$, 0.05 **M** $\Vert$ KCl, 1.0 **M**, $Hg_2Cl_2 \mid$ Hg		0.625	04Lab
Co $\mid CoCl_2$, 0.5 **M** $\Vert$ KCl, 1.0 **M**, $Hg_2Cl_2 \mid$ Hg		0.590	04Lab
Co $\mid CoCl_2$, 0.01 **M** $\Vert$ KCl, 1.0 **M**, $Hg_2Cl_2 \mid$ Hg		0.582	20Lam
Co $\mid CoCl_2$, 0.02 **M** $\Vert$ KCl, 1.0 **M**, $Hg_2Cl_2 \mid$ Hg		0.570	20Lam
Co $\mid CoCl_2$, 0.05 **M** $\Vert$ KCl, 1.0 **M**, $Hg_2Cl_2 \mid$ Hg		0.556	20Lam
Co $\mid CoCl_2$, 1.0 **M** $\Vert$ KCl, 1.0 **M**, $Hg_2Cl_2 \mid$ Hg		0.539	20Lam
Co $\mid CoCl_2$, 2.0 **M** $\Vert$ KCl, 1.0 **M**, $Hg_2Cl_2 \mid$ Hg		0.525	20Lam
Co $\mid CoCl_2$, 0.0670 **m**, $Hg_2Cl_2 \mid$ Hg		0.657	34Har4
Co $\mid CoCl_2$, 0.0597 **m**, $Hg_2Cl_2 \mid$ Hg		0.659	34Har4
Co $\mid CoCl_2$, 0.0670 **m**, KCl, 1.0 **M**, $Hg_2Cl_2 \mid$ Hg		0.598	34Har4

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Cerium (Ce)			
Ce | $CeBr_3$,sat. ‖ NH_4NO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.74	29Mül2
$Ce_{1.55wt\%}Hg$ | $CeBr_3$,sat. ‖ NH_4NO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.94	29Mül2
Pt | $Ce(SO_4)_2$, 0.0104 **m**, $Ce_2(SO_4)_3$, 0.0052 **m**, H_2SO_4, 0.5 **m** ‖ H_2SO_4, 0.5 **m** | H_2 | Pt		1.4617	31Kun
Pt | $Ce(SO_4)_2$, 0.00102 **m**, $Ce_2(SO_4)_3$, 0.00051 **m**, H_2SO_4, 0.5 **m** ‖ H_2SO_4, 0.5 **m** | H_2 | Pt		1.4604	31Kun
Pt | $Ce(SO_4)_2$, 0.00790 **m**, $Ce_2(SO_4)_3$, 0.00395 **m**, H_2SO_4, 1.0 **m** ‖ H_2SO_4, 0.5 **m** | H_2 | Pt		1.4437	31Kun
Pt | $Ce(SO_4)_2$, 0.00099 **m**, $Ce_2(SO_4)_3$, 0.000495 **m**, H_2SO_4, 1.0 **m** ‖ H_2SO_4, 0.5 **m** | H_2 | Pt		1.4433	31Kun
Copper (Cu)			
Cu | CuCN, 0.01 **M**, KCN, 0.0025 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	18 °C	0.506	05Spi
Cu | CuCN, 0.01 **M**, KCN, 0.005 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	18 °C	0.635	05Spi
Cu | CuCN, 0.01 **M**, KCN, 0.01 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	18 °C	0.830	05Spi
Cu | CuCN, 0.01 **M**, KCN, 0.02 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg,	18 °C	0.973	05Spi
Cu | CuCN, 0.01 **M**, KCN, 0.05 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	18 °C	1.125	05Spi
Cu | CuCN, 0.01 **M**, KCN, 0.1 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	18 °C	1.226	05Spi
Cu | CuCN, 0.01 **M**, KCN, 0.5 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	18 °C	1.453	05Spi
Cu | CuCN, 0.025 **M**, KCN, 0.025 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	18 °C	0.600	05Spi
Cu | CuCN, 0.025 **M**, KCN, 0.05 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	18 °C	0.838	05Spi
Cu | CuCN, 0.025 **M**, KCN, 0.1 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	18 °C	1.138	05Spi
Cu | CuCN, 0.025 **M**, KCN, 0.25 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	18 °C	1.343	05Spi
Cu | CuCN, 0.025 **M**, KCN, 0.5 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	18 °C	1.441	05Spi
Cu | CuCN, 0.025 **M**, KCN, 0.1 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	18 °C	1.138	05Spi
Cu | CuCN, 0.1 **M**, KCN, 0.1 **M** ‖ KCl, 0.1 **M**, Hg_2CL_2 | Hg	18 °C	0.663	05Spi
Cu | CuCN, 0.1 **M**, KCN, 0.2 **M** ‖ KCl, 0.1 **M**, Hg_2CL_2 | Hg	18 °C	0.946	05Spi
Cu | CuCN, 0.1 **M**, KCN, 0.4 **M** ‖ KCl, 0.1 **M**, Hg_2CL_2 | Hg	18 °C	1.300	05Spi
Cu | CuCN, 0.1 **M**, KCN, 1.0 **M** ‖ KCl, 0.1 **M**, Hg_2CL_2 | Hg	18 °C	1.505	05Spi
Cu | $CuSO_4$, 0.1 **M** | $Cu_{10wt\%}Hg$	0 °C	0.00631	19Öho
	15 °C	0.00533	19Öho
	20 °C	0.00471	19Öho
	25 °C	0.00378	19Öho
	30 °C	0.00263	19Öho
Cu | Cu_2O, NaOH, 1.027 **m** ‖ NaOH, 1.027 **m**, HgO | Hg	0 °C	0.4587	27Ish
	15 °C	0.4602	27Ish
	25 °C	0.4617	27Ish
$Cu_{sat.}Hg$ | $CuSO_4$, 0.5 **m** | Cu	0 °C	0.00694	33Oku
	15 °C	0.00584	33Oku
	25°C	0.00510	33Oku
	35 °C	0.00436	33Oku
$Cu_{sat.}Hg$ | $CuSO_4$,sat., $CuSO_4$*5 H_2O,solid, Hg_2SO_4 | Hg	20 °C	0.3499	33Ish3
	25 °C	0.3467	33Ish3
	30 °C	0.3434	33Ish3
	35 °C	0.3401	33Ish3
$Cu_{sat.}Hg$ | $CuSO_4$, 0.00420 **m**, Hg_2SO_4 | Hg		0.4125	30Get
$Cu_{sat.}Hg$ | $CuSO_4$, 0.01049 **m**, Hg_2SO_4 | Hg		0.4107	30Get
$Cu_{sat.}Hg$ | $CuSO_4$, 0.02098 **m**, Hg_2SO_4 | Hg		0.4007	30Get

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
$Cu_{sat.}Hg \mid CuSO_4$, 0.04193 **m**, $Hg_2SO_4 \mid Hg$		0.3906	30Get
$Cu_{sat.}Hg \mid CuSO_4$, 0.10488 **m**, $Hg_2SO_4 \mid Hg$		0.3785	30Get
$Cu_{sat.}Hg \mid CuSO_4$, 0.20984 **m**, $Hg_2SO_4 \mid Hg$		0.3701	30Get
Hydrogen (H) [4])			
$H_2 \mid$ HBr, 1 **m** $\mid$ AgBr $\mid$ Ag	273 K	0.08168	37Har1
	278 K	0.07994	37Har1
	283 K	0.07804	37Har1
	288 K	0.07596	37Har1
	293 K	0.07379	37Har1
	298 K	0.07129	37Har1
	303 K	0.06874	37Har1
	308 K	0.06604	37Har1
	313 K	0.06302	37Har1
	318 K	0.05997	37Har1
	323 K	0.05668	37Har1
$H_2 \mid$ HBr, 1 **m** $\mid$ AgBr $\mid$ Ag	278 K	0.07989	36Owe
	283 K	0.07803	36Owe
	288 K	0.07599	36Owe
	293 K	0.07376	36Owe
	298 K	0.07136	36Owe
	303 K	0.06878	36Owe
	308 K	0.06602	36Owe
	313 K	0.06306	36Owe
Pt, $H_2 \mid$ HBr, 0.001 **m**, AgBr $\mid$ Ag	303 K	0.43118	36Har2
(U_T [V]=U_0+0.682·10^{-3}(T–303 K)+2.930·10^{-6}(T–303 K)^2, 273<T<333 K)			
Pt, $H_2 \mid$ HBr, 0.005 **m**, AgBr $\mid$ Ag	303 K	0.34910	36Har2
(U_T [V]=U_0+0.416·10^{-3}(T–303 K)+2.860·10^{-6}(T–303 K)^2, 273<T<333 K)			
Pt, $H_2 \mid$ HBr, 0.010 **m**, AgBr $\mid$ Ag	303 K	0.31421	36Har2
(U_T [V]=U_0+0.3032·10^{-3}(T–303 K)+2.818·10^{-6}(T–303 K)^2, 273<T<333 K)			
Pt, $H_2 \mid$ HBr, 0.020 **m**, AgBr $\mid$ Ag	303 K	0.27958	36Har2
(U_T [V]=U_0+0.1923·10^{-3}(T–303 K)+2.742·10^{-6}(T–303 K)^2, 273<T<333 K)			
Pt, $H_2 \mid$ HBr, 0.05 **m**, AgBr $\mid$ Ag	303 K	0.23427	36Har2
(U_T [V]=U_0+0.0482·10^{-3}(T–303 K)+2.739·10^{-6}(T–303 K)^2, 273<T<333 K)			
Pt, $H_2 \mid$ HBr, 0.1 **m**, AgBr $\mid$ Ag	303 K	0.20021	36Har2
(U_T [V]=U_0–0.0570·10^{-3}(T–303 K)+2.571·10^{-6}(T–303 K)^2, 273<T<333 K)			
Pt, $H_2 \mid$ HBr, 0.2 **m**, AgBr $\mid$ Ag	303 K	0.16550	36Har2
(U_T [V]=U_0–0.1625·10^{-3}(T–303 K)+2.620·10^{-6}(T–303 K)^2, 273<T<333 K)			
Pt, $H_2 \mid$ HBr, 0.5 **m**, AgBr $\mid$ Ag	303 K	0.11735	36Har2
(U_T [V]=U_0–0.3039·10^{-3}(T–303 K)+2.662·10^{-6}(T–303 K)^2, 273<T<333 K)			
Pt, $H_2 \mid$ HBr, 1.0 **m**, AgBr $\mid$ Ag	303 K	0.07612	36Har2
(U_T [V]=U_0–0.4167·10^{-3}(T–303 K)+2.595·10^{-6}(T–303 K)^2, 273<T<333 K)			
Pt, $H_2 \mid$ HBr, AgBr $\mid$ Ag	273 K	0.08127	36Har2
	278 K	0.07955	36Har2
	283 K	0.07770	36Har2
	288 K	0.07561	36Har2
	293 K	0.07340	36Har2
	298 K	0.07103	36Har2
	303 K	0.0685	36Har2
Pt, $H_2 \mid$ HBr, 0.01 **m**, KBr, 0.0 **m**, AgBr $\mid$ Ag		0.3136	33Har1

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | HBr, 0.01 **m**, KBr, 0.01 **m**, AgBr | Ag		0.2976	33Har1
Pt, H_2 | HBr, 0.01 **m**, KBr, 0.02 **m**, AgBr | Ag		0.2884	33Har1
Pt, H_2 | HBr, 0.01 **m**, KBr, 0.05 **m**, AgBr | Ag		0.2728	33Har1
Pt, H_2 | HBr, 0.01 **m**, KBr, 0.1 **m**, AgBr | Ag		0.2595	33Har1
Pt, H_2 | HBr, 0.01 **m**, KBr, 0.2 **m**, AgBr | Ag		0.2451	33Har1
Pt, H_2 | HBr, 0.01 **m**, KBr, 0.5 **m**, AgBr | Ag		0.2246	33Har1
Pt, H_2 | HBr, 0.01 **m**, KBr, 1.0 **m**, AgBr | Ag		0.2063	33Har1
Pt, H_2 | HBr, 0.01 **m**, KBr, 3.0 **m**, AgBr | Ag		0.1658	33Har1
Pt, H_2 | HBr, 0.01 **m**, LiBr, 0.0 **m**, AgBr | Ag		0.3134	33Van2
Pt, H_2 | HBr, 0.01 **m**, LiBr, 0.02511 **m**, AgBr | Ag		0.2840	33Van2
Pt, H_2 | HBr, 0.01 **m**, LiBr, 0.04827 **m**, AgBr | Ag		0.2726	33Van2
Pt, H_2 | HBr, 0.01 **m**, LiBr, 0.1004 **m**, AgBr | Ag		0.2577	33Van2
Pt, H_2 | HBr, 0.01 **m**, LiBr, 0.1994**m**, AgBr | Ag		0.2427	33Van2
Pt, H_2 | HBr, 0.01 **m**, LiBr, 0.5016 **m**, AgBr | Ag		0.2194	33Van2
Pt, H_2 | HBr, 0.01 **m**, LiBr, 0.9996 **m**, AgBr | Ag		0.1972	33Van2
Pt, H_2 | HBr, 0.01 **m**, LiBr, 1.9837 **m**, AgBr | Ag		0.1654	33Van2
Pt, H_2 | HBr, 0.01 **m**, LiBr, 2.4893 **m**, AgBr | Ag		0.1514	33Van2
Pt, H_2 | $B(OH)_3$, 5.466 m**m**, Na_3BO_3, 4.640 m**m**, NaCl, 4.785 m**m**, AgCl | Ag		0.9016	34Owe
Pt, H_2 | $B(OH)_3$, 9.688 m**m**, Na_3BO_3, 8.225 m**m**, NaCl, 8.481 m**m**, AgCl | Ag		0.8869	34Owe
Pt, H_2 | $B(OH)_3$, 26.890 m**m**, Na_3BO_3, 22.831 m**m**, NaCl, 23.54 m**m**, AgCl | Ag		0.8608	34Owe
Pt, H_2 | $B(OH)_3$, 65.33 m**m**, Na_3BO_3, 55.895 m**m**, NaCl, 57.63 m**m**, AgCl | Ag		0.8282	34Owe
Pt, H_2 | $B(OH)_3$, 12.066 m**m**, Na_3BO_3, 4.46 m**m**, NaCl, 3.79 m**m**, AgCl | Ag		0.8862	34Owe
Pt, H_2 | $B(OH)_3$, 22.096 m**m**, Na_3BO_3, 8.216 m**m**, NaCl, 6.94 m**m**, AgCl | Ag		0.8702	34Owe
Pt, H_2 | $B(OH)_3$, 53.043 m**m**, Na_3BO_3, 19.723 m**m**, NaCl, 16.66 m**m**, AgCl | Ag		0.8458	34Owe
Pt, H_2 | $B(OH)_3$, 109.876 m**m**, Na_3BO_3, 40.856 m**m**, NaCl, 34.51 m**m**, AgCl | Ag		0.8234	34Owe
Pt, H_2 | $HClO_4$ ‖ $HClO_4$, As_2O_3,sol. | As		0.234	24Sch2
Pt, H_2 | HCl, 0.000983 **M** ‖ HCl, 0.000983 **M**, BiOCl,sol. | Bi		0.2825	18Noy
Pt, H_2 | HCl, 0.009022 **M** ‖ HCl, 0.009022 **M**, BiOCl,sol. | Bi		0.2429	18Noy
Pt, H_2 | HCl, 0.1041 **M** ‖ HCl, 0.1041 **M**, BiOCl,sol. | Bi		0.2022	18Noy
Pt, H_2 | HCl, 0.5005 **M** ‖ HCl, 0.5005 **M**, BiOCl,sol. | Bi		0.1756	18Noy
Pt, H_2 | HCl, 1.0 **M** | Cl_2, C (U_T [V]=1.360–1.0·10^{-3}(T–20 °C))	20 °C	1.360	24Sch3
Pt, H_2 | $NaHCO_3$, 0.443 **M**, NaCl, 1.0 **M** ‖ NaCl, 0.95 **M**, $NaHCO_3$, 0.422 **M**, HOCl, 0.0084 **M** | Pt	18 °C	1.608	04Ner
Pt, H_2 | $NaHCO_3$, 0.0755 **M**, NaCl, 1.0 **M** ‖ NaCl, 1.0 **M**, $NaHCO_3$, 0.0755 **M**, HOCl, 0.150 **M** | Pt	15 °C	1.734	04Ner
Pt, H_2 | HCl, 0.01447 **m** ‖ HCl, 0.01447 **m**, CuCl | Cu		0.3472	18Noy
Pt, H_2 | HCl, 0.0434 **m** ‖ HCl, 0.0434 **m**, CuCl | Cu		0.2923	18Noy
Pt, H_2 | HCl, 0.09596 **m** ‖ HCl, 0.09596 **m**, CuCl | Cu		0.2520	18Noy
Pt, H_2 | HCl, 0.1876 **m** ‖ HCl, 0.1876 **m**, CuCl | Cu		0.1916	18Noy
Pt, H_2 | HCl, 0.3253 **m** ‖ HCl, 0.3253 **m**, CuCl | Cu		0.1580	18Noy

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | HBr, 0.1 **m**, CuBr | $Cu_{sat.}Hg$	20 °C	0.1732	34Ish1
	25 °C	0.1726	34Ish1
	30 °C	0.1721	34Ish1
Pt, H_2 | HI, 0.05104 **m**, CuI | $Cu_{sat.}Hg$	20 °C	0.01407	34Ish1
	25 °C	0.01295	34Ish1
	30 °C	0.01210	34Ish1
	35 °C	0.01127	34Ish1
Pt, H_2 | $Ba(OH)_2$, 0.05058 **m**, $BaCl_2$, 0.01067 **m**, AgCl | H_2, Pt		1.0883	32Har3
Pt, H_2 | $Ba(OH)_2$, 0.05043 **m**, $BaCl_2$, 0.02027 **m**, AgCl | H_2, Pt		1.0715	32Har3
Pt, H_2 | $Ba(OH)_2$, 0.05092 **m**, $BaCl_2$, 0.04988 **m**, AgCl | H_2, Pt		1.0483	32Har3
Pt, H_2 | $Ba(OH)_2$, 0.05088 **m**, $BaCl_2$, 0.1028 **m**, AgCl | H_2, Pt		1.0287	32Har3
Pt, H_2 | $Ba(OH)_2$, 0.05073 **m**, $BaCl_2$, 0.2053 **m**, AgCl | H_2, Pt		1.0090	32Har3
Pt, H_2 | $Ba(OH)_2$, 0.05020 **m**, $BaCl_2$, 0.5120 **m**, AgCl | H_2, Pt		0.9808	32Har3
Pt, H_2 | $Ba(OH)_2$, 0.05081 **m**, $BaCl_2$, 1.0079 **m**, AgCl | H_2, Pt		0.9554	32Har3
Pt, H_2 | $Ba(OH)_2$, 0.04920 **m**, $BaCl_2$, 1.3171 **m**, AgCl | H_2, Pt		0.9455	32Har3
Pt, H_2 | $B(OH)_3$, 209.277 m**m**, Na_3BO_3, 77.817 m**m**, NaCl, 65.73 m**m**, AgCl | Ag		0.8009	34Owe
Pt, H_2 | HCOOH, 6.54 m**m**, KHCOOH, 8.253 m**m**, KCl,7.056 m**m**, AgCl | Ag		0.5784	34Har2
Pt, H_2 | HCOOH, 11.76 m**m**, KHCOOH, 14.661 m**m**, KCl, 12.689 m**m**, AgCl | Ag		0.5629	34Har2
Pt, H_2 | HCOOH, 24.42 m**m**, KHCOOH, 30.444 m**m**, KCl, 26.349 m**m**, AgCl | Ag		0.5439	34Har2
Pt, H_2 | HCOOH, 35.75 m**m**, KHCOOH, 44.569 m**m**, KCl, 38.574 m**m**, AgCl | Ag		0.5343	34Har2
Pt, H_2 | HCOOH, 48.63 m**m**, KHCOOH, 60.627 m**m**, KCl, 52.471 m**m**, AgCl | Ag		0.5265	34Har2
Pt, H_2 | HCOOH, 98.76 m**m**, KHCOOH, 123.124 m**m**, KCl, 106.562 m**m**, AgCl | Ag		0.5088	34Har2
Pt, H_2 | CH_3COOH, 6.313 m**m**, CH_3COONa, 6.313 m**m**, NaCl, 6.313 **m**, AgCl | Ag	308.15 K	0.6414	33Har6
	333.15 K	0.6598	33Har6
Pt, H_2 | CH_3COOH, 12.87 m**m**, CH_3COONa, 12.87 m**m**, NaCl, 12.87 **m**, AgCl | Ag	308.15 K	0.6227	33Har6
	333.15 K	0.6394	33Har6
Pt, H_2 | CH_3COOH, 21.02 m**m**, CH_3COONa, 21.02 m**m**, NaCl, 21.02 **m**, AgCl | Ag	308.15 K	0.6097	33Har6
	333.15 K	0.6253	33Har6
Pt, H_2 | CH_3COOH, 47.97 m**m**, CH_3COONa, 47.97 m**m**, NaCl, 47.97 **m**, AgCl | Ag	308.15 K	0.5878	33Har6
	333.15 K	0.6015	33Har6
Pt, H_2 | CH_3COOH, 76.13 m**m**, CH_3COONa, 76.13 m**m**, NaCl, 76.13 **m**, AgCl | Ag	308.15 K	0.5755	33Har6
	333.15 K	0.5882	33Har6
Pt, H_2 | CH_3COOH, 100.3 m**m**, CH_3COONa, 100.3 m**m**, NaCl, 100.3 **m**, AgCl | Ag	308.15 K	0.5684	33Har6
	333.15 K	0.5804	33Har6
Pt, H_2 | CH_3COOH, 4.779 m**m**, CH_3COONa, 4.599 m**m**, NaCl, 4.896 **m**, AgCl | Ag		0.6396	32Har1
Pt, H_2 | CH_3COOH, 12.035 m**m**, CH_3COONa, 11.582 m**m**, NaCl, 12.426 **m**, AgCl | Ag		0.6158	32Har1
Pt, H_2 | CH_3COOH, 21.006 m**m**, CH_3COONa, 20.216 m**m**, NaCl, 21.516 **m**, AgCl | Ag		0.6015	32Har1
Pt, H_2 | CH_3COOH, 49.22 m**m**, CH_3COONa, 47.37 m**m**, NaCl, 50.42 **m**, AgCl | Ag		0.5798	32Har1

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | CH_3COOH, 81.01 m**m**, CH_3COONa, 77.96 m**m**, NaCl, 82.97 **m**, AgCl | Ag		0.5671	32Har1
Pt, H_2 | CH_3COOH, 90.56 m**m**, CH_3COONa, 87.16 m**m**, NaCl, 92.76 **m**, AgCl | Ag		0.5642	32Har1
Pt, H_2 | CH_3COOH, 0.1003 **m**, NaCl, 0.05 **m**, AgCl | Ag		0.4744	28Har1
Pt, H_2 | CH_3COOH, 0.1006 **m**, NaCl, 0.5 **m**, AgCl | Ag		0.4184	28Har1
Pt, H_2 | CH_3COOH, 0.1050 **m**, NaCl, 1.0 **m**, AgCl | Ag		0.3986	28Har1
Pt, H_2 | CH_3COOH, 0.1050 **m**, NaCl, 2.0 **m**, AgCl | Ag		0.3773	28Har1
Pt, H_2 | CH_3COOH, 0.2018 **m**, NaCl, 0.1 **m**, AgCl | Ag		0.4484	28Har1
Pt, H_2 | CH_3COOH, 0.2018 **m**, NaCl, 0.2 **m**, AgCl | Ag		0.4317	28Har1
Pt, H_2 | CH_3COOH, 0.2018 **m**, NaCl, 0.5 **m**, AgCl | Ag		0.4088	28Har1
Pt, H_2 | CH_3COOH, 0.2018 **m**, NaCl, 1.066 **m**, AgCl | Ag		0.3877	28Har1
Pt, H_2 | CH_3COOH, 0.2018 **m**, NaCl, 2.0 **m**, AgCl | Ag		0.3682	28Har1
Pt, H_2 | CH_3COOH, 0.2018 **m**, NaCl, 3.0 **m**, AgCl | Ag		0.3519	28Har1
Pt, H_2 | CH_3COOH, 1.002 **m**, NaCl, 0.01997 **m**, AgCl | Ag		0.4646	28Har1
Pt, H_2 | CH_3COOH, 1.002 **m**, NaCl, 0.049 **m**, AgCl | Ag		0.4429	28Har1
Pt, H_2 | CH_3COOH, 1.002 **m**, NaCl, 0.1 **m**, AgCl | Ag		0.4255	28Har1
Pt, H_2 | CH_3COOH, 1.002 **m**, NaCl, 0.2 **m**, AgCl | Ag		0.4085	28Har1
Pt, H_2 | CH_3COOH, 1.002 **m**, NaCl, 0.5 **m**, AgCl | Ag		0.3865	28Har1
Pt, H_2 | CH_3COOH, 1.002 **m**, NaCl, 1.0 **m**, AgCl | Ag		0.3672	28Har1
Pt, H_2 | CH_3COOH, 1.002 **m**, NaCl, 2.43 **m**, AgCl | Ag		0.3377	28Har1
Pt, H_2 | CH_3COOH, 0.2 **m**, NaCl, 0.05 **m**, AgCl | Ag		0.4648	31Har1
Pt, H_2 | CH_3COOH, 0.2 **m**, NaCl, 0.1 **m**, AgCl | Ag		0.4486	31Har1
Pt, H_2 | CH_3COOH, 0.2 **m**, NaCl, 0.2 **m**, AgCl | Ag		0.4319	31Har1
Pt, H_2 | CH_3COOH, 0.2 **m**, NaCl, 0.5 **m**, AgCl | Ag		0.4094	31Har1
Pt, H_2 | CH_3COOH, 0.2 **m**, NaCl, 1.0 **m**, AgCl | Ag		0.3902	31Har1
Pt, H_2 | CH_3COOH, 0.2 **m**, NaCl, 3.0 **m**, AgCl | Ag		0.3522	31Har1
Pt, H_2 | CH_3COOH, 0.2 **m**, KCl, 0.05 **m**, AgCl | Ag		0.4654	31Har1
Pt, H_2 | CH_3COOH, 0.2 **m**, KCl, 0.1 **m**, AgCl | Ag		0.4496	31Har1
Pt, H_2 | CH_3COOH, 0.2 **m**, KCl, 0.2 **m**, AgCl | Ag		0.4329	31Har1
Pt, H_2 | CH_3COOH, 0.2 **m**, KCl, 0.5 **m**, AgCl | Ag		0.4117	31Har1
Pt, H_2 | CH_3COOH, 0.2 **m**, KCl, 1.0 **m**, AgCl | Ag		0.3941	31Har1
Pt, H_2 | CH_3COOH, 0.2 **m**, KCl, 3.0 **m**, AgCl | Ag		0.3649	31Har1
Pt, H_2 | C_3H_7COOH, 5.64 m**m**, C_3H_7COONa, 6.41 m**m**, NaCl, 6.06 m**m**, AgCl | Ag	273.15 K	0.6202	34Har3
Pt, H_2 | C_3H_7COOH, 7.17 m**m**, C_3H_7COONa, 6.87 m**m**, NaCl, 7.06 m**m**, AgCl | Ag	298.15 K	0.6339	34Har3
	313.15 K	0.6465	34Har3
Pt,H_2 | C_3H_7COOH, 13.74 m**m**, C_3H_7COONa, 15.63 m**m**, NaCl, 14.77 m**m**, AgCl | Ag	273.15 K	0.5993	34Har3
Pt,H_2 | C_3H_7COOH, 15.15 m**m**, C_3H_7COONa, 14.53 m**m**, NaCl, 14.93 m**m**, AgCl | Ag	298.15 K	0.6150	34Har3
	313.15 K	0.6265	34Har3
Pt, H_2 | C_3H_7COOH, 22.32 m**m**, C_3H_7COONa, 21.39 m**m**, NaCl, 21.98 m**m**, AgCl | Ag	298.15 K	0.6051	34Har3
	313.15 K	0.6161	34Har3
Pt, H_2 | C_3H_7COOH, 23.81 m**m**, C_3H_7COONa, 27.07 m**m**, NaCl, 25.59 m**m**, AgCl | Ag	273.15 K	0.5867	34Har3
Pt, H_2 | C_3H_7COOH, 40.40 m**m**, C_3H_7COONa, 45.94 m**m**, NaCl, 43.43 m**m**, AgCl | Ag	273.15 K	0.5746	34Har3

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | C_3H_7COOH, 51.06 m**m**, C_3H_7COONa, 48.95 m**m**, NaCl, 50.29 m**m**, AgCl | Ag	298.15 K	0.5840	34Har3
	313.15 K	0.5941	34Har3
Pt, H ‖ quinone, hydroquinone, HCl, 0.01 **m**, AgCl | Ag	273.33 K	0.2599	33Har2
	283.15 K	0.2498	33Har2
	288.15 K	0.2449	33Har2
	293.15 K	0.2401	33Har2
	298.15 K	0.2355	33Har2
	303.15 K	0.2310	33Har2
	308.15 K	0.2267	33Har2
	313.15 K	0.22225	33Har2
U_T [V]=(0.2355–911.2·10^{-6}(T–298.15 K)+2.939·10^{-7}(T–298.15 K)), 273.15<T<313.15 K			33Har2
Pt, H_2 | Na_2CO_3, 0.05 **M**, $NaHCO_3$, 0.15 **M** | KCl, 0.1 **M**, HgCl | Ag	291.15 K	0.884	22Men
Pt, H_2 | Na_2CO_3, 0.1 **M**, $NaHCO_3$, 0.1 **M** | KCl, 0.1 **M**, HgCl | Ag	291.15 K	0.908	22Men
Pt, H_2 | Na_2CO_3, 0.18 **M**, $NaHCO_3$, 0.01 **M** | KCl, 0.1 **M**, HgCl | Ag	291.15 K	0.959	22Men
Pt, H_2 | Na_2CO_3, 0.01 **M**, $NaHCO_3$, 0.001 **M** | KCl, 0.1 **M**, HgCl | Ag	291.15 K	0.920	22Men
Pt, H_2 | Na_2CO_3, 0.016 **M**, $NaHCO_3$, 0.004 **M** | KCl, 0.1 **M**, HgCl | Ag	291.15 K	0.985	22Men
Pt, H_2 | NaH_2PO_4, 2.689 m**m**, Na_2HPO_4, 2.689 m**m**, NaCl, 2.689 m**m**, AgCl | Ag		0.7943	33Nim
Pt, H_2 | NaH_2PO_4, 5.468 m**m**, Na_2HPO_4, 5.468 m**m**, NaCl, 5.468 m**m**, AgCl | Ag		0.7738	33Nim
Pt, H_2 | NaH_2PO_4, 6.272 m**m**, Na_2HPO_4, 6.272 m**m**, NaCl, 6.272 m**m**, AgCl | Ag		0.7697	33Nim
Pt, H_2 | NaH_2PO_4, 8.75 m**m**, Na_2HPO_4, 8.75 m**m**, NaCl, 8.75 m**m**, AgCl | Ag		0.7599	33Nim
Pt, H_2 | NaH_2PO_4, 11.9 m**m**, Na_2HPO_4, 11.9 m**m**, NaCl, 11.9 m**m**, AgCl | Ag		0.7506	33Nim
Pt, H_2 | NaH_2PO_4, 20.87 m**m**, Na_2HPO_4, 20.87 m**m**, NaCl, 20.87 m**m**, AgCl | Ag		0.7333	33Nim
Pt, H_2 | NaH_2PO_4, 28.74 m**m**, Na_2HPO_4, 28.74 m**m**, NaCl, 28.74 m**m**, AgCl | Ag		0.7233	33Nim
Pt, H_2 | HCl, 0.0001 **m**, AgCl | Ag (U_T [V]=U_0+0.94533·10^{-3}(T–298 K) –3.26·10^{-6}(T–298 K)^2+8.89·10^{-9}(T–298 K)^3, 273<T<333 K)		0.69622	33Har3
Pt, H_2 | HCl, 0.0002 **m**, AgCl | Ag (U_T [V]=U_0+0.82711·10^{-3}(T–298 K) –3.27·10^{-6}(T–298 K)^2+9.37·10^{-9}(T–298 K)^3, 273<T<333 K)		0.66084	33Har3
Pt, H_2 | HCl, 0.0005 **m**, AgCl | Ag (U_T [V]=U_0+0.67216·10^{-3}(T–298 K) –3.23·10^{-6}(T–298 K)^2+8.37·10^{-9}(T–298 K)^3, 273<T<333 K)		0.61422	33Har3
Pt, H_2 | HCl, 0.001 **m**, AgCl | Ag (U_T [V]=U_0+0.55453·10^{-3}(T–298 K) –3.23·10^{-6}(T–298 K)^2+9.33·10^{-9}(T–298 K)^3, 273<T<333 K)		0.57912	33Har3
Pt, H_2 | HCl, 0.002 **m**, AgCl | Ag (U_T [V]=U_0+0.43844·10^{-3}(T–298 K) –3.26·10^{-6}(T–298 K)^2+9.67·10^{-9}(T–298 K)^3, 273<T<333 K)		0.54423	33Har3
Pt, H_2 | HCl, 0.005 **m**, AgCl | Ag (U_T [V]=U_0+0.28700·10^{-3}(T–298 K) –3.16·10^{-6}(T–298 K)^2+8.45 10^{-9}(T–298 K)^3, 273<T<333 K)		0.49843	33Har3
Pt, H_2 | HCl, 0.01 **m**, AgCl | Ag (U_T [V]=U_0+0.17384·10^{-3}(T–298 K) –3.11·10^{-6}(T–298 K)^2+7.00 10^{-9}(T–298 K)^3, 273<T<333 K)		0.46418	33Har3
Pt, H_2 | HCl, 0.02 **m**, AgCl | Ag (U_T [V]=U_0+0.06424·10^{-3}(T–298 K) –2.98·10^{-6}(T–298 K)^2+4.25 10^{-9}(T–298 K)^3, 273<T<333 K)		0.43023	33Har3
Pt, H_2 | HCl, 0.05 **m**, AgCl | Ag (U_T [V]=U_0–0.07742·10^{-3}(T–298 K) –3.01·10^{-6}(T–298 K)^2+6.05 10^{-9}(T–298 K)^3, 273<T<333 K)		0.38588	33Har3

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | HCl, 0.1 **m**, AgCl | Ag (U_T [V]=U_0–0.18157·10^{-3}(T–298 K) –2.90·10^{-6}(T–298 K)^2+1.79 10^{-9}(T–298 K)^3, 273<T<333 K)		0.35240	33Har3
Pt, H_2 | HCl, 0.2 **m**, AgCl | Ag (U_T [V]=U_0–0.28361·10^{-3}(T–298 K) –2.85·10^{-6}(T–298 K)^2+5.65 10^{-9}(T–298 K)^3, 273<T<333 K)		0.31874	33Har3
Pt, H_2 | HCl, 0.5 **m**, AgCl | Ag (U_T [V]=U_0–0.41705·10^{-3}(T–298 K) –2.82·10^{-6}(T–298 K)^2+5.25 10^{-9}(T–298 K)^3, 273<T<333 K)		0.27230	33Har3
Pt, H_2 | HCl, 1.0 **m**, AgCl | Ag (U_T [V]=U_0–0.51764·10^{-3}(T–298 K) –2.64·10^{-6}(T–298 K)^2+9.61 10^{-9}(T–298 K)^3, 273<T<333 K)		0.23328	33Har3
Pt, H_2 | HCl, 1.5 **m**, AgCl | Ag (U_T [V]=U_0–0.57285·10^{-3}(T–298 K) –2.44·10^{-6}(T–298 K)^2+7.15 10^{-9}(T–298 K)^3, 273<T<333 K)		0.20719	33Har3
Pt, H_2 | HCl, 2.0 **m**, AgCl | Ag (U_T [V]=U_0–0.60872·10^{-3}(T–298 K) –2.31·10^{-6}(T–298 K)^2–2.54 10^{-9}(T–298 K)^3, 273<T<333 K)		0.18634	33Har3
Pt, H_2 | HCl, 3.0 **m**, AgCl | Ag (U_T [V]=U_0–0.660·10^{-3}(T–298 K) –2.34·10^{-6}(T–298 K)^2–1.08 10^{-9}(T–298 K)^3, 273<T<333 K)		0.15184	33Har3
Pt, H_2 | HCl, 4.0 **m**, AgCl | Ag (U_T [V]=U_0–0.67076·10^{-3}(T–298 K) –3.81·10^{-6}(T–298 K)^2+2.46 10^{-9}(T–298 K)^3, 273<T<333 K)		0.12215	33Har3
Pt, H_2 | HCl, AgCl | Ag	0 °C	0.23647	52Swi
	5 °C	0.23406	52Swi
	10 °C	0.23145	52Swi
	15 °C	0.22865	52Swi
	18 °C	0.22689	52Swi
	20 °C	0.22568	52Swi
	25 °C	0.22254	52Swi
	30 °C	0.21924	52Swi
	35 °C	0.21578	52Swi
	37 °C	0.21432	52Swi
	40 °C	0.21216	52Swi
	45 °C	0.20841	52Swi
	50 °C	0.20452	52Swi
	55 °C	0.20050	52Swi
	60 °C	0.19635	52Swi
Pt, H_2 | HCl, AgCl | Ag	0 °C	0.23642	33Har3
	5 °C	0.23400	33Har3
	10 °C	0.23134	33Har3
	15 °C	0.22854	33Har3
	20 °C	0.22558	33Har3
	25 °C	0.22246	33Har3
	30 °C	0.21919	33Har3
	35 °C	0.21570	33Har3
	40 °C	0.21207	33Har3
	45 °C	0.20828	33Har3
	50 °C	0.20444	33Har3
	55 °C	0.20042	33Har3
	60 °C	0.19627	33Har3
Pt, H_2 | HCl, AgCl | Ag	0 °C	0.23655	54Bat2
	5 °C	0.23413	54Bat2
	10 °C	0.23142	54Bat2
	15 °C	0.22857	54Bat2

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
	20 °C	0.22557	54Bat2
	25 °C	0.22234	54Bat2
	30 °C	0.21904	54Bat2
	35 °C	0.21565	54Bat2
	40 °C	0.21208	54Bat2
	45 °C	0.20835	54Bat2
	50 °C	0.20449	54Bat2
	55 °C	0.20056	54Bat2
	60 °C	0.19649	54Bat2
	70 °C	0.18782	54Bat2
	80 °C	0.1787	54Bat2
	90 °C	0.1695	54Bat2
	95 °C	0.1651	54Bat2
Pt, H_2 | HCl, AgCl | Ag	0 °C	0.23652	53Har
	5 °C	0.23405	53Har
	10 °C	0.23137	53Har
	15 °C	0.22849	53Har
	20 °C	0.22549	53Har
	25 °C	0.22239	53Har
	30 °C	0.21908	53Har
	35 °C	0.21570	53Har
	40 °C	0.21207	53Har
	45 °C	0.20833	53Har
	50 °C	0.20449	53Har
Pt, H_2 | HCl, AgCl | Ag	60 °C	0.19676	60Gre
	90 °C	0.1696	60Gre
	125 °C	0.1330	60Gre
	150 °C	0.1032	60Gre
	175 °C	0.0708	60Gre
	200 °C	0.0348	60Gre
	225 °C	–0.0051	60Gre
	250 °C	–0.054	60Gre
	275 °C	–0.090	60Gre
Pt, H_2 | HCl, 0.05 **m**, $AlCl_3$, 0.0 **m**, AgCl | Ag		0.3863	31Har2
Pt, H_2 | HCl, 0.05 **m**, $AlCl_3$, 0.005 **m**, AgCl | Ag		0.3809	31Har2
Pt, H_2 | HCl, 0.05 **m**, $AlCl_3$, 0.01 **m**, AgCl | Ag		0.3769	31Har2
Pt, H_2 | HCl, 0.05 **m**, $AlCl_3$, 0.02 **m**, AgCl | Ag		0.3704	31Har2
Pt, H_2 | HCl, 0.05 **m**, $AlCl_3$, 0.05 **m**, AgCl | Ag		0.3574	31Har2
Pt, H_2 | HCl, 0.05 **m**, $AlCl_3$, 0.1 **m**, AgCl | Ag		0.3445	31Har2
Pt, H_2 | HCl, 0.05 **m**, $AlCl_3$, 0.45 **m**, AgCl | Ag		0.3010	31Har2
Pt, H_2 | HCl, 0.1 **m**, $AlCl_3$, 0.4 **m**, AgCl | Ag		0.2859	31Har2
Pt, H_2 | HCl, 0.2 **m**, $AlCl_3$, 0.3 **m**, AgCl | Ag		0.2742	31Har2
Pt, H_2 | HCl, 0.3 **m**, $AlCl_3$, 0.2 **m**, AgCl | Ag		0.2708	31Har2
Pt, H_2 | HCl, 0.45 **m**, $AlCl_3$, 0.05 **m**, AgCl | Ag		0.2712	31Har2
Pt, H_2 | HCl, 0.5 **m**, $AlCl_3$, 0.0 **m**, AgCl | Ag		0.3024	31Har2
Pt, H_2 | HCl, 0.01 **m**, $CeCl_3$, 0.0 **m**, AgCl | Ag		0.4647	34Mas
Pt, H_2 | HCl, 0.01 **m**, $CeCl_3$, 0.005 **m**, AgCl | Ag		0.4451	34Mas

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | HCl, 0.01 **m**, $CeCl_3$, 0.01 **m**, AgCl | Ag		0.4350	34Mas
Pt, H_2 | HCl, 0.01 **m**, $CeCl_3$, 0.05 **m**, AgCl | Ag		0.4054	34Mas
Pt, H_2 | HCl, 0.01 **m**, $CeCl_3$, 0.10 **m**, AgCl | Ag		0.3904	34Mas
Pt, H_2 | HCl, 0.01 **m**, $CeCl_3$, 0.25 **m**, AgCl | Ag		0.3700	34Mas
Pt, H_2 | HCl, 0.01 **m**, $CeCl_3$, 0.5 **m**, AgCl | Ag		0.3491	34Mas
Pt, H_2 | HCl, 0.01 **m**, $CeCl_3$, 1.0 **m**, AgCl | Ag		0.3210	34Mas
Pt, H_2 | HCl, 0.1 **m**, $CeCl_3$, 0.150 **m**, AgCl | Ag		0.3157	34Mas
Pt, H_2 | HCl, 0.3 **m**, $CeCl_3$, 0.1167 **m**, AgCl | Ag		0.2811	34Mas
Pt, H_2 | HCl, 0.5 **m**, $CeCl_3$, 0.0833 **m**, AgCl | Ag		0.2625	34Mas
Pt, H_2 | HCl, 0.7 **m**, $CeCl_3$, 0.050 **m**, AgCl | Ag		0.2490	34Mas
Pt, H_2 | HCl, 1.0 **m**, $CeCl_3$, 0.0 **m**, AgCl | Ag		0.2334	34Mas
Pt, H_2 | HCl, 0.02498 **m**, $HClO_4$, 0.02498 **m**, AgCl | Ag		0.4030	33Mur
Pt, H_2 | HCl, 0.0421 **m**, $HClO_4$, 0.04821 **m**, AgCl | Ag		0.3714	33Mur
Pt, H_2 | HCl, 0.1493 **m**, $HClO_4$, 0.1493 **m**, AgCl | Ag		0.3163	33Mur
Pt, H_2 | HCl, 0.2570 **m**, $HClO_4$, 0.2570 **m**, AgCl | Ag		0.2884	33Mur
Pt, H_2 | HCl, 0.4961 **m**, $HClO_4$, 0.4961 **m**, AgCl | Ag		0.2506	33Mur
Pt, H_2 | HCl, 0.00988 **m**, $HClO_4$, 0.08985 **m**, AgCl | Ag		0.4107	33Mur
Pt, H_2 | HCl, 0.04005 **m**, $HClO_4$, 0.3604 **m**, AgCl | Ag		0.3416	33Mur
Pt, H_2 | HCl, 0.07087 **m**, $HClO_4$, 0.6378 **m**, AgCl | Ag		0.3112	33Mur
Pt, H_2 | HCl, 0.09994 **m**, $HClO_4$, 0.8995 **m**, AgCl | Ag		0.2906	33Mur
Pt, H_2 | HCl, 0.02498 **m**, $HClO_4$, 0.02498 **m**, AgCl | Ag		0.4030	33Mur
Pt, H_2 | HCl, 3.842 m**m**, H_2SO_4, 4.143 m**m**, AgCl | Ag	278 K	0.4810	52Dav
	283 K	0.4829	52Dav
	288 K	0.4847	52Dav
Pt, H_2 | HCl, 5.736 m**m**, H_2SO_4, 7.784 m**m**, AgCl | Ag	278 K	0.4600	52Dav
	283 K	0.4618	52Dav
	288 K	0.4633	52Dav
Pt, H_2 | HCl, 7.085 m**m**, H_2SO_4, 7.641 m**m**, AgCl | Ag	278 K	0.4540	52Dav
	283 K	0.4555	52Dav
	288 K	0.4569	52Dav
Pt, H_2 | HCl, 13.982 m**m**, H_2SO_4, 13.254 m**m**, AgCl | Ag	278 K	0.4259	52Dav
	283 K	0.4268	52Dav
	288 K	0.4278	52Dav
Pt, H_2 | HCl, 2.831 m**m**, H_2SO_4, 3.34 m**m**, AgCl | Ag	293 K	0.4996	52Dav
	298 K	0.5015	52Dav
	303 K	0.5031	52Dav
	308 K	0.5047	52Dav
Pt, H_2 | HCl, 2.913 m**m**, H_2SO_4, 3.403 m**m**, AgCl | Ag	293 K	0.4984	52Dav
	298 K	0.5003	52Dav
	303 K	0.5020	52Dav
	308 K	0.5035	52Dav
Pt, H_2 | HCl, 3.575 m**m**, H_2SO_4, 4.209 m**m**, AgCl | Ag	293 K	0.4887	52Dav
	298 K	0.4904	52Dav
	303 K	0.4920	52Dav
	308 K	0.4934	52Dav
Pt, H_2 | HCl, 9.872 m**m**, H_2SO_4, 9.487 m**m**, AgCl | Ag	293 K	0.4865	52Dav
	298 K	0.4882	52Dav
	303 K	0.4897	52Dav

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | HCl, 0.01 **m**, HgCl | Ag	291 K	0.4631	22Har
	298 K	0.4643	22Har
	303 K	0.4651	22Har
Pt, H_2 | HCl, 0.1 **m**, HgCl | Ag	291 K	0.5478	22Har
	298 K	0.5788	22Har
	303 K	0.5814	22Har
Pt, H_2 | HCl, 0.01 **m**, KCl, 0.0 **m**, AgCl | Ag	273 K	0.4579	33Har5
	298 K	0.4642	33Har5
	333 K	0.4667	33Har5
Pt, H_2 | HCl, 0.01 **m**, KCl, 0.01 **m**, AgCl | Ag	273 K	0.4432	33Har5
	298 K	0.4482	33Har5
	333 K	0.4491	33Har5
Pt, H_2 | HCl, 0.01 **m**, KCl, 0.05 **m**, AgCl | Ag	273 K	0.4204	33Har5
	298 K	0.4235	33Har5
	333 K	0.4218	33Har5
Pt, H_2 | HCl, 0.01 **m**, KCl, 0.1 **m**, AgCl | Ag	273 K	0.4081	33Har5
	298 K	0.4101	33Har5
	333 K	0.4072	33Har5
Pt, H_2 | HCl, 0.01 **m**, KCl, 0.5 **m**, AgCl | Ag	273 K	0.3766	33Har5
	298 K	0.3760	33Har5
	333 K	0.3697	33Har5
Pt, H_2 | HCl, 0.01 **m**, KCl, 1.0 **m**, AgCl | Ag	273 K	0.3594	33Har5
	298 K	0.3574	33Har5
	333 K	0.3497	33Har5
Pt, H_2 | HCl, 0.01 **m**, KCl, 2.0 **m**, AgCl | Ag	273 K	0.3389	33Har5
	298 K	0.3356	33Har5
	333 K	0.3267	33Har5
Pt, H_2 | HCl, 0.01 **m**, KCl, 3.5 **m**, AgCl | Ag	273 K	0.3183	33Har5
	298 K	0.3140	33Har5
	333 K	0.3042	33Har5
Pt, H_2 | HCl, 0.1 **m**, KCl, 3.9 **m**, AgCl | Ag		0.2474	32Haw
Pt, H_2 | HCl, 0.5 **m**, KCl, 3.5 **m**, AgCl | Ag		0.2043	32Haw
Pt, H_2 | HCl, 1.0 **m**, KCl, 3.0 **m**, AgCl | Ag		0.1816	32Haw
Pt, H_2 | HCl, 2.0 **m**, KCl, 2.0 **m**, AgCl | Ag		0.1566	32Haw
Pt, H_2 | HCl, 4.0 **m**, KCl, 0.0 **m**, AgCl | Ag		0.1222	32Haw
Pt, H_2 | HCl, 5.0 **m**, KCl, 0.0 **m**, AgCl | Ag		0.0955	32Haw
Pt, H_2 | HCl, 6.037 m**m**, $La_2(SO_4)_3$, 0.5865 m**m**, AgCl | Ag	293 K	0.4899	52Jon
	298 K	0.4915	52Jon
	303 K	0.4929	52Jon
	308 K	0.4941	52Jon
Pt, H_2 | HCl, 5.961 m**m**, $La_2(SO_4)_3$, 0.7618 m**m**, AgCl | Ag	293 K	0.4915	52Jon
	298 K	0.4931	52Jon
	303 K	0.4945	52Jon
	308 K	0.4959	52Jon
Pt, H_2 | HCl, 6.036 m**m**, $La_2(SO_4)_3$, 0.9803 m**m**, AgCl | Ag	293 K	0.4910	52Jon
	298 K	0.4926	52Jon
	303 K	0.4940	52Jon
	308 K	0.4953	52Jon

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | HCl, 6.030 m**m**, $La_2(SO_4)_3$, 2.990 m**m**, AgCl | Ag	293 K	0.4945	52Jon
	298 K	0.4964	52Jon
	303 K	0.4981	52Jon
	308 K	0.4997	52Jon
Pt, H_2 | HCl, 0.01 **m**, LiCl, 0.01 **m**, AgCl | Ag	288 K	0.4464	33Har4
	293 K	0.4472	33Har4
	298 K	0.4478	33Har4
	303 K	0.4484	33Har4
	308 K	0.4488	33Har4
Pt, H_2 | HCl, 0.01 **m**, LiCl, 0.05 **m**, AgCl | Ag	288 K	0.4222	33Har4
	293 K	0.4226	33Har4
	298 K	0.4228	33Har4
	303 K	0.4230	33Har4
	308 K	0.4231	33Har4
Pt, H_2 | HCl, 0.01 **m**, LiCl, 0.1 **m**, AgCl | Ag	288 K	0.4088	33Har4
	293 K	0.4091	33Har4
	298 K	0.4092	33Har4
	303 K	0.4092	33Har4
	308 K	0.4092	33Har4
Pt, H_2 | HCl, 0.01 **m**, LiCl, 0.5 **m**, AgCl | Ag	288 K	0.3730	33Har4
	293 K	0.3727	33Har4
	298 K	0.3724	33Har4
	303 K	0.3719	33Har4
	308 K	0.3712	33Har4
Pt, H_2 | HCl, 0.01 **m**, LiCl, 1.0 **m**, AgCl | Ag	288 K	0.3529	33Har4
	293 K	0.3525	33Har4
	298 K	0.3519	33Har4
	303 K	0.3513	33Har4
	308 K	0.3505	33Har4
Pt, H_2 | HCl, 0.01 **m**, LiCl, 1.5 **m**, AgCl | Ag	288 K	0.3380	33Har4
	293 K	0.3375	33Har4
	298 K	0.3368	33Har4
	303 K	0.3360	33Har4
	308 K	0.3351	33Har4
Pt, H_2 | HCl, 0.01 **m**, LiCl, 2.0 **m**, AgCl | Ag	288 K	0.3249	33Har4
	293 K	0.3243	33Har4
	298 K	0.3235	33Har4
	303 K	0.3227	33Har4
	308 K	0.3217	33Har4
Pt, H_2 | HCl, 0.01 **m**, LiCl, 2.5 **m**, AgCl | Ag	288 K	0.3125	33Har4
	293 K	0.3119	33Har4
	298 K	0.3111	33Har4
	303 K	0.3102	33Har4
	308 K	0.3092	33Har4
Pt, H_2 | HCl, 0.01 **m**, LiCl, 3.0 **m**, AgCl | Ag	288 K	0.3011	33Har4
	293 K	0.3004	33Har4
	298 K	0.2996	33Har4
	303 K	0.2987	33Har4
	308 K	0.2976	33Har4

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | HCl, 0.01 **m**, LiCl, 3.5 **m**, AgCl | Ag	288 K	0.2902	33Har4
	293 K	0.2894	33Har4
	298 K	0.2885	33Har4
	303 K	0.2875	33Har4
	308 K	0.2864	33Har4
Pt, H_2 | HCl, 0.01 **m**, LiCl, 4.0 **m**, AgCl | Ag	288 K	0.2794	33Har4
	293 K	0.2785	33Har4
	298 K	0.2775	33Har4
	303 K	0.2765	33Har4
	308 K	0.2753	33Har4
Pt, H_2 | HCl, 0.01 **m**, LiCl, 4.5 **m**, AgCl | Ag	288 K	0.2681	33Har4
	293 K	0.2671	33Har4
	298 K	0.2660	33Har4
	303 K	0.2647	33Har4
	308 K	0.2634	33Har4
Pt, H_2 | HCl, 0.01 **m**, LiCl, 5.0 **m**, AgCl | Ag	288 K	0.2567	33Har4
	293 K	0.2555	33Har4
	298 K	0.2543	33Har4
	303 K	0.2530	33Har4
	308 K	0.2526	33Har4
Pt, H_2 | HCl, 0.1 **m**, LiCl, 3.9 **m**, AgCl | Ag		0.2147	32Haw
Pt, H_2 | HCl, 0.55 **m**, LiCl, 3.5 **m**, AgCl | Ag		0.1735	32Haw
Pt, H_2 | HCl, 1.0 **m**, LiCl, 3.0 **m**, AgCl | Ag		0.1578	32Haw
Pt, H_2 | HCl, 2.0 **m**, LiCl, 2.0 **m**, AgCl | Ag		0.1390	32Haw
Pt, H_2 | HCl, 3.0 **m**, LiCl, 1.0 **m**, AgCl | Ag		0.1289	32Haw
Pt, H_2 | HCl, 7.762 m**m**, $MgSO_4$, 3.795 m**m**, AgCl | Ag	293 K	0.4806	52Jon
	298 K	0.4821	52Jon
	303 K	0.4835	52Jon
	308 K	0.4837	52Jon
Pt, H_2 | HCl, 5.893 m**m**, $MgSO_4$, 4.596 m**m**, AgCl | Ag	293 K	0.4953	52Jon
	298 K	0.4971	52Jon
	303 K	0.4988	52Jon
	308 K	0.5004	52Jon
Pt, H_2 | HCl, 5.294 m**m**, $MgSO_4$, 5.762 m**m**, AgCl | Ag	293 K	0.5017	52Jon
	298 K	0.5037	52Jon
	303 K	0.5057	52Jon
	308 K	0.5073	52Jon
Pt, H_2 | HCl, 7.66 m**m**, $MgSO_4$, 6.933 m**m**, AgCl | Ag	293 K	0.4840	52Jon
	298 K	0.4857	52Jon
	303 K	0.4872	52Jon
	308 K	0.4886	52Jon
Pt, H_2 | HCl, 0.1 **m**, NaCl, 3.9 **m**, AgCl | Ag		0.2311	32Haw
Pt, H_2 | HCl, 0.5 **m**, NaCl, 3.5 **m**, AgCl | Ag		0.1884	32Haw
Pt, H_2 | HCl, 1.0 **m**, NaCl, 3.0 **m**, AgCl | Ag		0.1686	32Haw
Pt, H_2 | HCl, 2.0 **m**, NaCl, 2.0 **m**, AgCl | Ag		0.1472	32Haw
Pt, H_2 | HCl, 3.0 **m**, NaCl, 1.0 **m**, AgCl | Ag		0.1337	32Haw
Pt, H_2 | HCl, 0.01 **m**, NaCl, 0.05 **m**, AgCl | Ag		0.4236	31Har1

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | HCl, 0.01 **m**, NaCl, 0.1 **m**, AgCl | Ag		0.4103	31Har1
Pt, H_2 | HCl, 0.01 **m**, NaCl, 0.2**m**, AgCl | Ag		0.3955	31Har1
Pt, H_2 | HCl, 0.01 **m**, NaCl,0.5 **m**, AgCl | Ag		0.3746	31Har1
Pt, H_2 | HCl, 0.01 **m**, NaCl,1.0 **m**, AgCl | Ag		0.3547	31Har1
Pt, H_2 | HCl, 0.01 **m**, NaCl,3.0 **m**, AgCl | Ag		0.3087	31Har1
Pt, H_2 | HCl, 0.2480 **m**, $NaClO_4$, 0.0 **m**, AgCl | Ag		0.3078	33Bat
Pt, H_2 | HCl, 0.2480 **m**, $NaClO_4$, 0.1450 **m**, AgCl | Ag		0.3085	33Bat
Pt, H_2 | HCl, 0.2480 **m**, $NaClO_4$, 0.2459 **m**, AgCl | Ag		0.3088	33Bat
Pt, H_2 | HCl, 0.2480 **m**, $NaClO_4$, 0.7208 **m**, AgCl | Ag		0.3070	33Bat
Pt, H_2 | HCl, 0.1015 **m**, $NaClO_4$, 0.0 **m**, AgCl | Ag		0.3513	33Bat
Pt, H_2 | HCl, 0.1015 **m**, $NaClO_4$, 0.1684 **m**, AgCl | Ag		0.3543	33Bat
Pt, H_2 | HCl, 0.1015 **m**, $NaClO_4$, 0.5409 **m**, AgCl | Ag		0.3548	33Bat
Pt, H_2 | HCl, 0.1015 **m**, $NaClO_4$, 0.8902 **m**, AgCl | Ag		0.3532	33Bat
Pt, H_2 | HCl, 0.04639 **m**, $NaClO_4$, 0.0 **m**, AgCl | Ag		0.3899	33Bat
Pt, H_2 | HCl, 0.04639 **m**, $NaClO_4$, 0.1241 **m**, AgCl | Ag		0.3944	33Bat
Pt, H_2 | HCl, 0.04639 **m**, $NaClO_4$, 0.2633 **m**, AgCl | Ag		0.3956	33Bat
Pt, H_2 | HCl, 0.04639 **m**, $NaClO_4$, 0.4446 **m**, AgCl | Ag		0.3961	33Bat
Pt, H_2 | HCl, 0.04639 **m**, $NaClO_4$, 0.6230 **m**, AgCl | Ag		0.3955	33Bat
Pt, H_2 | HCl, 0.04639 **m**, $NaClO_4$, 0.8460 **m**, AgCl | Ag		0.3946	33Bat
Pt, H_2 | HCl, 0.1036 **m**, $Na_2S_2O_6$, 0.0 **m**, AgCl | Ag		0.3505	33Mur
Pt, H_2 | HCl, 0.1036 **m**, $Na_2S_2O_6$, 0.05 **m**, AgCl | Ag		0.3521	33Mur
Pt, H_2 | HCl, 0.1035 **m**, $Na_2S_2O_6$, 0.1 **m**, AgCl | Ag		0.3531	33Mur
Pt, H_2 | HCl, 0.1034 **m**, $Na_2S_2O_6$, 0.2 **m**, AgCl | Ag		0.3544	33Mur
Pt, H_2 | HCl, 0.1029 **m**, $Na_2S_2O_6$, 0.6 **m**, AgCl | Ag		0.3564	33Mur
Pt, H_2 | HCl, 0.1023 **m**, $Na_2S_2O_6$, 1.10 **m**, AgCl | Ag		0.3559	33Mur
Pt, H_2 | HCl, 0.05 **m**, $Na_2S_2O_6$, 0.0 **m**, AgCl | Ag		0.3861	33Mur
Pt, H_2 | HCl, 0.04985 **m**, $Na_2S_2O_6$, 0.25 **m**, AgCl | Ag		0.3923	33Mur
Pt, H_2 | HCl, 0.04963 **m**, $Na_2S_2O_6$, 0.65 **m**, AgCl | Ag		0.3944	33Mur
Pt, H_2 | HCl, 0.04919 **m**, $Na_2S_2O_6$, 1.4291 **m**, AgCl | Ag		0.3934	33Mur
Pt, H_2 | HCl, 0.01 **m**, $SrCl_2$, 0.0 **m**, AgCl | Ag		0.4645	33Van1
Pt, H_2 | HCl, 0.01 **m**, $SrCl_2$, 0.025 **m**, AgCl | Ag		0.4248	33Van1
Pt, H_2 | HCl, 0.01 **m**, $SrCl_2$, 0.05 **m**, AgCl | Ag		0.4116	33Van1
Pt, H_2 | HCl, 0.01 **m**, $SrCl_2$, 0.1 **m**, AgCl | Ag		0.3970	33Van1
Pt, H_2 | HCl, 0.01 **m**, $SrCl_2$, 0.2 **m**, AgCl | Ag		0.3816	33Van1
Pt, H_2 | HCl, 0.01 **m**, $SrCl_2$, 0.5 **m**, AgCl | Ag		0.3563	33Van1
Pt, H_2 | HCl, 0.01 **m**, $SrCl_2$, 1.0 **m**, AgCl | Ag		0.3290	33Van1
Pt, H_2 | HCl, 0.01 **m**, $SrCl_2$, 2.0 **m**, AgCl | Ag		0.2857	33Van1
Pt, H_2 | HCl, 0.01 **m**, $SrCl_2$, 2.75 **m**, AgCl | Ag		0.2551	33Van1
Pt, H_2 | HCl, 0.533 **m**, cane sugar, 0.0 **m**, AgCl | Ag		0.3158	35Kri
Pt, H_2 | HCl, 1.068 **m**, cane sugar, 0.0 **m**, AgCl | Ag		0.2759	35Kri
Pt, H_2 | HCl, 2.188 **m**, cane sugar, 0.0 **m**, AgCl | Ag		0.2258	35Kri
Pt, H_2 | HCl, 4.110 **m**, cane sugar, 0.0 **m**, AgCl | Ag		0.1640	35Kri
Pt, H_2 | HCl, 0.533 **m**, cane sugar, 0.292 **m**, AgCl | Ag		0.3136	35Kri
Pt, H_2 | HCl, 1.068 **m**, cane sugar, 0.292 **m**, AgCl | Ag		0.2719	35Kri
Pt, H_2 | HCl, 2.188 **m**, cane sugar, 0.292 **m**, AgCl | Ag		0.2217	35Kri
Pt, H_2 | HCl, 4.110 **m**, cane sugar, 0.292 **m**, AgCl | Ag		0.1601	35Kri
Pt, H_2 | HCl, AgCl, D-glucose, 5wt% | Ag (molar standard potential)		0.2186	50Wil

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | HCl, AgCl, D-glucose, 5wt% | Ag (molal standard potential)		0.2195	50Wil
Pt, H_2 | HCl, AgCl, D-glucose, 10wt% | Ag (molar standard potential)		0.2141	50Wil
Pt, H_2 | HCl, AgCl, D-glucose, 10wt% | Ag (molal standard potential)		0.2160	50Wil
Pt, H_2 | HCl, AgCl, D-glucose, 20wt% | Ag (molar standard potential)		0.2045	50Wil
Pt, H_2 | HCl, AgCl, D-glucose, 20wt% | Ag (molal standard potential)		0.2083	50Wil
Pt, H_2 | HCl, AgCl, D-glucose, 30wt% | Ag (molar standard potential)		0.1935	50Wil
Pt, H_2 | HCl, AgCl, D-glucose, 30wt% | Ag (molal standard potential)		0.1995	50Wil
Pt, H_2 | KOH, 0.01 **m**, KBr, 0.01 **m**, AgBrCl | Ag		0.8997	33Har1
Pt, H_2 | KOH, 0.01 **m**, KBr, 0.02 **m**, AgBrCl | Ag		0.8919	33Har1
Pt, H_2 | KOH, 0.01 **m**, KBr, 0.05 **m**, AgBrCl | Ag		0.8585	33Har1
Pt, H_2 | KOH, 0.01 **m**, KBr, 0.1 **m**, AgBrCl | Ag		0.8410	33Har1
Pt, H_2 | KOH, 0.01 **m**, KBr, 0.2 **m**, AgBrCl | Ag		0.8234	33Har1
Pt, H_2 | KOH, 0.01 **m**, KBr, 0.5 **m**, AgBrCl | Ag		0.8002	33Har1
Pt, H_2 | KOH, 0.01 **m**, KBr, 1.0 **m**, AgBrCl | Ag		0.7841	33Har1
Pt, H_2 | KOH, 0.01 **m**, KBr, 3.0 **m**, AgBrCl | Ag		0.7569	33Har1
Pt, H_2 | KOH, 0.01 **m**, KCl, 0.01 **m**, AgCl | Ag	273 K	1.0462	33Har5
	298 K	1.0503	33Har5
	333 K	1.0567	33Har5
Pt, H_2 | KOH, 0.02 **m**, KCl, 0.01 **m**, AgCl | Ag	273 K	1.0299	33Har5
	298 K	1.0326	33Har5
	333 K	1.0370	33Har5
Pt, H_2 | KOH, 0.05 **m**, KCl, 0.01 **m**, AgCl | Ag	273 K	1.0084	33Har5
	298 K	1.00092	33Har5
	333 K	1.0107	33Har5
Pt, H_2 | KOH, 0.1 **m**, KCl, 0.01 **m**, AgCl | Ag	273 K	0.9922	33Har5
	298 K	0.9916	33Har5
	333 K	0.9914	33Har5
Pt, H_2 | KOH, 0.2 **m**, KCl, 0.01 **m**, AgCl | Ag	273 K	0.9761	33Har5
	298 K	0.9742	33Har5
	333 K	0.9717	33Har5
Pt, H_2 | KOH, 0.5 **m**, KCl, 0.01 **m**, AgCl | Ag	273 K	0.9554	33Har5
	298 K	0.9517	33Har5
	333 K	0.9465	33Har5
Pt, H_2 | KOH, 1.0 **m**, KCl, 0.01 **m**, AgCl | Ag	273 K	0.9407	33Har5
	298 K	0.9356	33Har5
	333 K	0.9284	33Har5
Pt, H_2 | KOH, 2.0 **m**, KCl, 0.01 **m**, AgCl | Ag	273 K	0.9270	33Har5
	298 K	0.9207	33Har5
	333 K	0.9114	33Har5
Pt, H_2 | KOH, 3.5 **m**, KCl, 0.01 **m**, AgCl | Ag	273 K	0.9175	33Har5
	298 K	0.9101	33Har5
	333 K	0.8992	33Har5
Pt, H_2 | LiOH, 0.01 **m**, LiCl, 0.01 **m**, AgCl | Ag	288 K	1.0481	33Har4
	293 K	1.0490	33Har4
	298 K	1.0498	33Har4
	303 K	1.0506	33Har4
	308 K	1.0515	33Har4

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | LiOH, 0.01 **m**, LiCl, 0.02 **m**, AgCl | Ag	288 K	1.0073	33Har4
	293 K	1.0074	33Har4
	298 K	1.0076	33Har4
	303 K	1.0076	33Har4
	308 K	1.0077	33Har4
Pt, H_2 | LiOH, 0.01 **m**, LiCl, 0.1 **m**, AgCl | Ag	288 K	0.9891	33Har4
	293 K	0.9890	33Har4
	298 K	0.9888	33Har4
	303 K	0.9886	33Har4
	308 K	0.9884	33Har4
Pt, H_2 | LiOH, 0.01 **m**, LiCl ,0.5 **m**, AgCl | Ag	288 K	0.9443	33Har4
	293 K	0.9437	33Har4
	298 K	0.9428	33Har4
	303 K	0.9418	33Har4
	308 K	0.9409	33Har4
Pt, H_2 | LiOH, 0.01 **m**, LiCl, 1.0 **m**, AgCl | Ag	288 K	0.9224	33Har4
	293 K	0.9212	33Har4
	298 K	0.9199	33Har4
	303 K	0.9186	33Har4
	308 K	0.9173	33Har4
Pt, H_2 | LiOH, 0.01 **m**, LiCl, 1.5 **m**, AgCl | Ag	288 K	0.9074	33Har4
	293 K	0.9060	33Har4
	298 K	0.9045	33Har4
	303 K	0.9029	33Har4
	308 K	0.9014	33Har4
Pt, H_2 | LiOH, 0.01 **m**, LiCl, 2.0 **m**, AgCl | Ag	288 K	0.8953	33Har4
	293 K	0.8937	33Har4
	298 K	0.8920	33Har4
	303 K	0.8903	33Har4
	308 K	0.8886	33Har4
Pt, H_2 | LiOH, 0.01 **m**, LiCl, 2.5 **m**, AgCl | Ag	288 K	0.8847	33Har4
	293 K	0.8831	33Har4
	298 K	0.8813	33Har4
	303 K	0.8794	33Har4
	308 K	0.8776	33Har4
Pt, H_2 | LiOH, 0.01 **m**, LiCl, 3.0 **m**, AgCl | Ag	288 K	0.8752	33Har4
	293 K	0.8734	33Har4
	298 K	0.8715	33Har4
	303 K	0.8696	33Har4
	308 K	0.8676	33Har4
Pt, H_2 | LiOH, 0.01 **m**, LiCl, 3.5 **m**, AgCl | Ag	288 K	0.8665	33Har4
	293 K	0.8646	33Har4
	298 K	0.8625	33Har4
	303 K	0.8604	33Har4
	308 K	0.8583	33Har4

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | LiOH, 0.01 **m**, LiCl, 4.0 **m**, AgCl | Ag	288 K	0.8582	33Har4
	293 K	0.8562	33Har4
	298 K	0.8541	33Har4
	303 K	0.8519	33Har4
	308 K	0.8496	33Har4
Pt, H_2 | LiOH, 0.01 **m**, LiCl, 4.5 **m**, AgCl | Ag	288 K	0.8503	33Har4
	293 K	0.8482	33Har4
	298 K	0.8460	33Har4
	303 K	0.8437	33Har4
	308 K	0.8414	33Har4
Pt, H_2 | LiOH, 0.01 **m**, LiCl, 5.0 **m**, AgCl | Ag	288 K	0.8406	33Har4
	293 K	0.8383	33Har4
	298 K	0.8359	33Har4
	303 K	0.8334	33Har4
	308 K	0.8327	33Har4
Pt, H_2 | LiOH, 9.913 m**m**, LiBr, 20.02 m**m**, AgCl | Ag		0.8801	33Van2
Pt, H_2 | LiOH, 9.963 m**m**, LiBr, 49.99 m**m**, AgCl | Ag		0.8563	33Van2
Pt, H_2 | LiOH, 9.995 m**m**, LiBr, 99.95 m**m**, AgCl | Ag		0.8372	33Van2
Pt, H_2 | LiOH, 9.969 m**m**, LiBr, 197.6 m**m**, AgCl | Ag		0.8182	33Van2
Pt, H_2 | LiOH, 9.985 m**m**, LiBr, 499.4 m**m**, AgCl | Ag		0.7900	33Van2
Pt, H_2 | LiOH, 10.0 m**m**, LiBr, 1.0007 **m**, AgCl | Ag		0.7651	33Van2
Pt, H_2 | LiOH, 10.0 m**m**, LiBr, 1.4993 **m**, AgCl | Ag		0.7486	33Van2
H_2 | HI, 1 **m** | AgI | Ag	278 K	–0.14717	35Owe
	283 K	–0.14810	35Owe
	288 K	–0.14925	35Owe
	293 K	–0.15067	35Owe
	298 K	–0.15230	35Owe
	303 K	–0.15401	35Owe
	308 K	–0.15591	35Owe
	313 K	–0.15792	35Owe
Pt, H_2 | HI, 0.030 **m** | AgI | Ag		0.0368	20Noy
Pt, H_2 | HI, 0.0980 **m** | AgI | Ag		–0.0228	20Noy
Pt, H_2 | $H_2C_2O_4$, $Ag_2C_2O_4$,sat. | Ag	25 °C	0.4723	48Fer
	30 °C	0.4679	48Fer
	35 °C	0.4624	48Fer
Pt, H_2 | Na_2CO_3, 0.005821 **m**, $NaHCO_3$, 0.005843 **m**, NaCl, 0.005821 **m**, AgCl | Ag	273 K	0.9259	41Har
	288 K	0.9443	41Har
	298 K	0.9565	41Har
	308 K	0.9685	41Har
	323 K	0.9860	41Har
Pt, H_2 | Na_2CO_3, 0.011556 **m**, $NaHCO_3$, 0.011602 **m**, NaCl, 0.011556 **m**, AgCl | Ag	273 K	0.9076	41Har
	288 K	0.9251	41Har
	298 K	0.9370	41Har
	308 K	0.9488	41Har
	323 K	0.9665	41Har

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | Na_2CO_3, 0.02230 **m**, $NaHCO_3$, 0.02239 **m**, NaCl, 0.02230 **m**, AgCl | Ag	273 K	0.8853	41Har
	288 K	0.9019	41Har
	298 K	0.9130	41Har
	308 K	0.9254	41Har
	323 K	0.9419	41Har
Pt, H_2 | $NaHC_2O_4$, 0.8589 **m**, $Na_2C_2O_4$, 0.045 **m**, NaCl, 0.000657 **m**, AgCl | Ag	273 K	0.5618	39Har2
	298 K	0.5807	39Har2
	323 K	0.6005	39Har2
Pt, H_2 | $NaHC_2O_4$, 0.001857 **m**, $Na_2C_2O_4$, 0.008346 **m**, NaCl, 0.01214 **m**, AgCl | Ag	273 K	0.5447	39Har2
	298 K	0.5621	39Har2
	323 K	0.5802	39Har2
Pt, H_2 | $NaHC_2O_4$, 0.02003**m**, $Na_2C_2O_4$, 0.01053 **m**, NaCl, 0.001532 **m**, AgCl | Ag	273 K	0.5385	39Har2
	298 K	0.5551	39Har2
	323 K	0.5728	39Har2
Pt, H_2 | $NaHC_2O_4$, 0.0231143 **m**, $Na_2C_2O_4$, 0.01217 **m**, NaCl, 0.00177 **m**, AgCl | Ag	273 K	0.5346	39Har2
	298 K	0.5507	39Har2
	323 K	0.5643	39Har2
Pt, H_2 | $NaHC_2O_4$, 0.027776 **m**, $Na_2C_2O_4$, 0.01460 **m**, NaCl, 0.0212434 **m**, AgCl | Ag	273 K	0.5294	39Har2
	298 K	0.5451	39Har2
	323 K	0.5616	39Har2
Pt, H_2 | $NaHC_2O_4$, 0.03591 **m**, $Na_2C_2O_4$, 0.01889 **m**, NaCl, 0.002747 **m**, AgCl | Ag	273 K	0.5222	39Har2
	298 K	0.5372	39Har2
	323 K	0.5533	39Har2
Pt, H_2 | $NaHC_2O_4$, 0.04226 **m**, $Na_2C_2O_4$, 0.02222 **m**, NaCl, 0.003232 **m**, AgCl | Ag	273 K	0.5174	39Har2
	298 K	0.5321	39Har2
	323 K	0.5475	39Har2
Pt, H_2 | $NaHC_2O_4$, 0.04945 **m**, $Na_2C_2O_4$, 0.02600 **m**, NaCl, 0.003782 **m**, AgCl | Ag	273 K	0.5132	39Har2
	298 K	0.5321	39Har2
	323 K	0.5420	39Har2
Pt, H_2 | $NaHSO_4$, 0.9173 m**m**, Na_2SO_4, 0.9780 m**m**, NaCl, 0.9173 m**m**, AgCl | Ag		0.58622	34Ham
Pt, H_2 | $NaHSO_4$, 1.947 m**m**, Na_2SO_4, 1.957 m**m**, NaCl, 1.947 m**m**, AgCl | Ag		0.55388	34Ham
Pt, H_2 | $NaHSO_4$, 9.724 m**m**, Na_2SO_4, 9.776 m**m**, NaCl, 9.724 m**m**, AgCl | Ag		0.51260	34Ham
Pt, H_2 | $NaHSO_4$, 19.57 m**m**, Na_2SO_4, 19.68 m**m**, NaCl, 19.57 m**m**, AgCl | Ag		0.48356	34Ham
Pt, H_2 | $NaHSO_4$, 50.10 **mm**, Na_2SO_4, 50.54 **mm**, NaCl, 5.10 **mm**, AgCl | Ag		0.45420	34Ham
Pt, H_2 | $NaHSO_4$, 69.19 **mm**, Na_2SO_4, 69.56 **mm**, NaCl, 69.19 **mm**, AgCl | Ag		0.40388	34Ham
Pt, H_2 | $Sr(OH)_2$, 0.01369 **m**, $SrCl_2$, 0.02009 **m**, AgCl | Ag		1.0390	33Van1
Pt, H_2 | $Sr(OH)_2$, 0.01369 **m**, $SrCl_2$, 0.05010 **m**, AgCl | Ag		1.0131	33Van1
Pt, H_2 | $Sr(OH)_2$, 0.01329 **m**, $SrCl_2$, 0.09993 **m**, AgCl | Ag		0.9939	33Van1
Pt, H_2 | $Sr(OH)_2$, 0.01312 **m**, $SrCl_2$, 0.2013 **m**, AgCl | Ag		0.9730	33Van1
Pt, H_2 | $Sr(OH)_2$, 0.01338 **m**, $SrCl_2$, 0.5003 **m**, AgCl | Ag		0.9435	33Van1
Pt, H_2 | $Sr(OH)_2$, 0.01323 **m**, $SrCl_2$, 0.9988 **m**, AgCl | Ag		0.9154	33Van1
Pt, H_2 | $Sr(OH)_2$, 0.01371 **m**, $SrCl_2$, 1.9810 **m**, AgCl | Ag		0.8803	33Van1
Pt, H_2 | $Sr(OH)_2$, 0.01344 **m**, $SrCl_2$, 2.4855 **m**, AgCl | Ag		0.8641	33Van1

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | HCl, 0.01 **M** ‖ HCl, 0.01 **M**, quinone, hydroquinone | Au		0.69935	35Hov
Pt, H_2 | HCl, 0.04 **M** ‖ HCl, 0.04 **M**, quinone, hydroquinone | Au		0.69926	35Hov
Pt, H_2 | HCl, 0.1 **M** ‖ HCl, 0.1 **M**, quinone, hydroquinone | Au		0.69906	35Hov
Pt, H_2 | HCl, 0.5 **M** ‖ HCl, 0.5 **M**, quinone, hydroquinone | Au		0.69766	35Hov
Pt, H_2 | HCl, 1 **M** ‖ HCl, 1 **M**, quinone, hydroquinone | Au		0.69583	35Hov
Pt, H_2 | HCl, 2 **M** ‖ HCl, 2 **M**, quinone, hydroquinone | Au		0.69200	35Hov
Pt, H_2 | HCl, 0.01 **M**, LiCl, 1.0 **M**, quinone, hydroquinone | Au		0.6974	35Hov
Pt, H_2 | HCl, 0.01 **M**, NaCl, 1.0 **M**, quinone, hydroquinone | Au		0.6970	35Hov
Pt, H_2 | HCl, 0.01 **M**, KCl, 1.0 **M**, quinone, hydroquinone | Au		0.6972	35Hov
Pt, H_2 | HCl, 0.01 **M**, $MgCl_2$, 0.5 **M**, quinone, hydroquinone | Au		0.6974	35Hov
Pt, H_2 | HCl, 0.01 **M**, $CaCl_2$, 0.5 **M**, quinone, hydroquinone | Au		0.6974	35Hov
Pt, H_2 | HCl, 0.01 **M**, $SrCl_2$, 0.5 **M**, quinone, hydroquinone | Au		0.6973	35Hov
Pt, H_2 | HCl, 0.01 **M**, $BaCl_2$, 0.5 **M**, quinone, hydroquinone | Au		0.6969	35Hov
Pt, H_2 | HCl, 0.01 **M**, H_2SO_4, 0.5 **M**, quinone, hydroquinone | Au		0.6977	35Hov
Pt, H_2 | HCl, 0.01 **M**, Li_2SO_4, 0.5 **M**, quinone, hydroquinone | Au		0.7010	35Hov
Pt, H_2 | HCl, 0.01 **M**, Na_2SO_4, 0.5 **M**, quinone, hydroquinone | Au		0.7007	35Hov
Pt, H_2 | HCl, 0.01 **M**, K_2SO_4, 0.5 **M**, quinone, hydroquinone | Au		0.7008	35Hov
Pt, H_2 | HCl, 0.01 **M**, $MgSO_4$, 0.5 **M**, quinone, hydroquinone | Au		0.7007	35Hov
Pt, H_2 | $Ba(OH)_2$, 0.00245 **m** | Ba_xHg | $Ba(OH)_2$, 0.12 **m** | H_2, Pt		–0.1245	32Har2
Pt, H_2 | $Ba(OH)_2$, 0.0050 **m** | Ba_xHg | $Ba(OH)_2$, 0.12 **m** | H_2, Pt		–0.0976	32Har2
Pt, H_2 | $Ba(OH)_2$, 0.01132 **m** | Ba_xHg | $Ba(OH)_2$, 0.12 **m** | H_2, Pt		–0.0725	32Har2
Pt, H_2 | $Ba(OH)_2$, 0.02015 **m** | Ba_xHg | $Ba(OH)_2$, 0.12 **m** | H_2, Pt		–0.0540	32Har2
Pt, H_2 | $Ba(OH)_2$, 0.05323 **m** | Ba_xHg | $Ba(OH)_2$, 0.12 **m** | H_2, Pt		–0.0233	32Har2
Pt, H_2 | $Ba(OH)_2$, 0.07983 **m** | Ba_xHg | $Ba(OH)_2$, 0.12 **m** | H_2, Pt		–0.0121	32Har2
Pt, H_2 | $Ba(OH)_2$, 0.1055 **m** | Ba_xHg | $Ba(OH)_2$, 0.12 **m** | H_2, Pt		–0.0036	32Har2
Pt, H_2 | $Ba(OH)_2$, 0.2296 **m** | Ba_xHg | $Ba(OH)_2$, 0.12 **m** | H_2, Pt		0.0185	32Har2
Pt, H_2 | HCl, 0.1 **M** ‖ HCl, 0.01 **M** | H_2, Pt		0.0201	27Erd
Ir, H_2 | HCl, 0.1 **m** ‖ HCl, 0.01 **m** | H_2, Ir		0.0173	24Dru
Ir, H_2 | HCl, 0.02 **m** ‖ HCl, 0.01 **m** | H_2, Ir		0.0053	24Dru
Pt, H_2 | HCl, 0.1005 **M** ‖ $HClO_4$, 0.1 **M** | H_2, Pt		0.0006	24Sch2
Pt, H_2 | NaH-phthalat, 0.1 **M** ‖ NH_4Cl, 0.0 **M** ‖ $HClO_4$, 0.1 **M** | H_2, Pt		0.1493	31Mur
Pt, H_2 | NaH-phthalat, 0.1 **M** ‖ NH_4Cl, 0.1 **M** ‖ $HClO_4$, 0.1 **M** | H_2, Pt		0.1518	31Mur
Pt, H_2 | NaH-phthalat, 0.1 **M** ‖ NH_4Cl, 0.5 **M** ‖ $HClO_4$, 0.1 **M** | H_2, Pt		0.1625	31Mur
Pt, H_2 | NaH-phthalat, 0.1 **M** ‖ NH_4Cl, 1.0 **M** ‖ $HClO_4$, 0.1 **M** | H_2, Pt		0.1685	31Mur
Pt, H_2 | NaH-phthalat, 0.1 **M** ‖ NH_4Cl, 1.75 **M** ‖ $HClO_4$, 0.1 **M** | H_2, Pt		0.1719	31Mur
Pt, H_2 | NaH-phthalat, 0.1 **M** ‖ NH_4Cl, 2.5 **M** ‖ $HClO_4$, 0.1 **M** | H_2, Pt		0.1740	31Mur
Pt, H_2 | NaH-phthalat, 0.1 **M** ‖ NH_4Cl, 3.5 **M** ‖ $HClO_4$, 0.1 **M** | H_2, Pt		0.1763	31Mur
Pt, H_2 | NaH-phthalat, 0.1 **M** ‖ NH_4Cl, 4.1 **M** ‖ $HClO_4$, 0.1 **M** | H_2, Pt		0.1769	31Mur
Pt, H_2 | NaH-phthalat, 0.1 **M** ‖ NH_4Cl, 5.66 **M** [5]) ‖ $HClO_4$, 0.1 **M** | H_2, Pt		0.1773	31Mur
Pt, H_2 | NaOH, 0.25 **m** | $Na_{0.01wt\%}Hg$, NaOH, 0.05 **m** | H_2, Pt	273 K	0.06929	33Har7
	298 K	0.0760	33Har7
	308 K	0.0784	33Har7
U_T [V]=0.06929+285·10^{-6}(T–273 K)–0.73·10^{-6}(T–273 K)^2, 273<T<308 K			33Har7
Pt, H_2 | NaOH, 0.5 **m** | $Na_{0.01wt\%}Hg$, NaOH, 0.05 **m** | H_2, Pt	273 K	0.10018	33Har7
	298 K	0.1102	33Har7
	308 K	0.1139	33Har7
U_T [V]=0.10018+422·10^{-6}(T–273 K)–0.90·10^{-6}(T–273 K)^2, 273<T<308K			33Har7

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | NaOH, 1.0 **m** | $Na_{0.01wt\%}Hg$, NaOH, 0.05 **m** | H_2, Pt	273 K	0.13156	33Har7
	298 K	0.1452	33Har7
	308 K	0.1501	33Har7
U_T [V]=0.13156+585·10^{-6}(T–273 K)–1.59·10^{-6}(T–273 K)^2, 273<T<308 K			33Har7
Pt, H_2 | NaOH, 1.5 **m** | $Na_{0.01wt\%}Hg$, NaOH, 0.05 **m** | H_2, Pt	273 K	0.1511	33Har7
U_T [V]=0.1511+672·10^{-6}(T–273 K)–1.79·10^{-6}(T–273 K)^2, 273<T<308 K			
Pt, H_2 | NaOH, 2.0 **m** | $Na_{0.01wt\%}Hg$, NaOH, 0.05 **m** | H_2, Pt	273 K	0.16600	33Har7
	298 K	0.1832	33Har7
	308 K	0.1894	33Har7
U_T [V]=0.1660+736·10^{-6}(T–273 K)–1.91·10^{-6}(T–273 K)^2, 273<T<308 K			33Har7
Pt, H_2 | NaOH, 2.5 **m** | $Na_{0.01wt\%}Hg$, NaOH, 0.05 **m** | H_2, Pt	273 K	0.1785	33Har7
U_T [V]=0.1785+829·10^{-6}(T–273 K)–2.97·10^{-6}(T–273 K)^2, 273<T<308 K			
Pt, H_2 | NaOH, 3.0 **m** | $Na_{0.01wt\%}Hg$, NaOH, 0.05 **m** | H_2, Pt	273 K	0.1903	33Har7
U_T [V]=0.1903+895·10^{-6}(T–273 K)–2.52·10^{-6}(T–273 K)^2, 273<T<308 K			
Pt, H_2 | NaOH, 3.5 **m** | $Na_{0.01wt\%}Hg$, NaOH, 0.05 **m** | H_2, Pt	273 K	0.2017	33Har7
U_T [V]=0.2017+898·10^{-6}(T–273 K)–2.82·10^{-6}(T–273 K)^2, 273<T<308 K			
Pt, H_2 | NaOH, 4.0 **m** | $Na_{0.01wt\%}Hg$, NaOH, 0.05 **m** | H_2, Pt	273 K	0.2126	33Har7
	298 K	0.2338	33Har7
	308 K	0.2413	33Har7
U_T [V]=0.2126+915·10^{-6}(T–273 K)–2.76·10^{-6}(T–273 K)^2, 273<T 308 K			33Har7
Pt, H_2 | NaOH, 0.108 **m** | Na_xHg | NaCl, 4.5 **m**, NaOH, 0.5 **m** | H_2, Pt		0.1385	28Har2
Pt, H_2 | NaOH, 0.108 **m** | Na_xHg | NaCl, 4.0 **m**, NaOH, 1.0 **m** | H_2, Pt		0.1586	28Har2
Pt, H_2 | NaOH, 0.108 **m** | Na_xHg | NaCl, 3.5 **m**, NaOH, 1.5 **m** | H_2, Pt		0.1700	28Har2
Pt, H_2 | NaOH, 0.108 **m** | Na_xHg | NaCl, 3.0 **m**, NaOH, 2.0 **m** | H_2, Pt		0.1795	28Har2
Pt, H_2 | NaOH, 0.108 **m** | Na_xHg | NaCl, 2.5 **m**, NaOH, 2.5 **m** | H_2, Pt		0.1876	28Har2
Pt, H_2 | NaOH, 0.108 **m** | Na_xHg | NaCl, 2.0 **m**, NaOH, 3.0 **m** | H_2, Pt		0.1943	28Har2
Pt, H_2 | NaOH, 0.108 **m** | Na_xHg | NaCl, 1.5 **m**, NaOH, 3.5 **m** | H_2, Pt		0.2003	28Har2
Pt, H_2 | NaOH, 0.108 **m** | Na_xHg | NaCl, 1.0 **m**, NaOH, 4.0 **m** | H_2, Pt		0.2071	28Har2
Pt, H_2 | NaOH, 0.108 **m** | Na_xHg | NaCl, 0.5 **m**, NaOH, 4.5 **m** | H_2, Pt		0.2137	28Har2
Pt, H_2 | NaOH, 0.108 **m** | Na_xHg | NaCl, 0.0 **m**, NaOH, 5.0 **m** | H_2, Pt		0.2184	28Har2
Pt, H_2 | NaOH, 0.1349 **m** | Na_xHg | NaCl, 0.0 **m**, NaOH, 1.0 **m** | H_2, Pt		0.0991	37Har3
Pt, H_2 | NaOH, 0.1349 **m** | Na_xHg | NaCl, 0.0826 **m**, NaOH, 0.9174 **m** | H_2, Pt		0.0961	37Har3
Pt, H_2 | NaOH, 0.1349 **m** | Na_xHg | NaCl, 0.4298 **m**, NaOH, 0.5702 **m** | H_2, Pt		0.0819	37Har3
Pt, H_2 | NaOH, 0.1349 **m** | Na_xHg | NaCl, 0.8146 **m**, NaOH, 0.1854 **m** | H_2, Pt		0.0530	37Har3
Pt, H_2 | NaOH, 0.1349 **m** | Na_xHg | NaCl, 0.8968 **m**, NaOH, 0.1032 **m** | H_2, Pt		0.0378	37Har3
Pt, H_2 | NaOH, 0.1349 **m** | Na_xHg | NaCl, 0.0 **m**, NaOH, 0.5 **m** | H_2, Pt		0.0640	37Har3
Pt, H_2 | NaOH, 0.1349 **m** | Na_xHg | NaCl, 0.2782 **m**, NaOH, 0.2218 **m** | H_2, Pt		0.0415	37Har3
Pt, H_2 | NaOH, 0.1349 **m** | Na_xHg | NaCl, 0.4115 **m**, NaOH, 0.0885 **m** | H_2, Pt		0.0176	37Har3
Pt, H_2 | KOH, 0.01080 **m** | HgK_x | HgK_x | KCl, 0.0 **m**, KOH, 3.5 **m** | H_2, Pt		0.2042	28Har2
Pt, H_2 | KOH, 0.01080 **m** | HgK_x | HgK_x | KCl, 0.5 **m**, KOH, 3.0 **m** | H_2, Pt		0.1950	28Har2
Pt, H_2 | KOH, 0.01080 **m** | HgK_x | HgK_x | KCl, 1.0 **m**, KOH, 2.5 **m** | H_2, Pt		0.1861	28Har2
Pt, H_2 | KOH, 0.01080 **m** | HgK_x | HgK_x | KCl, 1.5 **m**, KOH, 2.0 **m** | H_2, Pt		0.1764	28Har2
Pt, H_2 | KOH, 0.01080 **m** | HgK_x | HgK_x | KCl, 2.0 **m**, KOH, 1.5 **m** | H_2, Pt		0.1644	28Har2
Pt, H_2 | KOH, 0.01080 **m** | HgK_x | HgK_x | KCl, 2.5 **m**, KOH, 1.0 **m** | H_2, Pt		0.1497	28Har2
Pt, H_2 | KOH, 0.01080 **m** | HgK_x | HgK_x | KCl, 3.0 **m**, KOH, 0.5 **m** | H_2, Pt		0.1282	28Har2
Pt, H_2 | CsOH, 0.01016 **m** | Cs_xHg | CsOH, 0.05 **m** | H_2, Pt		–0.0764	30Har1
Pt, H_2 | CsOH, 0.02550 **m** | Cs_xHg | CsOH, 0.05 **m** | H_2, Pt		–0.0431	30Har1
Pt, H_2 | CsOH, 0.02044 **m** | Cs_xHg | CsOH, 0.05 **m** | H_2, Pt		–0.03246	30Har1

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | CsOH, 0.05 **m** | Cs_xHg | CsOH, 0.05 **m** | H_2, Pt		0.0000	30Har1
Pt, H_2 | CsOH, 0.09065 **m** | Cs_xHg | CsOH, 0.05 **m** | H_2, Pt		0.02888	30Har1
Pt, H_2 | CsOH, 0.232 **m** | Cs_xHg | CsOH, 0.05 **m** | H_2, Pt		0.07379	30Har1
Pt, H_2 | CsOH, 0.5402 **m** | Cs_xHg | CsOH, 0.05 **m** | H_2, Pt		0.1168	30Har1
Pt, H_2 | CsOH, 1.3205 **m** | Cs_xHg | CsOH, 0.05 **m** | H_2, Pt		0.1687	30Har1
Pt, H_2 | CsOH, 0.05 **m**, CsCl, 0.09563 **m** | Cs_xHg | CsOH, 0.05 **m** | H_2, Pt		0.0224	30Har2
Pt, H_2 | CsOH, 0.05 **m**, CsCl, 0.3511 **m** | Cs_xHg | CsOH, 0.05 **m** | H_2, Pt		0.0435	30Har2
Pt, H_2 | CsOH, 0.05 **m**, CsCl, 0.9411 **m** | Cs_xHg | CsOH, 0.05 **m** | H_2, Pt		0.0634	30Har2
Pt, H_2 | CsOH, 0.05 **m**, CsCl, 1.7077 **m** | Cs_xHg | CsOH, 0.05 **m** | H_2, Pt		0.0779	30Har2
Pt, H_2 | CsOH, 0.05 **m**, CsCl, 3.9896 **m** | Cs_xHg | CsOH, 0.05 **m** | H_2, Pt		0.1025	30Har2
Pt, H_2 | KOH, 0.2240 **m** | K_xHg | KOH, 0.09154 **m** | H_2, Pt	273 K	0.0397	34Ham
	293 K	0.0423	34Ham
	313 K	0.046	34Ham
	333 K	0.0482	34Ham
Pt, H_2 | KOH, 0.9956 **m** | K_xHg | KOH, 0.09154 **m** | H_2, Pt	273 K	0.1104	34Ham
	293 K	0.1183	34Ham
	313 K	0.1256	34Ham
	333 K	0.1325	34Ham
Pt, H_2 | KOH, 4.0565 **m** | K_xHg | KOH, 0.09154 **m** | H_2, Pt	273 K	0.2107	34Ham
	293 K	0.2239	34Ham
	313 K	0.2380	34Ham
	333 K	0.2471	34Ham
Pt, H_2 | KOH, 10.052 **m** | K_xHg | KOH, 0.09154 **m** | H_2, Pt	273 K	0.3483	34Ham
	293 K	0.3637	34Ham
	313 K	0.3769	34Ham
	333 K	0.3867	34Ham
Pt, H_2 | KOH, 17.544 **m** | K_xHg | KOH, 0.09154 **m** | H_2, Pt	273 K	0.4948	34Ham
	293 K	0.5085	34Ham
	313 K	0.5224	34Ham
	333 K	0.5308	34Ham
Pt, H_2 | CH_3COOH, 0.4992 **M**, CH_3COOK, 0.4992 **M**, $(CH_3COO)_2Hg_2$,sol. | Hg		0.8100	39Lar
Pt, H_2 | CH_3COOH, 0.5087 **M**, CH_3COOK, 0.5087 **M**, $(CH_3COO)_2Hg_2$,sol. | Hg		0.8096	39Lar
Pt, H_2 | CH_3COOH, 0.5170 **M**, CH_3COOK, 0.5170 **M**, $(CH_3COO)_2Hg_2$,sol. | Hg		0.8092	39Lar
Pt, H_2 | CH_3COOH, 0.7433 **M**, CH_3COOK, 0.7433 **M**, $(CH_3COO)_2Hg_2$,sol. | Hg		0.7999	39Lar
Pt, H_2 | CH_3COOH, 1.035 **M**, CH_3COOK, 1.035 **M**, $(CH_3COO)_2Hg_2$,sol. | Hg, U_{00} [V]=0.5109		0.7914	39Lar
Pt [6]), H_2 | HBr, 0.1012 **M**, Hg_2Br_2 | Hg		0.2698	33Ish2
Pt, H_2 | HCl, 0.01006 **M**, Hg_2Cl_2 | Hg		0.5108	17Lin
Pt, H_2 | HCl, 0.1009 **M**, Hg_2Cl_2 | Hg		0.3990	17Lin
Pt, H_2 | HCl, 1.029 **M**, Hg_2Cl_2 | Hg		0.2780	17Lin
Pt, H_2 | HCl, 2.25 **M**, Hg_2Cl_2 | Hg		0.2231	17Lin
Pt, H_2 | HCl, 6.75 **M**, Hg_2Cl_2 | Hg		0.0972	17Lin
Pt, H_2 | HCl, 10.0 **M**, Hg_2Cl_2 | Hg		0.0302	17Lin
Pt, H_2 | HCl, 16.0 **M**, Hg_2Cl_2 | Hg		–0.0662 [7])	17Lin

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2, p=1 atm | HCl, 0.1 **M** ‖ HCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.3990	24Hai
Pt, H_2, p=110.2 atm | HCl,0.1 **M** ‖ HCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.4596	24Hai
Pt, H_2, p=556.8 atm | HCl,0.1 **M** ‖ HCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.4844	24Hai
Pt, H_2, p=1035.2 atm | HCl,0.1 **M** ‖ HCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.4975	24Hai
Pt, H_2 | $HClO_4$, 0.0817 **m** ‖ $HClO_4$, 0.0817 **m**, $HgClO_4$, 0.00550 **m** | Hg		0.7777	16Lin
Pt, H_2 | $HClO_4$, 0.0817 **m** ‖ $HClO_4$, 0.0817 **m**, $HgClO_4$, 0.0010 **m** | Hg		0.7569	16Lin
Pt, H_2 | $HClO_4$, 0.0817 **m** ‖ $HClO_4$, 0.0817 **m**, $HgClO_4$, 0.00055 **m** | Hg		0.7480	16Lin
Pt, H_2 | HF, 0.03 **M** | KCl,sat., HgCl | Hg		0.3820	24Wyn
Pt, H_2 | HF, 0.5 **M** | KCl,sat., HgCl | Hg		0.33318	24Wyn
Pt, H_2 | HF, 1 **M** | KCl,sat., HgCl | Hg		0.3120	24Wyn
Pt, H_2 | KOH, 15wt%, HgO | Hg	285 K	0.93	19Hof
Pt, H_2 | KCl, 1.49 **M**, KOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.01 **M**, HgCl | Hg	273 K	1.0286	29Bje
	298 K	1.0361	29Bje
Pt, H_2 | KCl, 0.99 **M**, KOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.01 **M**, HgCl | Hg	273 K	1.0279	29Bje
	298 K	1.0350	29Bje
Pt H_2 | KCl, 0.29 **M**, KOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.01 **M**, HgCl | Hg	273 K	1.0200	29Bje
	298 K	1.0358	29Bje
Pt, H_2 | KCl, 0.00 **M**, KOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.001 **M**, HgCl | Hg	298 K	1.0955	29Bje
Pt, H_2 | KCl, 0.09 **M**, KOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.01 **M**, HgCl | Hg	273 K	1.0305	29Bje
	298 K	1.0375	29Bje
Pt, H_2 | KCl, 0.099 **M**, KOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.001 **M**, HgCl | Hg	273 K	0.9768	29Bje
	298 K	0.9790	29Bje
Pt, H_2 | KCl, 0.029 **M**, KOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.001 **M**, HgCl | Hg	273 K	0.9787	29Bje
	298 K	0.9803	29Bje
Pt, H_2 | KCl, 0.02 **M**, KOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.01 **M**, HgCl | Hg	273 K	1.0323	29Bje
	298 K	1.0398	29Bje
Pt, H_2 | KCl, 0.0 **M**, KOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.01 **M**, HgCl | Hg	273 K	1.0333	29Bje
	298 K	1.0433	29Bje
Pt, H_2 | KCl, 0.009 **M**, KOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.001 **M**, HgCl | Hg	273 K	0.9807	29Bje
	298 K	0.9836	29Bje
Pt, H_2 | KCl, 0.002 **M**, KOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.001 **M**, HgCl | Hg	273 K	0.9808	29Bje
	298 K	0.9838	29Bje
Pt, H_2 | KCl, 0.00 **M**, KOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.001 **M**, HgCl | Hg	273 K	0.9810	29Bje
	298 K	0.9839	29Bje
Pt, H_2 | $Ca(OH)_2$, 0.05 g dm^{-3} ‖ KCl,sat., Hg_2Cl_2 | Hg		0.9062	33Fli
Pt, H_2 | $Ca(OH)_2$, 0.1024 g dm^{-3} ‖ KCl,sat., Hg_2Cl_2 | Hg		0.9235	33Fli
Pt, H_2 | $Ca(OH)_2$, 0.2004 g dm^{-3} ‖ KCl,sat., Hg_2Cl_2 | Hg		0.9393	33Fli
Pt, H_2 | $Ca(OH)_2$, 0.4002 g dm^{-3} ‖ KCl,sat., Hg_2Cl_2 | Hg		0.9588	33Fli
Pt, H_2 | $Ca(OH)_2$, 0.8004 g dm^{-3} ‖ KCl,sat., Hg_2Cl_2 | Hg		0.9720	33Fli
Pt, H_2 | $Ca(OH)_2$, 1.199 g dm^{-3} ‖ KCl,sat., Hg_2Cl_2 | Hg		0.9810	33Fli
Pt, H_2 | $H_2C_2O_4$, 0.05380 **M**, $Hg_2C_2O_4$,sol. | Hg		0.6315	41Lar
Pt, H_2 | $H_2C_2O_4$, 0.1022 **M**, $Hg_2C_2O_4$,sol. | Hg		0.6200	41Lar
Pt, H_2 | $H_2C_2O_4$, 0.2054 **M**, $Hg_2C_2O_4$,sol. | Hg		0.6085	41Lar
Pt, H_2 | $H_2C_2O_4$, 0.3987 **M**, $Hg_2C_2O_4$,sol. | Hg		0.5981	41Lar
Pt, H_2 | $H_2C_2O_4$, 0.4778 **M**, $Hg_2C_2O_4$,sol. | Hg		0.5953	41Lar
Pt, H_2 | $H_2C_2O_4$, 0.6127 **M**, $Hg_2C_2O_4$,sol. | Hg		0.5917	41Lar
Pt, H_2 | H_3PO_4 ,0.001026 **m**, Hg_2HPO_4,sol. | Hg		0.7992	49deV

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | H_3PO_4, 0.01026 **m**, Hg_2HPO_4,sol. | Hg		0.7740	49deV
Pt, H_2 | H_3PO_4, 0.02052 **m**, Hg_2HPO_4,sol. | Hg		0.7600	49deV
Pt, H_2 | H_3PO_4, 0.05130 **m**, Hg_2HPO_4,sol. | Hg		0.7443	49deV
Pt, H_2 | H_3PO_4, 0.1032 **m**, Hg_2HPO_4,sol. | Hg		0.7339	49deV
Pt, H_2 | H_2SO_4, 0.05 **m**, Hg_2SO_4,sat. | Hg	298 K	0.75434	35Har
(U_T [V]=U_0–0.010·10^{-3}(T–298 K)–0.0583·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 0.10 **m**, Hg_2SO_4,sat. | Hg	298 K	0.73714	35Har
(U_T [V]=U_0–0.0317·10^{-3}(T–298 K)–0.0347·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 0.20 **m**, Hg_2SO_4,sat. | Hg	298 K	0.7197	35Har
(U_T [V]=U_0–0.07305·10^{-3}(T–298 K)–0.0272·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 0.50 **m**, Hg_2SO_4,sat. | Hg	298 K	0.69609	35Har
(U_T [V]=U_0–0.13994·10^{-3}(T–298 K)–0.0133·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 1.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.67600	35Har
(U_T [V]=U_0–0.18577·10^{-3}(T–298 K)+0.0622·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 1.50 **m**, Hg_2SO_4,sat. | Hg	298 K	0.66213	35Har
(U_T [V]=U_0–0.21635·10^{-3}(T–298 K)+0.0947·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 2.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.65088	35Har
(U_T [V]=U_0–0.235·10^{-3}(T–298 K)+0.1031·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 3.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.63041	35Har
(U_T [V]=U_0–0.2418·10^{-3}(T–298 K)+0.1174·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 4.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.61201	35Har
(U_T [V]=U_0–0.2406·10^{-3}(T–298 K)+0.1301·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 5.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.59498	35Har
(U_T [V]=U_0–0.2343·10^{-3}(T–298 K)+0.1427·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 6.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.57952	35Har
(U_T [V]=U_0–0.2323·10^{-3}(T–298 K)+0.1583·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 7.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.56550	35Har
(U_T [V]=U_0–0.2268·10^{-3}(T–298 K)+0.1689·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 8.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.55289	35Har
(U_T [V]=U_0–0.2217·10^{-3}(T–298 K)+0.2116·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 9.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.54153	35Har
(U_T [V]=U_0–0.2154·10^{-3}(T–298 K)+0.2944·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 10.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.53149	35Har
(U_T [V]=U_0–0.2098·10^{-3}(T–298 K)+0.3273·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 11.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.52189	35Har
(U_T [V]=U_0–0.2033·10^{-3}(T–298 K)+0.3657·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 12.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.51304	35Har
(U_T [V]=U_0–0.1991·10^{-3}(T–298 K)+0.3859·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 13.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.50475	35Har
(U_T [V]=U_0–0.1960·10^{-3}(T–298 K)+0.3198·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 14.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.49688	35Har
(U_T [V]=U_0–0.1870·10^{-3}(T–298 K)+0.2909·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 15.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.48950	35Har
(U_T [V]=U_0–0.2098·10^{-3}(T–298 K)+0.3273·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 16.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.48236	35Har
(U_T [V]=U_0–0.1762·10^{-3}(T–298 K)+0.1506·10^{-6}(T–298 K)^2)			
Pt, H_2 | H_2SO_4, 17.0 **m**, Hg_2SO_4,sat. | Hg	298 K	0.47552	35Har
(U_T [V]=U_0–0.1628·10^{-3}(T–298 K)+0.1181·10^{-6}(T–298 K)^2)			

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | H_2SO_4, 17.5 **m**, Hg_2SO_4,sat. | Hg	298 K	0.47213	35Har
(U_T [V]=U_0–0.1655·10^{-3}(T–298 K)+0.1170·10^{-6}(T–298 K)^2)			
Pt, H_2, p=742.5 Torr | H_2SO_4, 0.00482 **m**, Hg_2SO_4 | Hg	298.7 K	0.8241	10Brö
Pt, H_2, p=750.7 Torr | H_2SO_4, 0.06209 **m**, Hg_2SO_4 | Hg	298.2 K	0.7491	10Brö
Pt, H_2, p=765.5 Torr | H_2SO_4, 1.181 **m**, Hg_2SO_4 | Hg	286.9 K	0.6403	10Brö
Pt, H_2, p=754.2 Torr | H_2SO_4, 1.181 **m**, Hg_2SO_4 | Hg	298.7 K	0.6370	10Brö
Pt, H_2, p=733.5 Torr | H_2SO_4, 1.181 **m**, Hg_2SO_4 | Hg	310.4 K	0.6328	10Brö
Pt, H_2, p=640.3 Torr | H_2SO_4, 1.181 **m**, Hg_2SO_4 | Hg	333.5 K	0.6242	10Brö
Pt, H_2, p=458.1 Torr | H_2SO_4, 1.181 **m**, Hg_2SO_4 | Hg	353.2 K	0.6134	10Brö
Pt, H_2 | H_2SO_4, Hg_2SO_4 | Hg	273 K	0.63495	35Har
	278 K	0.63097	35Har
	283 K	0.62704	35Har
	288 K	0.62307	35Har
	293 K	0.61930	35Har
	298 K	0.61515	35Har
	303 K	0.61107	35Har
	308 K	0.60701	35Har
	313 K	0.60305	35Har
	318 K	0.59900	35Har
	323 K	0.59487	35Har
	328 K	0.59051	35Har
	333 K	0.58659	35Har
Pt, H_2 | H_2SO_4, 0.01 **m**, $(NH_4)_2SO_4$, 0.0 **m**, Hg_2SO_4 | Hg		0.7971	34Cro
Pt, H_2 | H_2SO_4, 0.1 **m**, $(NH_4)_2SO_4$, 0.0 **m**, Hg_2SO_4 | Hg		0.7372	34Cro
Pt, H_2 | H_2SO_4, 0.01 **m**, $(NH_4)_2SO_4$, 0.01 **m**, Hg_2SO_4 | Hg		0.7952	34Cro
Pt, H_2 | H_2SO_4, 0.1 **m**, $(NH_4)_2SO_4$, 0.01 **m**, Hg_2SO_4 | Hg		0.7379	34Cro
Pt, H_2 | H_2SO_4, 0.01 **m**, $(NH_4)_2SO_4$, 0.04 **m**, Hg_2SO_4 | Hg		0.7964	34Cro
Pt, H_2 | H_2SO_4, 0.1 **m**, $(NH_4)_2SO_4$, 0.04 **m**, Hg_2SO_4 | Hg		0.7382	34Cro
Pt, H_2 | H_2SO_4, 0.01 **m**, $(NH_4)_2SO_4$, 0.1 **m**, Hg_2SO_4 | Hg		0.7997	34Cro
Pt, H_2 | H_2SO_4, 0.1 **m**, $(NH_4)_2SO_4$, 0.1 **m**, Hg_2SO_4 | Hg		0.7397	34Cro
Pt, H_2 | H_2SO_4, 0.01 **m**, $(NH_4)_2SO_4$, 0.2 **m**, Hg_2SO_4 | Hg		0.8037	34Cro
Pt, H_2 | H_2SO_4, 0.1 **m**, $(NH_4)_2SO_4$, 0.2 **m**, Hg_2SO_4 | Hg		0.7456	34Cro
Pt, H_2 | H_2SO_4, 0.01 **m**, $(NH_4)_2SO_4$, 0.5 **m**, Hg_2SO_4 | Hg		0.8102	34Cro
Pt, H_2 | H_2SO_4, 0.1 **m**, $(NH_4)_2SO_4$, 0.5 **m**, Hg_2SO_4 | Hg		0.7497	34Cro
Pt, H_2 | H_2SO_4, 0.01 **m**, $(NH_4)_2SO_4$, 1.0 **m**, Hg_2SO_4 | Hg		0.8153	34Cro
Pt, H_2 | H_2SO_4, 0.1 **m**, $(NH_4)_2SO_4$, 1.0 **m**, Hg_2SO_4 | Hg		0.7558	34Cro
Pt, H_2 | H_2SO_4, 0.01 **m**, $(NH_4)_2SO_4$, 2.0 **m**, Hg_2SO_4 | Hg		0.8251	34Cro
Pt, H_2 | H_2SO_4, 0.1 **m**, $(NH_4)_2SO_4$, 2.0 **m**, Hg_2SO_4 | Hg		0.7616	34Cro
Pt, H_2 | H_2SO_4, 0.0504 **M**, acetic acid, 0.0 **M**, Hg_2SO_4 | Hg		0.7544	33Mac
Pt, H_2 | H_2SO_4, 0.1369 **M**, acetic acid, 0.0 **M**, Hg_2SO_4 | Hg		0.7290	33Mac
Pt, H_2 | H_2SO_4, 0.2741 **M**, acetic acid, 0.0 **M**, Hg_2SO_4 | Hg		0.7115	33Mac
Pt, H_2 | H_2SO_4, 0.5472 **M**, acetic acid, 0.0 **M**, Hg_2SO_4 | Hg		0.6929	33Mac
Pt, H_2 | H_2SO_4, 1.093 **M**, acetic acid, 0.0 **M**, Hg_2SO_4 | Hg		0.6712	33Mac
Pt, H_2 | H_2SO_4, 1.642 **M**, acetic acid, 0.0 **M**, Hg_2SO_4 | Hg		0.6555	33Mac
Pt, H_2 | H_2SO_4, 2.184 **M**, acetic acid, 0.0 **M**, Hg_2SO_4 | Hg		0.6418	33Mac
Pt, H_2 | H_2SO_4, 0.0545 **M**, acetic acid, 0.0991 **M**, Hg_2SO_4 | Hg		0.7517	33Mac
Pt, H_2 | H_2SO_4, 0.1004 **M**, acetic acid, 0.0991 **M**, Hg_2SO_4 | Hg		0.7367	33Mac
Pt, H_2 | H_2SO_4, 0.2008 **M**, acetic acid, 0.0991 **M**, Hg_2SO_4 | Hg		0.7193	33Mac

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | H_2SO_4, 0.4016 **M**, acetic acid, 0.0991 **M**, Hg_2SO_4 | Hg		0.7012	33Mac
Pt, H_2 | H_2SO_4, 0.8197 **M**, acetic acid, 0.0991 **M**, Hg_2SO_4 | Hg		0.6802	33Mac
Pt, H_2 | H_2SO_4, 1.642 **M**, acetic acid, 0.0991 **M**, Hg_2SO_4 | Hg		0.6550	33Mac
Pt, H_2 | H_2SO_4, 3.278 **M**, acetic acid, 0.0991 **M**, Hg_2SO_4 | Hg		0.6153	33Mac
Pt, H_2 | H_2SO_4, 0.0547 **M**, acetic acid, 0.548 **M**, Hg_2SO_4 | Hg		0.7496	33Mac
Pt, H_2 | H_2SO_4, 0.1366 **M**, acetic acid, 0.548 **M**, Hg_2SO_4 | Hg		0.7269	33Mac
Pt, H_2 | H_2SO_4, 0.2733 **M**, acetic acid, 0.548 **M**, Hg_2SO_4 | Hg		0.7095	33Mac
Pt, H_2 | H_2SO_4, 0.5453 **M**, acetic acid, 0.548 **M**, Hg_2SO_4 | Hg		0.6908	33Mac
Pt, H_2 | H_2SO_4, 1.092 **M**, acetic acid, 0.548 **M**, Hg_2SO_4 | Hg		0.6690	33Mac
Pt, H_2 | H_2SO_4, 1.637 **M**, acetic acid, 0.548 **M**, Hg_2SO_4 | Hg		0.6531	33Mac
Pt, H_2 | H_2SO_4, 2.179 **M**, acetic acid, 0.548 **M**, Hg_2SO_4 | Hg		0.6392	33Mac
Pt, H_2 | H_2SO_4, 0.0546 **M**, acetic acid, 1.365 **M**, Hg_2SO_4 | Hg		0.7463	33Mac
Pt, H_2 | H_2SO_4, 0.1364 **M**, acetic acid, 1.365 **M**, Hg_2SO_4 | Hg		0.7237	33Mac
Pt, H_2 | H_2SO_4, 0.2738 **M**, acetic acid, 1.365 **M**, Hg_2SO_4 | Hg		0.7061	33Mac
Pt, H_2 | H_2SO_4, 0.5458 **M**, acetic acid, 1.365 **M**, Hg_2SO_4 | Hg		0.6873	33Mac
Pt, H_2 | H_2SO_4, 1.091 **M**, acetic acid, 1.365 **M**, Hg_2SO_4 | Hg		0.6653	33Mac
Pt, H_2 | H_2SO_4, 1.637 **M**, acetic acid, 1.365 **M**, Hg_2SO_4 | Hg		0.6492	33Mac
Pt, H_2 | H_2SO_4, 2.177 **M**, acetic acid, 1.365 **M**, Hg_2SO_4 | Hg		0.6353	33Mac
Pt, H_2 | H_2SO_4, 0.0546 **M**, acetic acid, 2.186 **M**, Hg_2SO_4 | Hg		0.7430	33Mac
Pt, H_2 | H_2SO_4, 0.1087 **M**, acetic acid, 2.186 **M**, Hg_2SO_4 | Hg		0.7260	33Mac
Pt, H_2 | H_2SO_4, 0.2731 **M**, acetic acid, 2.186 **M**, Hg_2SO_4 | Hg		0.7027	33Mac
Pt, H_2 | H_2SO_4, 0.5478 **M**, acetic acid, 2.186 **M**, Hg_2SO_4 | Hg		0.6838	33Mac
Pt, H_2 | H_2SO_4, 1.093 **M**, acetic acid, 2.186 **M**, Hg_2SO_4 | Hg		0.6616	33Mac
Pt, H_2 | H_2SO_4, 1.640 **M**, acetic acid, 2.186 **M**, Hg_2SO_4 | Hg		0.6449	33Mac
Pt, H_2 | H_2SO_4, 2.184 **M**, acetic acid, 2.186 **M**, Hg_2SO_4 | Hg		0.6308	33Mac
Pt, H_2 | H_2SO_4, 0.0543 **M**, acetic acid, 3.278 **M**, Hg_2SO_4 | Hg		0.7379	33Mac
Pt, H_2 | H_2SO_4, 0.1089 **M**, acetic acid, 3.278 **M**, Hg_2SO_4 | Hg		0.7209	33Mac
Pt, H_2 | H_2SO_4, 0.2732 **M**, acetic acid, 3.278 **M**, Hg_2SO_4 | Hg		0.6979	33Mac
Pt, H_2 | H_2SO_4, 0.5471 **M**, acetic acid, 3.278 **M**, Hg_2SO_4 | Hg		0.6789	33Mac
Pt, H_2 | H_2SO_4, 1.091 **M**, acetic acid, 3.278 **M**, Hg_2SO_4 | Hg		0.6567	33Mac
Pt, H_2 | H_2SO_4, 1.639 **M**, acetic acid, 3.278 **M**, Hg_2SO_4 | Hg		0.6398	33Mac
Pt, H_2 | H_2SO_4, 2.181 **M**, acetic acid, 3.278 **M**, Hg_2SO_4 | Hg		0.6248	33Mac
Pt, H_2 | H_2SO_4, 0.0545 **M**, acetic acid, 4.38 **M**, Hg_2SO_4 | Hg		0.7325	33Mac
Pt, H_2 | H_2SO_4, 0.1086 **M**, acetic acid, 4.38 **M**, Hg_2SO_4 | Hg		0.7156	33Mac
Pt, H_2 | H_2SO_4, 0.2729 **M**, acetic acid, 4.38 **M**, Hg_2SO_4 | Hg		0.6928	33Mac
Pt, H_2 | H_2SO_4, 0.5482 **M**, acetic acid, 4.38 **M**, Hg_2SO_4 | Hg		0.6734	33Mac
Pt, H_2 | H_2SO_4, 1.093 **M**, acetic acid, 4.38 **M**, Hg_2SO_4 | Hg		0.6505	33Mac
Pt, H_2 | H_2SO_4, 1.643 **M**, acetic acid, 4.38 **M**, Hg_2SO_4 | Hg		0.6334	33Mac
Pt, H_2 | H_2SO_4, 2.185 **M**, acetic acid, 4.38 **M**, Hg_2SO_4 | Hg		0.6181	33Mac
Pt, H_2 | NaCl, 1.49 **M**, NaOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	273 K	1.0164	29Bje
	298 K	1.0246	29Bje
Pt, H_2 | NaCl, 0.99 **M**, NaOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	273 K	1.0185	29Bje
	298 K	1.0267	29Bje
Pt, H_2 | NaCl, 0.29 **M**, NaOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	273 K	1.0252	29Bje
	298 K	1.0323	29Bje
Pt, H_2 | NaCl, 0.0 **M**, NaOH, 0.1 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	273 K	1.0808	29Bje
	298 K	1.0938	29Bje

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | NaCl, 0.05 **M**, NaOH, 0.05 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	298 K	1.0764	29Bje
Pt, H_2 | NaCl, 0.07 **M**, NaOH, 0.03 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	273 K	1.0543	29Bje
	298 K	1.0640	29Bje
Pt, H_2 | NaCl, 0.09 **M**, NaOH, 0.01 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	273 K	1.0291	29Bje
	298 K	1.0363	29Bje
Pt, H_2 | NaCl, 0.0 **M**, NaOH, 0.03 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	273 K	1.0563	29Bje
	298 K	1.0666	29Bje
Pt, H_2 | NaCl, 0.015 **M**, NaOH, 0.015 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	273 K	1.0412	29Bje
	298 K	1.0490	29Bje
Pt, H_2 | NaCl, 0.021 **M**, NaOH, 0.009 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	273 K	1.0296	29Bje
	298 K	1.0371	29Bje
Pt, H_2 | NaCl, 0.027 **M**, NaOH, 0.003 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	273 K	1.0326	29Bje
	298 K	1.0408	29Bje
Pt, H_2 | NaCl, 0.0021 **M**, NaOH, 0.0009 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	273 K	0.9776	29Bje
	298 K	0.9536	29Bje
Pt, H_2 | NaCl, 0.0027 **M**, NaOH, 0.0003 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	273 K	0.9522	29Bje
	298 K	0.9536	29Bje
Pt, H_2 | NaCl, 0.0007 **M**, NaOH, 0.0003 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	273 K	0.9527	29Bje
	298 K	0.9542	29Bje
Pt, H_2 | NaCl, 0.0000 **M**, NaOH, 0.0003 **M** ‖ KCl, 3.5 **M** ‖ KCl, 0.1 **M**, HgCl | Hg	273 K	0.9512	29Bje
	298 K	0.9535	29Bje
Pt, H_2 | NaOH, 0.673 **M**, HgO | Hg	298.65 K	0.9287	08Brö
	302.45 K	0.9248	08Brö
	314.85 K	0.9208	08Brö
	327.45 K	0.9163	08Brö
Pt, H_2 | HCl, 1.0 **M**, $IrCl_6^{2-}$, 4.86 m**M**, $IrCl_6^{3-}$, 5.16 m**M** | Ir		1.0240	31Sho
Pt, H_2 | HCl, 1.0 **M**, $IrCl_6^{2-}$, 4.84 m**M**, $IrCl_6^{3-}$, 5.19 m**M** | Ir		1.0237	31Sho
Pt, H_2 | HCl, 1.0 **M**, $IrCl_6^{2-}$, 2.48 m**M**, $IrCl_6^{3-}$, 7.63 m**M** | Ir		0.9969	31Sho
Pt, H_2 | HCl, 1.0 **M**, $IrCl_6^{2-}$, 1.98 m**M**, $IrCl_6^{3-}$, 8.15 m**M** | Ir		0.9891	31Sho
Pt, H_2 | H_2SO_4, 0.513 **M** ‖ H_2SO_4, 0.513 **M**, $PbSO_4$,sol. | PbO_2	0 °C	1.610	99Dol
Pt, H_2 | H_2SO_4, 1.01 **M** ‖ H_2SO_4, 1.01 **M**, $PbSO_4$,sol. | PbO_2	0 °C	1.617	99Dol
Pt, H_2 | H_2SO_4, 2.30 **M** ‖ H_2SO_4, 2.30 **M**, $PbSO_4$,sol. | PbO_2	0 °C	1.654	99Dol
Pt, H_2 | H_2SO_4, 7.70 **M** ‖ H_2SO_4, 7.70 **M**, $PbSO_4$,sol. | PbO_2	0 °C	1.801	99Dol
Pt, H_2 | HI, 0.0891 **m** | PbI_2,sat. | Pb_{Hg}		0.2252	28Get1
Pt, H_2 | H_2SO_4, 0.001 **m**, $PbSO_4$ | Pb_{Hg}	273 K	0.0954	34Shr2
	298 K	0.0959	34Shr2
Pt, H_2 | H_2SO_4, 0.002 **m**, $PbSO_4$ | Pb_{Hg}	273 K	0.1173	34Shr2
	298 K	0.1190	34Shr2
Pt, H_2 | H_2SO_4, 0.005 **m**, $PbSO_4$ | Pb_{Hg}	273 K	0.1449	34Shr2
	298 K	0.1475	34Shr2
Pt, H_2 | H_2SO_4, 0.01 **m**, $PbSO_4$ | Pb_{Hg}	273 K	0.1644	34Shr2
	298 K	0.1674	34Shr2
Pt, H_2 | H_2SO_4, 0.02 **m**, $PbSO_4$ | Pb_{Hg}	273 K	0.1830	34Shr2
	298 K	0.1864	34Shr2
Pt, H_2 | $HClO_4$, 1.0 **M** ‖ HNO_3, 1.0 **M**, $Co(NO_3)_2$, $Co(NO_3)_3$ | Pt	0 °C	1.800	37Noy
Pt, H_2 | $HClO_4$, 1.0 **M** ‖ HNO_3, 3.0 **M**, $Co(NO_3)_2$, $Co(NO_3)_3$ | Pt	25 °C	1.842	37Noy

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | $HClO_4$, 1.0 **M** ‖ HNO_3, 4.0 **M**, $Co(NO_3)_2$, $Co(NO_3)_3$ | Pt	0 °C	1.816	37Noy
	25 °C	1.850	37Noy
Pt, H_2 | $HClO_4$, 0.19841 **M** ‖ $HClO_4$, 0.19460 **M**, $Ce(ClO_4)_3$, 0.74 m**M** $Ce(ClO_4)_4$, 9.33 m**M** | Pt		1.7331	43She
Pt, H_2 | $HClO_4$, 0.8071 **M** ‖ $HClO_4$, 0.8024 **M**, $Ce(ClO_4)_3$, 1.54 m**M** $Ce(ClO_4)_4$, 8.43 m**M** | Pt		1.6998	43She
Pt, H_2 | $HClO_4$, 0.19841 **M** ‖ $HClO_4$, 0.19649 **M**, $Ce(ClO_4)_3$, 2.95 m**M** $Ce(ClO_4)_4$, 7.11 m**M** | Pt		1.6982	43She
Pt, H_2 | $HClO_4$, 2.4472 **M** ‖ $HClO_4$, 2.3855 **M**, $Ce(ClO_4)_3$, 3.55 m**M** $Ce(ClO_4)_4$, 6.58 m**M** | Pt		1.7175	43She
Pt, H_2 | $HClO_4$, 1.6353 **M** ‖ $HClO_4$, 1.6310 **M**, $Ce(ClO_4)_3$, 4.98 m**M** $Ce(ClO_4)_4$, 4.75 m**M** | Pt		1.6998	43She
Pt, H_2 | $HClO_4$, 0.15 **m** ‖ $HClO_4$, 0.1 **m**, $Hg_2(ClO_4)_2$, 0.0900 **m**, $Hg(ClO_4)_2$, 0.00810 **m** | Pt		0.9034	31Pop
Pt, H_2 | $HClO_4$, 0.15 **m** ‖ $HClO_4$, 0.1 **m**, $Hg_2(ClO_4)_2$, 0.0632 **m**, $Hg(ClO_4)_2$, 0.0040 **m** | Pt		0.9032	31Pop
Pt, H_2 | $HClO_4$, 0.15 **m** ‖ $HClO_4$, 0.1 **m**, $Hg_2(ClO_4)_2$, 0.0449 **m**, $Hg(ClO_4)_2$, 0.0020 **m** | Pt		0.9024	31Pop
Pt, H_2 | $HClO_4$, 0.15 **m** ‖ $HClO_4$, 0.1 **m**, $Hg_2(ClO_4)_2$, 0.0316 **m**, $Hg(ClO_4)_2$, 0.0010 **m** | Pt		0.9017	31Pop
Pt, H_2 | $HClO_4$, 0.08 **m** ‖ $HClO_4$, 0.08 **m**, $Hg_2(ClO_4)_2$, 0.004 **m**, $Hg(ClO_4)_2$, 0.004 **m** | Pt		0.9020	31Pop
Pt, H_2 | $HClO_4$, 0.2 **m** ‖ $HClO_4$, 0.2 **m**, $Hg_2(ClO_4)_2$, 0.004 **m**, $Hg(ClO_4)_2$, 0.004 **m** | Pt		0.9007	31Pop
Pt, H_2 | $HClO_4$, 1.0 **m** ‖ $HClO_4$, 1.0 **m**, $Hg_2(ClO_4)_2$, 0.004 **m**, $Hg(ClO_4)_2$, 0.004 **m** | Pt		0.9072	31Pop
Pt, H_2 | $HClO_4$, 0.4948 **M** ‖ $HClO_4$, 0.4948 **M**, $TlClO_4$, 0.0383 **M**, $Tl(ClO_4)_3$, 0.001474 **M** | Pt		1.2379	36She
Pt, H_2 | $HClO_4$, 0.7160 **M** ‖ $HClO_4$, 0.7160 **M**, $TlClO_4$, 0.0533 **M**, $Tl(ClO_4)_3$, 0.002119 **M** | Pt		1.2301	36She
Pt, H_2 | $HClO_4$, 0.9496 **M** ‖ $HClO_4$,0.9496 **M**, $TlClO_4$, 0.1682 **M**, $Tl(ClO_4)_3$, 0.005804 **M** | Pt		1.2244	36She
Pt, H_2 | $HClO_4$, 1.2270 **M** ‖ $HClO_4$,1.2270 **M**, $TlClO_4$, 0.1090 **M**, $Tl(ClO_4)_3$, 0.009230 **M** | Pt		1.2318	36She
Pt, H_2 | $HClO_4$, 0.1077 **M**, VO_2ClO_4, 0.5293 m**m**, $VO(ClO_4)_2$, 0.004420 **m** | Pt		0.9031	34Car
Pt, H_2 | $HClO_4$, 0.1077 **M**, VO_2ClO_4, 0.2.489 m**m**, $VO(ClO_4)_2$, 0.004420 **m** | Pt		0.9395	34Car
Pt, H_2 | $HClO_4$, 0.1077 **M**, VO_2ClO_4, 9.855 m**m**, $VO(ClO_4)_2$, 0.004420 **m** | Pt		0.9723	34Car
Pt, H_2 | $HClO_4$, 0.1077 **M**, VO_2ClO_4, 19.67 m**m**, $VO(ClO_4)_2$, 0.004420 **m** | Pt		0.9875	34Car
Pt, H_2 | $HClO_4$, 0.02398 **M**, VO_2ClO_4, 4.947 m**m**, $VO(ClO_4)_2$, 0.004420 **m** | Pt		0.9098	34Car
Pt, H_2 | $HClO_4$, 0.1077 **M**, VO_2ClO_4, 4.947 m**m**, $VO(ClO_4)_2$, 0.004420 **m** | Pt		0.9554	34Car
Pt, H_2 | $HClO_4$, 0.2338 **M**, VO_2ClO_4, 4.947 m**m**, $VO(ClO_4)_2$, 0.004420 **m** | Pt		0.9783	34Car
Pt, H_2 | $HClO_4$, 0.4442 **M**, VO_2ClO_4, 4.947 m**m**, $VO(ClO_4)_2$, 0.004420 **m** | Pt		0.99774	34Car
Pt, H_2 | $HClO_4$, 0.900 **M**, VO_2ClO_4, 4.947 m**m**, $VO(ClO_4)_2$, 0.004420 **m** | Pt		1.0198	34Car
Pt, H_2 | $HClO_4$, 0.1077 **M**, VO_2ClO_4, 0.004947 m**m**, $VO(ClO_4)_2$, 0.03511 **m** | Pt		0.9084	34Car
Pt, H_2 | $HClO_4$, 0.025 **M** ‖ $HClO_4$, 0.025 **M**, $KMnO_4$, 0.0237 **M**, MnO_2 | Pt		1.4799	26Bro
Pt, H_2 | $HClO_4$, 0.025 **M** ‖ $HClO_4$, 0.025 **M**, $KMnO_4$, 0.0237 **M**, MnO_2 | Pt		1.4975	26Bro
Pt, H_2 | $HClO_4$, 0.025 **M** ‖ $HClO_4$, 0.025 **M**, $KMnO_4$, 0.0237 **M**, MnO_2 | Pt		1.5114	26Bro
Pt, H_2 | KOH, 5.5 **M** ‖ $Ni_2O_3*(H_2O)_x$	10 °C	1.305	06Zed
	65 °C	1.266	06Zed

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | $HClO_4$, 1.027 **m** ‖ $HClO_4$, 1.027 **m**, (Np^{IV}/Np^{III}=0.279) | Pt		0.1216	52Coh
Pt, H_2 | $HClO_4$, 1.027 **m** ‖ $HClO_4$, 1.027 **m**, (Np^{IV}/Np^{III}=0.656 | Pt		0.1428	52Coh
Pt, H_2 | $HClO_4$, 1.027 **m** ‖ $HClO_4$, 1.027 **m**, (Np^{IV}/Np^{III}=1.028) | Pt		0.1546	52Coh
Pt, H_2 | $HClO_4$, 1.027 **m** ‖ $HClO_4$, 1.027 **m**, (Np^{IV}/Np^{III}=1.725) | Pt		0.1676	52Coh
Pt, H_2 | $HClO_4$, 1.027 **m** ‖ $HClO_4$, 1.027 **m**, (Np^{IV}/Np^{III}=2.300) | Pt		0.1746	52Coh
Pt, H_2 | $HClO_4$, 1.027 **m** ‖ $HClO_4$, 1.027 **m**, (Np^{IV}/Np^{III}=6.874) | Pt		0.2039	52Coh
Pt, H_2 | $HClO_4$, 1.0 **m** ‖ $HClO_4$, 1.0 **m**, (Np^{V}/Np^{IV}=0.3896) | Pt		0.7154	52Coh
Pt, H_2 | $HClO_4$, 1.0 **m** ‖ $HClO_4$, 1.0 **m**, (Np^{V}/Np^{IV}=0.7801) | Pt		0.7329	52Coh
Pt, H_2 | $HClO_4$, 1.0 **m** ‖ $HClO_4$, 1.0 **m**, (Np^{V}/Np^{IV}=1.171) | Pt		0.7444	52Coh
Pt, H_2 | $HClO_4$, 1.0 **m** ‖ $HClO_4$, 1.0 **m**, (Np^{V}/Np^{IV}=1.560) | Pt		0.7518	52Coh
Pt, H_2 | $HClO_4$, 1.0 **m** ‖ $HClO_4$, 1.0 **m**, (Np^{V}/Np^{IV}=2.117) | Pt		0.7581	52Coh
Pt, H_2 | $HClO_4$, 1.0 **m** ‖ $HClO_4$, 1.0 **m**, (Np^{V}/Np^{IV}=2.498) | Pt		0.7631	52Coh
Pt, H_2 | $HClO_4$, 1.0 **m** ‖ $HClO_4$, 1.0 **m**, (Np^{V}/Np^{IV}=3.120) | Pt		0.7689	52Coh
Pt, H_2 | $HClO_4$, 1.027 **m** ‖ $HClO_4$, 1.027 **m**, (Np^{VI}/Np^{V}=0.3175) | Pt		1.1067	52Coh
Pt, H_2 | $HClO_4$, 1.027 **m** ‖ $HClO_4$, 1.027 **m**, (Np^{VI}/Np^{V}=0.6207) | Pt		1.1239	52Coh
Pt, H_2 | $HClO_4$, 1.027 **m** ‖ $HClO_4$, 1.027 **m**, (Np^{VI}/Np^{V}=1.336) | Pt		1.1434	52Coh
Pt, H_2 | $HClO_4$, 1.027 **m** ‖ $HClO_4$, 1.027 **m**, (Np^{VI}/Np^{V}=2.448) | Pt		1.1591	52Coh
Pt, H_2 | $HClO_4$, 1.027 **m** ‖ $HClO_4$, 1.027 **m**, (Np^{VI}/Np^{V}=5.024) | Pt		1.1721	52Coh
Pt, H_2 | $HClO_4$, 1.0 **M** ‖ $TlNO_3$,$Tl(NO_3)_3$, HNO_3, 0.5 **M** | Pt		1.2303	36Noy
Pt, H_2 | $HClO_4$, 1.0 **M** ‖ $TlNO_3$,$Tl(NO_3)_3$, HNO_3, 1.0 **M** | Pt		1.2306	36Noy
Pt, H_2 | $HClO_4$, 1.0 **M** ‖ $TlNO_3$,$Tl(NO_3)_3$, HNO_3, 1.0 **M** | Pt	0 °C	1.193	36Noy
Pt, H_2 | $HClO_4$, 2.0 **M** ‖ $TlNO_3$, $Tl(NO_3)_3$, HNO_3, 2.0 | Pt		1.2300	36Noy
Pt, H_2 | CH_3COOH, 0.01 **M**, CH_3COONa, 0.01 **M**, $SrCl_2$, 0.996 **M**, quinhydrone | Pt		0.0012	31Lar
Pt, H_2 | CH_3COOH, 0.01 **M**, CH_3COONa, 0.01 **M**, $SrCl_2$, 1.793 **M**, quinhydrone | Pt		0.0037	31Lar
Pt, H_2 | CH_3COOH, 0.01 **M**, CH_3COONa, 0.01 **M**, $BaCl_2$, 0.5 **M**, quinhydrone | Pt		0.0007	31Lar
Pt, H_2 | CH_3COOH, 0.01 **M**, CH_3COONa, 0.01 **M**, $BaCl_2$, 0.9 **M**, quinhydrone | Pt		0.0025	31Lar
Pt, H_2 | CH_3COOH, 0.01 **M**, CH_3COONa, 0.01 **M**, $BaCl_2$, 0.15 **M**, quinhydrone | Pt		0.0043	31Lar
Pt, H_2 | HCl, 6.078 **m** ‖ HCl, 6.078 **m** ‖ N_2+Cl_2, 0.00597 **M**, Pt		1.1478	25For
Pt, H_2 | HCl, 6.078 **m** ‖ HCl, 6.078 **m** ‖ N_2+Cl_2, 0.01492 **M**, Pt		1.1595	25For
Pt, H_2 | HCl, 6.078 **m** ‖ HCl, 6.078 **m** ‖ N_2+Cl_2, 0.04633 **M**, Pt		1.1741	25For
Pt, H_2 | HCl, 6.078 **m** ‖ HCl, 6.078 **m** ‖ N_2+Cl_2, 0.09905 **M**, Pt		1.1840	25For
Pt, H_2 | HI, 5.8 **M** ‖ HI, 5.8 **M**, I_2, 0.29 **M** ‖ Pt	31.6 °C	0.2372	10Ste
(U_T [V]=0.2372–11.0·10^{-4}(T–31.6 °C))			
Pt, H_2 | HCl, 0.0100 **m** ‖ HIO_3, 0.0100 **m**, I_2,sol. | Pt		1.1425	31Lun
Pt, H_2 | HCl, 0.0950 **m** ‖ HIO_3, 0.0950 **m**, I_2,sol. | Pt		1.1621	31Lun
Pt, H_2 | $HClO_4$, 0.001 **m**, $Pb(ClO_4)_2$, 0.001 **m**, PbO_2,sol. | Pt		1.378	33Bat
Pt, H_2 | $HClO_4$, 0.0025 **m**, $Pb(ClO_4)_2$, 0.0025 **m**, PbO_2,sol. | Pt		1.389	33Bat
Pt, H_2 | $HClO_4$, 0.01 **m**, $Pb(ClO_4)_2$, 0.01 **m**, PbO_2,sol. | Pt		1.407	33Bat
Pt, H_2 | $HClO_4$, 0.025 **m**, $Pb(ClO_4)_2$, 0.025 **m**, PbO_2,sol. | Pt		1.417	33Bat
Pt, H_2 | $HClO_4$, 0.1 **m**, $Pb(ClO_4)_2$, 0.1 **m**, PbO_2,sol. | Pt		1.432	33Bat
Pt, H_2 | $HClO_4$, 0.25 **m**, $Pb(ClO_4)_2$, 0.01 **m**, PbO_2,sol. | Pt		1.501	33Bat
Pt, H_2 | $HClO_4$, 0.25 **m**, $Pb(ClO_4)_2$, 0.001 **m**, PbO_2,sol. | Pt		1.523	33Bat
Pt, H_2 | $HClO_4$, 0.25 **m**, $Pb(ClO_4)_2$, 0.0001 **m**, PbO_2,sol. | Pt		1.549	33Bat

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt, H_2 | H_2SO_4, 0.0005 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.42151	35Har
(U_T [V]=U_0–0.68479·10^{-3}(T–298 K)+0.9111·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 0.0007 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.43276	35Har
(U_T [V]=U_0–0.65906·10^{-3}(T–298 K)+0.9028·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 0.001 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.444458	35Har
(U_T [V]=U_0–0.63680·10^{-3}(T–298 K)+0.9222·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 0.002 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.46706	35Har
(U_T [V]=U_0–0.60396·10^{-3}(T–298 K)+0.9028·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 0.003 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.47976	35Har
(U_T [V]=U_0–0.58896·10^{-3}(T–298 K)+0.9077·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 0.005 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.49528	35Har
(U_T [V]=U_0–0.58172·10^{-3}(T–298 K)+0.8994·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 0.007 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.50502	35Har
(U_T [V]=U_0–0.57454·10^{-3}(T–298 K)+0.9111·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 0.01 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.51540	35Har
(U_T [V]=U_0–0.56682·10^{-3}(T–298 K)+0.9206·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 0.02 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.53427	35Har
(U_T [V]=U_0–0.53635·10^{-3}(T–298 K)+0.9127·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 0.03 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.54464	35Har
(U_T [V]=U_0–0.51416·10^{-3}(T–298 K)+0.9063·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 0.05 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.55733	35Har
(U_T [V]=U_0–0.48853·10^{-3}(T–298 K)+0.9050·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 0.10 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.57398	35Har
(U_T [V]=U_0–0.46565·10^{-3}(T–298 K)+0.8812·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 0.20 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.59038	35Har
(U_T [V]=U_0–0.42392·10^{-3}(T–298 K)+0.8696·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 0.50 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.61231	35Har
(U_T [V]=U_0–0.35634·10^{-3}(T–298 K)+0.8600·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 1.0 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.63121	35Har
(U_T [V]=U_0–0.30673·10^{-3}(T–298 K)+0.7847·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 1.50 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.64429	35Har
(U_T [V]=U_0–0.27453·10^{-3}(T–298 K)+0.7512·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 2.0 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.65507	35Har
(U_T [V]=U_0–0.25546·10^{-3}(T–298 K)+0.7438·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 3.0 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.67536	35Har
(U_T [V]=U_0–0.24794·10^{-3}(T–298 K)+0.7292·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 4.0 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.69379	35Har
(U_T [V]=U_0–0.24851·10^{-3}(T–298 K)+0.7166·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 5.0 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.71096	35Har
(U_T [V]=U_0–0.25419·10^{-3}(T–298 K)+0.7042·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 6.0 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.72646	35Har
(U_T [V]=U_0–0.25540·10^{-3}(T–298 K)+0.6884·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, 7.0 **m**, $PbSO_4$,sat. | PbO_2 | Pt	298 K	1.74061	35Har
(U_T [V]=U_0–0.26037·10^{-3}(T–298 K)+0.6778·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pt, H_2 | H_2SO_4, $PbSO_4$ | PbO_2 | Pt	273 K	1.6769	35Har
	288 K	1.6816	35Har
	298 K	1.6849	35Har
	313 K	1.6923	35Har
	333 K	1.6986	35Har

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
H_2 | $HClO_4$ | Sb_2O_3 | Sb	313 K	0.1524	24Sch1
Indium (In)			
In | $InCl_3$, 0.001515 **M**, HCl, 0.02 **M**, AgCl | Ag		0.7370	30Hak
In | $InCl_3$, 0.00533 **M**, HCl, 0.02 **M**, AgCl | Ag		0.7264	30Hak
In | $InCl_3$, 0.01104 **M**, HCl, 0.02 **M**, AgCl | Ag		0.7191	30Hak
In | $InCl_3$, 0.02744 **M**, HCl, 0.02 **M**, AgCl | Ag		0.7061	30Hak
In | $InCl_3$, 0.01 **M** ‖ KNO_3, 0.1 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.739	04Thi
In | $InCl_3$, 0.1 **M** ‖ KNO_3,0.1 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.728	04Thi
In | $InCl_3$, 1.0 **M** ‖ KNO_3,0.1 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.714	04Thi
$In_{0.016wt\%}Hg$ | $In_2(SO_4)_3$ | $In_{0.008wt\%}Hg$	0 °C	0.0058	10Ric
$In_{0.242wt\%}Hg$ | $In_2(SO_4)_3$ | $In_{0.120wt\%}Hg$	0 °C	0.0057	10Ric
$In_{0.384wt\%}Hg$ | $In_2(SO_4)_3$ | $In_{0.242wt\%}Hg$	0 °C	0.0038	10Ric
$In_{1.92wt\%}Hg$ | $In_2(SO_4)_3$ | $In_{0.384wt\%}Hg$	0 °C	0.0145	10Ric
Iodine (I)			
Pt | I_2, 0.005 **M**, KI, 0.115 **M** ‖ I_2, 0.05 **M**, KI, 0.115 **M** | Pt	20.4 °C	0.0444	08Lau
Pt | I_2, 0.005 **M**, KI, 0.23 **M** ‖ I_2, 0.05 **M**, KI, 0.23 **M** | Pt	20.4 °C	0.0352	08Lau
Pt | I_2, 0.005 **M**, KI, 0.344 **M** ‖ I_2, 0.05 **M**, KI, 0.344 **M** | Pt	20.4 °C	0.0326	08Lau
Pt | I_2, 0.0005 **M**, KI, 0.115 **M** ‖ I_2, 0.005 **M**, KI 0.005 **M** | Pt	20.4 °C	0.0304	08Lau
Pt | I_2, 0.0005 **M**, KI, 0.23 **M** ‖ I_2, 0.005 **M**, KI, 0.23 **M** | Pt	20.4 °C	0.0297	08Lau
Pt | I_2, 0.0005 **M**, KI, 0.344 **M** ‖ I_2, 0.005 **M**, KI, 0.344 **M** | Pt	20.4 °C	0.0295	08Lau
Pt | I_2, 0.0139 **M**, KI, 1.0 **M** ‖ I_2, 0.056 **M**, KI, 1.0 **M** | Pt	20.4 °C	0.0184	08Lau
Pt | I_2, 0.0139 **M**, KI, 1.0 **M** ‖ I_2, 0.109 **M**, KI, 1.0 **M** | Pt	20.4 °C	0.0273	08Lau
Pt | I_2, 0.0139 **M**, KI, 1.0 **M** ‖ I_2, 0.209 **M**, KI, 1.0 **M** | Pt	20.4 °C	0.0376	08Lau
Pt | I_2, 0.0139 **M**, KI, 1.0 **M** ‖ I_2, 0.508 **M**, KI, 1.0 **M** | Pt	20.4 °C	0.0588	08Lau
Iron (Fe)			
Fe | $FeSO_4$, 0.123 **M**, H_2SO_4, 0.0 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.753	24Ric
Fe | $FeSO_4$, 0.123 **M**, H_2SO_4, 0.0005 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.753	24Ric
Fe | $FeSO_4$, 0.131 **M**, H_2SO_4, 0.05 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.724	24Ric
Fe | $FeSO_4$, 0.123 **M**, H_2SO_4, 0.125 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.669	24Ric
Fe | $FeSO_4$, 0.130 **M**, H_2SO_4, 0.25 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.634	24Ric
Lanthanum (La)			
La | $LaBr_3$,sat. ‖ KCl, 1 **M**, Hg_2Cl_2 | Hg		0.71	29Mül1
$La_{1.5wt\%}Hg$ | $LaBr_3$,sat. ‖ KCl, 1 **M**, Hg_2Cl_2 | Hg		0.932	29Mül1
Lead (Pb)			
Pb | $PbCl_2$ | $Pb_{sat.}Hg$		0.00584	29Car
Pb | PbI_2,sol., KI, 0.01 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.5277	17Lew3
Pb | PbI_2,sol., KI, 0.1 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.5812	17Lew3
Pb | $Pb(NO_3)_2$, 0.0036 **M** ‖ KNO_3, 0.5 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.487	03Sac
Pb | $Pb(NO_3)_2$, 0.011 **M** ‖ KNO_3, 0.5 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.472	03Sac
Pb | $Pb(NO_3)_2$, 0.05 **M** ‖ KNO_3, 0.5 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.456	03Sac
Pb | $Pb(NO_3)_2$, 0.36 **M** ‖ KNO_3, 0.5 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.438	03Sac
Pb | PbO(yellow),sol., NaOH, 1.0 **M** ‖ NaOH, 1.0 **M**, HgO | Hg	20 °C	0.6734	22App
Pb | PbO(red),sol., NaOH, 1.0 **M** ‖ NaOH, 1.0 **M**, HgO | Hg	20 °C	0.6808	22App
Pb | PbO,sol., NaOH, HgO | Hg	0 °C	0.673	26Fri
	23.5 °C	0.675	26Fri
	39 °C	0.675	26Fri
Pb | PbS,sol., Na_2S, 0.005 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	10.5 °C	0.700	22Jel

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pb | PbS,sol., Na_2S, 0.01 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	10.5 °C	0.712	22Jel
Pb | PbS,sol., Na_2S, 0.05 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	10.5 °C	0.743	22Jel
Pb | PbS,sol., Na_2S, 0.1 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	10.5 °C	0.764	22Jel
Pb | PbS,sol., Na_2S, 0.5 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	10.5 °C	0.787	22Jel
Pb | $PbSO_4$,sol., Na_2SO_4, Hg_2SO_4 | Hg		0.9697	18Hen
Pb | $PbSO_4$,sol., $ZnSO_4$,sat., Hg_2SO_4 | $Pb_{sat.}Hg$		0.00586	34Hoy
Pb | $PbSO_4$,sat., H_2SO_4, 0.00046 **M** | PbO_2	0 °C	1.472	99Dol
Pb | $PbSO_4$,sat., H_2SO_4, 0.0124 **M** | PbO_2	0 °C	1.672	99Dol
Pb | $PbSO_4$,sat., H_2SO_4, 0.0505 **M** | PbO_2	0 °C	1.745	99Dol
Pb | $PbSO_4$,sat., H_2SO_4, 0.111 **M** | PbO_2	0 °C	1.782	99Dol
Pb | $PbSO_4$,sat., H_2SO_4, 0.180 **M** | PbO_2	0 °C	1.808	99Dol
Pb | $PbSO_4$,sat., H_2SO_4, 0.360 **M** | PbO_2	0 °C	1.844	99Dol
Pb | $PbSO_4$,sat., H_2SO_4, 1.00 **M** | PbO_2	0 °C	1.896	99Dol
Pb | $PbSO_4$ | H_2SO_4, 0.05 **m** | $PbSO_4$ | PbO_2 | Pt	298 K	1.76874	35Har
(U_T [V]=U_0–0.310·10^{-3}(T–298 K)+0.134·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pb | $PbSO_4$ | H_2SO_4, 0.1 **m** | $PbSO_4$ | PbO_2 | Pt	298 K	1.80207	35Har
(U_T [V]=U_0–0.265·10^{-3}(T–298 K)+0.129·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pb | $PbSO_4$ | H_2SO_4, 0.5 **m** | $PbSO_4$ | PbO_2 | Pt	298 K	1.87910	35Har
(U_T [V]=U_0–0.045·10^{-3}(T–298 K)+0.126·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pb | $PbSO_4$ | H_2SO_4, 1.0 **m** | $PbSO_4$ | PbO_2 | Pt	298 K	1.91737	35Har
(U_T [V]=U_0+0.0561·10^{-3}(T–298 K)+0.108·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pb | $PbSO_4$ | H_2SO_4, 2.0 **m** | $PbSO_4$ | PbO_2 | Pt	298 K	1.96637	35Har
(U_T [V]=U_0+0.159·10^{-3}(T–298 K)+0.108·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pb | $PbSO_4$ | H_2SO_4, 3.0 **m** | $PbSO_4$ | PbO_2 | Pt	298 K	2.00874	35Har
(U_T [V]=U_0+0.178·10^{-3}(T–298 K)+0.097·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pb | $PbSO_4$ | H_2SO_4, 4.0 **m** | $PbSO_4$ | PbO_2 | Pt	298 K	2.04789	35Har
(U_T [V]=U_0+0.177·10^{-3}(T–298 K)+0.091·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pb | $PbSO_4$ | H_2SO_4, 5.0 **m** | $PbSO_4$ | PbO_2 | Pt	298 K	2.08502	35Har
(U_T [V]=U_0+0.167·10^{-3}(T–298 K)+0.087·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pb | $PbSO_4$ | H_2SO_4, 6.0 **m** | $PbSO_4$ | PbO_2 | Pt	298 K	2.11910	35Har
(U_T [V]=U_0+0.162·10^{-3}(T–298 K)+0.085·10^{-6}(T–298 K)^2, 273<T<333 K)			
Pb | $PbSO_4$ | H_2SO_4, 7.0 **m** | $PbSO_4$ | PbO_2 | Pt	298 K	2.15071	35Har
(U_T [V]=U_0+0.153·10^{-3}(T–298 K)+0.080·10^{-6}(T–298 K)^2, 273<T<333 K)			
$Pb_{sat.}Hg$ | $PbCl_2$, 0.125 g/kg H_2O, KCl, 3.75 g/kg H_2O, AgCl | Ag		0.5480	33All
$Pb_{sat.}Hg$ | $PbCl_2$, 0.125 g/kg H_2O, KCl, 35 g/kg H_2O, AgCl | Ag		0.5311	33All
$Pb_{sat.}Hg$ | $PbCl_2$, 0.125 g/kg H_2O, KCl, 68 g/kg H_2O, AgCl | Ag		0.5311	33All
$Pb_{sat.}Hg$ | $PbCl_2$, 0.125 g/kg H_2O, KCl, 292 g/kg H_2O, AgCl | Ag		0.5661	33All
$Pb_{sat.}Hg$ | $PbCl_2$, 0.25 g/kg H_2O, KCl, 3.75 g/kg H_2O, AgCl | Ag		0.5362	33All
$Pb_{sat.}Hg$ | $PbCl_2$, 0.25 g/kg H_2O, KCl, 35 g/kg H_2O, AgCl | Ag		0.5176	33All
$Pb_{sat.}Hg$ | $PbCl_2$, 0.25 g/kg H_2O, KCl, 68 g/kg H_2O, AgCl | Ag		0.5188	33All
$Pb_{sat.}Hg$ | $PbCl_2$, 0.25 g/kg H_2O, KCl, 292 g/kg H_2O, AgCl | Ag		0.5521	33All
$Pb_{sat.}Hg$ | $PbCl_2$, 0.5 g/kg H_2O, KCl, 3.75 g/kg H_2O, AgCl | Ag		0.5256	33All
$Pb_{sat.}Hg$ | $PbCl_2$, 0.5 g/kg H_2O, KCl, 35 g/kg H_2O, AgCl | Ag		0.5057	33All
$Pb_{sat.}Hg$ | $PbCl_2$, 0.5 g/kg H_2O, KCl, 68 g/kg H_2O, AgCl | Ag		0.5071	33All
$Pb_{sat.}Hg$ | $PbCl_2$, 0.5 g/kg H_2O, KCl, 292 g/kg H_2O, AgCl | Ag		0.5392	33All
$Pb_{sat.}Hg$ | $PbCl_2$, 1.0 g/kg H_2O, KCl, 3.75 g/kg H_2O, AgCl | Ag		0.5149	33All
$Pb_{sat.}Hg$ | $PbCl_2$, 1.0 g/kg H_2O, KCl, 35 g/kg H_2O, AgCl | Ag		0.4939	33All
$Pb_{sat.}Hg$ | $PbCl_2$, 1.0 g/kg H_2O, KCl, 68 g/kg H_2O, AgCl | Ag		0.4946	33All

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
$Pb_{sat.}Hg \mid PbCl_2$, 1.0 g/kg H_2O, KCl, 292 g/kg H_2O, AgCl $\mid$ Ag		0.5272	33All
$Pb_{sat.}Hg \mid PbCl_2$, 0.125 g/kg H_2O, LiCl, 2 g/kg $H_2O \parallel$ LiCl, 2 g/kg H_2O, AgCl $\mid$ Ag		0.5407	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.125 g/kg H_2O, LiCl, 10 g/kg $H_2O \parallel$ LiCl, 10 g/kg H_2O, AgCl $\mid$ Ag		0.5209	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.125 g/kg H_2O, LiCl, 50 g/kg $H_2O \parallel$ LiCl, 50 g/kg H_2O, AgCl $\mid$ Ag		0.5117	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.125 g/kg H_2O, LiCl, 200 g/kg $H_2O \parallel$ LiCl, 200 g/kg H_2O, AgCl $\mid$ Ag		0.5239	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.125 g/kg H_2O, LiCl, 600 g/kg $H_2O \parallel$ LiCl, 600 g/kg H_2O, AgCl $\mid$ Ag		0.5658	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.125 g/kg H_2O, LiCl,sat. $\parallel$ LiCl,sat., AgCl $\mid$ Ag		0.5679	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.25 g/kg H_2O, LiCl, 2 g/kg $H_2O \parallel$ LiCl, 2 g/kg H_2O, AgCl $\mid$ Ag		0.5315	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.25 g/kg H_2O, LiCl, 10 g/kg $H_2O \parallel$ LiCl, 10 g/kg H_2O, AgCl $\mid$ Ag		0.5122	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.25 g/kg H_2O, LiCl, 50 g/kg $H_2O \parallel$ LiCl, 50 g/kg H_2O, AgCl $\mid$ Ag		0.5032	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.25 g/kg H_2O, LiCl, 200 g/kg $H_2O \parallel$ LiCl, 200 g/kg H_2O, AgCl $\mid$ Ag		0.5163	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.25 g/kg H_2O, LiCl, 600 g/kg $H_2O \parallel$ LiCl, 600 g/kg H_2O, AgCl $\mid$ Ag		0.5575	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.25 g/kg H_2O, LiCl,sat. $\parallel$ LiCl,sat., AgCl $\mid$ Ag		0.5594	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.5 g/kg H_2O, LiCl, 2 g/kg $H_2O \parallel$ LiCl, 2 g/kg H_2O, AgCl $\mid$ Ag		0.5221	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.5 g/kg H_2O, LiCl, 10 g/kg $H_2O \parallel$ LiCl, 10 g/kg H_2O, AgCl $\mid$ Ag		0.5032	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.5 g/kg H_2O, LiCl, 50 g/kg $H_2O \parallel$ LiCl, 50 g/kg H_2O, AgCl $\mid$ Ag		0.4946	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.5 g/kg H_2O, LiCl, 200 g/kg $H_2O \parallel$ LiCl, 200 g/kg H_2O, AgCl $\mid$ Ag		0.5084	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.5 g/kg H_2O, LiCl, 600 g/kg $H_2O \parallel$ LiCl, 600 g/kg H_2O, AgCl $\mid$ Ag		0.5485	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.5 g/kg H_2O, LiCl,sat. $\parallel$ LiCl,sat., AgCl $\mid$ Ag		0.5510	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 1.0 g/kg H_2O, LiCl, 2 g/kg $H_2O \parallel$ LiCl, 2 g/kg H_2O, AgCl $\mid$ Ag		0.5128	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 1.0 g/kg H_2O, LiCl, 10 g/kg $H_2O \parallel$ LiCl, 10 g/kg H_2O, AgCl $\mid$ Ag		0.494	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 1.0 g/kg H_2O, LiCl, 50 g/kg $H_2O \parallel$ LiCl, 50 g/kg H_2O, AgCl $\mid$ Ag		0.4867	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 1.0 g/kg H_2O, LiCl, 200 g/kg $H_2O \parallel$ LiCl, 200 g/kg H_2O, AgCl $\mid$ Ag		0.5002	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 1.0 g/kg H_2O, LiCl, 600 g/kg $H_2O \parallel$ LiCl, 600 g/kg H_2O, AgCl $\mid$ Ag		0.5397	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 1.0 g/kg H_2O, LiCl,sat. $\parallel$ LiCl,sat., AgCl $\mid$ Ag		0.5435	35Par
$Pb_{sat.}Hg \mid PbCl_2$,sat., LiCl, 2 g/kg $H_2O \parallel$ LiCl, 2 g/kg H_2O, AgCl $\mid$ Ag		0.4842	35Par
$Pb_{sat.}Hg \mid PbCl_2$,sat., LiCl, 10 g/kg $H_2O \parallel$ LiCl, 10 g/kg H_2O, AgCl $\mid$ Ag		0.4842	35Par
$Pb_{sat.}Hg \mid PbCl_2$,sat., LiCl, 50 g/kg $H_2O \parallel$ LiCl, 50 g/kg H_2O, AgCl $\mid$ Ag		0.4842	35Par
$Pb_{sat.}Hg \mid PbCl_2$,sat., LiCl, 200 g/kg $H_2O \parallel$ LiCl, 200 g/kg H_2O, AgCl $\mid$ Ag		0.4842	35Par

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
$Pb_{sat.}Hg \mid PbCl_2$,sat., LiCl, 600 g/kg $H_2O \parallel$ LiCl, 600 g/kg H_2O, AgCl $\mid$ Ag		0.4842	35Par
$Pb_{sat.}Hg \mid PbCl_2$,sat., LiCl,sat. $\parallel$ LiCl,sat., AgCl $\mid$ Ag		0.4842	35Par
$Pb_{sat.}Hg \mid PbCl_2$, 0.125 g/kg H_2O, NaCl, 5 g/kg $H_2O \parallel$ NaCl, 5 g/kg H_2O, AgCl $\mid$ Ag		0.5369	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.125 g/kg H_2O, NaCl, 10 g/kg $H_2O \parallel$ NaCl, 10 g/kg H_2O, AgCl $\mid$ Ag		0.5276	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.125 g/kg H_2O, NaCl, 50 g/kg $H_2O \parallel$ NaCl, 50 g/kg H_2O, AgCl $\mid$ Ag		0.5187	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.125 g/kg H_2O, NaCl, 100 g/kg $H_2O \parallel$ NaCl, 100 g/kg H_2O, AgCl $\mid$ Ag		0.5215	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.125 g/kg H_2O, NaCl, 250 g/kg $H_2O \parallel$ NaCl, 250 g/kg H_2O, AgCl $\mid$ Ag		0.5408	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.125 g/kg H_2O, NaCl,sat. $\parallel$ NaCl,sat., AgCl $\mid$ Ag		0.5557	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.25 g/kg H_2O, NaCl, 5 g/kg $H_2O \parallel$ NaCl, 5 g/kg H_2O, AgCl $\mid$ Ag		0.5272	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.25 g/kg H_2O, NaCl, 10 g/kg $H_2O \parallel$ NaCl, 10 g/kg H_2O, AgCl $\mid$ Ag		0.5184	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.25 g/kg H_2O, NaCl, 50 g/kg $H_2O \parallel$ NaCl, 50 g/kg H_2O, AgCl $\mid$ Ag		0.5092	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.25 g/kg H_2O, NaCl, 100 g/kg $H_2O \parallel$ NaCl, 100 g/kg H_2O, AgCl $\mid$ Ag		0.5128	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.25 g/kg H_2O, NaCl, 250 g/kg $H_2O \parallel$ NaCl, 250 g/kg H_2O, AgCl $\mid$ Ag		0.5315	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.25 g/kg H_2O, NaCl,sat. $\parallel$ NaCl,sat., AgCl $\mid$ Ag		0.5469	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.5 g/kg H_2O, NaCl, 5 g/kg $H_2O \parallel$ NaCl, 5 g/kg H_2O, AgCl $\mid$ Ag		0.5174	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.5 g/kg H_2O, NaCl, 10 g/kg $H_2O \parallel$ NaCl, 10 g/kg H_2O, AgCl $\mid$ Ag		0.5092	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.5 g/kg H_2O, NaCl, 50 g/kg $H_2O \parallel$ NaCl, 50 g/kg H_2O, AgCl $\mid$ Ag		0.5002	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.5 g/kg H_2O, NaCl, 100 g/kg $H_2O \parallel$ NaCl, 100 g/kg H_2O, AgCl $\mid$ Ag		0.5038	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.5 g/kg H_2O, NaCl, 250 g/kg $H_2O \parallel$ NaCl, 250 g/kg H_2O, AgCl $\mid$ Ag		0.5222	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 0.5 g/kg H_2O, NaCl,sat. $\parallel$ NaCl,sat., AgCl $\mid$ Ag		0.5376	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 1.0 g/kg H_2O, NaCl, 5 g/kg $H_2O \parallel$ NaCl, 5 g/kg H_2O, AgCl $\mid$ Ag		0.5078	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 1.0 g/kg H_2O, NaCl, 10 g/kg $H_2O \parallel$ NaCl, 10 g/kg H_2O, AgCl $\mid$ Ag		0.5001	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 1.0 g/kg H_2O, NaCl, 50 g/kg $H_2O \parallel$ NaCl, 50 g/kg H_2O, AgCl $\mid$ Ag		0.4914	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 1.0 g/kg H_2O, NaCl, 100 g/kg $H_2O \parallel$ NaCl, 100 g/kg H_2O, AgCl $\mid$ Ag		0.4945	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 1.0 g/kg H_2O, NaCl, 250 g/kg $H_2O \parallel$ NaCl, 250 g/kg H_2O, AgCl $\mid$ Ag		0.5126	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 1.0 g/kg H_2O, NaCl,sat. $\parallel$ NaCl,sat., AgCl $\mid$ Ag		0.5280	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 2.0 g/kg H_2O, NaCl, 5 g/kg $H_2O \parallel$ NaCl, 5 g/kg H_2O, AgCl $\mid$ Ag		0.4984	35Bra
$Pb_{sat.}Hg \mid PbCl_2$, 2.0 g/kg H_2O, NaCl, 10 g/kg $H_2O \parallel$ NaCl, 10 g/kg H_2O,		0.4909	35Bra

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
AgCl | Ag			
$Pb_{sat.}Hg$ | $PbCl_2$, 2.0 g/kg H_2O, NaCl, 100 g/kg H_2O ‖ NaCl, 100 g/kg H_2O, AgCl | Ag		0.4858	35Bra
$Pb_{sat.}Hg$ | $PbCl_2$, 2.0 g/kg H_2O, NaCl, 250 g/kg H_2O ‖ NaCl, 250 g/kg H_2O, AgCl | Ag		0.5040	35Bra
$Pb_{sat.}Hg$ | $PbCl_2$, 2.0 g/kg H_2O, NaCl,sat. ‖ NaCl,sat., AgCl | Ag		0.5183	35Bra
$Pb_{sat.}Hg$ | $PbCl_2$,sat., NaCl, 5 g/kg H_2O ‖ NaCl, 5 g/kg H_2O, AgCl | Ag		0.4842	35Bra
$Pb_{sat.}Hg$ | $PbCl_2$,sat., NaCl, 10 g/kg H_2O ‖ NaCl, 10 g/kg H_2O, AgCl | Ag		0.4842	35Bra
$Pb_{sat.}Hg$ | $PbCl_2$,sat., NaCl, 50 g/kg H_2O ‖ NaCl, 50 g/kg H_2O, AgCl | Ag		0.4842	35Bra
$Pb_{sat.}Hg$ | $PbCl_2$,sat., NaCl, 100 g/kg H_2O ‖ NaCl, 100 g/kg H_2O, AgCl | Ag		0.4842	35Bra
$Pb_{sat.}Hg$ | $PbCl_2$,sat., NaCl, 250 g/kg H_2O ‖ NaCl, 250 g/kg H_2O, AgCl | Ag		0.4842	35Bra
$Pb_{sat.}Hg$ | $PbCl_2$,sat., NaCl,sat. ‖ NaCl,sat., AgCl | Ag		0.4842	35Bra
$Pb_{sat.}Hg$ | $PbBr_2$,sat., KBr ‖ KBr, AgBr | Ag		0.3465	32Can
(U_T [V]=0.3465–2.89·10^{-4}(T–25 °C))			
$Pb_{sat.}Hg$ | $PbBr_2$, 1.536 m**m**, AgBr | Ag		0.4334	42Bat
$Pb_{sat.}Hg$ | $PbBr_2$, 3.626 m**m**, AgBr | Ag		0.4053	42Bat
(U_T [V]=0.4053+61.2·10^{-5}(T–25 °C)–3.97·10^{-6}(T–25 °C)^2)			
$Pb_{sat.}Hg$ | $PbBr_2$, 7.196 m**m**, AgBr | Ag		0.3835	42Bat
(U_T [V]=0.3835+54.6·10^{-5}(T–25 °C)–3.20·10^{-6}(T–25 °C)^2)			
$Pb_{sat.}Hg$ | $PbBr_2$, 9.551 m**m**, AgBr | Ag		0.3758	42Bat
(U_T [V]=0.3758+51.1·10^{-5}(T–25 °C)–2.90·10^{-6}(T–25 °C)^2)			
$Pb_{sat.}Hg$ | $PbBr_2$, 12.736 m**m**, AgBr | Ag		0.3682	42Bat
(U_T [V]=0.3682+50.2·10^{-5}(T–25 °C)–2.70·10^{-6}(T–25 °C)^2)			
$Pb_{sat.}Hg$ | $PbBr_2$, 18.200 m**m**, AgBr | Ag		0.3584	42Bat
(U_T [V]=0.3584+46.4·10^{-5}(T–25 °C)–2.37·10^{-6}(T–25 °C)^2)			
$Pb_{sat.}Hg$ | $PbCl_2$, 0.2116 m**m**, AgCl | Ag		0.6537	29Car
$Pb_{sat.}Hg$ | $PbCl_2$, 0.6197 m**m**, AgCl | Ag		0.6143	29Car
$Pb_{sat.}Hg$ | $PbCl_2$, 1.034 m**m**, AgCl | Ag		0.5960	29Car
$Pb_{sat.}Hg$ | $PbCl_2$, 1.337 m**m**, AgCl | Ag		0.5870	29Car
$Pb_{sat.}Hg$ | $PbCl_2$, 2.348 m**m**, AgCl | Ag		0.5677	29Car
$Pb_{sat.}Hg$ | $PbCl_2$, 2.62 m**m**, AgCl | Ag		0.5639	29Car
$Pb_{sat.}Hg$ | $PbCl_2$, 5.16 m**m**, AgCl | Ag		0.5419	29Car
$Pb_{sat.}Hg$ | $PbCl_2$, 10.39 m**m**, AgCl | Ag		0.5205	29Car
$Pb_{sat.}Hg$ | $PbCl_2$, 20.48 m**m**, AgCl | Ag		0.5012	29Car
$Pb_{sat.}Hg$ | $PbCl_2$, 29.55 m**m**, AgCl | Ag		0.4913	29Car
$Pb_{sat.}Hg$ | $PbCl_2$, 39.05 m**m**, AgCl | Ag		0.4842	29Car
$Pb_{sat.}Hg$ | $PbCl_2$, 0.001028 m**m**, $Ba(NO_3)_2$, 0.002135 **m**, AgCl | Ag		0.5994	36Cro
$Pb_{sat.}Hg$ | $PbCl_2$, 0.002056 m**m**, $Ba(NO_3)_2$, 0.00427 **m**, AgCl | Ag		0.5764	36Cro
$Pb_{sat.}Hg$ | $PbCl_2$, 0.004112 m**m**, $Ba(NO_3)_2$, 0.01709 **m**, AgCl | Ag		0.5333	36Cro
$Pb_{sat.}Hg$ | $PbCl_2$, 0.01097 m**m**, $Ba(NO_3)_2$, 0.02279 **m**, AgCl | Ag		0.5248	36Cro
$Pb_{sat.}Hg$ | $PbCl_2$, 0.01729 m**m**, $Ba(NO_3)_2$, 0.03575 **m**, AgCl | Ag		0.5130	36Cro
$Pb_{sat.}Hg$ | $PbCl_2$, 0.019037 m**m**, $Ba(NO_3)_2$, 0.00503 **m**, AgCl | Ag		0.5043	36Cro
$Pb_{sat.}Hg$ | $PbCl_2$, 0.019037 m**m**, $Ba(NO_3)_2$, 0.01003 **m**, AgCl | Ag		0.5052	36Cro
$Pb_{sat.}Hg$ | $PbCl_2$, 0.019037 m**m**, $Ba(NO_3)_2$, 0.02008 **m**, AgCl | Ag		0.5072	36Cro
$Pb_{sat.}Hg$ | $PbCl_2$, 0.019037 m**m**, $Ba(NO_3)_2$, 0.04013 **m**, AgCl | Ag		0.5096	36Cro
$Pb_{sat.}Hg$ | $PbCl_2$, 0.019037 m**m**, $Ba(NO_3)_2$, 0.08027 **m**, AgCl | Ag		0.5143	36Cro
$Pb_{sat.}Hg$ | $PbCl_2$, 0.019037 m**m**, $Ba(NO_3)_2$, 0.10025 **m**, AgCl | Ag		0.5154	36Cro
$Pb_{sat.}Hg$ | $PbCl_2$, 0.000446 m**m**, $Ba(NO_3)_2$, 0.02004 **m**, AgCl | Ag		0.6464	36Cro

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
$Pb_{sat.}Hg \mid PbCl_2$, 0.000892 m**m**, $Ba(NO_3)_2$, 0.02004 **m**, AgCl | Ag		0.6182	36Cro
$Pb_{sat.}Hg \mid PbCl_2$, 0.003567 m**m**, $Ba(NO_3)_2$, 0.02004 **m**, AgCl | Ag		0.5647	36Cro
$Pb_{sat.}Hg \mid PbCl_2$, 0.007136 m**m**, $Ba(NO_3)_2$, 0.02004 **m**, AgCl | Ag		0.5392	36Cro
$Pb_{sat.}Hg \mid PbCl_2$, 0.01427 m**m**, $Ba(NO_3)_2$, 0.02004 **m**, AgCl | Ag		0.5161	36Cro
$Pb_{sat.}Hg \mid PbCl_2$, 0.02856 m**m**, $Ba(NO_3)_2$, 0.02004 **m**, AgCl | Ag		0.4950	36Cro
PbHg | PbI_2,sol., PbI_2,sat., KI, 0.025 **m**, AgI | Ag	15.3 °C	0.21043	37Can2
	23.6 °C	0.20845	37Can2
	25 °C	0.20810	37Can2
	26.1 °C	0.20708	37Can2
	30.0 °C	0.20760	37Can2
	32.1 °C	0.20623	37Can2
	35 °C	0.20515	37Can2
PbHg | PbI_2,sol., PbI_2,sat., KI, 0.05 **m**, AgI | Ag	15.3 °C	0.2095	37Can2
	17 °C	0.2091	37Can2
	23.6°C	0.2078	37Can2
	25 °C	0.2075	37Can2
	26.1 °C	0.2072	37Can2
	30.0 °C	0.2065	37Can2
	32.2 °C	0.2063	37Can2
	35 °C	0.2053	37Can2
	36 °C	0.2053	37Can2
	40 °C	0.2043	37Can2
PbHg | PbI_2,sol., PbI_2,sat., KI, 0.1 **m**, AgI | Ag	15.3 °C	0.2092	37Can2
	18.7 °C	0.2081	37Can2
	23.6 °C	0.2074	37Can2
	25 °C	0.2070	37Can2
	26.1 °C	0.2070	37Can2
	30.0 °C	0.2056	37Can2
	32.2 °C	0.2060	37Can2
	35 °C	0.2046	37Can2
	36 °C	0.2049	37Can2
	40 °C	0.2035	37Can2
$Pb_{sat.}Hg \mid PbHPO_4$,sol., H_3PO_4, 0.001 **m** | H_2, Pt		0.0623	47Mas
$Pb_{sat.}Hg \mid PbHPO_4$,sol., H_3PO_4, 0.005 **m** | H_2, Pt		0.0978	47Mas
$Pb_{sat.}Hg \mid PbHPO_4$,sol., H_3PO_4, 0.01 **m** | H_2, Pt		0.1113	47Mas
$Pb_{sat.}Hg \mid PbHPO_4$,sol., H_3PO_4, 0.05 **m** | H_2, Pt		0.1383	47Mas
$Pb_{sat.}Hg \mid PbHPO_4$,sol., H_3PO_4, 0.1 **m** | H_2, Pt		0.1492	47Mas
$Pb_{sat.}Hg \mid PbHPO_4$,sol., H_3PO_4, 0.5 **m** | H_2, Pt		0.1724	47Mas
$Pb_{sat.}Hg \mid PbHPO_4$,sol., H_3PO_4, 1.0 **m** | H_2, Pt		0.1828	47Mas
$Pb_{sat.}Hg \mid PbHPO_4$,sol., H_3PO_4, 3.0 **m** | H_2, Pt		0.2018	47Mas
$Pb_{sat.}Hg \mid PbHPO_4$,sol., H_3PO_4, 5.0 **m** | H_2, Pt		0.2134	47Mas
$Pb_{sat.}Hg \mid PbHPO_4$,sol., H_3PO_4, 7.0 **m** | H_2, Pt		0.2229	47Mas
$Pb_{sat.}Hg \mid PbHPO_4$,sol., H_3PO_4, 10.0 **m** | H_2, Pt		0.2343	47Mas
$Pb_{sat.}Hg \mid PbO$(red),sol., $Ba(OH)_2$, 0.132 **M** | H_2, Pt	20 °C	0.2599	29Ish
	25 °C	0.2579	29Ish
	30 °C	0.2561	29Ish
	35 °C	0.2545	29Ish

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
	40 °C	0.2531	29Ish
$Pb_{sat.}Hg \mid PbBr_2$,sol., $PbBr_2$,sat., $Hg_2Br_2 \mid Hg$	20 °C	0.4136	33Ish2
	25 °C	0.4138	33Ish2
	30 °C	0.4140	33Ish2
	35 °C	0.4142	33Ish2
$Pb_{sat.}Hg \mid PbI$,sol., $PbBr_2$,sat., KI, 0.01 **M** ‖ KI, 0.01 **M**, PbI,sol., Hg_2I_2,sol. | Hg (U_T [V]=0.3175+1.29·10^{-4}(T–25 °C), 20<T<40 °C)		0.3175	28Vos
$Pb_{sat.}Hg \mid PbI$,sol., $PbBr_2$,sat., CdI_2,sat. ‖ CdI_2,sat., PbI,sol., Hg_2I_2,sol. | Hg (U_T [V]=0.4170+3.60·10^{-4}(T–25 °C), 25<T<40 °C)		0.4170	28Vos
$Pb_{sat.}Hg \mid PbO$(red),sol., $Ba(OH)_2$, 0.2 **m**, HgO | Hg		0.6714	32Spe2
$Pb_{sat.}Hg \mid PbO$(yellow),sol., $Ba(OH)_2$, 0.2 **m**, HgO | Hg		0.6683	32Spe2
$Pb_{sat.}Hg \mid PbO$(white,hydrate),sol., $Ba(OH)_2$, 0.1 **m**, HgO | Hg		0.6664	32Spe2
$PbHg \mid PbSO_4 \mid Na_2SO_4 \cdot 10H_2O \mid Na_2SO_4$,sat. | $Hg_2SO_4 \mid Hg$	298 K	0.96466	18Hen
$Pb_{sat.}Hg \mid PbSO_4$,sol., Li_2SO_4*H_2O,sol., Li_2SO_4,sat., $Hg_2SO_4 \mid Hg$	20 °C	0.9635	33Ued2
	25 °C	0.9644	33Ued2
	30 °C	0.9653	33Ued2
	35 °C	0.9660	33Ued2
$Pb_{(Hg)} \mid Pb(NO_3)_2$, 0.0005 **M** ‖ $Pb(NO_3)_2$, 0.5 **M** ‖ $Pb_{(Hg)}$		0.104	03Con
$Pb_{(Hg)} \mid Pb(NO_3)_2$, 0.0005 **M** ‖ $Pb(NO_3)_2$, 0.005 **M** ‖ $Pb_{(Hg)}$		0.043	03Con
$Pb_{(Hg)} \mid Pb(NO_3)_2$, 0.005 **M** ‖ $Pb(NO_3)_2$, 0.05 **M** ‖ $Pb_{(Hg)}$		0.035	03Con
$Pb_{(Hg)} \mid Pb(NO_3)_2$, 0.05 **M** ‖ $Pb(NO_3)_2$, 0.5 **M** ‖ $Pb_{(Hg)}$		0.026	03Con
Lithium (Li)			
$Li_{0.035wt\%}Hg \mid$ LiOH, 0.1**m** ‖ LiCl, 0.1 **m** ‖ KCl, 0.1 **m** ‖ KCl, 0.1**m**, $Hg_2Cl_2 \mid Hg$		2.3952	13Lew
$Li_{0.00282wt\%}Hg \mid$ LiCl, 0.0217**m** ‖ KCl, 3.5 **m** ‖ KCl, 1.0 **m**, $Hg_2Cl_2 \mid Hg$		2.3871	24Dru
$Li_{0.00282wt\%}Hg \mid Li_2SO_4$, 0.0098**m** ‖ KCl, 3.5 **m** ‖ KCl, 1.0 **m**, $Hg_2Cl_2 \mid Hg$		2.3861	24Dru
$Li_{0.035wt\%}Hg \mid$ LiCl,sat., $Hg_2Cl_2 \mid Hg$		1.9790	33Ued1
$Li_{0.00282wt\%}Hg \mid$ LiCl, 0.0271 **m** | KCl, 3.5 **m** | KCl, 1.0 **m**, $Hg_2Cl_2 \mid Hg$		2.3871	24Dru
$Li_{0.035wt\%}Hg \mid$ LiOH,sat., HgO | Hg	293 K	2.1288	33Ued1
	298 K	2.1245	33Ued1
	303 K	2.1203	33Ued1
$Li_{0.035wt\%}Hg \mid$ LiOH, 0.1 **m** | LiCl, 0.1 **m** | KCl, 0.1 **m** | KCl, 1.0 **m**, $Hg_2Cl_2 \mid Hg$		2.3952	13Lew
$Li_{0.00282wt\%}Hg \mid Li_2SO_4$, 0.00998 **m** | KCl, 3.5 **m** | KCl, 1.0 **m**, $Hg_2Cl_2 \mid Hg$		2.3861	24Dru
$Li_{0.035wt\%}Hg \mid Li_2SO_4$,sat., $PbSO_4$,sat. | $Pb_{6wt\%}Hg$	293 K	2.6760	33Ued2
	298 K	2.6748	33Ued2
	303 K	2.6736	33Ued2
Magnesium (Mg)			
$Mg \mid MgBr_2$,sat., diethylether ‖ Hg_2Br_2,sol., diethylether | Hg		1.561	34Sch
Manganese (Mn)			
$Mn_{0.41wt\%}Hg \mid MnSO_4$, 0.5 **M** | KCl, 1.0 **M**, $Hg_2Cl_2 \mid Hg$	20 °C	1.428	26Roy
$Mn_{0.92wt\%}Hg \mid MnSO_4$, 0.5 **M** | KCl, 1.0 **M**, $Hg_2Cl_2 \mid Hg$	20 °C	1.445	26Roy
$Mn_{1.83wt\%}Hg \mid MnSO_4$, 0.5 **M** | KCl, 1.0 **M**, $Hg_2Cl_2 \mid Hg$	20 °C	1.449	26Roy
$Mn_{3.36wt\%}Hg \mid MnSO_4$, 0.5 **M** | KCl, 1.0 **M**, $Hg_2Cl_2 \mid Hg$	20 °C	1.451	26Roy
Mercury (Hg)			
$Hg \mid Hg_2Cl_2$, KCl, 0.1 **m** ‖ KCl,sat. ‖ $K_2C_2O_4$, 0.0066 **m**, $Ag_2C_2O_4 \mid Ag$		0.213	32Has
$Hg \mid Hg_2Cl_2$, KCl, 0.1 **m** ‖ KCl,sat. ‖ $K_2C_2O_4$, 0.0131 **m**, $Ag_2C_2O_4 \mid Ag$		0.207	32Has

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Hg | Hg_2Cl_2, KCl, 0.1 **m** ‖ KCl,sat. ‖ $K_2C_2O_4$, 0.0656 **m**, $Ag_2C_2O_4$ | Ag		0.196	32Has
Hg | Hg_2Cl_2, KCl, 0.1 **m** ‖ KCl,sat. ‖ $K_2C_2O_4$, 0.1314 **m**, $Ag_2C_2O_4$ | Ag		0.190	32Has
Hg | Hg_2Cl_2, KCl, 0.1 **m** ‖ KCl,sat. ‖ $K_2C_2O_4$, 0.656 **m**, $Ag_2C_2O_4$ | Ag		0.182	32Has
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ $NaNO_3$, 0.3 **M**, KNO_3, 1.7 **M** ‖ $AgClO_4$, 0.00271 **M** | Ag		0.3114	27Bre
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ $NaNO_3$, 0.3 **M**, KNO_3, 1.7 **M** ‖ $AgClO_4$, 0.0271 **M** | Ag		0.3681	27Bre
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ $NaNO_3$, 0.3 **M**, KNO_3, 1.7 **M** ‖ $AgClO_4$, 0.271 **M** | Ag		0.4221	27Bre
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl,sat. ‖ K_2CrO_4, 0.005 **M**, Ag_2CrO_4 | Ag		0.203	32Has
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl,sat. ‖ K_2CrO_4, 0.01 **M**, Ag_2CrO_4 | Ag		0.195	32Has
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl,sat. ‖ K_2CrO_4, 0.05 **M**, Ag_2CrO_4 | Ag		0.182	32Has
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl,sat. ‖ K_2CrO_4, 0.1 **M**, Ag_2CrO_4 | Ag		0.175	32Has
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl,sat. ‖ K_2CrO_4, 0.5 **M**, Ag_2CrO_4 | Ag		0.161	32Has
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ $NaNO_3$, 0.3 **M**, KNO_3, 1.7 **M** ‖ $AgNO_3$, 0.001 **M** | Ag		0.2856	27Bre
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ $NaNO_3$, 0.3 **M**, KNO_3, 1.7 **M** ‖ $AgNO_3$, 0.01 **M** | Ag		0.3433	27Bre
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ $NaNO_3$, 0.3 **M**, KNO_3, 1.7 **M** ‖ $AgNO_3$, 0.1 **M** | Ag		0.3980	27Bre
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ $NaNO_3$, 0.3 **M**, KNO_3, 1.7 **M** ‖ Ag_2SO_4, 0.002 **M** | Ag		0.3031	27Bre
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ $NaNO_3$, 0.3 **M**, KNO_3, 1.7 **M** ‖ Ag_2SO_4, 0.02 **M** | Ag		0.3579	27Bre
Hg | Hg_2SO_4,sol., $CdSO_4$, 0.005 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.399	32Has
Hg | Hg_2SO_4,sol., $CdSO_4$, 0.01 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.393	32Has
Hg | Hg_2SO_4,sol., $CdSO_4$, 0.05 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.387	32Has
Hg | Hg_2SO_4,sol., $CdSO_4$, 0.1 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.3855	32Has
Hg | Hg_2SO_4,sol., $CdSO_4$, 0.5 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.384	32Has
Hg | Hg_2SO_4,sol., $CdSO_4$, 1.0 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.382	32Has
Hg | Hg_2SO_4,sol., $CdSO_4$, 1.5 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.381	32Has
Hg | Hg_2SO_4,sol., K_2SO_4, 0.005 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.398	32Has
Hg | Hg_2SO_4,sol., K_2SO_4, 0.01 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.392	32Has
Hg | Hg_2SO_4,sol., K_2SO_4, 0.05 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.3775	32Has
Hg | Hg_2SO_4,sol., K_2SO_4, 0.1 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.372	32Has
Hg | Hg_2SO_4,sol., K_2SO_4, 0.5 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.3635	32Has
Hg | Hg_2SO_4,sol., $MgSO_4$, 0.005 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.399	32Has
Hg | Hg_2SO_4,sol., $MgSO_4$, 0.01 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.393	32Has
Hg | Hg_2SO_4,sol., $MgSO_4$, 0.05 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.380	32Has
Hg | Hg_2SO_4,sol., $MgSO_4$, 0.1 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.375	32Has
Hg | Hg_2SO_4,sol., $MgSO_4$, 0.5 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.366	32Has
Hg | Hg_2SO_4,sol., $MgSO_4$, 1.0 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.360	32Has
Hg | Hg_2SO_4,sol., Na_2SO_4, 0.005 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.398	32Has
Hg | Hg_2SO_4,sol., Na_2SO_4, 0.01 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.392	32Has
Hg | Hg_2SO_4,sol., Na_2SO_4, 0.05 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.378	32Has
Hg | Hg_2SO_4,sol., Na_2SO_4, 0.1 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.3735	32Has
Hg | Hg_2SO_4,sol., Na_2SO_4, 0.3 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.369	32Has
Hg | Hg_2SO_4,sol., Na_2SO_4, 0.6 **m** ‖ KCl,sat. ‖ KCl, 0.1 **M**, AgCl | Ag		0.3675	32Has
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $AuCl_3$, 0.01 **M** | Au		0.894	06Faw
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ NH_4NO_3,sat. ‖ HNO_3, 3.32 **M**, Au_2O | Au		0.835	07Cam
Hg | Hg_2SO_4, H_2SO_4, 0.0506 **m** ‖ H_2SO_4, 0.0506 **m**, Au_2O_3 | Au		0.599	27Ger

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Hg | Hg_2SO_4, H_2SO_4, 0.1031 **m** ‖ H_2SO_4, 0.1031 **m**, Au_2O_3 | Au		0.629	27Ger
Hg | Hg_2SO_4, H_2SO_4, 1.031 **m** ‖ H_2SO_4, 1.031 **m**, Au_2O_3 | Au		0.690	27Ger
Hg | PbS,sat. | $PbSO_4$,sat. | $Pb_{sat.}Hg$		0.0670	29Nim
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 1.0 **M**, CuCl, 0.1016 **M**, $CuCl_2$, 0.1001 **M** | Pt		0.1489	09Pom
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 1.0 **M**, CuCl, 0.1254 **M**, $CuCl_2$, 0.2002 **M** | Pt		0.1673	09Pom
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 1.0 **M**, CuCl, 0.1449 **M**, $CuCl_2$, 0.3003 **M** | Pt		0.1841	09Pom
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 1.0 **M**, CuCl, 0.1703 **M**, $CuCl_2$, 0.4004 **M** | Pt		0.1920	09Pom
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 1.0 **M**, CuCl, 0.1919 **M**, $CuCl_2$, 0.5005 **M** | Pt		0.2033	09Pom
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ $FeSO_4$, 0.5 **M** | Fe (ultrapure)	20 °C	0.761	07Ric
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ $FeSO_4$, 0.5 **M** | Fe (technical purity)	20 °C	0.765	07Ric
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ $FeSO_4$, 0.5 **M** | Fe (reduced)	20 °C	0.761	07Ric
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ $FeSO_4$, 0.5 **M** | Fe (hydrogen loaded)	20 °C	0.9	07Ric
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 0.1 **M**, $FeCl_2$, 0.0995 **M**, $FeCl_3$, 0.0005 **M** | Pt	17 °C	0.296	98Pet
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 0.1 **M**, $FeCl_2$, 0.099 **M**, $FeCl_3$, 0.001 **M** | Pt	17 °C	0.312	98Pet
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 0.1 **M**, $FeCl_2$, 0.09 **M**, $FeCl_3$, 0.01 **M** | Pt	17 °C	0.375	98Pet
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 0.1 **M**, $FeCl_2$, 0.05 **M**, $FeCl_3$, 0.05 **M** | Pt	17 °C	0.427	98Pet
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 0.1 **M**, $FeCl_2$, 0.01 **M**, $FeCl_3$, 0.09 **M** | Pt	17 °C	0.483	98Pet
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 0.1 **M**, $FeCl_2$, 0.004 **M**, $FeCl_3$, 0.096 **M** | Pt	17 °C	0.506	98Pet
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 0.1 **M**, $FeCl_2$, 0.002 **M**, $FeCl_3$, 0.098 **M** | Pt	17 °C	0.522	98Pet
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 0.5 **M**, $FeCl_2$, 0.45 **M**, $FeCl_3$, 0.05 **M** | Pt	17 °C	0.346	98Pet
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 0.25 **M**, $FeCl_2$, 0.225 **M**, $FeCl_3$, 0.025 **M** | Pt	17 °C	0.362	98Pet
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ HCl, 0.75 **M**, $FeCl_2$, 0.225 **M**, $FeCl_3$, 0.025 **M** | Pt	17 °C	0.341	98Pet
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ NH_4Cl, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.6972	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ NH_4Cl, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.6890	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ NaCl, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7001	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ NaCl, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.6978	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.6983	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.6913	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ LiCl, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7047	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ LiCl, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7016	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $MgCl_2$, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7038	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $MgCl_2$, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7014	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $CaCl_2$, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7033	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $CaCl_2$, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7026	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $BaCl_2$, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7032	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $BaCl_2$, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**,		0.7025	29Isg

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
$FeCl_2$, 0.1 **M** | Pt			
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $ZnCl_2$, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7055	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $ZnCl_2$, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7100	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $CdCl_2$, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7112	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $CdCl_2$, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7144	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ NH_4NO_3, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7123	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ NH_4NO_3, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7141	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $NaNO_3$, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7139	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $NaNO_3$, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7164	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KNO_3, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7130	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KNO_3, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7138	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $LiNO_3$, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7176	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $LiNO_3$, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7258	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $Mg(NO_3)_2$, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7195	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $Mg(NO_3)_2$, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7292	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $Ca(NO_3)_2$, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7171	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $Ca(NO_3)_2$, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7254	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $Sr(NO_3)_2$, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7171	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $Sr(NO_3)_2$, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7225	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $Cd(NO_3)_2$, 1.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7260	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $Cd(NO_3)_2$, 2.0 **M**, HCl, 0.1 **M**, $FeCl_3$, 0.1 **M**, $FeCl_2$, 0.1 **M** | Pt		0.7382	29Isg
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl, 3.5 **M** ‖ HNO_3, 0.05 **M**, $Fe(NO_3)_2$, 0.05 **M**, $Fe(NO_3)_3$, 0.05 **M** | Pt		0.4494	12Noy
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl, 1.75 **M** ‖ HNO_3, 0.05 **M**, $Fe(NO_3)_2$, 0.05 **M**, $Fe(NO_3)_3$, 0.05 **M** | Pt		0.4479	12Noy
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl, 3.5 **M** ‖ HNO_3, 0.0125 **M**, $Fe(NO_3)_2$, 0.01 **M**, $Fe(NO_3)_3$, 0.05 **M** | Pt		0.4519	12Noy
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl, 1.75 **M** ‖ HNO_3, 0.0125 **M**, $Fe(NO_3)_2$, 0.01 **M**, $Fe(NO_3)_3$, 0.056 **M** | Pt		0.4517	12Noy

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl, 3.5 **M** ‖ HNO_3, 0.05 **M**, $Fe(NO_3)_2$, 0.0045 **M**, $Fe(NO_3)_3$, 0.0056 **M** | Pt		0.4649	12Noy
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl, 1.75 **M** ‖ HNO_3, 0.05 **M**, $Fe(NO_3)_2$, 0.0045 **M**, $Fe(NO_3)_3$, 0.0056 **M** | Pt		0.4649	12Noy
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl, 3.5 **M** ‖ HNO_3, 0.05 **M**, $Fe(NO_3)_2$, 0.00257 **M**, $Fe(NO_3)_3$, 0.0064 **M** | Pt		0.4131	12Noy
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl, 1.75 **M** ‖ HNO_3, 0.05 **M**, $Fe(NO_3)_2$, 0.0257 **M**, $Fe(NO_3)_3$, 0.0064 **M** | Pt		0.4121	12Noy
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $K_4Fe(CN)_6$, 0.09 **M**, $K_3Fe(CN)_6$, 0.01 **M** | Pt		0.133	02Fre
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $K_4Fe(CN)_6$, 0.08 **M**, $K_3Fe(CN)_6$, 0.02 **M** | Pt		0.153	02Fre
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $K_4Fe(CN)_6$, 0.05 **M**, $K_3Fe(CN)_6$, 0.05 **M** | Pt		0.189	02Fre
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $K_4Fe(CN)_6$, 0.02 **M**, $K_3Fe(CN)_6$, 0.08 **M** | Pt		0.222	02Fre
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $K_4Fe(CN)_6$, 0.01 **M**, $K_3Fe(CN)_6$, 0.09 **M** | Pt		0.241	02Fre
Hg | Hg_2Cl_2, KCl, 0.1 **m** ‖ KCN, 0.0 **m**, $K_4Fe(CN)_6$, 0.01 **m**, $K_3Fe(CN)_6$, 0.01 **m** | Pt		0.0684	32Mas1
Hg | Hg_2Cl_2, KCl, 0.1 **m** ‖ KCN, 0.0321 **m**, $K_4Fe(CN)_6$, 0.01 **m**, $K_3Fe(CN)_6$, 0.01 **m** | Pt		0.0747	32Mas1
Hg | Hg_2Cl_2, KCl, 0.1 **m** ‖ KCN, 0.1284 **m**, $K_4Fe(CN)_6$, 0.01 **m**, $K_3Fe(CN)_6$, 0.01 **m** | Pt		0.0892	32Mas1
Hg | Hg_2Cl_2, KCl, 0.1 **m** ‖ KCN, 0.2766 **m**, $K_4Fe(CN)_6$, 0.01 **m**, $K_3Fe(CN)_6$, 0.01 **m** | Pt		0.1040	32Mas1
Hg | Hg_2Cl_2, KCl, 0.1 **m** ‖ KCN, 0.5532 **m**, $K_4Fe(CN)_6$, 0.01 **m**, $K_3Fe(CN)_6$, 0.01 **m** | Pt		0.1182	32Mas1
Hg | Hg_2Cl_2, KCl, 0.1 **m** ‖ KCN, 0.8685 **m**, $K_4Fe(CN)_6$, 0.01 **m**, $K_3Fe(CN)_6$, 0.01 **m** | Pt		0.1242	32Mas1
Hg | Hg_2Cl_2, KCl, 0.1 **m** ‖ KCN, 1.1580 **m**, $K_4Fe(CN)_6$, 0.01 **m**, $K_3Fe(CN)_6$, 0.01 **m** | Pt		0.1257	32Mas1
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl,sat. ‖ KI, 0.000978 **M**, I_2,sol. | Pt		0.440	22Keo
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl,sat. ‖ KI, 0.00391 **M**, I_2,sol. | Pt		0.4080	22Keo
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl,sat. ‖ KI, 0.01 **M**, I_2,sol. | Pt		0.3872	22Keo
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl,sat. ‖ KI, 0.02 **M**, I_2,sol. | Pt		0.3780	22Keo
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl,sat. ‖ KI, 0.0625 **M**, I_2,sol. | Pt		0.3456	22Keo
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl,sat. ‖ KI, 0.1 **M**, I_2,sol. | Pt		0.3340	22Keo
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl,sat. ‖ KI, 0.5 **M**, I_2,sol. | Pt		0.2978	22Keo
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl,sat. ‖ KI, 1.0 **M**, I_2,sol. | Pt		0.2892	22Keo
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ CuI,sol., I_2,sol. | Pt	20...21 °C	0.358	10Fed
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KNO_3, 0.1 **M**, PbI_2,sol., I_2,sol. | Pt	20...21 °C	0.400	10Fed
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KCl, 0.5 **M** ‖ HIO_3, 1.0 **M**, I_2,sol. | Pt	18 °C	0.841	05Lut1, 05Lut2
Hg | Hg_2Cl_2, HCl, 0.1 **M** ‖ HCl, 0.1 **M** | air+Cl_2, 0.00293 atm, PtIr		1.0150	11Lew
Hg | Hg_2Cl_2, HCl, 0.1 **M** ‖ HCl, 0.1 **M** | air+Cl_2, 0.00629 atm, PtIr		1.0242	11Lew
Hg | Hg_2Cl_2, HCl, 0.1 **M** ‖ HCl, 0.1 **M** | air+Cl_2, 0.0124 atm, PtIr		1.0330	11Lew
Hg | Hg_2Cl_2, HCl, 0.1 **M** ‖ HCl, 0.1 **M** | air+Cl_2, 0.0249 atm, PtIr		1.0424	11Lew
Hg | Hg_2Cl_2, HCl, 0.1 **M** ‖ HCl, 0.1 **M** | air+Cl_2, 0.0490 atm, PtIr		1.0508	11Lew
Hg | Hg_2Cl_2, HCl, 1.0 **M** ‖ HCl, 1.0 **M** | Cl_2, PtIr(15wt%)	15.0 °C	1.0996	22Ger
	24.9 °C	1.0903	22Ger
	35.0 °C	1.0813	22Ger
Hg | $CrCl_2$, 2.163 **M**, $CrCl_3$, 0.241 **M**, HCl, 0.1 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.801	17For

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Hg | $CrCl_2$, 3.153 **M**, $CrCl_3$, 1.457 **M**, HCl, 0.1 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.757	17For
Hg | $CrCl_2$, 0.774 **M**, $CrCl_3$, 1.200 **M**, HCl, 0.1 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.727	17For
Hg | $CrCl_2$, 0.773 **M**, $CrCl_3$, 3.627 **M**, HCl, 0.1 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.693	17For
Hg | Hg_2Cl_2, $CaCl_2$, 0.002 **M** ‖ $CaCl_2$, 0.001 **M**, Hg_2Cl_2 | Hg	30 °C	0.01853	32Mas2
Hg | Hg_2Cl_2, $CaCl_2$, 0.0025 **M** ‖ $CaCl_2$, 0.00125 **M**, Hg_2Cl_2 | Hg	30 °C	0.01875	32Mas2
Hg | Hg_2Cl_2, $CaCl_2$, 0.0025 **M** ‖ $CaCl_2$, 0.002 **M**, Hg_2Cl_2 | Hg	30 °C	0.00593	32Mas2
Hg | Hg_2Cl_2, $CaCl_2$, 0.005 **M** ‖ $CaCl_2$, 0.0025 **M**, Hg_2Cl_2 | Hg	30 °C	0.01799	32Mas2
Hg | Hg_2Cl_2, $SrCl_2$, 0.002 **M** ‖ $SrCl_2$, 0.001 **M**, Hg_2Cl_2 | Hg	30 °C	0.0185	32Mas2
Hg | Hg_2Cl_2, $SrCl_2$, 0.0025 **M** ‖ $SrCl_2$, 0.00125 **M**, Hg_2Cl_2 | Hg	30 °C	0.0185	32Mas2
Hg | Hg_2Cl_2, $SrCl_2$, 0.0025 **M** ‖ $SrCl_2$, 0.002 **M**, Hg_2Cl_2 | Hg	30 °C	0.00589	32Mas2
Hg | Hg_2Cl_2, $BaCl_2$, 0.002 **M** ‖ $BaCl_2$, 0.001 **M**, Hg_2Cl_2 | Hg	30 °C	0.0238	32Mas2
Hg | Hg_2Cl_2, $BaCl_2$, 0.0025 **M** ‖ $BaCl_2$, 0.001 **M**, Hg_2Cl_2 | Hg	30 °C	0.0179	32Mas2
Hg | Hg_2Cl_2, $BaCl_2$, 0.002 **M** ‖ $BaCl_2$, 0.00125 **M**, Hg_2Cl_2 | Hg	30 °C	0.0175	32Mas2
Hg | Hg_2Cl_2, $BaCl_2$, 0.005 **M** ‖ $BaCl_2$, 0.0025 **M**, Hg_2Cl_2 | Hg	30 °C	0.0172	32Mas2
Hg | Hg_2Cl_2, $BaCl_2$, 0.01 **M** ‖ $BaCl_2$, 0.005 **M**, Hg_2Cl_2 | Hg	30 °C	0.0162	32Mas2
Hg | Hg_2Cl_2, HCl, 0.1 **M** ‖ HCl, 0.01 **M**, Hg_2Cl_2 | Hg		0.0948	27Erd
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl, 0.01 **M**, Hg_2Cl_2 | Hg	25 °C	0.0576	17Lew2
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	25 °C	–0.0529	17Lew2
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg	25 °C	–0.0831	17Lew2
Hg | Hg_2Cl_2, $ZnCl_2$, 0.00125 **M** ‖ $ZnCl_2$, 0.001 **M**, Hg_2Cl_2 | Hg	30 °C	0.00579	32Mas2
Hg | Hg_2Cl_2, $ZnCl_2$, 0.0025 **M** ‖ $ZnCl_2$, 0.00125 **M**, Hg_2Cl_2 | Hg	30 °C	0.0188	32Mas2
Hg | Hg_2Cl_2, $ZnCl_2$, 0.005 **M** ‖ $ZnCl_2$, 0.0025 **M**, Hg_2Cl_2 | Hg	30 °C	0.0184	32Mas2
Hg | Hg_2Cl_2, $ZnCl_2$, 0.01 **M** ‖ $ZnCl_2$, 0.005 **M**, Hg_2Cl_2 | Hg	30 °C	0.0179	32Mas2
Hg | Hg_2I_2,sol., KI, 0.05 **m** ‖ KCl, 0.1 **m**, Hg_2Cl_2 | Hg		0.2927	31McB
Hg | Hg_2I_2,sol., KI, 0.1 **m** ‖ KCl, 0.1 **m**, Hg_2Cl_2 | Hg		0.3094	31McB
Hg | $HgNO_3$, 0.01 **M**, HNO_3, 0.001 **M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.4449	12Dru
Hg | $HgNO_3$, 0.01 **M**, HNO_3, 0.009 **M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.4424	12Dru
Hg | $HgNO_3$, 0.01 **M**, HNO_3, 0.098 **M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		0.4311	12Dru
Hg | Hg_2SO_4, H_2SO_4, 0.05 **m** ‖ H_2SO_4, 0.1 **m**, Hg_2SO_4 | Hg		0.0145	35Ham2
Hg | Hg_2SO_4, H_2SO_4, 0.1 **m** ‖ H_2SO_4, 0.2 **m**, Hg_2SO_4 | Hg		0.0143	35Ham2
Hg | Hg_2SO_4, H_2SO_4, 0.2**m** ‖ H_2SO_4, 0.3 **m**, Hg_2SO_4 | Hg		0.00852	35Ham2
Hg | Hg_2SO_4, H_2SO_4, 0.3 **m** ‖ H_2SO_4, 0.5 **m**, Hg_2SO_4 | Hg		0.0110	35Ham2
Hg | Hg_2SO_4, H_2SO_4, 0.5 **m** ‖ H_2SO_4, 1.0 **m**, Hg_2SO_4 | Hg		0.0163	35Ham2
Hg | Hg_2SO_4, H_2SO_4, 1.0 **m** ‖ H_2SO_4, 2.0 **m**, Hg_2SO_4 | Hg		0.0200	35Ham2
Hg | Hg_2Cl_2, NaCl, 1.0 **M** ‖ NaCl, 1.0 **M**, Na_2IrCl_6, 0.984 m**M**, Na_3IrCl_6, 2.132 m**M** | Ir		0.728	30Ter
Hg | Hg_2Cl_2, NaCl, 1.0 **M** ‖ NaCl, 1.0 **M**, Na_2IrCl_6, 0.984 m**M**, Na_3IrCl_6, 8.757 m**M** | Ir		0.764	30Ter
Hg | Hg_2Cl_2, NaCl, 1.0 **M** ‖ NaCl, 1.0 **M**, Na_2IrCl_6, 0.984 m**M**, Na_3IrCl_6, 18.48 m**M** | Ir		0.777	30Ter
Hg | Hg_2Cl_2, NaCl, 1.0 **M** ‖ NaCl, 1.0 **M**, Na_2IrCl_6, 8.330 m**M**, Na_3IrCl_6, 0.874 m**M** | Ir		0.638	30Ter
Hg | Hg_2Cl_2 | KCl, 1.0 **M** ‖ HNO_3, 0.1 **M**, $Mn(NO_3)_2$, 0.005 **M** | MnO_2	20 °C	0.9813	00Tow
Hg | Hg_2Cl_2 | KCl, 1.0 **M** ‖ HNO_3, 0.00078 **M**, $Mn(NO_3)_2$, 0.005 **M** | MnO_2	20 °C	0.7685	00Tow
Hg | Hg_2Cl_2 | KCl, 1.0 **M** ‖ HNO_3, 0.05 **M**, $Mn(NO_3)_2$, 0.1 **M** | MnO_2	20 °C	0.9442	00Tow
Hg | Hg_2Cl_2 | KCl, 1.0 **M** ‖ HNO_3, 0.05**M**, $Mn(NO_3)_2$, 0.00078 **M** | MnO_2	20 °C	0.9942	00Tow
Hg | Hg_2Cl_2 | KCl, 0.1 **M** ‖ KCl,sat. ‖ H_2SO_4, 0.05 **M**, $KMnO_4$,0.002 **M** | MnO_2		1.145	03Ing
Hg | Hg_2Cl_2 | KCl, 0.1 **M** ‖ KCl,sat. ‖ H_2SO_4, 0.01 **M**, $KMnO_4$, 0.01 **M** | MnO_2		1.115	03Ing

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Hg | Hg_2Cl_2 | KCl, 0.1 **M** ‖ KCl,sat. ‖ H_2SO_4, 0.02 **M**, $KMnO_4$, 0.02 **M** | MnO_2		1.165	03Ing
Hg | Hg_2Cl_2 | KCl, 0.1 **M** ‖ KCl,sat. ‖ H_2SO_4, 0.05 **M**, $KMnO_4$, 0.05 **M** | MnO_2		1.171	03Ing
Hg | Hg_2Cl_2 | KCl, 0.1 **M** ‖ KCl,sat. ‖ H_2SO_4, 0.005 **M**, $MnSO_4$, 0.001 **M** | MnO_2		0.851	03Ing
Hg | Hg_2Cl_2 | KCl, 0.1 **M** ‖ KCl,sat. ‖ H_2SO_4, 0.005 **M**, $MnSO_4$, 0.01 **M** | MnO_2		0.821	03Ing
Hg | Hg_2Cl_2 | KCl, 0.1 **M** ‖ KCl,sat. ‖ H_2SO_4, 0.05 **M**, $MnSO_4$, 0.001 **M** | MnO_2		0.958	03Ing
Hg | Hg_2Cl_2 | KCl, 0.1 **M** ‖ KCl,sat. ‖ H_2SO_4, 0.05 **M**, $MnSO_4$, 0.01 **M** | MnO_2		0.935	03Ing
Hg | Hg_2Cl_2 | KCl, 1.0 **M** ‖ $NiCl_2$, 0.5 **M** | Ni		0.596	09Sch
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl,sat. ‖ $NiSO_4$, 0.005 **M** | Ni		0.6607	28Mur
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl,sat. ‖ $NiSO_4$, 0.05 **M** | Ni		0.6385	28Mur
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KNO_3, 1.0 **M**, $Ni(NO_3)_2$, 0.005 **M**, $Ni(OH)_2$, HgO | Hg		0.182	04Lab
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ KNO_3, 1.0 **M**, $Ni(NO_3)_2$, 0.05 **M**, $Ni(OH)_2$, HgO | Hg		0.266	04Lab
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.2578, HCl, 0.1 **M** | Pt	20 °C	0.4056	32Car
	60 °C	0.4355	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.5066, HCl, 0.1 **M** | Pt	20 °C	0.4339	32Car
	60 °C	0.4652	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.0634, HCl, 1.0 **M** | Pt	20 °C	0.3219	32Car
	60 °C	0.3294	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.2642, HCl, 1.0 **M** | Pt	20 °C	0.3640	32Car
	60 °C	0.3808	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.500, HCl, 1.0 **M** | Pt	20 °C	0.3923	32Car
	60 °C	0.4118	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.7449, HCl, 1.0 **M** | Pt	20 °C	0.3909	32Car
	60 °C	0.4417	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.1579, HCl, 3.0 **M** | Pt	20 °C	0.3037	32Car
	60 °C	0.3022	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.5078, HCl, 3.0 **M** | Pt	20 °C	0.3481	32Car
	60 °C	0.3528	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.7533, HCl, 3.0 **M** | Pt	20 °C	0.3783	32Car
	60 °C	0.3876	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.8762, HCl, 3.0 **M** | Pt	20 °C	0.3971	32Car
	60 °C	0.4055	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.2103, HCl, 4.9 **M** | Pt	20 °C	0.2724	32Car
	60 °C	0.2651	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.4227, HCl, 4.9 **M** | Pt	20 °C	0.2958	32Car
	60 °C	0.2926	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.5206, HCl, 4.9 **M** | Pt	20 °C	0.3090	32Car
	60 °C	0.3084	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.6471, HCl, 4.9 **M** | Pt	20 °C	0.3203	32Car
	60 °C	0.3206	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.6811, HCl, 4.9 **M** | Pt	20 °C	0.3272	32Car
	60 °C	0.3285	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.7750, HCl, 4.9 **M** | Pt	20 °C	0.3368	32Car
	60 °C	0.3387	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.2892, HCl, 6.66 **M** | Pt	20 °C	0.2409	32Car
	60 °C	0.2313	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.4496, HCl, 6.66 **M** | Pt	20 °C	0.2557	32Car
	60 °C	0.2473	32Car

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.5243, HCl, 6.66 **M** | Pt	20 °C	0.2679	32Car
	60 °C	0.2572	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.7829, HCl, 6.66 **M** | Pt	20 °C	0.2977	32Car
	60 °C	0.2934	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.4084, HCl, 10.02 **M** | Pt	20 °C	0.1777	32Car
	60 °C	0.1757	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.5759, HCl, 10.02 **M** | Pt	20 °C	0.1912	32Car
	60 °C	0.1733	32Car
Hg | Hg_2Cl_2, KCl, 1.0 **M** ‖ $FeCl_3$, c_1, $FeCl_2$, c_2, $c_1/(c_1+c_2)$=0.8054, HCl, 10.02 **M** | Pt	20 °C	0.2219	32Car
	60 °C	0.2075	32Car
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl, 0.1 **M**, K_2PtCl_6, 0.520 m**m**, K_2PtCl_4, 0.483 m**m** | Pt	35 °C	0.4585	30Smi
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl, 0.1 **M**, K_2PtCl_6, 0.752 m**m**, K_2PtCl_4, 0.252 m**m** | Pt	35 °C	0.4720	30Smi
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl, 0.1 **M**, K_2PtCl_6, 0.555 m**m**, K_2PtCl_4, 1.956 m**m** | Pt	35 °C	0.4410	30Smi
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl, 0.1 **M**, K_2PtCl_6, 1.251 m**m**, K_2PtCl_4, 1.248 m**m** | Pt	35 °C	0.4581	30Smi
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl, 0.1 **M**, K_2PtCl_6, 2.009 m**m**, K_2PtCl_4, 0.500 m**m** | Pt	35 °C	0.4762	30Smi
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl, 0.1 **M**, K_2PtCl_6, 2.511 m**m**, K_2PtCl_4, 2.488 m**m** | Pt	35 °C	0.4585	30Smi
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl, 0.1 **M**, K_2PtCl_6, 1.017 m**m**, K_2PtCl_4, 3.957 m**m** | Pt	35 °C	0.4404	30Smi
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl, 0.1 **M**, K_2PtCl_6, 1.026 m**m**, K_2PtCl_4, 8.999 m**m** | Pt	35 °C	0.4303	30Smi
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ KCl, 0.1 **M**, K_2PtCl_6, 2.483 m**m**, K_2PtCl_4, 7.504 m**m** | Pt	35 °C	0.4446	30Smi
Hg | Hg_2Cl_2, HCl, 3.5 **m** ‖ HCl, 3.5 **m**, Sb^{III},Sb^{V} | Au on Pt		0.746	49Bro
Hg | Hg_2Cl_2, HCl, 4.5 **m** ‖ HCl, 4.5 **m**, Sb^{III},Sb^{V} | Au on Pt		0.748	49Bro
Hg | Hg_2Cl_2, HCl, 6.0 **m** ‖ HCl, 6.0 **m**, Sb^{III},Sb^{V} | Au on Pt		0.818	49Bro
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ HCl, 0.1053 **M**, TlCl, 0.161 m**M**, $TlCl_3$, 0.0087 **M** | Pt		0.5723	05Abe
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ HCl, 0.1053 **M**, TlCl, 0.644 m**M**, $TlCl_3$, 0.0087 **M** | Pt		0.5570	05Abe
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ HCl, 0.1053 **M**, TlCl, 3.22 m**M**, $TlCl_3$, 0.0087 **M** | Pt		0.5393	05Abe
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ HCl, 0.001 **M**, KBr, 0.1003 **M** | Br_2, x=0.0122 in CCl_4 | Pt		0.8009	17Lew1
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ HCl, 0.001 **M**, KBr, 0.03009 **M** | Br_2, x=0.0125 in CCl_4 | Pt		0.8276	17Lew1
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ HCl, 0.001 **M**, KBr, 0.01003 **M** | Br_2, x=0.0126 in CCl_4 | Pt		0.8519	17Lew1
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ HCl, 0.001 **M**, KBr, 0.1003 **M** | Br_2, x=0.00463 in CCl_4 | Pt		0.7886	17Lew1
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ HCl, 0.001 **M**, KBr, 0.03009 **M** | Br_2, x=0.0047 in CCl_4 | Pt		0.8163	17Lew1
Hg | Hg_2Cl_2, KCl, 0.1 **M** ‖ HCl, 0.001 **M**, KBr, 0.01003 **M** | Br_2, x=0.0050 in CCl_4 | Pt		0.842	17Lew1

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Hg | Hg_2SO_4, H_2SO_4, 1.0 **M** ‖ H_2SO_4, 1.0 **M**, $CoSO_4$, 0.465 **M**, $Co_2(SO_4)_3$, 0.0173 **M** | Pt	0 °C	1.03	03Obe
Hg | Hg_2SO_4, H_2SO_4, 1.0 **M** ‖ H_2SO_4, 1.0 **M**, $CoSO_4$, 0.0580 **M**, $Co_2(SO_4)_3$, 0.0175 **M** | Pt	0 °C	1.07	03Obe
Hg | Hg_2SO_4, H_2SO_4, 1.0 **M** ‖ H_2SO_4, 1.0 **M**, $CoSO_4$, 0.0881 **M**, $Co_2(SO_4)_3$, 0.0486 **M** | Pt	0 °C	1.09	03Obe
Hg | Hg_2Cl_2, NH_4Cl, 0.5 **M** ‖ NH_4Cl, 0.5 **M** ‖ NH_3, H_2 | Pt (hydrogen passed through aqueous solution of 0.05 **M** NH_3)	18 °C	0.8200	10Brö
Hg | Hg_2Cl_2, NH_4Cl,sat. ‖ NH_4Cl,sat. ‖ NH_3, H_2 | Pt (hydrogen passed through aqueous solution of 0.05 **M** NH_3)	18 °C	0.6924	10Brö
Hg | Hg_2SO_4, Li_2SO_4, 0.025 **M** ‖ Li_xHg | Li_2SO_4, 0.05 **M**, Hg_2SO_4 | Hg		–0.0205	26Ake
Hg | Hg_2SO_4, Li_2SO_4, 0.5 **M** ‖ Li_xHg | Li_2SO_4, 0.05 **M**, Hg_2SO_4 | Hg		0.0684	26Ake
Hg | Hg_2SO_4, Na_2SO_4, 0.025 **M** ‖ Na_xHg | Na_2SO_4, 0.05 **M**, Hg_2SO_4 | Hg		–0.0198	26Ake
Hg | Hg_2SO_4, Na_2SO_4, 0.25 **M** ‖ Na_xHg | Na_2SO_4, 0.05 **M**, Hg_2SO_4 | Hg		0.0452	26Ake
Hg | Hg_2SO_4, Na_2SO_4, 1.0 **M** ‖ Na_xHg | Na_2SO_4, 0.05 **M**, Hg_2SO_4 | Hg		0.0809	26Ake
Hg | $K_2Hg(CN)_4$, 4.355 m**M**, KCN, 2.54 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	12 °C	0.520	32Bri
Hg | $K_2Hg(CN)_4$, 8.33 m**M**, KCN, 7.72 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	12 °C	0.459	32Bri
Hg | $K_2Hg(CN)_4$, 62.5 m**M**, KCN, 5.79 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	12 °C	0.455	32Bri
Hg | $K_2Hg(CN)_4$, 57.1 m**M**, KCN, 13.7 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	12 °C	0.497	32Bri
Hg | Nb^{3+}, 0.238 m**m**, Nb^{5+}, 23.76 m**m**, H_2SO_4, 3.14 **M** ‖ KCl,sat. ‖ KCl,sat., Hg_2Cl_2 | Hg		0.3125	28Kie
Hg | Nb^{3+}, 0.470 m**m**, Nb^{5+}, 23.53 m**m**, H_2SO_4, 3.14 **M** ‖ KCl,sat. ‖ KCl,sat., Hg_2Cl_2 | Hg		0.3225	28Kie
Hg | Nb^{3+}, 1.14 m**m**, Nb^{5+}, 22.85 m**m**, H_2SO_4, 3.14 **M** ‖ KCl,sat. ‖ KCl,sat., Hg_2Cl_2 | Hg		0.3348	28Kie
Hg | Nb^{3+}, 2.2 m**m**, Nb^{5+}, 21.81 m**m**, H_2SO_4, 3.14 **M** ‖ KCl,sat. ‖ KCl,sat., Hg_2Cl_2 | Hg		0.3443	28Kie
Hg | Nb^{3+}, 7.1 m**m**, Nb^{5+}, 16.90 m**m**, H_2SO_4, 3.14 **M** ‖ KCl,sat. ‖ KCl,sat., Hg_2Cl_2 | Hg		0.3618	28Kie
Hg | Nb^{3+}, 12 m**m**, Nb^{5+}, 12 m**m**, H_2SO_4, 3.14 **M** ‖ KCl,sat. ‖ KCl,sat., Hg_2Cl_2 | Hg		0.3730	28Kie
Hg | Nb^{3+}, 16.73 m**m**, Nb^{5+}, 7.27 m**m**, H_2SO_4, 3.14 **M** ‖ KCl,sat. ‖ KCl,sat., Hg_2Cl_2 | Hg		0.3840	28Kie
Hg | Nb^{3+}, 21.55 m**m**, Nb^{5+}, 2.4 m**m**, H_2SO_4, 3.14 **M** ‖ KCl,sat. ‖ KCl,sat., Hg_2Cl_2 | Hg		0.4014	28Kie
Hg | Nb^{3+}, 22.72 m**m**, Nb^{5+}, 1.27 m**m**, H_2SO_4, 3.14 **M** ‖ KCl,sat. ‖ KCl,sat., Hg_2Cl_2 | Hg		0.4102	28Kie
Hg | Nb^{3+}, 23.5 m**m**, Nb^{5+}, 0.515 m**m**, H_2SO_4, 3.14 **M** ‖ KCl,sat. ‖ KCl,sat., Hg_2Cl_2 | Hg		0.4211	28Kie
Hg | Hg_2Cl_2, KCl,sat. ‖ HCl, 2.5 **M**, $TeCl_4$, 0.402 m**M** | Te		0.2724	33Get
Hg | Hg_2Cl_2, KCl,sat. ‖ HCl, 2.5 **M**, $TeCl_4$, 0.902 m**M** | Te		0.2754	33Get
Hg | Hg_2Cl_2, KCl,sat. ‖ HCl, 2.5 **M**, $TeCl_4$, 2.16 m**M** | Te		0.2828	33Get
Hg | Hg_2Cl_2, KCl,sat. ‖ HCl, 2.5 **M**, $TeCl_4$, 4.02 m**M** | Te		0.2867	33Get
Hg | Hg_2Cl_2, KCl,sat. ‖ HCl, 2.5 **M**, $TeCl_4$, 8.04 m**M** | Te		0.2913	33Get
Hg | Hg_2Cl_2, KCl,sat. ‖ HCl, 2.5 **M**, $TeCl_4$, 11.5 m**M** | Te		0.2933	33Get

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Hg | Hg_2Cl_2 sat. acetic acid | chloranil, anhydrous acetic acid | Pt	298 K	0.9095	56Bru
Hg | Hg_2Cl_2 sat. acetic acid | chloride solution in anhydrous acetic acid | AgCl | Ag	298 K	0.3928	56Bru
Hg | Hg_2Cl_2 | KCl,sat., formic acid ‖ quinhydrone, 0.05 **M**, sodium formate 0.25 **M** in formic acid | Pt	298 K	0.5384	56Pin
Palladium (Pd)			
Pd | $PdCl_2$, HCl, 0.9952 **M** ‖ HCl | H_2, Pt		0.623	43Tem
Pd | $PdCl_2$, HCl, $HClO_4$ ‖ HCl | H_2, Pt (formation of $PdCl_4^{2-}$)		0.629	43Tem
Pd | $PdCl_2$, $HClO_4$, 4.0 **M** ‖ $HClO_4$ | H_2, Pt		0.987	43Tem
Platinum (Pt)			
Pt | $K_2Fe(C_2O_4)_2$, 0.0425 **M**, $K_3Fe(C_2O_4)_2$, 0.075 **M**, $K_2C_2O_4$ | KCl, 1.0 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg	17 °C	0.274	10Sch1
Pt | $K_2Fe(C_2O_4)_2$, 0.025 **M**, $K_3Fe(C_2O_4)_2$, 0.025 **M**, $K_2C_2O_4$ | KCl, 1.0 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg	17 °C	0.239	10Sch1
Pt | $K_2Fe(C_2O_4)_2$, 0.0075 **M**, $K_3Fe(C_2O_4)_2$, 0.0425 **M**, $K_2C_2O_4$ | KCl, 1.0 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg	17 °C	0.195	10Sch1
Pt | $K_2Fe(C_2O_4)_2$, 0.00213 **M**, $K_3Fe(C_2O_4)_2$, 0.025 **M**, $K_2C_2O_4$ | KCl, 0.1 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg	17 °C	0.195	10Sch1
Pt | $K_2Fe(C_2O_4)_2$, 0.00085 **M**, $K_3Fe(C_2O_4)_2$, 0.004 **M**, $K_2C_2O_4$ | KCl, 0.1 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg	17 °C	0.160	10Sch1
Pt | $K_3Mo(CN)_8$, 0.0853 **M**, $K_4Mo(CN)_8$, 0.0147 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	0 °C	0.6062	27Gru
Pt | $K_3Mo(CN)_8$, 0.050 **M**, $K_4Mo(CN)_8$, 0.05 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	0 °C	0.5643	27Gru
Pt | $K_3Mo(CN)_8$, 0.010 **M**, $K_4Mo(CN)_8$, 0.09 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	0 °C	0.5118	27Gru
Pt | $SnCl_2$, 0.05 **M**, $SnCl_4$, 0.45 **M**, NH_4Cl, 1.0 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	18 °C	0.122	11För
Pt | $SnCl_2$, 0.45 **M**, $SnCl_4$, 0.05 **M**, NH_4Cl, 1.0 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	18 °C	0.181	11För
Pt | $SnCl_2$, 0.05 **M**, $SnCl_4$, 0.45 **M**, HCl, 0.5 **M** ‖ NH_4NO_3,sat. ‖ H_2SO_4, 0.5 **M**, Hg_2SO_4 | Hg	18 °C	0.506	11För
Pt | $SnCl_2$, 0.45 **M**, $SnCl_4$, 0.05 **M**, HCl, 0.5 **M** ‖ NH_4NO_3,sat. ‖ H_2SO_4, 0.5 **M**, Hg_2SO_4 | Hg	18 °C	0.552	11För
Pt | UO_2SO_4, 0.0131 **M**, $U(SO_4)_2$, 0.0069 **M**, H_2SO_4, 0.052 **M** ‖ H_2SO_4, 0.052 **M**, Hg_2SO_4 | Hg		0.403	08Lut
Pt | UO_2SO_4, 0.0127 **M**, $U(SO_4)_2$, 0.0273 **M**, H_2SO_4, 0.053 **M** ‖ H_2SO_4, 0.053 **M**, Hg_2SO_4 | Hg		0.418	08Lut
Pt | UO_2SO_4, 0.0941 **M**, $U(SO_4)_2$, 0.0259 **M**, H_2SO_4, 0.054 **M** ‖ H_2SO_4, 0.054 **M**, Hg_2SO_4 | Hg		0.393	08Lut
Pt | UO_2SO_4, 0.0917 **M**, $U(SO_4)_2$, 0.0883 **M**, H_2SO_4, 0.050 **M** ‖ H_2SO_4, 0.050 **M**, Hg_2SO_4 | Hg		0.407	08Lut
Pt | UO_2SO_4, 0.0911 **M**, $U(SO_4)_2$, 0.00886 **M**, H_2SO_4, 0.050 **M** ‖ H_2SO_4, 0.050 **M**, Hg_2SO_4 | Hg		0.380	08Lut
Pt | U^{VI}, U^{V}, HCl, 0.1 **m**, AgCl | Ag		0.55	49Kra
Pt | U^{IV}, U^{V}, HCl, 0.1 **m**, AgCl | Ag		0.31	49Kra
Pt | $K_3W(CN)_8$, 8.87 m**M**, $K_4W(CN)_8$, 0.986 m**M**, KCl, 1.0 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	0 °C	0.3733	24Col

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Pt | $K_3W(CN)_8$, 4.83 m**M**, $K_4W(CN)_8$, 4.83 m**M**, KCl, 1.0 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	0 °C	0.3218	24Col
Pt | $K_3W(CN)_8$, 0.986 m**M**, $K_4W(CN)_8$, 8.87 m**M**, KCl, 1.0 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	0 °C	0.2702	24Col
Pt | $K_3W(CN)_8$, 9.64 m**M**, $K_4W(CN)_8$, 86.7 m**M**, KCl, 0.0 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	0 °C	0.2426	24Col
Pt | $K_3W(CN)_8$, 95.5 m**M**, $K_4W(CN)_8$, 10.6 m**M**, KCl, 0.0 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	0 °C	0.3481	24Col
Pt | V^{III}, 0.0138 **N**, V^{IV}, 0.0138 **N**, H_2SO_4, 1.5 **M** ‖ H_2SO_4, 1.5 **M**, Hg_2SO_4 | Hg	18 °C	0.299	30Foe
Pt | V^{III}, 0.125 **N**, V^{IV}, 0.125 **N**, H_2SO_4, 0.375 **M** ‖ H_2SO_4, 1.0 **M**, Hg_2SO_4 | Hg	18 °C	0.377	30Foe
Pt | V^{III}, 0.125 **N**, V^{IV}, 0.125 **N**, H_2SO_4, 0.625 **M** ‖ H_2SO_4, 1.0 **M**, Hg_2SO_4 | Hg	18 °C	0.354	30Foe
Pt | V^{III}, 0.125 **N**, V^{IV}, 0.125 **N**, H_2SO_4, 1.5 **M** ‖ H_2SO_4, 1.5 **M**, Hg_2SO_4 | Hg	18 °C	0.299	30Foe
Pt | V^{IV}, 0.0138 **N**, V^{V}, 0.0138 **N**, H_2SO_4, 0.375 **M** ‖ H_2SO_4, 1.5 **M**, Hg_2SO_4 | Hg	18 °C	0.315	30Foe
Pt | V^{IV}, 0.0138 **N**, V^{V}, 0.0138 **N**, H_2SO_4, 0.625 **M** ‖ H_2SO_4, 1.5 **M**, Hg_2SO_4 | Hg	18 °C	0.327	30Foe
Pt | V^{IV}, 0.0138 **N**, V^{V}, 0.0138 **N**, H_2SO_4, 1.5 **M** ‖ H_2SO_4, 1.5 **M**, Hg_2SO_4 | Hg	18 °C	0.384	30Foe
Pt | V^{IV}, 0.125 **N**, V^{V}, 0.125 **N**, H_2SO_4, 0.375 **M** ‖ H_2SO_4, 1.0 **M**, Hg_2SO_4 | Hg	18 °C	0.316	30Foe
Pt | V^{IV}, 0.125 **N**, V^{V}, 0.125 **N**, H_2SO_4, 0.625 **M** ‖ H_2SO_4, 1.0 **M**, Hg_2SO_4 | Hg	18 °C	0.327	30Foe
Pt | V^{IV}, 0.125 **N**, V^{V}, 0.125 **N**, H_2SO_4,1.5 **M** ‖ H_2SO_4,1.5 **M**, Hg_2SO_4 | Hg	18 °C	0.384	30Foe
Pt | TlCl, 0.1814 m**M**, $TlCl_3$, 0.01886 m**M**, HCl+$HClO_4$, 0.9952 **M** ‖ HCl, 0.9952 **M** | H_2, Pt		0.7593	42Hug
Pt | TlCl, 0.1016 m**M**, $TlCl_3$, 0.09430 m**M**, HCl+$HClO_4$, 0.9952 **M** ‖ HCl, 0.9952 **M** | H_2, Pt		0.7862	42Hug
Pt | TlCl, 0.1097 m**M**, $TlCl_3$, 1.053 m**M**, HCl+$HClO_4$, 0.9952 **M** ‖ HCl, 0.9952 **M** | H_2, Pt		0.8148	42Hug
Pt | TlCl, 0.1452 m**M**, $TlCl_3$, 4.715 m**M**, HCl+$HClO_4$, 0.9952 **M** ‖ HCl, 0.9952 **M** | H_2, Pt		0.8328	42Hug
Pt | TlCl, 0.1806 m**M**, $TlCl_3$, 1.886 m**M**, HCl+$HClO_4$, 0.9952 **M** ‖ HCl, 0.9952 **M** | H_2, Pt		0.8468	42Hug
Pt | TlCl, 0.01806 m**M**, $TlCl_3$, 1.886 m**M**, HCl+$HClO_4$, 0.9952 **M** ‖ HCl, 0.9952 **M** | H_2, Pt		0.8468	42Hug
Potassium (K)			
K_zHg | KCl, 0.001 **m** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		2.3121	30Bir
K_zHg | KCl, 0.01 **m** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		2.2784	30Bir
K_zHg | KCl, 0.02 **m** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		2.2610	30Bir
K_zHg | KCl, 0.05 **m** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		2.2414	30Bir
K_zHg | KCl, 0.1 **m** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		2.2246	30Bir
K_zHg | KCl, 0.2 **m** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		2.2080	30Bir
K_zHg | KCl, 0.5 **m** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		2.1875	30Bir
K_zHg | KCl, 1.0 **m** ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		2.1818	30Bir
$K_{x=0.002559}$Hg | KCl, 0.10245 **m**, Hg_2Cl_2 | Hg		2.2217	34Arm

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
$K_{x=0.002559}Hg \mid KCl$, 1.0168 **m**, $Hg_2Cl_2 \mid Hg$		2.0704	34Arm
$K_{0.1wt\%}Hg \mid KI$, 0.002 **m**, AgI $\mid Ag \mid$ AgI, KI, 0.02 **m** $\mid K_{0.1wt\%}Hg$		0.1155	33Gel
$K_{0.1wt\%}Hg \mid KI$, 0.005 **m**, AgI $\mid Ag \mid$ AgI, KI, 0.05 **m** $\mid K_{0.1wt\%}Hg$		0.1108	33Gel
$K_{0.1wt\%}Hg \mid KI$, 0.01 **m**, AgI $\mid Ag \mid$ AgI, KI, 0.05 **m** $\mid K_{0.1wt\%}Hg$		0.0762	33Gel
$K_{0.1wt\%}Hg \mid KI$, 0.01 **m**, AgI $\mid Ag \mid$ AgI, KI, 0.12 **m** $\mid K_{0.1wt\%}Hg$		0.1093	33Gel
$K_{0.1wt\%}Hg \mid KI$, 0.02 **m**, AgI $\mid Ag \mid$ AgI, KI, 0.1 **m** $\mid K_{0.1wt\%}Hg$		0.0765	33Gel
$K_{0.1wt\%}Hg \mid KI$, 0.02 **m**, AgI $\mid Ag \mid$ AgI, KI, 0.2 **m** $\mid K_{0.1wt\%}Hg$		0.1087	33Gel
$K_{0.1wt\%}Hg \mid KI$, 0.05 **m**, AgI $\mid Ag \mid$ AgI, KI, 0.5 **m** $\mid K_{0.1wt\%}Hg$		0.1082	33Gel
$K_{0.1wt\%}Hg \mid KI$, 0.02 **m** $\Vert$ KI, 0.002 **m** $\mid K_{0.1wt\%}Hg$		0.0559	33Gel
$K_{0.1wt\%}Hg \mid KI$, 0.05 **m** $\Vert$ KI, 0.005 **m** $\mid K_{0.1wt\%}Hg$		0.0549	33Gel
$K_{0.1wt\%}Hg \mid KI$, 0.05 **m** $\Vert$ KI, 0.01 **m** $\mid K_{0.1wt\%}Hg$		0.0380	33Gel
$K_{0.1wt\%}Hg \mid KI$, 0.1 **m** $\Vert$ KI, 0.01 **m** $\mid K_{0.1wt\%}Hg$		0.0552	33Gel
$K_{0.1wt\%}Hg \mid KI$, 0.1 **m** $\Vert$ KI, 0.02 **m** $\mid K_{0.1wt\%}Hg$		0.0388	33Gel
$K_{0.1wt\%}Hg \mid KI$, 0.2 **m** $\Vert$ KI, 0.02 **m** $\mid K_{0.1wt\%}Hg$		0.0551	33Gel
$K_{0.1wt\%}Hg \mid KI$, 0.5 **m** $\Vert$ KI, 0.05 **m** $\mid K_{0.1wt\%}Hg$		0.0551	33Gel
$K_{0.2216wt\%}Hg \mid KOH$, 0.2026 **m** $\Vert$ KCl, 1.0 **m**, $Hg_2Cl_2 \mid Hg$		2.1891	12Lew
Rubidium (Rb)			
$Rb_zHg \mid RbOH$, 0.1 **m** $\Vert$ RbCl, 0.1 **m**, $Hg_2Cl_2 \mid Hg$		2.1805	15Lew
$Rb_zHg \mid RbOH$, 0.1 **m**, RbCl, 0.1 **m** $\Vert$ KCl, 1.0 **m**, $Hg_2Cl_2 \mid Hg$		2.1805	15Lew
Silver (Ag)			
$Ag \mid Ag_2SO_4$, K_2SO_4,sat. $\Vert$ K_2SO_4,sat., $Hg_2SO_4 \mid Hg$		0.0407	33Ish1
$Ag \mid$ AgBr, KBr, 0.05 **m** $\mid K_xHg \mid$ KBr, 0.1 **m**, AgBr $\mid Ag$		0.0323	29Har2
$Ag \mid$ AgBr, KBr, 0.1 **m** $\mid K_xHg \mid$ KBr, 0.1 **m**, AgBr $\mid Ag$		0.0	29Har2
$Ag \mid$ AgBr, KBr, 0.2 **m** $\mid K_xHg \mid$ KBr, 0.1 **m**, AgBr $\mid Ag$		–0.0321	29Har2
$Ag \mid$ AgBr, KBr, 0.5 **m** $\mid K_xHg \mid$ KBr, 0.1 **m**, AgBr $\mid Ag$		–0.0746	29Har2
$Ag \mid$ AgBr, KBr, 1.0 **m** $\mid K_xHg \mid$ KBr, 0.1 **m**, AgBr $\mid Ag$		–0.1069	29Har2
$Ag \mid$ AgBr, KBr, 1.5 **m** $\mid K_xHg \mid$ KBr, 0.1 **m**, AgBr $\mid Ag$		–0.1262	29Har2
$Ag \mid$ AgBr, KBr, 2.0 **m** $\mid K_xHg \mid$ KBr, 0.1 **m**, AgBr $\mid Ag$		–0.1405	29Har2
$Ag \mid$ AgBr, KBr, 2.5 **m** $\mid K_xHg \mid$ KBr, 0.1 **m**, AgBr $\mid Ag$		–0.1510	29Har2
$Ag \mid$ AgBr, KBr, 3.0 **m** $\mid K_xHg \mid$ KBr, 0.1 **m**, AgBr $\mid Ag$		–0.1614	29Har2
$Ag \mid$ AgBr, KBr, 3.5 **m** $\mid K_xHg \mid$ KBr, 0.1 **m**, AgBr $\mid Ag$		–0.1701	29Har2
$Ag \mid$ AgBr, KBr, 4.0 **m** $\mid K_xHg \mid$ KBr, 0.1 **m**, AgBr $\mid Ag$		–0.1778	29Har2
$Ag \mid$ AgCl, $1.43 \cdot 10^{-4}$ **M**, KCl, 1.824 **M** $\Vert$ KNO_3,sat. $\Vert$ $AgNO_3$, 0.01 **M** $\mid Ag$		0.4591	15For
$Ag \mid$ AgCl, $1.43 \cdot 10^{-4}$ **M**, KCl, 2.968 **M** $\Vert$ KNO_3,sat. $\Vert$ $AgNO_3$, 0.01 **M** $\mid Ag$		0.5100	15For
$Ag \mid$ AgCl, $1.43 \cdot 10^{-4}$ **M**, KCl, 4.064 **M** $\Vert$ KNO_3,sat. $\Vert$ $AgNO_3$, 0.01 **M** $\mid Ag$		0.5385	15For
$Ag \mid$ AgCl, 4.7006 m**m** $\Vert$ HCl, 0.100313 **m**, AgCl $\mid Ag$		0.12341	39Har4
(U_T [V]=0.12341+2.48·10^{-4}(T–25 °C)–0.56·10^{-6}(T– 25 °C)^2)			
$Ag \mid$ AgCl, 6.188 m**m** $\Vert$ HCl, 0.100313 **m**, AgCl $\mid Ag$		0.11217	39Har4
(U_T [V]=0.11217+2.25·10^{-4}(T–25 °C)–0.51·10^{-6}(T–25 °C)^2)			
$Ag \mid$ AgCl, 18.028 m**m** $\Vert$ HCl, 0.100313 **m**, AgCl $\mid Ag$		0.06885	39Har4
(U_T [V]=0.06885+1.38·10^{-4}(T–25 °C)–0.31·10^{-6}(T–25 °C)^2)			
$Ag \mid$ AgCl, 58.486 m**m** $\Vert$ HCl, 0.100313 **m**, AgCl $\mid Ag$		0.02171	39Har4
(U_T [V]=0.02171+0.44·10^{-4}(T–25 °C)–0.10·10^{-6}(T–25 °C)^2)			
$Ag \mid$ AgCl, 0.179939 **m** $\Vert$ HCl, 0.100313 **m**, AgCl $\mid Ag$		–0.02351	39Har4
(–U_T [V]=0.02351+0.48·10^{-4}(T–25 °C)+0.10·10^{-6}(T–25 °C)^2)			
$Ag \mid$ AgCl, 0.55472 **m** $\Vert$ HCl, 0.100313 **m**, AgCl $\mid Ag$		–0.07131	39Har4
(–U_T [V]=0.07131+1.38·10^{-4}(T–25 °C)+0.31·10^{-6}(T–25 °C)^2)			

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Ag | AgCl, 0.97192 **m** ‖ HCl, 0.100313 **m**, AgCl | Ag		–0.09787	39Har4
($-U_T$ [V]=0.09787+1.78·10^{-4}(T–25 °C)+0.41·10^{-6}(T–25 °C)^2)			
Ag | AgCl, 1.69357 **m** ‖ HCl, 0.100313 **m**, AgCl | Ag		–0.12844	39Har4
($-U_T$ [V]=0.12844+2.15·10^{-4}(T–25 °C)+0.51·10^{-6}(T–25 °C)^2)			
Ag | AgCl, 2.99000 **m** ‖ HCl, 0.100313 **m**, AgCl | Ag		–0.16695	39Har4
($-U_T$ [V]=0.16695+2.50·10^{-4}(T–25 °C)+0.62·10^{-6}(T–25 °C)^2)			
Ag | AgCl, KCl, 0.05 **m** | K_xHg | KCl, 0.1 **m**, AgCl | Ag		0.0323	29Har2
Ag | AgCl, KCl, 0.1 **m** | K_xHg | KCl, 01 **m**, AgCl | Ag		0.0	29Har2
Ag | AgCl, KCl, 0.2 **m** | K_xHg | KCl, 0.1 **m**, AgCl | Ag		–0.0319	29Har2
Ag | AgCl, KCl, 0.5 **m** | K_xHg | KCl, 0.1 **m**, AgCl | Ag		–0.0739	29Har2
Ag | AgCl, KCl, 1.0 **m** | K_xHg | KCl, 0.1 **m**, AgCl | Ag		–0.1056	29Har2
Ag | AgCl, KCl, 1.5 **m** | K_xHg | KCl, 0.1 **m**, AgCl | Ag		–0.1246	29Har2
Ag | AgCl, KCl, 2.0 **m** | K_xHg | KCl, 0.1 **m**, AgCl | Ag		–0.1388	29Har2
Ag | AgCl, KCl, 2.5 **m** | K_xHg | KCl, 0.1 **m**, AgCl | Ag		–0.1502	29Har2
Ag | AgCl, KCl, 3.0 **m** | K_xHg | KCl, 0.1 **m**, AgCl | Ag		–0.1598	29Har2
Ag | AgCl, KCl, 3.5 **m** | K_xHg | KCl, 0.1 **m**, AgCl | Ag		–0.1677	29Har2
Ag | AgCl, KCl, 4.0 **m** | K_xHg | KCl, 0.1 **m**, AgCl | Ag		–0.1755	29Har2
Ag | AgBr, $CdBr_2$, 0.1 **m** | $CdBr_2$, 0.01 **m**, AgBr | Ag		0.0229	29Luc2
Ag | AgBr, $CdBr_2$, 0.1 **m** | $CdBr_2$, 0.01396 **m**, AgBr | Ag		0.0194	29Luc2
Ag | AgBr, $CdBr_2$, 0.1 **m** | $CdBr_2$, 0.02801 **m**, AgBr | Ag		0.0119	29Luc2
Ag | AgBr, $CdBr_2$, 0.1 **m** | $CdBr_2$, 0.03997 **m**, AgBr | Ag		0.00832	29Luc2
Ag | AgBr, $CdBr_2$, 0.1 **m** | $CdBr_2$, 0.05629 **m**, AgBr | Ag		0.00508	29Luc2
Ag | AgBr, $CdBr_2$, 0.1 **m** | $CdBr_2$, 0.08114 **m**, AgBr | Ag		0.00176	29Luc2
Ag | AgBr, $CdBr_2$, 0.1 **m** | $CdBr_2$, 0.2026 **m**, AgBr | Ag		–0.00529	29Luc2
Ag | AgBr, $CdBr_2$, 0.1 **m** | $CdBr_2$, 0.4069 **m**, AgBr | Ag		–0.00915	29Luc2
Ag | AgBr, $CdBr_2$, 0.1 **m** | $CdBr_2$, 0.9996 **m**, AgBr | Ag		–0.0117	29Luc2
Ag | AgBr, $CdBr_2$, 0.1 **m** | $CdBr_2$, 3.0815 **m**, AgBr | Ag		–0.0100	29Luc2
Ag | AgBr, KBr, 0.001 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	18 °C	0.0342	21Kol
Ag | AgBr, KBr, 0.01 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	18 °C	0.0851	21Kol
Ag | AgBr, KBr, 0.1 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	18 °C	0.1414	21Kol
Ag | AgBr, LiBr, 0.1 **m** | Li_xHg | LiBr, 0.1 **m**, AgBr | Ag		0.00	29Har2
Ag | AgBr, LiBr, 0.1 **m** | Li_xHg | LiBr, 0.2 **m**, AgBr | Ag		0.0340	29Har2
Ag | AgBr, LiBr, 0.1 **m** | Li_xHg | LiBr, 0.5 **m**, AgBr | Ag		0.0803	29Har2
Ag | AgBr, LiBr, 0.1 **m** | Li_xHg | LiBr, 1.0 **m**, AgBr | Ag		0.1191	29Har2
Ag | AgBr, LiBr, 0.1 **m** | Li_xHg | LiBr, 2.0 **m**, AgBr | Ag		0.1657	29Har2
Ag | AgBr, LiBr, 0.1 **m** | Li_xHg | LiBr, 2.5 **m**, AgBr | Ag		0.1845	29Har2
Ag | AgBr, LiBr, 0.1 **m** | Li_xHg | LiBr, 3.0 **m**, AgBr | Ag		0.1998	29Har2
Ag | AgBr, LiBr, 0.1 **m** | Li_xHg | LiBr, 3.5 **m**, AgBr | Ag		0.2160	29Har2
Ag | AgBr, LiBr, 0.1 **m** | Li_xHg | LiBr, 4.0 **m**, AgBr | Ag		0.2324	29Har2
Ag | AgCl, LiCl, 0.1 **m** | Li_xHg | LiCl, 0.1 **m**, AgCl | Ag		0.00	29Har2
Ag | AgCl, LiCl, 0.1 **m** | Li_xHg | LiCl, 0.2 **m**, AgCl | Ag		0.0336	29Har2
Ag | AgCl, LiCl, 0.1 **m** | Li_xHg | LiCl, 0.5 **m**, AgCl | Ag		0.0790	29Har2
Ag | AgCl, LiCl, 0.1 **m** | Li_xHg | LiCl, 1.0 **m**, AgCl | Ag		0.1168	29Har2
Ag | AgCl, LiCl, 0.1 **m** | Li_xHg | LiCl, 2.0 **m**, AgCl | Ag		0.1624	29Har2
Ag | AgCl, LiCl, 0.1 **m** | Li_xHg | LiCl, 2.5 **m**, AgCl | Ag		0.1862	29Har2
Ag | AgCl, LiCl, 0.1 **m** | Li_xHg | LiCl, 3.0 **m**, AgCl | Ag		0.1958	29Har2
Ag | AgCl, LiCl, 0.1 **m** | Li_xHg | LiCl, 3.5 **m**, AgCl | Ag		0.2112	29Har2
Ag | AgCl, LiCl, 0.1 **m** | Li_xHg | LiCl, 4.0 **m**, AgCl | Ag		0.2250	29Har2

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Ag | AgBr, NaBr, 0.1 **m** | Na_xHg | NaBr, 0.1 **m**, AgBr | Ag		0.00	29Har2
Ag | AgBr, NaBr, 0.1 **m** | Na_xHg | NaBr, 0.2 **m**, AgBr | Ag		0.0327	29Har2
Ag | AgBr, NaBr, 0.1 **m** | Na_xHg | NaBr, 0.5 **m**, AgBr | Ag		0.0769	29Har2
Ag | AgBr, NaBr, 0.1 **m** | Na_xHg | NaBr, 1.0 **m**, AgBr | Ag		0.1118	29Har2
Ag | AgBr, NaBr, 0.1 **m** | Na_xHg | NaBr, 1.5 **m**, AgBr | Ag		0.1333	29Har2
Ag | AgBr, NaBr, 0.1 **m** | Na_xHg | NaBr, 2.0 **m**, AgBr | Ag		0.1507	29Har2
Ag | AgBr, NaBr, 0.1 **m** | Na_xHg | NaBr, 2.5 **m**, AgBr | Ag		0.1650	29Har2
Ag | AgBr, NaBr, 0.1 **m** | Na_xHg | NaBr, 3.0 **m**, AgBr | Ag		0.1775	29Har2
Ag | AgBr, NaBr, 0.1 **m** | Na_xHg | NaBr, 3.5 **m**, AgBr | Ag		0.1884	29Har2
Ag | AgBr, NaBr, 0.1 **m** | Na_xHg | NaBr, 4.0 **m**, AgBr | Ag		0.1987	29Har2
Ag | AgCl, $BaCl_2$, 0.03016 **m** | Ba_xHg | $BaCl_2$, 0.01 **m**, AgCl | Ag		0.0355	25Luc
Ag | AgCl, $BaCl_2$, 0.1004 **m** | Ba_xHg | $BaCl_2$, 0.01 **m**, AgCl | Ag		0.0742	25Luc
Ag | AgCl, $BaCl_2$, 0.2978 **m** | Ba_xHg | $BaCl_2$, 0.01 **m**, AgCl | Ag		0.1092	25Luc
Ag | AgCl, $BaCl_2$, 0.9710 **m** | Ba_xHg | $BaCl_2$, 0.01 **m**, AgCl | Ag		0.1532	25Luc
Ag | AgCl, $BaCl_2$, 1.521 **m** | Ba_xHg | $BaCl_2$, 0.01 **m**, AgCl | Ag		0.1718	25Luc
Ag | AgCl, $BaCl_2$, 1.550 **m** | Ba_xHg | $BaCl_2$, 0.01 **m**, AgCl | Ag		0.1771	25Luc
Ag | AgCl, $BaCl_2$, 0.01923 **m** ‖ $BaCl_2$, 0.01 **m**, AgCl | Ag		0.0093	25Luc
Ag | AgCl, $BaCl_2$, 0.07971 **m** ‖ $BaCl_2$, 0.01 **m**, AgCl | Ag		0.0185	25Luc
Ag | AgCl, $BaCl_2$, 0.1791 **m** ‖ $BaCl_2$, 0.01 **m**, AgCl | Ag		0.0280	25Luc
Ag | AgCl, $BaCl_2$, 0.3927 **m** ‖ $BaCl_2$, 0.01 **m**, AgCl | Ag		0.0394	25Luc
Ag | AgCl, $BaCl_2$, 0.8122 **m** ‖ $BaCl_2$, 0.01 **m**, AgCl | Ag		0.0491	25Luc
Ag | AgCl, $BaCl_2$, 1.715 **m** ‖ $BaCl_2$, 0.01 **m**, AgCl | Ag		0.0600	25Luc
Ag | AgCl, $CaCl_2$, 0.01547 **m** ‖ $CaCl_2$, 0.01 **m**, AgCl | Ag		0.0060	25Luc
Ag | AgCl, $CaCl_2$, 0.02006 **m** ‖ $CaCl_2$, 0.01 **m**, AgCl | Ag		0.0097	25Luc
Ag | AgCl, $CaCl_2$, 0.04966 **m** ‖ $CaCl_2$, 0.01 **m**, AgCl | Ag		0.0220	25Luc
Ag | AgCl, $CaCl_2$, 0.08053 **m** ‖ $CaCl_2$, 0.01 **m**, AgCl | Ag		0.0284	25Luc
Ag | AgCl, $CaCl_2$, 0.1612 **m** ‖ $CaCl_2$, 0.01 **m**, AgCl | Ag		0.0368	25Luc
Ag | AgCl, $CaCl_2$, 0.3208 **m** ‖ $CaCl_2$, 0.01 **m**, AgCl | Ag		0.0460	25Luc
Ag | AgCl, $CaCl_2$, 0.6162 **m** ‖ $CaCl_2$, 0.01 **m**, AgCl | Ag		0.0549	25Luc
Ag | AgCl, $CaCl_2$, 1.013 **m** ‖ $CaCl_2$, 0.01 **m**, AgCl | Ag		0.0676	25Luc
Ag | AgCl, $CaCl_2$, 2.586 **m** ‖ $CaCl_2$, 0.01 **m**, AgCl | Ag		0.0836	25Luc
Ag | AgCl, $CaCl_2$, 0.03504 **m** | $Ca_{0.02wt\%}Hg$ | $CaCl_2$, 0.01 **m**, AgCl | Ag		0.0411	25Luc
Ag | AgCl, $CaCl_2$, 0.06294 **m** | $Ca_{0.02wt\%}Hg$ | $CaCl_2$, 0.01 **m**, AgCl | Ag		0.0606	25Luc
Ag | AgCl, $CaCl_2$, 0.1032 **m** | $Ca_{0.02wt\%}Hg$ | $CaCl_2$, 0.01 **m**, AgCl | Ag		0.0772	25Luc
Ag | AgCl, $CaCl_2$, 0.2066 **m** | $Ca_{0.02wt\%}Hg$ | $CaCl_2$, 0.01 **m**, AgCl | Ag		0.1014	25Luc
Ag | AgCl, $CaCl_2$, 0.3713 **m** | $Ca_{0.02wt\%}Hg$ | $CaCl_2$, 0.01 **m**, AgCl | Ag		0.1235	25Luc
Ag | AgCl, $CaCl_2$, 0.4655 **m** | $Ca_{0.02wt\%}Hg$ | $CaCl_2$, 0.01 **m**, AgCl | Ag		0.1335	25Luc
Ag | AgCl, $CaCl_2$, 0.6627 **m** | $Ca_{0.02wt\%}Hg$ | $CaCl_2$, 0.01 **m**, AgCl | Ag		0.1543	25Luc
Ag | AgCl, $CaCl_2$, 1.003 **m** | $Ca_{0.02wt\%}Hg$ | $CaCl_2$, 0.01 **m**, AgCl | Ag		0.1773	25Luc
Ag | AgCl, $CaCl_2$, 1.960 **m** | $Ca_{0.02wt\%}Hg$ | $CaCl_2$, 0.01 **m**, AgCl | Ag		0.2313	25Luc
Ag | AgCl, $CaCl_2$, 3.507 **m** | $Ca_{0.02wt\%}Hg$ | $CaCl_2$, 0.01 **m**, AgCl | Ag		0.3005	25Luc
Ag | AgCl, $CaCl_2$, 0.0999 **m** | Ca_xHg | $CaCl_2$, 0.00992 **m**, AgCl | Ag		0.0761	30Sca1
Ag | AgCl, $CaCl_2$, 0.0998 **m** | Ca_xHg | $CaCl_2$, 0.01051 **m**, AgCl | Ag		0.0742	30Sca1
Ag | AgCl, $CaCl_2$, 0.1004 **m** | Ca_xHg | $CaCl_2$, 0.02594 **m**, AgCl | Ag		0.0439	30Sca1
Ag | AgCl, $CaCl_2$, 0.0998 **m** | Ca_xHg | $CaCl_2$, 0.02758 **m**, AgCl | Ag		0.0412	30Sca1
Ag | AgCl, $CaCl_2$, 0.0999 **m** | Ca_xHg | $CaCl_2$, 0.04113 **m**, AgCl | Ag		0.0289	30Sca1
Ag | AgCl, $CaCl_2$, 0.0999 **m** | Ca_xHg | $CaCl_2$, 0.1263 **m**, AgCl | Ag		–0.0078	30Sca1

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Ag | AgCl, $CaCl_2$, 0.1003 **m** | Ca_xHg | $CaCl_2$, 0.4420 **m**, AgCl | Ag		–0.0558	30Sca1
Ag | AgCl, $CaCl_2$, 0.0997 **m** | Ca_xHg | $CaCl_2$, 0.6589 **m**, AgCl | Ag		–0.0768	30Sca1
Ag | AgCl, $CaCl_2$, 0.1006 **m** | Ca_xHg | $CaCl_2$, 0.9301 **m**, AgCl | Ag		–0.0962	30Sca1
Ag | AgCl, $CdCl_2$, 0.1 **m** | $CdCl_2$, 0.0101 **m**, AgCl | Ag		0.0265	29Luc2
Ag | AgCl, $CdCl_2$, 0.1 **m** | $CdCl_2$, 0.02015 **m**, AgCl | Ag		0.0177	29Luc2
Ag | AgCl, $CdCl_2$, 0.1 **m** | $CdCl_2$, 0.03030 **m**, AgCl | Ag		0.0129	29Luc2
Ag | AgCl, $CdCl_2$, 0.1 **m** | $CdCl_2$, 0.05056 **m**, AgCl | Ag		0.00729	29Luc2
Ag | AgCl, $CdCl_2$, 0.1 **m** | $CdCl_2$, 0.08145 **m**, AgCl | Ag		0.00214	29Luc2
Ag | AgCl, $CdCl_2$, 0.1 **m** | $CdCl_2$, 0.2014 **m**, AgCl | Ag		–0.00697	29Luc2
Ag | AgCl, $CdCl_2$, 0.1 **m** | $CdCl_2$, 0.5101 **m**, AgCl | Ag		–0.0152	29Luc2
Ag | AgCl, $CdCl_2$, 0.1 **m** | $CdCl_2$, 0.9999 **m**, AgCl | Ag		–0.0203	29Luc2
Ag | AgCl, $CdCl_2$, 0.1 **m** | $CdCl_2$, 4.227 **m**, AgCl | Ag		–0.0274	29Luc2
Ag | AgCl, $CdCl_2$, 0.1 **m** | $CdCl_2$, 6.146 **m**, AgCl | Ag		–0.0265	29Luc2
Ag | AgCl, CsCl, 0.05 **m** | Cs_xHg | CsCl, 0.1 **m**, AgCl | Ag		–0.03201	30Har1
Ag | AgCl, CsCl, 0.1 **m** | Cs_xHg | CsCl, 0.1 **m**, AgCl | Ag		0.0	30Har1
Ag | AgCl, CsCl, 0.2 **m** | Cs_xHg | CsCl, 0.1 **m**, AgCl | Ag		0.0311	30Har1
Ag | AgCl, CsCl, 0.5 **m** | Cs_xHg | CsCl, 0.1 **m**, AgCl | Ag		0.0711	30Har1
Ag | AgCl, CsCl, 1.0 **m** | Cs_xHg | CsCl, 0.1 **m**, AgCl | Ag		0.1013	30Har1
Ag | AgCl, CsCl, 2.0 **m** | Cs_xHg | CsCl, 0.1 **m**, AgCl | Ag		0.1323	30Har1
Ag | AgCl, CsCl, 3.0 **m** | Cs_xHg | CsCl, 0.1 **m**, AgCl | Ag		0.1502	30Har1
Ag | AgCl | NaCl, 0.01<**m**<0.1 ‖ KCl,sat., NaN_3, 0.01<**m**<0.1 | AgN_3 | Ag	278 K	0.0620	38Tay
	288 K	0.0657	38Tay
	298 K	0.0695	38Tay
	308 K	0.0733	38Tay
	318 K	0.0772	38Tay
Ag | AgCl, NaCl, 0.1 **m** | Na_xHg | NaCl, 0.1 **m**, AgCl | Ag		0.00	29Har2
Ag | AgCl, NaCl, 0.1 **m** | Na_xHg | NaCl, 0.2 **m**, AgCl | Ag		0.0324	29Har2
Ag | AgCl, NaCl, 0.1 **m** | Na_xHg | NaCl, 0.5 **m**, AgCl | Ag		0.0756	29Har2
Ag | AgCl, NaCl, 0.1 **m** | Na_xHg | NaCl, 1.0 **m**, AgCl | Ag		0.1095	29Har2
Ag | AgCl, NaCl, 0.1 **m** | Na_xHg | NaCl, 1.5 **m**, AgCl | Ag		0.1303	29Har2
Ag | AgCl, NaCl, 0.1 **m** | Na_xHg | NaCl, 2.0 **m**, AgCl | Ag		0.1462	29Har2
Ag | AgCl, NaCl, 0.1 **m** | Na_xHg | NaCl, 2.5 **m**, AgCl | Ag		0.1592	29Har2
Ag | AgCl, NaCl, 0.1 **m** | Na_xHg | NaCl, 3.0 **m**, AgCl | Ag		0.1703	29Har2
Ag | AgCl, NaCl, 0.1 **m** | Na_xHg | NaCl, 3.5 **m**, AgCl | Ag		0.1818	29Har2
Ag | AgCl, NaCl, 0.1 **m** | Na_xHg | NaCl, 4.0 **m**, AgCl | Ag		0.1907	29Har2
Ag | AgCl, NaCl, 0.0953 **m** | Na_zHg | NaOH, 0.9174 **m**, NaCl, 0.0826 **m**, AgCl | Ag		0.0469	37Har3
Ag | AgCl, NaCl, 0.0953 **m** | Na_zHg | NaOH, 0.5702 **m**, NaCl, 0.4298 **m**, AgCl | Ag		0.0898	37Har3
Ag | AgCl, NaCl, 0.0953 **m** | Na_zHg | NaOH, 0.1854 **m**, NaCl, 0.8146 **m**, AgCl | Ag		0.1064	37Har3
Ag | AgCl, NaCl, 0.0953 **m** | Na_zHg | NaOH, 0.1032 **m**, NaCl, 0.8968 **m**, AgCl | Ag		0.1094	37Har3
Ag | AgCl, NaCl, 0.0953 **m** | Na_zHg | NaOH, 0.00 **m**, NaCl, 1.0 **m**, AgCl | Ag		0.1120	37Har3
Ag | AgCl, NaCl, 0.0953 **m** | Na_zHg | NaOH, 0.2218 **m**, NaCl, 0.2782 **m**, AgCl | Ag		0.0628	37Har3

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Ag | AgCl, NaCl, 0.0953 **m** | Na_zHg | NaOH, 0.0885 **m**, NaCl, 0.4115 **m**, AgCl | Ag		0.0733	37Har3
Ag | AgCl, NaCl, 0.0953 **m** | Na_zHg | NaOH, 0.00 **m**, NaCl, 0.5 **m**, AgCl | Ag		0.0744	37Har3
Ag | AgI, KI, 0.005 **m** ‖ KCl, 0.005 **m** | Ag		0.3724	37Can1
Ag | AgI, KI, 0.05 **m** ‖ KCl, 0.05 **m** | Ag		0.3728	37Can1
Ag | AgI, KI, 0.50 **m** ‖ KCl, 0.50 **m** | Ag		0.3724	37Can1
Ag | AgI, KI, 1.0 **m** ‖ KCl, 1.0 **m** | Ag		0.3723	37Can1
Ag | $KAg(CN)_2$, 0.0005 **M**, KCN, 0.02 **M** ‖ KCN, 0.01 **M**, $KAg(CN)_2$, 0.0005 **M** | Ag		0.0344	04Bod
Ag | $KAg(CN)_2$, 0.005 **M**, KCN, 0.2 **M** ‖ KCN, 0.1 **M**, $KAg(CN)_2$, 0.0005 **M** | Ag		0.0431	04Bod
Ag | $KAg(CN)_2$, 0.005 **M**, KCN, 0.1 **M** ‖ KCN, 0.1 **M**, $KAg(CN)_2$, 0.00125 **M** | Ag		–0.0351	04Bod
Ag | $KAg(CN)_2$, 0.05 **M**, KCN, 1.0 **M** ‖ KCN, 1.0 **M**, $KAg(CN)_2$, 0.0125 **M** | Ag		–0.0375	04Bod
Ag | Ag_2S | HCl, 0.5 **M**, H_2S,sat., 1 atm ‖ HCl, 0.5 **M** | H_2, Pt	278 K	0.03805	20Noy
	283 K	0.03767	20Noy
	298 K	0.03658	20Noy
	308 K	0.0357	20Noy
Ag | $KAg(CN)_2$, 0.05 **M**, KCN, 1.0 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.894	04Bod
Ag | $KAg(CN)_2$, 3.44 m**M**, KCN, 5.51 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		1.1542	32Bri
Ag | $KAg(CN)_2$, 4.86 m**M**, KCN, 3.89 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		1.1302	32Bri
Ag | $KAg(CN)_2$, 9.78 m**M**, KCN, 14.1 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg		1.0634	32Bri
Ag | AgBr, HBr, 0.1 **M** ‖ HBr, 0.1 **M** ‖ HBr, 0.1 **M** | Br_2, $3.82\cdot10^{-5}$ **M**, Pt		0.860	99Lut
Ag | AgBr, HBr, 0.1 **M** ‖ HBr, 0.1 **M** ‖ HBr, 0.1 **M** | Br_2, $2.28\cdot10^{-4}$ **M**, Pt		0.881	99Lut
Ag | AgBr, HBr, 0.1 **M** ‖ HBr, 0.1 **M** ‖ HBr, 0.1 **M** | Br_2, 0.00186 **M**, Pt		0.909	99Lut
Ag | AgBr, HBr, 0.1 **M** ‖ HBr, 0.1 **M** ‖ HBr, 0.1 **M** | Br_2, 0.00645 **M**, Pt		0.925	99Lut
Ag | AgBr, HBr, 0.1 **M** ‖ HBr, 0.1 **M** ‖ HBr, 0.1 **M** | Br_2, 0.0192 **M**, Pt		0.941	99Lut
Ag | AgBr, HBr, 0.1 **M** ‖ HBr, 0.1 **M** ‖ HBr, 0.1 **M** | Br_2, 0.106 **M**, Pt		0.980	99Lut
Ag | AgBr, HBr, 0.1 **M** ‖ HBr, 0.1 **M** ‖ HBr, 0.1 **M** | Br_2, 0.214 **M**, Pt		1.007	99Lut
Ag | AgBr, KBr, 0.009995 **M** ‖ KBr, 0.009995 **M** ‖ Br_2, a=$4.815\cdot10^{-3}$, PtIr		0.9468	34Jon
Ag | AgBr, KBr, 0.019999 **M** ‖ KBr, 0.009995 **M** ‖ Br_2, a=$4.819\cdot10^{-3}$, PtIr		0.9471	34Jon
Ag | AgBr, KBr, 0.050480 **M** ‖ KBr, 0.009995 **M** ‖ Br_2, a=$6.159\cdot10^{-3}$, PtIr		0.9471	34Jon
Ag | AgBr, KBr, 0.10095 **M** ‖ KBr, 0.009995 **M** ‖ Br_2, a=$4.927\cdot10^{-3}$, PtIr		0.9471	34Jon
Ag | AgBr, KBr, 0.20193 **M** ‖ KBr, 0.009995 **M** ‖ Br_2, a=$11.823\cdot10^{-3}$, PtIr		0.9471	34Jon
Ag | AgBr, KBr, 0.5003 **M** ‖ KBr, 0.009995 **M** ‖ Br_2, a=$10.355\cdot10^{-3}$, PtIr		0.9471	34Jon
Ag | AgBr, KBr, 1.007 **M** ‖ KBr, 0.009995 **M** ‖ Br_2, a=$1.801\cdot10^{-3}$, PtIr		0.9471	34Jon
Ag | AgBr, KBr, 0.9996 **M** ‖ KBr, 0.009995 **M** ‖ Br_2, a=$7.533\cdot10^{-3}$, PtIr		0.9471	34Jon
Ag | AgCl | HCl, 0.01 **m**, quinhydrone | Pt	0.18 °C	0.2599	33Har2
	283 K	0.2498	33Har2
	288 K	0.2449	33Har2
	293 K	0.2401	33Har2
	298 K	0.2355	33Har2
	303K	0.2310	33Har2
	308 K	0.2267	33Har2

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
	313 K	0.2225	33Har2
Ag | AgCl | HCl, 0.1 **M** | Hg_2Cl_2 | Hg	293 K	0.041	55Lie
	328 K	0.055	55Lie
	333 K	0.058	55Lie
	373 K	0.058	55Lie
	397 K	0.078	55Lie
	413 K	0.103	55Lie
	448 K	0.138	55Lie
	468 K	0.098	55Lie
	538 K	0.264	55Lie
Ag | AgCl | HCl, 0.5 **M** | Hg_2Cl_2 | Hg	293 K	0.041	55Lie
Ag | AgCl | HCl, 1 **M** | Hg_2Cl_2 | Hg	293 K	0.041	55Lie
	344 K	0.058	55Lie
	383 K	0.065	55Lie
	433 K	0.088	55Lie
	493 K	0.110	55Lie
	533 K	0.130	55Lie
Ag | AgCl, $MgCl_2$, 0.0227 **m** ‖ KCl,sat. ‖ KCl, 0.1 **m**, Hg_2Cl_2 | Hg		0.021	28Jon
Ag | AgCl, $MgCl_2$, 0.0399 **m** ‖ KCl,sat. ‖ KCl, 0.1 **m**, Hg_2Cl_2 | Hg		0.0335	28Jon
Ag | AgCl, $MgCl_2$, 0.0875 **m** ‖ KCl,sat. ‖ KCl, 0.1 **m**, Hg_2Cl_2 | Hg		0.0500	28Jon
Ag | AgCl, $MgCl_2$, 0.202 **m** ‖ KCl,sat. ‖ KCl, 0.1 **m**, Hg_2Cl_2 | Hg		0.0675	28Jon
Ag | AgCl, $MgCl_2$, 0.386 **m** ‖ KCl,sat. ‖ KCl, 0.1 **m**, Hg_2Cl_2 | Hg		0.0805	28Jon
Ag | AgCl, $MgCl_2$, 0.968 **m** ‖ KCl,sat. ‖ KCl, 0.1 **m**, Hg_2Cl_2 | Hg		0.1025	28Jon
Ag | AgCl, $MgCl_2$, 2.175 **m** ‖ KCl,sat. ‖ KCl, 0.1 **m**, Hg_2Cl_2 | Hg		0.118	28Jon
Ag | AgCl, $MgCl_2$, 4.217 **m** ‖ KCl,sat. ‖ KCl, 0.1 **m**, Hg_2Cl_2 | Hg		0.147	28Jon
Ag | AgI, ZnI_2, 0.01075 **m** ‖ ZnI_2, 0.01075 **m**, Hg_2I_2 | Hg	15 °C	0.1076	39Vos
	20 °C	0.1091	39Vos
	25 °C	0.1106	39Vos
	30 °C	0.1122	39Vos
	35 °C	0.1139	39Vos
Ag | AgI, ZnI_2, 0.01474 **m** ‖ ZnI_2, 0.01075 **m**, Hg_2I_2 | Hg	15 °C	0.1081	39Vos
	20 °C	0.1096	39Vos
	25 °C	0.1112	39Vos
	30 °C	0.1127	39Vos
	35 °C	0.1144	39Vos
Ag | AgI, ZnI_2, 0.01801 **m** ‖ ZnI_2, 0.01075 **m**, Hg_2I_2 | Hg	15 °C	0.1086	39Vos
	20 °C	0.1102	39Vos
	25 °C	0.1117	39Vos
	30 °C	0.1134	39Vos
	35 °C	0.1150	39Vos
Ag | AgI, ZnI_2, 0.03373 **m** ‖ ZnI_2, 0.01075 **m**, Hg_2I_2 | Hg	25 °C	0.1125	39Vos
	30 °C	0.1141	39Vos
Ag | AgI, ZnI_2, 0.06068 **m** ‖ ZnI_2, 0.01075 **m**, Hg_2I_2 | Hg	15 °C	0.1117	39Vos
	20 °C	0.1132	39Vos
	25 °C	0.1147	39Vos
	30 °C	0.1163	39Vos
	35 °C	0.1179	39Vos
Ag | Ag_2O | KOH | HgO | Hg	298 K	0.2440	57Ham

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Ag | AgCl, $PbCl_2$, 0.00262 **m**, D-mannitol, 0.10 **m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.5659	38Cro
Ag | AgCl, $PbCl_2$, 0.00528 **m**, D-mannitol, 0.10 **m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.5419	38Cro
Ag | AgCl, $PbCl_2$, 0.01056 **m**, D-mannitol, 0.10 **m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.5197	38Cro
Ag | AgCl, $PbCl_2$, 0.01580 **m**, D-mannitol, 0.10 **m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.5096	38Cro
Ag | AgCl, $PbCl_2$, 0.02109 **m**, D-mannitol, 0.10 **m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.5013	38Cro
Ag | AgCl, $PbCl_2$, 0.01055 **m**, D-mannitol, 0.0 **m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.5200	38Cro
Ag | AgCl, $PbCl_2$, 0.01055 **m**, D-mannitol, 0.1012 **m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.5205	38Cro
Ag | AgCl, $PbCl_2$, 0.01055 **m**, D-mannitol, 0.2897 **m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.5217	38Cro
Ag | AgCl, $PbCl_2$, 0.01055 **m**, D-mannitol, 0.5609 **m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.5254	38Cro
Ag | AgCl, $PbCl_2$, 0.01055 **m**, D-mannitol, 0.7313 **m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.5270	38Cro
Ag | AgCl, $PbCl_2$, 0.023455 **m**, D-mannitol, 0.0 **m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.4972	38Cro
Ag | AgCl, $PbCl_2$, 0.023455 **m**, D-mannitol, 0.0990 **m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.4983	38Cro
Ag | AgCl, $PbCl_2$, 0.023455 **m**, D-mannitol, 0.2899 **m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.5000	38Cro
Ag | AgCl, $PbCl_2$, 0.023455 **m**, D-mannitol, 0.5604 **m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.50224	38Cro
Ag | AgCl, $PbCl_2$, 0.023455 **m**, D-mannitol, 0.7304**m**, $Ba(NO_3)_2$, 0.0 **m** | $Pb_{sat.}Hg$		0.5047	38Cro
Ag | AgCl, $PbCl_2$, 0.01090 **m**, D-mannitol, 0.20698 **m**, $Ba(NO_3)_2$, 0.02339 **m** | $Pb_{sat.}Hg$		0.5269	38Cro
Ag | AgCl, $PbCl_2$, 0.01083 **m**, D-mannitol, 0.15708 **m**, $Ba(NO_3)_2$, 0.02324 **m** | $Pb_{sat.}Hg$		0.5266	38Cro
Ag | AgCl, $PbCl_2$, 0.01076**m**, D-mannitol, 0.1002 **m**, $Ba(NO_3)_2$, 0.02339 **m** | $Pb_{sat.}Hg$		0.5261	38Cro
Ag | AgCl, $PbCl_2$, 0.01069 **m**, D-mannitol, 0.05009 **m**, $Ba(NO_3)_2$, 0.02294 **m** | $Pb_{sat.}Hg$		0.5258	38Cro
Ag | AgCl, $PbCl_2$, 0.01063 **m**, D-mannitol, 0.00 **m**, $Ba(NO_3)_2$, 0.02281 **m** | $Pb_{sat.}Hg$		0.5256	38Cro
Sodium (Na)			
$Na_{0.2065wt\%}Hg$ | NaOH, 0.2 **m** ‖ NaCl, 0.2 **m** ‖ KCl, 0.2 **m** ‖ KCl, 1.0 **m**, Hg_2Cl_2 | Hg		2.1749	10Lew
$Na_{0.551wt\%}Hg$ | NaOH | $Na_{0.1815wt\%}Hg$		0.0565	22Ric
$Na_{0.1978wt\%}Hg$ | NaOH | $Na_{0.0396wt\%}Hg$		0.0531	22Ric
$Na_{0.0396wt\%}Hg$ | NaOH | $Na_{0.0197wt\%}Hg$		0.0194	22Ric
$Na_{0.206wt\%}Hg$ | NaCl, 0.001 **m** ‖ KCl, 1.0 **m**, Hg_2Cl_2 | Hg		2.2967	30Bir
$Na_{0.206wt\%}Hg$ | NaCl, 0.01 **m** ‖ KCl, 1.0 **m**, Hg_2Cl_2 | Hg		2.2665	30Bir
$Na_{0.206wt\%}Hg$ | NaCl, 0.02 **m** ‖ KCl, 1.0 **m**, Hg_2Cl_2 | Hg		2.2556	30Bir
$Na_{0.206wt\%}Hg$ | NaCl, 0.05 **m** ‖ KCl, 1.0 **m**, Hg_2Cl_2 | Hg		2.2296	30Bir
$Na_{0.206wt\%}Hg$ | NaCl, 0.1 **m** ‖ KCl, 1.0 **m**, Hg_2Cl_2 | Hg		2.2147	30Bir
$Na_{0.206wt\%}Hg$ | NaCl, 0.2 **m** ‖ KCl, 1.0 **m**, Hg_2Cl_2 | Hg		2.1956	30Bir
$Na_{0.206wt\%}Hg$ | NaCl, 0.5 **m** ‖ KCl, 1.0 **m**, Hg_2Cl_2 | Hg		2.1750	30Bir
$Na_{0.206wt\%}Hg$ | NaCl, 1.0 **m** ‖ KCl, 1.0 **m**, Hg_2Cl_2 | Hg		2.1531	30Bir
Sulfur (S)			
Pt | S,sat., Na_2S, 0.00781 **M** ‖ KCl, 1.0 M, Hg_2Cl_2 | Hg		0.5523	02Küs
Pt | S,sat., Na_2S, 0.125 **M** ‖ KCl, 1.0 M, Hg_2Cl_2 | Hg		0.5836	02Küs
Pt | S,sat., Na_2S, 2.0 **M** ‖ KCl, 1.0 M, Hg_2Cl_2 | Hg		0.6211	02Küs

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Thallium (Tl)			
Tl | TlCl,sol., TlCl.sat., Hg_2Cl_2 | Hg	20 °C	0.8230	34Öho
(U_T [V]=0.8230–0.32·10^{-3}(T–20 °C))			
Tl | TlCl, 0.002 **M** ‖ NH_4NO_3,sat. ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	17 °C	0.813	09Bri
Tl | TlCl, 0.005 **M** ‖ NH_4NO_3,sat. ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	17 °C	0.788	09Bri
Tl | TlCl, 0.01 **M** ‖ NH_4NO_3,sat. ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	17 °C	0.773	09Bri
Tl | $TlNO_3$, 0.002 **M** ‖ NH_4NO_3,sat. ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	17 °C	0.813	09Bri
Tl | $TlNO_3$, 0.01 **M** ‖ NH_4NO_3,sat. ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	17 °C	0.775	09Bri
Tl | $TlNO_3$, 0.1 **M** ‖ NH_4NO_3,sat. ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	17 °C	0.724	09Bri
Tl | Tl_2SO_4, 0.005 **M**, Hg_2SO_4 | Hg	14.9 °C	1.1198	34Öho
	25 °C	1.1310	34Öho
Tl | Tl_2SO_4, 0.0125 **M**, Hg_2SO_4 | Hg	14.9 °C	1.0967	34Öho
	25 °C	1.1065	34Öho
Tl | Tl_2SO_4, 0.025 **M**, Hg_2SO_4 | Hg	14.9 °C	1.0768	34Öho
	25 °C	1.0885	34Öho
Tl | Tl_2SO_4, 0.05 **M**, Hg_2SO_4 | Hg	14.9 °C	1.0609	34Öho
	25 °C	1.0690	34Öho
$Tl_{sat.}$Hg | TlBr,sol., TlBr,sat., Hg_2Br_2 | Hg	20 °C	0.7920	34Öho
(U_T [V]=0.7920–0.00023(T–20 °C))			
$Tl_{sat.}$Hg | TlBr,sol., KBr, 0.1 **M** ‖ KBr, 0.1 **M**, Hg_2Br_2 | Hg		0.7945	33Ish5
$Tl_{sat.}$Hg | TlI,sol., TlI,sat., Hg_2I_2 | Hg	20 °C	0.7572	34Öho
(U_T [V]=0.7572–0.00023(T–20 °C))			
$Tl_{5wt\%}$Hg | TlCl, 0.0005 **m**, AgCl | Ag		0.8307	34Cow
$Tl_{5wt\%}$Hg | TlCl, 0.001 **m**, AgCl | Ag		0.7956	34Cow
$Tl_{5wt\%}$Hg | TlCl, 0.002 **m**, AgCl | Ag		0.7609	34Cow
$Tl_{5wt\%}$Hg | TlCl, 0.004 **m**, AgCl | Ag		0.7266	34Cow
$Tl_{5wt\%}$Hg | TlCl, 0.006 **m**, AgCl | Ag		0.7068	34Cow
$Tl_{5wt\%}$Hg | TlCl, 0.008 **m**, AgCl | Ag		0.6929	34Cow
$Tl_{5wt\%}$Hg | TlCl, 0.01 **m**, AgCl | Ag		0.6821	34Cow
$Tl_{sat.}$Hg | $TlNO_3$, 0.433 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.7105	05Abe
$Tl_{sat.}$Hg | $TlNO_3$, 0.0183 **M** ‖ $TlNO_3$, 0.0107 **M** | Tl_xHg	17 °C	0.0099	21Dru
$Tl_{sat.}$Hg | $TlNO_3$, 0.0568 **M** ‖ $TlNO_3$, 0.0107 **M** | Tl_xHg	17 °C	0.0325	21Dru
$Tl_{sat.}$Hg | $TlNO_3$, 0.1530 **M** ‖ $TlNO_3$, 0.0107 **M** | Tl_xHg	17 °C	0.0536	21Dru
$Tl_{sat.}$Hg | Tl_2SO_4, 0.1101 **M** ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.7330	05Abe
$Tl_{5wt\%}$Hg | TlCl, AgCl | Tl		0.1196	34Cow
Tl | Tl_2SO_4 | $Tl_{49.48wt\%}$Hg		0.0025	19Ric
$Tl_{49.42wt\%}$Hg | Tl_2SO_4 | $Tl_{34.03wt\%}$Hg		0.00852	19Ric
$Tl_{34.03wt\%}$Hg | Tl_2SO_4 | $Tl_{27.36wt\%}$Hg		0.00868	19Ric
$Tl_{27.36wt\%}$Hg | Tl_2SO_4 | $Tl_{20.97wt\%}$Hg		0.01157	19Ric
$Tl_{20.78wt\%.}$Hg | Tl_2SO_4 | $Tl_{10.02wt\%}$Hg		0.03377	19Ric
$Tl_{10.02wt\%}$Hg | Tl_2SO_4 | $Tl_{4.930wt\%}$Hg		0.002948	19Ric
$Tl_{4.935wt\%}$Hg | Tl_2SO_4 | $Tl_{3.788wt\%}$Hg		0.00976	19Ric
$Tl_{3.788wt\%}$Hg | Tl_2SO_4 | $Tl_{1.704wt\%}$Hg		0.0264	19Ric
$Tl_{1.704wt\%}$Hg | Tl_2SO_4 | $Tl_{0.3315wt\%}$Hg		0.04555	19Ric
Tin (Sn)			
Sn | $SnCl_2$, 0.01 **M**, HCl, 0.01 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg	24.5 °C	0.4645	28Pry
Sn | $SnCl_2$, 0.01 **M**, HCl, 0.1 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg	24.5 °C	0.4770	28Pry
Sn | $SnCl_2$, 0.01 **M**, HCl, 1.0 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg	24.5 °C	0.5194	28Pry

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Sn | $SnCl_2$, 0.1 **M**, HCl, 1.0 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg	24.5 °C	0.4892	28Pry
Sn | $SnCl_2$, 0.092 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	18.5 °C	0.480	11För
Sn | $SnCl_2$, 0.897 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	18.5 °C	0.471	11För
Sn | $SnCl_4$, 0.01 **M**, HCl, 0.0109 **M**, KCl, 0.0 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5000	34Pry
Sn | $SnCl_4$, 0.01 **M**, HCl, 0.0109 **M**, KCl, 0.25 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5021	34Pry
Sn | $SnCl_4$, 0.01 **M**, HCl, 0.0109 **M**, KCl, 0.5 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5300	34Pry
Sn | $SnCl_4$, 0.01 **M**, HCl, 0.0109 **M**, KCl, 1.149 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5518	34Pry
Sn | $SnCl_4$, 0.01 **M**, HCl, 0.0109 **M**, KCl, 1.445 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5562	34Pry
Sn | $SnCl_4$, 0.01 **M**, HCl, 0.02 **M**, KCl, 0.0 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5091	34Pry
Sn | $SnCl_4$, 0.01 **M**, HCl, 0.12 **M**, KCl, 0.12 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5231	34Pry
Sn | $SnCl_4$, 0.01 **M**, HCl, 0.43 **M**, KCl, 0.43**M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5441	34Pry
Sn | $SnCl_4$, 0.01 **M**, HCl, 1.03 **M**, KCl, 1.03 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5671	34Pry
Sn | $SnCl_4$, 0.01 **M**, HCl, 0.5 **M**, KCl, 0.0 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5331	34Pry
Sn | $SnCl_4$, 0.01 **M**, HCl, 0.5 **M**, KCl, 1.11 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5599	34Pry
Sn | $SnCl_4$, 0.01 **M**, HCl, 0.5 **M**, KCl, 2.46 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5831	34Pry
Sn | $SnCl_4$, 0.01 **M**, HCl, 1.48 **M**, KCl, 1.48 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5837	34Pry
Sn | $Sn(NO_3)_2$, 0.05 **M** ‖ KNO_3, 0.5 **M** ‖ KCl, 1.0 M, Hg_2Cl_2 | Hg		0.440	03Sac
Sn | $Sn(ClO_4)_4$, 0.02 **M**, $HClO_4$, 0.5 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5026	34Pry
Sn | $Sn(ClO_4)_4$, 0.05 **M**, $HClO_4$, 0.5 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5011	34Pry
Sn | $Sn(ClO_4)_4$, 0.1 **M**, $HClO_4$, 0.876 **M** ‖ KCl, 3.5 **M**, Hg_2Cl_2 | Hg		0.5007	34Pry
Sn | $SnSO_4$, 0.010 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	18 °C	0.494	11För
Sn | $SnSO_4$, 0.492 **M** ‖ KCl,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	18 °C	0.473	11För
Sn | $SnSO_4$, 0.096 **M**, H_2SO_4, 0.5 **M** ‖ NH_4NO_3,sat. ‖ H_2SO_4, 0.5 **M**, Hg_2SO_4 | Hg	18 °C	0.887	11För
Sn | $SnSO_4$, 0.0105 **M**, H_2SO_4, 0.5 **M** ‖ NH_4NO_3,sat. ‖ H_2SO_4, 0.5 **M**, Hg_2SO_4 | Hg	18 °C	0.913	11För
Zinc (Zn)			
Zn | ZnO, $Ba(OH)_2$, 0.1<**m**<0.2 | H_2, Pt	44.9 °C	0.4198	26Mai
Zn | $K_2(CN)_4$, 9.09 m**M**, KCN, 3.74 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	12.5 °C	1.344	32Bri
Zn | $K_2(CN)_4$, 9.09 m**M**, KCN, 17.11 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	12.5 °C	1.429	32Bri
Zn | $K_2(CN)_4$, 7.41 m**M**, KCN, 25.08 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	12.5 °C	1.410	32Bri
Zn | $K_2(CN)_4$, 4.76 m**M**, KCN, 27.86 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	12.5 °C	1.434	32Bri
Zn | $K_2(CN)_4$, 2.91 m**M**, KCN, 36.17 m**M** ‖ KNO_3,sat. ‖ KCl, 1.0 **M**, Hg_2Cl_2 | Hg	12.5 °C	1.437	32Bri
Zn | $ZnCl_2$, 0.005 **m**, Hg_2Cl_2 | Hg		1.2244	31Get1
Zn | $ZnCl_2$, 0.01 **m**, Hg_2Cl_2 | Hg		1.2035	31Get1
Zn | $ZnCl_2$, 0.0199 **m**, Hg_2Cl_2 | Hg		1.1800	31Get1
Zn | $ZnCl_2$, 0.05003 **m**, Hg_2Cl_2 | Hg		1.1474	31Get1
Zn | $ZnCl_2$, 0.10019 **m**, Hg_2Cl_2 | Hg		1.1308	31Get1
Zn | $ZnCl_2$, 0.25148 **m**, Hg_2Cl_2 | Hg		1.1009	31Get1
Zn | $ZnCl_2$, 0.50786 **m**, Hg_2Cl_2 | Hg		1.0782	31Get1
Zn | $ZnCl_2$, 1.0375 **m**, Hg_2Cl_2 | Hg		1.0568	31Get1
Zn | $Zn(OH)_2$,sol., NaOH, 1.0 **M** ‖ NaOH, 1.0 **M**, HgO | Hg		1.344	27Die
Zn | $ZnSO_4$, 0.005 **M** ‖ KCl,sat. ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	22 °C	1.163	21Moo

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
Zn | $ZnSO_4$, 0.05 **M** ‖ KCl,sat. ‖ KCl, 0.1 **M**, Hg_2Cl_2 | Hg	22 °C	1.138	21Moo
$Zn_{10wt\%}Hg$ | $Zn(OH)_2$, 0.2 **M**, KOH, 4.2 **M** ‖ KOH, 4.2 **M** | H_2, Pt		0.424	07Fau
$Zn_{10wt\%}Hg$ | $ZnBr_2$,sat., Hg_2Br_2 | Hg	20 °C	0.6896	32Ish1
	25 °C	0.6874	32Ish1
	30 °C	0.6850	32Ish1
	35 °C	0.6826	32Ish1
	37.5 °C	0.6830	32Ish1
	40 °C	0.6833	32Ish1
	45 °C	0.6840	32Ish1
$Zn_{10wt\%}Hg$ | $ZnCl_2$,sat., Hg_2Cl_2 | Hg	25 °C	0.8231	33Ish1
	30 °C	0.8238	33Ish1
	35 °C	0.8244	33Ish1
	40 °C	0.8251	33Ish1
	42.5 °C	0.8255	33Ish1
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.00963 **m**, AgCl | Ag	15 °C	1.1568	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.0304 **m**, AgCl | Ag	15 °C	1.1180	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.1011 **m**, AgCl | Ag	15 °C	1.0799	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.1444 **m**, AgCl | Ag	15 °C	1.0678	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.3590 **m**, AgCl | Ag	15 °C	1.0387	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.5285 **m**, AgCl | Ag	15 °C	1.0262	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 1.202 **m**, AgCl | Ag	15 °C	1.0007	32Ham
$Zn_{sat.}Hg$ | $ZnCl_2$, 0.002941 **m**, AgCl | Ag		1.1983	30Sca2
$Zn_{sat.}Hg$ | $ZnCl_2$, 0.00781 **m**, AgCl | Ag		1.1650	30Sca2
$Zn_{sat.}Hg$ | $ZnCl_2$, 0.01236 **m**, AgCl | Ag		1.1495	30Sca2
$Zn_{sat.}Hg$ | $ZnCl_2$, 0.02144 **m**, AgCl | Ag		1.1310	30Sca2
$Zn_{sat.}Hg$ | $ZnCl_2$, 0.04242 **m**, AgCl | Ag		1.1090	30Sca2
$Zn_{sat.}Hg$ | $ZnCl_2$, 0.9048 **m**, AgCl | Ag		1.0844	30Sca2
$Zn_{sat.}Hg$ | $ZnCl_2$, 0.2111 **m**, AgCl | Ag		1.0566	30Sca2
$Zn_{sat.}Hg$ | $ZnCl_2$, 0.4499 **m**, AgCl | Ag		1.0328	30Sca2
$Zn_{sat.}Hg$ | $ZnCl_2$, 0.6404 **m**, AgCl | Ag		1.0221	30Sca2
$Zn_{sat.}Hg$ | $ZnCl_2$, 1.480 **m**, AgCl | Ag		0.9978	30Sca2
Zn_zHg | $ZnCl_2$, 0.000625 **m** ‖ $ZnCl_2$, 0.00125 **m** | Zn_zHg	30 °C	0.000638	32Mas2
Zn_zHg | $ZnCl_2$, 0.00125 **m** ‖ $ZnCl_2$, 0.0025 **m** | Zn_zHg	30 °C	0.000655	32Mas2
Zn_zHg | $ZnCl_2$, 0.0025 **m** ‖ $ZnCl_2$, 0.005 **m** | Zn_zHg	30 °C	0.000615	32Mas2
Zn_zHg | $ZnCl_2$, 0.005 **m** ‖ $ZnCl_2$, 0.01 **m** | Zn_zHg	30 °C	0.00060	32Mas2
ZnHg | $Zn(ClO_4)_2$, 0.1 **m** ‖ $Zn(ClO_4)_2$, 0.1 **m** | ZnHg		0.0000	46Sto1
ZnHg | $Zn(ClO_4)_2$, 0.1 **m** ‖ $Zn(ClO_4)_2$, 0.5 **m** | ZnHg		0.0400	46Sto1
ZnHg | $Zn(ClO_4)_2$, 0.1 **m** ‖ $Zn(ClO_4)_2$, 1.0 **m** | ZnHg		0.0673	46Sto1
ZnHg | $Zn(ClO_4)_2$, 0.1 **m** ‖ $Zn(ClO_4)_2$, 1.5 **m** | ZnHg		0.0909	46Sto1
ZnHg | $Zn(ClO_4)_2$, 0.1 **m** ‖ $Zn(ClO_4)_2$, 2.0 **m** | ZnHg		0.1139	46Sto1
ZnHg | $Zn(ClO_4)_2$, 0.1 **m** ‖ $Zn(ClO_4)_2$, 2.5 **m** | ZnHg		0.1372	46Sto1
ZnHg | $Zn(ClO_4)_2$, 0.1 **m** ‖ $Zn(ClO_4)_2$, 3.0 **m** | ZnHg		0.1604	46Sto1
ZnHg | $Zn(ClO_4)_2$, 0.1 **m** ‖ $Zn(ClO_4)_2$, 3.5 **m** | ZnHg		0.1836	46Sto1
ZnHg | $Zn(ClO_4)_2$, 0.1 **m** ‖ $Zn(ClO_4)_2$, 4.0 **m** | ZnHg		0.2068	46Sto1
$Zn_{sat.}Hg$ | ZnI_2, 0.3 **m** ‖ ZnI_2, 0.05 **m** | $Zn_{sat.}Hg$		0.0418	46Sto2
$Zn_{sat.}Hg$ | ZnI_2, 0.3 **m** ‖ ZnI_2, 0.1 **m** | $Zn_{sat.}Hg$		0.0269	46Sto2
$Zn_{sat.}Hg$ | ZnI_2, 0.3 **m** ‖ ZnI_2, 0.3 **m** | $Zn_{sat.}Hg$		0.0	46Sto2
$Zn_{sat.}Hg$ | ZnI_2, 0.3 **m** ‖ ZnI_2, 1.0 **m** | $Zn_{sat.}Hg$		–0.0409	46Sto2

Cell composition [2])	T [K, °C]	U_0 [V]	Ref.
$Zn_{sat.}Hg \mid ZnI_2$, 0.3 **m** $\Vert ZnI_2$, 2.5 **m** $\mid Zn_{sat.}Hg$		–0.0769	46Sto2
$Zn_{sat.}Hg \mid ZnI_2$, 0.3 **m** $\Vert ZnI_2$, 5.0 **m** $\mid Zn_{sat.}Hg$		–0.1168	46Sto2
$Zn_{sat.}Hg \mid ZnI_2$, 0.3 **m** $\Vert ZnI_2$, 10.0 **m** $\mid Zn_{sat.}Hg$		–0.2143	46Sto2
$Zn_{10wt\%}Hg \mid ZnSO_4$*6 H_2O,sol., $ZnSO_4$,sat., $Hg_2SO_4 \mid Hg$		1.4137	11Coh
$Zn_{10wt\%}Hg \mid ZnSO_4$*6 H_2O,sol., $ZnSO_4$,sat., $Hg_2SO_4 \mid Hg$	35 °C	1.4046	11Coh
$Zn_{10wt\%}Hg \mid ZnSO_4$*7 H_2O,sol., $ZnSO_4$,sat., $Hg_2SO_4 \mid Hg$		1.4197	11Coh
$Zn_{10wt\%}Hg \mid ZnSO_4$*7 H_2O,sol., $ZnSO_4$,sat., $Hg_2SO_4 \mid Hg$	35 °C	1.4062	11Coh
$Zn_{sat.}Hg \mid ZnSO_4$, 0.0008558 **m**, $Hg_2SO_4 \mid Hg$		1.5331	32Mas3
$Zn_{sat.}Hg \mid ZnSO_4$, 0.0020793 **m**, $Hg_2SO_4 \mid Hg$		1.5281	32Mas3
$Zn_{sat.}Hg \mid ZnSO_4$, 0.0039726 **m**, $Hg_2SO_4 \mid Hg$		1.5233	32Mas3
$Zn_{sat.}Hg \mid ZnSO_4$, 0.010109 **m**, $Hg_2SO_4 \mid Hg$		1.5137	32Mas3
$Zn_{sat.}Hg \mid ZnSO_4$, 0.04413 **m**, $Hg_2SO_4 \mid Hg$		1.4930	32Mas3
$Zn_{sat.}Hg \mid ZnSO_4$, 0.06099 **m**, $Hg_2SO_4 \mid Hg$		1.4897	32Mas3
$Zn_{sat.}Hg \mid ZnSO_4$, 0.10539 **m**, $Hg_2SO_4 \mid Hg$		1.4830	32Mas3
$Zn_{sat.}Hg \mid ZnSO_4$, 0.23255 **m**, $Hg_2SO_4 \mid Hg$		1.4729	32Mas3
$Zn_{sat.}Hg \mid ZnSO_4$, 0.46574 **m**, $Hg_2SO_4 \mid Hg$		1.4655	32Mas3
$Zn_{sat.}Hg \mid ZnSO_4$, 1.0 **m**, $Hg_2SO_4 \mid Hg$		1.4560	32Mas3
$Zn_{sat.}Hg \mid ZnSO_4$, 2.0 **m**, $Hg_2SO_4 \mid Hg$		1.4438	32Mas3
$Zn_{sat.}Hg \mid ZnSO_4$, 3.0 **m**, $Hg_2SO_4 \mid Hg$		1.4305	32Mas3
$Zn_{sat.}Hg \mid ZnSO_4$, 0.0005 **m**, Pb_2SO_4,sat. $\mid Pb_{sat.}Hg$		0.6114	31Cow
$Zn_{sat.}Hg \mid ZnSO_4$, 0.001 **m**, Pb_2SO_4,sat. $\mid Pb_{sat.}Hg$		0.5971	31Cow
$Zn_{sat.}Hg \mid ZnSO_4$, 0.002 **m**, Pb_2SO_4,sat. $\mid Pb_{sat.}Hg$		0.5832	31Cow
$Zn_{sat.}Hg \mid ZnSO_4$, 0.005 **m**, Pb_2SO_4,sat. $\mid Pb_{sat.}Hg$		0.5660	31Cow
$Zn_{sat.}Hg \mid ZnSO_4$, 0.01 **m**, Pb_2SO_4,sat. $\mid Pb_{sat.}Hg$		0.5535	31Cow
$Zn_{sat.}Hg \mid ZnSO_4$, 0.02 **m**, Pb_2SO_4,sat. $\mid Pb_{sat.}Hg$		0.5425	31Cow
$Zn_{sat.}Hg \mid ZnSO_4$, 0.05 **m**, Pb_2SO_4,sat. $\mid Pb_{sat.}Hg$		0.5287	31Cow
$Zn_{sat.}Hg \mid ZnSO_4$, 0.0512 **m**, Pb_2SO_4,sat. $\mid Pb_{sat.}Hg$	15 °C	0.5343	36Kie
	25 °C	0.5283	36Kie
	35 °C	0.5222	36Kie
$Zn_{sat.}Hg \mid ZnSO_4$, 0.150 **m**, Pb_2SO_4,sat. $\mid Pb_{sat.}Hg$	15 °C	0.5196	36Kie
	25 °C	0.5135	36Kie
	35 °C	0.5074	36Kie
$Zn_{sat.}Hg \mid ZnSO_4$, 0.510 **m**, Pb_2SO_4,sat. $\mid Pb_{sat.}Hg$	15 °C	0.5055	36Kie
	25 °C	0.4992	36Kie
	35 °C	0.4929	36Kie
$Zn_{sat.}Hg \mid ZnSO_4$, 1.501 **m**, Pb_2SO_4,sat. $\mid Pb_{sat.}Hg$	15 °C	0.4920	36Kie
	25 °C	0.4850	36Kie
	35 °C	0.4780	36Kie

[1]) Description of the cell composition is based on the Stockholm convention (1953). All data refer to room temperature RT (T = 298 K) if not stated otherwise.

[2]) Metal leads (e.g. platinum contacts) used solely as electrical connector and being identical at both cell terminals are not always explicitly mentioned.

[3]) Cell voltage in international Volt; in absolute Volt U = 1.01865 V.

[4]) If not stated otherwise hydrogen is present at standard pressure.

[5]) Saturated.

[6]) Most frequently platinum is used as electronically conducting phase of the hydrogen electrode, thus it is mentioned only when explicitly stated also in the original report.

[7]) Cell voltage with negative terminal at right-hand side of cell description (for details see introduction).

References

98Pet Peters, R.: Z. Phys. Chem. **26** (1898) 193.
99Dol Dolezalek, F.: Z. Elektrochem. **6** (1899) 517.
99Lut Luther, R.: Z. Phys. Chem. **30** (1899) 647.
00Tow Tower, O.F.: Z. Phys. Chem. **32** (1900) 566.
02Fre Fredenhagen, C.: Z. Anorg. Allg. Chem. **29** (1902) 396.
02Küs Küster, F.W., Hommel, W.: Z. Elektrochem. **8** (1902) 496.
03Con Conrad, W.: Dissertation Universität Göttingen, 1903.
03Ing Inglis, J.K.H.: Z. Elektrochem. **9** (1903) 226.
03Obe Oberer, E.: Dissertation ETH Zürich, 1903.
03Sac Sackur, O.: Arb. Gesund. **20** (1903) 539.
04Bod Bodländer, G., Eberlein, W.: Z. Anorg. Allg. Chem. **39** (1904) 197.
04Lab Labendzinski, S., Abegg, R.: Z. Elektrochem. **10** (1904) 77.
04Ner Nernst, W., Sand, J.: Z. Phys. Chem. **48** (1904) 601.
04Thi Thiel, A.: Z. Anorg. Allg. Chem. **39** (1904) 333.
05Abe Abegg, R., Spencer, J.F.: Z. Anorg. Allg. Chem. **46** (1905) 406.
05Lut1 Luther, R., Sammet, V.: Z. Elektrochem. **11** (1905) 293.
05Lut2 Luther, R., Sammet, V.: Z. Phys. Chem. **53** (1905) 645.
05Spi Spiker, F.: Z. Elektrochem. **11** (1905) 345.
06Faw Fawsitt, C.E.: J. Chem. Ind. **25** (1906) 1133.
06Zed Zedner, J.: Z. Elektrochem. **12** (1906) 463.
07Cam Campbell, F.H.: Chem. News J. Ind. Sci. **96** (1907) 15.
07Fau Faust, O.: Z. Elektrochem. **13** (1907) 161.
07Ric Richards, T.W., Behr, G.E.: Z. Phys. Chem. **58** (1907) 301.
08Brö Brönsted, J.N.: Z. Phys. Chem. (Leipzig) **65** (1908) 84.
08Lau Laurie, A.P.: Z. Phys. Chem. (Leipzig) **64** (1908) 615.
08Lut Luther, R., Michie, A.C.: Z. Elektrochem. **14** (1908) 826.
09Bri Brisle, F.J.: Trans. Faraday Soc. **4** (1909) 159.
09Pom Poma, G.: Atti Accad. Naz. Lincei **181** (1909) 192.
09Sch Schweitzer, A.: Z. Elektrochem. **15** (1909) 607.
10Brö Brönsted, J.N.: Z. Phys. Chem. **68** (1910) 703.
10Fed Fedotiew, P.: Z. Anorg. Allg. Chem. **69** (1910) 35.
10Lew Lewis, G.N., Kraus, Ch.A.: J. Am. Chem. Soc. **32** (1910) 1459.
10Ric Richards, T.W., Wilson, J.H.: Z. Phys. Chem. **72** (1910) 129.
10Sch1 Schaper, C.: Z. Phys. Chem. **72** (1910) 308.
10Ste Stegmüller, P.: Z. Elektrochem. **16** (1910) 90.
11Coh Cohen, E., Inouye, K., Euwen, C.: Z. Phys. Chem. **75** (1911) 1.
11För Förster, F., Yamasaki, J.: Z. Elektrochem. **17** (1911) 361.
11Lew Lewis, G.N., Rupert, F.F.: J. Am. Chem. Soc. **33** (1911) 299.
12Dru Drucker, C.: Z. Elektrochem. **18** (1912) 236.
12Lew Lewis, G.N., Keyes, F.G.: J. Am. Chem. Soc. **34** (1912) 119.
12Noy Noyes, A.A., Brann, B.F.: J. Am. Chem. Soc. **34** (1912) 1023.
13Lew Lewis, G.N., Keyes, F.G.: J. Am. Chem. Soc. **35** (1913) 340.
15For Forbes, G.S., Anderegg, F.O.: J. Am. Chem. Soc. **37** (1915) 1676.
15Lew Lewis, G.N., Argo, W.L.: J. Am. Chem. Soc. **37** (1915) 1983.
16Lin Linhart, G.A.: J. Am. Chem. Soc. **38** (1916) 2356
17For Forbes, G.S., Richter, H.W. : J. Am. Chem. Soc. **39** (1917) 1140.
17Lew1 Lewis, G.N., Storch, H.: J. Am. Chem. Soc. **39** (1917) 2544.
17Lew2 Lewis, G.N., Brighton, T.B.: J. Am. Chem. Soc. **39** (1917) 2245.
17Lew3 Lewis, G.N., Brighton, T.B.: J. Am. Chem. Soc. **39** (1917) 1906.
17Lin Linhart, G.A.: J. Am. Chem. Soc. **39** (1917) 2601.

18Hen Henderson, W.E., Stegeman, G.: J. Am. Chem. Soc. **40** (1918) 84.
18Noy Noyes, A.A., Chow, M.: J. Am. Chem. Soc. **40** (1918) 739.
18Oba Obata, J.: Proc. Tokyo Math. Phys. Soc. **9** (1918) 416; see also: Proc. Phys. Math. Soc. Jpn. **2** (1920) 3; **3** (1921) 136.
19Hof Hofmann, K.A., Wurthmann, B.: Ber. Dtsch. Chem. Ges. **52** (1919) 1185.
19Ric Richards, T.W., Daniels, F.: J. Am. Chem. Soc. **41** (1919) 1732.
19Öho Öholm, L.W.: Medd. K. Vetenskapsakad. Nobelinst. **5** (1919) 41.
20Lam Lamp, A.B., Larson, A.T.: J. Am. Chem. Soc. **42** (1920) 2024.
20Noy Noyes, A.A., Freed, E.S.: J. Am. Chem. Soc. **42** (1920) 476.
21Dru Drucker, C.: Z. Phys. Chem. **96** (1921) 381.
21Kol Kolthoff, J.M.: Z. Anorg. Allg. Chem. **119** (1921) 202.
21Moo Moore, W.C.: J. Am. Chem. Soc. **43** (1921) 81.
22App Applebey, M.P., Reid, R.D.: J. Chem. Soc. **121** (1922) 2129.
22Ger Gerke, R.H.: J. Am. Chem. Soc. **44** (1922) 1684.
22Har Harned, H.S., Brumbaugh, N.J.: J. Am. Chem. Soc. **44** (1922) 2729.
22Jel Jellinek, K., Czerwinski, J.: Z. Phys. Chem. **102** (1922) 438.
22Keo McKeown, A.: Trans. Faraday Soc. **17** (1922) 517.
22Men Menzel, H.: Z. Phys. Chem. **100** (1922) 276.
22Ric Richards, Th.W., Conant, J.B.: J. Am. Chem. Soc. **44** (1922) 601.
24Col Collenberg, O.: Z. Phys. Chem. **109** (1924) 353.
24Dru Drucker, C., Riethof, G.: Z. Phys. Chem. **111** (1924) 1.
24Hai Hainsworth, W.R., Rowley, H.J., McInnes, D.A.: J. Am. Chem. Soc. **46** (1924) 1437.
24Ric Richards, Th.W., Richards, W.T.: J. Am. Chem. Soc. **46** (1924) 89.
24Sch1 Schuhmann, R.: J. Am. Chem. Soc. **46** (1924) 52.
24Sch2 Schuhmann, R.: J. Am. Chem. Soc. **46** (1924) 58; 1444.
24Sch3 Schmid, A.: Helv. Chim. Acta **7** (1924) 370.
24Wyn Wynne-Jones, W.F.K., Hudlestone, H.: J. Chem. Soc. **125** (1924) 1031.
25For Forbes, G.S., Glass, S.W., Fuoss, R.M.: J. Am. Chem. Soc. **47** (1925) 2892.
25Luc Lucasse, W.W.: J. Am. Chem. Soc. **47** (1925) 743.
25Wil Wilke, E., Kieninger, O.: Z. Phys. Chem. **116** (1925) 215.
26Ake Akerlöf, G.: J. Am. Chem. Soc. **48** (1926) 1160; 1169.
26Bro Brown, D.J., Tefft, R.F.: J. Am. Chem. Soc. **48** (1926) 1128.
26Dru1 Drucker, C., Luft, F.: Z. Phys. Chem. **121** (1926) 307.
26Fri Fried, F.: Z. Phys. Chem. **123** (1926) 406.
26Gru Grube, G., Schlecht, L.: Z. Elektrochem. **32** (1926) 178.
26Mai Maier, Ch.G., Parks, G.S., Anderson, C.Fr.: J. Am. Chem. Soc. **48** (1926) 2564.
26Roy Royce, H.D., Kahlenberg, L.: Trans. Am. Electrochem. Soc. **50** (1926) 281.
26Vos Vosburgh, W.C.: J. Opt. Soc. Am. **12** (1926) 511.
27Bod Bodfors, P.: Z. Phys. Chem. **130** (1927) 82.
27Bre Brester, A.: Recl. Trav. Chim. Pays Bas **46** (1927) 328.
27Die Dietrich, H.G., Johnston, J.: J. Am. Chem. Soc. **49** (1927) 1419.
27Erd Erdey-Gruz, T.: Z. Phys. Chem. **131** (1927) 81.
27Ger Gerke, R.H., Rourke, M.D.: J. Am. Chem. Soc. **49** (1927) 1855.
27Gru Grube, G., Breitinger, G.: Z. Elektrochem. **33** (1927) 112.
27Ish Ishikawa, F., Kimura, G.: Chem. Inst. Dept. Kyoto Imp. Univ. (1927) 255.
28Get1 Getmann, F.H.: J. Phys. Chem. **32** (1928) 940.
28Get2 Getmann, F.H.: J. Phys. Chem. **32** (1928) 91.
28Har1 Harned, H.S., Robinson, R.A.: J. Am. Chem. Soc. **50** (1928) 3157.
28Har2 Harned, H.S., McHarris jr., J.A.: J. Am. Chem. Soc. **50** (1928) 2633.
28Jon Jones, G., Kaplan, B.B.: J. Am. Chem. Soc. **50** (1928) 1845; 2066.
28Kie Kiehl, S.J., Hart, D.: J. Am. Chem. Soc. **50** (1928) 1845; 2337.
28Mur Murata, K.: Bull. Chem. Soc. Jpn. **3** (1928) 57.
28Pry Prytz, M.: Z. Anorg. Allg. Chem. **172** (1928) 142.
28Vos Vosburgh, W.C.: J. Am. Chem. Soc. **50** (1928) 2386.

28Yos Yoshida, T.: Sci. Rep. Tohoku Univ. Ser. 1 **17** (1928) 1279.
29Bje Bjerrum, N., Unmack, A.: K. Dan. Vidensk. Selsk. Mat. Fys. Medd. **9** (1929) 5-206.
29Car Carmody, W.R.: J. Am. Chem. Soc. **51** (1929) 2901; 2905; 5908.
29Fos Fosbinder, R.J.: J. Am. Chem. Soc. **51** (1929) 1345.
29Har2 Harned, H.S.: J. Am. Chem. Soc. **51** (1929) 416.
29Isg Isgarischew, N., Turkowskaja, A.: Z. Phys. Chem. Abt. A **140** (1929) 227.
29Ish Ishikawa, F., Shibata, E.: Sci. Rep. Tohoku Univ. Ser. 1 **18** (1929) 109.
29Luc1 Lucasse, W.W.: J. Am. Chem. Soc. **51** (1929) 2597.
29Luc2 Lucasse, W.W.: J. Am. Chem. Soc. **51** (1929) 2605.
29Mül1 Müller, R., Kreiner, F., Schmidt, H.J.: Monatsh. Chem. **53** (1929) 215.
29Mül2 Müller, R., Schmidt, H.J.: Monatsh. Chem. **53** (1929) 224.
29Nim Nims, L.F., Bonner, W.D.: J. Phys. Chem. **33** (1929) 586.
30Bir Bird, E.W., Hixon, R.M.: J. Phys. Chem. **34** (1930) 1412.
30Foe Foerster, F., Böttcher, F.: Z. Phys. Chem. Abt. A **151** (1930) 321.
30Get Getmann, F.H.: J. Phys. Chem. **34** (1930) 454.
30Hak Hakomori, S.I.: J. Am. Chem. Soc. **52** (1930) 2373.
30Har1 Harned, H.S., Schupp, O.E.: J. Am. Chem. Soc. **52** (1930) 3886.
30Har2 Harned, H.S., Schupp, O.E.: J. Am. Chem. Soc. **52** (1930) 3892.
30Pri Priepke, R.J., Vosburg, W.: J. Am. Chem. Soc. **52** (1930) 4831.
30Sca1 Scatchard, G., Tefft, R.: J. Am. Chem. Soc. **52** (1930) 2265.
30Sca2 Scatchard, G., Tefft, R.: J. Am. Chem. Soc. **52** (1930) 2272.
30Smi Smith, E.R.: J. Res. Natl. Bur. Stand. **5** (1930) 735.
30Ter Terry, H., Baker, H.C.: J. Am. Chem. Soc. **52** (1930) 2583.
31Cow Cowperthwaite, J.A., La Mer, V.K.: J. Am. Chem. Soc. **53** (1931) 4333.
31Get1 Getmann, F.H.: J. Phys. Chem. **35** (1931) 2749.
31Get2 Getmann, F.H.: J. Phys. Chem. **35** (1931) 588.
31Har1 Harned, H.S., Murphy, G.M.: J. Am. Chem. Soc. **53** (1931) 8.
31Har2 Harned, H.S., Mason, C.M.: J. Am. Chem. Soc. **53** (1931) 3377.
31Kun Kunz, A.H.: J. Am. Chem. Soc. **53** (1931) 98.
31LaM La Mer, V.K., Parks, W.G.: J. Am. Chem. Soc. **53** (1931) 2040.
31Lar Larsson, E., Adell, B.: Z. Phys. Chem. Abt. A **156** (1931) 352.
31Lun Lundberg, W.O., Vestling, C.S., Ahlberg, J.E.: J. Am. Chem. Soc. **53** (1931) 352.
31McB McBain, J.W., van Rysselberghe, P.J., Squance, W.A.: J. Phys. Chem. **35** (1931) 999.
31Mur Murray, E.N., Acree, S.F.: J. Res. Natl. Bur. Stand. **7** (1931) 713.
31Pop Popoff, S., Riddick, J.A., Wirth, V.J., Ough, L.D.: J. Am. Chem. Soc. **53** (1931) 1195.
31Sho Show-Chow-Woo: J. Am. Chem. Soc. **53** (1931) 469.
32Bri Britton, H.T.S., Dodd, E.N.: J. Am. Chem. Soc. **54** (1932) 1940.
32Can Cann, J.Y., Sumner, R.A.: J. Phys. Chem. **36** (1932) 2615.
32Car Carter, S.R., Glover, T.J.: J. Phys. Chem. **36** (1932) 679.
32Ham Hamilton, R.T., Butler, J.A.V.: Proc. R. Soc. London **138** (1932) 450.
32Har1 Harned, H.S., Ehlers, R.W.: J. Am. Chem. Soc. **54** (1932) 1350.
32Har2 Harned, H.S., Mason, C.M.: J. Am. Chem. Soc. **54** (1932) 1439.
32Har3 Harned, H.S., Mason, C.M.: J. Am. Chem. Soc. **54** (1932) 3112.
32Has Hass, K., Jellinek, K.: Z. Phys. Chem. Abt. A **162** (1932) 153.
32Haw Hawkins, J.E.: J. Am. Chem. Soc. **54** (1932) 4480.
32Ish1 Ishikawa, F., Yoshida, T.: Sci. Rep. Tohoku Imp. Univ. Ser. 1 **21** (1932) 474.
32Ish2 Ishikawa, F., Shibata, E.: Sci. Rep. Tohoku Imp. Univ. Ser. 1 **21** (1932) 499.
32Mas1 Masaki, K., Ikkatai, T.: Bull. Chem. Soc. Jpn. **7** (1932) 233.
32Mas2 Masaki, K., Ikkatai, T.: Bull. Chem. Soc. Jpn. **7** (1932) 35.
32Mas3 Masaki, K., Ikkatai, T.: Bull. Chem. Soc. Jpn. **7** (1932) 238.
32Spe1 Spencer, H.M., Selden, R.F.: J. Am. Chem. Soc. **54** (1932) 4504.
32Spe2 Spencer, H.M., Mote, J.H.: J. Am. Chem. Soc. **54** (1932) 4618.
32Wil Wilke, E., Martin, W.: Z. Phys. Chem. Abt. A **160** (1932) 39.
33All Allmand, A.J., Burrage, L.J.: Trans. Faraday Soc. **29** (1933) 692.

33Bat Bates, St.J., Urmston, J.W.: J. Am. Chem. Soc. **55** (1933) 4068.
33Fli Flint, E.P., Wells, L.S.: J. Res. Natl. Bur. Stand. **11** (1933) 163.
33Gel Gelbach, R.W.: J. Am. Chem. Soc. **55** (1933) 4857.
33Get Getmann, F.H.: Trans. Am. Chem. Soc. **64** (1933) 201.
33Har1 Harned, H.S., Hamer, W.J.: J. Am. Chem. Soc. **55** (1933) 4496.
33Har2 Harned, H.S., Wright, D.D.: J. Am. Chem. Soc. **55** (1933) 4849.
33Har3 Harned, H.S., Ehlers, R.W.: J. Am. Chem. Soc. **55** (1933) 2179.
33Har4 Harned, H.S., Copson, H.R.: J. Am. Chem. Soc. **55** (1933) 2206.
33Har5 Harned, H.S., Hamer, W.J.: J. Am. Chem. Soc. **55** (1933) 2194.
33Har6 Harned, H.S., Ehlers, R.W.: J. Am. Chem. Soc. **55** (1933) 652.
33Har7 Harned, H.S., Hecker, J.C.: J. Am. Chem. Soc. **55** (1933) 4838.
33Ish1 Ishikawa, F., Kimura, G., Murooka, T.: Sci. Rep. Tohoku Imp. Univ. Ser. 1 **22** (1933) 131.
33Ish2 Ishikawa, F., Ueda, Y.: Sci. Rep. Tohoku Imp. Univ. Ser. 1 **22** (1933) 249.
33Ish3 Ishikawa, F., Murooka, T.: Sci. Rep. Tohoku Imp. Univ. Ser. 1 **22** (1933) 138.
33Ish4 Ishikawa, F., Ueda, Y.: Sci. Rep. Tohoku Imp. Univ. Ser. 1 **22** (1933) 263.
33Ish5 Ishikawa, F., Ferui, Y.: Bull. Inst. Phys. Chem. Res. - Rikagaku Kenkyusho Iho **12** (1933) 45.
33LaM La Mer, V.K., Parks, W.G.: J. Am. Chem. Soc. **55** (1933) 4343.
33Mac MacDougall, F.H., Blumer, D.R.: J. Am. Chem. Soc. **55** (1933) 2236.
33Mur Murdock, P.G., Barton, R.C.: J. Am. Chem. Soc. **55** (1933) 4074.
33Nim Nims, L.F: J. Am. Chem. Soc. **55** (1933) 1946.
33Oku Oku, M.: Sci. Rep. Tohoku Imp. Univ. Ser. 1 **22** (1933) 288.
33Ued1 Ueda, Y.: Sci. Rep. Tohoku Imp. Univ. Ser. 1 **22** (1933) 448.
33Ued2 Ueda, Y.: Sci. Rep. Tohoku Imp. Univ. Ser. 1 **22** (1933) 879.
33Van1 Vance, J.E.: J. Am. Chem. Soc. **55** (1933) 2729.
33Van2 Vance, J.E.: J. Am. Chem. Soc. **55** (1933) 4519.
34Arm Armbruster, M.H., Crenshaw, J.L.: J. Am. Chem. Soc. **56** (1934) 2525.
34Car Carpenter, J.E.: J. Am. Chem. Soc. **56** (1934) 1874.
34Cow Cowperthwaite, J.A., La Mer, V.K., Barksdale, J.: J. Am. Chem. Soc. **56** (1934) 544.
34Cro Crockford, H.D., Simmons, N.L.: J. Am. Chem. Soc. **56** (1934) 1537.
34Ham Hamer, W.J.: J. Am. Chem. Soc. **56** (1934) 860.
34Har2 Harned, H.S., Embree, N.D.: J. Am. Chem. Soc. **56** (1934) 1042.
34Har3 Harned, H.S., Sutherland, R.O.: J. Am. Chem. Soc. **56** (1934) 2039.
34Har4 Haring, M.M., Westfall, B.B.: Trans. Am. Electrochem. Soc. **63** (1934) 61.
34Hoy Hoyt, C.S., Stegemann, G.: J. Phys. Chem. **38** (1934) 753.
34Ish1 Ishikawa, F., Yamasaki, S., Murooko, T.: Sci. Rep. Tohoku Imp. Univ. Ser. 1 **23** (1934) 115.
34Jon Jones, G., Baeckström, S.: J. Am. Chem. Soc. **56** (1934) 1524.
34Mas Mason, Ch.M., Kellam, D.B.: J. Phys. Chem. **38** (1934) 689.
34Owe Owen, B.B.: J. Am. Chem. Soc. **56** (1934) 1695.
34Par Parks, W.G., La Mer, V.K.: J. Am. Chem. Soc. **56** (1934) 90.
34Pry Prytz, M.: Z. Anorg. Allg. Chem. **219** (1934) 89.
34Sch Scherer, G.A., Newton, R.E.: J. Am. Chem. Soc. **56** (1934) 18.
34Shr2 Shrawder jr., J., Cowperthwaite, J.A.: J. Am. Chem. Soc. **56** (1934) 2340.
34Tip Tippetts, E.A., Newton, R.F.: J. Am. Chem. Soc. **56** (1934) 1675.
34Öho Öholm, L.W.: Finska Kemistsamf. Medd. **43** (1934) 25.
35Bra Brasher, D.M., Parton, H.N.: Trans. Faraday Soc. **31** (1935) 681.
35Ham2 Hamer, W.J.: J. Am. Chem. Soc. **57** (1935) 662.
35Har Harned, H.S., Hamer, W.J.: J. Am. Chem. Soc. **57** (1935) 27.
35Hov Hovorka, F., Dearing, W.C.: J. Am. Chem. Soc. **57** (1935) 446.
35Kri Krieble, V.K., Reinhart, F.M.: J. Am. Chem. Soc. **57** (1935) 19.
35Owe Owen, B.B.: J. Am. Chem. Soc. **57** (1935) 1926.
35Par Parton, H.N.: Trans. Faraday Soc. **31** (1935) 686.
36Cro Crockford, H.D., Farr jr., H.O.: J. Am. Chem. Soc. **58** (1936) 87.
36Har1 Harned, H.S., Fitzgerald, M.: J. Am. Chem. Soc. **58** (1936) 2624.
36Har2 Harned, H.S., Keston, A.S., Donelson, J.G.: J. Am. Chem. Soc. **58** (1936) 989.

36Kie Kielland, J.: J. Am. Chem. Soc. **58** (1936) 1855.
36Noy Noyes, A.A., Garner, C.S.: J. Am. Chem. Soc. **58** (1936) 1268.
36Owe Owen, B.B., Foering, L.: J. Am. Chem. Soc. **58** (1936) 1575.
36She Sherill, M.S., Haas jr., A.J.: J. Am. Chem. Soc. **58** (1936) 952.
37Bat Bates, R.G., Vosburg, W.C.: J. Am. Chem. Soc. **59** (1937) 1188.
37Can1 Cann, J.Y., Taylor, A.C.: J. Am. Chem. Soc. **59** (1937) 1484.
37Can2 Cann, J.Y., Taylor, A.C.: J. Am. Chem. Soc. **59** (1937) 1987.
37Har1 Harned, H.S., Donelson, J.G.: J. Am. Chem. Soc. **59** (1937) 1280.
37Har3 Harned, H.S., Cook, M.A.: J. Am. Chem. Soc. **59** (1937) 1890.
37Noy Noyes, A.A., Deahl, T.J.: J. Am. Chem. Soc. **59** (1937) 1337.
38Cro Crockford, H.D., Loffin, J.C.: J. Am. Chem. Soc. **60** (1938) 1606.
38Tay Taylor, A.C., Nims, L.F.: J. Am. Chem. Soc. **60** (1938) 262.
39Har2 Harned, H.S., Fallon, L.D.: J. Am. Chem. Soc. **61** (1939) 3111.
39Har4 Harned, H.S., Dreby, E.C.: J. Am. Chem. Soc. **61** (1939) 3113.
39Lar Larson, W.D., Tomsicek, W.J.: J. Am. Chem. Soc. **61** (1939) 65.
39Vos Vosburgh, W.C., Derr, P.F., Cooper, G.R., Bates, R.G.: J. Am. Chem. Soc. **61** (1939) 2592.
41Har Harned, H.S., Scholes jr., S.R.: J. Am. Chem. Soc. **63** (1941) 1706.
41Lar Larson, W.D., Tomsicek, W.J.: J. Am. Chem. Soc. **63** (1941) 3329.
42Bat Bates, R.G.: J. Am. Chem. Soc. **64** (1942) 1136.
42Hug Hughes, R.H., Garner, C.S.: J. Am. Chem. Soc. **64** (1942) 1644.
43She Sherrill, M.S., King, C.B., Spooner, R.C.: J. Am. Chem. Soc. **65** (1943) 170.
43Tem Templeton, D.H., Watt, G.W., Gardner, C.S.: J. Am. Chem. Soc. **65** (1943) 1608.
46Sto1 Stokes, R.H., Levien, B.J.: J. Am. Chem. Soc. **68** (1946) 333.
46Sto2 Stokes, R.H., Levien, B.J.: J. Am. Chem. Soc. **68** (1946) 1853.
47Mas Mason, C.M., Blum, W.M.: J. Am. Chem. Soc. **69** (1947) 1246.
48Fer Ferrell jr., D.T., Blackburn, I., Vosburgh, W.C.: J. Am. Chem. Soc. **70** (1948) 3812.
49Bro Brown, R.A., Swift, E.H.: J. Am. Chem. Soc. **71** (1949) 2719.
49Kra Kraus, K.A., Nelson, F., Johnson, G.L.: J. Am. Chem. Soc. **71** (1949) 2520.
49deV deVries, T., Cohen, D.: J. Am. Chem. Soc. **71** (1949) 1114.
50Wil Williams, J.P., Knight, S.B., Crockford, H.D.: J. Am. Chem. Soc. **72** (1950) 1277.
52Coh Cohen, D., Hindman, J.C.: J. Am. Chem. Soc. **74** (1952) 4679; 4682.
52Dav Davis, C.W., Jones, H.W., Monk, C.B.: Trans. Faraday Soc. **48** (1952) 921.
52Jon Jones, H.W., Monk, C.B.: Trans. Faraday Soc. **48** (1952) 929.
52Swi Swinehart, D.F.: J. Am. Chem. Soc. **74** (1952) 1100.
53Har Harned, H.S., Paxton, T.R.: J. Phys. Chem. **57** (1953) 531.
54Bat2 Bates, R.G., Bowers, V.E.: J. Res. Natl. Bur. Stand. U.S. **53** (1954) 282.
55Lie Lietzke, M.H., Vaughan, J.V.: J. Am. Chem. Soc. **77** (1955) 876.
55Sti Stille, U.: Messen und Rechnen in der Physik, Braunschweig: Vieweg, 1955, p. 217.
56Bru Bruckenstein, S., Kolthoff, I.M.: J. Am. Chem. Soc. **78** (1956) 2974.
56Pin Pinfold, T.A., Sebba, F.: J. Am. Chem. Soc. **78** (1956) 2095; 5193.
57Ham Hamer, W.G., Craig, D.N.: J. Electrochem. Soc. **104** (1957) 206.
58Tre1 Treumann, W.B., Ferris, L.M.: J. Am. Chem. Soc. **80** (1958) 5048.
58Tre2 Treumann, W.B., Ferris, L.M.: J. Am. Chem. Soc. **80** (1958) 5050.
60Gre Greeley, R.S., Smith, W.T., Stoughton, R.W., Lietzke, M.H.: J. Phys. Chem. **64** (1960) 652; 1445.

Table 2.2. Cell voltages with nonaqueous [1]) electrolyte systems. [2])

Cell composition	T [°C]	U_0 [V]	Ref.
Barium (Ba)			
$Ba_{0.2629wt\%}Hg \mid BaCl_2$, hydrazine $\mid Ba_{0.01083wt\%}Hg$		0.04712	26And
$Ba_{0.2629wt\%}Hg \mid BaCl_2$, hydrazine $\mid Ba_{0.01894wt\%}Hg$		0.03953	26And
$Ba_{0.2629wt\%}Hg \mid BaCl_2$, hydrazine $\mid Ba_{0.0503wt\%}Hg$		0.02624	26And
$Ba_{0.2629wt\%}Hg \mid BaCl_2$, hydrazine $\mid Ba_{0.1160wt\%}Hg$		0.01396	26And
$Ba_{0.2629wt\%}Hg \mid BaCl_2$, hydrazine $\mid Ba_{0.2085wt\%}Hg$		0.00427	26And
$Ba_{0.2629wt\%}Hg \mid BaCl_2$, hydrazine $\mid Ba_{0.1228wt\%}Hg$	15	0.01256	26And
	25	0.01291	26And
	35	0.01326	26And
$Ba_{0.1228wt\%}Hg \mid BaCl_2$, hydrazine $\mid Ba_{0.03106wt\%}Hg$	15	0.01935	26And
	25	0.01998	26And
	35	0.02061	26And
$Ba_{0.2679wt.\%}Hg \mid BaCl_2$,pyridine $\mid Ba_{0.01083wt.\%}Hg$		0.0471	26And
$Ba_{0.2679wt.\%}Hg \mid BaCl_2$,pyridine $\mid Ba_{0.1160wt.\%}Hg$		0.0134	26And
$Ba_{0.2679wt.\%}Hg \mid BaCl_2$,pyridine $\mid Ba_{0.2085wt.\%}Hg$		0.0043	26And
Cadmium (Cd)			
$Cd \mid Cd(NO_3)_2$, 0.05 **M**, $NH_3 \parallel AgNO_3$, 0.1 **M**, $NH_3 \mid Ag$	–35	0.963	26Cos
(U_T [V]=0.963–0.33·10^{-3} (T+35 °C), –65<T<–35 °C)			
$Cd_{sat.}Hg \mid CdF_2$,sol. $\parallel$ HF(NaF), Hg_2F_2,sol. $\mid$ Hg	10	1.0337	52Koe
$Cd_{sat.}Hg \mid CdF_2$,sol. $\parallel$ HF(NaF), Hg_2F_2,sol. $\parallel PbF_2$*2.5 H_2O,sol. $\mid Pb_{sat.}Hg$	0	0.030	52Koe
$Cd \mid Cd(NO_3)_2$, 0.05 **M**, pyridine $\mid AgNO_3$, 0.1 **M**, pyridine $\mid$ Ag		0.746	99Kah
$Cd \mid CdI_2$, 0.00465 **M**, acetonitrile $\parallel CdI_2$, 0.0204 **M**, acetonitrile $\parallel$ Cd		0.032	00Kah
$Cd \mid Cd(ClOO_4)_2$, 0.01 **M**, methanol $\parallel NaClO_4$, 0.01 **m**, methanol $\parallel$ NaCl, 0.01 **m**, AgCl, methanol $\mid$ Ag		0.570	29Buc
$Cd_{9at\%}Hg \mid CdI_2$,sol., CdI_2,sat., Hg_2I_2,sol., methanol $\mid$ Hg	25	0.4172	28Yos
	30	0.4190	28Yos
	35	0.4208	28Yos
$Cd_{9at\%}Hg \mid CdI_2$,sol., CdI_2,sat., Hg_2I_2,sol., ethanol $\mid$ Hg	25	0.4172	28Yos
	30	0.4192	28Yos
	35	0.4208	28Yos
$Cd_zHg \mid CdI_2$,sol., CdI_2,sat., Hg_2I_2,sol., acetone $\mid$ Hg	22.5	0.4169	28Yos
	25.0	0.4178	28Yos
	30	0.4195	28Yos
Calcium (Ca)			
$Ca \mid CaI_2$, pyridine $\mid Ca_{0.0253wt\%}Hg$		0.0843	26Dru1
$Ca_{1at\%}Hg \mid CaI_2$, 0.0093 **M**, pyridine $\mid AgNO_3$, 0.1 **M**, pyridine $\mid$ Ag		1.85	15Cam
$Ca_{5at\%}Hg \mid CaI_2$, 0.0093 **M**, pyridine $\mid AgNO_3$, 0.1 **M**, pyridine $\mid$ Ag		1.82	15Cam
$Ca_{10at\%}Hg \mid CaI_2$, 0.0093 **M**, pyridine $\mid AgNO_3$, 0.1 **M**, pyridine $\mid$ Ag		1.84	15Cam
$Ca_{20at\%}Hg \mid CaI_2$, 0.0093 **M**, pyridine $\mid AgNO_3$, 0.1 **M**, pyridine $\mid$ Ag		1.82	15Cam
$Ca_{25at\%}Hg \mid CaI_2$, 0.0093 **M**, pyridine $\mid AgNO_3$, 0.1 **M**, pyridine $\mid$ Ag		1.99	15Cam
$Ca \mid CaCl_2$, 0.25 **m**, methanol, $Hg_2Cl_2 \mid$ Hg	–80	2.20	15Cam
$Ca_{0.25wt\%}Hg \mid CaCl_2$, 0.25 **m**, methanol, $Hg_2Cl_2 \mid$ Hg	–80	1.37	15Cam
$Ca_{1.06wt\%}Hg \mid CaCl_2$, 0.25 **m**, methanol, $Hg_2Cl_2 \mid$ Hg	–80	1.70	15Cam
$Ca_{4.06wt\%}Hg \mid CaCl_2$, 0.25 **m**, methanol, $Hg_2Cl_2 \mid$ Hg	–80	1.82	15Cam
$Ca_{7.90wt\%}Hg \mid CaCl_2$, 0.25 **m**, methanol, $Hg_2Cl_2 \mid$ Hg	–80	2.02	15Cam

Cell composition	T [°C]	U_0 [V]	Ref.
Cerium (Ce)			
Ce | $CeBr_3$, pyridine | $Ce_{5wt\%}Hg$		0.245	26Mül
Ce | $CeBr_3$,sat., pyridine | $Ce_{5wt\%}Hg$		0.24	29Mül2
$Ce_{5wt\%}Hg$ | $CeBr_3$, pyridine | $AgNO_3$, 0.1 **M**, pyridine | Ag		0.965	29Mül2
Ce | $CeBr_3$,sat., pyridine | $AgNO_3$, 0.1 **M**, pyridine | Ag		0.72	29Mül2
Copper (Cu)			
Cu | $Cu(NO_3)_2$,sat., pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		0.210	99Kah
Cu | $Cu(NO_3)_2$, 0.05 **M**, aniline ‖ $AgNO_3$, 0.1 **M**, aniline | Ag		0.210	99Kah
Hydrogen (H)			
Pt, H_2 | HBr, 0.474 m**m**, methanol, AgBr | Ag		0.2639	42Kan
Pt, H_2 | HBr, 1.274 m**m**, methanol, AgBr | Ag		0.2155	42Kan
Pt, H_2 | HBr, 9.36 m**m**, methanol, AgBr | Ag		0.1217	42Kan
Pt, H_2 | HBr, 54.30 m**m**, methanol, AgBr | Ag		0.0439	42Kan
Ir, H_2 | HCl, 0.4955 m**m**, methanol, AgCl | Ag		0.3861	25Non
Ir, H_2 | HCl, 1.652 m**m**, methanol, AgCl | Ag		0.3274	25Non
Ir, H_2 | HCl, 4.939 m**m**, methanol, AgCl | Ag		0.2757	25Non
Ir, H_2 | HCl, 7.113 m**m**, methanol, AgCl | Ag		0.2698	25Non
Ir, H_2 | HCl, 14.44 m**m**, methanol, AgCl | Ag		0.2285	25Non
Ir, H_2 | HCl, 25.49 m**m**, methanol, AgCl | Ag		0.2049	25Non
Ir, H_2 | HCl, 53.12 m**m**, methanol, AgCl | Ag		0.1758	25Non
Ir, H_2 | HCl, 75.07 m**m**, methanol, AgCl | Ag		0.1612	25Non
Ir, H_2 | HCl, 0.1155 **m**, methanol, AgCl | Ag		0.1439	25Non
Ir, H_2 | HCl, 0.5574 **m**, methanol, AgCl | Ag		0.0844	25Non
H_2 | HCl, 1 **m**, methanol(10wt%)/water | AgCl |	20	0.21818	61Ive
H_2 | HCl, 1 **m**, methanol(20wt%)/water | AgCl |	20	0.21151	61Ive
H_2 | HCl, 1 **m**, methanol(43.3wt%)/water | AgCl |	25	0.1941	61Ive
H_2 | HCl, 1 **m**, methanol(64.0wt%)/water | AgCl |	25	0.1764	61Ive
H_2 | HCl, 1 **m**, methanol(84.2wt%)/water | AgCl |	25	0.1319	61Ive
H_2 | HCl, 1 **m**, methanol(94.2wt%)/water | AgCl |	25	0.0840	61Ive
H_2 | HCl, 1 **m**, methanol | AgCl |	20	–0.0099	61Ive
H_2 | HCl, 1 **m**, ethanol(10wt%)/water | AgCl |	20	0.21900	61Ive
H_2 | HCl, 1 **m**, ethanol(20wt%)/water | AgCl |	20	0.21025	61Ive
H_2 | HCl, 1 **m**, ethanol | AgCl |	20	–0.0813	61Ive
H_2 | HCl, 1 **m**, 2-propanol(5wt%)/water | AgCl |	20	0.22110	61Ive
H_2 | HCl, 1 **m**, 2-propanol(10wt%)/water | AgCl |	20	0.21660	61Ive
H_2 | HCl, 1 **m**, 2-propanol(20wt%)/water | AgCl |	20	0.20905	61Ive
H_2 | HCl, 1 **m**, glycerol(10wt%)/water | AgCl |	25	0.21650	61Ive
H_2 | HCl, 1 **m**, glycerol(30wt%)/water | AgCl |	25	0.20221	61Ive
H_2 | HCl, 1 **m**, glycerol(50wt%)/water | AgCl |	25	0.18392	61Ive
H_2 | HCl, 1 **m**, acetone | AgCl |	20	–0.53	61Ive
H_2 | HCl, 1 **m**, 1,4-dioxane(20wt%)/water | AgCl |	25	0.20303	61Ive
H_2 | HCl, 1 **m**, 1,4-dioxane(45wt%)/water | AgCl |	25	0.16352	61Ive
H_2 | HCl, 1 **m**, 1,4-dioxane(70wt%)/water | AgCl |	25	0.06395	61Ive
H_2 | HCl, 1 **m**, 1,4-dioxane(82wt%)/water | AgCl |	25	0.0611	61Ive
H_2 | HCl, 1 **m**, formamide | AgCl |	20	0.204	58Man
H_2 | HCl, 1 **m**, formic acid | AgCl |	20	–0.1200	61Ive
H_2 | HCl, 1 **m**, acetic acid | AgCl |	20	–0.6180	61Ive

Cell composition	T [°C]	U_0 [V]	Ref.
Pt, H_2 | CH_3COOH, 0.01 **m**, methanol, CH_3COONa, 0.00976 **m** ‖ NaCl, 0.01 **m**, methanol ‖ HCl, 0.00976 **m** | H_2, Pt		0.4080	24Bje
Pt, H_2 | CH_3COOH, 0.1 **m**, methanol, CH_3COONa, 0.0966 **m** ‖ NaCl, 0.1 **m**, methanol ‖ HCl, 0.0966 **m** | H_2, Pt		0.4452	24Bje
Pt, H_2 | CH_3COOH, 0.0382 **M**, methanol ‖ NaCl, 0.1 **M**, Hg_2Cl_2 | Hg		0.6688	34Goo
Pt, H_2 | Na-methylat, 0.0217 **m**, methanol ‖ NaCl, 0.01 **m**, AgCl | Ag		0.9930	29Buc
Pt, H_2 | HCl, 0.02 **m**, methanol(20wt%)/water, AgCl | Ag		0.4191	30Ake
Pt, H_2 | HCl, 0.02 **m**, methanol(40wt%)/water, AgCl | Ag		0.4075	30Ake
Pt, H_2 | HCl, 0.02 **m**, methanol(60wt%)/water, AgCl | Ag		0.3915	30Ake
Pt, H_2 | HCl, 0.02 **m**, methanol(90wt%)/water, AgCl | Ag		0.3292	30Ake
Pt, H_2 | HCl, 0.05 **m**, methanol(20wt%)/water, AgCl | Ag		0.3753	30Ake
Pt, H_2 | HCl, 0.05 **m**, methanol(40wt%)/water, AgCl | Ag		0.3644	30Ake
Pt, H_2 | HCl, 0.05 **m**, methanol(60wt%)/water, AgCl | Ag		0.3496	30Ake
Pt, H_2 | HCl, 0.05 **m**, methanol(90wt%)/water, AgCl | Ag		0.2902	30Ake
Pt, H_2 | HCl, 0.5 **m**, methanol(20wt%)/water, AgCl | Ag		0.2621	30Ake
Pt, H_2 | HCl, 0.5 **m**, methanol(40wt%)/water, AgCl | Ag		0.2534	30Ake
Pt, H_2 | HCl, 0.5 **m**, methanol(60wt%)/water, AgCl | Ag		0.2417	30Ake
Pt, H_2 | HCl, 0.5 **m**, methanol(90wt%)/water, AgCl | Ag		0.1910	30Ake
Pt, H_2 | HCl, 4.393 m**m**, H_2SO_4, 4.654 m**m**, methanol(10wt%)/water, AgCl | Ag	20	0.4749	53Eva
	25	0.4767	53Eva
	30	0.4784	53Eva
	35	0.4798	53Eva
Pt, H_2 | HCl, 5.549 m**m**, H_2SO_4, 6.182 m**m**, methanol(10wt%)/water, AgCl | Ag	20	0.4633	53Eva
	25	0.4651	53Eva
	30	0.4665	53Eva
	35	0.4678	53Eva
Pt, H_2 | HCl, 8.866 m**m**, H_2SO_4, 18.901 m**m**, methanol(10wt%)/water, AgCl | Ag	20	0.4326	53Eva
	25	0.4339	53Eva
	30	0.4349	53Eva
	35	0.4356	53Eva
Pt, H_2 | HCl, 8.795 m**m**, H_2SO_4, 24.623 m**m**, methanol(10wt%)/water, AgCl | Ag	20	0.4287	53Eva
	25	0.4299	53Eva
	30	0.4309	53Eva
	35	0.4319	53Eva
Pt, H_2 | HCl, 13.268 m**m**, H_2SO_4, 14.008 m**m**, methanol(10wt%)/water, AgCl | Ag	20	0.4237	53Eva
	25	0.4248	53Eva
	35	0.4261	53Eva
Pt, H_2 | HCl, 23.935 m**m**, H_2SO_4, 31.290 m**m**, methanol(10wt%)/water, AgCl | Ag	20	0.3940	53Eva
	25	0.3945	53Eva
	30	0.3948	53Eva
	35	0.3948	53Eva
Pt, H_2 | HCl, 4.426 m**m**, H_2SO_4, 28.239 m**m**, methanol(20wt%)/water, AgCl | Ag	20	0.4414	53Eva
	25	0.4429	53Eva
	30	0.4443	53Eva
	35	0.4456	53Eva
Pt, H_2 | HCl, 4.586 m**m**, H_2SO_4, 6.224 m**m**, methanol(20wt%)/water, AgCl | Ag	20	0.4648	53Eva
	25	0.4668	53Eva
	30	0.4683	53Eva
	35	0.4698	53Eva

Cell composition	T [°C]	U_0 [V]	Ref.
Pt, H_2 | HCl, 5.435 m**m**, H_2SO_4, 6.437 m**m**, methanol(20wt%)/water, AgCl | Ag	20	0.4588	53Eva
	25	0.4605	53Eva
	30	0.4620	53Eva
	35	0.4634	53Eva
Pt, H_2 | HCl, 6.110 m**m**, H_2SO_4, 4.335 m**m**, methanol(20wt%)/water, AgCl | Ag	20	0.4592	53Eva
	25	0.4609	53Eva
	30	0.4623	53Eva
	35	0.4637	53Eva
Pt, H_2 | HCl, 6.509 m**m**, H_2SO_4, 9.975 m**m**, methanol(20wt%)/water, AgCl | Ag	20	0.4469	53Eva
	25	0.4485	53Eva
	30	0.4499	53Eva
	35	0.4510	53Eva
Pt, H_2 | HCl, 14.179 m**m**, H_2SO_4, 4.286 m**m**, methanol(20wt%)/water, AgCl | Ag	20	0.4267	53Eva
	25	0.4277	53Eva
Pt, H_2 | HCl, 0.1 **M** ‖ HCl, 0.01 **M**, ethanol(16.18vol%)/water | H_2, Pt		0.0192	27Erd
Pt, H_2 | HCl, 0.1 **M** ‖ HCl, 0.01 **M**, ethanol(32.35vol%)/water | H_2, Pt		0.0169	27Erd
Pt, H_2 | HCl, 0.1 **M** ‖ HCl, 0.01 **M**, ethanol(52.35vol%)/water | H_2, Pt		0.0216	27Erd
Pt, H_2 | HCl, 0.1 **M** ‖ HCl, 0.01 **M**, ethanol(67.64vol%)/water | H_2, Pt		0.0275	27Erd
Pt, H_2 | HCl, 0.1 **M** ‖ HCl, 0.01 **M**, ethanol | H_2, Pt		0.044	27Erd
Pt, H_2 | HCl, AgCl, ethanol(10wt%)/water | Ag	20	0.21898	42Pat
(U_T [V]=0.21898+4.89247·10^{-4} (T–20 °C)–1.40006·10^{-4} (T–20 °C)^2)			
Pt, H_2 | HCl, AgCl, ethanol(20wt%)/water | Ag	20	0.21011	42Pat
(U_T [V]=0.21011+4.08604·10^{-4} (T–20 °C)–5.57748·10^{-4} (T–20 °C)^2)			
Pt, H_2 | HCl, 0.005124 **m**, AgCl, ethanol(30wt%)/water | Ag		0.4769	54Har
Pt, H_2 | HCl, 0.009768 **m**, AgCl, ethanol(30wt%)/water | Ag		0.4453	54Har
Pt, H_2 | HCl, 0.016212 **m**, AgCl, ethanol(30wt%)/water | Ag		0.4212	54Har
Pt, H_2 | HCl, 0.029336 **m**, AgCl, ethanol(30wt%)/water | Ag		0.3931	54Har
Pt, H_2 | HCl, 0.045850 **m**, AgCl, ethanol(30wt%)/water | Ag		0.3720	54Har
Pt, H_2 | HCl, 0.053154 **m**, AgCl, ethanol(30wt%)/water | Ag		0.3652	54Har
Pt, H_2 | HCl, 0.006306 **m**, AgCl, ethanol(40wt%)/water | Ag		0.4618	54Har
Pt, H_2 | HCl, 0.013185 **m**, AgCl, ethanol(40wt%)/water | Ag		0.4263	54Har
Pt, H_2 | HCl, 0.019214 **m**, AgCl, ethanol(40wt%)/water | Ag		0.4084	54Har
Pt, H_2 | HCl, 0.021231 **m**, AgCl, ethanol(40wt%)/water | Ag		0.4038	54Har
Pt, H_2 | HCl, 0.022995 **m**, AgCl, ethanol(40wt%)/water | Ag		0.3999	54Har
Pt, H_2 | HCl, 0.041587 **m**, AgCl, ethanol(40wt%)/water | Ag		0.3732	54Har
Pt, H_2 | HCl, 0.006027 **m**, AgCl, ethanol(50wt%)/water | Ag		0.4566	54Har
Pt, H_2 | HCl, 0.011901 **m**, AgCl, ethanol(50wt%)/water | Ag		0.4241	54Har
Pt, H_2 | HCl, 0.021062 **m**, AgCl, ethanol(50wt%)/water | Ag		0.3975	54Har
Pt, H_2 | HCl, 0.028877 **m**, AgCl, ethanol(50wt%)/water | Ag		0.3827	54Har
Pt, H_2 | HCl, 0.051775 **m**, AgCl, ethanol(50wt%)/water | Ag		0.3561	54Har
Pt, H_2 | HCl, 0.070456 **m**, AgCl, ethanol(50wt%)/water | Ag		0.3420	54Har
Pt, H_2 | HCl, 0.01 **m**, AgCl, ethanol(25mol%)/water | Ag		0.4361	29But
Pt, H_2 | HCl, 0.05 **m**, AgCl, ethanol(25mol%)/water | Ag		0.3606	29But
Pt, H_2 | HCl, 0.1 **m**, AgCl, ethanol(25mol%)/water | Ag		0.3281	29But
Pt, H_2 | HCl, 0.3 **m**, AgCl, ethanol(25mol%)/water | Ag		0.2778	29But
Pt, H_2 | HCl, 0.6 **m**, AgCl, ethanol(25mol%)/water | Ag		0.2431	29But
Pt, H_2 | HCl, 1.0 **m**, AgCl, ethanol(25mol%)/water | Ag		0.2141	29But
Pt, H_2 | HCl, 0.01 **m**, AgCl, ethanol(50mol%)/water | Ag		0.4033	29But

Cell composition	T [°C]	U_0 [V]	Ref.
Pt, H_2 | HCl, 0.05 **m**, AgCl, ethanol(50mol%)/water | Ag		0.3331	29But
Pt, H_2 | HCl, 0.1 **m**, AgCl, ethanol(50mol%)/water | Ag		0.3034	29But
Pt, H_2 | HCl, 0.3 **m**, AgCl, ethanol(50mol%)/water | Ag		0.2550	29But
Pt, H_2 | HCl, 0.6 **m**, AgCl, ethanol(50mol%)/water | Ag		0.2209	29But
Pt, H_2 | HCl, 1.0 **m**, AgCl, ethanol(50mol%)/water | Ag		0.1926	29But
Pt, H_2 | HCl, 0.01 **m**, AgCl, ethanol(75mol%)/water | Ag		0.3563	29But
Pt, H_2 | HCl, 0.05 **m**, AgCl, ethanol(75mol%)/water | Ag		0.2902	29But
Pt, H_2 | HCl, 0.1 **m**, AgCl, ethanol(75mol%)/water | Ag		0.2622	29But
Pt, H_2 | HCl, 0.3 **m**, AgCl, ethanol(75mol%)/water | Ag		0.2156	29But
Pt, H_2 | HCl, 0.6 **m**, AgCl, ethanol(75mol%)/water | Ag		0.1825	29But
Pt, H_2 | HCl, 1.0 **m**, AgCl, ethanol(75mol%)/water | Ag		0.1524	29But
Pt, H_2 | HCl, 0.01 **m**, AgCl, ethanol(95mol%)/water | Ag		0.2752	29But
Pt, H_2 | HCl, 0.05 **m**, AgCl, ethanol(95mol%)/water | Ag		0.2119	29But
Pt, H_2 | HCl, 0.1 **m**, AgCl, ethanol(95mol%)/water | Ag		0.1839	29But
Pt, H_2 | HCl, 0.3 **m**, AgCl, ethanol(95mol%)/water | Ag		0.1374	29But
Pt, H_2 | HCl, 0.6 **m**, AgCl, ethanol(95mol%)/water | Ag		0.1013	29But
Pt, H_2 | HCl, 1.0 **m**, AgCl, ethanol(95mol%)/water | Ag		0.0691	29But
Pt, H_2 | HCl, 0.01 **m**, AgCl, ethanol | Ag		0.1938	29But
Pt, H_2 | HCl, 0.05 **m**, AgCl, ethanol | Ag		0.1390	29But
Pt, H_2 | HCl, 0.1 **m**, AgCl, ethanol | Ag		0.1144	29But
Pt, H_2 | HCl, 0.3 **m**, AgCl, ethanol | Ag		0.0806	29But
Pt, H_2 | HCl, 0.6 **m**, AgCl, ethanol | Ag		0.0591	29But
Pt, H_2 | HCl, 1.0 **m**, AgCl, ethanol | Ag		0.0405	29But
Pt, H_2 | C_2H_5COOH, 8.97 m**m**, C_2H_5COONa, 10.21 m**m**, NaCl, 9.84 m**m**,	0	0.6138	42Pat
AgCl, methanol(10wt%)/water | Ag	20	0.6317	42Pat
	40	0.6492	42Pat
Pt, H_2 | C_2H_5COOH, 44.86 m**m**, C_2H_5COONa, 51.06 m**m**, NaCl, 49.22 m**m**,	0	0.5779	42Pat
AgCl, methanol(10wt%)/water | Ag	20	0.5921	42Pat
	40	0.6062	42Pat
Pt, H_2 | C_2H_5COOH, 78.73 m**m**, C_2H_5COONa, 89.61 m**m**, NaCl, 86.39 m**m**,	0	0.5674	42Pat
AgCl, methanol(10wt%)/water | Ag	20	0.5780	42Pat
	40	0.5914	42Pat
Pt, H_2 | C_2H_5COOH, 6.07 m**m**, C_2H_5COONa, 7.01 m**m**, NaCl, 6.076 m**m**,	0	0.5720	42Pat
AgCl, methanol(20wt%)/water | Ag	20	0.6469	42Pat
	40	0.6662	42Pat
Pt, H_2 | C_2H_5COOH, 37.36 m**m**, C_2H_5COONa, 43.15 m**m**, NaCl, 37.39 m**m**,	0	0.5862	42Pat
AgCl, methanol(20wt%)/water | Ag	20	0.6015	42Pat
	40	0.6175	42Pat
Pt, H_2 | C_2H_5COOH, 67.29 m**m**, C_2H_5COONa, 77.71 m**m**, NaCl, 67.35 m**m**,	0	0.5720	42Pat
AgCl, methanol(20wt%)/water | Ag	20	0.5874	42Pat
	40	0.6019	42Pat
Ir, H_2 | HCl, 0.0003045 **m**, AgCl, ethanol | Ag		0.3466	28Woo
Ir, H_2 | HCl, 0.0005158 **m**, AgCl, ethanol | Ag		0.3216	28Woo
Ir, H_2 | HCl, 0.001198 **m**, AgCl, ethanol | Ag		0.2827	28Woo
Ir, H_2 | HCl, 0.002193 **m**, AgCl, ethanol | Ag		0.2550	28Woo
Ir, H_2 | HCl, 0.005442 **m**, AgCl, ethanol | Ag		0.2185	28Woo
Ir, H_2 | HCl, 0.01114 **m**, AgCl, ethanol | Ag		0.1905	28Woo

Cell composition	T [°C]	U_0 [V]	Ref.
Ir, H_2 | HCl, 0.02537 **m**, AgCl, ethanol | Ag		0.1601	28Woo
Ir, H_2 | HCl, 0.05188 **m**, AgCl, ethanol | Ag		0.1356	28Woo
Ir, H_2 | HCl, 0.1147 **m**, AgCl, ethanol | Ag		0.1090	28Woo
Ir, H_2 | HCl, 0.2189 **m**, AgCl, ethanol | Ag		0.0880	28Woo
Ir, H_2 | HCl, 0.4732 **m**, AgCl, ethanol | Ag		0.0636	28Woo
Ir, H_2 | HCl, 1.0 **m**, AgCl, ethanol | Ag		0.0405	28Woo
Ir, H_2 | HCl, 1.343 **m**, AgCl, ethanol | Ag		0.0292	28Woo
Pt, H_2 | HCl, 0.084 m**m**, AgCl, ethanol | Ag		0.5082	54Muk
Pt, H_2 | HCl, 0.149 m**m**, AgCl, ethanol | Ag		0.4795	54Muk
Pt, H_2 | HCl, 0.771 m**m**, AgCl, ethanol | Ag		0.3962	54Muk
Pt, H_2 | HCl, 1.030 m**m**, AgCl, ethanol | Ag		0.3817	54Muk
Pt, H_2 | HCl, 1.290 m**m**, AgCl, ethanol | Ag		0.3710	54Muk
Pt, H_2 | HCl, 1.780 m**m**, AgCl, ethanol | Ag		0.3570	54Muk
Pt, H_2 | HCl, 3.530 m**m**, AgCl, ethanol | Ag		0.3250	54Muk
Pt, H_2 | HCl, 6.154 m**m**, AgCl, ethanol | Ag		0.2992	54Muk
Pt, H_2 | HCl, 9.070 m**m**, AgCl, ethanol | Ag		0.2794	54Muk
Pt, H_2 | HCl, 11.834 m**m**, AgCl, ethanol | Ag		0.2704	54Muk
Pt, H_2 | HCl, 15.130 m**m**, AgCl, ethanol | Ag		0.2659	54Muk
Pt, H_2 | HCl, 21.790 m**m**, AgCl, ethanol | Ag		0.2440	54Muk
Pt, H_2 | HCl, 37.241 m**m**, AgCl, ethanol | Ag		0.2216	54Muk
Pt, H_2 | HCl, 65.090 m**m**, AgCl, ethanol | Ag		0.1956	54Muk
Pt, H_2 | HCl, 130.180 m**m**, AgCl, ethanol | Ag		0.1632	54Muk
Pt, H_2 | HCl, 0.00886 **m**, Hg_2Cl_2, ethanol | Hg		0.2406	22Dan
Pt, H_2 | HCl, 0.019 **m**, Hg_2Cl_2, ethanol | Hg		0.2232	22Dan
Pt, H_2 | HCl, 0.0274 **m**, Hg_2Cl_2, ethanol | Hg		0.1957	22Dan
Pt, H_2 | HCl, 0.0877 **m**, Hg_2Cl_2, ethanol | Hg		0.1538	22Dan
Pt, H_2 | HCl, 0.1703 **m**, Hg_2Cl_2, ethanol | Hg		0.1326	22Dan
Pt, H_2 | HCl, 0.3020 **m**, Hg_2Cl_2, ethanol | Hg		0.1175	22Dan
Pt, H_2 | HCl, 0.01 **M** ‖ HCl, 0.01 **M**, Hg_2Cl_2,sol., ethanol(16.18vol%)/water | Hg		0.4992	27Erd
Pt, H_2 | HCl, 0.01 **M** ‖ HCl, 0.01 **M**, Hg_2Cl_2,sol., ethanol(32.35vol%)/water | Hg		0.4858	27Erd
Pt, H_2 | HCl, 0.01 **M** ‖ HCl, 0.01 **M**, Hg_2Cl_2,sol., ethanol(52.35vol%)/water | Hg		0.4697	27Erd
Pt, H_2 | HCl, 0.01 **M** ‖ HCl, 0.01 **M**, Hg_2Cl_2,sol., ethanol(67.64vol%)/water | Hg		0.4402	27Erd
Pt, H_2 | HCl, 0.01 **M** ‖ HCl, 0.01 **M**, Hg_2Cl_2,sol., ethanol | Hg		0.322	27Erd
Pt, H_2 | HCl, 0.1 **M** ‖ HCl, 0.01 **M**, Hg_2Cl_2,sol., ethanol(16.18vol%)/water | Hg		0.3832	27Erd
Pt, H_2 | HCl, 0.1 **M** ‖ HCl, 0.01 **M**, Hg_2Cl_2,sol., ethanol(32.35vol%)/water | Hg		0.3742	27Erd
Pt, H_2 | HCl, 0.1 **M** ‖ HCl, 0.01 **M**, Hg_2Cl_2,sol., ethanol(52.35vol%)/water | Hg		0.3545	27Erd
Pt, H_2 | HCl, 0.1 **M** ‖ HCl, 0.01 **M**, Hg_2Cl_2,sol., ethanol(67.64vol%)/water | Hg		0.3329	27Erd
Pt, H_2 | HCl, 0.1 **M** ‖ HCl, 0.01 **M**, Hg_2Cl_2,sol., ethanol | Hg		0.227	27Erd
Pt, H_2 | H_2SO_4, 0.05 **m** ‖ Hg_2SO_4,sat., ethanol(10wt%)/water | Hg		0.7470	46Cro
Pt, H_2 | H_2SO_4, 0.10 **m** ‖ Hg_2SO_4,sat., ethanol(10wt%)/water | Hg		0.7303	46Cro
Pt, H_2 | H_2SO_4, 0.50 **m** ‖ Hg_2SO_4,sat., ethanol(10wt%)/water | Hg		0.6899	46Cro
Pt, H_2 | H_2SO_4, 1.0 **m** ‖ Hg_2SO_4,sat., ethanol(10wt%)/water | Hg		0.6700	46Cro
Pt, H_2 | H_2SO_4, 2.0 **m** ‖ Hg_2SO_4,sat., ethanol(10wt%)/water | Hg		0.6450	46Cro
Pt, H_2 | H_2SO_4, 4.0 **m** ‖ Hg_2SO_4,sat., ethanol(10wt%)/water | Hg		0.6028	46Cro
Pt, H_2 | H_2SO_4, 6.0 **m** ‖ Hg_2SO_4,sat., ethanol(10wt%)/water | Hg		0.5757	46Cro
Pt, H_2 | H_2SO_4, 8.0 **m** ‖ Hg_2SO_4,sat., ethanol(10wt%)/water | Hg		0.5491	46Cro
Pt, H_2 | H_2SO_4, 0.05 **m** ‖ Hg_2SO_4,sat., ethanol(20wt%)/water | Hg		0.7419	46Cro

Cell composition	T [°C]	U_0 [V]	Ref.
Pt, H_2 | H_2SO_4, 0.10 **m** ‖ Hg_2SO_4,sat., ethanol(20wt%)/water | Hg		0.7254	46Cro
Pt, H_2 | H_2SO_4, 0.50 **m** ‖ Hg_2SO_4,sat., ethanol(20wt%)/water | Hg		0.6853	46Cro
Pt, H_2 | H_2SO_4, 1.0 **m** ‖ Hg_2SO_4,sat., ethanol(20wt%)/water | Hg		0.6658	46Cro
Pt, H_2 | H_2SO_4, 2.0 **m** ‖ Hg_2SO_4,sat., ethanol(20wt%)/water | Hg		0.6409	46Cro
Pt, H_2 | H_2SO_4, 4.0 **m** ‖ Hg_2SO_4,sat., ethanol(20wt%)/water | Hg		0.6042	46Cro
Pt, H_2 | H_2SO_4, 6.0 **m** ‖ Hg_2SO_4,sat., ethanol(20wt%)/water | Hg		0.5737	46Cro
Pt, H_2 | H_2SO_4, 8.0 **m** ‖ Hg_2SO_4,sat., ethanol(20wt%)/water | Hg		0.5471	46Cro
Pt, H_2 | H_2SO_4, 0.05 **m** ‖ Hg_2SO_4,sat., ethanol(35wt%)/water | Hg		0.7359	46Cro
Pt, H_2 | H_2SO_4, 0.10 **m** ‖ Hg_2SO_4,sat., ethanol(35wt%)/water | Hg		0.7159	46Cro
Pt, H_2 | H_2SO_4, 0.50 **m** ‖ Hg_2SO_4,sat., ethanol(35wt%)/water | Hg		0.6801	46Cro
Pt, H_2 | H_2SO_4, 1.0 **m** ‖ Hg_2SO_4,sat., ethanol(35wt%)/water | Hg		0.6602	46Cro
Pt, H_2 | H_2SO_4, 2.0 **m** ‖ Hg_2SO_4,sat., ethanol(35wt%)/water | Hg		0.6355	46Cro
Pt, H_2 | H_2SO_4, 4.0 **m** ‖ Hg_2SO_4,sat., ethanol(35wt%)/water | Hg		0.5977	46Cro
Pt, H_2 | H_2SO_4, 6.0 **m** ‖ Hg_2SO_4,sat., ethanol(35wt%)/water | Hg		0.5663	46Cro
Pt, H_2 | H_2SO_4, 8.0 **m** ‖ Hg_2SO_4,sat., ethanol(35wt%)/water | Hg		0.5408	46Cro
Pt, H_2 | H_2SO_4, 0.05 **m** ‖ Hg_2SO_4,sat., ethanol(50wt%)/water | Hg		0.7297	46Cro
Pt, H_2 | H_2SO_4, 0.10 **m** ‖ Hg_2SO_4,sat., ethanol(50wt%)/water | Hg		0.7135	46Cro
Pt, H_2 | H_2SO_4, 0.50 **m** ‖ Hg_2SO_4,sat., ethanol(50wt%)/water | Hg		0.6745	46Cro
Pt, H_2 | H_2SO_4, 1.0 **m** ‖ Hg_2SO_4,sat., ethanol(50wt%)/water | Hg		0.6548	46Cro
Pt, H_2 | H_2SO_4, 2.0 **m** ‖ Hg_2SO_4,sat., ethanol(50wt%)/water | Hg		0.6297	46Cro
Pt, H_2 | H_2SO_4, 4.0 **m** ‖ Hg_2SO_4,sat., ethanol(50wt%)/water | Hg		0.5915	46Cro
Pt, H_2 | H_2SO_4, 6.0 **m** ‖ Hg_2SO_4,sat., ethanol(50wt%)/water | Hg		0.5587	46Cro
Pt, H_2 | H_2SO_4, 8.0 **m** ‖ Hg_2SO_4,sat., ethanol(50wt%)/water | Hg		0.5314	46Cro
Pt, H_2 | H_2SO_4, 0.05 **m** ‖ Hg_2SO_4,sat., ethanol(65wt%)/water | Hg		0.7194	46Cro
Pt, H_2 | H_2SO_4, 0.10 **m** ‖ Hg_2SO_4,sat., ethanol(65wt%)/water | Hg		0.7038	46Cro
Pt, H_2 | H_2SO_4, 0.50 **m** ‖ Hg_2SO_4,sat., ethanol(65wt%)/water | Hg		0.6659	46Cro
Pt, H_2 | H_2SO_4, 1.0 **m** ‖ Hg_2SO_4,sat., ethanol(65wt%)/water | Hg		0.6460	46Cro
Pt, H_2 | H_2SO_4, 2.0 **m** ‖ Hg_2SO_4,sat., ethanol(65wt%)/water | Hg		0.6209	46Cro
Pt, H_2 | H_2SO_4, 4.0 **m** ‖ Hg_2SO_4,sat., ethanol(65wt%)/water | Hg		0.5834	46Cro
Pt, H_2 | H_2SO_4, 6.0 **m** ‖ Hg_2SO_4,sat., ethanol(65wt%)/water | Hg		0.5533	46Cro
Pt, H_2 | H_2SO_4, 0.05 **m** ‖ Hg_2SO_4,sat., ethanol(80wt%)/water | Hg		0.7033	46Cro
Pt, H_2 | H_2SO_4, 0.10 **m** ‖ Hg_2SO_4,sat., ethanol(80wt%)/water | Hg		0.6881	46Cro
Pt, H_2 | H_2SO_4, 0.50 **m** ‖ Hg_2SO_4,sat., ethanol(80wt%)/water | Hg		0.6516	46Cro
Pt, H_2 | H_2SO_4, 1.0 **m** ‖ Hg_2SO_4,sat., ethanol(80wt%)/water | Hg		0.6355	46Cro
Pt, H_2 | H_2SO_4, 2.0 **m** ‖ Hg_2SO_4,sat., ethanol(80wt%)/water | Hg		0.6148	46Cro
Pt, H_2 | H_2SO_4, 4.0 **m** ‖ Hg_2SO_4,sat., ethanol(80wt%)/water | Hg		0.5818	46Cro
Pt, H_2 | H_2SO_4, 6.0 **m** ‖ Hg_2SO_4,sat., ethanol(80wt%)/water | Hg		0.5485	46Cro
Pt, H_2 | H_2SO_4, 8.0 **m** ‖ Hg_2SO_4,sat., ethanol(80wt%)/water | Hg		0.5209	46Cro
Pt, H_2 | HCl, AgCl, ethane-1,2-diol(5wt%)/water | Ag		0.2191	50Cro2, 46Kni
Pt, H_2 | HCl, AgCl, ethane-1,2-diol(10wt%)/water | Ag		0.2164	50Cro2, 46Kni
Pt, H_2 | HCl, AgCl, ethane-1,2-diol(15wt%)/water | Ag		0.2133	50Cro2, 46Kni
Pt, H_2 | HCl, AgCl, ethane-1,2-diol(20wt%)/water | Ag		0.2102	50Cro2, 46Kni
Pt, H_2 | HCl, AgCl, ethane-1,2-diol(30wt%)/water | Ag		0.2036	50Cro2, 46Kni

Cell composition	T [°C]	U_0 [V]	Ref.
Pt, H_2 | HCl, AgCl, ethane-1,2-diol(40wt%)/water | Ag		0.1972	50Cro2, 46Kni
Pt, H_2 | HCl, AgCl, ethane-1,2-diol(60wt%)/water | Ag		0.1807	50Cro2, 46Kni
Pt, H_2 | H_2SO_4, 0.005 **M**, Hg_2SO_4,sol., ethane-1,2-diol, 10 **M**, water | Hg		0.7834	33Tri
Pt, H_2 | H_2SO_4, 0.025 **M**, Hg_2SO_4,sol., ethane-1,2-diol, 10 **M**, water | Hg		0.7453	33Tri
Pt, H_2 | H_2SO_4, 0.05 **M**, Hg_2SO_4,sol., ethane-1,2-diol, 10 **M**, water | Hg		0.7290	33Tri
Pt, H_2 | H_2SO_4, 0.25 **M**, Hg_2SO_4,sol., ethane-1,2-diol, 10 **M**, water | Hg		0.6904	33Tri
Pt, H_2 | H_2SO_4, 0.5 **M**, Hg_2SO_4,sol., ethane-1,2-diol, 10 **M**, water | Hg		0.6730	33Tri
Pt, H_2 | H_2SO_4, 1.0 **M**, Hg_2SO_4,sol., ethane-1,2-diol, 10 **M**, water | Hg		0.6505	33Tri
Pt, H_2 | HCl, 0.001862 **m**, AgCl,2-propanol(10wt%)/water | Ag		0.53930	39Har3
Pt, H_2 | HCl, 0.0208 **m**, AgCl,2-propanol(10wt%)/water | Ag		0.42025	39Har3
Pt, H_2 | HCl, 0.11188 **m**, AgCl,2-propanol(10wt%)/water | Ag		0.33984	39Har3
Pt, H_2 | HCl, 0.4451 **m**, AgCl,2-propanol(10wt%)/water | Ag		0.27180	39Har3
Pt, H_2 | HCl, 0.8863 **m**, AgCl,2-propanol(10wt%)/water | Ag		0.23417	39Har3
Pt, H_2 | HCl, 1.0 **m**, AgCl,2-propanol(10wt%)/water | Ag		0.22704	39Har3
Pt, H_2 | HCl, AgCl, 2-propanol(5wt%)/water | Ag	15	0.2239	47Moo
(U_T [V]=0.22110–5.7425·10^{-4} (T–20 °C)–3.8357·10^{-6} (T–20 °C)^2, 0<T<40 °C)	20	0.2211	47Moo
	25	0.2181	47Moo
Pt, H_2 | HCl, AgCl, 2-propanol(10wt%)/water | Ag	15	0.2192	47Moo
(U_T [V]=0.21666–5.3324·10^{-4}(T–20 °C)–4.7405·10^{-6}(T–20 °C)^2, 0<T<40 °C)	20	0.2167	47Moo
	25	0.2138	47Moo
Pt, H_2 | HCl, AgCl, 2-propanol(20wt%)/water | Ag	15	0.2114	47Moo
(U_T [V]=0.20905–4.9001·10^{-4} (T–20 °C)–6.8362·10^{-6} (T–20 °C)^2, 0<T<40 °C)	20	0.2091	47Moo
	25	0.2064	47Moo
Pt, H_2 | H_2SO_4, 0.05 **m**, Hg_2SO_4,sol., 2-propanol(5wt%)/water | Hg		0.7492	50Cro3
Pt, H_2 | H_2SO_4, 0.10 **m**, Hg_2SO_4,sol., 2-propanol(5wt%)/water | Hg		0.7327	50Cro3
Pt, H_2 | H_2SO_4, 0.50 **m**, Hg_2SO_4,sol., 2-propanol(5wt%)/water | Hg		0.6922	50Cro3
Pt, H_2 | H_2SO_4, 1.0 **m**, Hg_2SO_4,sol., 2-propanol(5wt%)/water | Hg		0.6720	50Cro3
Pt, H_2 | H_2SO_4, 0.05 **m**, Hg_2SO_4,sol., 2-propanol(10wt%)/water | Hg		0.7470	50Cro3
Pt, H_2 | H_2SO_4, 0.10 **m**, Hg_2SO_4,sol., 2-propanol(10wt%)/water | Hg		0.7304	50Cro3
Pt, H_2 | H_2SO_4, 0.50 **m**, Hg_2SO_4,sol., 2-propanol(10wt%)/water | Hg		0.6908	50Cro3
Pt, H_2 | H_2SO_4, 1.0 **m**, Hg_2SO_4,sol., 2-propanol(10wt%)/water | Hg		0.6709	50Cro3
Pt, H_2 | H_2SO_4, 0.05 **m**, Hg_2SO_4,sol., 2-propanol(20wt%)/water | Hg		0.7414	50Cro3
Pt, H_2 | H_2SO_4, 0.10 **m**, Hg_2SO_4,sol., 2-propanol(20wt%)/water | Hg		0.7249	50Cro3
Pt, H_2 | H_2SO_4, 0.50 **m**, Hg_2SO_4,sol., 2-propanol(20wt%)/water | Hg		0.6862	50Cro3
Pt, H_2 | H_2SO_4, 1.0 **m**, Hg_2SO_4,sol., 2-propanol(20wt%)/water | Hg		0.6661	50Cro3
Pt, H_2 | HCl, AgCl, glycerol(50wt%)/water | Ag		0.18932	46Har
(U_T=0.18392–7.45·10^{-4} (T–25 °C)–3·10^{-6} (T–25 °C)^2 [V], 0 < T < 40 °C)			
Pt, H_2 | HCl, 0.001 **m**, AgCl, glycerol(1mol%)/water | Ag		0.5777	26Luc
Pt, H_2 | HCl, 0.002 **m**, AgCl, glycerol(1mol%)/water | Ag		0.5415	26Luc
Pt, H_2 | HCl, 0.005 **m**, AgCl, glycerol(1mol%)/water | Ag		0.4599	26Luc
Pt, H_2 | HCl, 0.01 **m**, AgCl, glycerol(1mol%)/water | Ag		0.4615	26Luc
Pt, H_2 | HCl, 0.02 **m**, AgCl, glycerol(1mol%)/water | Ag		0.4276	26Luc
Pt, H_2 | HCl, 0.05 **m**, AgCl, glycerol(1mol%)/water | Ag		0.3834	26Luc
Pt, H_2 | HCl, 0.1 **m**, AgCl, glycerol(1mol%)/water | Ag		0.3495	26Luc
Pt, H_2 | HCl, 0.2 **m**, AgCl, glycerol(1mol%)/water | Ag		0.3161	26Luc
Pt, H_2 | HCl, 0.5 **m**, AgCl, glycerol(1mol%)/water | Ag		0.2696	26Luc

Cell composition	T [°C]	U_0 [V]	Ref.
Pt, H_2 | HCl, 1.0 **m**, AgCl, glycerol(1mol%)/water | Ag		0.2304	26Luc
Pt, H_2 | HCl, 2.0 **m**, AgCl, glycerol(1mol%)/water | Ag		0.1830	26Luc
Pt, H_2 | HCl, 4.0 **m**, AgCl, glycerol(1mol%)/water | Ag		0.1184	26Luc
Pt, H_2 | HCl, 0.001 **m**, AgCl, glycerol(5mol%)/water | Ag		0.5681	26Luc
Pt, H_2 | HCl, 0.002 **m**, AgCl, glycerol(5mol%)/water | Ag		0.5323	26Luc
Pt, H_2 | HCl, 0.005 **m**, AgCl, glycerol(5mol%)/water | Ag		0.4860	26Luc
Pt, H_2 | HCl, 0.01 **m**, AgCl, glycerol(5mol%)/water | Ag		0.4512	26Luc
Pt, H_2 | HCl, 0.02 **m**, AgCl, glycerol(5mol%)/water | Ag		0.4172	26Luc
Pt, H_2 | HCl, 0.05 **m**, AgCl, glycerol(5mol%)/water | Ag		0.3730	26Luc
Pt, H_2 | HCl, 0.1 **m**, AgCl, glycerol(5mol%)/water | Ag		0.3397	26Luc
Pt, H_2 | HCl, 0.2 **m**, AgCl, glycerol(5mol%)/water | Ag		0.3061	26Luc
Pt, H_2 | HCl, 0.5 **m**, AgCl, glycerol(5mol%)/water | Ag		0.2596	26Luc
Pt, H_2 | HCl, 1.0 **m**, AgCl, glycerol(5mol%)/water | Ag		0.2197	26Luc
Pt, H_2 | HCl, 2.0 **m**, AgCl, glycerol(5mol%)/water | Ag		0.1711	26Luc
Pt, H_2 | HCl, 4.0 **m**, AgCl, glycerol(5mol%)/water | Ag		0.1037	26Luc
Pt, H_2 | HCl, 0.01 **M**, Hg_2Cl_2, glycerol(23.50vol%)/water | Ag		0.4994	27Erd
Pt, H_2 | HCl, 0.1 **M**, Hg_2Cl_2, glycerol(23.50vol%)/water | Ag		0.3868	27Erd
Pt, H_2 | HCl, 0.01 **M**, Hg_2Cl_2, glycerol(47.96vol%)/water | Ag		0.4848	27Erd
Pt, H_2 | HCl, 0.1 **M**, Hg_2Cl_2, glycerol(47.96vol%)/water | Ag		0.3738	27Erd
Pt, H_2 | HCl, 0.01 **M**, Hg_2Cl_2, glycerol(71.32vol%)/water | Ag		0.4653	27Erd
Pt, H_2 | HCl, 0.1 **M**, Hg_2Cl_2, glycerol(71.32vol%)/water | Ag		0.3593	27Erd
Pt, H_2 | HCl, 0.01 **M**, Hg_2Cl_2, glycerol(94.60vol%)/water | Ag		0.4416	27Erd
Pt, H_2 | HCl, 0.1 **M**, Hg_2Cl_2, glycerol(94.60vol%)/water | Ag		0.3388	27Erd
Pt, H_2 | HCl, 0.01 **M**, Hg_2Cl_2, glycerol | Ag		0.403	27Erd
Pt, H_2 | HCl, 0.1 **M**, Hg_2Cl_2, glycerol | Ag		0.287	27Erd
Pt, H_2 | HCl, 0.01 **M**, glycerol(23.50vol%)/water ‖ HCl, 0.1 **M**, glycerol(23.50vol%)/water | H_2, Pt		0.0181	27Erd
Pt, H_2 | HCl, 0.01 **M**, glycerol(47.96vol%)/water ‖ HCl, 0.1 **M**, glycerol(47.96vol%)/water | H_2, Pt		0.0166	27Erd
Pt, H_2 | HCl, 0.01 **M**, glycerol(71.32vol%)/water ‖ HCl, 0.1 **M**, glycerol(71.32vol%)/water | H_2, Pt		0.0129	27Erd
Pt, H_2 | HCl, 0.01 **M**, glycerol(94.60vol%)/water ‖ HCl, 0.1 **M**, glycerol(94.60vol%)/water | H_2, Pt		0.0096	27Erd
Pt, H_2 | HCl, 0.01 **M**, glycerol ‖ HCl, 0.1 **M**, glycerol | H_2, Pt		0.025	27Erd
Pt, H_2 | quinhydrone, picric acid, 0.01 **M**, glycerol(23.32vol%)/water ‖ quinhydrone, picric acid, 0.1 **M**, glycerol(23.32vol%)/water | Pt		0.0084	27Erd
Pt, H_2 | quinhydrone, picric acid, 0.01 **M**, glycerol(47.87vol%)/water ‖ quinhydrone, picric acid, 0.1 **M**, glycerol(47.87vol%)/water | Pt		0.0069	27Erd
Pt, H_2 | quinhydrone, picric acid, 0.01 **M**, glycerol(71.18vol%)/water ‖ quinhydrone, picric acid, 0.1 **M**, glycerol(71.18vol%)/water | Pt		0.0062	27Erd
Pt, H_2 | quinhydrone, picric acid, 0.01 **M**, glycerol(95.56vol%)/water ‖ quinhydrone, picric acid, 0.1 **M**, glycerol(95.56vol%)/water | Pt		0.017	27Erd
Pt, H_2 | quinhydrone, picric acid, 0.01 **M**, glycerol ‖ quinhydrone, picric acid, 0.1 **M**, glycerol | Pt		0.148	27Erd
Pt, H_2 | HCl, 0.005 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag (U_T [V]=0.48096–1.87·10^{-4} (T–25 °C)–3.61·10^{-6} (T–25 °C)^2)		0.48096	38Har2, 38Har3, 39Har5, 36Har3

Cell composition	T [°C]	U_0 [V]	Ref.
Pt, H_2 | HCl, 0.01 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag (U_T [V]=0.4722–0.84·10^{-4} (T–25 °C)–3.58·10^{-6} (T–25 °C)^2)		0.44722	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.05 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag (U_T [V]=0.37074+1.54·10^{-4} (T–25 °C)–3.51·10^{-6} (T–25 °C)^2)		0.37074	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.10 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag (U_T [V]=0.33829–2.56·10^{-4} (T–25 °C)–3.46·10^{-6} (T–25 °C)^2)		0.33829	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.5 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag (U_T [V]=0.25991–4.90·10^{-4} (T–25 °C)–3.10·10^{-6} (T–25 °C)^2)		0.25991	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 1.0 **m** ,AgCl, 1,4-dioxane(20wt%)/water | Ag (U_T [V]=0.22099–5.95·10^{-4} (T–25 °C)–2.78·10^{-6} (T–25 °C)^2)		0.22099	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 1.5 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag (U_T [V]=0.19453–6.58·10^{-4} (T–25 °C)–2.52·10^{-6} (T–25 °C)^2)		0.19453	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 2.0 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag (U_T [V]=0.17291–7.02·10^{-4} (T–25 °C)–2.31·10^{-6} (T–25 °C)^2)		0.17291	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 3.0 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag (U_T [V]=0.13689–7.67·10^{-4} (T–25 °C)–1.99·10^{-6} (T–25 °C)^2)		0.13689	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.003 **m**, AgCl, 1,4-dioxane(45wt%)/water | Ag (U_T [V]=0.47082+0.78·10^{-4} (T–25 °C)–3.30·10^{-6} (T–25 °C)^2)		0.47082	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.005 **m**, AgCl, 1,4-dioxane(45wt%)/water | Ag (U_T [V]=0.44670–1.38·10^{-4} (T–25 °C)–3.24·10^{-6} (T–25 °C)^2)		0.44670	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.01 **m**, AgCl, 1,4-dioxane(45wt%)/water | Ag (U_T [V]=0.41470–2.23·10^{-4} (T–25 °C)–3.17·10^{-6} (T–25 °C)^2)		0.41470	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.05 **m**, AgCl, 1,4-dioxane(45wt%)/water | Ag (U_T [V]=0.34316+4.23·10^{-4} (T–25 °C)–2.99·10^{-6} (T–25 °C)^2)		0.34316	38Har2, 38Har3, 39Har5, 36Har3

Cell composition	T [°C]	U_0 [V]	Ref.
Pt, H_2 | HCl, 0.10 **m**, AgCl, 1,4-dioxane(45wt%)/water | Ag (U_T [V]=0.31290–5.08·10^{-4} (T–25 °C)–2.91·10^{-6} (T–25 °C)^2)		0.31290	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.5 **m**, AgCl, 1,4-dioxane(45wt%)/water | Ag (U_T [V]=0.23855–7.03·10^{-4} (T–25 °C)–2.52·10^{-6} (T–25 °C)^2)		0.23855	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 1.0 **m**, AgCl, 1,4-dioxane(45wt%)/water | Ag (U_T [V]=0.19951–7.91·10^{-4} (T–25 °C)–2.12·10^{-6} (T–25 °C)^2)		0.19951	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 1.5 **m**, AgCl, 1,4-dioxane(45wt%)/water | Ag (U_T [V]=0.17162–8.44·10^{-4} (T–25 °C)–1.84·10^{-6} (T–25 °C)^2)		0.17162	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 2.0 **m**, AgCl, 1,4-dioxane(45wt%)/water | Ag (U_T [V]=0.14811–8.82·10^{-4} (T–25 °C)–1.62·10^{-6} (T–25 °C)^2)		0.14811	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 3.0 **m**, AgCl, 1,4-dioxane(45wt%)/water | Ag (U_T [V]=0.10711–9.38·10^{-4} (T–25 °C)–1.29·10^{-6} (T–25 °C)^2)		0.10711	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.001 **m**, AgCl, 1,4-dioxane(70wt%)/water | Ag (U_T [V]=0.43716+4.50·10^{-4} (T–25 °C)–3.14·10^{-6} (T–25 °C)^2)		0.43716	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.005 **m**, AgCl, 1,4-dioxane(70wt%)/water | Ag (U_T [V]=0.37130–6.20·10^{-4} (T–25 °C)–2.71·10^{-6} (T–25 °C)^2)		0.37130	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.01 **m**, AgCl, 1,4-dioxane(70wt%)/water | Ag (U_T [V]=0.34525–6.80·10^{-4} (T–25 °C)–2.60·10^{-6} (T–25 °C)^2)		0.34525	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.05 **m**, AgCl, 1,4-dioxane(70wt%)/water | Ag (U_T [V]=0.28739+8.11·10^{-4} (T–25 °C)–2.33·10^{-6} (T–25 °C)^2)		0.28739	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.10 **m**, AgCl, 1,4-dioxane(70wt%)/water | Ag (U_T [V]=0.26195–8.67·10^{-4} (T–25 °C)–2.22·10^{-6} (T–25 °C)^2)		0.26195	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.5 **m**, AgCl, 1,4-dioxane(70wt%)/water | Ag (U_T [V]=0.19274–9.95·10^{-4} (T–25 °C)–1.80·10^{-6} (T–25 °C)^2)		0.19274	38Har2, 38Har3, 39Har5, 36Har3

Cell composition	T [°C]	U_0 [V]	Ref.
Pt, H_2 | HCl, 1.0 **m**, AgCl, 1,4-dioxane(70wt%)/water | Ag (U_T [V]=0.150000–10.54·10^{-4} (T–25 °C)–1.45·10^{-6} (T–25 °C)^2)		0.15000	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 1.5 **m**, AgCl, 1,4-dioxane(70wt%)/water | Ag (U_T [V]=0.11638–10.90·10^{-4} (T–25 °C)–1.19·10^{-6} (T–25 °C)^2)		0.11638	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.001 **m**, AgCl, 1,4-dioxane(82wt%)/water | Ag (U_T [V]=0.36094+8.03·10^{-4} (T–25 °C)–5.54·10^{-6} (T–25 °C)^2)		0.36094	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.005 **m**, AgCl, 1,4-dioxane(82wt%)/water | Ag (U_T [V]=0.30913–9.43·10^{-4} (T–25 °C)–3.45·10^{-6} (T–25 °C)^2)		0.30913	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.01 **m**, AgCl, 1,4-dioxane(82wt%)/water | Ag (U_T [V]=0.28844–9.85·10^{-4} (T–25 °C)–2.80·10^{-6} (T–25 °C)^2)		0.28844	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.05 **m**, AgCl, 1,4-dioxane(82wt%)/water | Ag (U_T [V]=0.24060+10.73·10^{-4} (T–25 °C)–2.29·10^{-6} (T–25 °C)^2)		0.24060	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.10 **m**, AgCl, 1,4-dioxane(82wt%)/water | Ag (U_T [V]=0.21870–11.14·10^{-4} (T–25 °C)–1.85·10^{-6} (T–25 °C)^2)		0.21870	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.3 **m**, AgCl, 1,4-dioxane(82wt%)/water | Ag (U_T [V]=0.17636–11.70·10^{-4} (T–25 °C)–1.30·10^{-6} (T–25 °C)^2)		0.17636	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.5 **m**, AgCl, 1,4-dioxane(82wt%)/water | Ag (U_T [V]=0.14792–11.71·10^{-4} (T–25 °C)–0.65·10^{-6} (T–25 °C)^2)		0.14792	38Har2, 38Har3, 39Har5, 36Har3
Pt, H_2 | HCl, 0.0056 **m**, AgCl, acetone | Ag		0.0938	54Eve
Pt, H_2 | HCl, 0.0090 **m**, AgCl, acetone | Ag		0.0811	54Eve
Pt, H_2 | HCl, 0.0219 **m**, AgCl, acetone | Ag		0.0483	54Eve
Pt, H_2 | HCl, 0.0290 **m**, AgCl, acetone | Ag		0.0392	54Eve
Pt, H_2 | HCl, 0.055 **m**, AgCl, acetone | Ag		0.0226	54Eve
Pt, H_2 | HCl, 0.079 **m**, AgCl, acetone | Ag		0.0126	54Eve
Pt, H_2 | HCl, 0.093 **m**, AgCl, acetone | Ag		0.0072	54Eve
Pt, H_2 | HCl, 0.128 **m**, AgCl, acetone | Ag		–0.0001	54Eve
Pt, H_2 | HCl, 0.149 **m**, AgCl, acetone | Ag		–0.0049	54Eve
Pt, H_2 | NaOH, 0.02426 **m**, NaCl, 0.02426 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag	0	1.0635	39Har6
	5	1.0644	39Har6
	10	1.0652	39Har6
	15	1.0661	39Har6
	20	1.0670	39Har6

Cell composition	T [°C]	U_0 [V]	Ref.
Pt, H_2 | NaOH, 0.02426 **m**, NaCl, 0.02426 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag	25	1.0678	39Har6
	30	1.0687	39Har6
	35	1.0695	39Har6
	40	1.0704	39Har6
	50	1.0721	39Har6
Pt, H_2 | NaOH, 0.04344 **m**, NaCl, 0.04344 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag	0	1.0636	39Har6
	5	1.0645	39Har6
	10	1.0653	39Har6
	15	1.0662	39Har6
	20	1.0670	39Har6
	25	1.0678	39Har6
	30	1.0687	39Har6
	35	1.0695	39Har6
	40	1.0704	39Har6
	50	1.0720	39Har6
Pt, H_2 | NaOH, 0.07621 **m**, NaCl, 0.07621 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag	0	1.0638	39Har6
	5	1.0646	39Har6
	10	1.0654	39Har6
	15	1.0663	39Har6
	20	1.0671	39Har6
	25	1.0679	39Har6
	30	1.0688	39Har6
	35	1.0696	39Har6
	40	1.0704	39Har6
	50	1.0721	39Har6
Pt, H_2 | NaOH, 0.01 **m**, NaCl, 0.03 **m**, AgCl, 1,4-dioxane(45wt%)/water | Ag	0	1.0660	39Har6
	5	1.0659	39Har6
	10	1.0658	39Har6
	15	1.0656	39Har6
	20	1.0655	39Har6
	25	1.0654	39Har6
	30	1.0652	39Har6
	35	1.0651	39Har6
	40	1.0650	39Har6
	50	1.0647	39Har6
Pt, H_2 | NaOH, 0.01 **m**, NaCl, 0.06 **m**, AgCl, 1,4-dioxane(45wt%)/water | Ag	0	1.0488	39Har6
	5	1.0484	39Har6
	10	1.0479	39Har6
	15	1.0475	39Har6
	20	1.0471	39Har6
	25	1.0467	39Har6
	30	1.0463	39Har6
	35	1.0459	39Har6
	40	1.0455	39Har6
	50	1.0447	39Har6

Cell composition	T [°C]	U_0 [V]	Ref.
Pt, H_2 | NaOH, 0.01 **m**, NaCl, 0.09 **m**, AgCl, 1,4-dioxane(45wt%)/water | Ag	0	1.0385	39Har6
	5	1.0380	39Har6
	10	1.0374	39Har6
	15	1.0369	39Har6
	20	1.0363	39Har6
	25	1.0358	39Har6
	30	1.0352	39Har6
	35	1.0347	39Har6
	40	1.0341	39Har6
	50	1.0330	39Har6
Pt, H_2 | NaOH, 0.01 **m**, NaCl, 0.15 **m**, AgCl, 1,4-dioxane(70wt%)/water | Ag	0	1.0520	39Har6
	5	1.0505	39Har6
	10	1.0489	39Har6
	15	1.0473	39Har6
	20	1.0457	39Har6
	25	1.0441	39Har6
	30	1.0426	39Har6
	35	1.0410	39Har6
	40	1.0394	39Har6
	50	1.0362	39Har6
Pt, H_2 | NaOH, 0.01 **m**, NaCl, 0.30 **m**, AgCl, 1,4-dioxane(70wt%)/water | Ag	0	1.0299	39Har6
	5	1.0280	39Har6
	10	1.0260	39Har6
	15	1.0240	39Har6
	20	1.0220	39Har6
	25	1.0200	39Har6
	30	1.0181	39Har6
	35	1.0161	39Har6
	40	1.0141	39Har6
	50	1.0101	39Har6
Pt, H_2 | CH_3COOH, 0.01 **m**, CH_3COONa, 0.01 **m**, AgCl, 1,4-dioxane(82wt%)/water | Ag	5	0.6903	39Har6
	10	0.6911	39Har6
	15	0.6929	39Har6
	20	0.6945	39Har6
	25	0.6962	39Har6
	30	0.6978	39Har6
	35	0.6994	39Har6
	40	0.7011	39Har6
	45	0.7028	39Har6
Pt, H_2 | CH_3COOH, 0.05 **m**, CH_3COONa, 0.05 **m**, AgCl, 1,4-dioxane(82wt%)/water | Ag	5	0.6501	39Har6
	10	0.6508	39Har6
	15	0.6519	39Har6
	20	0.6532	39Har6
	25	0.6540	39Har6
	30	0.6549	39Har6
	35	0.6559	39Har6
	40	0.6568	39Har6
	45	0.6577	39Har6

Cell composition	T [°C]	U_0 [V]	Ref.
Pt, H_2 | CH_3COOH, 0.08 **m**, CH_3COONa, 0.08 **m**, AgCl,	5	0.6362	39Har6
1,4-dioxane(82wt%)/water | Ag	10	0.6372	39Har6
	15	0.6379	39Har6
	20	0.6384	39Har6
	25	0.6391	39Har6
	30	0.6399	39Har6
	35	0.6406	39Har6
	40	0.6413	39Har6
	45	0.6429	39Har6
Pt, H_2 | H_2SO_4, 0.002461 **m**, Hg_2SO_4, acetic acid | Hg		0.537	31Hut
Pt, H_2 | H_2SO_4, 0.006053 **m**, Hg_2SO_4, acetic acid | Hg		0.525	31Hut
Pt, H_2 | H_2SO_4, 0.007730 **m**, Hg_2SO_4, acetic acid | Hg		0.521	31Hut
Pt, H_2 | H_2SO_4, 0.01377 **m**, Hg_2SO_4, acetic acid | Hg		0.512	31Hut
Pt, H_2 | H_2SO_4, 0.02662 **m**, Hg_2SO_4, acetic acid | Hg		0.504	31Hut
Pt, H_2 | H_2SO_4, 0.03796 **m**, Hg_2SO_4, acetic acid | Hg		0.498	31Hut
Pt, H_2 | H_2SO_4, 0.1437 **m**, Hg_2SO_4, acetic acid | Hg		0.480	31Hut
Pt, H_2 | H_2SO_4, 0.5014 **m**, Hg_2SO_4, acetic acid | Hg		0.460	31Hut
Pt, H_2 | H_2SO_4, 0.8715 **m**, Hg_2SO_4, acetic acid | Hg		0.452	31Hut
Lanthanum (La)			
La | $LaBr_3$,sat., pyridine ‖ $AgNO_3$, 0.1 **M** | Ag		0.71	29Mül1
$La_{5wt\%}Hg$ | $LaBr_3$,sat., pyridine ‖ $AgNO_3$, 0.1 **M** | Ag		0.93	29Mül1
La | $LaBr_3$,sat., pyridine | $La_{5wt\%}Hg$		0.255	29Mül1
Lead (Pb)			
$Pb_{sat.}Hg$ | PbF_2*2.5 H_2O,sol. | HF(NaF), Hg_2F_2,sol. | Hg	0	1.0637	52Koe
Pb | $Pb(NO_3)_2$, 0.05 **M** in NH_3 ‖ $AgNO_3$, 0.1 **M** in NH_3 | Ag	–35	0.467	26Cos
(U_T [V]=0.467–0.5·10^{-3} (T+35 °C), –65<T<–35 °C)			
Pb_zHg | $PbCl_2$, 0.000874 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag		0.5788	44Nob
Pb_zHg | $PbCl_2$, 0.00175 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag		0.5561	44Nob
Pb_zHg | $PbCl_2$, 0.00525 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag		0.5209	44Nob
Pb_zHg | $PbCl_2$, 0.01049 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag		0.5017	44Nob
Pb_zHg | $PbCl_2$, 0.02173 **m**, AgCl, 1,4-dioxane(20wt%)/water | Ag		0.4842	44Nob
Pb_zHg | $PbCl_2$, 0.000951 **m**, AgCl, 1,4-dioxane(40wt%)/water | Ag		0.5462	44Nob
Pb_zHg | $PbCl_2$, 0.00214 **m**, AgCl, 1,4-dioxane(40wt%)/water | Ag		0.5120	44Nob
Pb_zHg | $PbCl_2$, 0.00428 **m**, AgCl, 1,4-dioxane(40wt%)/water | Ag		0.5016	44Nob
Pb_zHg | $PbCl_2$, 0.00932 **m**, AgCl, 1,4-dioxane(40wt%)/water | Ag		0.4842	44Nob
Pb_zHg | $PbCl_2$, 0.000424 **m**, AgCl, 1,4-dioxane(60wt%)/water | Ag		0.5207	44Nob
Pb_zHg | $PbCl_2$, 0.000894 **m**, AgCl, 1,4-dioxane(60wt%)/water | Ag		0.5081	44Nob
Pb_zHg | $PbCl_2$, 0.00153 **m**, AgCl, 1,4-dioxane(60wt%)/water | Ag		0.4929	44Nob
Pb_zHg | $PbCl_2$, 0.00323 **m**, AgCl, 1,4-dioxane(60wt%)/water | Ag		0.4842	44Nob
Pb | $Pb(NO_3)_2$, 0.05 **M**, aniline ‖ $AgNO_3$, 0.1 **M**, aniline | Ag		0.327	99Kah
Lithium (Li)			
$Li_{0.035wt\%}Hg$ | LiCl, 0.01 **m**, AgCl, ethanol | Ag		2.287	32McF
$Li_{0.00581wt\%}Hg$ | LiCl, 0.0064 **m**, AgCl, ethanol | Ag		2.1995	26Dru2
$Li_{0.00581wt\%}Hg$ | LiCl, 0.0122 **m**, AgCl, ethanol | Ag		2.1760	26Dru2
$Li_{0.00581wt\%}Hg$ | LiCl, 0.0244 **m**, AgCl, ethanol | Ag		2.1511	26Dru2
$Li_{0.00581wt\%}Hg$ | LiCl, 0.0487 **m**, AgCl, ethanol | Ag		2.1264	26Dru2
$Li_{0.00581wt\%}Hg$ | LiCl, 0.0974 **m**, AgCl, ethanol | Ag		2.1040	26Dru2
Li | LiI, propylamine | $Li_{0.0350wt\%}Hg$		0.9502	13Lew

Cell composition	T [°C]	U_0 [V]	Ref.
Magnesium (Mg)			
Mg | $MgCl_2$, 0.35 **m**, methanol, Hg_2Cl_2 | Hg	–80	1.8	15Cam
$Mg_{0.12wt\%}Hg$ | $MgCl_2$, 0.35 **m**, methanol, Hg_2Cl_2 | Hg	–80	1.22	15Cam
$Mg_{3.11wt\%}Hg$ | $MgCl_2$, 0.35 **m**, methanol, Hg_2Cl_2 | Hg	–80	1.83	15Cam
$Mg_{9.02wt\%}Hg$ | $MgCl_2$, 0.35 **m**, methanol, Hg_2Cl_2 | Hg	–80	1.95	15Cam
Mg | $Mg(NO_3)_2$, 0.05 **M**, pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		0.929	99Kah
$Mg_{5at\%}Hg$ | MgI_2, 0.038 **M**, pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		1.68	15Cam
$Mg_{10at\%}Hg$ | MgI_2, 0.038 **M**, pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		1.67	15Cam
$Mg_{21at\%}Hg$ | MgI_2, 0.038 **M**, pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		1.70	15Cam
$Mg_{25at\%}Hg$ | MgI_2, 0.038 **M**, pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		1.85	15Cam
$Mg_{38at\%}Hg$ | MgI_2, 0.038 **M**, pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		1.58	15Cam
$Mg_{1at\%}Hg$ | $MgBr_2$, sat.pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		1.65	23Mül
$Mg_{10at\%}Hg$ | $MgBr_2$, sat.pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		1.72	23Mül
$Mg_{30at\%}Hg$ | $MgBr_2$, sat.pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		1.72	23Mül
$Mg_{40at\%}Hg$ | $MgBr_2$, sat.pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		1.73	23Mül
$Mg_{50at\%}Hg$ | $MgBr_2$, sat.pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		1.69	23Mül
$Mg_{52at\%}Hg$ | $MgBr_2$, sat.pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		1.31	23Mül
$Mg_{60at\%}Hg$ | $MgBr_2$, sat.pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		1.33	23Mül
$Mg_{90at\%}Hg$ | $MgBr_2$, sat.pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		1.34	23Mül
$Mg_{99at\%}Hg$ | $MgBr_2$, sat.pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		1.25	23Mül
Mg | $Mg(NO_3)_2$, 0.05 **M**, aniline ‖ $AgNO_3$, 0.1 **M**, aniline | Ag		1.151	99Kah
Mercury (Hg)			
Hg | Hg_2F_2,sol., HF(NaF) | CuF_2,sol. | $Cu_{sat.}Hg$	0	0.274	52Koe
	10	0.283	52Koe
Hg | Hg_2F_2,sol. | NaF, x=0.010 in HF ‖ NaF, HF*H_2O, AgF, x=0.00154 in HF | Ag	0	0.0543	52Koe
Hg | Hg_2F_2,sol. | NaF, x=0.020 in HF ‖ NaF, HF*H_2O, AgF, x=0.00308 in HF | Ag	0	0.0737	52Koe
Hg | Hg_2F_2,sol. | NaF, x=0.0320 in HF ‖ NaF, HF*H_2O, AgF, x=0.0020 in HF | Ag	0	0.0642	52Koe
Hg | Hg_2F_2,sol. | NaF, x=0.034 in HF ‖ NaF, HF*H_2O, AgF, x=0.0040 in HF | Ag	0	0.0763	52Koe
Hg | Hg_2F_2,sol. | NaF, x=0.036 in HF ‖ NaF, HF*H_2O, AgF, x=0.0060 in HF | Ag	0	0.0901	52Koe
Hg | Hg_2F_2,sol. | NaF, x=0.040 in HF ‖ NaF, HF*H_2O, AgF, x=0.00615 in HF | Ag	0	0.0915	52Koe
Hg | Hg_2F_2,sol. | NaF, x=0.050 in HF ‖ NaF, HF*H_2O, AgF, x=0.00769 in HF | Ag	0	0.0964	52Koe
Hg | Hg_2SO_4, H_2SO_4, methanol | H_2, Pt	20	0.544431	46Kan
	25	0.53920	46Kan
	30	0.53510	46Kan
	35	0.53177	46Kan
Hg | $HgNO_3$, 0.00108 **m**, methanol ‖ $HgNO_3$, 0.00108 **m**, methanol | Hg		0.006	09Neu
Hg | $HgNO_3$, 0.0043 **m**, methanol ‖ $HgNO_3$, 0.0173 **m**, methanol | Hg		0.003	09Neu
Hg | $HgNO_3$, 0.0173 **m**, methanol ‖ $HgNO_3$, 0.069 **m**, methanol | Hg		0.002	09Neu
Hg | $HgNO_3$, 0.0345 **m**, methanol ‖ $HgNO_3$, 0.069 **m**, methanol | Hg		0.001	09Neu
Hg | Hg_2Cl_2, HCl, 0.1 **M**, glycerin(23.50vol%)/water ‖ Hg_2Cl_2, HCl, 0.01 **M**, glycerin(23.50vol%)/water | H_2, Pt		0.0946	27Erd

Cell composition	T [°C]	U_0 [V]	Ref.
Hg | Hg_2Cl_2, HCl, 0.1 **M**, glycerol(47.96vol%)/water ‖ Hg_2Cl_2, HCl, 0.01 **M**, glycerol(47.96vol%)/water | H_2, Pt		0.0944	27Erd
Hg | Hg_2Cl_2, HCl, 0.1 **M**, glycerol(71.32vol%)/water ‖ Hg_2Cl_2, HCl, 0.01 **M**, glycerol(71.32vol%)/water | H_2, Pt		0.0926	27Erd
Hg | Hg_2Cl_2, HCl, 0.1 **M**, glycerol(94.60vol%)/water ‖ Hg_2Cl_2, HCl, 0.01 **M**, glycerol(94.60vol%)/water | H_2,Pt		0.0924	27Erd
Hg | Hg_2Cl_2, HCl, 0.1 **M**, glycerol ‖ Hg_2Cl_2, HCl, 0.01 **M**, glycerol | H_2, Pt		0.087	27Erd
Hg | Hg_2picrate,sol., picric acid, 0.1 **M**, glycerol(23.32vol%)/water ‖ picric acid, 0.01 **M**, Hg_2picrate,sol., glycerol(23.32vol%)/water | Hg		0.1017	27Erd
Hg | Hg_2picrate,sol., picric acid, 0.1 **M**, glycerol(47.87vol%)/water ‖ picric acid, 0.1 **M**, Hg_2picrate,sol., glycerol(47.87vol%)/water | Hg		0.1012	27Erd
Hg | Hg_2picrate,sol., picric acid, 0.1 **M**, glycerol(71.18vol%)/water ‖ picric acid, 0.1 **M**, Hg_2picrate,sol., glycerol(71.18vol%)/water | Hg		0.1008	27Erd
Hg | Hg_2picrate,sol., picric acid, 0.1 **M**, glycerol(95.56vol%)/water ‖ picric acid, 0.1 **M**, Hg_2picrate,sol., glycerol(95.56vol%)/water | Hg		0.1084	27Erd
Hg | Hg_2picrate,sol., picric acid, 0.1 **M**, glycerol ‖ picric acid, 0.1 **M**, Hg_2picrate,sol., glycerol | Hg		0.186	27Erd
Hg | Hg_2picrate,sol., picric acid, 0.01 **M**, glycerol(23.32vol%)/water ‖ quinhydrone, picric acid, 0.01 **M**, glycerol(23.32vol%)/water | Pt		0.0854	27Erd
Hg | Hg_2picrate,sol., picric acid, 0.01 **M**, glycerol(47.87vol%)/water ‖ quinhydrone, picric acid, 0.01 **M**, glycerol(47.87vol%)/water | Pt		0.0800	27Erd
Hg | Hg_2picrate,sol., picric acid, 0.01 **M**, glycerol(71.18vol%)/water ‖ quinhydrone, picric acid, 0.01 **M**, glycerol(71.18vol%)/water | Pt		0.0766	27Erd
Hg | Hg_2picrate,sol., picric acid, 0.01 **M**, glycerol(95.56vol%)/water ‖ quinhydrone, picric acid, 0.01 **M**, glycerol(95.56vol%)/water | Pt		0.092	27Erd
Hg | Hg_2picrate,sol., picric acid, 0.01 **M**, glycerol ‖ quinhydrone, picric acid, 0.01 **M**, glycerol | Pt		0.268	27Erd
Hg | Hg_2picrate,sol., picric acid, 0.1 **M**, glycerol(23.32vol%)/water ‖ quinhydrone, picric acid, 0.1 **M**, glycerol(23.32vol%)/water | Pt		0.246	27Erd
Hg | Hg_2picrate,sol., picric acid, 0.1 **M**, glycerol(47.87vol%)/water ‖ quinhydrone, picric acid, 0.1 **M**, glycerol(47.87vol%)/water | Pt		0.282	27Erd
Hg | Hg_2picrate,sol., picric acid, 0.1 **M**, glycerol(71.18vol%)/water ‖ quinhydrone, picric acid, 0.1 **M**, glycerol(71.18vol%)/water | Pt		0.305	27Erd
Hg | Hg_2picrate,sol., picric acid 0.1 **M**, glycerol(95.56vol%)/water ‖ quinhydrone, picric acid, 0.1 **M**, glycerol(95.56vol%)/water | Pt		0.328	27Erd
Hg | Hg_2picrate,sol., picric acid, 0.1 **M**, glycerol ‖ quinhydrone, picric acid, 0.1 **M**, glycerol | Pt		0.059	27Erd
Platinum (Pt)			
Pt | quinhydrone, picric acid, 0.01 **M** ethanol(16.02vol%)/water ‖ quinhydrone, picric acid, 0.1 **M**, ethanol(16.02vol%)/water | Pt		0.0079	27Erd
Pt | quinhydrone, picric acid, 0.01 **M**, ethanol(33.78vol%)/water ‖ quinhydrone, picric acid, 0.1 **M**, ethanol(33.78vol%)/water | Pt		0.0080	27Erd
Pt | quinhydrone, picric acid, 0.01 **M**, ethanol(50.18vol%)/water ‖ quinhydrone, picric acid, 0.1 **M**, ethanol(50.18vol%)/water | Pt		0.0116	27Erd
Pt | quinhydrone, picric acid, 0.01 **M**, ethanol(67.46vol%)/water ‖ quinhydrone, picric acid, 0.1 **M**, ethanol(67.46vol%)/water | Pt		0.0238	27Erd
Pt | quinhydrone, picric acid, 0.01 **M**, ethanol ‖ quinhydrone, picric acid, 0.1 **M**, ethanol | Pt		0.0079	27Erd

Cell composition	T [°C]	U_0 [V]	Ref.
Pt | quinhydrone, picric acid, 0.01 **M**, ethanol(16.02vol%)/water ‖ picric acid, 0.01 **M**, Hg_2picrate, ethanol(16.02vol%)/water | Hg		0.0834	27Erd
Pt | quinhydrone, picric acid, 0.01 **M**, ethanol(33.78vol%)/water ‖ picric acid, 0.01 **M**, Hg_2picrate, ethanol(33.78vol%)/water | Hg		0.1063	27Erd
Pt | quinhydrone, picric acid, 0.01 **M**, ethanol(50.18vol%)/water ‖ picric acid, 0.01 **M**, Hg_2picrate, ethanol(50.18vol%)/water | Hg		0.1427	27Erd
Pt | quinhydrone, picric acid, 0.01 **M**, ethanol(67.46vol%)/water ‖ picric acid, 0.01 **M**, Hg_2picrate, ethanol(67.46vol%)/water | Hg		0.1506	27Erd
Pt | quinhydrone, picric acid, 0.01 **M**, ethanol ‖ picric acid, 0.01 **M**, Hg_2picrate, ethanol | Hg		0.094	27Erd
Pt | quinhydrone, picric acid, 0.1 **M**, ethanol(16.02vol%)/water ‖ picric acid, 0.1 **M**, Hg_2picrate, ethanol(16.02vol%)/water | Hg		–0.00243	27Erd
Pt | quinhydrone, picric acid, 0.1 **M**, ethanol(33.78vol%)/water ‖ picric acid, 0.1 **M**, Hg_2picrate, ethanol(33.78vol%)/water | Hg		0.0012	27Erd
Pt | quinhydrone, picric acid, 0.1 **M**, ethanol(50.18vol%)/water ‖ picric acid, 0.1 **M**, Hg_2picrate, ethanol(50.18vol%)/water | Hg		0.0294	27Erd
Pt | quinhydrone, picric acid, 0.1 **M**, ethanol(67.46vol%)/water ‖ picric acid, 0.1 **M**, Hg_2picrate, ethanol(67.46vol%)/water | Hg		0.0424	27Erd
Pt | quinhydrone, picric acid, 0.1 **M**, ethanol ‖ picric acid, 0.1 **M**, Hg_2picrate, ethanol | Hg		0.005	27Erd
Potassium (K)			
$K_{0.2756wt\%}Hg$ | KCl, 1.492 **m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	–36	0.3834	51Sed
$K_{0.2756wt\%}Hg$ | KCl, 1.923 **m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	–36	0.3750	51Sed
$K_{0.2756wt\%}Hg$ | KCl, 3.204 **m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	–36	0.3575	51Sed
$K_{0.2756wt\%}Hg$ | KCl, 4.696 **m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	–36	0.3457	51Sed
$K_{0.2756wt\%}Hg$ | KCl, 13.98 **m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	–36	0.3147	51Sed
$K_{0.2756wt\%}Hg$ | KCl, 14.26 **m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	–36	0.3137	51Sed
$K_{0.2756wt\%}Hg$ | KCl, 21.51 **m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	–36	0.3042	51Sed
$K_{0.2756wt\%}Hg$ | KCl, 22.66 **m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	–36	0.3027	51Sed
$K_{0.2216wt.\%}Hg$ | KI,sol., KI,sat., AgI, ethanol | Ag		1.6271	34Ish3
$K_{0.2216wt.\%}Hg$ | KI,sol., KI,sat., CuI, ethanol | Cu		1.5904	34Ish3
$K_{0.2216wt\%}Hg$ | KI,sol., KI,sat., CuI, ethanol | Cu_zHg		1.5993	34Ish3
$K_{0.2216wt\%}Hg$ | KBr,sol., KBr,sat., AgBr,sol., ethanol(85wt%)/water | Ag (from Ag_2O)		1.8793	34Ish2
K | KI, ethyl amine | $K_{0.2216wt\%}Hg$		1.0481	12Lew
$K_{1.127mol\%}Hg$ | KI, ethyl amine | $K_{0.01950mol\%}Hg$		0.1237	33Ben
$K_{1.127mol\%}Hg$ | KI, ethyl amine | $K_{0.1924mol\%}Hg$		0.06156	33Ben
$K_{1.127mol\%}Hg$ | KI, ethyl amine | $K_{0.6721mol\%}Hg$		0.02099	33Ben
$K_{1.127mol\%}Hg$ | KI, ethyl amine | $K_{1.539mol\%}Hg$		–0.01552	33Ben
$K_{1.127mol\%}Hg$ | KI, ethyl amine | $K_{1.944mol\%}Hg$		–0.06156	33Ben
$K_{1.127mol\%}Hg$ | KI, ethyl amine | $K_{2.193mol\%}Hg$		–0.03672	33Ben
$K_{1.127mol\%}Hg$ | KI, ethyl amine | $K_{2.444mol\%}Hg$		–0.04418	33Ben
$K_{1.127mol\%}Hg$ | KI, ethyl amine | $K_{2.627mol\%}Hg$		–0.05084	33Ben
Rubidium (Rb)			
Rb | RbI, ethyl amine with 7.9mol% ammonia | $Rb_{0.02698\ eq/kg}Hg$		1.0745	15Lew

Cell composition	T [°C]	U_0 [V]	Ref.
Silver (Ag)			
Ag | $AgNO_3$, 0.298 m**M** in NH_3 ‖ $AgNO_3$, 1.87 m**M** in NH_3 | Ag	–38	0.0308	05Cad
Ag | $AgNO_3$, 0.0487 m**M** in NH_3 ‖ $AgNO_3$, 1.82 m**M** in NH_3 | Ag	–38	0.0592	05Cad
Ag | $AgNO_3$, 0.0080 m**M** in NH_3 ‖ $AgNO_3$, 1.82 m**M** in NH_3 | Ag	–38	0.0845	05Cad
Ag | $AgNO_3$, 0.743 m**M** in NH_3 ‖ $AgNO_3$, 4.54 m**M** in NH_3 | Ag	–38	0.0341	05Cad
Ag | AgBr, NaBr, 0.1 **m**, methanol ‖ NaCl, 0.1 **m**, methanol, Hg_2Cl_2 | Hg		0.0206	29Buc
Ag | AgBr, NaBr, 0.00125 **m**, ethanol ‖ NaBr, 0.00625 **m**, ethanol, AgBr | Ag		0.0127	26Dru2
Ag | AgBr, NaBr, 0.00313 **m**, ethanol ‖ NaBr, 0.00156 **m**, ethanol, AgBr | Ag		0.0138	26Dru2
Ag | AgBr, NaBr, 0.00625 **m**, ethanol ‖ NaBr, 0.00313 **m**, ethanol, AgBr | Ag		0.0134	26Dru2
Ag | AgBr, NaBr, 0.025 **m**, ethanol ‖ NaBr, 0.0125 **m**, ethanol, AgBr | Ag		0.0120	26Dru2
Ag | AgBr, NaBr, 0.05 **m**, ethanol ‖ NaBr ,0.025 **m**, ethanol, AgBr | Ag		0.0115	26Dru2
Ag | AgBr, NaBr, 0.1 **m**, ethanol ‖ NaBr, 0.00156 **m**, ethanol, AgBr | Ag		0.0737	26Dru2
Ag | AgBr, NaBr, 0.1 **m**, ethanol ‖ NaBr, 0.00313 **m**, ethanol, AgBr | Ag		0.0599	26Dru2
Ag | AgBr, NaBr, 0.1 **m**, ethanol ‖ NaBr, 0.00625 **m**, ethanol, AgBr | Ag		0.0466	26Dru2
Ag | AgBr, NaBr, 0.1 **m**, ethanol ‖ NaBr, 0.0125 **m**, ethanol, AgBr | Ag		0.0341	26Dru2
Ag | AgBr, NaBr, 0.1 **m**, ethanol ‖ NaBr, 0.025 **m**, ethanol, AgBr | Ag		0.0222	26Dru2
Ag | AgBr, NaBr, 0.1 **m**, ethanol ‖ NaBr, 0.05 **m**, ethanol, AgBr | Ag		0.0104	26Dru2
Ag | AgBr, KBr, 0.001013 **M**, ethanol(85wt%)/water ‖ AgCl, KCl, 0.001013 **M**, ethanol(24.67wt%)/water | Ag	15.9	0.1419	30Afa
Ag | AgBr, KBr, 0.01006 **M**, ethanol(85wt%)/water ‖ AgCl, KCl, 0.01006 **M**, ethanol(24.67wt%)/water | Ag	15.2	0.1447	30Afa
Ag | AgBr, KBr, 0.1001 **M**, ethanol(85wt%)/water ‖ AgCl, KCl, 0.1001 **M**, ethanol(24.67wt%)/water | Ag	15.1	0.1469	30Afa
Ag | AgBr, KBr, 0.001011 **M**, ethanol(85wt%)/water ‖ AgCl, KCl, 0.001011 **M**, ethanol(50.6wt%)/water | Ag	16.6	0.1447	30Afa
Ag | AgBr, KBr, 0.01006 **M**, ethanol(85wt%)/water ‖ AgCl, KCl, 0.01006 **M**, ethanol(50.6wt%)/water | Ag	16.6	0.1371	30Afa
Ag | AgBr, KBr, 0.10012 **M**, ethanol(85wt%)/water ‖ AgCl, KCl, 0.10012 **M**, ethanol(50.6wt%)/water | Ag	16.6	0.1392	30Afa
Ag | AgCl, HCl, 0.01571 **m**, methanol, AgCl ‖ AgCl, HCl, 0.001274 **m**, methanol, AgCl | Ag		0.0844	25Non
Ag | AgCl, HCl, 0.07332 **m**, methanol, AgCl ‖ AgCl, HCl, 0.004939 **m**, methanol, AgCl | Ag		0.0852	25Non
Ag | AgCl, HCl, 0.07333 **m**, methanol, AgCl ‖ AgCl, HCl, 0.008396 **m**, methanol, AgCl | Ag		0.0672	25Non
Ag | AgCl, HCl, 0.1155 **m**, methanol, AgCl ‖ AgCl, HCl, 0.009857 **m**, methanol, AgCl | Ag		0.0751	25Non
Ag | AgCl, HCl, 0.4802 **m**, methanol, AgCl ‖ AgCl, HCl, 0.04261 **m**, methanol, AgCl | Ag		0.0709	25Non
Ag | AgCl, HCl, 0.1 **m**, methanol ‖ AgCl, NaCl, 0.1 **m**, methanol | Ag		0.0206	29Buc
Ag | AgCl, LiCl, 0.0630 **m**, methanol ‖ AgCl, LiCl, 0.0063 **m**, methanol | Ag		0.0316	22Pea
Ag | AgCl, LiCl, 0.1262 **m**, methanol ‖ AgCl, LiCl, 0.0136 **m**, methanol | Ag		0.0291	22Pea
Ag | AgCl, LiCl, 0.3812 **m**, methanol ‖ AgCl, LiCl, 0.0378 **m**, methanol | Ag		0.0275	22Pea
Ag | AgCl, LiCl, 0.6374 **m**, methanol ‖ AgCl, LiCl, 0.0630 **m**, methanol | Ag		0.0285	22Pea
Ag | AgCl, NaCl, 0.0180 **m**, methanol ‖ AgCl, NaCl, 0.00204 **m**, methanol | Ag		0.0451	27Wol
Ag | AgCl, NaCl, 0.0991 **m**, methanol ‖ AgCl, NaCl, 0.00999 **m**, methanol | Ag		0.0441	27Wol

Cell composition	T [°C]	U_0 [V]	Ref.
Ag | AgCl, NaCl, 0.1790 **m**, methanol ‖ AgCl, NaCl, 0.0180 **m**, methanol | Ag		0.0422	27Wol
Ag | AgCl, NaCl, 0.1 **m**, methanol ‖ NaCl, 0.1 **m**, methanol, Hg_2Cl_2 | Hg		0.0496	29Buc
Ag | AgCl, NaCl, 0.01 **m**, methanol ‖ $NaClO_4$, 0.01 **m**, methanol ‖ $AgClO_4$, 0.01 **m** | Ag		0.5243	29Buc
Ag | AgI, KI, 0.054 **m**, methanol ‖ NaCl, 0.1 **m**, Hg_2Cl_2, methanol | Hg		0.3381	29Buc
Ag | AgI,sol., KI, 0.005 **M**, ethanol ‖ KI, 0.001 m**M**, AgI,sol., ethanol | Ag	18...20	0.0335	29Isa
Ag | AgI,sol., KI, 0.02 **M**, ethanol ‖ KI, 0.001 m**M**, AgI,sol., ethanol | Ag	18...20	0.0595	29Isa
Ag | AgI,sol., KI, 0.02 **M**, ethanol ‖ KI, 0.002 m**M**, AgI,sol., ethanol | Ag	18...20	0.0455	29Isa
Ag | AgI,sol., KI, 0.04 **M**, ethanol ‖ KI, 0.001 m**M**, AgI,sol., ethanol | Ag	18...20	0.0705	29Isa
Ag | AgI,sol., KI, 0.04 **M**, ethanol ‖ KI, 0.005 m**M**, AgI,sol., ethanol | Ag	18...20	0.0380	29Isa
Ag | AgI,sol., KI, 0.05 **M**, ethanol ‖ KI, 0.001 m**M**, AgI,sol., ethanol | Ag	18...20	0.0745	29Isa
Ag | AgI,sol., KI, 0.05 **M**, ethanol ‖ KI, 0.002 m**M**, AgI,sol., ethanol | Ag	18...20	0.0600	29Isa
Ag | AgI,sol., KI, 0.05 **M**, ethanol ‖ KI, 0.005 m**M**, AgI,sol., ethanol | Ag	18...20	0.0415	29Isa
Ag | AgI,sol., KI, 0.05 **M**, ethanol ‖ KI, 0.01 m**M**, AgI,sol., ethanol | Ag	18...20	0.0275	29Isa
Ag | $AgNO_3$, 0.001 **m**, ethanol ‖ $AgNO_3$, 0.01 **m**, ethanol | Ag		0.0597	13Bel
Ag | $AgNO_3$, 0.001 **m**, ethanol ‖ $AgNO_3$, 0.1 **m**, ethanol | Ag		0.1066	13Bel
Ag | $AgNO_3$, 0.01 **m**, ethanol ‖ $AgNO_3$, 0.1 **m**, ethanol | Ag		0.0470	13Bel
Ag | $AgNO_3$, 0.00156 **m**, methanol ‖ $AgNO_3$, 0.00625 **m**, methanol | Ag		0.032	23Gra
Ag | $AgNO_3$, 0.00156 **m**, methanol ‖ $AgNO_3$, 0.025 **m**, methanol | Ag		0.065	23Gra
Ag | $AgNO_3$, 0.00625 **m**, methanol ‖ $AgNO_3$, 0.025 **m**, methanol | Ag		0.034	23Gra
Ag | $AgNO_3$, 0.00156 **m**, methanol ‖ $AgNO_3$, 0.1 **m**, methanol | Ag		0.093	23Gra
Ag | $AgNO_3$, 0.00625 **m**, methanol ‖ $AgNO_3$, 0.1 **m**, methanol | Ag		0.063	23Gra
Ag | $AgNO_3$, 0.025 **m**, methanol ‖ $AgNO_3$, 0.1 **m**, methanol | Ag		0.027	23Gra
Ag | $AgNO_3$, 0.00156 **m**, methanol ‖ $AgNO_3$, 0.025 **m**, methanol | Ag		0.065	34Mül, 06Wil
Ag | $AgNO_3$, 0.00625 **m**, methanol ‖ $AgNO_3$, 0.025 **m**, methanol | Ag		0.031	34Mül, 06Wil
Ag | $AgNO_3$, 0.00156 **m**, methanol ‖ $AgNO_3$, 0.1 **m**, methanol | Ag		0.086	34Mül, 06Wil
Ag | $AgNO_3$, 0.00625 **m**, methanol ‖ $AgNO_3$, 0.1 **m**, methanol | Ag		0.044	34Mül, 06Wil
Ag | $AgNO_3$, 0.025 **m**, methanol ‖ $AgNO_3$, 0.1 **m**, methanol | Ag		0.021	34Mül, 06Wil
Ag | $AgNO_3$, 0.000099 **m**, methanol ‖ $AgNO_3$, 0.00099 **m**, methanol | Ag		0.054	34Mül, 06Wil
Ag | $AgNO_3$, 0.00099 **m**, methanol ‖ $AgNO_3$, 0.0099 **m**, methanol | Ag		0.061	34Mül, 06Wil
Ag | $AgNO_3$, 0.0099 **m**, methanol ‖ $AgNO_3$, 0.099 **m**, methanol | Ag		0.073	34Mül, 06Wil
Ag | $AgNO_3$, 0.099 **m**, methanol ‖ $AgNO_3$, 0.99 **m**, methanol | Ag		0.044	34Mül, 06Wil
Ag | $AgNO_3$, 0.0005 **m**, acetone ‖ $AgNO_3$, 0.007 **m**, acetone | Ag		0.0555	12Ros, 13Ros
Ag | $AgNO_3$, 0.001 **m**, acetone ‖ $AgNO_3$, 0.01 **m**, acetone | Ag		0.0536	12Ros, 13Ros
Ag | $AgNO_3$, 0.001 **m**, acetone ‖ $AgNO_3$, 0.02 **m**, acetone | Ag		0.0621	12Ros, 13Ros

Cell composition	T [°C]	U_0 [V]	Ref.
Ag | $AgNO_3$, 0.002 **m**, acetone ‖ $AgNO_3$, 0.02 **m**, acetone | Ag		0.0490	12Ros, 13Ros
Ag | $AgNO_3$, 0.005 **m**, acetone ‖ $AgNO_3$, 0.02 **m**, acetone | Ag		0.0289	12Ros, 13Ros
Ag | $AgNO_3$, 0.01 **m**, acetone ‖ $AgNO_3$, 0.02 **m**, acetone | Ag		0.0114	12Ros, 13Ros
Ag | AgCl, KCl, 0.02 **m** | K_zHg | KCl, 0.02 **m**, AgCl, methanol(10wt%)/water | Ag		0.0181	30Ake
Ag | AgCl, KCl, 0.02 **m** | K_zHg | KCl, 0.02 **m**, AgCl, methanol(20wt%)/water | Ag		0.0364	30Ake
Ag | AgCl, KCl, 0.02 **m** | K_zHg | KCl, 0.02 **m**, AgCl, methanol(40wt%)/water | Ag		0.0747	30Ake
Ag | AgCl, KCl, 0.02 **m** | K_zHg | KCl, 0.02 **m**, AgCl, methanol(60wt%)/water | Ag		0.1158	30Ake
Ag | AgCl, KCl, 0.02 **m** | K_zHg | KCl, 0.02 **m**, AgCl, methanol(90wt%)/water | Ag		0.1966	30Ake
Ag | AgCl, KCl, 0.05 **m** | K_zHg | KCl, 0.05 **m**, AgCl, methanol(10wt%)/water | Ag		0.0175	30Ake
Ag | AgCl, KCl, 0.05 **m** | K_zHg | KCl, 0.05 **m**, AgCl, methanol(20wt%)/water | Ag		0.0351	30Ake
Ag | AgCl, KCl, 0.05 **m** | K_zHg | KCl, 0.05 **m**, AgCl, methanol(40wt%)/water | Ag		0.0715	30Ake
Ag | AgCl, KCl, 0.05 **m** | K_zHg | KCl, 0.05 **m**, AgCl, methanol(60wt%)/water | Ag		0.1116	30Ake
Ag | AgCl, KCl, 0.05 **m** | K_zHg | KCl, 0.05 **m**, AgCl, methanol(90wt%)/water | Ag		0.1883	30Ake
Ag | AgCl, KCl, 0.1 **m** | K_zHg | KCl, 0.1 **m**, AgCl, methanol(10wt%)/water | Ag		0.0171	30Ake
Ag | AgCl, KCl, 0.1 **m** | K_zHg | KCl, 0.1 **m**, AgCl, methanol(20wt%)/water | Ag		0.0344	30Ake
Ag | AgCl, KCl, 0.1 **m** | K_zHg | KCl, 0.1 **m**, AgCl, methanol(40wt%)/water | Ag		0.0701	30Ake
Ag | AgCl, KCl, 0.1 **m** | K_zHg | KCl, 0.1 **m**, AgCl, methanol(60wt%)/water | Ag		0.1091	30Ake
Ag | AgCl, KCl, 0.1 **m** | K_zHg | KCl, 0.1 **m**, AgCl, methanol(90wt%)/water | Ag		0.1854	30Ake
Ag | AgCl, KCl, 0.5 **m** | K_zHg | KCl, 0.5 **m**, AgCl, methanol(10wt%)/water | Ag		0.0161	30Ake
Ag | AgCl, KCl, 0.5 **m** | K_zHg | KCl, 0.5 **m**, AgCl, methanol(20wt%)/water | Ag		0.0324	30Ake
Ag | AgCl, KCl, 0.5 **m** | K_zHg | KCl, 0.5 **m**, AgCl, methanol(40wt%)/water | Ag		0.0660	30Ake
Ag | AgCl, KCl, 0.5 **m** | K_zHg | KCl, 0.5 **m**, AgCl, methanol(60wt%)/water | Ag		0.1028	30Ake
Ag | AgCl, HCl, 0.02493 **m**, ethanol ‖ HCl, 0.002461 **m**, AgCl, ethanol | Ag		0.0646	28Woo
Ag | AgCl, HCl, 0.06239 **m**, ethanol ‖ HCl, 0.004111 **m**, AgCl, ethanol | Ag		0.0705	28Woo
Ag | AgCl, HCl, 0.08238 **m**, ethanol ‖ HCl, 0.008224 **m**, AgCl, ethanol | Ag		0.0579	28Woo
Ag | AgCl, HCl, 0.1764 **m**, ethanol ‖ HCl, 0.01062 **m**, AgCl, ethanol | Ag		0.0644	28Woo
Ag | AgCl, HCl, 0.4753 **m**, ethanol ‖ HCl, 0.04171 **m**, AgCl, ethanol | Ag		0.0581	28Woo

Cell composition	T [°C]	U_0 [V]	Ref.
Ag | AgCl, HCl, 0.7457 **m**, ethanol ‖ HCl, 0.07319 **m**, AgCl, ethanol | Ag		0.0535	28Woo
Ag | AgCl, HCl, 0.8388 **m**, ethanol ‖ HCl, 0.08238 **m**, AgCl, ethanol | Ag		0.0540	28Woo
Ag | AgCl, HCl, 1.0 **m**, ethanol ‖ HCl, 0.1 **m**, AgCl, ethanol | Ag		0.0542	28Woo
Ag | AgCl, HCl, 1.0 **m**, ethanol ‖ HCl, 0.01035 **m**, AgCl, ethanol | Ag		0.1129	28Woo
Ag | AgCl, HCl, 1.2411 **m**, ethanol ‖ HCl, 0.1364 **m**, AgCl, ethanol | Ag		0.0524	28Woo
Ag | AgCl, HCl, 1.2411 **m**, ethanol ‖ HCl, 0.004111 **m**, AgCl, ethanol | Ag		0.1418	28Woo
Ag | AgCl, LiCl, 0.00295 **m**, ethanol ‖ LiCl, 0.00295 **m**, AgCl, ethanol | Ag		0.0121	26Dru2
Ag | AgCl, LiCl, 0.00590 **m**, ethanol ‖ LiCl, 0.00295 **m**, AgCl, ethanol | Ag		0.0116	26Dru2
Ag | AgCl, LiCl, 0.0120 **m**, ethanol ‖ LiCl, 0.00590 **m**, AgCl, ethanol | Ag		0.0110	26Dru2
Ag | AgCl, LiCl, 0.0240 **m**, ethanol ‖ LiCl, 0.0120 **m**, AgCl, ethanol | Ag		0.0105	26Dru2
Ag | AgCl, LiCl, 0.0480 **m**, ethanol ‖ LiCl, 0.0240 **m**, AgCl, ethanol | Ag		0.0100	26Dru2
Ag | AgCl, LiCl, 0.0965 **m**, ethanol ‖ LiCl, 0.00147 **m**, AgCl, ethanol | Ag		0.0647	26Dru2
Ag | AgCl, LiCl, 0.0965 **m**, ethanol ‖ LiCl, 0.00295 **m**, AgCl, ethanol | Ag		0.0526	26Dru2
Ag | AgCl, LiCl, 0.0965 **m**, ethanol ‖ LiCl, 0.00590 **m**, AgCl, ethanol | Ag		0.0408	26Dru2
Ag | AgCl, LiCl, 0.0965 **m**, ethanol ‖ LiCl, 0.0120 **m**, AgCl, ethanol | Ag		0.0295	26Dru2
Ag | AgCl, LiCl, 0.0965 **m**, ethanol ‖ LiCl, 0.0240 **m**, AgCl, ethanol | Ag		0.0192	26Dru2
Ag | AgCl, LiCl, 0.0965 **m**, ethanol ‖ LiCl, 0.0480 **m**, AgCl, ethanol | Ag		0.0093	26Dru2
Ag | AgCl, LiCl, 0.005 **m**, ethanol ‖ $LiClO_4$, 0.005 **m**, ethanol ‖ $AgClO_4$, 0.005 **m**, ethanol | Ag		0.5311	32McF
Ag | AgCl, LiCl, 0.01 **m**, ethanol ‖ $LiClO_4$, 0.01 **m**, ethanol ‖ $AgClO_4$, 0.01 **m**, ethanol | Ag		0.5626	32McF
Ag | AgCl, HCl, 0.0020142 **m**, 1,4-dioxane(20wt%)/water ‖ AgCl, HCl, 0.038278 **m**, 1,4-dioxane(20wt%)/water | Ag (U_T=0.11700–2.47·10^{-4} (T–25 °C)–0.64·10^{-6} (T–25 °C)^2 [V])		0.11700	39Har4
Ag | AgCl, HCl, 0.0038779 **m**, 1,4-dioxane(20wt%)/water ‖ AgCl, HCl, 0.038278 **m**, 1,4-dioxane(20wt%)/water | Ag (U_T [V]=0.09090–1.82·10^{-4} (T–25 °C)–0.44·10^{-6} (T–25 °C)^2)		0.09090	39Har4
Ag | AgCl, HCl, 0.0085061 **m**, 1,4-dioxane(20wt%)/water ‖ AgCl, HCl, 0.038278 **m**, 1,4-dioxane(20wt%)/water | Ag (U_T [V]=0.05954–1.19·10^{-4} (T–25 °C)–0.29·10^{-6} (T–25 °C)^2)		0.05954	39Har4
Ag | AgCl, HCl, 0.018333 **m**, 1,4-dioxane(20wt%)/water ‖ AgCl, HCl, 0.038278 **m**, 1,4-dioxane(20wt%)/water | Ag (U_T [V]=0.11700–0.58·10^{-4} (T–25 °C)–0.15·10^{-6} (T–25 °C)^2)		0.02900	39Har4
Ag | AgCl, HCl, 0.080510 **m**, 1,4-dioxane(20wt%)/water ‖ AgCl, HCl, 0.038278 **m**, 1,4-dioxane(20wt%)/water | Ag (–U_T [V]=0.02935+0.59·10^{-4} (T–25 °C)+0.15·10^{-6} (T–25 °C)^2)		–0.02935	39Har5
Ag | AgCl, HCl, 0.165267 **m**, 1,4-dioxane(20wt%)/water ‖ AgCl, HCl, 0.038278 **m**, 1,4-dioxane(20wt%)/water | Ag (–U_T [V]=0.05796+1.10·10^{-4} (T–25 °C)+0.29·10^{-6} (T–25 °C)^2)		–0.05796	39Har4
Ag | AgCl, HCl, 0.33243 **m**, 1,4-dioxane(20wt%)/water ‖ AgCl, HCl, 0.038278 **m**, 1,4-dioxane(20wt%)/water | Ag (–U_T [V]=0.08634+1.60·10^{-4} (T–25 °C)+0.46·10^{-6} (T–25 °C)^2)		–0.08634	39Har4
Ag | AgCl, HCl, 0.70003 **m**, 1,4-dioxane(20wt%)/water ‖ AgCl, HCl, 0.038278 **m**, 1,4-dioxane(20wt%)/water | Ag (–U_T [V]=0.11908+2.13·10^{-4} (T–25 °C)+0.66·10^{-6} (T–25 °C)^2)		–0.11908	39Har4
Ag | AgCl, HCl, 1.49548 **m**, 1,4-dioxane(20wt%)/water ‖ AgCl, HCl, 0.038278 **m**, 1,4-dioxane(20wt%)/water | Ag (–U_T [V]=0.15870+2.67·10^{-4} (T–25 °C)+1.07·10^{-6} (T–25 °C)^2)		–0.15870	39Har4

Cell composition	T [°C]	U_0 [V]	Ref.
Ag | AgCl, HCl, 2.78348 **m**, 1,4-dioxane(20wt%)/water ‖ AgCl, HCl, 0.038278 **m**, 1,4-dioxane(20wt%)/water | Ag ($-U_T$ [V]=0.20077+3.10·10^{-4} (T–25 °C)+1.69·10^{-6} (T–25 °C)^2)		–0.20077	39Har4
Ag | AgCl, HCl, 0.0035521 **m**, 1,4-dioxane(45wt%)/water ‖ AgCl, HCl, 0.037604 **m**, 1,4-dioxane(45wt%)/water | Ag (U_T [V]=0.08765–1.74·10^{-4} (T–25 °C)–0.44·10^{-6} (T–25 °C)^2)		0.08765	39Har4
Ag | AgCl, HCl, 0.0084476 **m**, 1,4-dioxane(45wt%)/water ‖ AgCl, HCl, 0.037604 **m**, 1,4-dioxane(45wt%)/water | Ag (U_T [V]=0.05442–0.95·10^{-4} (T–25 °C)–0.34·10^{-6} (T–25 °C)^2)		0.05442	39Har4
Ag | AgCl, HCl, 0.018603 **m**, 1,4-dioxane(45wt%)/water ‖ AgCl, HCl, 0.037604 **m**, 1,4-dioxane(45wt%)/water | Ag (U_T [V]=0.02516–0.44·10^{-4} (T–25 °C)–0.19·10^{-6} (T–25 °C)^2)		0.02516	39Har4
Ag | AgCl, HCl, 0.076748 **m**, 1,4-dioxane(45wt%)/water ‖ AgCl, HCl, 0.037604 **m**, 1,4-dioxane(45wt%)/water | Ag ($-U_T$ [V]=0.02546+0.45·10^{-4} (T–25 °C)+0.20·10^{-6} (T–25 °C)^2)		–0.02546	39Har4
Ag | AgCl, HCl, 0.16490 **m**, 1,4-dioxane(45wt%)/water ‖ AgCl, HCl, 0.037604 **m**, 1,4-dioxane(45wt%)/water | Ag ($-U_T$ [V]=0.05313+0.95·10^{-4} (T–25 °C)+0.44·10^{-6} (T–25 °C)^2)		–0.05313	39Har4
Ag | AgCl, HCl, 0.29646 **m**, 1,4-dioxane(45wt%)/water ‖ AgCl, HCl, 0.037604 **m**, 1,4-dioxane(45wt%)/water | Ag ($-U_T$ [V]=0.07540+1.32·10^{-4} (T–25 °C)+0.62·10^{-6} (T–25 °C)^2)		–0.07540	39Har4
Ag | AgCl, HCl, 0.68931 **m**, 1,4-dioxane(45wt%)/water ‖ AgCl, HCl, 0.037604 **m**, 1,4-dioxane(45wt%)/water | Ag ($-U_T$ [V]=0.11058+1.86·10^{-4} (T–25 °C)+0.87·10^{-6} (T–25 °C)^2)		–0.11058	39Har4
Ag | AgCl, HCl, 1.50730 **m**, 1,4-dioxane(45wt%)/water ‖ AgCl, HCl, 0.037604 **m**, 1,4-dioxane(45wt%)/water | Ag ($-U_T$ [V]=0.15212+2.29·10^{-4} (T–25 °C)+1.11·10^{-6} (T–25 °C)^2)		–0.15212	39Har4
Ag | AgCl, HCl, 2.44143 **m**, 1,4-dioxane(45wt%)/water ‖ AgCl, HCl, 0.037604 **m**, 1,4-dioxane(45wt%)/water | Ag ($-U_T$ [V]=0.18579+2.36·10^{-4} (T–25 °C)+1.26·10^{-6} (T–25 °C)^2)		–0.18579	39Har4
Ag | AgCl, HCl, 2.99791 **m**, 1,4-dioxane(45wt%)/water ‖ AgCl, HCl, 0.037604 **m**, 1,4-dioxane(45wt%)/water | Ag ($-U_T$ [V]=0.20336+2.35·10^{-4} (T–25 °C)+1.32·10^{-6} (T–25 °C)^2)		–0.20336	39Har4
Ag | AgCl, HCl, 0.0038857 **m**, 1,4-dioxane(70wt%)/water ‖ AgCl, HCl, 0.036576 **m**, 1,4-dioxane(70wt%)/water | Ag (U_T [V]=0.06200–0.82·10^{-4} (T–25 °C)–0.40·10^{-6} (T–25 °C)^2)		0.06200	39Har4
Ag | AgCl, HCl, 0.0085917 **m**, 1,4-dioxane(70wt%)/water ‖ AgCl, HCl, 0.036576 **m**, 1,4-dioxane(70wt%)/water | Ag (U_T [V]=0.03905–0.60·10^{-4} (T–25 °C)–0.31·10^{-6} (T–25 °C)^2)		0.03905	39Har4
Ag | AgCl, HCl, 0.0199096 **m**, 1,4-dioxane(70wt%)/water ‖ AgCl, HCl, 0.036576 **m**, 1,4-dioxane(70wt%)/water | Ag (U_T [V]=0.01673–0.28·10^{-4} (T–25 °C)–0.17·10^{-6} (T–25 °C)^2)		0.01673	39Har4
Ag | AgCl, HCl, 0.077959 **m**, 1,4-dioxane(70wt%)/water ‖ AgCl, HCl, 0.036576 **m**, 1,4-dioxane(70wt%)/water | Ag ($-U_T$ [V]=0.02090+0.37·10^{-4} (T–25 °C)+0.28·10^{-6} (T–25 °C)^2)		–0.02090	39Har4
Ag | AgCl, HCl, 0.15798 **m**, 1,4-dioxane(70wt%)/water ‖ AgCl, HCl,0.036576 **m**, 1,4-dioxane(70wt%)/water | Ag ($-U_T$ [V])=0.04211+0.71·10^{-4} (T–25 °C)+0.57·10^{-6} (T–25 °C)^2		–0.04211	39Har4

Cell composition	T [°C]	U_0 [V]	Ref.
Ag | AgCl, HCl, 0.33140 **m**, 1,4-dioxane(70wt%)/water ‖ AgCl, HCl, 0.036576 **m**, 1,4-dioxane(70wt%)/water | Ag ($-U_T$ [V]=0.06499+1.08·10^{-4} (T–25 °C)+0.91·10^{-6} (T–25 °C)^2)		–0.06499	39Har4
Ag | AgCl, HCl, 0.35639 **m**, 1,4-dioxane(70wt%)/water ‖ AgCl, HCl, 0.036576 **m**, 1,4-dioxane(70wt%)/water | Ag ($-U_T$ [V]=0.06796+1.10·10^{-4} (T–25 °C)+0.92·10^{-6} (T–25 °C)^2)		–0.06796	39Har4
Ag | AgCl, HCl, 0.63909 **m**, 1,4-dioxane(70wt%)/water ‖ AgCl, HCl, 0.036576 **m**, 1,4-dioxane(70wt%)/water | Ag ($-U_T$ [V]=0.09225+1.39·10^{-4} (T–25 °C)+1.23·10^{-6} (T–25 °C)^2)		–0.09225	39Har4
Ag | AgCl, HCl, 0.91788 **m**, 1,4-dioxane(70wt%)/water ‖ AgCl, HCl, 0.036576 **m**, 1,4-dioxane(70wt%)/water | Ag ($-U_T$ [V]=0.11065+1.57·10^{-4} (T–25 °C)+1.42·10^{-6} (T–25 °C)^2)		–0.11065	39Har4
Ag | AgCl, HCl, 1.30623 **m**, 1,4-dioxane(70wt%)/water ‖ AgCl, HCl, 0.036576 **m**, 1,4-dioxane(70wt%)/water | Ag ($-U_T$ [V]=0.12916+1.76·10^{-4} (T–25 °C)+1.62·10^{-6} (T–25 °C)^2)		–0.12916	39Har4
Ag | AgCNS,sol., LiCNS, 0.01 **m**,ethanol ‖ LiCl, 0.01 **m**, AgCl, ethanol | Ag		0.157	32McF
Ag | AgI, NaI, 0.0989 **m**, ethanol | Na_xHg | NaI, 0.004133 **m**, AgI, ethanol | Ag		0.1328	30Par1
Ag | AgI, NaI, 0.1003 **m**, ethanol | Na_xHg | NaI, 0.00609 **m**, AgI, ethanol | Ag		0.1177	30Par1
Ag | AgI, NaI, 0.0996 **m**, ethanol | Na_xHg | NaI, 0.00920 **m**, AgI, ethanol | Ag		0.1012	30Par1
Ag | AgI, NaI, 0.0991 **m**, ethanol | Na_xHg | NaI, 0.01576 **m**, AgI, ethanol | Ag		0.0778	30Par1
Ag | AgI, NaI, 0.0998 **m**, ethanol | Na_xHg | NaI, 0.01945 **m**, AgI, ethanol | Ag		0.0690	30Par1
Ag | AgI, NaI, 0.1002 **m**, ethanol | Na_xHg | NaI, 0.02970 **m**, AgI, ethanol | Ag		0.0513	30Par1
Ag | AgI, NaI, 0.1003 **m**, ethanol | Na_xHg | NaI, 0.0761 **m**, AgI, ethanol | Ag		0.0116	30Par1
Ag | AgI, NaI, 0.1015 **m**, ethanol | Na_xHg | NaI, 0.002217 **m**, AgI, ethanol | Ag		0.1494	30Par1
Ag | AgI, NaI, 0.1968 **m**, ethanol | Na_xHg | NaI, 0.009923 **m**, AgI, ethanol | Ag		0.1257	30Par1
Ag | AgI ,NaI, 0.5704 **m**, ethanol | Na_xHg | NaI, 0.09544 **m**, AgI, ethanol | Ag		0.0720	30Par1
Ag | AgI, NaI, 0.5704 **m**, ethanol | Na_xHg | NaI, 0.01010 **m**, AgI, ethanol | Ag		0.1678	30Par1
Ag | AgI, NaI, 0.0970 **M**, ethanol | NaI, 0.0066 **M**, KI, 0.0031 **M**, AgI, ethanol | Ag		0.0341	30Par2
Ag | AgI, NaI, 0.1002 **M**, ethanol | NaI, 0.0051 **M**, KI, 0.0049 **M**, AgI, ethanol | Ag		0.0390	30Par2
Ag | AgI, NaI, 0.0987 **M**, ethanol | NaI, 0.0036 **M**, KI, 0.0064 **M**, AgI, ethanol | Ag		0.0430	30Par2
Ag | AgI, NaI, 0.0970 **M**, ethanol | NaI, 0.0017 **M**, KI, 0.0087 **M**, AgI, ethanol | Ag		0.0462	30Par2
Ag | AgI, NaI, 0.0980 **M**, ethanol | NaI, 0.0010 **M**, KI, 0.0082 **M**, AgI, ethanol | Ag		0.0502	30Par2
Ag | AgI, NaI, 0.0993 **M**, ethanol | NaI, 0.0166 **M**, KI, 0.0041 **M**, AgI, ethanol | Ag		0.0278	30Par2
Ag | AgI, NaI, 0.0993 **M**, ethanol | NaI, 0.0162 **M**, KI, 0.0047 **M**, AgI, ethanol | Ag		0.0279	30Par2

Cell composition	T [°C]	U_0 [V]	Ref.
Ag | AgI, NaI, 0.0997 **M**, ethanol | NaI, 0.0126 **M**, KI, 0.0068 **M**, AgI, ethanol | Ag		0.0300	30Par2
Ag | AgI, NaI, 0.0997 **M**, ethanol | NaI, 0.0074 **M**, KI, 0.0127 **M**, AgI, ethanol | Ag		0.0314	30Par2
Ag | AgI, NaI, 0.0970 **M**, ethanol | NaI, 0.0018 **M**, KI, 0.0172 **M**, AgI, ethanol | Ag		0.0344	30Par2
Ag | $AgNO_3$, 0.00205 **m**, acetonitrile ‖ $AgNO_3$, 0.00412 **m**, acetonitrile | Ag		0.012	28Koc
Ag | $AgNO_3$, 0.00205 **m**, acetonitrile ‖ $AgNO_3$, 0.0083 **m**, acetonitrile | Ag		0.030	28Koc
Ag | $AgNO_3$, 0.00205 **m**, acetonitrile ‖ $AgNO_3$, 0.0152 **m**, acetonitrile | Ag		0.047	28Koc
Ag | $AgNO_3$, 0.00205 **m**, acetonitrile ‖ $AgNO_3$, 0.0305 **m**, acetonitrile | Ag		0.059	28Koc
Ag | $AgNO_3$, 0.00205 **m**, acetonitrile ‖ $AgNO_3$, 0.061 **m**, acetonitrile | Ag		0.075	28Koc
Ag | $AgNO_3$, 0.00205 **m**, acetonitrile ‖ $AgNO_3$, 0.122 **m**, acetonitrile | Ag		0.086	28Koc
Ag | $AgNO_3$, 0.0078 **m**, acetonitrile ‖ $AgNO_3$, 0.125 **m**, acetonitrile | Ag		0.046	00Kah
Ag | $AgNO_3$, 0.001 **m**, pyridine ‖ $AgNO_3$, 0.1 **m**, pyridine | Ag		0.076	00Kah
Ag | $AgNO_3$, 0.002 **m**, pyridine ‖ $AgNO_3$, 0.1 **m**, pyridine | Ag		0.061	00Kah
Ag | $AgNO_3$, 0.01 **m**, pyridine ‖ $AgNO_3$, 0.1 **m**, pyridine | Ag		0.035	00Kah
Ag | $AgNO_3$, 0.00156 **m**, benzonitrile ‖ $AgNO_3$, 0.00312 **m**, benzonitrile | Ag		0.012	28Koc
Ag | $AgNO_3$, 0.00156 **m**, benzonitrile ‖ $AgNO_3$, 0.00625 **m**, benzonitrile | Ag		0.023	28Koc
Ag | $AgNO_3$, 0.00156 **m**, benzonitrile ‖ $AgNO_3$, 0.0125 **m**, benzonitrile | Ag		0.036	28Koc
Ag | $AgNO_3$, 0.00156 **m**, benzonitrile ‖ $AgNO_3$, 0.025 **m**, benzonitrile | Ag		0.049	28Koc
Ag | $AgNO_3$, 0.00156 **m**, benzonitrile ‖ $AgNO_3$, 0.05 **m**, benzonitrile | Ag		0.061	28Koc
Ag | $AgNO_3$, 0.00156 **m**, benzonitrile ‖ $AgNO_3$, 0.1 **m**, benzonitrile | Ag		0.076	28Koc
Sodium (Na)			
$Na_{0.2wt\%}Hg$ | NaCl, 2.148 m**m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	36	–0.4835	51Sed
$Na_{0.2wt\%}Hg$ | NaCl, 4.968 m**m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	36	–0.4559	51Sed
$Na_{0.2wt\%}Hg$ | NaCl, 8.551 m**m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	36	–0.4421	51Sed
$Na_{0.2wt\%}Hg$ | NaCl, 15.18 m**m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	36	–0.4262	51Sed
$Na_{0.2wt\%}Hg$ | NaCl, 30.14 m**m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	36	–0.4106	51Sed
$Na_{0.2wt\%}Hg$ | NaCl, 74.97 m**m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	36	–0.3952	51Sed
$Na_{0.2wt\%}Hg$ | NaCl, 0.1058 **m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	36	–0.3865	51Sed
$Na_{0.2wt\%}Hg$ | NaCl, 0.2666 **m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	36	–0.3752	51Sed
$Na_{0.2wt\%}Hg$ | NaCl, 0.4316 **m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	36	–0.3702	51Sed
$Na_{0.2wt\%}Hg$ | NaCl, 0.5283 **m** in NH_3, $ZnCl_2$*6 NH_3,sol. | $Zn_{sat.}Hg$	36	–0.3672	51Sed
Na_zHg | $NaNO_3$, 0.060 m**M** in NH_3 ‖ $NaNO_3$, 2.245 m**M** in NH_3 | Na_zHg	–36	0.0603	05Cad
Na_zHg | $NaNO_3$, 0.367 m**M** in NH_3 ‖ $NaNO_3$, 2.245 m**M** in NH_3 | Na_zHg	–36	0.0307	05Cad
Na_zHg | $NaNO_3$, 0.312 m**M** in NH_3 ‖ $NaNO_3$, 11.7 m**M** in NH_3 | Na_zHg	–38	0.0572	05Cad
Na_zHg | $NaNO_3$, 1.91 m**M** in NH_3 ‖ $NaNO_3$, 11.7 m**M** in NH_3 | Na_zHg	–38	0.0276	05Cad
Na_zHg | $NaNO_3$, 1.22 m**M** in NH_3 ‖ $NaNO_3$, 45.8 m**M** in NH_3 | Na_zHg	–39	0.0465	05Cad
Na_zHg | $NaNO_3$, 7.48 m**M** in NH_3 ‖ $NaNO_3$, 45.8 m**M** in NH_3 | Na_zHg	–39	0.0236	05Cad
$Na_{0.206wt\%}Hg$ | NaBr,sol., NaBr,sat., AgBr, ethanol | Ag (silver electrolytically deposited)		1.7640	34Ish2
$Na_{0.206wt\%}Hg$ | NaBr,sol., NaBr,sat., AgBr, ethanol | Ag (silver produced by reduction of Ag_2O)		1.7631	34Ish2
$Na_{0.247wt\%}Hg$ | $NaClO_4$, 0.01 **m**, ethanol ‖ $LiClO_4$, 0.01 **m**, ethanol ‖ LiCl, 0.01 **m**, AgCl, ethanol | Ag		1.9943	32McF
$Na_{0.466wt\%}Hg$ | $NaClO_4$, 0.01 **m**, ethanol ‖ $LiClO_4$, 0.01 **m**, ethanol ‖ LiCl, 0.01 **m**, AgCl, ethanol | Ag		1.9943	32McF

Cell composition	T [°C]	U_0 [V]	Ref.
$Na_{0.206wt\%}Hg$ | NaI,sol., NaI,sat., CuI,sol., 3-methyl-1-butanol | Cu_yHg		1.4102	34Ish3
Na | NaI, ethyl amine | $Na_{0.206wt\%}Hg$		0.8456	10Lew
Na_xGa | NaI, ethyl amine | Na (mol fraction x=0.568 10^{-5})		0.7116	34Gil
Na_xGa | NaI, ethyl amine | Na (mol fraction x=3.493 10^{-5})		0.7025	34Gil
Na_xGa | NaI, ethyl amine | Na (mol fraction x=7.96 10^{-5})		0.7012	34Gil
Na_xGa | NaI, ethyl amine | Na (mol fraction x=13 10^{-5})		0.7001	34Gil
Thallium (Tl)			
$Tl_{sat.}Hg$ | TlCl,sol., NaCl, 0.1 **m**, methanol ‖ NaCl, 0.1 **m**, Hg_2Cl_2, methanol | Hg		0.642	29Buc
$Tl_{sat.}Hg$ | $TlClO_4$, 0.00987 **m**, methanol ‖ NaCl, 0.01 **m**, methanol ‖ NaCl, 0.1 **m**, Hg_2Cl_2, methanol | Hg		0.6237	29Buc
$Tl_{sat.}Hg$ | TlCl,sol., LiCl, 0.01 **m**, ethanol ‖ LiCl, 0.01 **m**, ethanol, AgCl | Ag		0.5639	32McF
$Tl_{sat.}Hg$ | $TlClO_4$,sol., LiCl, 0.03 **m**, ethanol ‖ LiCl, 0.03 **m**, ethanol, AgCl | Ag		0.5710	32McF
Tl | $TlNO_3$, 0.1 **M**, pyridine ‖ $AgNO_3$, 0.1 **M**, pyridine | Ag		0.699	99Kah
Zinc (Zn)			
$Zn_{0.0289at.fract.}Hg$ | $ZnCl_2$*10 NH_3, NH_4Cl, TlCl,sol. | $Tl_{0.0265at.fract.}Hg$		0.9394	34Ell
$Zn_{0.0289at.fract.}Hg$ | $ZnCl_2$*10 NH_3, NH_4Cl, TlCl,sol. | $Tl_{0.0913at.fract.}Hg$		0.8900	34Ell
$Zn_{0.0289at.fract.}Hg$ | $ZnCl_2$*10 NH_3, NH_4Cl, TlCl,sol. | $Tl_{0.428at.fract.}Hg$		0.8231	34Ell
$Zn_{0.0301at.fract.}Hg$ | $ZnCl_2$*10 NH_3, NH_4Cl, TlCl,sol. | $Tl_{0.0265at.fract.}Hg$	1.75	0.9680	34Ell
$Zn_{10wt\%}Hg$ | $ZnBr_2$,sol., $ZnBr_2$,sat., Hg_2Br_2,sol., ethanol | Hg	20	0.6806	32Ish1
	25	0.6813	32Ish1
	30	0.6820	32Ish1
	35	0.6826	32Ish1
	40	0.6833	32Ish1
	45	0.6841	32Ish1
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.01263 **m**, AgCl, acetone | Ag		1.007	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.02997 **m**, AgCl, acetone | Ag		1.001	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.08820 **m**, AgCl, acetone | Ag		0.9888	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.2176 **m**, AgCl, acetone | Ag		0.9783	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.4616 **m**, AgCl, acetone | Ag		0.9653	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 1.047 **m**, AgCl, acetone | Ag		0.9521	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.002115 **m**, AgCl, acetone(50mol%)/water | Ag		1.0839	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.02451 **m**, AgCl, acetone(50mol%)/water | Ag		1.0431	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.04656 **m**, AgCl, acetone(50mol%)/water | Ag		1.0314	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.1331 **m**, AgCl, acetone(50mol%)/water | Ag		1.0177	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.2479 **m**, AgCl, acetone(50mol%)/water | Ag		1.0061	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.4475 **m**, AgCl, acetone(50mol%)/water | Ag		0.9981	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 0.7099 **m**, AgCl, acetone(50mol%)/water | Ag		0.9894	32Ham
$Zn_{1wt\%}Hg$ | $ZnCl_2$, 1.4490 **m**, AgCl, acetone(50mol%)/water | Ag		0.9722	32Ham
Zn | $Zn(NO_3)_2$, 0.1 **M**, pyridine ‖ $Cd(NO_3)_2$, 0.05 **M**, pyridine | Cd	17	0.340	99Kah
Zn | $Zn(NO_3)_2$, 0.05 **M**, aniline ‖ $AgNO_3$, 0.1 **M**, aniline | Ag		0.916	99Kah

[1]) All mixed electrolyte solvent systems including those containing water are listed below, for a given element cells with inorganic solvents are listed first followed by solutions based on organic solvents.

[2]) Description of the cell composition is based on the Stockholm convention (1953). All data refer to room temperature RT if not stated otherwise.

References

99Kah Kahlenberg, L.: J. Phys. Chem. **3** (1899) 709.
00Kah Kahlenberg, L.: J. Phys. Chem. **4** (1900) 379.
05Cad Cady, H.P.: J. Phys. Chem. **9** (1905) 477.
06Wil Wilson, J.H.: Am. Chem. J. **35** (1906) 78.
09Neu Neustadt, J., Abegg, R.: Z. Phys. Chem. (Leipzig) **69** (1909) 486.
10Lew Lewis, G.N., Kraus, Ch.A.: J. Am. Chem. Soc. **32** (1910) 1459.
12Lew Lewis, G.N., Keyes, F.G.: J. Am. Chem. Soc. **34** (1912) 119.
12Ros Roshdestwensky, A., Lewis, W.C.M.: Proc. Chem. Soc. **27** (1912) 266.
13Bel Bell, J.M., Feild, A.L.: J. Am. Chem. Soc. **35** (1913) 715.
13Lew Lewis, G.N., Keyes, F.G.: J. Am. Chem. Soc. **35** (1913) 340.
13Ros Roshdestwensky, A., Lewis, W.C.M.: Proc. Chem. Soc. **28** (1913) 239.
15Cam Cambi, L.: Atti Accad. Naz. Lincei Cl. Sci. Fis. Mat. Nat. Rend. **24**(I) (1915) 932; **23**(II) (1915) 606.
15Lew Lewis, G.N., Argo, W.L.: J. Am. Chem. Soc. **37** (1915) 1983.
22Dan Danner, P.S.: J. Am. Chem. Soc. **44** (1922) 2832.
22Pea Pearce, J.N., Hart, H.B.: J. Am. Chem. Soc. **44** (1922) 2411.
23Gra Grant, J., Partington, J.R.: Trans. Faraday Soc. **19** (1923) 414.
23Mül Müller, R., Knaus, W.: Z. Anorg. Allg. Chem. **130** (1923) 173.
24Bje Bjerrum, N., Unmack, A., Zechmeister, L.: K. Dan. Vidensk. Selsk. Mat. Fys. Medd. **5** (1924) 11.
25Non Nonhebel, G., Hartley, H.: Philos. Mag. **50** (1925) 729.
26And Anderson, P.A.: J. Am. Chem. Soc. **48** (1926) 2285.
26Cos Costeanu, G.J.: C. R. Acad. Sci. (Paris) **197** (1926) 2285.
26Dru1 Drucker, C., Luft, F.: Z. Phys. Chem. **121** (1926) 307.
26Dru2 Drucker, C., Schingnitz, R.: Z. Phys. Chem. **122** (1926) 149.
26Luc Lucasse, W.W.: J. Am. Chem. Soc. **48** (1926) 626.
26Mül Müller, R.: Z. Anorg. Allg. Chem. **156** (1926) 65.
27Erd Erdey-Gruz, T.: Z. Phys. Chem. **131** (1927) 81.
27Wol Wolfenden, J.H., Wright, C.P., Rosskana, N.L., Buckley, P.S.: Trans. Faraday Soc. **23** (1927) 491.
28Koc Koch, K.W.: J. Am. Chem. Soc. **50** (1928) 924.
28Woo Woolcock, J.W., Hartley, H.: Philos. Mag. **5** (1928) 1133.
28Yos Yoshida, T.: Sci. Rep. Tohoku Univ. Ser. 1 **17** (1928) 1279.
29Buc Buckley, P.S., Hartley, H.: Philos. Mag. **8** (1929) 320.
29But Butler, J.A.V., Robertson, C.M.: Proc. R. Soc. London, Ser. A **125** (1929) 694.
29Isa Isaaks, G.F., Partington, J.R.: Trans. Faraday Soc. **25** (1929) 53.
29Mül1 Müller, R., Kreiner, F., Schmidt, H.J.: Monatsh. Chem. **53** (1929) 215.
29Mül2 Müller, R., Schmidt, H.J.: Monatsh. Chem. **53** (1929) 224.
30Afa Afanasiev, A.S.: J. Am. Chem. Soc. **52** (1930) 3477.
30Ake Akerlöf, G.: J. Am. Chem. Soc. **52** (1930) 2353.
30Par1 Partington, J.R., Simpson, H.G.: Trans. Faraday Soc. **26** (1930) 625.
30Par2 Partington, J.R., Simpson, H.G.: Trans. Faraday Soc. **26** (1930) 147.
31Hut Hutchinson, A.W., Chandlee, G.C.: J. Am. Chem. Soc. **53** (1931) 2881.
32Ham Hamilton, R.T., Butler, J.A.V.: Proc. R. Soc. London **138** (1932) 450.
32Ish1 Ishikawa, F., Yoshida, T.: Sci. Rep. Tohoku Imp. Univ. Ser. 1 **21** (1932) 474.
32McF McFarlane, A., Hartley, H.: Philos. Mag. **13** (1932) 425.
33Ben Bent, H.E., Gilfillan jr., E.S.: J. Am. Chem. Soc. **55** (1933) 3989.
33Tri Trimble, H.M., Ebert, P.F.: J. Am. Chem. Soc. **55** (1933) 958.
34Ell Elliot, N., Yost, D.M.: J. Am. Chem. Soc. **56** (1934) 1057.
34Gil Gilfillan jr., E.S., Bent, H.E.: J. Am. Chem. Soc. **56** (1934) 1661.
34Goo Goodhue, L.D., Hixon, R.M.: J. Am. Chem. Soc. **56** (1934) 1924.
34Ish2 Ishikawa, F., Tachiki, K.: Sci. Rep. Tohoku Imp. Univ. Ser. 1 **23** (1934) 103.
34Ish3 Ishikawa, F., Tachiki, K., Murooko, T.: Sci. Rep. Tohoku Imp. Univ. Ser. 1 **23** (1934) 147.
34Mül Müller, R., Kunzmüller, H., Pinter, E., von Seebach, B.: Monatsh. Chem. **63** (1934) 317.
36Har3 Harned, H.S., Morrison, J.O.: J. Am. Chem. Soc. **58** (1936) 1908.
38Har2 Harned, H.S., Donelson, J.G.: J. Am. Chem. Soc. **60** (1938) 2128.
38Har3 Harned, H.S., Calmon, C.: J. Am. Chem. Soc. **60** (1938) 2130.
39Har3 Harned, H.S., Calmon, C.: J. Am. Chem. Soc. **61** (1939) 1891.
39Har4 Harned, H.S., Dreby, E.C.: J. Am. Chem. Soc. **61** (1939) 3113.

39Har5 Harned, H.S., Walker, F.: J. Am. Chem. Soc. **61** (1939) 48.
39Har6 Harned, H.S., Fallon, L.D.: J. Am. Chem. Soc. **61** (1939) 2377.
42Kan Kanning, E.W., Campbell, A.W.: J. Am. Chem. Soc. **64** (1942) 1672.
42Pat Patterson, A., Felsing, W.A.: J. Am. Chem. Soc. **64** (1942) 1480.
44Nob Noble, M.V., Garret, A.B.: J. Am. Chem. Soc. **66** (1944) 231.
46Cro Crockford, H.D., Wideman, S.A.: J. Phys. Chem. **50** (1946) 418.
46Har Harned, H.S., Nestler, F.H.M.: J. Am. Chem. Soc. **68** (1946) 665.
46Kan Kanning, E.W., Bowman, M.G.: J. Am. Chem. Soc. **68** (1946) 2042.
46Kni Knight, S.B., Mason, J.F., Roesel, D.: J. Am. Chem. Soc. **68** (1946) 661.
47Moo Moore, R.L., Felsing, W.A.: J. Am. Chem. Soc. **69** (1947) 1076.
50Cro2 Crockford, H.D., Knight, S.B., Staton, H.A.: J. Am. Chem. Soc. **72** (1950) 2164.
50Cro3 Crockford, H.D., Land, J.E.: J. Am. Chem. Soc. **72** (1950) 1895.
51Sed Sedlet, J., De Vries, T.: J. Am. Chem. Soc. **73** (1951) 5808.
52Koe Koerber, G.G., de Vries, T.: J. Am. Chem. Soc. **74** (1952) 5008.
53Eva Evans, J.I., Monk, C.B.: Trans. Faraday Soc. **49** (1953) 415.
54Eve Everett, D.M., Rasmussen, S.E.: J. Chem. Soc. (1954) 2812.
54Har Harned, H.S., Allen, D.S.: J. Phys. Chem. **58** (1954) 191.
54Muk Mukherjee, L.M.: J. Phys. Chem. **58** (1954) 1042.
58Man Mandel, M., Decroly, P.: Nature (London) **182** (1958) 794.
61Ive Ives, D.J.G., Janz, G.J.: Reference Electrodes, New York: Academic Press, 1961.

Table 2.3. Cell voltages with molten salt electrolyte systems. [1])

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Aluminum (Al)			
Al | NaCl, 13.74 wt%, KCl, 8.85 wt%, $AlCl_3$, 77.41 wt% | Cl_2, 720.5 Torr, C	96	2.092	33Tre
	102	2.084	33Tre
	124	2.070	33Tre
	135	2.062	33Tre
	166	2.039	33Tre
	196	2.001	33Tre
	220	1.975	33Tre
Bismuth (Bi)			
Bi | $BiCl_3$, KCl, LiCl | Bi_xAu, x=0.3192 (dU/dT=0.075·10^{-3} V/°C)	600	0.0357	51Kle
Bi | $BiCl_3$, KCl, LiCl | Bi_xAu, x=0.4475 (dU/dT=0.0335·10^{-3} V/°C)	600	0.02115	51Kle
Bi | $BiCl_3$, KCl, LiCl | Bi_xAu, x=0.5125 (dU/dT=0.0268·10^{-3} V/°C)	600	0.01788	51Kle
Bi | $BiCl_3$, KCl, LiCl | Bi_xAu, x=0.6015 (dU/dT=0.0183·10^{-3} V/°C)	600	0.01343	51Kle
Bi | $BiCl_3$, KCl, LiCl | Bi_xAu, x=0.7483 (dU/dT=0.0088·10^{-3} V/°C)	600	0.00743	51Kle
Bi | $BiCl_3$, KCl, LiCl | Bi_xAu, x=0.8926 (dU/dT=0.0044·10^{-3} V/°C)	600	0.00301	51Kle
Cadmium (Cd)			
Cd | $CdCl_2$, KCl, LiCl | Cd_xAu, x=0.158 (dU/dT= –0.258·10^{-3} V/°C)	602	0.2222	32Öla
Cd | $CdCl_2$, KCl, LiCl | Cd_xAu, x=0.227 (dU/dT=0.031·10^{-3} V/°C)	602	0.2112	32Öla
Cd | $CdCl_2$, KCl, LiCl | Cd_xAu, x=0.241 (dU/dT=0.007 mV/°C)	602	0.2097	32Öla
Cd | $CdCl_2$, KCl, LiCl | Cd_xAu, x=0.326 (dU/dT= –0.013 mV/°C)	602	0.1627	32Öla
Cd | $CdCl_2$, KCl, LiCl | Cd_xAu, x=0.418 (dU/dT= –0.063 mV/°C)	602	0.1531	32Öla
Cd | $CdCl_2$, KCl, LiCl | Cd_xAu, x=0.440 (dU/dT= –0.099 mV/°C)	602	0.1501	32Öla
Cd | $CdCl_2$, KCl, LiCl | Cd_xAu, x=0.493 (dU/dT= –0.110 mV/°C)	602	0.1283	32Öla
Cd | $CdCl_2$, KCl, LiCl | Cd_xAu, x=0.555 (dU/dT= –0.003 mV/°C)	602	0.0518	32Öla
Cd | $CdCl_2$, KCl, LiCl | Cd_xAu, x=0.599 (dU/dT= –0.095 mV/°C)	602	0.0448	32Öla
Cd | $CdCl_2$, KCl, LiCl | Cd_xAu, x=0.698 (dU/dT= –0.116 mV/°C)	602	0.0408	32Öla
Cd | $CdCl_2$, KCl, LiCl | Cd_xAu, x=0.664 (dU/dT= –0.0060 mV/°C)	602	0.0432	32Öla
Cd | $CdCl_2$, KCl, LiCl | Cd_xAu, x=0.714 (dU/dT= –0.0113 mV/°C)	602	0.0372	32Öla
Cd | $CdCl_2$, KCl, LiCl | Cd_xAu, x=0.763 (dU/dT=0.017 mV/°C)	602	0.0060	32Öla
Cd | $CdCl_2$, KCl, LiCl | Cd_xBi, x=0.1134	422	0.0673	23Tay
Cd | $CdCl_2$, KCl, LiCl | Cd_xBi, x=0.1134	528	0.0789	23Tay
Cd | $CdCl_2$, KCl, LiCl | Cd_xBi, x=0.282	422	0.0416	23Tay
Cd | $CdCl_2$, KCl, LiCl | Cd_xBi, x=0.282	528	0.0509	23Tay
Cd | $CdCl_2$, KCl, LiCl | Cd_xBi, x=0.472	422	0.0257	23Tay
Cd | $CdCl_2$, KCl, LiCl | Cd_xBi, x=0.472	528	0.0305	23Tay
Cd | $CdCl_2$, KCl, LiCl | Cd_xBi, x=0.864	422	0.0039	23Tay
Cd | $CdCl_2$, KCl, LiCl | Cd_xBi, x=0.864	528	0.0043	23Tay
Cd | $CdCl_2$, 4-5%, KCl, LiCl | Cd_xCu, x=0.463	602	0.0269	54Ric
Cd | $CdCl_2$, 4-5%, KCl, LiCl | Cd_xCu, x=0.507	602	0.023	54Ric
Cd | $CdCl_2$, 4-5%, KCl, LiCl | Cd_xCu, x=0.599	602	0.0155	54Ric
Cd | $CdCl_2$, 4-5%, KCl, LiCl | Cd_xCu, x=0.662	602	0.0120	54Ric
Cd | $CdCl_2$, 4-5%, KCl, LiCl | Cd_xCu, x=0.782	602	0.0073	54Ric
Cd | $CdCl_2$, 4-5%, KCl, LiCl | Cd_xCu, x=0.859	602	0.00460	54Ric
Cd | $CdCl_2$, 4-5%, KCl, LiCl | Cd_xCu, x=0.915	602	0.00328	54Ric
Cd | $CdCl_2$, KCl, LiCl | Cd_xPb, x=0.155	436	0.0155	23Tay

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Cd | $CdCl_2$, KCl, LiCl | Cd_xPb, x=0.401	436	0.0108	23Tay
Cd | $CdCl_2$, KCl, LiCl | Cd_xPb, x=0.584	436	0.0073	23Tay
Cd | $CdCl_2$, KCl, LiCl | Cd_xPb, x=0.786	436	0.00043	23Tay
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.1095, $x2/x3$=0.1974 (dU/dT=0.1052 mV/°C)	500	0.05079	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.1995, $x2/x3$=0.1974 (dU/dT=0.0780 mV/°C)	500	0.03433	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.3143, $x2/x3$=0.1974 (dU/dT=0.0602 mV/°C)	500	0.02377	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.4199, $x2/x3$=0.1974 (dU/dT=0.0466 mV/°C)	500	0.01729	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.5699, $x2/x3$=0.1974 (dU/dT=0.0291 mV/°C)	500	0.01101	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.6914, $x2/x3$=0.1974 (dU/dT=0.0219 mV/°C)	500	0.00788	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.8119, $x2/x3$=0.1974 (dU/dT=0.0123 mV/°C)	500	0.00509	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.9376, $x2/x3$=0.1974 (dU/dT=0.00351 mV/°C)	500	0.00179	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.1182, $x2/x3$=0.5 (dU/dT=0.101 mV/°C)	500	0.05962	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.2333, $x2/x3$=0.5 (dU/dT=0.0769 mV/°C)	500	0.03972	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.4291, $x2/x3$=0.5 (dU/dT=0.0449 mV/°C)	500	0.02241	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.5523, $x2/x3$=0.5 (dU/dT=0.0377 mV/°C)	500	0.01597	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.6291, $x2/x3$=0.5 (dU/dT=0.0304 mV/°C)	500	0.01247	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.7793, $x2/x3$=0.5 (dU/dT=0.0115 mV/°C)	500	0.00446	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.8221, $x2/x3$=0.5 (dU/dT=0.0112 mV/°C)	500	0.00513	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Bi_{x3}$, $x1$=0.9503, $x2/x3$=0.5 (dU/dT=0.00245 mV/°C)	500	0.00150	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.0963, $x2/x3$=0.4480 (dU/dT=0.1036 mV/°C)	500	0.09454	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.2326, $x2/x3$=0.4936 (dU/dT=0.0654 mV/°C)	500	0.06347	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.3980, $x2/x3$=0.4894 (dU/dT=0.0454 mV/°C)	500	0.04027	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.4984, $x2/x3$=0.5009 (dU/dT=0.0403 mV/°C)	500	0.02851	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.5910, $x2/x3$=0.4932 (dU/dT=0.0320 mV/°C)	500	0.02018	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.6985, $x2/x3$=0.4480 (dU/dT=0.0231 mV/°C)	500	0.01236	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.8096, $x2/x3$=0.5075 (dU/dT=0.0135 mV/°C)	500	0.00674	51Ell

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.9581, $x2/x3$=0.5258 (dU/dT=0.00262 mV/°C)	500	0.00163	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.0772, $x2/x3$=1.896 (dU/dT=0.0985 mV/°C)	500	0.07050	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.2497, $x2/x3$=1.932 (dU/dT=0.0677 mV/°C)	500	0.04099	51Ell
Cd | $CdCl_2$, 5% ,KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.3731, $x2/x3$=2.548 (dU/dT=0.0517 mV/°C)	500	0.02523	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.4926, $x2/x3$=1.846 (dU/dT=0.0396 mV/°C)	500	0.01898	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.5928, $x2/x3$=1.900 (dU/dT=0.0293 mV/°C)	500	0.01328	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.6758, $x2/x3$=2.188 (dU/dT=0.0220 mV/°C)	500	0.00969	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.8002, $x2/x3$=1.842 (dU/dT=0.0132 mV/°C)	500	0.00602	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sb_{x3}$, $x1$=0.9498, $x2/x3$=1.789 (dU/dT=0.00270 mV/°C)	500	0.00152	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.0992, $x2/x3$=0.5 (dU/dT=0.1138 mV/°C)	500	0.0603	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.2738, $x2/x3$=0.5079 (dU/dT=0.0680 mV/°C)	500	0.03182	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.4023, $x2/x3$=0.5 (dU/dT=0.0486 mV/°C)	500	0.02222	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.4983, $x2/x3$=0.4981 (dU/dT=0.0382 mV/°C)	500	0.01703	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.5966, $x2/x3$=0.5 (dU/dT=0.0288 mV/°C)	500	0.01309	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.6928, $x2/x3$=0.5 (dU/dT=0.0210 mV/°C)	500	0.00963	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.8042, $x2/x3$=0.5 (dU/dT=0.0121 mV/°C)	500	0.00608	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.9055, $x2/x3$=0.5 (dU/dT=0.00512 mV/°C)	500	0.00298	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.9403, $x2/x3$=0.527 (dU/dT=0.00290 mV/°C)	500	0.00186	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.1067, $x2/x3$=2.0 (dU/dT=0.1138 mV/°C)	500	0.05338	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.2491, $x2/x3$=2.0 (dU/dT=0.0735 mV/°C)	500	0.03098	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.4041, $x2/x3$=2.0 (dU/dT=0.0484 mV/°C)	500	0.01950	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.5160, $x2/x3$=2.0 (dU/dT=0.0352 mV/°C)	500	0.01433	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.6078, $x2/x3$=2.0 (dU/dT=0.0268 mV/°C)	500	0.01106	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.6986, $x2/x3$=2.0 (dU/dT=0.0192 mV/°C)	500	0.00840	51Ell

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.7967, $x2/x3$=2.0 (dU/dT=0.0123 mV/°C)	500	0.00571	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.8930, $x2/x3$=2.14 (dU/dT=0.00627 mV/°C)	500	0.00329	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | $Cd_{x1}Pb_{x2}Sn_{x3}$, $x1$=0.9546, $x2/x3$=2.0 (dU/dT=0.00243 mV/°C)	500	0.00146	51Ell
Cd | $CdCl_2$, KCl, LiCl | Cd_xSb, x=0.3444 ($dU/dT = 46.7\cdot10^{-3}$ mV/°C, 495<T<515 °C)	480	0.06395	38Sel
Cd | $CdCl_2$, KCl, LiCl | Cd_xSb, x=0.3745 (dU/dT=$41.9\cdot10^{-3}$ mV/°C, 480<T<500 °C)	480	0.05990	38Sel
Cd | $CdCl_2$,KCl,LiCl | Cd_xSb, x=0.4006 (dU/dT=$38.0\cdot10^{-3}$ mV/°C, 460<T<480 °C)	480	0.05651	38Sel
Cd | $CdCl_2$,KCl,LiCl | Cd_xSb, x=0.4340 (dU/dT=$32.0\cdot10^{-3}$ mV/°C, 450<T<480 °C)	480	0.05200	38Sel
Cd | $CdCl_2$,KCl,LiCl | Cd_xSb, x=0.5590 (dU/dT=$31.4\cdot10^{-3}$ mV/°C, 450<T<480 °C)	480	0.03385	38Sel
Cd | $CdCl_2$,KCl,LiCl | Cd_xSb, x=0.5880 (dU/dT=$35.6\cdot10^{-3}$ mV/°C, 450<T<485 °C)	480	0.02969	38Sel
Cd | $CdCl_2$,KCl,LiCl | Cd_xSb, x=0.6760 (dU/dT=$48.0\cdot10^{-3}$ mV/°C, 400<T<450 °C)	480	0.01896	38Sel
Cd | $CdCl_2$,KCl,LiCl | Cd_xSb, x=0.7497 (dU/dT=$25.4\cdot10^{-3}$ mV/°C, 420<T<470 °C)	480	0.01156	38Sel
Cd | $CdCl_2$,KCl,LiCl | Cd_xSb, x=0.8180 (dU/dT=$17.5\cdot10^{-3}$ mV/°C, 370<T<460 °C)	480	0.00720	38Sel
Cd | $CdCl_2$,KCl,LiCl | Cd_xSb, x=0.8960 (dU/dT=$6.4\cdot10^{-3}$ mV/°C, 370<T<400 °C)	480	0.00331	38Sel
Cd | $CdCl_2$, KCl, LiCl | Cd_xSn, x=0.0835	430	0.0578	23Tay
	585	0.0749	23Tay
Cd | $CdCl_2$, KCl, LiCl | Cd_xSn, x=0.258	430	0.0292	23Tay
	585	0.0395	23Tay
Cd | $CdCl_2$, KCl, LiCl | Cd_xSn, x=0.569	430	0.0129	23Tay
	585	0.0173	23Tay
Cd | $CdCl_2$, KCl, LiCl | Cd_xSn, x=0.816	430	0.0051	23Tay
	585	0.0064	23Tay
Cd | $CdCl_2$ | Cl_2, C	600	1.340	29Lor
	650	1.309	29Lor
	700	1.279	29Lor
	750	1.249	29Lor
Cd | $CdCl_2$, x=0.4, $BaCl_2$ | Cl_2, C (dU/dT= –0.604 mV/°C)	600	1.3323	53Lan1
Cd | $CdCl_2$, x=0.5, $BaCl_2$ | Cl_2, C (dU/dT= –0.567 mV/°C)	600	1.3142	53Lan1
Cd | $CdCl_2$, x=0.6, $BaCl_2$ | Cl_2, C (dU/dT= –0.555 mV/°C)	600	1.3001	53Lan1
Cd | $CdCl_2$, x=0.7, $BaCl_2$ | Cl_2, C (dU/dT= –0.549 mV/°C)	600	1.2097	53Lan1
Cd | $CdCl_2$, x=0.8, $BaCl_2$ | Cl_2, C (dU/dT= –0.538 mV/°C)	600	1.2874	53Lan1
Cd | $CdCl_2$, x=0.3, KCl | Cl_2, C (dU/dT= –0.579 mV/°C)	600	1.5094	53Lan1
Cd | $CdCl_2$, x=0.4, KCl | Cl_2, C (dU/dT= –0.578 mV/°C)	600	1.4522	53Lan1
Cd | $CdCl_2$, x=0.5, KCl | Cl_2, C (dU/dT= –0.5830 mV/°C)	600	1.4022	53Lan1
Cd | $CdCl_2$, x=0.6, KCl | Cl_2, C (dU/dT= –0.5715 mV/°C)	600	1.3775	53Lan1
Cd | $CdCl_2$, x=0.7, KCl | Cl_2, C (dU/dT= –0.560 mV/°C)	600	1.3609	53Lan1
Cd | $CdCl_2$, x=0.9, KCl | Cl_2, C (dU/dT= –0.528 mV/°C)	600	1.3405	53Lan1

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Cd | $CdCl_2$, x=0.3, NaCl | Cl_2, C (dU/dT= –0.5010 mV/°C)	600	1.4199	53Lan1
Cd | $CdCl_2$, x=0.4, NaCl | Cl_2, C (dU/dT= –0.5720 mV/°C)	600	1.3978	53Lan1
Cd | $CdCl_2$, x=0.5, NaCl | Cl_2, C (dU/dT= –0.5638 mV/°C)	600	1.3798	53Lan1
Cd | $CdCl_2$, x=0.6, NaCl | Cl_2, C (dU/dT= –0.5518 mV/°C)	600	1.3597	53Lan1
Cd | $CdCl_2$, x=0.7, NaCl | Cl_2, C (dU/dT= –0.5553 mV/°C)	600	1.3491	53Lan1
Cd | $CdCl_2$, x=0.8, NaCl | Cl_2, C (dU/dT= –0.5438 mV/°C)	600	1.3436	53Lan1
Cd | $CdCl_2$, x=1.0, NaCl | Cl_2, C (dU/dT= –0.5684 mV/°C)	600	1.3382	53Lan1
Cd | $CdCl_2$, 5%, KCl, LiCl | Pb_xBi, x=0.1357 (dU/dT=0.0894 mV/°C)	500	0.08608	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | Pb_xBi, x=0.3333 (dU/dT=0.0474 mV/°C)	500	0.04723	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | Pb_xBi, x=0.5017 (dU/dT=0.0312 mV/°C)	500	0.02960	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | Pb_xBi, x=0.5791 (dU/dT=0.0223 mV/°C)	500	0.02271	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | Pb_xBi, x=0.9011 (dU/dT=0.00434 mV/°C)	500	0.00357	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | Pb_xSb, x=0.6902 (dU/dT=0.0173 mV/°C)	500	0.01290	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | Pb_xSb, x=0.7668 (dU/dT=0.0128 mV/°C)	500	0.00908	51Ell
Cd | $CdCl_2$, 5%, KCl, LiCl | Pb_xSb, x=0.8952 (dU/dT=0.0051 mV/°C)	500	0.00369	51Ell
Lead (Pb)			
Pb | $PbBr_2$ | Br_2, C	400	1.0928	33Sal
	450	1.0624	33Sal
	500	1.0321	33Sal
	550	1.0017	33Sal
Pb | $PbBr_2$, x=0.200, $ZnBr_2$ | Br_2, C	400	1.1297	33Sal
	450	1.1045	33Sal
	500	1.0810	33Sal
Pb | $PbBr_2$, x=0.210, $ZnBr_2$ | Br_2, C	400	1.1268	33Sal
	450	1.1033	33Sal
	500	1.0798	33Sal
Pb | $PbBr_2$, x=0.288, $ZnBr_2$ | Br_2, C	400	1.1211	33Sal
	450	1.0956	33Sal
	500	1.0700	33Sal
Pb | $PbBr_2$, x=0.392, $ZnBr_2$ | Br_2, C	400	1.1155	33Sal
	450	1.0884	33Sal
	500	1.0613	33Sal
Pb | $PbBr_2$, x=0.486, $ZnBr_2$ | Br_2, C	400	1.1114	33Sal
	450	1.0832	33Sal
	500	1.0550	33Sal
Pb | $PbBr_2$, x=0.604, $ZnBr_2$ | Br_2, C	400	1.1058	33Sal
	450	1.0772	33Sal
	500	1.0487	33Sal
Pb | $PbBr_2$, x=0.704, $ZnBr_2$ | Br_2, C	400	1.1016	33Sal
	450	1.0726	33Sal
	500	1.0436	33Sal
Pb | $PbBr_2$, x=0.794, $ZnBr_2$ | Br_2, C	400	1.0986	33Sal
	450	1.0691	33Sal
	500	1.0396	33Sal
Pb | $PbBr_2$, x=0.902, $ZnBr_2$ | Br_2, C	400	1.0951	33Sal
	450	1.0652	33Sal
	500	1.0352	33Sal

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Pb | $PbBr_2$, x=1.0, $ZnBr_2$ | Br_2, C	400	1.0928	33Sal
	450	1.0624	33Sal
	500	1.0321	33Sal
Pb | $PbCl_2$ | Cl_2, C	500	1.273 [3])	35Hol, 29Lor, 33Sal, 30Wac
		1.275	53Lan1
	550	1.244 [3])	35Hol, 29Lor, 33Sal, 30Wac
		1.248	53Lan1
	600	1.213 [3])	35Hol, 29Lor, 33Sal, 30Wac
		1.221	53Lan1
	650	1.183 [3])	35Hol, 29Lor, 33Sal, 30Wac
		1.195	53Lan1
	700	1.153 [3])	35Hol, 29Lor, 33Sal, 30Wac
		1.168	53Lan1
	750	1.121 [3])	35Hol, 29Lor, 33Sal, 30Wac
		1.142	53Lan1
	800	1.093 [3])	35Hol, 29Lor, 33Sal, 30Wac
		1.115	53Lan1
Pb | $PbCl_2$, x=0.450, $PbBr_2$ | Br_2, C	450	1.1038	30Sal
	500	1.0768	30Sal
	550	1.0469	30Sal
Pb | $PbCl_2$, x=0.500, $PbBr_2$ | Br_2, C	450	1.0959	30Sal
	500	1.0687	30Sal
	550	1.0415	30Sal
Pb | $PbCl_2$, x=0.600, $PbBr_2$ | Br_2, C	450	1.0866	30Sal
	500	1.0593	30Sal
	550	1.0318	30Sal
Pb | $PbCl_2$, x=0.800, $PbBr_2$ | Br_2, C	450	1.0736	30Sal
	500	1.0447	30Sal
	550	1.0156	30Sal

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Pb | $PbCl_2$, x=1.0, $PbBr_2$ | Br_2, C	450	1.0624	30Sal
	500	1.0321	30Sal
	550	1.0017	30Sal
Pb | $PbCl_2$, x=0.6, $BaCl_2$ | Cl_2, C (dU/dT= –0.480 mV/°C)	700	1.1919	53Lan2
Pb | $PbCl_2$, x=0.8, $BaCl_2$ | Cl_2, C (dU/dT= –0.505 mV/°C)	700	1.1777	53Lan2
Pb | $PbCl_2$, x=0.6, $CaCl_2$ | Cl_2, C (dU/dT= –0.4935 mV/°C)	700	1.1860	53Lan2
Pb | $PbCl_2$, x=0.8, $CaCl_2$ | Cl_2, C (dU/dT= –0.5212 mV/°C)	700	1.1757	53Lan2
Pb | $PbCl_2$, x=0.4, KCl | Cl_2, C (dU/dT= –0.5144 mV/°C)	500	1.3506	53Lan2
Pb | $PbCl_2$, x=0.5, KCl | Cl_2, C (dU/dT= –0.5216 mV/°C)	500	1.3290	53Lan2
Pb | $PbCl_2$, x=0.6, KCl | Cl_2, C (dU/dT= –0.5365 mV/°C)	500	1.3133	53Lan2
Pb | $PbCl_2$, x=0.7, KCl | Cl_2, C (dU/dT= –0.5445 mV/°C)	500	1.3019	53Lan2
Pb | $PbCl_2$, x=0.8, KCl | Cl_2, C (dU/dT= –0.5260 mV/°C)	500	1.2851	53Lan2
Pb | $PbCl_2$, x=0.9, KCl | Cl_2, C (dU/dT= –0.5150 mV/°C)	500	1.2754	53Lan2
Pb | $PbCl_2$, x=0.466, LiCl | Cl_2, C (dU/dT= –0.5020 mV/°C)	500	1.2980	53Lan2
Pb | $PbCl_2$, x=0.6625, LiCl | Cl_2, C (dU/dT= –0.5126 mV/°C)	500	1.2867	53Lan2
Pb | $PbCl_2$, x=0.840, LiCl | Cl_2, C (dU/dT= –0.5235 mV/°C)	500	1.2800	53Lan2
Pb | $PbCl_2$, x=1.0, LiCl | Cl_2, C (dU/dT= –0.531 mV/°C)	500	1.2745	53Lan2
Pb | $PbCl_2$, x=0.50, NaCl | Cl_2, C (dU/dT= –0.492 mV/°C)	500	1.3021	53Lan2
Pb | $PbCl_2$, x=0.60, NaCl | Cl_2, C (dU/dT= –0.5050 mV/°C)	500	1.2952	53Lan2
Pb | $PbCl_2$, x=0.75, NaCl | Cl_2, C (dU/dT= –0.5245 mV/°C)	500	1.2876	53Lan2
Pb | $PbCl_2$, x=0.9, NaCl | Cl_2, C (dU/dT= –0.5285 mV/°C)	500	1.2784	53Lan2
Pb | $PbCl_2$, x=0.6, $SrCl_2$ | Cl_2, C (dU/dT= –0.446 mV/°C)	700	1.1877	53Lan2
Pb | $PbCl_2$, x=0.8, $SrCl_2$ | Cl_2, C (dU/dT= –0.505 mV/°C)	700	1.1773	53Lan2
Pb | $PbCl_2$, x=0.301, $ZnCl_2$ | Cl_2, C	500	1.3315	30Wac
	600	1.2800	30Wac
Pb | $PbCl_2$, x=0.490, $ZnCl_2$ | Cl_2, C	500	1.3095	30Wac
	600	1.2520	30Wac
Pb | $PbCl_2$, x=0.595, $ZnCl_2$ | Cl_2, C	500	1.2990	30Wac
	600	1.2405	30Wac
Pb | $PbCl_2$, x=0.688, $ZnCl_2$ | Cl_2, C	500	1.2905	30Wac
	600	1.2310	30Wac
Pb | $PbCl_2$, x=0.885, $ZnCl_2$ | Cl_2, C	500	1.2800	30Wac
	600	1.2185	30Wac
Pb | $PbCl_2$, x=1.0, $ZnCl_2$ | Cl_2, C	500	1.2730	30Wac
	600	1.2105	30Wac
Pb | $PbCl_2$, 5 wt%, KCl, LiCl | Pb_xAu, x=0.2135 (dU/dT=0.101·10^{-3} V/°C, 730<T<830 °C)	600	0.0912	49Kle
Pb | $PbCl_2$, 5 wt%, KCl, LiCl | Pb_xAu, x=0.3336 (dU/dT=0.0625·10^{-3} V/°C, 600<T<700 °C)	600	0.06025	49Kle
Pb | $PbCl_2$, 5 wt%, KCl, LiCl | Pb_xAu, x=0.4247 (dU/dT = 0.0468·10^{-3} V/°C, 500<T<700 °C)	600	0.04580	49Kle
Pb | $PbCl_2$, 5 wt%, KCl, LiCl | Pb_xAu, x=0.5078 (dU/dT = 0.0349·10^{-3} V/°C, 500<T<700 °C)	600	0.03515	49Kle
Pb | $PbCl_2$, 5 wt%, KCl, LiCl | Pb_xAu, x=0.5935 (dU/dT = 0.0235·10^{-3} V/°C, 450<T<700 °C)	600	0.02584	49Kle
Pb | $PbCl_2$, 5 wt%, KCl, LiCl | Pb_xAu, x=0.6715 (dU/dT = 0.0166·10^{-3} V/°C, 450<T<700 °C)	600	0.01886	49Kle

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Pb | $PbCl_2$, 5 wt%, KCl, LiCl | Pb_xAu, x=0.8205 ($dU/dT = 0.0083 \cdot 10^{-3}$ V/°C, 450<T<700 °C)	600	0.00881	49Kle
Pb | $PbCl_2$, 5 wt%, KCl, LiCl | Pb_xAu, x=0.9665 ($dU/dT = 0.0012 \cdot 10^{-3}$ V/°C, 450<T<700 °C)	600	0.000140	49Kle
Pb | $PbCl_2$, KCl, LiCl | $PbBi_x$, x=0.152 (dU/dT=0.0074·10^{-3} V/°C)	700	0.00532	36Sel
Pb | $PbCl_2$, KCl, LiCl | $PbBi_x$, x=0.280 (dU/dT=0.0144·10^{-3} V/°C)	700	0.01148	36Sel
Pb | $PbCl_2$, KCl, LiCl | $PbBi_x$, x=0.400 (dU/dT=0.0208·10^{-3} V/°C)	700	0.01929	36Sel
Pb | $PbCl_2$, KCl, LiCl | $PbBi_x$, x=0.504 (dU/dT=0.0278·10^{-3} V/°C)	700	0.02782	36Sel
Pb | $PbCl_2$, KCl, LiCl | $PbBi_x$, x=0.586 (dU/dT=0.0376·10^{-3} V/°C)	700	0.03594	36Sel
Pb | $PbCl_2$, KCl, LiCl | $PbBi_x$, x=0.672 (dU/dT=0.0464·10^{-3} V/°C)	700	0.04540	36Sel
Pb | $PbCl_2$, KCl, LiCl | $PbBi_x$, x=0.770 (dU/dT=0.0644·10^{-3} V/°C)	700	0.05976	36Sel
Pb | $PbCl_2$, KCl, LiCl | $PbBi_x$, x=0.889 (dU/dT=0.1020·10^{-3} V/°C)	700	0.08615	36Sel
Pb | $PbCl_2$, KCl, LiCl | Pb_xSb, x=0.1095 (dU/dT=0.0970·10^{-3} V/°C, 600<T<630 °C)	475	0.07956	39Sel
Pb | $PbCl_2$, KCl, LiCl | Pb_xSb, x=0.1992 (dU/dT=0.0711·10^{-3} V/°C, 560<T<610 °C)	475	0.05848	39Sel
Pb | $PbCl_2$, KCl, LiCl | Pb_xSb, x=0.3057 (dU/dT=0.0554·10^{-3} V/°C, 550<T<605 °C)	475	0.04068	39Sel
Pb | $PbCl_2$, KCl, LiCl | Pb_xSb, x=0.4012 (dU/dT=0.0434·10^{-3} V/°C, 480<T<520 °C)	475	0.03225	39Sel
Pb | $PbCl_2$, KCl, LiCl | Pb_xSb, x=0.4997 (dU/dT=0.0344·10^{-3} V/°C, 420<T<470 °C)	475	0.02452	39Sel
Pb | $PbCl_2$, KCl, LiCl | Pb_xSb, x=0.6145 (dU/dT=0.0234·10^{-3} V/°C, 390<T<490 °C)	475	0.01682	39Sel
Pb | $PbCl_2$, KCl, LiCl | Pb_xSb, x=0.7024 (dU/dT=0.01725·10^{-3} V/°C, 385<T<490 °C)	475	0.01194	39Sel
Pb | $PbCl_2$, KCl, LiCl | Pb_xSb, x=0.7926 (dU/dT=0.01122·10^{-3} V/°C, 390<T<490 °C)	475	0.00767	39Sel
Pb | $PbCl_2$, KCl, LiCl | Pb_xSb, x=0.8710 (dU/dT=0.00689·10^{-3} V/°C, 370<T<460 °C)	475	0.00463	39Sel
Magnesium (Mg)			
Mg | $MgCl_2$ | Cl_2,C	704	2.5120	29Lor
	750	2.4730	29Lor
	790	2.4460	29Lor
Mg | $MgCl_2$ | Cl_2, C (U [V]=2.544–0.66·10^{-3} (T–718 °C), 660<T<725 °C)	718	2.544	55Mar2
Mg | $MgCl_2$, x=0.20, KCl | Cl_2, C (dU/dT= –0.41 mV/°C)	718	2.852	55Mar2
Mg | $MgCl_2$, x=0.30, KCl | Cl_2, C (dU/dT= –0.48 mV/°C)	718	2.766	55Mar2
Mg | $MgCl_2$, x=0.33, KCl | Cl_2, C (dU/dT= –0.50 mV/°C)	718	2.742	55Mar2
Mg | $MgCl_2$, x=0.40, KCl | Cl_2, C (dU/dT= –0.53 mV/°C)	718	2.688	55Mar2
Mg | $MgCl_2$, x=0.50, KCl | Cl_2, C (dU/dT= –0.58 mV/°C)	718	2.612	55Mar2
Mg | $MgCl_2$, x=0.60, KCl | Cl_2, C (dU/dT= –0.62 mV/°C)	718	2.570	55Mar2
Mg | $MgCl_2$, x=0.70, KCl | Cl_2, C (dU/dT= –0.63 mV/°C)	718	2.559	55Mar2
Mg | $MgCl_2$, x=0.04, LiCl | Cl_2, C (dU/dT= –0.45 mV/°C)	718	2.710	55Mar2
Mg | $MgCl_2$, x=0.10, LiCl | Cl_2, C (dU/dT= –0.51 mV/°C)	718	2.652	55Mar2
Mg | $MgCl_2$, x=0.20, LiCl | Cl_2, C (dU/dT= –0.49 mV/°C)	718	2.616	55Mar2
Mg | $MgCl_2$, x=0.30, LiCl | Cl_2, C (dU/dT= –0.52 mV/°C)	718	2.588	55Mar2
Mg | $MgCl_2$, x=0.33, LiCl | Cl_2, C (dU/dT= –0.50 mV/°C)	718	2.581	55Mar2

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Mg | $MgCl_2$, x=0.4, LiCl | Cl_2, C (dU/dT= –0.51 mV/°C)	718	2.567	55Mar2
Mg | $MgCl_2$, x=0.5, LiCl | Cl_2, C (dU/dT= –0.54 mV/°C)	718	2.556	55Mar2
Mg | $MgCl_2$, x=0.6, LiCl | Cl_2, C (dU/dT= –0.54 mV/°C)	718	2.550	55Mar2
Mg | $MgCl_2$, x=0.7, LiCl | Cl_2, C (dU/dT= –0.56 mV/°C)	718	2.547	55Mar2
Mg | $MgCl_2$, x=1.0, LiCl | Cl_2, C (dU/dT= –0.66 mV/°C)	718	2.544	55Mar2
Mg | $MgCl_2$, x=0.222, NaCl | Cl_2, C (dU/dT= –0.535 mV/°C)	718	2.692	55Mar2
Mg | $MgCl_2$, x=0.242, NaCl | Cl_2, C (dU/dT= –0.584 mV/°C)	718	2.682	55Mar2
Mg | $MgCl_2$, x=0.323, NaCl | Cl_2, C (dU/dT= –0.545 mV/°C)	718	2.641	55Mar2
Mg | $MgCl_2$, x=0.403, NaCl | Cl_2, C (dU/dT= –0.60 mV/°C)	718	2.613	55Mar2
Mg | $MgCl_2$, x=0.559, NaCl | Cl_2, C (dU/dT= –0.524 mV/°C)	718	2.571	55Mar2
Mg | $MgCl_2$, x=0.343, RbCl | Cl_2, C (dU/dT= –0.49 mV/°C)	718	2.795	55Mar2
Mg | $MgCl_2$, x=0.41, RbCl | Cl_2, C (dU/dT= –0.50 mV/°C)	718	2.718	55Mar2
Mg | $MgCl_2$, x=0.477, RbCl | Cl_2, C (dU/dT= –0.52 mV/°C)	718	2.669	55Mar2
Mg | $MgCl_2$, x=0.59, RbCl | Cl_2, C (dU/dT= –0.61 mV/°C)	718	2.590	55Mar2
Mg_xPb, x=at.frac.=1 | $MgCl_2$ | Mg_xPb, x=at.frac.=0.617	680	0.050	32Wag
Mg_xPb, x=at.frac.=1 | $MgCl_2$ | Mg_xPb, x=at.frac.=0.709	680	0.031	32Wag
Mg_xPb, x=at.frac.=1 | $MgCl_2$ | Mg_xPb, x=at.frac.=0.895	680	0.010	32Wag
Mg_xPb, x=at.frac.=0.697 | $MgCl_2$ | Mg_xPb, x=at.frac.=0.288	680	0.0932	32Wag
Mg_xPb, x=at.frac.=0.697 | $MgCl_2$ | Mg_xPb, x=at.frac.=0.358	680	0.0784	32Wag
Mg_xPb, x=at.frac.=0.697 | $MgCl_2$ | Mg_xPb, x=at.frac.=0.479	680	0.0551	32Wag
MgPb, x=at.frac.=0.697 | $MgCl_2$ | Mg_xPb, x=at.frac.=0.675	680	0.0125	32Wag
Mg_xPb, x=at.frac.=0.261 | $MgCl_2$ | Mg_xPb, x=at.frac.=0.059	680	0.0655	32Wag
Mg_xPb, x=at.frac.=0.261 | $MgCl_2$ | Mg_xPb, x=at.frac.=0.114	680	0.0400	32Wag
Mg_xPb, x=at.frac.=0.261 | $MgCl_2$ | Mg_xPb, x=at.frac.=0.174	680	0.0229	32Wag
Mg_xPb, x=at.frac.=0.261 | $MgCl_2$ | Mg_xPb, x=at.frac.=0.223	680	0.0134	32Wag
Silver (Ag)			
Ag | AgBr | Br_2, C	450	0.8010	32Sal1
	500	0.7865	32Sal1
	550	0.7720	32Sal1
Ag | AgBr, x=0.3538, KBr | Br_2, C	500	0.8815	31Sal2
	550	0.8727	31Sal2
	600	0.8640	31Sal2
Ag | AgBr, x=0.4502, KBr | Br_2, C	500	0.8588	31Sal2
	550	0.8483	31Sal2
	600	0.8374	31Sal2
Ag | AgBr, x=0.6005, KBr | Br_2, C	500	0.8297	31Sal2
	550	0.8177	31Sal2
	600	0.8057	31Sal2
Ag | AgBr, x=0.800, KBr | Br_2, C	500	0.8016	31Sal2
	550	0.7884	31Sal2
	600	0.7753	31Sal2
Ag | AgBr, x=1.0, KBr | Br_2, C	500	0.7863	31Sal2
	550	0.7720	31Sal2
	600	0.7577	31Sal2
Ag | AgBr, x=0.1100, LiBr | Br_2, C	500	0.8686	30Sal
	550	0.8648	30Sal
	600	0.8610	30Sal

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Ag | AgBr, x=0.2548, LiBr | Br_2, C	500	0.8301	30Sal
	550	0.8216	30Sal
	600	0.8128	30Sal
Ag | AgBr, x=0.4086, LiBr | Br_2, C	500	0.8202	30Sal
	550	0.8100	30Sal
	600	0.7993	30Sal
Ag | AgBr, x=0.5937, LiBr | Br_2, C	500	0.8085	30Sal
	550	0.7961	30Sal
	600	0.7836	30Sal
Ag | AgBr, x=1.0, LiBr | Br_2, C	500	0.7865	30Sal
	550	0.7720	30Sal
	600	0.7577	30Sal
Ag | AgBr, x=0.5130, NaBr | Br_2, C	550	0.8088	31Sal1
	600	0.7979	31Sal1
Ag | AgBr, x=0.6045, NaBr | Br_2, C	550	0.8002	31Sal1
	600	0.7883	31Sal1
Ag | AgBr, x=0.7486, NaBr | Br_2, C	550	0.7877	31Sal1
	600	0.7749	31Sal1
Ag | AgBr, x=1.0, NaBr | Br_2, C	550	0.7720	31Sal1
	600	0.7577	31Sal1
Ag | AgBr, x=0.00, $PbBr_2$ | Br_2, C	450	1.0624	32Sal1
	500	1.0321	32Sal1
	550	1.0017	32Sal1
Ag | AgBr, x=0.10, $PbBr_2$ | Br_2, C	450	0.9407	32Sal1
	500	0.9352	32Sal1
	550	0.9295	32Sal1
Ag | AgBr, x=0.20, $PbBr_2$ | Br_2, C	450	0.8976	32Sal1
	500	0.8897	32Sal1
	550	0.8818	32Sal1
Ag | AgBr, x=0.30, $PbBr_2$ | Br_2, C	450	0.8733	32Sal1
	500	0.8637	32Sal1
	550	0.8542	32Sal1
Ag | AgBr, x=0.40, $PbBr_2$ | Br_2, C	450	0.8580	32Sal1
	500	0.8470	32Sal1
	550	0.8358	32Sal1
Ag | AgBr, x=0.50, $PbBr_2$ | Br_2, C	450	0.8446	32Sal1
	500	0.8328	32Sal1
	550	0.8212	32Sal1
Ag | AgBr, x=0.60, $PbBr_2$ | Br_2, C	450	0.8336	32Sal1
	500	0.8207	32Sal1
	550	0.8078	32Sal1
Ag | AgBr, x=0.70, $PbBr_2$ | Br_2, C	450	0.8238	32Sal1
	500	0.8106	32Sal1
	550	0.7972	32Sal1
Ag | AgBr, x=0.80, $PbBr_2$ | Br_2, C	450	0.8149	32Sal1
	500	0.8015	32Sal1
	550	0.7882	32Sal1

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Ag | AgBr, x=1.0, $PbBr_2$ | Br_2, C	450	0.8010	32Sal1
	500	0.7865	32Sal1
	550	0.7720	32Sal1
Ag | AgBr, x=0.352, RbBr | Br_2, C	500	0.9040	32Sal2
	550	0.8937	32Sal2
	600	0.8833	32Sal2
Ag | AgBr, x=0.467, RbBr | Br_2, C	500	0.8697	32Sal2
	550	0.8582	32Sal2
	600	0.8468	32Sal2
Ag | AgBr, x=0.569, RbBr | Br_2, C	500	0.8398	32Sal2
	550	0.8275	32Sal2
	600	0.8151	32Sal2
Ag | AgBr, x=0.747, RbBr | Br_2, C	500	0.8116	32Sal2
	550	0.7986	32Sal2
	600	0.7856	32Sal2
Ag | AgBr, x=1.0, RbBr | Br_2, C	500	0.7865	32Sal2
	550	0.7720	32Sal2
	600	0.7577	32Sal2
Ag | AgCl | Cl_2, C	450	0.9145	33Sal
	500	0.9001	33Sal
	550	0.8855	33Sal
	600	0.8709	33Sal
Ag | AgCl, x=0.136, LiCl | Cl_2, C	500	0.9626	36Sal
	600	0.9497	36Sal
Ag | AgCl, x=0.252, LiCl | Cl_2, C	500	0.9424	36Sal
	600	0.9260	36Sal
Ag | AgCl, x=0.469, LiCl | Cl_2, C	500	0.9249	36Sal
	600	0.9038	36Sal
Ag | AgCl, x=0.573, LiCl | Cl_2, C	500	0.9199	36Sal
	600	0.98965	36Sal
Ag | AgCl, x=0.690, LiCl | Cl_2, C	500	0.9156	36Sal
	600	0.8901	36Sal
Ag | AgCl, x=0.804, LiCl | Cl_2, C	500	0.9085	36Sal
	600	0.8816	36Sal
Ag | AgCl, x=1.0, LiCl | Cl_2, C	500	0.9001	36Sal
	600	0.8709	36Sal
Ag | AgCl, x=0.10, $PbCl_2$ | Cl_2, C	500	1.0518	34Sal
	600	1.0424	34Sal
Ag | AgCl, x=0.20, $PbCl_2$ | Cl_2, C	500	1.0086	34Sal
	600	0.9930	34Sal
Ag | AgCl, x=0.30, $PbCl_2$ | Cl_2, C	500	0.9817	34Sal
	600	0.9631	34Sal
Ag | AgCl, x=0.45, $PbCl_2$ | Cl_2, C	500	0.9549	34Sal
	600	0.9329	34Sal
Ag | AgCl, x=0.60, $PbCl_2$ | Cl_2, C	500	0.9352	34Sal
	600	0.9107	34Sal
Ag | AgCl, x=0.80, $PbCl_2$ | Cl_2, C	500	0.9148	34Sal
	600	0.8878	34Sal

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Ag | AgCl, x=1.0, $PbCl_2$ | Cl_2, C	500	0.9001	34Sal
	600	0.8709	34Sal
Ag | AgCl, KCl | Ag_xAu, x=0.263	1087	0.198	32Wag
Ag | AgCl, KCl | Ag_xAu, x=0.471	1087	0.114	32Wag
Ag | AgCl, KCl | Ag_xAu, x=0.602	1087	0.074	32Wag
Ag | AgCl, KCl | Ag_xAu, x=0.716	1087	0.045	32Wag
Ag | AgCl, KCl | Ag_xAu, x=0.84	1087	0.021	32Wag
Ag | AgCl, KCl | Ag_xAu, x=0.030	729<T<886	5.6–0.023T [mV]	32Wag
Ag | AgCl, KCl | Ag_xAu, x=0.0910	720<T<847	18.02–0.0065T [mV]	32Wag
Ag | AgCl, KCl | Ag_xAu, x=0.250	720<T<847	17.2+0.0243T [mV]	32Wag
Ag | AgCl, KCl | Ag_xAu, x=0.400	720<T<847	47.9+0.0271T [mV]	32Wag
Ag | AgCl, KCl | Ag_xAu, x=0.500	759<T<873	69.7+0.0452T [mV]	32Wag
Ag | AgCl, KCl | Ag_xAu, x=0.750	759<T<873	131.2+0.0467T [mV]	32Wag
Thallium (Tl)			
Tl | TlCl, LiCl, KCl | Tl_xAu, x=0.382 (dU/dT=0.120 mV/°C)	600	0.1042	51Kle
Tl | TlCl, LiCl, KCl | Tl_xAu, x=0.485 (dU/dT=0.0950 mV/°C)	600	0.04880	51Kle
Tl | TlCl, LiCl, KCl | Tl_xAu, x=0.587 (dU/dT=0.0640 mV/°C)	600	0.05460	51Kle
Tl | TlCl, LiCl, KCl | Tl_xAu, x=0.681 (dU/dT=0.0350 mV/°C)	600	0.03785	51Kle
Tl | TlCl, LiCl, KCl | Tl_xAu, x=0.794 (dU/dT=0.0245 mV/°C)	600	0.02050	51Kle
Tl | TlCl, LiCl, KCl | Tl_xAu, x=0.895 (dU/dT=0.0117 mV/°C)	600	0.00915	51Kle
Tl | TlCl, $MgCl_2$, KCl | Tl_xBi, x=0.205	481	0.2240	29Hil
Tl | TlCl, $MgCl_2$, KCl | Tl_xBi, x=0.460	481	0.1261	29Hil
Tl | TlCl, $MgCl_2$, KCl | Tl_xBi, x=0.585	481	0.0900	29Hil
Tl | TlCl, $MgCl_2$, KCl | Tl_xBi, x=0.720	481	0.0502	29Hil
Tl | TlCl, $MgCl_2$, KCl | Tl_xBi, x=0.826	481	0.0194	29Hil
Tl | TlCl, CH_3COOK, CH_3COONa | Tl_xBi, x=0.086	271<T<274	0.219	32Wag
Tl | TlCl, CH_3COOK, CH_3COONa | Tl_xBi, x=0.154	271<T<274	0.176	32Wag
Tl | TlCl, CH_3COOK, CH_3COONa | Tl_xBi, x=0.239	271<T<274	0.147	32Wag
Tl | TlCl, CH_3COOK, CH_3COONa | Tl_xBi, x=0.352	271<T<274	0.117	32Wag
Tl | TlCl, CH_3COOK, CH_3COONa | Tl_xBi, x=0.733	271<T<274	0.028	32Wag
Tl | TlCl, LiCl, KCl | Tl_xPb, x=0.0454	438	0.1840	29Hil
	500	0.2020	29Hil
	563	0.2122	29Hil
Tl | TlCl, LiCl, KCl | Tl_xPb, x=0.0812	438	0.1660	29Hil
	500	0.1777	29Hil
	563	0.1885	29Hil
Tl | TlCl, LiCl, KCl | Tl_xPb, x=0.1316	438	0.1413	29Hil
	500	0.1511	29Hil
	563	0.1597	29Hil

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Tl | TlCl, LiCl, KCl | Tl_xPb, x=0.3079	438	0.0865	29Hil
Tl | TlCl, LiCl, KCl | Tl_xPb, x=0.6087	438	0.0381	29Hil
Tl | TlCl, LiCl, KCl | Tl_xPb, x=0.7653	438	0.0180	29Hil
	500	0.0183	29Hil
	563	0.0192	29Hil
Tl | TlCl, LiCl, KCl | Tl_xPb, x=0.8848	438	0.0048	29Hil
Tl | TlCl, LiCl, KCl | Tl_xSn, x=0.0794	353	0.0893	29Hil
	414	0.1047	29Hil
	478	0.1217	29Hil
Tl | TlCl, LiCl, KCl | Tl_xSn, x=0.1242	414	0.0867	29Hil
	478	0.0979	29Hil
Tl | TlCl, LiCl, KCl | Tl_xSn, x=0.1502	353	0.0627	29Hil
	414	0.0756	29Hil
	478	0.0884	29Hil
Tl | TlCl, LiCl, KCl | Tl_xSn, x=0.2107	414	0.0629	29Hil
	478	0.0712	29Hil
Tl | TlCl, LiCl, KCl | Tl_xSn, x=0.3407	414	0.0422	29Hil
	478	0.0507	29Hil
Tl | TlCl, LiCl, KCl | Tl_xSn, x=0.3944	353	0.0301	29Hil
	414	0.0374	29Hil
Tl | TlCl, LiCl, KCl | Tl_xSn, x=0.5781	353	0.0172	29Hil
	414	0.0225	29Hil
Tl | TlCl, LiCl, KCl | Tl_xSn, x=0.7128	414	0.0157	29Hil
	478	0.0194	29Hil
Thorium (Th)			
Th | $ThCl_4$, x=0.00825, NaCl, KCl(equimol.) | Cl_2, C	730	2.64	59Dro
Th | $ThCl_4$, x=0.00789, NaCl, KCl(equimol.) | Cl_2, C	730	2.61	59Dro
Th | $ThCl_4$, x=0.0171, NaCl, KCl(equimol.) | Cl_2, C	730	2.58	59Dro
Th | $KThF_5$, x=0.00623, NaCl, KCl(equimol.) | Cl_2, C	743	2.67	59Dro
Th | $KThF_5$, x=0.00868, NaCl, KCl(equimol.) | Cl_2, C	743	2.63	59Dro
Th | $KThF_5$, 0.8 wt%, $SrCl_2$, KCl(equimol.) | Cl_2, C	747	2.49	59Dro
Th | $KThF_5$, 2.0 wt%, $SrCl_2$, KCl(equimol.) | Cl_2, C	747	2.47	59Dro
Tin (Sn)			
Sn | $SnCl_2$, KCl, LiCl | Sn_xAu, x=0.2589	500	0.1898	50Kle
Sn | $SnCl_2$, KCl, LiCl | Sn_xAu, x=0.3596	500	0.1317	50Kle
Sn | $SnCl_2$, KCl, LiCl | Sn_xAu, x=0.5091	500	0.0690	50Kle
Sn | $SnCl_2$, KCl, LiCl | Sn_xAu, x=0.6815	500	0.02765	50Kle
Sn | $SnCl_2$, KCl, LiCl | Sn_xAu, x=0.8125	500	0.01116	50Kle
Sn | $SnCl_2$ | Sn_xBi, x=0.046	329<T<330	0.0762	32Wag
Sn | $SnCl_2$ | Sn_xBi, x=0.095	329<T<330	0.0585	32Wag
Sn | $SnCl_2$ | Sn_xBi, x=0.143	329<T<330	0.0479	32Wag
Sn | $SnCl_2$ | Sn_xBi, x=0.189	329<T<330	0.0406	32Wag
Sn | $SnCl_2$ | Sn_xBi, x=0.236	329<T<330	0.0351	32Wag

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Sn | $SnCl_2$ | Sn_xBi, x=0.292	329<T<330	0.0302	32Wag
Sn | $SnCl_2$ | Sn_xBi, x=0.367	329<T<330	0.0248	32Wag
Sn | $SnCl_2$ | Sn_xBi, x=0.453	329<T<330	0.0198	32Wag
Sn | $SnCl_2$ | Sn_xBi, x=0.549	329<T<330	0.0151	32Wag
Zinc (Zn)			
Zn | $ZnBr_2$ | Br_2, C	400	1.3382	33Sal, 30Wac
	450	1.3041	33Sal, 30Wac
	500	1.270	33Sal, 30Wac
Zn | $ZnCl_2$ | Cl_2, C	417.8	1.611	30Wac
	450	1.607	30Wac
	500	1.572	30Wac
	550	1.537	30Wac
	600	1.503	30Wac
	650	1.468	30Wac
	700	1.433	30Wac
Zn | $ZnCl_2$, x=0.4, NaCl | Cl_2, C (dU/dT= –0.561 mV/°C)	500	1.7003	54Lan
Zn | $ZnCl_2$, x=0.5, NaCl | Cl_2, C (dU/dT= –0.572 mV/°C)	500	1.6398	54Lan
Zn | $ZnCl_2$, x=0.6, NaCl | Cl_2, C (dU/dT= –0.644 mV/°C)	500	1.6032	54Lan
Zn | $ZnCl_2$, x=0.7, NaCl | Cl_2, C (dU/dT= –0.710 mV/°C)	500	1.5853	54Lan
Zn | $ZnCl_2$, x=0.8, NaCl | Cl_2, C (dU/dT= –0.707 mV/°C)	500	1.5751	54Lan
Zn | $ZnCl_2$, x=0.9, NaCl | Cl_2, C (dU/dT= –0.714 mV/°C)	500	1.5721	54Lan
Zn | $ZnCl_2$, x=0.5, NaCl | Cl_2, C (dU/dT= –0.624 mV/°C)	500	1.6163	54Lan
Zn | $ZnCl_2$, x=0.6, NaCl | Cl_2, C (dU/dT= –0.683 mV/°C)	500	1.5981	54Lan
Zn | $ZnCl_2$, x=0.7, NaCl | Cl_2, C (dU/dT= –0.700 mV/°C)	500	1.5787	54Lan
Zn | $ZnCl_2$, x=0.8, NaCl | Cl_2, C (dU/dT= –0.700 mV/°C)	500	1.5723	54Lan
Zn | $ZnCl_2$, x=0.9, NaCl | Cl_2, C (dU/dT= –0.708 mV/°C)	500	1.5695	54Lan
Zn | $ZnCl_2$, x=1.0, NaCl | Cl_2, C (dU/dT= –0.697 mV/°C)	500	1.5635	54Lan
Zn | LiCl, KCl, $ZnCl_2$ | Zn_xCd, x=0.150	436	0.0275	23Tay
	572	0.0396	23Tay
Zn | LiCl, KCl, $ZnCl_2$ | Zn_xCd, x=0.344	436	0.0136	23Tay
	572	0.0200	23Tay
Zn | LiCl, KCl, $ZnCl_2$ | Zn_xCd, x=0.536	436	0.0083	23Tay
	572	0.0126	23Tay
Zn | LiCl, KCl, $ZnCl_2$ | Zn_xCd, x=0.690	436	0.0058	23Tay
	572	0.0080	23Tay
Zn | $ZnCl_2$, KCl, LiCl | Zn_xGa, x=0.042 (dU/dT=0.1574·10^{-3} V/°C)	488.2	0.0932	55Gen
Zn | $ZnCl_2$, KCl, LiCl | Zn_xGa, x=0.150 (dU/dT=0.1001·10^{-3} V/°C)	488.2	0.0530	55Gen
Zn | $ZnCl_2$, KCl, LiCl | Zn_xGa, x=0.260 (dU/dT=0.0720·10^{-3} V/°C)	488.2	0.0367	55Gen
Zn | $ZnCl_2$, KCl, LiCl | Zn_xGa, x=0.370 (dU/dT=0.0565·10^{-3} V/°C)	488.2	0.0268	55Gen
Zn | $ZnCl_2$, KCl, LiCl | Zn_xGa, x=0.480 (dU/dT=0.0500·10^{-3} V/°C)	488.2	0.0200	55Gen
Zn | $ZnCl_2$, KCl, LiCl | Zn_xGa, x=0.590 (dU/dT=0.0363·10^{-3} V/°C)	488.2	0.0146	55Gen

Cell composition [2])	T [°C]	U_0 [V]	Ref.
Zn | $ZnCl_2$, KCl, LiCl | Zn_xGa, x=0.697 (dU/dT=0.0264·10^{-3} V/°C)	488.2	0.0103	55Gen
Zn | $ZnCl_2$, KCl, LiCl | Zn_xGa, x=0.775 (dU/dT=0.0167·10^{-3} V/°C)	488.2	0.0077	55Gen
Zn | $ZnCl_2$, KCl, LiCl | Zn_xGa, x=0.841 (dU/dT=0.0082·10^{-3} V/°C)	488.2	0.0051	55Gen
Zn | $ZnCl_2$, KCl, LiCl | Zn_xGa, x=0.910 (dU/dT=0.0054·10^{-3} V/°C)	488.2	0.0030	55Gen
Zn | $ZnCl_2$, KCl, LiCl | Zn_xGa, x=0.957 (dU/dT=0.0033·10^{-3} V/°C)	488.2	0.00145	55Gen
Zn | $ZnCl_2$, 5 wt% in KCl+LiCl eutectic | Zn_xSn, x=0.097 (dU/dT=0.1361·10^{-3} V/°C)	476.8	0.0554	55Fio
Zn | $ZnCl_2$, 5 wt% in KCl+LiCl eutectic | Zn_xSn, x=0.161 (dU/dT=0.1096·10^{-3} V/°C)	476.8	0.0406	55Fio
Zn | $ZnCl_2$, 5 wt% in KCl+LiCl eutectic | Zn_xSn, x=0.255 (dU/dT=0.0862·10^{-3} V/°C)	476.8	0.0278	55Fio
Zn | $ZnCl_2$, 5 wt% in KCl+LiCl eutectic | Zn_xSn, x=0.361 (dU/dT=0.0665·10^{-3} V/°C)	476.8	0.0191	55Fio
Zn | $ZnCl_2$, 5 wt% in KCl+LiCl eutectic | Zn_xSn, x=0.463 (dU/dT=0.0498·10^{-3} V/°C)	476.8	0.0130	55Fio
Zn | $ZnCl_2$, 5 wt% in KCl+LiCl eutectic | Zn_xSn, x=0.579 (dU/dT=0.0371·10^{-3} V/°C)	476.8	0.00857	55Fio
Zn | $ZnCl_2$, 5 wt% in KCl+LiCl eutectic | Zn_xSn, x=0.687 (dU/dT=0.0255·10^{-3} V/°C)	476.8	0.00571	55Fio
Zn | $ZnCl_2$, 5 wt% in KCl+LiCl eutectic | Zn_xSn, x=0.767 (dU/dT=0.0180·10^{-3} V/°C)	476.8	0.00417	55Fio
Zn | $ZnCl_2$, 5 wt% in KCl+LiCl eutectic | Zn_xSn, x=0.840 (dU/dT=0.0125·10^{-3} V/°C)	476.8	0.00355	55Fio
Zn | $ZnCl_2$, 5 wt% in KCl+LiCl eutectic | Zn_xSn, x=0.905 (dU/dT=0.0073·10^{-3} V/°C)	476.8	0.00249	55Fio
Zn | $ZnCl_2$, KCl, LiCl | Zn_xBi, x=0.0639	600	0.0666	52Kle
Zn | $ZnCl_2$, KCl, LiCl | Zn_xBi, x=0.1284	600	0.0428	52Kle
Zn | $ZnCl_2$, KCl, LiCl | Zn_xBi, x=0.5523	600	0.00312	52Kle
Zn | $ZnCl_2$, KCl, LiCl | Zn_xBi, x=0.7642	600	0.00101	52Kle
Zn | $ZnCl_2$, KCl, LiCl | Zn_xBi, x=0.9741	600	0.00055	52Kle
Zn | $ZnCl_2$, KCl, LiCl | Zn_xPb, x=0.0388	653	0.0487	52Kle
Zn | $ZnCl_2$, KCl, LiCl | Zn_xPb, x=0.1229	653	0.0116	52Kle
Zn | $ZnCl_2$, KCl, LiCl | Zn_xPb, x=0.2043	653	0.00126	52Kle
Zn | $ZnCl_2$, KCl, LiCl | Zn_xPb, x=0.5058	653	0.00090	52Kle
Zn | $ZnCl_2$, KCl, LiCl | Zn_xPb, x=0.8986	653	0.00080	52Kle
Zn | $ZnCl_2$, KCl, LiCl | Zn_xPb, x=0.9894	653	0.00048	52Kle
Zn | $ZnCl_2$, KCl, LiCl | Zn_xSn, x=0.221	431	0.0267	23Tay
	572	0.0401	23Tay
Zn | $ZnCl_2$, KCl, LiCl | Zn_xSn, x=0.384	431	0.0138	23Tay
	572	0.0226	23Tay
Zn | $ZnCl_2$, KCl, LiCl | Zn_xSn, x=0.583	431	0.0065	23Tay
	572	0.0106	23Tay
Zn | $ZnCl_2$, KCl, LiCl | Zn_xSn, x=0.845	431	0.0024	23Tay
	572	0.0038	23Tay
Zn | $ZnCl_2$, KCl, LiCl, NaCl | $Zn_{10at\%}Sn_{90at\%}Sb_{0at\%}$	600	0.0805	54Ala
Zn | $ZnCl_2$, KCl, LiCl, NaCl | $Zn_{8.33at\%}Sn_{75at\%}Sb_{16.67at\%}$	600	0.0903	54Ala
Zn | $ZnCl_2$, KCl, LiCl, NaCl | $Zn_{6.25at\%}Sn_{56.25at\%}Sb_{37.5at\%}$	600	0.1218	54Ala
Zn | $ZnCl_2$, KCl, LiCl, NaCl | $Zn_{5at\%}Sn_{45at\%}Sb_{50at\%}$	600	0.1424	54Ala

[1]) Description of the cell composition is based - as far as possible - on the Stockholm convention (1953), i.e. the left-hand electrode constitutes the negative terminal of the cell. Cells are listed according to the metallic constituent of the electrode mentioned first which is involved in the electrode reaction establishing the respective electrode potential. Contact materials and conductive additives may be mentioned first before the actual element of interest only for the sake of correct materials sequence. The sequence of electrode components is stated as reported in the original publications.

[2]) Carbon electrodes as employed with e.g. halogens are designated as C only irrespective of the actual composition.

[3]) Average value of the quoted references.

References

23Tay Taylor, N.W.: J. Am. Chem. Soc. **45** (1923) 2865.
29Hil Hildebrand, J.H., Sharma, J.N.: J. Am. Chem. Soc. **51** (1929) 462.
29Lor Lorenz, R., Velde H.: Z. Anorg. Allg. Chem. **183** (1929) 81.
30Sal Salstrom, E.: J. Am. Chem. Soc. **52** (1930) 4641; 4650.
30Wac Wachter, A., Hildebrand, J.H.: J. Am. Chem. Soc. **52** (1930) 4655.
31Sal1 Salstrom, E.: J. Am. Chem. Soc. **53** (1931) 1794.
31Sal2 Salstrom, E.: J. Am. Chem. Soc. **53** (1931) 3388.
32Sal1 Salstrom, E.: J. Am. Chem. Soc. **54** (1932) 2653.
32Sal2 Salstrom, E.: J. Am. Chem. Soc. **54** (1932) 4255.
32Wag Wagner, C., Engelhardt, G.: Z. Phys. Chem. Abt. A **159** (1932) 241.
32Öla Ölander, A.: J. Am. Chem. Soc. **54** (1932) 3819.
33Sal Salstrom, E.: J. Am. Chem. Soc. **55** (1933) 1029; 2426.
33Tre Treadwell, W.D., Terebesi, L.: Helv. Chim. Acta **16** (1933) 931.
34Sal Salstrom, E.: J. Am. Chem. Soc. **56** (1934) 1272.
35Hol Holub, L., Neubert, F., Sauerwald, F.: Z. Phys. Chem. Abt. A **174** (1935) 161.
36Sal Salstrom, E., Kew, E.T., Powell, T.M.: J. Am. Chem. Soc. **58** (1936) 1848.
36Sel Seltz, H., Strickler, H.S.: J. Am. Chem. Soc. **58** (1936) 2084.
38Sel Seltz, H., DeWitt, J.: J. Am. Chem. Soc. **60** (1938) 1305.
39Sel Seltz, H., DeWitt, J.: J. Am. Chem. Soc. **61** (1939) 2594.
49Kle Kleppa, O.J.: J. Am. Chem. Soc. **71** (1949) 3275.
50Kle Kleppa, O.J.: J. Am. Chem. Soc. **72** (1950) 3346.
51Ell Elliot, J.F., Chipman, J.: J. Am. Chem. Soc. **73** (1951) 2682.
51Kle Kleppa, O.J.: J. Am. Chem. Soc. **73** (1951) 385.
52Kle Kleppa, O.J.: J. Am. Chem. Soc. **74** (1952) 6052.
53Lan1 Lantratov, M.F., Alabyshev, A.F.: Zh. Prikl. Khim. **26** (1953) 321; 353.
53Lan2 Lantratov, M.F., Alabyshev, A.F.: Zh. Prikl. Khim. **26** (1953) 263.
54Ala Alabyshev, A.F., Lantratov, M.F.: Zh. Prikl. Khim. **27** (1954) 851.
54Lan Lantratov, M.F., Alabyshev, A.F.: Zh. Prikl. Khim. **27** (1954) 722.
54Ric Riccoboni, L., Genta, V., Fiorani, M., Valenti, V.: Gazz. Chim. Ital. **84** (1954) 982.
55Fio Fiorani, M., Valenti, V.: Gazz. Chim. Ital. **85** (1955) 607.
55Gen Genta, V., Fiorini, M., Valenti, V.: Gazz. Chim. Ital. **85** (1955) 103.
55Mar2 Markov, B.F., Delimarskii, Yu.K., Panchenko, I.D.: Zh. Fiz. Khim. **29** (1955) 51.
59Dro Drossbach, P., Neumayr, F.: Z. Elektrochem. **63** (1959) 516.

Table 2.4. Cell voltages with solid state electrolyte systems. [1])

Cell composition	T [K, °C]	U_0 [V]	Ref.
Aluminum (Al)			
Al | Al_2O_3 | α- and β-alumina | NbO | Nb	850 K	0.6806	96Sub
(U_T±0.00007 [V]=0.8209–0.00016400 T, 795<T<954 K)			
Pt, Al, Al_2O_3 | α- and β-alumina | NbO, Ta, Pt	840 K	0.6822	97Sub2
(U_T±0.00007 [V]=0.820–0.000164 T, 820<T<860 K)			
Pt, O_2, Al_2O_3, $NaAl_{11}O_{17}$ | Na-β"-alumina | $Na_2Mo_2O_7$, Na_2MoO_4,O_2, Pt	800 K	0.2977	02Sub
(U_T±0.004 [V]=–0.00999+0.0003847 T, 740<T<800 K)			
Antimony (Sb)			
Sb | CoO | $CoSb_2O_4$ | NiO | Ni	900 K	0.07642	96Swa1
(U_T±0.00048 [V]=0.1207–0.00004920 T, 795<T<954 K)			
Chromel, Mo, Sb, CoO, $CoSb_2O_4$ | 15CSZ | NiO, Ni, Mo, Chromel [6])	900 K	0.07642	01Swa
(U_T±0.0005 [V]=0.1207–0.00004920 T, 795<T<954 K)			
Sb | Fe_3O_4 | $FeSb_2$ | NiO | Ni	800 K	0.37226	92Pan
(U_T±0.00028 [V]=–0.1090+0.00032907 T, 758<T<901 K)			
Sb | NiO | $NiSb_2O_4$ | NiO | Ni	800 K	0.00486	96Swa2
(U_T±0.00029 [V]=0.0883–0.00010430 T, 771<T<931 K)			
Sb | Sb_2O_3 | O_2(air)	800 K	0.8182	86Aza
(U_T±0.00242 [V]=1.1774–0.00044901 T, 629<T<895 K)			
Sb(liquid) | Sb_2O_3 | O_2(air)	1000 K	0.733	86Aza
(U_T±0.00112 [V]=1.1459–0.00041290 T, 948<T<1051 K)			
Pt, Sb_2O_3(ortho), Sb_2O_4(ortho), Pt black | 15YSZ | O_2(air), Pt [5])	600 K	0.57165	87Pan
(U_T±0.00042 [V]=0.8723–0.00050108 T, 585<T<828 K)			
Barium (Ba)			
Pt, O_2, $BaCuO_2$, CuO, BaF_2 | CaF_2 | BaF_2, $BaZrO_3$, ZrO_2, O_2, Pt	900 K	0.1454	91Aza2
(U_T±0.0016 [V]=0.3674–0.000261199 T, 736<T<1102 K)			
$BaCu_2O_2$ | $BaCuO_2$ | Ba_2CuO_3 | O_2(air), Pt	1200 K	0.10443	96Pan
(U_T±0.0022 [V]=1.267–0.00096947 T, 1175<T<1235 K)			
$BaMoO_3$ | $BaMoO_4$ | O_2(air), Pt	1100 K	0.74498	89Pan3
(U_T±0.00141 [V]=1.2961–0.00050102 T, 1053<T<1196 K)			
Pt, SS, $BaPbO_3$, $Ba_4Pb_3O_{10}$, PbO | 15CSZ | O_2, Pt, p_{O_2} = 0.21 atm	800 K	0.24486	95Mal1
(U_T±0.00291 [V]=0.59412–0.00043657 T, 756<T<968 K) [2])			
Pt, SS, $Ba_4Pb_3O_{10}$, Ba_2PbO_4, PbO | 15CSZ | O_2, Pt [2]), p_{O_2} = 0.21 atm	800 K	0.28408	95Mal1
(U_T±0.00243 [V]=0.52915–0.00030634 T, 797<T<1097 K)			
Pt, SS, O_2, BaO, BaF_2 | CaF_2 or BaF_2 | BaF_2, $BaPbO_3$, Pb_3O_4, O_2, Pt [3]), p_{O_2} = 1 atm (U_T±0.00193 [V]=0.596011–0.00025508 T, 700<T<880 K)	800 K	0.39194	95Mal1
Pt, O_2, $BaZrO_3$, ZrO_2, BaF_2 | CaF_2 | BaF_2, Y_2BaCuO_5, $Y_2Cu_2O_5$, CuO, O_2, Pt	900 K	0.03305	91Aza2
(U_T±0.0009 [V]=–0.2248+0.000286507199 T, 853<T<1066 K)			
Bismuth (Bi)			
Bi_2O_3, Bi_5CuO_8 | 8YSZ | Cu_2O, Cu [4])	800 K	0.08387	94Mal1
(U_T±0.00199 [V]=0.18021–0.00012042 T, 701<T<941 K)			
Bi_5CuO_8, Bi_4CuO_7, Cu_2O | 8YSZ | CuO, Cu_2O, Cu [4])	900 K	0.26136	94Mal1
(U_T±0.00325 [V]=0.11273+0.00016515 T, 822<T<989 K)			
Pt, Bi_2CuO_4, $Bi_2Cu_2O_5$, Cu_2O | 15YSZ | O_2, Pt [5]), p_{O_2} = 0.21 atm	800 K	0.44029	94Mal1
(U_T±0.00113 [V]=0.88504–0.00060593 T, 772<T<950 K)			

Cell composition	T [K, °C]	U_0 [V]	Ref.
Pt, O_2, $Bi_2Cu_2O_5$, Bi_2CuO_4, CuF_2 | CaF_2 | CuF_2, Bi_2CuO_4, Bi_4CuO_7, O_2, Pt, p_{O_2} = 1 atm (U_T±0.00212 [V]=0.13015–0.00011476 T, 662<T<854 K)	800 K	0.00383	94Mal1
Cadmium (Cd)			
Cd | $CdCl_2$ ‖ $PbCl_2$ | Pb	600 °C	0.1228	30Opp
	650 °C	0.1243	30Opp
	700 °C	0.1265	30Opp
Carbon (C)			
Pt, O_2, CO_2, Na_2CO_3 | Na-β"-alumina | Na_2ZrO_3, ZrO_2, O_2, Pt (U_T±0.004 [V]=–0.2024+0.0002487 T, 850<T<940 K)	900 K	0.02143	02Sub
Cesium (Cs)			
Cs_3CrO_4 | Cr_2O_3 | Cs_2CrO_4 | O_2(air), Pt (U_T±0.0046 [V]=1.3423–0.00052787 T, 860<T<968; U_T±0.0011 [V]=1.2665–0.00045346 T, 1052<T<1098 K)	900 K	0.86722	97Sre
$Cs_2Cr_2O_7$(liquid) | Cr_2O_3 | β-Cs_2CrO_4 | O_2(air), Pt (U_T±0.0018 [V]=0.3443–0.00025709 T, 1025<T<1190 K)	1100 K	0.06150	97Sre
Pt, Cs_2CrO_4(α or β) | Cr_2O_3 | Cs_3CrO_4 | 15CSZ | O_2(air), Pt [6]) (U_T±0.00368 [V]=1.34588–0.00053174 T, 860<T<967 K; U_T±0.0015 [V]=1.26647–0.00045345 T, 1052<T<1097 K)	900 K	0.8673	97Pan2
Pt, Cs_2CrO_4(β) | Cr_2O_3 | Cs_2Cr_2O7 | 15CSZ | O_2(air), Pt [6]) (U_T±0.00184 [V]=0.34442–0.00025715 T, 1025<T<1189 K)	1100 K	0.06155	97Pan2
Cs_2RuO_4 | Cs_2RuO_3 | Cs_2O(liquid) | O_2(air), Pt (U_T±0.001802 [V]=0.7657–0.00048774 T, 946<T<1205 K)	1100 K	0.23677	92Mal
Chromium (Cr)			
Pt, Cr, CrF_2 | CaF_2 | CrF_2, Fe-40Cr, Pt (U_T±0.0014 [V]=–0.0070+0.0000177 T, 920<T<1065 K)	1000 K	0.017	87Aza2
Pt, Cr, CrF_2 | CaF_2 | CrF_2, Ni-25Cr, Pt (U_T±0.00092 [V]=–0.0345+0.0000892 T, 920<T<1065 K)	1000 K	0.0547	87Aza2
Cr | CrF_2 | NiF_2 | Ni (U_T±0.0025 [V]=0.4098–0.0003281 T, 497<T<1063 K)	900 K	0.70509	88Aza
Pt, Cr, CrF_2 | CaF_2 | NiF_2, Ni, Pt (U_T±0.0025 [V]=0.4098–0.0003281 T, 497<T<1063 K)	900 K	0.70509	87Aza1
Pt, Cr, CrF_2, NaF | β-alumina | NaF, NiF_2, Ni, Pt (U_T±0.0025 [V]=0.4098–0.0003281 T, 497<T<1063 K)	900 K	0.70509	87Aza1
Pt, Cr_2O_3,sol., Sb_2O_3,sol., $CrSbO_4$,sol. | 15CSZ | air [6]), p_{O_2} = 0.21 atm Pt (U_T±0.0005 [V]=0.87425–0.00045952 T, 770<T<920 K; U_T±0.0002 [V]=1.00358–0.00059859 T, 931<T<982 K)	800 K	0.5066	99Swa1
CrF_2 | Cr_7Te_8 | Cr_5Te_6 | CrF_2 | Cr (U_T±0.0013 [V]=–0.3675+0.00058087 T, 765<T<1026 K)	900 K	0.15528	89Aza1
CrF_2 | Cr_5Te_6 | Cr_3Te_4 | CrF_2 | Cr (U_T±0.0038 [V]=–0.7042+0.00102135 T, 763<T<948 K)	900 K	0.21501	89Aza1
CrF_2 | Cr_3Te_4 | Cr_2Te_3 | CrF_2 | Cr (U_T±0.0035 [V]=–0.2092+0.00053928 T, 734<T<878 K)	800 K	0.22222	89Aza1
Pt, $LiCrO_2$ | Li_2CrO_4 | Cr_2O_3 | O_2, p_{O_2} = 0.21 atm Pt (U_T±0.0014 [V]=0.33003–0.00045406 T, 626<T<721 K)	700 K	0.01218	85Pan
Cr_2O_3 | K_2CrO_4 | $KCrO_2$ | O_2(air), Pt (U_T±0.00382 [V]=1.2152–0.00066155 T, 771<T<1200 K)	900 K	0.61980	84Sre1

Cell composition	T [K, °C]	U_0 [V]	Ref.
Cr_2O_3 | Rb_2CrO_4 | Rb_3CrO_4 | O_2(air), Pt (U_T±0.00257 [V]=1.2168–0.00050062 T, 798<T<967 K; U_T±0.00345 [V]=1.1671–0.00044982 T, 994<T<1189 K)	1000 K	0.71728	94Pan
Cr_2O_3 | Rb_2CrO_4 | $Rb_2Cr_2O_7$ (liquid) | O_2(air), Pt (U_T±0.0017 [V]=0.3357–0.00025697 T, 973<T<1153 K)	1000 K	0.07873	94Pan
Cobalt (Co)			
Pt, Co, CoF_2 | CaF_2 | Ni, NiF, Pt (U_T±0.0023 [V]=0.1187–0.0000180 T, 688<T<957 K)	1000 K	0.1349	84Bha
Co | CoO | O_2(air), Pt (U_T±0.0024 [V]=1.2760–0.00045492 T, 963<T<1113 K)	1000 K	0.8211	86Sre1
Pt | Co, CoO | 15CSZ | NiO, Co | Pt [6]) (U_T±0.001 [V]=0.0175–0.000620 T, 1175<T<1375 K)	1323 K	0.0965	75Ram
CoO | Co_3O_4 | O_2(air), Pt (U_T±0.0024 [V]=1.1163–0.0009669 T, 716<T<1095 K)	1000 K	0.1494	86Sre1
CoO | Co_3O_4, Cu_2O | Cu (U_T±0.0022 [V]=–0.0028+0.0000308 T, 947<T<1145 K)	1000 K	0.0280	86Sre1
CoO | $CoSb_2O_4$ | CoSbO | O_2(air), Pt (U_T±0.00052 [V]=1.1063+0.00039920 T, 790<T<1039 K)	900 K	0.74702	96Swa1
Pt, CoO, $Co_7Sb_2O_{12}$, $CoSb_2O_4$ | 15CSZ | CoSbO | O_2(air), Pt [6]) (U_T±0.00052 [V]=1.1063+0.00039920 T, 790<T<1040 K)	900 K	0.74702	01Swa
Pt, $Co_7Sb_2O_{12}$, $CoSb_2O_6$, $CoSb_2O_4$ | 15CSZ | O_2(air), Pt [6]) (U_T±0.0004 [V]=0.9679+0.0004395 T, 855<T<1035 K)	900 K	0.57235	01Swa
Co_3O_3 | $Co_7Sb_2O_{12}$ | $CoSb_2O_6$ | O_2(air), Pt (U_T±0.00038 [V]=0.9679+0.00043950 T, 855<T<1035 K)	900 K	0.57235	96Swa1
Copper (Cu)			
Pt, Cu, Cu_2O | 8YSZ | RuO_2, Ru, Pt [4]) (U_T±0.00133 [V]=0.0803+0.00007035 T, 751<T<1200 K)	1100 K	0.11577	90Mal2
Pt, Cu, CuO_2 | 15YSZ | O_2(air), Pt [5]) (U_T±0.0023 [V]=0.85682–0.000401 T, 925<T<1225 K)	1000 K	0.4558	97Sub3
Pt, Cu, Cu_2O | 8YSZ | RuO_2, La_2O_3, Pt [4])	942.5 K	0.2647	90Mal2
Pt, Cu, Cu_2O | 8YSZ | $LaRuO_3$, RuO_2, La_2O_3, Pt [4])	985 K	0.1173	90Mal2
Cu | Sm_2O_3 | $SmCuO_2$ | O_2(air) (U_T±0.00298 [V]=1.1933–0.00067700 T, 967<T<1196 K)	1000 K	0.5163	97Sre
Pt, Cu, Sm_2O_3, $SmCuO_2$ | 15YSZ | O_2(air), Pt [5]) (U_T±0.00298 [V]=1.19334–0.000677 T, 967<T<1196 K)	1000 K	0.5163	97Sub3
Cu | $YCuO_2$ | Y_2O_3 | O_2(air) (U_T±0.0021 [V]=0.8877–0.00042274 T, 1113<T<1255 K)	1200 K	0.3804	96Pan
Cu_2O | Y_2BaO_4 | $OBaCu_5$ | O_2(air) (U_T±0.00508 [V]=1.0361–0.00084942 T, 717<T<1021 K)	800 K	0.35656	89Pan1
Cu_2O | $BaCu_2O_2$ | $BaCuO_2$ | O_2(air) (U_T±0.0017 [V]=0.4427–0.00024715 T, 1003<T<1132 K)	1100 K	0.17084	96Pan
Cu_2O | Sm_2O_3 | Sm_2CuO_4 | O_2(air) (U_T±0.0017 [V]=0.8592–0.0005720 T, 968<T<1203 K)	1100 K	0.230	97Sre
Cu_2O | Sm_2O_3 | Sm_2CuO_4 | CuO | Cu_2O (U_T±0.00018 [V]=0.1814–0.00005150 T, 1017<T<1088 K)	1050 K	0.12732	97Sre
CuO | $YCuO_2$ | Y_2O_3 | O_2(air) (U_T±0.0018 [V]=0.6473–0.00050294 T, 782<T<1122 K)	800 K	0.24531	96Pan

Cell composition	T [K, °C]	U_0 [V]	Ref.
Germanium (Ge)			
Ge | GeO_2(tetragonal) | O_2, p_{O_2} = 0.0001 atm	850 K	1.0696	79Sre
(U_T±0.003 [V]=1.4805–0.00048342 T, 800<T<1150 K)			
Ge | GeO_2(hexagonal) | O_2, p_{O_2} = 0.0001 atm	850 K	1.0541	79Sre
(U_T±0.003 [V]=1.4247–0.00043599 T, 810<T<1180 K)			
Hafnium (Hf)			
Pt, HfO_2, MoO_2, $HfMo_2O_8$ | 15YSZ | O_2(air), Pt [5])	1000 K	0.7193	90Pan2
(U_T±0.00205 [V]=0.8733–0.00029325 T, 982<T<1217 K)			
Iridium (Ir)			
Ir | IrO_2 | O_2(air)	800 K	0.2388	85Mal
(U_T±0.00492 [V]=0.6116–0.00046596 T, 769<T<1263 K)			
Pt | Ir | Cu, CuO_2 | 15CSZ | IrO_2, Ir | Pt [6])	1015 K	0.5967	75Ram
(U_T±0.004 [V]=0.2360–0.00785 T, 970<T<1175 K)			
Pt | Ir | Ni, NiO | 15CSZ | IrO_2, Ir | Pt [6])	1087 K	0.313	75Ram
(U_T±0.0035 [V]=0.5679–0.0038 T, 920<T<1175 K)			
Iron (Fe)			
Fe | FeF_2 | NiF_2 | Ni	800 K	0.35484	88Aza
(U_T±0.0035 [V]=0.2446–0.0001378 T, 506<T<1063 K)			
Pt, Fe, FeF_2 | CaF_2 | NiF_2, Ni, Pt	800 K	0.35484	87Aza1
(U_T±0.0035 [V]=0.2446–0.0001378 T, 506<T<1063 K)			
Pt, Fe, FeF_2, NaF, CaF_2 | β-alumina | NaF, NiF_2, Ni, Pt	800 K	0.35484	87Aza1
(U_T±0.0035 [V]=0.2446–0.0001378 T, 506<T<1063 K)			
Pt, Fe, Fe_xO | 15YSZ | SnO, SnO_2, C, Pt [5])	1206 K	0.11172	01Mal
Pt | Fe, $Fe_{0.95}O$(wustite) | 15CSZ | CoO, Co | Pt [6])	1323 K	0.1965	75Ram
(U_T±0.001 [V]=0.0880–0.000820 T, 1175<T<1375 K)			
Pt | Fe, $Fe_{0.95}O$(wustite) | 15CSZ | NiO, Ni | Pt [6])	1323 K	0.2966	75Ram
(U_T±0.001 [V]=0.1024–0.000820 T, 1175<T<1375 K)			
Pt | Fe, Fe_3O_4 | 15YDT | Cu, CuO_2 | 15YSZ | O_2(air) [7]) [5])	750 K	1.1219	86Mal2
(U_T±0.0005 [V]=2.0697–0.0012637 T, 727<T<794 K)			
Pt | Fe, $Fe_{0.95}O$ | 15YDT | Cu, CuO_2 | 15YSZ | O_2(air) [7]) [5])	800 K	1.0671	86Mal2
(U_T±0.0005 [V]=1.3515–0.0003555 T, 779<T<834 K)			
Pt, $LaFeO_3$, Fe, La_2O_3 | 15CSZ | $FeO_{0.95}$ | Fe [6])	1100 K	0.03441	86Sre2
(U_T±0.00117 [V]=0.07831–0.00003991 T, 1094<T<1299 K)			
Fe_3O_4, $FeSbO_4$ | 15CSZ | air [6]), p_{O_2} = 0.21 atm	800 K	0.676	03Swa
(U_T±0.00022=1.010–0.0004172 T [V], 771<T<921 K)			
$FeSb_2O_4$, Sb_2O_3 | 15CSZ | air [6]), p_{O_2} = 0.21 atm	800 K	0.577	03Swa
(U_T±0.00046 [V]=0.8701–0.0003665 T; U_T±0.00031 [V]=0.8093–0.0002992 T, 787<T<981 K)			
Lanthanum (La)			
La_2CoO_4, La_2O_3,Co | 15YSZ | O_2, Pt [5]), p_{O_2} = 0.001 atm	1000 K	0.7349	84Sre2
(U_T±0.0037 [V]=1.1675–0.0004326 T, 973<T<1375K)			
Pt | O_2, La_2O_3, LaF_3 | CaF_2 | LaF_3, $LaGaO_3$, Ga_2O_3, O_2 | Pt	900 K	0.0696	91Aza1
(U_T±0.0012 [V]=0.08481536–0.00001685 T, 840<T<1105 K)			

Cell composition	T [K, °C]	U_0 [V]	Ref.
Pt, MnO, $LaMnO_3$, La_2O_3 | 15YSZ | O_2, Pt [5]), p_{O_2} = 0.001 atm	1100 K	0.8625	83Sre1
(U_T±0.0042 [V]=1.7402–0.0007979 T, 1064<T<1308 K)			
Pt, O_2, La_2O_3, LaF_3 | CaF_2 | LaF_3, $LaCrO_3$, Cr_2O_3, O_2 | Pt	900 K	0.1470	90Aza2
(U_T±0.0014 [V]=0.1536–0.0000073 T, 855<T<1073 K)			
Pt, La_2O_3, LaF_3, O_2 | CaF_2 | LaF_3, $LaVO_4$, V_2O_5,sol., O_2 | Pt	900 K	0.3593	95Pan2
(U_T±0.0011 [V]=0.3167+0.00004731 T, 887<T<960 K)			
Pt, La_2O_3, LaF_3, O_2 | CaF_2 | LaF_3, $LaVO_4$, V_2O_5(liq.), O_2 | Pt	1000 K	0.5114	95Pan2
(U_T±0.0027 [V]=0.4351+0.00007634 T, 982<T<1059 K)			
Pt | O_2, $La_{0.95}Sr_{0.05}CrO_3$, SrF_2 | SrF_2 | $SrZrO_3$, ZrO_2, SrF_2, O_2 | Pt	1000 K	0.0701	03Sub1
(U_T±0.0011 [V]=0.2306–0.0001605 T, 900<T<1060 K)			
Pt | O_2, $La_{0.85}Sr_{0.15}CrO_3$, SrF_2 | SrF_2 | $SrZrO_3$, ZrO_2, SrF_2, O_2 | Pt	1000 K	0.00328	03Sub1
(U_T±0.0002 [V]=–0.010+0.0000428 T, 890<T<1000)			
Lead (Pb)			
Pb | $PbCl_2$ | glass | Cl_2, C	516 °C	1.093	34Gru
	566 °C	1.057	34Gru
	616 °C	1.111	34Gru
Pb | $PbCl_2$ | glass | $PbCl_2$ | Cl_2, C	515 °C	1.235	34Gru
	555 °C	1.217	34Gru
	595 °C	1.188	34Gru
	635 °C	1.168	34Gru
Pb | $PbCl_2$, KCl | glass | $PbCl_2$, KCl | Cl_2, C	410 °C	1.387	34Gru
	460 °C	1.358	34Gru
	510 °C	1.332	34Gru
	560 °C	1.306	34Gru
	630 °C	1.270	34Gru
Pb | $PbCl_2$ ‖ $CdCl_2$, $PbCl_2$, x=0.1 | Pb	570 °C	0.0597	28Lor
Pb | $PbCl_2$ ‖ $CdCl_2$, $PbCl_2$, x=0.176 | Pb	570 °C	0.0412	28Lor
Pb | $PbCl_2$ ‖ $CdCl_2$, $PbCl_2$, x=0.25 | Pb	570 °C	0.0373	28Lor
Pb | $PbCl_2$ ‖ $CdCl_2$, $PbCl_2$, x=0.45 | Pb	570 °C	0.0322	28Lor
Pb | $PbCl_2$ ‖ $CdCl_2$, $PbCl_2$, x=0.75 | Pb	570 °C	0.0130	28Lor
Pb | $PbCl_2$, KCl, 2:1 ‖ $PbCl_2$ | Pb	490 °C	0.027	34Gru
	540 °C	0.034	34Gru
	630 °C	0.048	34Gru
Pb | $PbCl_2$, KCl ‖ $PbCl_2$ | Pb	470 °C	0.047	34Gru
	520 °C	0.065	34Gru
	580 °C	0.072	34Gru
	630 °C	0.079	34Gru
Pb | $PbCl_2$, KCl, 2:1 | glass | $PbCl_2$, KCl, 2:1 | Cl_2, C	426 °C	1.346	34Gru
	476 °C	1.315	34Gru
	526 °C	1.288	34Gru
	576 °C	1.260	34Gru
	626 °C	1.234	34Gru
Pb | $PbCl_2$, KCl | glass | AgCl | Ag	430 °C	0.673	34Gru
	480 °C	0.622	34Gru
	530 °C	0.619	34Gru
Pb | $PbCl_2$, KCl | glass | AgCl | Ag	630 °C	0.616	34Gru

Cell composition	T [K, °C]	U_0 [V]	Ref.
Pb | $PbCl_2$, KCl, 2:1 | glass | AgCl | Ag	426 °C	0.635	34Gru
	476 °C	0.536	34Gru
	526 °C	0.510	34Gru
	626 °C	0.511	34Gru
Pb(liquid) | PbO | O_2(air)	850 K	1.2066	95Mal2
(U_T±0.00093 [V]=1.1019–0.00051324 T, 767<T<901 K)			
PbO | Pb_3O_4 | O_2(air)	750 K	0.0412	95Mal2
(U_T±0.00101 [V]=0.3738–0.00044341 T, 696<T<812 K)			
[Pb-Na] | α- and β-alumina | FeO | Fe	900 K	1.1101	91Jac
(U_T±0.002 [V]=1.77992–0.00076560 T, 790<T<980 K)			
[Pb-Na] | α- and β"-alumina | FeO | Fe	900 K	0.7623	91Jac
(U_T±0.0005 [V]=1.6013–0.00093220 T, 790<T<980 K)			
[Pb-Na] | β"-alumina | $NaAlO_2$ | FeO | Fe	900 K	0.51847	91Jac
(U_T±0.0006 [V]=1.0744–0.00061770 T, 790<T<980 K)			
Lithium (Li)			
Pt, $LiCrO_2$, Cr, Cr_2O_3, LiF | CaF_2 | Co, CoF_2, Pt	900 K	0.7952	84Bha
(U_T±0.0059 [V]=0.1714–0.0006931 T, 771<T<971 K)			
Pt, $LiCrO_2$, Cr, Cr_2O_3, LiF | CaF_2 | Ni, NiF_2, Pt	900 K	0.9163	84Bha
(U_T±0.0174 [V]=0.1534–0.0008477 T, 688<T<1003 K)			
Magnesium (Mg)			
Mg | $MgCl_2$ ‖ $ZnCl_2$ | Zn	720 °C	0.869	28Lor
	750 °C	0.891	28Lor
Mg | $MgCl_2$ ‖ KCl,NaCl ‖ $CdCl_2$ | Cd	720 °C	0.964	28Lor
	750 °C	0.996	28Lor
Mg | $MgCl_2$ ‖ $PbCl_2$ | Pb	720 °C	1.078	28Lor
	740 °C	1.103	28Lor
	760 °C	1.121	28Lor
Mg | $MgCl_2$ ‖ TlCl | Tl	720 °C	0.530	28Lor
	750 °C	0.517	28Lor
MgU_2O_6 | $MgU_2O_{6.67}$ | O_2(air)	850 K	0.45694	89Mad
(U_T±0.01219 [V]=1.1189–0.00077878 T, 779<T<1093 K)			
$MgU_2O_{6.67}$ | MgU_2O_7 | O_2(air)	850 K	0.82157	89Mad
(U_T±0.000497 [V]=1.0390–0.00081660 T, 826<T<1208 K)			
$Mg_{0.2}UO_{2+x}$ | MgU_2O_6 | O_2(air)	850 K	0.58374	89Mad
(U_T±0.00633 [V]=0.8293–0.00028889 T, 8146<T<1294 K)			
Manganese (Mn)			
Mn | MnF_2 | NiF_2 | Ni	900 K	1.0406	88Aza
(U_T±0.00121 [V]=0.7206–0.0003556 T, 840<T<1090 K)			
Molybdenum (Mo)			
Mo | MoO_2 | FeO | Fe	900 K	0.03090	90Mal1
(U_T±0.00149 [V]=0.1206–0.00009966 T, 837<T<1151 K)			
Pt | C | MoO | Mo_4O_{11} | 15YSZ | O_2(air), Pt [5])	800 K	0.49461	88Pan1
(U_T±0.0016 [V]=0.7893–0.00036836 T, 768<T<1058 K)			
Pt | C | MoO | Mo_4O_3 | 15YSZ | O_2(air), Pt [5])	1100 K	0.39271	88Pan1
(U_T±0.0016 [V]=0.7893–0.00036836 T, 1075<T<1150 K)			
MoO_2 | Mo_3Te_4, $MoTe_2$(α) | FeO | Fe	900 K	0.02429	90Mal1
(U_T±0.0016 [V]=–0.0556+0.00008877 T, 775<T<1196 K)			

Cell composition	T [K, °C]	U_0 [V]	Ref.
$MoTe_2(\alpha)$ | Te(solid or liquid), MoO_2 | O_2(air), Pt	700 K	0.94160	88Mal
(U_T±0.0044 [V]=1.3571–0.00059357 T, 654<T<713 K; U_T±0.00335 [V]=1.1685–0.00033937 T, 728<T<1065 K)			
Mo, Sb, Fe_3O_4, $FeSb_2O_4$ | 15CSZ | NiO, Ni, Mo [6])	800 K	0.115	03Swa
(U_T±0.0039 [V]=0.1528–0.00004728 T, 771<T<911 K)			
Neodymium (Nd)			
Pt, $Nd_4Co_3O_{10}$, Nd_2O_3, CoO | 15CSZ | O_2(air), Pt [6])	1100 K	0.3172	98Sub2
(U_T±0.003 [V]=0.8958–0.000526 T, 1080<T<1223 K)			
Pt, $Nd_4Co_3O_{10}$, $NdCoO_3$, CoO | 15CSZ | O_2(air), Pt [6])	1150 K	0.2587	98Sub2
(U_T±0.0024 [V]=0.9269–0.000581 T, 1100<T<1200 K)			
Pt, $SmCoO_3$, Sm_2O_3, CoO | 15CSZ | O_2(air), Pt [6])	1100 K	0.2252	98Sub2
(U_T±0.0011 [V]=0.5563–0.000304728 T, 1080<T<1180 K)			
Pt, $GdCoO_3$, Gd_2O_3, CoO | 15CSZ | O_2(air), Pt [6])	1150 K	0.1773	98Sub2
(U_T±0.0022 [V]=0.6741–0.000432 T, 1100<T<1200 K)			
Pt, $DyCoO_3$, Dy_2O_3, CoO | 15CSZ | O_2(air), Pt [6])	1150 K	0.1508	98Sub2
(U_T±0.0034 [V]=0.7902–0.000556 T, 1100<T<1200 K)			
Nickel (Ni)			
Ni | La_2NiO_4 | La_2O_3 | NiO | Ni	1200 K	0.05518	76Sre
(U_T±0.001 [V]=–0.1325+0.0001564 T, 1123<T<1373 K)			
Pt, Ni, NiF_2 | CaF_2 | CrF_2, Ni-25Cr, Pt	1000 K	0.0116	87Aza2
(U_T±0.00092 [V]=–0.0345+0.0000892 T, 973<T<1123 K)			
Pt, Ni, NiO | 15CSZ | CuO_2, Cu | Pt [6])	1000 K	0.284	75Ram
(U_T±0.002 [V]=0.3640–0.0008 T, 950<T<1275 K)			
Pt, Ni, NiO | 15CSZ | CoO, La_2O_3, $LaCoO_3$ | Pt [6])	1161 K	0.2196	72Sre
(U_T [V]=–0.0238+0.0002079 T, 1100<T<1200 K)			
Pt | Ni, NiO | 15CSZ | CoO, La_2O_3, $LaCoO_3$ | Pt [6])	1318 K	0.2342	72Sre
(U_T [V]=0.1352+0.000076 T, 1250<T<1325 K)			
NiO | $NiSb_2O_4$ | $NiSb_2O_6$ | O_2(air), Pt	900 K	0.55454	96Swa2
(U_T±0.0003 [V]=1.0535–0.00055440 T, 894<T<1061 K)			
NiO | $NiTeO_3$ | Tellur(liquid) | O_2(air), Pt	900 K	0.43485	91Mal2
(U_T±0.00148 [V]=0.8887–0.00050427 T, 833<T<1104 K)			
NiO | Ni_3TeO_6 | $NiTeO_3$ | O_2(air), Pt	900 K	0.16141	91Mal2
(U_T±0.00421 [V]=0.8953–0.00081543 T, 624<T<964 K)			
Chromel, Mo, Sb, NiO, $NiSb_2O_4$ | 15CSZ | NiO, Ni, Mo, Chromel [6])		0.00486	99Swa2
(U_T±0.0003 [V]=0.0883–0.0001043 T, 769<T<931 K)			
Pt, NiO, $NiSb_2O_4$, $NiSb_2O_6$ | 15CSZ | air, p_{O_2} = 0.21 atm, Pt [6])		0.4991	99Swa2
(U_T±0.0003 [V]=1.0535–0.0005544 T, 894<T<1138 K)			
Niobium (Nb)			
Pt, Nb, NbO | α- and β-alumina | Ta_2O_5, Ta, Pt	1100 K	0.0369	97Sub2
(U_T±0.0005 [V]=0.115–0.000071 T, 970<T<1120 K)			
Pt, Nb, Li_3NbO_4, Li_2O | α- and β-alumina | NbO | Nb	1100 K	0.2874	97Sub2
(U_T±0.003 [V]=0.7395–0.00041100 T, 1009<T<1145 K)			
Nitrogen (N)			
Pt, N_2, $p_{O_2} \approx 10^{-3}$ atm | 15YSZ | air, Pt [5])	900 K	0.08069	01Mal
(U_T±0.0026 [V]=–0.0022+0.0000921 T, 788<T<965 K)			

Cell composition	T [K, °C]	U_0 [V]	Ref.
Palladium (Pd)			
Pd | PdO | 15YSZ | O_2(air) [5])	800 K	0.14378	83Mal
(U_T±0.00142 [V]=0.5845–0.00055090 T, 699<T<1060 K)			
PdTe | $PdTe_2$ | TeO_2 | O_2(air)	900 K	0.34171	86Mal3
(U_T±0.00399 [V]=0.7303–0.00043176 T, 613<T<995 K)			
Potassium (K)			
$K_2Cr_2O_7$(liquid) | Cr_2O_3 | β-K_2CrO_4 | O_2(air), Pt	1100 K	0.04063	97Sre
(U_T±0.0026 [V]=0.3120–0.00024670 T, 978<T<1149 K)			
$KCrO_2$ | K_3CrO_4 | K_2CrO_4 | O_2(air), Pt	1100 K	0.60586	97Pan1
(U_T±0.0035 [V]=1.1521–0.00049658 T, 950<T<1134 K)			
Rhodium (Rh)			
Rh | Rh_2O_3 | O_2(air), Pt	900 K	0.19770	85Mal
(U_T±0.00255 [V]=0.6082–0.00045611 T, 699<T<1060 K)			
Ruthenium (Ru)			
Ru, $CaRuO_3$, CaO | 15CSZ | O_2(air), Pt [6])	1100 K	0.35095	91Mal1
(U_T±0.00170 [V]=0.7542–0.00036659 T, 971<T<1312 K)			
Ru | Cs_2RuO_3 | Cs_2O(liquid) | RuO_2, Ru	1100 K	0.001555	92Mal
(U_T±0.00122 [V]=0.0614–0.00004168 T, 1001<T<1189 K)			
Pt, $LaRuO_3$, La_2O_3, Ru | 15YSZ | O_2(air), Pt [6])	1100 K	0.29513	90Mal2
(U_T±0.00240 [V]=0.8441–0.00049906 T, 1032<T<1256 K)			
Pt, SrO, Sr_2RuO_4, Ru | 15CSZ | O_2(air), Pt [6])	1100 K	0.56478	93Mal
(U_T±0.00337 [V]=1.0093–0.00040411 T, 1044<T<1235 K)			
Pt, $Sr_3Ru_2O_7$, $SrRuO_3$, Ru | 15CSZ | O_2(air), Pt [6])	1000 K	0.57601	93Mal
(U_T±0.00378 [V]=1.0105–0.00043449 T, 991<T<1282 K)			
Pt, Sr_2RuO_4, $Sr_3Ru_2O_7$, Ru | 15CSZ | O_2(air), Pt [6])	1000 K	0.59174	93Mal
(U_T±0.00326 [V]=1.0607–0.00046896 T, 926<T<1164 K)			
Pt, Ru, RuO_2 | 15YSZ | O_2(air), Pt [5])	1100 K	0.2609	90Mal2
(U_T±0.00252 [V]=0.8136–0.00050242 T, 1005<T<1106 K)			
Samarium (Sm)			
Sm_2O_3 | $SmCuO_2$ | Sm_2CuO_4 | O_2(air)	1100 K	0.0958	97Sre
(U_T±0.0036 [V]=0.6491–0.000503 T, 619<T<1152 K)			
Pt, Sm_2CuO_4, Sm_2O_3, CuO_2 | 15YSZ | O_2(air), Pt [5])	1100 K	0.230	97Sub3
(U_T±0.0037 [V]=0.85922–0.000572 T, 968<T<1203 K)			
Pt, Sm_2CuO_4, Sm_2O_3, CuO_2 | 10CSZ | Cu_2O, CuO, Pt [8])	1050 K	0.1273	97Sub3
(U_T±0.00018 [V]=0.18135–0.0000515 T, 1017<T<1088 K)			
Pt, Sm_2CuO_4, $SmCuO_2$, Sm_2O_3 | 15YSZ | O_2(air), Pt [5])	1000 K	0.1461	97Sub3
(U_T±0.0036 [V]=0.64906–0.000503 T, 619<T<1152 K)			
Sm_2O_3 | Sm_2BaO_4 | BaF_2 | $BaZrO_3$ | ZrO_2	1000 K	0.05840	98Sub1
(U_T±0.0006 [V]=0.0227+0.000035070 T, 906<T<1137 K)			
Silver (Ag)			
Ag | AgCl | Jena glass | AgCl | Cl_2	480 °C	0.895	34Gru
	530 °C	0.882	34Gru
	630 °C	0.857	34Gru
Ag | AgCl | fels glass | AgCl | Cl_2, C	614 °C	0.870	34Gru
	664 °C	0.867	34Gru
	704 °C	0.866	34Gru
Ag | AgCl | glass | $PbCl_2$ | Cl_2, C	520 °C	1.113	34Gru
	570 °C	1.119	34Gru
	630 °C	1.128	34Gru

Cell composition	T [K, °C]	U_0 [V]	Ref.
Sodium (Na)			
Pt, $NaCrO_2$, Cr, Cr_2O_3, NaF | CaF_2 | Co, CoF_2, Pt (U_T±0.0055 [V]=0.7458–0.0003613 T, 782<T<1014 K)	900 K	1.0707	84Bha
Pt, $NaCrO_2$, Cr, Cr_2O_3, NaF | CaF_2 | Ni, NiF_2, Pt (U_T±0.0265 [V]=0.3606–0.0000265 T, 809<T<1040 K)	900 K	1.160	84Bha
$NaCrO_2$ | Cr_2O_3 | Na_2CrO_4 | O_2(air), Pt (U_T±0.0044 [V]=0.48367–0.00034155 T, 784<T<1012 K)	900 K	0.17627	83Sre2
$Na_2Cr_2O_7$(liquid) | Cr_2O_3 | β-Na_2CrO_4 | O_2(air) (U_T±0.0012 [V]=0.2620–0.00022930 T, 960<T<1069 K)	1000 K	0.03270	97Sre
Pt, O_2, Na_2ZrO_3, ZrO_2 | Na-β"-alumina | Al_2O_3, $NaAl_{11}O_{17}$, O_2, Pt (U_T±0.004 [V]=0.26026–0.0001394 T, 880<T<1045 K)	1000 K	0.3996	02Sub
Pt, O_2, Na_2ZrO_3, ZrO_2 | Na-β"-alumina | $Na_2Mo_2O_7$, Na_2MoO_4, O_2, Pt (U_T±0.002 [V]=–0.28824+0.0004975 T, 700<T<800 K)	800 K	0.6862	02Sub
Strontium (Sr)			
Pt, SrO, SrF_2, O_2 | SrF_2 | O_2, SrF_2, Y_2O_3, SrY_2O_4 (U_T±0.00076 [V]=0.23646–0.00016198 T, 935<T<1080 K)	1000 K	0.07448	02Aru
Pt | O_2, $SrZrO_3$, ZrO_2, SrF_2 | SrF_2 | $La_{1.7}Sr_{0.3}CrO_3$, SrF_2, O_2 | Pt (U_T±0.0003 [V]=0.1059–0.0000835 T, 940<T<1040 K)	1000 K	0.0224	03Sub1
Pt, SrO, SrF_2, O_2 | SrF_2 | O_2, SrF_2, $La_{0.85}Sr_{0.15}MnO_{3-y}$ | Pt (U_T±0.00013 [V]=0.05166–0.000010106 T, 720<T<1048 K)	1000 K	0.06176	01Aru
Pt, $La_{0.85}Sr_{0.15}MnO_{3-y}$, SrF_2, O_2 | CaF_2 | O_2, SrF_2, $La_{0.935}Sr_{0.065}MnO_{3-y}$, Pt ($U_T$±0.00025 [V]=–0.02612–0.00007270 T, 745<T<989 K)	900 K	0.03931	01Aru
Pt, $La_{0.85}Sr_{0.15}MnO_{3-y}$, SrF_2, O_2 | SrF_2 | O_2, SrF_2, $La_{0.935}Sr_{0.065}MnO_{3-y}$, Pt ($U_T$±0.00053 [V]=–0.02605–0.00007306 T, 940<T<1040 K)	1000 K	0.04701	01Aru
Pt, $La_{0.8}Sr_{0.2}MnO_{3-y}$, SrF_2, O_2 | CaF_2 | O_2, SrF_2, $La_{0.935}Sr_{0.065}MnO_{3-y}$, Pt ($U_T$±0.00122 [V]=–0.03421–0.00010438 T, 744<T<986 K)	900 K	0.05973	01Aru
Pt, $La_{0.85}Sr_{0.15}MnO_{3-y}$, SrF_2, O_2 | CaF_2 | O_2, SrF_2, $La_{0.9}Sr_{0.1}MnO_{3-y}$, Pt ($U_T$±0.00013 [V]=–0.02155–0.00004533 T, 736<T<1004 K)	900 K	0.01925	01Aru
Pt, $La_{0.8}Sr_{0.2}MnO_{3-y}$, SrF_2, O_2 | CaF_2 | O_2, SrF_2, $La_{0.9}Sr_{0.1}MnO_{3-y}$, Pt	897 K	0.045	01Aru
Pt, $La_{0.8}Sr_{0.2}MnO_{3-y}$, SrF_2, O_2 | CaF_2 | O_2, SrF_2, $La_{0.9}Sr_{0.1}MnO_{3-y}$, Pt	986 K	0.0504	01Aru
Tantalum (Ta)			
Ta | $LiTaO_3$ | $LiTa_3O_8$ | α- and β-alumina | Ta_2O_5 | Ta (U_T±0.00076 [V]=0.0150+0.0001517 T, 1130<T<1213 K)	1200 K	0.19794	97Sub1
Ta | Ta_2O_5 | α- and β-alumina | NbO | Nb (U_T±0.00005 [V]=0.1150–0.000071 T, 795<T<954 K)	1000 K	0.0440	96Sub
Tellurium (Te)			
Te | Sb_2Te_3 | Sb_2O_3 | O_2(air) (U_T±0.0008 [V]=1.0277–0.00038396 T)	700 K	0.75893	86Mal1
Te | TeO_2 | O_2(air) (U_T±0.00314 [V]=0.8376–0.00048972 T, 588<T<722 K)	600 K	0.54377	86Mal1
Te(liquid) | TeO_2 | O_2(air) (U_T±0.00144 [V]=0.87917–0.00054830 T, 726<T<898 K)	800 K	0.44053	86Mal1
Pt, C, $[Te]_{Pd\text{-}Te5at\%}$, TeO_2 | 15YSZ | O_2, Pt [5]), p_{O_2} = 0.21 atm (U_T±0.00214 [V]=0.52270–0.00051321 T, 643<T<953 K)	800 K	0.11213	89Mal
Pt, C, $[Te]_{Pd\text{-}Te10at\%}$, TeO_2 | 15YSZ | O_2, Pt [5]), p_{O_2} = 0.21 atm (U_T±0.00214 [V]=0.52270–0.00051321 T, 643<T<953 K)	800 K	0.11213	89Mal
Pt, C, $[Te]_{Pd\text{-}Te21.5at\%}$, TeO_2 | 15YSZ | O_2, Pt [5]), p_{O_2} = 0.21 atm (U_T±0.00357 [V]=0.50725–0.00049013 T, 555<T<974 K)	800 K	0.11515	89Mal

Cell composition	T [K, °C]	U_0 [V]	Ref.
Pt, C, $[Te]_{Pd\text{-}Te25at\%}$, TeO_2 | 15YSZ | O_2, Pt [5]), p_{O_2} = 0.21 atm	800 K	0.11448	89Mal
(U_T±0.00261 [V]=0.54414–0.00053707 T, 612<T<972 K)			
Pt, C, $[Te]_{Pd\text{-}Te26.5at\%}$, TeO_2 | 15YSZ | O_2, Pt [5]), p_{O_2} = 0.21 atm	800 K	0.11736	89Mal
(U_T±0.00218 [V]=0.52717–0.00051226 T, 632<T<818 K)			
Pt, C, $[Te]_{Pd\text{-}Te28at\%}$, TeO_2 | 15YSZ | O_2, Pt [5]), p_{O_2} = 0.21 atm	800 K	0.17782	89Mal
(U_T±0.00225 [V]=0.53393–0.00052167 T, 540<T<736 K; U_T±0.00251 [V]=0.50796–0.00048654 T, 746<T<899 K)			
Pt, C, $[Te]_{Pd\text{-}Te32at\%}$, TeO_2 | 15YSZ | O_2, Pt [5]), p_{O_2} = 0.21 atm	800 K	0.29305	89Mal
(U_T±0.00175 [V]=0.69491–0.00050120 T, 578<T<750 K; U_T±0.00201 [V]=0.77336–0.00060681 T, 772<T<889 K)			
Pt, C, $[Te]_{Pd\text{-}Te35at\%}$, TeO_2 | 15YSZ | O_2, Pt [5]), p_{O_2} = 0.21 atm	800 K	0.32583	89Mal
(U_T±0.00304 [V]=0.57359–0.00030970 T, 584<T<730 K; U_T±0.00281 [V]=0.77709–0.00058433 T, 749<T<916 K)			
Pt, C, $[Te]_{Pd\text{-}Te38at\%}$, TeO_2 | 15YSZ | O_2, Pt [5]), p_{O_2} = 0.21 atm	800 K	0.32179	89Mal
(U_T±0.00138 [V]=0.71316–0.00048921 T, 546<T<768 K; U_T±0.00126 [V]=0.85994–0.00067843 T, 776<T<878 K)			
Pt, C, $[Te]_{Pd\text{-}Te43at\%}$, TeO_2 | 15YSZ | O_2, Pt [5]), p_{O_2} = 0.21 atm	800 K	0.32699	89Mal
(U_T±0.00217 [V]=0.66947–0.00042809 T, 547<T<771 K; U_T±0.00074 [V]=0.63742–0.00038612 T, 785<T<963 K)			
Pt, C, $[Te]_{Pd\text{-}Te52at\%}$, TeO_2 | 15YSZ | O_2, Pt [5]), p_{O_2} = 0.21 atm	800 K	0.33893	89Mal
(U_T±0.00168 [V]=0.60155–0.00032827 T, 549<T<958 K)			
Thallium (Tl)			
Tl | TlCl ‖ $PbCl_2$ | Pb	500 °C	0.365	28Lor
	600 °C	0.404	28Lor
Tl | TlCl ‖ $SnCl_2$ | Sn	440 °C	0.378	28Lor
	460 °C	0.386	28Lor
	480 °C	0.395	28Lor
	550 °C	0.403	28Lor
Tl | TlCl ‖ KCl, LiCl ‖ $SnCl_2$ | Sn	500 °C	0.367	28Lor
	600 °C	0.382	28Lor
Thorium (Th)			
Pt, O_2, ThO_2-2mol%SrO, SrF_2 | SrF_2 | ZrO_2, $SrZrO_3$, SrF_2, O_2, Pt	750 K	0.0384	03Sub2
(U_T±0.0008 [V]=–0.0141+0.0000700 T, 678<T<782 K)			
Pt, O_2, ThO_2-3mol%SrO, SrF_2 | SrF_2 | ZrO_2, $SrZrO_3$, SrF_2, O_2, Pt	750 K	0.0704	03Sub2
(U_T±0.0015 [V]=–0.0571+0.0001700 T, 657<T<795 K)			
Pt, O_2, ThO_2-3mol%SrO, SrF_2 | SrF_2 | ThO_2-2mol%SrO, SrF_2, O_2, Pt	750 K	0.0304	03Sub2
(U_T±0.0008 [V]=–0.0148+0.0000602 T, 688<T<780 K)			
Tin (Sn)			
Pt, C, SnO, SnO_2 | 15YSZ | N_2 [5]), $p_{O_2} \approx 10^{-3}$ atm	923 K	0.8802	01Mal
Uranium (U)			
Pt, $UMoO_5$ | $UMoO_6$ | O_2, p_{O_2} = 0.21 atm, Pt	900 K	0.30989	87Swa
(U_T±0.0043 [V]=0.7517–0.00049090 T, 776<T<1127 K)			

Cell composition	T [K, °C]	U_0 [V]	Ref.
Yttrium (Y)			
$YBa_2Cu_3O_{7-x}$ | O_2(air), Pt	700 K	0.11069	89Pan2
(U_T±0.0020 [V]=0.2982–0.00026787 T, 620<T<850 K; U_T±0.0003 [V]=0.1643–0.00010961 T, 860<T<910 K)			
Y_2BaO_4, $Y_2Ba_4O_7$, BaF_2 | CaF_2 | BaF_2, $BaZrO_3$, ZrO_2, O_2(air), Pt	900 K	0.1419	90Aza1
(U_T±0.003 [V]=0.2022–0.00006699 T, 776<T<1110 K)			
Pt, O_2, Y_2O_3, YF_3 | CaF_2 | YF_3, $Y_2Cu_2O_5$, CuO, O_2, Pt	900 K	0.00212	91Aza2
(U_T±0.0005 [V]=–0.0319+0.000037802 T, 861<T<980 K)			
Pt | Y_2O_3, Cu_2O, $Y_2Cu_2O_5$ | 15YSZ | O_2(air), Pt [5])	1100 K	0.12652	88Pan2
(U_T±0.0011 [V]=0.6443–0.00047071 T, 1097<T<1292 K)			
Y_2MoO_6 | Y_2MoO_5 | O_2, p_{O_2} = 0.21 atm, Pt	1100 K	0.71473	91Pan
(U_T±0.00271 [V]=1.2886–0.00052170 T, 1015<T<1222 K)			
Pt, O_2(1 atm) | Y_2O_3, Y_2BaO_4, BaF_2 | CaF_2 | BaF_2, $BaZrO_3$, ZrO_2, O_2(1 atm), Pt	1000 K	0.2996	89Aza2
(U_T±0.0021 [V]=0.3010–0.0000372 T, 974<T<1222 K)			
Zinc (Zn)			
Zn | $ZnCl_2$ ‖ $CdCl_2$ | Cd	600 °C	0.155	30Opp
	650 °C	0.150	30Opp
Zn | $ZnCl_2$ ‖ $PbCl_2$ | Pb	500 °C	0.276	28Lor
	530 °C	0.273	28Lor
	560 °C	0.270	28Lor
	600 °C	0.267	28Lor
Zn | $ZnCl_2$ ‖ $SnCl_2$ | Sn	430 °C	0.323	28Lor
	500 °C	0.315	28Lor
Zn | $ZnCl_2$ ‖ TlCl | Cd	430 °C	0.077	28Lor
	450 °C	0.086	28Lor
	470 °C	0.095	28Lor
	500 °C	0.110	28Lor
Zirconium (Zr)			
Pt, C, ZrO_2, $ZrTe_3O_8$, Te(liquid) | O_2(air)	800 K	0.47575	94Mal2
(U_T±0.00318 [V]=0.8981–0.00046928 T, 874<T<1135 K)			
Pt, O_2, ZrO_2, $SrZrO_3$, SrF_2 | SrF_2 | ThO_2-1 mol% SrO, SrF_2, O_2, Pt	750 K	0.0323	03Sub2
(U_T±0.0003 [V]=0.0684–0.0000482 T, 668<T<757 K)			
$ZrMo_2O_8$ | MoO_2 | ZrO_2 | O_2(air), Pt	1000 K	0.47463	90Pan1
(U_T±0.00156 [V]=0.7724–0.0002977 T, 986<T<1206 K)			

[1]) Description of the cell composition is based - as far as possible - on the Stockholm convention (1953), i.e. the left-hand electrode constitutes the negative terminal of the cell. Cells are listed according to the metallic constituent of the electrode mentioned first which is involved in the electrode reaction establishing the respective electrode potential. Contact materials and conductive additives may be mentioned first before the actual element of interest only for the sake of correct materials sequence. The sequence of electrode components is stated as reported in the original publications. When an oxygen electrode is used as reference electrode an oxygen partial pressure of 0.21 atm is assumed.

[2]) SS = stainless steel, inert separator (barrier) between Pt contact and electrode under investigation; 15CSZ = 15 wt% CaO-stabilized ZrO_2.

[3]) SS = stainless steel, inert separator (barrier) between Pt contact and electrode under investigation.

[4]) 8YSZ = 8 wt% Y_2O_3-stabilized ZrO_2.

[5]) 15YSZ = 15 wt% Y_2O_3-stabilized ZrO_2.

[6]) 15CSZ = 15 mol% CaO-stabilized ZrO_2.

[7]) 15YDT = 15 wt% yttria doped thoria.

[8]) 10CSZ = 10 mol% CaO-stabilized ZrO_2.

References

28Lor Lorenz, R., Michael, F.: Z. Phys. Chem. Abt. A **137** (1928) 1.

30Opp Oppenheimer, F.: Z. Anorg. Allg. Chem. **189** (1930) 297.

34Gru Grube, G., Rau, A.: Z. Elektrochem. **40** (1934) 352.

72Sre Sreedharan, O.M., Chandrasekharaiah, M.S.: Mater. Res. Bull. **7** (1972) 1135.

75Ram Ramakrishnan, E.S., Sreedharan, O.M., Chandrasekharaiah, M.S.: J. Electrochem. Soc. **122** (1975) 328.

76Sre Sreedharan, O.M., Chandrasekharaiah, M.S., Karkhanawala, M.D.: High Temp. Sci. **8** (1976) 179.

79Sre Sreedharan, O.M., Athiappan, E., Pankajavalli, R., Gnanamoorthy, J.B.: J. Less-Common Met. **68** (1979) 143.

83Mal Mallika, C., Sreedharan, O.M., Gnanamoorthy, J.B.: J. Less-Common Met. **95** (1983) 213.

83Sre1 Sreedharan, O.M., Pankajavalli, R., Gnanamoorthy, J.B.: High Temp. Sci. **16** (1983) 251.

83Sre2 Sreedharan, O.M., Madan, B.S., Gnanamoorthy, J.B.: J. Nucl. Mater. **119** (1983) 296.

84Bha Bhat, N.P., Swaminathan, K., Krishnamurthy, D., Sreedharan, O.M., Sundaresan, M.: Proc. 3rd Int. Conf. Liquid Metal Eng. and Technol., Oxford, 1984, Vol. 1, Brit. Nucl. Energ. Soc., London, p. 323.

84Sre1 Sreedharan, O.M., Madan, B.S., Pankajavalli, R., Gnanamoorthy, J.B.: Proc. 3rd International Conference on Liquid Metal Engineering and Technology, Vol. 1, in: Energy Production, London, 1984, p. 249.

84Sre2 Sreedharan, O.M., Pankajavalli, R.: J. Mater. Sci. Lett. **3** (1984) 388.

85Mal Mallika, C., Sreedharan, O.M., Chandrasekharaiah, M.S.: J. Less-Common Met. **107** (1985) 203.

85Pan Pankajavalli, R., Sreedharan, O.M., Gnanamoorthy, J.B.: J. Nucl. Mater. **127** (1985) 170.

86Aza Azad, A.M., Pankajavalli, R., Sreedharan, O.M.: J. Chem. Thermodyn. **18** (1986) 255.

86Mal1 Mallika, C., Sreedharan, O.M.: J. Chem. Thermodyn. **18** (1986) 727.

86Mal2 Mallika, C., Pankajavalli, R., Sreedharan, O.M.: Electrochim. Acta **31** (1986) 885.

86Mal3 Mallika, C., Sreedharan, O.M.: J. Mater. Sci. Lett. **5** (1986) 915.

86Sre1 Sreedharan, O.M., Mallika, C.: Mater. Chem. Phys. **14** (1986) 375.

86Sre2 Sreedharan, O.M., Chandrasekharaiah, M.S.: J. Mater. Sci. **21** (1986) 2581.

87Aza1 Azad, A.M., Sreedharan, O.M.: J. Appl. Electrochem. **17** (1987) 949.

87Aza2 Azad, A.M., Sreedharan, O.M.: Bull. Electrochem. **3** (1987) 331.

87Pan Pankajavalli, R., Sreedharan, O.M.: J. Mater. Sci. **22** (1987) 177.

87Swa Swaminathan, K., Mallika, C., Sreedharana, O.M.: J. Am. Ceram. Soc. **70** (1987) C-168.

88Aza Azad, A.M., Sreedharan, O.M.: Trans. SAEST **23** (1988) 45.

88Mal Mallika, C., Sreedharan, O.M.: J. Chem. Thermodyn. **20** (1988) 769.

88Pan1 Pankajavalli, R., Sreedharan, O.M.: J. Chem. Thermodyn. **20** (1988) 309.

88Pan2 Pankajavalli, R., Sreedharan, O.M.: J. Mater. Sci. Lett. **7** (1988) 714.

89Aza1 Azad, A.M., Sreedharan, O.M.: J. Nucl. Mater. **167** (1989) 89.

89Aza2 Azad, A.M., Sreedharan, O.M.: J. Mater. Sci. Lett. **8** (1989) 67.

89Mad Madan, B.S., Mallika, C., Sreedharan, O.M., Rodiguez, P., Paranjpe, S.R.: J. Nucl. Mater. **167** (1989) 278.

89Mal Mallika, C., Sreedharan, O.M.: J. Nucl. Mater. **167** (1989) 181.

89Pan1 Pankajavalli, R., Sreedharan, O.M.: J. Mater. Sci. Lett. **8** (1989) 225.

89Pan2 Pankajavalli, R., Sreedharan, O.M.: J. Mater. Sci. Lett. **8** (1989) 697.

89Pan3 Pankajavalli, R., Sreedharan, O.M.: International Symposium on Thermochemistry and Chemical Processing, IGCAR, Kalpakkam, 1989, p. 2.

90Aza1 Azad, A.M., Sreedharan, O.M.: J. Mater. Sci. Lett. **9** (1990) 304.

90Aza2 Azad, A.M., Sudha, R., Sreedharan, O.M.: J. Less-Common Met. **166** (1990) 57.

90Mal1 Mallika, C., Sreedharan, O.M.: J. Nucl. Mater. **170** (1990) 246.

90Mal2 Mallika, C., Sreedharan, O.M.: J. Less-Common Met. **162** (1990) 51.

90Pan1 Pankajavalli, R., Sreedharan, O.M.: J. Nucl. Mater. **172** (1990) 151.

90Pan2 Pankajavalli, R., Sreedharan, O.M.: J. Nucl. Mater. **175** (1990) 194.
91Aza1 Azad, A.M., Sudha, R., Sreedharan, O.M.: Mater. Res. Bull. **26** (1991) 97.
91Aza2 Azad, A.M., Sreedharan, O.M., Jacob, K.T.: J. Mater. Sci. **26** (1991) 3374.
91Jac Jacob, K.T., Swaminathan, K., Sreedharan, O.M.: Electrochim. Acta **36** (1991) 791.
91Mal1 Mallika, C., Sreedharan, O.M.: J. Alloys Compounds **177** (1991) 273.
91Mal2 Mallika, C., Sreedharan, O.M.: Thermochim. Acta **190** (1991) 217.
91Pan Pankajavalli, R., Sreedharan, O.M.: J. Less-Common Met. **171** (1991) 249.
92Mal Mallika, C., Sreedharan, O.M.: J. Nucl. Mater. **186** (1992) 250.
92Pan Pankajavalli, R., Sreedharan, O.M.: Natl. Workshop on Chemical Sensors, SAEST, Kalpakkam, IGCAR, 1992.
93Mal Mallika, C., Sreedharan, O.M.: J. Alloys Compounds **191** (1993) 219.
94Mal1 Mallika, C., Sreedharan, O.M.: J. Alloys Compounds **216** (1994) 47.
94Mal2 Mallika, C., Sreedharan, O.M.: J. Nucl. Mater. **211** (1994) 123.
94Pan Pankajavalli, R., Sreedharan, O.M., Gnanamoorthy, J.B.: J. Nucl. Mater. **217** (1994) 104.
95Mal1 Mallika, C., Rama Rao, G.V., Sreedharan, O.M., Varadaraju, U.V.: Mater. Sci. Eng. B **33** (1995) 96.
95Mal2 Mallika, C., Sreedharan, O.M.: Mater. Lett. **22** (1995) 5.
95Pan2 Pankajavalli, R., Sreedharan, O.M.: Mater. Lett. **24** (1995) 247.
96Pan Pankajavalli, R., Sreedharan, O.M.: J. Mater. Sci. **31** (1996) 3137.
96Sub Subasri, R., Sreedharan, O.M., Bhat, N.P.: Trans. SAEST **31** (1996) 59.
96Swa1 Swaminathan, K., Jacob, K.T., Sreedharan, O.M.: 50th Annual Technical Meeting, Ind. Inst. Metals, New Delhi, 1996, B1-0-01.
96Swa2 Swaminathan, K., Jacob, K.T., Sreedharan, O.M.: 50th Annual Technical Meeting, Ind. Inst. Metals, New Delhi, 1996, B1-P-03
97Pan1 Pankajavalli, R., Sreedharan, O.M.: Proceedings of the Symposium on Localized Corrosion and Environmental Cracking, IGCAR, Kalpakkam, 1979, C55.
97Pan2 Pankajavalli, R., Sreedharan, O.M., Gnanamoorthy, J.B.: J. Nucl. Mater. **250** (1997) 53.
97Sre Sreedharan, O.M.: private communication, 1997.
97Sub1 Subasri, R., Sreedharan, O.M.: Mater. Lett. **30** (1997) 289.
97Sub2 Subasri, R., Sreedharan, O.M.: Solid State Ionics **93** (1997) 341.
97Sub3 Subasri, R., Pankajavalli, R., Sreedharan, O.M.: Physica C **281** (1997) 85.
98Sub1 Subasri, R., Sreedharan, O.M.: J. Alloys Compounds **274** (1998) 153.
98Sub2 Subasri, R., Pankajavalli, R., Sreedharan, O.M.: J. Alloys Compounds **269** (1998) 71.
99Swa1 Swaminathan, K., Sreedharan, O.M.: J. Nucl. Mater. **275** (1999) 225.
99Swa2 Swaminathan, K., Sreedharan, O.M.: J. Alloys Compounds **292** (1999) 100.
01Aru Arul Anthony, S., Swaminathan, K., Nagaraja, K.S., Sreedharan, O.M.: J. Alloys Compounds **322** (2001) 113.
01Mal Mallika, C., Edwin Suresh Raj, A.M., Nagaraja, K.S., Sreedharan, O.M.: Thermochim. Acta **371** (2001) 95.
01Swa Swaminathan, K., Sreedharan, O.M.: J. Mater. Sci. **36** (2001) 1031.
02Aru Arul Antony, S., Pankajavalli, R., Nagaraja, K.S., Sreedharan, O.M.: Mater. Lett. **57** (2002) 469.
02Sub Subasri, R., Mathews, T., Swaminathan, K., Sreedharan, O.M.: J. Nucl. Mater. **300** (2002) 237.
03Sub1 Subasri, R., Mathews, T., Swaminathan, K., Sreedharan, O.M.: J. Alloys Compounds **354** (2003) 193.
03Sub2 Subasri, R., Mallika, C., Mathews, T., Sastry, V.S., Sreedharan, O.M.: J. Nucl. Mater. **312** (2003) 249.
03Swa Swaminathan, K., Sreedharan, O.M.: J. Alloys Compounds **358** (2003) 48.

3 Electrode potentials of zero charge

3.1 Thermodynamics of an electrode in contact with an electrolyte solution

When a metal used as an electrode [1]) is brought into contact with an ion-containing electrolyte solution or a molten salt acting as a liquid electrolyte, a double layer is established at the phase boundary (c.f. chapter 1, [10Gou1, 10Gou2] for reviews see [86Tra, 86Vor, 97Kor, 99Rit]). Depending on the ratio of the electrochemical potentials of the involved species (ions, electrons etc.) in both phases starting from the initially present nonequilibrium the composition of both phases at the interface is changed by dissolution or deposition of metal atoms until equilibrium is reached. This results in an excess charge on the metal compensated by an excess charge of ions of the opposite charge on the solution side in close distance to the electrode (i.e. inside the double layer). In case of electrode/solution systems where no such charge transfer is possible the unequal charge distribution will be established without metal dissolution or deposition. The local electroneutrality in the double layer results in an uneven distribution of the ions in the solution phase; nevertheless electroneutrality in the solution remains unchanged. The charge q_m on the metal is thus proportional to the excess of ions in the double layer according to $q_m = -F\sum_i \Gamma_i z_i$ with z_i being the ionic charge number, F the Faraday constant and Γ_i the surface excess. The charge q_m on the metal is equal to the charge $-q_s$ on the solution side with opposite sign $q_m = -q_s$. Assuming an ideally polarizable interface, i.e. an electrode where no charge transfer occurs upon immersion into the solution, the Gibbs adsorption isotherm can be written at constant pressure and temperature, $ds = -q_m dE + \sum_i \Gamma_i d\mu_i$.

The excess charge density on the metal is thus $\left(\frac{\partial\sigma}{\partial E}\right)_{T,p,\mu_i} = -q_m$ with μ_i being the chemical potential of species i, E the electrode potential and σ the interfacial tension. The interfacial tension can be measured as a function of the electrode potential by using a suitable reference electrode. The potential, where $\left(\frac{\partial\sigma}{\partial E}\right)_{T,p,\mu_i} = 0$, corresponds to the situation, where there is no excess charge on the electrode, the value of the electrode potential is called potential of zero charge E_{pzc} (this term was introduced 1928 [28Fru2]). If it is assumed, that there are no specific interactions between molecules or any other species dissolved in the electrolyte solution and the metal surface, the corresponding electrode potential should be called

[1]) The term „electrode“ is used primarily to designate an electronically (electron or hole) conducting material. According to W. Nernst the term „electrode“ should be used to name an electronically conducting materal in combination with a specific ionically conducting material. Although this definition is convincing it is not generally accepted. Thus within this chapter „electrode“ will be used in the generally accepted meaning.

differently, according to [63Ant, 66Lor] E_N. Although this distinction is certainly valid, most values reported below have not been discussed explicitly having this distinction in mind. A similar complication caused by the specific adsorption or chemisorption of oxygen and in particular hydrogen (at e.g. platinum group metals) has been discussed extensively by Frumkin and Petrii resulting in the definition of the potentials of zero total charge and zero free charge E_{pztc} and E_{pzfc} [75Fru]; see also [70Fru]; for a review see [99Tra]. Nevertheless most values have been determined under conditions, where these interactions (i.e. chemisorption of hydrogen in particular) are unlikely, thus E_{pzfc} is termed generally E_{pzc}. At first glance E_{pzc} may be considered as a specific property of the metal alone like e.g. the electronic work function of a metal at the vacuum/metal interface. Since the properties of the ionic species in the solution have to be taken into account, also the actual value of E_{pzc} depends on the composition of the solution. The conceivable relationship between E_{pzc} and the work function has been discussed [53Vas, 68Boc1, 71Tra, 77Tra, 76Ard] as well as the general importance of the electron work function in electrochemistry [77Tra]. Particular attention has been paid to these relationships for selected metals with different crystallographic orientations and all other parameters kept unchanged [82Lec]. It was observed, that an increase in the density of atoms in the topmost layer, which is accompanied by an increase in the relative surface energy, corresponds to a more positive value of E_{pzc}. General trends between crystal face orientation and interfacial parameters have been considered [88Bor2, 92Tra2, 99Tra]. A relationship between the enthalpy ΔH of the metal lattice, the free heat content and E_{pzc} has been derived [62Kho]. Further relations between material properties like density, atomic weight etc. and E_{pzc} have been discussed [55Uks2], a more general review of relationships between the nature of the metal surface and the value of E_{pzc} has been given elsewhere [67Fru]. Correlations between the passivation potential and E_{pzc} have been reviewed with particular attention to iron alloys and their components [67Gri1]. Attempts have been made to compare values of E_{pzc} with surface charge-potential information deduced from work function-composition data for the respective metallic surfaces as obtained with ultra-high vacuum surface analytical techniques (electrode emersion) and charge displacement data obtained during CO adsorption. A comparison focussing on the Pt(111) surface has been reported [98Wea]. Despite the considerable uncertainties seemingly involved in this approach these data are included in the following tables.

The influence of the temperature on the value of E_{pzc} has been discussed in detail [56Ran, 57Gra, 66But1, 72Pan1, 87Bac, 88Ham, 90Sil, 96Sil, 99Tra]. It is related to the structure of the electrochemical double layer, in particular to the extent of specific anion adsorption. With molten salt electrolytes the temperature dependence of values of the E_{pzc} has been attributed to shifts of the reference electrode potential [66Pri]. The influence of the pH-value on E_{pzc} has been discussed in detail [56Khe, 57Khe] The relationship between E_{pzc} and the kinetics of electrode processes, in particular the rate of the electron transfer in electrochemical reactions, has been pointed out [55Fru]. It has been the basis for a method to determine the value of E_{pzc}. The influence of the surface roughness important in particular at solid metal electrodes has been discussed extensively together with a listing of suitable methods for surface roughness determination [99Tra]. The influence of steps and kinks on values of E_{pzc} has been discussed based on a theoretical model [03Iba].

A review of recent progress in the understanding of E_{pzc} has been provided elsewhere [99Tra], numerous listings of values of E_{pzc} have been published starting with [33Fru]. Apparently no comprehensive review of data has been published after [69Per], more recently only a selective review has been published [99Tra]. Relationships between exchange current densities and E_{pzc} have been discussed [70Khe].

3.2 Determination of the electrode potential of zero charge

Various methods have been employed for the determination of E_{pzc} of liquid and solid metals. Besides purely electrochemical ones (e.g. measurement of the differential double layer capacity (see also chapter 4.2)) further techniques have been used for the investigation of the surface tension at the solid/electrolyte solution phase boundary. The employed methods can be grouped into several families based on the meas-

urement principle employed. Various methods are aimed at determining the interfacial tension or its change. They are generally regarded as electrocapillary methods. A small group of methods particularly often employed in studies of solid electrode materials employs the measurement of the interfacial capacity of the electrochemical double layer; they are summarized as capacitance methods. A broad variety of methods is based on measurements made during exposure of a fresh electrode surface to the electrolyte solution or molten electrolyte. These methods are briefly reviewed in the following section. Further details and more thorough reviews can be found in the literature [55Fru, 67Arg, 69Per, 72Con, 72Pay, 79Fru, 80Fru].

Electrocapillary and surface tension methods

The interfacial tension of an electrode surface being in contact with an electrolyte solution or molten electrolyte is related to the electrode potential and the charge on the electrode in a direct way:

$$\mathrm{d}\gamma = -q\,\mathrm{d}E - \sum_{\mathrm{i}} \Gamma_{\mathrm{i}}\,\mathrm{d}\mu_{\mathrm{i}} = -\sigma\,\mathrm{d}E - RT\sum_{\mathrm{i}} \Gamma_{\mathrm{i}}\,\mathrm{d}\ln a \tag{3.1}$$

with γ being the surface tension, $q = \sigma$ the (surface) charge density, a_{i} the chemical potential of component i. This makes measurements of the interfacial tension of liquid electrodes an attractive route towards the knowledge of E_{pzc}. Based on experimental evidence [96Hai, 97Iba1, 97Iba2, 98Hai] it has been pointed out, that interfacial tension and surface stress should not be regarded as equal [98Lip, 00Sie] (see also [99Tra]). Consequently measurements of surface stress employed to identify E_{pzc} should be considered with care. Since this debate is still open, no critical evaluation of the respective data collected in the tables is attempted.

Electrocapillary curves

For a liquid metal in particular the value of E_{pzc} can be measured most precisely by identifying the electrode potential, where no current flows during a change of the electrode surface area. This requires the absence of any Faradaic reaction causing a charge transfer at an appreciable rate in the investigated range of electrode potentials. Electrocapillary curves, i.e. relationships between the surface tension γ and the electrode potential E, can be obtained most directly with a capillary electrometer as described originally by Lippmann [73Lip, 75Lip], for a thorough review see [28Fru3, 47Gra]. With this instrument the electrolyte solution/electrode interface is made inside a thin capillary tube. The height of the metal column necessary to maintain the phase boundary at a given level is proportional to the interfacial tension. Calibration is performed by using solutions of known properties. Since the value of E_{pzc} is directly related to the maximum of the electrocapillary curve, calibration is not essential. Various designs of instruments have been reviewed [54Par, 52Gra]. The capillary can be substituted by a dropping metal electrode where the weight of the metal drop is used as a measure of the interfacial tension [03Kuc]. Using a sessile-drop Gouy measured the surface tension of the mercury/liquid interface [16Gou, 17Gou]. The same information can be obtained by measuring the drop time as a function of the electrode potential [49Gra]. (Data obtained with these methods are labelled EC).

Contact angle method

If a gas bubble adheres to an electrode surface being in contact with an electrolyte solution, the contact angle θ can be measured as an indicator of the interfacial tension and its change. The respective relationship is $\cos\theta = (\gamma_{\mathrm{g,m}} - \gamma_{\mathrm{s,m}})/\gamma_{\mathrm{g,s}}$ with g, s, m referring to the gas, solution and metal phase respectively. It was initially observed by Möller, that θ changes with E [08Möl]. Assuming that $\gamma_{\mathrm{g,s}}$ and $\gamma_{\mathrm{g,m}}$ do not depend on the electrode potential a plot of relationship follows:

$$\frac{\mathrm{d}\gamma_{\mathrm{g,s}}}{\mathrm{d}E} = -\gamma_{\mathrm{s,m}}\frac{\mathrm{d}\cos\theta}{\mathrm{d}E} \tag{3.2}$$

Thus an electrocapillary curve can be obtained. The assumption of $\gamma_{g,m}$ being independent from the electrode potential has been questioned by Frumkin et al. [32Fru]. Although closer inspection of the influence of the potential resulted in the conclusion, that the maximum of θ occurs at E_{pzc} (for further disscussions see [69Per]), the method has been used infrequently.

The contact angle between electrode and electrolyte solution can be determined using a solid and partially emersed electrode and observing the meniscus rise [68Mor, 69Mor, 71Mor]. (Data obtained with these methods are labelled CA).

Alternatively the capillary rise can be measured using a bulk acoustic wave sensor [97Che]. (Data obtained with these methods are labelled BAW).

Ribbon extension method

The extension of a metal ribbon made of the electrode material hanging as a working electrode in the electrolyte solution under investigation has been measured [69Bec, 76Lin, 79Bec]. In the used experimental setup the change of the tension of the ribbon held under slight tension is recorded. A change in ribbon length is directly related to a change of the surface stress. This in turn can be used to calculate changes of the surface free energy (surface tension) according to the Shuttleworth equation. The maximum in the surface tension vs. electrode potential curve is equal to the value of E_{pzc} as derived from the Lippmann equation. (Data obtained with this method are labelled R).

Surface stress measurement method

The piezoelectric technique determines the surface tension (surface stress) of an electrode [66Gok, 70Gok, 76Gok, 79Mal1, 79Mal2, 86Seo]. Whereas no difficulties are encountered in measuring the surface tension of liquid electrodes (see above), measurements on solid materials are more difficult. In this technique a thin film of the material under study is glued to a piezoelectric element. The electrode is brought into contact with the electrolyte solution. A small sinusoidal voltage is superimposed to the slow DC-ramp applied to the electrode. This voltage dE causes a corresponding change of the surface stress dγ of the electrode and subsequently a response from the piezoelectric element. Recording the amplitude of the signal of the piezoelectric element and its phase angle with respect to the perturbation as a function of the electrode potential results in diagrams showing a sharp minimum of the amplitude and a change of the phase angle by 180° at E_{pzc}. The use of the cantilever of an atomic force microscope as the sensing device has been reported [95Rai]. The bending of a glass sheet coated with a thin layer of the electrode metal to be investigated acting as the working electrode has been measured using a laser-beam deflection apparatus [99Uen]. (Data obtained with this method are labelled PE). Interfacial tension at the solid electrode-solution interface can also be measured with a thickness shear mode (TSW) bulk acoustic wave sensor [96Jin]. (Data obtained with this method are labelled BAW).

Film electrode bending method

As already discussed with the ribbon extension method a change of the surface tension can result in changes of the mechanical dimensions of the electrode. Besides direct measurements of a change of the length and a more indirect approach via the deformation of a piezoelectric element the bending of a thin film electrode exposed to the electrolyte solution on one side only can be utilized [71Fre]. This concept takes into account earlier observations of spontaneous bending of InSb wafers attributed to differences in interfacial tension [64Spa]. Although these measurements yield only differential values of the surface tension $\Delta\gamma$ according to

$$\frac{1}{2RT}\left(\frac{\partial\Delta\gamma}{\partial\ln a_{\pm}}\right)_{V,\,T,\,p,\,\ldots} = \Delta\Gamma \tag{3.3}$$

proper selection of reference states allows the determination of the desired information on the interface. With this technique the electrode material under study is coated by vapor deposition as a thin layer of approximately 100 nm on a glass strip. A change of the electrode potential applied to the metal film being in contact with an electrolyte solution results in a change of the interfacial tension $\Delta\gamma$, consequently the electrode bends. For small bending the curvature R is given by $R = EI/\boldsymbol{M}$ with E being Youngs modulus of the glass substrate, $\boldsymbol{M}$ being the bending moment as given by $\boldsymbol{M} = 0.5\ \delta\Delta\gamma d$ (d is the width, δ is the thickness of the electrode) and $I = (1/12)\delta^3 d$. The angular deflection AD of an electrode with the length of l is $AD = \Delta\gamma 6\ l\ /E\delta^2$. A plot of the differential surface tension thus measured by means of a suitable optical device can be plotted as a function of the electrode potential etc.. The maximum of the curve corresponds to the value of E_{pzc}. The measurement of changes in specific surface energy via Koesters laser interferometry as a function of electrode potential yielding electrocapillary curves has been suggested [97Heu] (Data obtained by this method are designated EB).

Capacitance methods

Measurement of the differential capacitance $C = \mathrm{d}q/\mathrm{d}E$ of the electrode/solution interface as a function of the electrode potential E results in a curve representing the influence of E on the value of C. The curves show an absolute minimum at E_{pzc} indicating a maximum in the effective thickness of the double layer as assumed in the simple model of a condenser [39Fru]. C is related to the electrocapillary curve and the surface tension according to $C = \mathrm{d}^2\gamma/\mathrm{d}E^2$. Certain conditions have to be met in order to allow the measured capacity of the electrochemical double to be identified with the differential capacity (see [69Per]). In dilute electrolyte solutions this is generally the case.

After initial difficulties in applying this method to a hanging mercury drop [60Fru] this method has been successfully adapted to a dropping mercury electrode [41Gra] and subsequently to solid electrodes [52Bre, 51Dos1, 51Dos2, 51Dos3, 51Dos4]. Experimental setups for measuring C can be classified into AC-techniques and pulse techniques [61Dam]. In early experiments the value of C was determined by comparing the AC-potential drop across an appropriately designed electrochemical cell including the electrochemical phase boundary under investigation with a calibration circuit containing a capacitor of known value [36Bor]. By comparing the potential drops the value of the double layer capacitance was determined. Using an impedance bridge Grahame [53Gra] determined the interfacial capacity. Since this setup is essentially a modified Wheatstone bridge, measurements are time consuming. More recently the setup of choice is a standard arrangement as used for cyclic voltammetry. The DC potential ramp applied to the electrode under investigation is modulated by adding a small amplitude AC voltage (a few millivolts, frequency ten to twenty Hz). The AC current flowing across the interface is detected by means of a frequency selective amplifier. Its rectified value can be recorded as a function of the electrode resulting in a C vs. E curve. Calibration of the capacity axis is done by inserting standard capacitors instead of the electrochemical cell. The AC current is directly proportional to C provided that the cell resistance is small compared to the resistive component of the interfacial impedance. This can be accomplished by a proper design of the cell and the electrode under investigation. Use of a phase sensitive amplifier suggested for detecting only the capacitive out-of-phase component has been suggested but found to be not necessary [81Fle]. Critical reviews of the method have been reported [73Lei]. (Data obtained with this method are labelled T).

In some cases the cyclic voltamogram itself recorded when ramping the electrode potential as described above shows a characteristic minimum around E_{pzc} (e.g. with single crystal gold electrodes [86Kol, 96Ham2]) (Data obtained with this method are labelled CV).

Instead of applying an AC potential modulation pulse methods have been used infrequently. If a direct current pulse I of sufficiently short duration is applied in the absence of a Faradaic reaction, the transferred charge will be entirely consumed for charging the double layer. From the electrode potential vs. time slope the capacity can be calculated according to $C = I/(\mathrm{d}V/\mathrm{d}t)$. Depending on the speed of any conceivable competitive Faradaic reaction the pulse can be generated by simple conventional circuitry, by square wave polarography, single pulses [58Pop, 59McM, 52Lov, 62Sch1] or triangular waves [47Bor]. A method employing potentiodynamic pulses has been described [78Brz]. An approach based on galvanostatic transients has been reported [59Mat3]. (Data obtained with this method are labelled PC).

The presence of a Faradaic electrode reaction of any kind competing with the double layer charging presents a problem in determining the purely capacitive current needed to calculate the surface charge. From a plot of $It^{1/2}$ vs. $I^{1/2}$ (I = total electrode current) with a fixed concentration of the ions of the electrode metal dissolved in solution, the surface charge can be obtained [65But1]. (Data obtained with this method are labelled TC).

Fresh-surface-exposure methods

If an electrode is brought into contact with an electrolyte solution or a molten electrolyte, the establishment of the electrochemical double layer will be accompanied by a transfer of electrical charge. In a suitable arrangement this charge can be measured as an external current. If the contact is made in a way which adjusts the electrode potential upon immersion exactly to the value of E_{pzc}, the current will be nil. Various methods briefly described below have been devised to detect exactly this situation.

Scratching and streaming electrode techniques

If it were possible to bring a clean, bare metal surface instantaneously into contact with an electrolyte solution, the potential difference between the metal and the electrolyte solution should change from zero (before contact is made) to a certain value caused by the formation of the electrochemical double layer and any Faradaic reaction. The change of this value being measurable as the electrode potential to a reference electrode versus time (relaxation time) would be indicative of the nature and the relative rate of these processes. If possible at all, in practice a finite time is needed to make the contact or to renew the electrode surface of an electrode being continuously in contact with the solution. If the rate of the double layer formation or reconstitution is fast compared to any Faradaic charge transfer process, the electrode potential observed immediately after the first, fast relaxation should correspond to the electrode potential of zero charge. Subsequent charge transfer processes will cause further potential shifts and a continuing relaxation.

In a different approach to obtain a self-renewing electrode a steady stream of the metal is again used as an electrode [49Gra]. If a jet of the streaming metal is used as an electrode, its length varies as a function of the electrode potential showing a minimum at E_{pzc} [50Riu]. Such an experimental approach is verified with a streaming mercury electrode and a solution of e.g. NaCl or KCl. The electrode potential measured with the streaming mercury showing a minimum of length corresponds to the value of E_{pzc}. If the flow is stopped, a further relaxation of the electrode potential is observed and the finally established electrode potential in the system described would correspond to the potential of a calomel electrode. A further experimental development has been proposed [84Cza].

Obviously this method is limited to liquid metals like mercury and gallium and their amalgams respectively alloys. Modifications of this method have been reported [86Hor]. At elevated temperatures with molten salt electrolytes alloys with an appropriately low melting point can be investigated, too.

If components of the solution phase are prone to electrochemical reactions (e.g. reduction of dissolved oxygen, reduction of oxidising anions) their presence may also cause Faradaic reactions and the subsequent establishment of an electrode potential different from E_{pzc}.

At first glance such a procedure seems to be highly theoretical for solid metals. The development of the scratching technique has provided an experimental approach to the ideal described above [64And, 66Bod, 67Bod, 67Per, 68And, 69And]. With this technique the electrode metal surface is scratched in situ (i.e. in the presence of an electrolyte solution) with a sapphire or corundum bit. The freshly exposed metal surface in contact with the electrolyte solution establishes a potential which is equivalent to E_{pzc} in the absence of any sufficiently fast Faradaic process. A record of the electrode potential of the continuously scraped metal surface versus time shows an initial value always established after complete renewal of the surface, which relaxes to a different value because of the said Faradaic processes. The initial value can be taken as E_{pzc} after careful consideration of the absence of any interfering Faradaic process. A comparison of the data obtained with this method with data obtained with the immersion method described below has been reported [74Cla2]. Although in both methods essentially the same situation - a

freshly exposed metal surface is exposed to the electrolyte solution - is present, the results tend to differ for various reasons. (Data obtained with this method are labelled S).

Immersion method

If the electrode under investigation is connected to an external voltage source before being immersed into the solution with the second output of the voltage source wired to a reference electrode already immersed in the solution, the charge necessary to establish the electrochemical double layer can be recorded as a transient current at the immersion of the test electrode. If the external voltage is adjusted to a value equal to E_{pzc} this current transient will be nil [62Jak1, 71Jak]. An advanced method taking into account the influence of gas molecules possibly adsorbed on the test electrode which may react upon immersion has been described by Kim [73Kim]. For a careful evaluation of the data and a comparison with results obtained with the open-circuit scrape method see [74Cla2]. (Data obtained with this method are labelled IM).

Miscellaneous methods

Various methods including the repulsion between two electrode wires immersed in a solution, further friction and ultrasonic methods etc., which have not gained practical importance, have been reviewed elsewhere [67Arg]. Some methods, which have been employed to some extent and cannot be fitted easily into one of the groups reviewed above, are briefly treated in the following.

Electron photoemission technique

Under certain conditions the illumination of the electrode/solution interface can result in an electric current. Various reasons can cause this phenomenon. Homogeneous photochemical reactions in the solution may result in electrochemically active particles, which in turn cause the flow of electric current across the interface. Absorption of electromagnetic radiation in the visible and the UV region by the metal electrode can cause the emission of electrons into the solution (photoemission, [65Bar1, 66Bar, 72Bro, 80Gur]). These electrons in turn can cause reactions at the interface termed 'heterogeneous photoreactions'. If the former process of photoemission is the central process, experimental data can be treated based on the theory developed by Brodsky and Gurevich [67Gur, 68Bro1]. At a fixed wavelength of the illuminating light the dependence of the photocurrent I_{pc} on the electrode potential E in concentrated solutions can be described by $I_{pc} = A\,(E_0 - E)^{5/2}$ with E_0 being the threshold electrode potential, where the photocurrent is observed first, and A being a constant depending on the strength of illumination and the nature of the metal. Because of the influence of the potential distribution within the electrochemical double layer at dilute electrolyte concentrations the value of E in the above equation has to be substituted by $E - E_1$ with E_1 representing this influence. Evaluation of dE_1/dE and of $\frac{dI_{pc}^{0.4}}{d(-E)}$ results in a value of E_{pzc} [69Ple, 73Rot]. (Data obtained with this method are labelled P).

Friction method

The mechanical friction between the electrode surface under investigation and a suitable probe depends on the electrode potential showing a maximum at E_{pzc}. Of various experimental setups reviewed previously [69Boc2] the most recent one is used to measure the friction between a test electrode in wire shape and a cylindrical slider sitting on the wire [69Boc1]. Results are interpreted in terms of the repulsion of double layers being present on the wire and the slider [69Boc2]. (Data obtained with this method are labelled F).

Vibrating electrode method

The interface between a liquid metal and an electrolyte solution can be vibrated by applying an oscillating external pressure variation. An electric signal picked up at the oscillating interface as a function of the applied electrode potential can be registered. It shows a particular dependency, at E_{pzc} the signal vanishes [83Mly]. (Data obtained with this method are labelled VE).

Double layer repulsion method

Variations of the repulsion between two electrodes being held at the same electrode potential can be correlated to changes in the diffuse double layer [59Vor, 62Vor, 63Vor]. Assuming that the repulsion is due to charges of the same sign being present on both electrodes the repulsion should pass a minimum at E_{pzc} provided that specific adsorption is absent. This method is analogous to the measurement of the friction (see above). (Data obtained with this method are labelled DR).

Electrode displacement method

Observation of the electrode under examination being exposed to an electric field may yield information about the value of E_{pzc}. Any charge on the electrode, which can be a wire or a colloidal particle, will result in a movement in the external field. Assuming that the movement is due to charges being present on the electrode the rate of the movement should pass through a minimum at E_{pzc} provided that specific adsorption is absent. (Data obtained with this method are labelled ED).

Acousto-electrochemical method

If sound vibrations are applied to an electrode being in contact with a solution, a variable electrode potential component ΔE is generated. Its amplitude depends on the electrode potential E. It shows a maximum at the E_{pzc}. Further details have been provided elsewhere [66Kuk2]. (Data obtained with this method are labelled AE).

Organic adsorption method

Given that certain conditions are met, an organic molecule can be used as a probe of the electrostatic field in the compact Helmholtz double layer. Assuming field/dipole interactions across this layer is the basis for an explanation of the adsorption of neutral organic molecules on the electrode surface. At E_{pzc} this electrostatic field across the double layer arising from excess charge on the electrode is zero. At a given fixed concentration of the organic molecule in the solution phase its concentration on the electrode surface will be independent of the concentration of the supporting electrolyte. Thus potential dependent curves of surface coverage with the organic material at various concentrations of the supporting electrolyte for different supporting electrolyte systems will intersect at the potential corresponding to E_{pzc}.The coverage is determined using a radiotracer technique (see also 4.2). Further details have been reported [63Dah, 63Gre]. (Data obtained with this method are labelled OA).

Ionic adsorption

A correlation between the amount of adsorbed ions and the electrode potential, in particular E_{pzc}, has been identified apparently for the first time by Frumkin and Obrutschewa [26Fru]. A minimum of ionic adsorption was found at E_{pzc}, this is equivalent to the absence of specific adsorption at E_{pzc}. The measurement of the amount of adsorbed ions was performed by measuring the ionic concentration in the solution as a function of the electrode potential or by measuring the surface concentration of adsorbed ions by e.g. radiotracer techniques (see also 4.2). (Data obtained with this method are labelled IA).

Isoelectronic shift of potential

By measuring the hydrogen electrode potential as a function of the concentration of hydrogen ions, i.e. of the pH-value of the electrolyte solution, at a constant charge on the electrode it is possible to calculate the dependence of the adsorption of hydrogen ions on the electrode potential. Thermodynamic treatment of the obtained data allows the calculation of E_{pzc} [66Fru2, 68Fru, 68Pet, 69Pet]. (Data obtained with this method are labelled ISS).

CO charge displacement method

Carbon monoxide CO is known to become strongly adsorbed on various metals both from the gas phase as well as from the liquid phase. When CO is present in an aqueous electrolyte solution it will displace other species less strongly adsorbed on a metal electrode surface of e.g. platinum [92Cla]. Charge vs. time transients recorded at various fixed electrode potentials outside the potential region, where electrooxidation of CO or any other Faradaic process occurs, yields data pertaining to the charge and thus the amount of species desorbed (displaced) during CO adsorption. Transients recorded at various CO-adsorption electrode potentials can be used to identify E_{pzc}. This method (which is sometimes called in abbreviated form 'charge displacement method') has been applied successfully also to the determination of local values of E_{pzc} on metals showing areas with significantly different CO adsorption properties [00Alv]. (Data obtained with this method are labelled CO).

Radiotracer technique

The relationship between the amount of adsorbed particles of a certain type (ion, organic molecule etc.) and the electrode potential can be used to identify E_{pzc}. This has been discussed in detail in the description of methods given above. The central aspect of all methods is the determination of the amount of species actually adsorbed. One method is the radiotracer technique. The species to be adsorbed has to be radioactive. In particular radioactively labelled species emitting β-radiation are employed. The amount of radiation emitted from the electrode is directly proportional to the amount of adsorbed species after careful consideration of the radiation of dissolved species and the thermal background. In order to increase the surface area accessible for adsorption electrodes to be investigated, rough or even porous electrodes with a considerably enlarged ratio of true to geometric surface area can be prepared. Since species in the solution phase as well as in the adsorbed state are both radioactive, the radiation emitted from the dissolved species has to be removed or at least decreased somehow. This can be achieved preferably by using a thin layer technique wherein the electrolyte solution is squeezed away from the electrode as soon as the adsorption equilibrium is attained. Numerous reviews describing details of this method are available [95Hor, 86Zel, 85Hor, 84Kaz, 65Kaz, 65Bal2, 64Boc1, 60Blo]; for general information see also [65Erw]. (Data obtained with this method are labelled RT.)

Salting out method

With some metals the value of E_{pzc} cannot be derived from e.g. tensammetric curves, because the value of E_{pzc} is positive against the standard potential. Consequently Faradaic currents caused by the dissolution of the metal would interfere with the registration of the supposedly purely capacitive current flowing across the interface. Calculation of the charge on the electrode as a prerequisite for the determination of the electrode potential value, where $\sigma = 0$, can be based on the dependence of the cathodic desorption potential E_{des} of an organic substance at a given fixed concentration on the variable concentration of the electrolyte in the electrolyte solution [75Bat, 74Dam]. This dependence shows a maximum, because in a dilute solution the value of E_{des} shifts towards anodic potentials with an increase of the electrolyte concentration because of an increase of the surface charge. At even higher electrolyte concentrations salting out will occur and the value of E_{des} will shift again towards cathodic values. From the observed maximum value of E_{des} it is possible to calculate the corresponding surface charge. Integration of the capacitance vs. the electrode potential curve measured in the range of electrode potentials, where no Faradaic current interferes,

results in a charge vs. electrode potential curve. Extrapolation towards $\sigma = 0$ results in a value of E_{pzc}. (Data obtained with this method are labelled SO). A related approach suggested by Frumkin [74Fru2] is based on the assumption, that the surface charges of metals at the potential of desorption E_{des} of the same organic substance (e.g. *n*-amyl alcohol) at the same concentration are equal. Further calculation is similar to the procedure outlined above. (Data obtained with this method are labelled DE).

Hardness measurements

Various mechanical properties of a metal like e.g. creep, deformation under mechanical stress, hardness, friction (see above) etc. are presumed to be dependent on the electrode potential. If measurements of a selected property as a function of electrode potential reveal an extreme value at a certain electrode potential identical with E_{pzc}, the method can be used for determinations of E_{pzc}. Although there does not seem to exist a firm theory linking the mechanical hardness of a solid with the electrode potential, intuitively a relationship between hardness and surface tension γ seems to be reasonable. In early experiments the mechanical damping of a pendulum resting on the electrode under investigation was used. The rate of change of the pendulum's swing amplitude was recorded as a function of the electrode potential [44Reb]. It was assumed, that the damping effect was directly related to the bulk metal hardness. Later work reviewed by Perkins and Andersen [69Per] related the damping to surface oxides, repulsion of the double layers on the test electrode and sampling device (pendulum etc.) and other effects. Results obtained with the pendulum method are nevertheless still assigned to measurements of surface hardness. The dependence of the erosion of a vibrating electrode immersed in an electrolyte solution containing an abrasive additive (e.g. corundum powder) on the electrode potential was also employed [66Kuk1] (Data obtained with this method are labelled H).

Permanent deformation has also been investigated. The rate of flow of gold wires has been found to depend on the electrode potential with a minimum of the rate at E_{pzc} [51Pfü]. The creep of a lead electrode has been studied [69Lik]. (Data obtained with this method are labelled FL).

Electroreduction of ions

If charged particles are involved in an electrode process, the rate of this process depends on the charge on the electrode surface, since adsorption or desorption of reaction participants is influenced by the charge on the electrode. It has been shown experimentally, that the rate of reaction for the reduction of $S_2O_8^{2-}$ drops sharply in the vicinity of E_{pzc} [55Fru]. The same has been observed for a few other electrode reactions including the hydrogen evolution reaction [69Kil]. More recently similar relationships have been observed for the electroreduction of N_2O [95Att]. From suitable kinetic measurements the value of E_{pzc} can be deduced. (Data obtained with this method are labelled ER).

Gravimetric method

Based on the theory of wetting [85deG] the electrocapillary action at the three phase boundary liquid/gas/partially immersed solid electrode can be understood. Youngs equation [52Goo, 59Joh] describes this situation for a smooth rigid isotropic metal surface: $\cos\theta = (\gamma_{g,s} - \gamma_{s,l})/\gamma_{g,l}$ with g,s,l referring to the gas, solid and liquid phase respectively and θ being the contact angle formed by the metal and the liquid. For a thin metal sheet of weight $m_0 g$ in air attached to one end of a microbalance beam and partially immersed in a liquid it can be demonstrated, that $mg = m_0 g - wdz\rho + 2(w + d)\ \gamma_{g,l}\cos\theta$ with m mass of the sheet, w width and d thickness of the sheet, z depth of immersion, ρ difference in density between gas and liquid, g gravitational constant. Since an accurate measurement of z is difficult, it is desirable to have the term $wdz\rho$ as small as possible in comparison to $2(w + d)\ \gamma_{g,l}\cos\theta$. This can be done by using thin sheets with small d. Accordingly the equation can be rewritten as $\gamma_{g,l}\cos\theta = (m - m_0)g/2w$. A change of the electrode potential E applied to the sheet will result in a change of the contact angle θ and a rise or drop of the meniscus [71Mor] with an accompanying change of the detectable weight of the sheet $(m - m_0)$. Assuming that $\gamma_{g,l}$ does not change with E a plot of $(m - m_0)$ vs. E will result in an electrocapillary curve.

According to Balashova and Frumkin [38Bal] also $\gamma_{g,s}$ can be assumed to be independent from the electrode potential. Thus partial differentiation with respect to E is possible:

$$\gamma \lg \frac{\partial \cos\theta}{\partial E} = -\left(\frac{\partial \gamma_{sl}}{\partial E}\right)_{T,\, p,\, \mu_i} \tag{3.4}$$

The right hand term is equal to the excess surface charge density q_M on the metal electrode of the Lippmann equation. Since Lippmann's equation was originally derived for the liquid metal/solution interface, the reversible work of formation of a unit surface of a solid electrode has to be reconsidered. This work now contains a contribution γ caused by the plastic cleavage and the surface stress Y caused by elastic stretching under equilibrium conditions [86Tra]. Accordingly Lippmann's equation has to be reformulated for the solid metal/liquid interface

$$-\left(\frac{\partial \gamma_{sl}}{\partial E}\right)_{T,\, p,\, m_i} = q_M + (\gamma - Y)\left(\frac{\partial \varepsilon_e}{\partial E}\right)_{T,\, p,\, \mu_i} \tag{3.5}$$

with ε_e being the elastic surface strain. The second term in this equation is not definitely known; nevertheless it has been demonstrated, that the elastic stress Y of a gold ribbon is essentially identical to the interfacial tension γ of an unstrained gold crystal [79Moh]. Thus the second term of the equation can be neglected. Plots of $(m\text{-}m_0)g/2w$ vs. E can be used to obtain q_M, the double layer capacity C_{DL} and the ionic excess $\Gamma_{\pm}$. The experimental setup and further details have been reported [87Mur, 88Mur, 89Mur]. (Data obtained with this method are designated G).

Null solution method

Based on the original hypothesis of the evolution of potential differences at the metal/solution interface proposed by Nernst [97Ner] Palmær devised an experiment which resulted in the value of E_{pzc} [98Pal]. It is based on the observation, that depending on the concentration of dissolved metal ions in the solution these ions will deposit on the metal if their 'osmotic pressure' in the solution is larger than in the metal. If their 'osmotic pressure' is smaller, the metal will be dissolved in part until equilibrium is attained.With an appropriate solution concentration neither deposition nor dissolution will happen. This solution is called null solution. Closer inspection of the situation at the interface in the null solution reveals, that exactly under these conditions there is no net charge on the electrode. This is exactly the condition for the registration of E_{pzc}. (Data obtained with this method are designated NS).

Spectroscopic methods

Various spectroscopic methods applied to studies of the electrochemical interface (spectroelectrochemical methods) provide information about structure and dynamics at this interface. Since e.g. coverage and degree/strength of the adsorbate/metal bonding depend on the potential drop across the interface and consequently on the charge on the metal surface, any spectroscopic information indicating zero charge on the surface will identify E_{pzc}. Sum frequency generation SFG has been employed successfully [95Tad]. (Data obtained with this method are designated SFG). Surface enhanced Raman spectroscopy SERS [90Sob, 92Joa] and interfacial infrared spectroscopy in various methodologies can be employed to probe vibrational spectra of species directly interacting with the electrode surface. The electric field at the electrode/solution interface as well as the charge on the electrode strongly influence position and intensity of vibrational bands. From a comparison of vibrational spectra of the adsorbed species in its free and in its adsorbed form (the latter as a function of applied electrode potential) the situation, where there is no charge on the electrode, i.e. E_{pzc}, can be deduced. (Data obtained with this method are designated VIB). Relationships between the surface plasmon excitation energy and the electrode potential have been stud-

ied [80Tad]. A shift of this energy with E_{pzc} has been found for the Ag(111) surface as compared with a polycristalline silver surface [80Tad]. No practical use of this observation has been reported.

Theoretical approaches

Based on the theory of electron emission from a heterogeneous surface E_{pzc} values have been calculated for two-phase systems (alloys). Calculated data show good agreement with experimental data for the studied alloys [65Gri], an influence of the mode of preparation (electrochemical deposition or melting) was observed. In one example this was attributed to the more developed texture of the alloy, i.e. more crystals of either alloy constituent were present [67Gri1]. Experimental values are tabulated below, calculated ones are not included in this listing. (Data obtained with this method are designated EE).

General remarks on the tables

All values taken from literature sources are given on the standard hydrogen electrode scale. If necessary, conversions were based on standard values for the involved reference electrodes. If a single report contains several values measured with different concentrations of the electrolyte under investigation, all values are included in the table in order to allow insight into the influence of the concentration.
Data are listed according to the alphabetical order of the electrode material. Polycrystalline materials are listed first followed by single crystal samples and amalgam electrodes for every element. Within a given electrode material data are arranged according to increasing concentration of the supporting electrolyte solution.

References

73Lip Lippmann, G.: Ann. Phys. (Leipzig) **149** (1873) 547.
75Lip Lippmann, G.: Ann. Chim. Phys. (Leipzig) **5** (1875) 494.
97Ner Nernst, W.: Z. Elektrochem. **4** (1897) 29.
98Pal Palmaer, W.: Z. Phys. Chem. **25** (1898) 257.
03Kuc Kucera, G.: Ann. Phys. **11** (1903) 529.
08Möl Möller, H.G.: Z. Phys. Chem. (Leipzig) **65** (1908) 226.
10Gou1 Gouy, G.: J. Phys. (Paris) **9** (1910) 457.
10Gou2 Gouy, G.: C. R. Acad. Sci. (Paris) **149** (1910) 654.
16Gou Gouy, G.: Ann. Phys. (Paris) Ser. 9 **6** (1916) 3.
17Gou Gouy, G.: Ann. Phys. (Paris) Ser. 9 **7** (1917) 129.
26Fru Frumkin, A., Obrutschewa, A.: Z. Anorg. Allg. Chem. **158** (1926) 84.
28Fru2 Frumkin, A., Gorodetzkaja, A.: Z. Phys. Chem. **136** (1928) 451.
28Fru3 Frumkin, A.: Ergeb. Exakten Naturwiss. **7** (1928) 235.
32Fru Frumkin, A., Gorodetzkaja, A., Kabanov, B., Nekrassov, N.: Phys. Z. Sowjetunion **1** (1932) 255.
33Fru Frumkin, A.: Phys. Z. Sowjetunion **4** (1933) 239.
36Bor Borisova, T., Proskurnin, M.: Acta Physicochim. URSS **4** (1936) 819.
38Bal Balashova, N., Frumkin, A.: C. R. Acad. Sci. URSS **20** (1938) 449.
39Fru Vorsina, M., Frumkin, A.: C. R. Acad. Sci. URSS **24** (1939) 918.
41Gra Grahame, D.: J. Am. Chem. Soc. **63** (1941) 1207.
44Reb Rebinder, P.A., Venstrem, E.K.: Acta Physicochim. URSS **19** (1944) 36.
47Bor Borisova, T.I., Proskurnin, M.A.: Zh. Fiz. Khim. **21** (1947) 463.
47Gra Grahame, D.C.: Chem. Rev. **41** (1947) 441.
49Gra Grahame, D., Larsen, R.P., Poth, M.A.: J. Am. Chem. Soc. **71** (1949) 2978.
50Riu Rius, A., Marin, A.: An. R. Soc. Esp. Fis. Quim. Ser. A **46** (1950) 93.
51Dos1 Doss, K.S.G., Kalyanasundaram, A.: Current Sci. **20** (1951) 199.
51Dos2 Doss, K.S.G., Agarwal, H.P.: Proc. Indian Acad. Sci. Sect. A **34** (1951) 229.
51Dos3 Doss, K.S.G., Agarwal, H.P.: Proc. Indian Acad. Sci. Sect. A **34** (1951) 263.
51Dos4 Doss, K.S.G., Agarwal, H.P.: Proc. Indian Acad. Sci. Sect. A **33** (1951) 66.
51Pfü Pfützenreuther, A., Masing, G.: Z. Metallkd. **42** (1951) 361.
52Bre Breyer, B., Hacobian, S.: Aust. J. Sci. Res. Ser. A **5** (1952) 500.
52Goo Good, R.J.: J. Am. Chem. Soc. **74** (1952) 5041.
52Gra Grahame, D.C., Coffin, E.M., Cummings, J.I., Poth, M.A.: J. Am. Chem. Soc. **74** (1952) 1207.
52Lov Loveland, J.W., Elving, P.J.: J. Phys. Chem. **56** (1952) 250; 255.
53Gra Grahame, D.: J. Chem. Phys. **21** (1953) 1054.
53Vas Vasenin, R.M.: Zh. Fiz. Khim. **27** (1953) 878.
54Par Parsons, R., in: Modern Aspects of Electrochemistry, Vol. 1, Bockris, J.O'M., Conway, B.E. (eds.), London: Butterworth, 1954.
55Fru Frumkin, A.: Z. Elektrochem. **59** (1955) 807.
55Uks2 Ukshe, E.A., Levin, A.I.: Zh. Fiz. Khim. **29** (1955) 219.
56Khe Kheifets, V.L., Krasikov, B.S.: Dokl. Akad. Nauk SSSR **109** (1956) 586.
56Ran Randles, J.E.B., Whiteley, K.S.: Trans. Faraday Soc. **52** (1956) 1509.
57Gra Grahame, D.C.: J. Am. Chem. Soc. **79** (1957) 2093.
57Khe Kheifets, V.L., Krasikov, B.S.: Zh. Fiz. Khim. **31** (1957) 1992.
58Pop Popat, P.V., Hackerman, N.: J. Phys. Chem. **62** (1958) 1198.
59Joh Johnson, R.E.: J. Phys. Chem. **63** (1959) 1655.
59Mat3 Matsuda, H., Oka, S., Delahay, P.: J. Am. Chem. Soc. **81** (1959) 5077.
59McM McMullen, J.J., Hackerman, N.: J. Electrochem. Soc. **106** (1959) 341.
59Vor Voropaeva, T.N., Deryagin, B.V., Kabanov, B.N.: Dokl. Akad. Nauk SSSR **128** (1959) 981.

60Blo Blomgren, E.A., Bockris, J.O'M.: Nature (London) **186** (1960) 305.
60Fru Frumkin, A.: J. Electrochem. Soc. **107** (1960) 461.
61Dam Damaskin, B.B.: Usp. Khim. **30** (1961) 220.
62Jak1 Jakuszewski, B., Kozlowski, Z.: Rocz. Chem. **36** (1962) 1873.
62Kho Khomutov, N.E.: Zh. Fiz. Khim. **36** (1962) 2721.
62Sch1 Schmid, G.M., Hackerman, N.: J. Electrochem. Soc. **109** (1962) 243.
62Vor Voropaeva, T.N., Deryagin, B.V., Kabanov, B.N.: Kolloidn. Zh. **24** (1962) 396.
63Ant Antropov, L.I.: Ukr. Khim. Zh. **29** (1963) 555.
63Dah Dahms, H., Green, M.: J. Electrochem. Soc. **110** (1963) 1075.
63Gre Green, M., Dahms, H.: J. Electrochem. Soc. **110** (1963) 466.
63Vor Voropaeva, T.N., Deryagin, B.V., Kabanov, B.N.: Izv. Akad. Nauk SSSR, Otd. Khim. Nauk (1963) 257.
64And Andersen, T.N., Perkins, R.S., Eyring, H.: J. Am. Chem. Soc. **86** (1964) 4496.
64Boc1 Bockris, J.O'M., Green, M., Swinkels, D.A.J.: J. Electrochem. Soc. **111** (1964) 743.
64Spa Sparnaay, M.J.: Surf. Sci. **1** (1964) 213.
65Bal2 Balashova, N.A., Kazarinov, V.E.: Usp. Khim. **34** (1965) 1721.
65Bar1 Barker, G.C., Gardner, A.W.: Osnovnye Voprosy Sovremennoi Teoreticheskoi Elektrokhimii, Izd. Mir, Moscow, 1965, p. 118.
65But1 Butler, J.N., Meehan, M.L.: J. Phys. Chem. **69** (1965) 4051.
65Erw Erwall, L.G., Forsberg, H.G., Ljunggren, K.: Radioaktive Isotope in der Technik, Braunschweig: Friedr. Vieweg & Sohn, 1965.
65Gri Grilikhes, M.S., Krasikov, B.S.: Vestn. Leningr. Univ. Ser. 4, Fiz. Khim. **4** (1965) 131.
65Kaz Kazarinov, V.E., Balashova, N.A., Kulesneva, M.I.: Elektrokhimiya **1** (1965) 975.
66Bar Barker, G.C., Gardner, A.W., Sammon, D.C.: J. Electrochem. Soc. **113** (1966) 1182.
66Bod Bodé, D.D.: Ph.D. Thesis, University of Utah, 1966.
66But1 Butler, J.N.: J. Phys. Chem. **70** (1966) 2312.
66Fru2 Frumkin, A., Petry, O., Marvet, R.: J. Electroanal. Chem. **12** (1966) 504.
66Gok Gokhstein, A.Y.: Elektrokhimiya **2** (1966) 1318.
66Kuk1 Kukoz, F.I., Semenchenko, S.A.: Elektrokhimiya **2** (1966) 74.
66Kuk2 Kukoz, F.I., Semenchenko, S.A., Bogdanov, V.I.: Issled. Obl. Khim. Istochnikov Toka (1966) 201.
66Lor Lorenz, W.J., Fischer, F.: Electrochim. Acta **11** (1966) 1597.
66Pri Prisekina, T.N., Kusnetzov, V.A., Maljutina, N.P.: Elektrokhimiya **2** (1966) 1307.
67Arg Argade, S.G., Gileadi, E., in: Electrosorption, Gileadi, E. (ed.), New York: Plenum Press, 1967, p. 87.
67Bod Bodé, D.D., Andersen, T.N., Eyring, H.: J. Phys. Chem. **71** (1967) 792.
67Fru Frumkin, A.N.: J. Res. Inst. Catal. Hokkaido Univ. **15** (1967) 61.
67Gri1 Grigor'ev, N.B.: Zashch. Met. **3** (1967) 357.
67Gur Gurevich, Yu.Ya., Brodsky, A.M., Levich, B.G.: Elektrokhimiya **3** (1967) 1302.
67Per Perkins, R.S.: Ph.D. Thesis, University of Utah, 1967.
68And Andersen, T.N., Anderson, J.L., Bodé, D.D.: J. Res. Inst. Catal. Hokkaido Univ. **16** (1968) 449.
68Boc1 Bockris, J.O'M., Argade, S.E.: J. Chem. Phys. **49** (1968) 5133.
68Bro1 Brodsky, A.M., Gurevich, Yu.Ya.: Zh. Eksp. Teor. Fiz. **54** (1968) 213.
68Fru Frumkin, A., Petry, O., Kossaya, A., Entina, V., Topolev, V.: J. Electroanal. Chem. **16** (1968) 175.
68Mor Morcos, I., Fisher, H.: J. Electroanal. Chem. **17** (1968) 7.
68Pet Petrii, O.A., Kotlov, Yu.G.: Elektrokhimiya **4** (1968) 774.
69And Andersen, T.N., Anderson, J.L., Eyring, H.: J. Phys. Chem. **73** (1969) 3562.
69Bec Beck, T.R.: J. Phys. Chem. **73** (1969) 466.
69Boc1 Bockris, J.O'M., Argade, S.E., Gileadi, E.: Electrochim. Acta **14** (1969) 1259.
69Boc2 Bockris, J.O'M., Argade, S.D.: J. Chem. Phys. **50** (1969) 1622.
69Kil Kilimnik, A.B., Rotinyan, A.L.: Elektrokhimiya **5** (1969) 1234.

69Lik Likhtman, V., Kochanova, L., Leikis, D., Shchukin, E.: Elektrokhimiya **5** (1969) 729.

69Mor Morcos, I.: J. Electroanal. Chem. **20** (1969) 479.

69Per Perkins, R.S., Andersen, T.N.: Modern Aspects of Electrochemistry, Vol. 5, Bockris, J.O'M., Conway, B.E. (eds.), New York: Plenum Press, 1969, p. 203.

69Pet Petrii, O.A., Frumkin, A.N., Kotlov, Yu.G.: Elektrokhimiya **5** (1969) 476.

69Ple Pleskov, Yu.V., Rotenberg, Z.A.: J. Electroanal. Chem. **20** (1969) 1.

70Fru Frumkin, A., Petry, O., Damaskin, B.: J. Electroanal. Chem. **27** (1970) 81.

70Gok Gokhstein, A.Y.: Electrochim. Acta **15** (1970) 219.

70Khe Kheifets, V.L., Krasikov, B.S., Rotinyan, A.L.: Elektrokhimiya **6** (1970) 916.

71Fre Fredlein, R.A., Damjanovic, A., Bockris, J.O'M.: Surf. Sci. **25** (1971) 261.

71Jak Jakuszewsi, B., Kozlowski, Z., Partyka, S., Przasnyski, M., Romanovski, S.: Elektrokhimiya **7** (1971) 804.

71Mor Morcos, I.: J. Colloid Interface Sci. **37** (1971) 410.

71Tra Trasatti, S.: J. Electroanal. Chem. **33** (1971) 351.

72Bro Brodsky, A., Pleskov, Yu.: Progress in Surface Science 2, Davidson, S. (ed.), Oxford: Pergamon, 1972, p. 2.

72Con Conway, B.E.: Techniques of Electrochemistry, Vol. 1, Yeager, E., Salkind, A.J. (eds.), New York: Wiley-Interscience, 1972, p. 389.

72Pan1 Panin, V.A., Rybalka, K.V., Leikis, D.I.: Elektrokhimiya **8** (1972) 390.

72Pay Payne, R.: Techniques of Electrochemistry, Vol. 1, Yeager, E., Salkind, A.J. (eds.), New York: Wiley-Interscience, 1972, p. 43.

73Kim Kim, S.H.: J. Phys. Chem. **77** (1973) 2787.

73Lei Leikis, D.I., Rybalka, K.V., Sevastyanov, E.S., Frumkin, A.N.: J. Electroanal. Chem. **46** (1973) 161.

73Rot Rotenberg, Z., Pleskov, V.Yu.: Elektrokhimiya **9** (1973) 1419.

74Cla2 Clark, G.J., Andersen, T.N., Valentine, R.S., Eyring, H.: J. Electrochem. Soc. **121** (1974) 618.

74Dam Damaskin, B.B., Batrakov, V.V.: Elektrokhimiya **10** (1974) 140.

74Fru2 Frumkin, A.N., Damaskin, B.B., Grigoryev, N.B., Bagotskaya, I.A.: Electrochim. Acta **19** (1974) 69.

75Bat Batrakov, V.V., Damaskin, B.B.: J. Electroanal. Chem. **65** (1975) 361.

75Fru Frumkin, A.N., Petrii, O.A.: Electrochim. Acta **20** (1975) 347.

76Ard Ardizzone, S., Longhi, P., Mussini, T., Rondinini, S.: J. Chem. Thermodyn. **8** (1976) 557.

76Gok Gokhstein, A.Y.: Surface Tension of Solids and Adsorption, Moscow: Nauka, 1976.

76Lin Lin, K.-F., Beck, T.R.: J. Electrochem. Soc. **123** (1976) 1145.

77Tra Trasatti, S.: Advances in Electrochemistry and Electrochemical Engineering, Vol. 10, Gerischer, H., Tobias, C.W. (eds.), New York: John Wiley, 1977, p. 213.

78Brz Brzostowska-Smolska, M., Milkowska, M., Kalinowski, A., Minc, S.: J. Electroanal. Chem. **89** (1978) 389.

79Bec Beck, T.R., Lin, K.-F.: J. Electrochem. Soc. **126** (1979) 252.

79Fru Frumkin, A.: Zero Charge Potentials (Potentsialye Nulevogo Zaryada), Nauka, Moscow, 1979.

79Mal1 Malpas, R.E., Fredlein, R.A., Bard, A.J.: J. Electroanal. Chem. **89** (1979) 171.

79Mal2 Malpas, R.E., Fredlein, R.A., Bard, A.J.: J. Electroanal. Chem. **89** (1979) 339.

79Moh Mohilner, D.M., Beck, T.R.: J. Phys. Chem. **83** (1979) 1160.

80Fru Frumkin, A.N., Petrii, O.A., Damaskin, B.B.: Comprehensive Treatise of Electrochemistry, Vol. 1, Bockris, J'O.M., Conway, B.E., Yeager, E. (eds.), New York: Plenum Press, 1980, p. 221.

80Gur Gurevich, Y.Y., Pleskov, Y.V., Rotenberg, Z.A.: Photoelectrochemistry, New York: Consultants, 1980.

80Tad Tadjeddine, A., Kolb, D.M., Kötz, R.: Surf. Sci. **101** (1980) 277.

81Fle Fleischmann, M., Robinson, J., Waser, R.: J. Electroanal. Chem. **117** (1981) 257.

82Lec Lecoeur, J., Andro, J., Parsons, R.: Surf. Sci. **114** (1982) 320.

83Mly Mlynarczyk, R., Figaszewski, Z., Koczorowski, Z.: Pol. J. Chem. **57** (1983) 1011.

84Cza Czajkowski, J.M., Blaszczyk, T., Kazmierczak, D.: Electrochim. Acta **29** (1984) 439.

84Kaz Kazarinov, V.E., Andreev, V.N.: Comprehensive Treatise of Electrochemistry, Vol. 9, Yeager, E., Bockris, J.O'M., Conway, B.E., Sarangapani, S. (eds.), New York: Plenum Press, 1984, p. 393.

85Hor Horanyi, G.: Wiss. Z. Humboldt-Univ. Berlin, Math. Naturwiss. Reihe **34** (1985) 37.

85deG de Gennes, P.G.: Rev. Mod. Phys. **57** (1985) 827.

86Hor Horanyi, T.S., Takács, M.: J. Electroanal. Chem. **215** (1986) 83.

86Kol Kolb, D.M., Schneider, J.: Electrochim. Acta **31** (1986) 929.

86Seo Seo, M., Makino, T., Sato, N.: J. Electrochem. Soc. **133** (1986) 1138.

86Tra Trasatti, S., Parsons, R.: Pure Appl. Chem. **58** (1986) 437.

86Vor Vorotyntsev, M.A.: Modern Aspects of Electrochemistry, Vol. 17, Bockris, J.O'M., Conway, B.E., White, R.E. (eds.), New York: Plenum Press, 1986, p. 131.

86Zel Zelenay, P., Habib, M.A., Bockris, J.O'M.: Langmuir **2** (1986) 393.

87Bac Bacchetta, M., Francesconi, A., Trasatti, S., Doubova, L., Hamelin, A.: J. Electroanal. Chem. **218** (1987) 355.

87Mur Murphy, O.J., Wainright, J.S.: J. Electrochem. Soc. **134** (1987) 267.

88Bor2 Borkowska, Z., Hamelin, A.: J. Electroanal. Chem. **241** (1988) 373.

88Ham Hamelin, A., Stoicoviciu, L., Dubova, N.M., Trasatti, S.: J. Electroanal. Chem. **244** (1988) 133.

88Mur Murphy, O.J., Wainright, J.S.: J. Electrochem. Soc. **135** (1988) 138.

89Mur Murphy, O.J., Wainright, J.S.: Langmuir **5** (1989) 519.

90Sil Silva, F., Sottomayor, M.J., Hamelin, A.: J. Electroanal. Chem. **294** (1990) 239.

90Sob Sobocinski, R.L., Pemberton, J.E.: Langmuir **6** (1990) 43.

92Cla Clavilier, J., Albalat, R., Gomez, R., Orts, J.M., Feliu, J.M., Aldaz, A.: J. Electroanal. Chem. **330** (1992) 489.

92Joa Joa, S.L., Pemberton, J.E.: Langmuir **8** (1992) 2301.

92Tra2 Trasatti, S.: J. Electroanal. Chem. **329** (1992) 237.

95Att Attard, G.A., Ahmadi, A.: J. Electroanal. Chem. **389** (1995) 175.

95Hor Horanyi, G., Kolics, A.: Sov. Electrochem. (English Transl.) **31** (1995) 941.

95Rai Raiteri, R., Butt, H.-J.: J. Phys. Chem. **99** (1995) 15728.

95Tad Tadjeddine, A., Peremans, A., Guyot-Sionnest, P.: Surf. Sci. **335** (1995) 210.

96Hai Haiss, W., Sass, J.K.: Langmuir **12** (1996) 4311.

96Ham2 Hamelin, A.: J. Electroanal. Chem. **407** (1996) 1.

96Jin Jin-Hua, C., Li-Hua, N., Zhuo, Y.S.: J. Electroanal. Chem. **414** (1996) 53.

96Sil Silva, F., Sottomayor, M.J., Martins, A.: J. Chem. Soc. Faraday Trans. **92** (1996) 3693.

97Che Chen, J.-H., Si, S.-H., Nie, L.-H., Yao, S.-Z.: Electrochim. Acta **42** (1997) 689.

97Heu Heusler, K.E., Lang, G.: Electrochim. Acta **42** (1997) 747.

97Iba1 Ibach, H.: Surf. Sci. Rep. **29** (1997) 193.

97Iba2 Ibach, H., Bach, C.E., Giesen, M., Grossmann, A.: Surf. Sci. **375** (1997) 107.

97Kor Korzeniewski, C., Conway, B.E. (eds.): The Electrochemical Double Layer, Pennington: The Electrochemical Society, 1997.

98Hai Haiss, W., Nichols, R.J., Sass, J.K., Charle, K.P.: J. Electroanal. Chem. **452** (1998) 199.

98Lip Lipkowski, J., Schmickler, W., Kolb, D.M., Parsons, R.: J. Electroanal. Chem. **452** (1998) 193.

98Wea Weaver, M.J.: Langmuir **14** (1998) 3932.

99Rit Ritchie, I.M., Bailey, S., Woods, R.: Advan. Colloid Interface Sci. **80** (1999) 183.

99Tra Trasatti, S., Lust, E.: Modern Aspects of Electrochemistry, Vol. 33, White, R.E., Conway, B.E., Bockris, J.O'M., (eds.) New York: Kluwer Academic/Plenum Publisher, 1999, p. 1.

99Uen Ueno, K., Seo, M.: J. Electrochem. Soc. **146** (1999) 1496.

00Alv Álvarez, B., Climent, V., Feliu, J.M., Aldaz, A.: Electrochem. Commun. **2** (2000) 427.

00Sie Sieradzki, K., Friesen, C., Dimitrov, N.: 198. Electrochemical Society Meeting, Phoenix, Arizona, USA, 22.-27.10.2000, Extended Abstracts.

03Iba Ibach, H., Schmickler, W.: Phys. Rev. Lett. **91** (2003) 016106.

Table 3.1. Electrode potentials of zero charge of metal electrodes in contact with electrolyte solutions.

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Aluminum (Al)	0.01 **M** KCl/H_2O		IM	–0.523	62Jak1
	0.01 **M** LiCl/methanol		IM	–0.623	64Jak2
	0.01 **N** LiCl/methanol		IM	–0.632	65Jak1
	0.01 **M** LiCl/ethanol		IM	–0.751	64Jak2
	0.01 **N** LiCl/ethanol		IM	–0.752	65Jak1
	0.01 **M** LiCl/ethanol		IM	–0.751	64Jak2
	0.01 **N** LiCl/2-butanol		IM	–0.81	65Jak1
	0.01 **N** LiCl/2-propanol		IM	–0.77	65Jak1
	5 m**M** $TBABF_4$/dimethylformamide		T	–2.05	91Saf
	5 m**M** $TBABF_4$/dimethylsulfoxide		T	–1.75	91Saf
Antimony (Sb)	0.01 **N** HCl/H_2O		T	–0.19	63Wu
	$HClO_4$/methanol		T	–0.02	86Pas
	0.002 **N** KF/H_2O		T	–0.14	69Kha1
	KF; NaF/H_2O		T	–0.17	75Pul3
	0.002 **N** KCl/H_2O		T	–0.14	69Kha1
	0.002 **N** $KClO_4/H_2O$		T	–0.15	69Kha1
	0.002 **N** $KClO_4/H_2O$		T	–0.15	69Kha2
	KNO_3/H_2O		T	–0.18	69Hag
	0.002 **N** K_2SO_4/H_2O		T	–0.14	69Kha1
	1 **N** K_2SO_4/H_2O	pH=7	S	–0.11	69And
	0.002 **N** KNO_3/H_2O		T	–0.14	69Kha1
	1 **N** $NaClO_4/H_2O$	pH=1	S	–0.11	69And
	0.002 **N** NaF/H_2O		T	–0.14	69Kha1
	0.002 **N** NaF/H_2O		T	–0.169	94Lus
	1 **N** Na_2SO_4/H_2O		H	–0.2	65Kuk
	1 **N** Na_2SO_4/H_2O		H	–0.2	66Kuk1
Sb (001)	$LiClO_4$/EtOH		T	0.05	91Lus1
	0.0005 **M** NaF+0.0005 **M** HCl/H_2O		T	–0.36	92Lus
	0.001 **M** NaF/H_2O		T	–0.129	97Lus2
	0.002 **N** NaF/H_2O		T	–0.119	94Lus
	0.001 **M** NaF/ethanol		T	0.051	97Lus2
Sb (01$\bar{1}$)	0.001 **M** NaF/H_2O		T	–0.149	97Lus2
	0.002 **N** NaF/H_2O		T	–0.139	94Lus
Sb (111)	$LiClO_4$/EtOH		T	–0.02	91Lus1
	0.0005 **M** NaF+0.0005 **M** HCl/H_2O		T	–0.45	92Lus
	0.001 **M** NaF/H_2O		T	–0.219	97Lus2
	0.001 **M** NaF/ethanol		T	–0.019	97Lus2
	0.002 **M** NaF/H_2O		T	–0.219 [5])	97Lus1
	0.002 **N** NaF/H_2O		T	–0.209	94Lus
	0.002 **M** NaF/H_2O		T	–0.209 [6])	97Lus1
Sb (2$\bar{1}\bar{1}$)	0.001 **M** NaF/H_2O		T	–0.199	97Lus2
	0.002 **N** NaF/H_2O		T	–0.089	94Lus

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Arsenic (As)	not reported		n.a.	+0.2	53Vas
Barium (Ba)	not reported		n.a.	–2.06	53Vas
Beryllium (Be)	not reported		n.a.	–0.88	53Vas
Bismuth (Bi)	0.001 **M** HCl+0.01 **M** KCl/H_2O		P	–0.64	74Rot
	0.001 **N** H_2SO_4/H_2O		T	–0.4	66Pal
	0.001 **M** KCl/H_2O		T	–0.4	66Pal
	0.01 **M** KCl/H_2O		IM	–0.358	62Jak1
	0.001 **M** KF/H_2O		T	–0.625	77Pal
	0.002 **M** KF/H_2O		T	–0.39	69Pal
	0.001 **M** KF/methanol		T	–0.22	72Pet1
	0.001 **M** KNO_3+0.01 **M** KF/H_2O		P	–0.63	74Rot
	0.001 **N** KNO_3+0.01 **N** KF		P	–0.39	74Bro1
	1 **N** K_2SO_4/H_2O	pH=7	S	–0.34	69And
	1 **N** K_2SO_4/H_2O	pH=11	S	–0.64	69And
	0.01 **M** LiCl/methanol		IM	–0.513	64Jak2
	0.01 **M** LiCl/ethanol		IM	–0.529	64Jak2
	0.01 **N** LiCl/1-butanol		IM	–0.67	68Jak1
	0.01 **N** LiCl/2-butanol		IM	–0.66	65Jak1
	0.01 **N** LiCl/2-propanol		IM	–0.609	65Jak1
	0.01 **N** LiCl/1-pentanol		IM	–0.67	65Jak1
	0.01 **N** LiCl/quinoline		IM	–1.184	68Jak1
	LiF/methanol		T	–0.259	71Pet
	0.005 **N** $LiClO_4$/methanol		T	–0.239	79Vaa1
	0.003 **N** $LiClO_4$/ethanol		T	–0.189	75Pal
	0.005 **N** $LiClO_4$/ethanol		T	–0.209	79Vaa1
	0.005 **N** $LiClO_4$/1-propanol		T	–0.209	78Vaa
	0.005 **N** $LiClO_4$/1-propanol		T	–0.209	79Vaa1
	0.001 **N** $LiClO_4$/2-propanol		T	–0.179	80Vae
	0.005 **N** $LiClO_4$/2-propanol		T	–0.179	79Vaa1
	0.005 **N** $LiClO_4$/1-butanol		T	–0.229	79Vaa1
	0.005 **N** $LiClO_4$/iso-butanol		T	–0.239	88Vae
	0.005 **N** $LiClO_4$/2-butanol		T	–0.199	88Vae
	0.005 **N** $LiClO_4$/dimethylsulfoxide		T	–0.239	76Pet
	1 **N** $NaClO_4$/H_2O	pH=7	S	–0.34	69And
	1 **N** $NaClO_4$/H_2O	pH=11	S	–0.34	69And
	0.002 **M** NaF/H_2O		T	–0.354 [5])	97Lus1
	0.002 **M** NaF/H_2O		T	–0.369 [7])	97Lus1
	0.002 **M** NaF/H_2O		T	–0.374 [8])	97Lus1
	0.002 **M** NaF/methanol		T	–0.249 [9])	97Lus1
	0.002 **M** NaF/acetonitrile		T	–0.229 [9])	97Lus1
	0.05 **M** Na_2SO_4/H_2O		T	–0.159	85Lus2
	n.a.		EE	–0.37	67Gri2

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Bi (001)	0.003 **M** KF/H_2O		T	–0.334	85Lus1
	0.001 **M** $LiClO_4/H_2O$		T	–0.414	90Lus
	0.001 **M** $LiClO_4$/acetonitrile		T	–0.169	91Lus2
	0.005 **M** $LiClO_4$/ethanol		T	–0.189	88Ann
	0.001 **M** $NaBF_4/H_2O$		T	–0.399	90Lus
	0.001 **M** NaF/H_2O		T	–0.344	90Lus
	0.001 **M** NaF/H_2O		T	–0.354	97Lus2
	0.001 **M** NaF/acetonitrile		T	–0.169	97Lus2
	0.001 **M** NaF/methanol		T	–0.209	97Lus2
	0.002 **M** NaF/methanol		T	–0.209	97Lus1
	0.001 **M** NaF/ethanol		T	–0.179	97Lus2
	0.001 **M** NaF/*i*-propanol		T	–0.169	97Lus2
	0.05 **M** Na_2SO_4/H_2O		T	–0.199	85Lus2
	0.05 **M** Na_2SO_4/H_2O		T	–0.219	86Lus1
	0.05 **M** Na_2SO_4/H_2O		T	–0.069	86Lus2
Bi $(01\bar{1})$	0.003 **M** KF/H_2O		T	–0.339	85Lus1
	0.001 **M** KF/H_2O		T	–0.359	77Pal
	0.001 **M** $LiClO_4/H_2O$		T	–0.429	90Lus
	0.001 **M** $LiClO_4$/acetonitrile		T	–0.179	91Lus2
	0.001 **M** $LiClO_4$/ethanol		T	–0.199	86Ann
	0.005 **M** $LiClO_4$/ethanol		T	–0.199	88Ann
	0.001 **M** $NaBF_4/H_2O$		T	–0.409	90Lus
	0.001 **M** NaF/H_2O		T	–0.349	90Lus
	0.001 **M** NaF/H_2O		T	–0.349	97Lus2
	0.001 **M** NaF/acetonitrile		T	–0.179	97Lus2
	0.001 **M** NaF/methanol		T	–0.229	97Lus2
	0.001 **M** NaF/ethanol		T	–0.199	97Lus2
	0.001 **M** NaF/*i*-propanol		T	–0.219	97Lus2
	0.05 **M** Na_2SO_4/H_2O		T	–0.149	85Lus2
	0.05 **M** Na_2SO_4/H_2O		T	–0.119	86Lus1
	0.05 **M** Na_2SO_4/H_2O		T	0.001	86Lus2
Bi $(\bar{1}0\bar{1})$	0.003 **M** KF/H_2O		T	–0.334	85Lus1
	0.001 **M** $LiClO_4/H_2O$		T	–0.409	90Lus
	0.001 **M** $LiClO_4$/acetonitrile		T	–0.179	91Lus2
	0.001 **M** $NaBF_4/H_2O$		T	–0.389	90Lus
	0.001 **M** NaF/H_2O		T	–0.334	90Lus
	0.001 **M** NaF/H_2O		T	–0.344	97Lus2
	0.001 **M** NaF/acetonitrile		T	–0.179	97Lus2
	0.05 **M** Na_2SO_4/H_2O		T	–0.099	85Lus2
	0.05 **M** Na_2SO_4/H_2O		T	–0.039	86Lus1
	0.05 **M** Na_2SO_4/H_2O		T	–0.049	86Lus2
Bi (111)	0.0025 **M** KF/H_2O		T	–0.409	74Fru1
	0.001 **M** KF/H_2O		T	–0.409	77Pal
	0.003 **M** KF/H_2O		T	–0.409	85Lus1

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Bi (111)	0.005 **M** KF/H_2O		T	–0.42	74Fru3
	0.005 **M** K_2SO_4/H_2O		T	–0.439	74Fru1
	0.001 **M** $LiClO_4/H_2O$		T	–0.459	90Lus
	0.013 **M** $LiClO_4/H_2O$		T	–0.239	81Ann
	0.001 **M** $LiClO_4$/acetonitrile		T	–0.169	91Lus2
	0.001 **M** $LiClO_4$/ethanol		T	–0.209	86Ann
	0.005 **M** $LiClO_4$/ethanol		T	–0.209	88Ann
	0.001 **M** $NaBF_4/H_2O$		T	–0.424	90Lus
	0.001 **M** NaF/H_2O		T	–0.409	90Lus
	0.001 **M** NaF/H_2O		T	–0.414	97Lus2
	0.001 **M** NaF/acetonitrile		T	–0.259	97Lus2
	0.001 **M** NaF/methanol		T	–0.279	97Lus2
	0.001 **M** NaF/*i*-propanol		T	–0.209	97Lus2
	0.002 **M** NaF/H_2O		T	–0.414 [5])	97Lus1
	0.002 **M** NaF/H_2O		T	–0.409 [6])	97Lus1
	0.002 **M** NaF/H_2O		T	–0.389 [7])	97Lus1
	0.002 **M** NaF/H_2O		T	–0.379 [8])	97Lus1
	0.002 **M** NaF/H_2O		T	–0.384 [10])	97Lus1
	0.002 **M** NaF/H_2O		T	–0.384 [9])	97Lus1
	0.002 **M** NaF/H_2O		T	–0.379 [11])	97Lus1
	0.002 **M** NaF/acetonitrile		T	–0.259 [5])	97Lus1
	0.002 **M** NaF/methanol		T	–0.279 [5])	97Lus1
	0.001 **M** NaF/ethanol		T	–0.209	97Lus2
	0.05 **M** Na_2SO_4/H_2O		T	–0.249	85Lus2
	0.05 **M** Na_2SO_4/H_2O		T	–0.249	86Lus1
	0.05 **M** Na_2SO_4/H_2O		T	–0.019	86Lus2
Bi ($2\bar{1}\bar{1}$)	0.001 **M** KF/H_2O		T	–0.329	77Pal
	0.001 **M** $LiClO_4/H_2O$		T	–0.409	90Lus
	0.001 **M** $LiClO_4$/ethanol		T	–0.219	86Ann
	0.001 **M** $NaBF_4/H_2O$		T	–0.384	90Lus
	0.001 **M** NaF/H_2O		T	–0.334	90Lus
	0.001 **M** NaF/H_2O		T	–0.329	97Lus2
	0.001 **M** NaF/acetonitrile		T	–0.169	97Lus2
	0.001 **M** NaF/ethanol		T	–0.219	97Lus2
	0.001 **M** $LiClO_4$/acetonitrile		T	–0.179	91Lus2
Bi-Cu alloy	n.a.	55 at% Bi	EE	–0.35	67Gri2
	n.a.	30 at% Bi	EE	–0.345	67Gri2
	n.a.	18 at% Bi	EE	–0.33	67Gri2
	n.a.	8 at% Bi	EE	–0.32	67Gri2
	n.a.	4 at% Bi	EE	–0.27	67Gri2
Cadmium (Cd)	0.001 **M** HCl+0.01 **M** KCl/H_2O		P	–1.025	74Rot
	0.011 **N** $K_2S_2O_8/H_2O$		ER	–0.95	52Nik
	0.005 **N** KCl/H_2O		T	–0.92	50Bor
	0.001 **N** KCl/H_2O		T	–0.91	67Ham1
	0.005 **N** KCl/H_2O		T	–0.92	50Bor

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	0.005 **N** KCl/H_2O		T	–0.9	55Fru
	0.005 **N** KCl/H_2O		H	–0.7	55Fru
	0.005 **N** KCl/H_2O		H	–0.70	56Ven
	0.01 **N** KCl/H_2O		T	–0.93	50Bor
	0.01 **N** KCl+0.001 **N** HCl/H_2O		P	–0.8	74Bro1
	1.0 **N** KCl/H_2O		T	–0.95	50Bor
	0.001 **N** $KClO_4/H_2O$		T	–0.91	67Ham1
	0.01 **N** KF/H_2O		T	–0.78 [12])	73Ryb3
	1.5 **M** $LiClO_4$/dimethylformamide		S	–0.63	80Kar
	0.01 **M** NaF/H_2O		T	–0.74 [13])	69Bar1
	0.01 **M** NaF/H_2O		T	–0.75	72Pan1
	0.001 **M** NaF/H_2O		T	–0.72	91Agl
	0.001 **M** NaF/H_2O		T	–0.72	68Bar
	0.001 **M** NaF/H_2O		T	–0.75	72Lei
	0.005 **M** NaF/H_2O		T	–0.803	91Obr
	0.01 **M** NaF/H_2O		S	–0.69	80Kar
	0.02 **M** NaF/H_2O		T	–0.8	72Ham
	0.026 **M** NaF/H_2O		T	–0.77	75Mis
	0.01 **M** $NaNO_3/H_2O$		T	–0.684	80Pan
	0.01 **N** Na_2SO_4/H_2O		T	–0.68	65Gri
	0.01 **M** Na_2SO_4/H_2O		T	–0.73	75Non
	0.01 **M** Na_2SO_4/H_2O		S	–0.66 [14])	75Non
	0.01 **M** Na_2SO_4/H_2O		T	–0.73 [15])	75Non
	0.02 **M** $(Na_2SO_4 + H_2SO_4)/H_2O$	pH=2	T	–0.9	57Khe
	0.02 **M** $(Na_2SO_4 + H_2SO_4)/H_2O$	pH=4	T	–0.9	57Khe
	0.02 **M** $(Na_2SO_4 + H_2SO_4)/H_2O$	pH=6	T	–0.9	57Khe
	0.02 **M** $(Na_2SO_4 + H_2SO_4)/H_2O$	pH=7	T	–0.9	57Khe
	0.02 **M** $(Na_2SO_4 + H_2SO_4)/H_2O$	pH=2	T	–0.7 [16])	57Khe
	0.02 **M** $(Na_2SO_4 + H_2SO_4)/H_2O$	pH=4	T	–0.7 [16])	57Khe
	0.02 **M** $(Na_2SO_4 + H_2SO_4)/H_2O$	pH=6	T	–0.7 [16])	57Khe
	0.02 **M** $(Na_2SO_4 + H_2SO_4)/H_2O$	pH=7	T	–0.7 [16])	57Khe
	0.8 **M** Na_2SO_4/H_2O		T	–0.913	91Obr
Cd (001)	0.001 **M** NaF/H_2O		T	–0.709	97Lus2
Cd (1$\bar{1}$0)	0.001 **M** NaF/H_2O		T	–0.769	97Lus2
Cd (0001)	0.005 **M** $LiClO_4/H_2O$		T	–0.759	96Lus1
	0.005 **M** $NaClO_4/H_2O$		T	–0.784	92Nan
	NaF/H_2O		T	–0.749	89Nan
	0.005 **M** NaF/H_2O		T	–0.709	96Lus1
	0.005 **M** NaF/H_2O		T	–0.729 [6])	96Lus1
	0.005 **M** NaF/H_2O		T	–0.729 [15])	96Lus1
	0.002 **M** NaF/H_2O		T	–0.709 [5])	97Lus1
	0.002 **M** NaF/H_2O		T	–0.729 [6])	97Lus1
	0.002 **M** NaF/H_2O		T	–0.729 [7])	97Lus1
	0.002 **M** NaF/H_2O		T	–0.759 [8])	97Lus1
	0.005 **M** $NaNO_2/H_2O$		T	–0.774	92Nan

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Cd (0001)	0.005 **M** $NaNO_3/H_2O$		T	–0.784	92Nan
Cd (10$\bar{1}$0)	0.005 **M** $LiClO_4/H_2O$		T	–0.799	96Lus1
	0.005 **M** NaF/H_2O		T	–0.759	96Lus1
	0.005 **M** NaF/H_2O		T	–0.739 [8])	96Lus1
Cd (10$\bar{1}$1)	0.005 **M** NaF/H_2O		T	–0.749	96Lus1
Cd (11$\bar{2}$0)	0.005 **M** $LiClO_4/H_2O$		T	–0.799	96Lus1
	0.002 **M** NaF/H_2O		T	–0.779 [5])	97Lus1
	0.002 **M** NaF/H_2O		T	–0.749 [7])	97Lus1
	0.001 **M** NaF/H_2O		T	–0.759	97Lus2
	0.005 **M** NaF/H_2O		T	–0.769	96Lus1
	0.005 **M** NaF/H_2O		T	–0.749 [8a])	96Lus1
Cd (11$\bar{2}$1)	0.005 **M** NaF/H_2O		T	–0.739	96Lus1
Cd amalgam [17])	1 **M** KBr/H_2O	0.0003 % Cd	S	–0.423	30Fru
	1 **M** KBr/H_2O	0.001 % Cd	S	–0.431	30Fru
	1 **M** KBr/H_2O	0.004 % Cd	S	–0.444	30Fru
	1 **M** KBr/H_2O	0.01 % Cd	S	–0.456	30Fru
	1 **M** KBr/H_2O	0.1 % Cd	S	–0.48	30Fru
	1 **M** KBr/H_2O	4.8 % Cd	S	–0.538	30Fru
	1 **M** KCl/H_2O	0.004 % Cd	S	–0.345	30Fru
	1 **M** KCl/H_2O	0.01 % Cd	S	–0.360	30Fru
	1 **M** KCl/H_2O	0.1 % Cd	S	–0.4	30Fru
	1 **M** KCl/H_2O	1 % Cd	S	–0.434	30Fru
	1 **M** KCl/H_2O	4.8 % Cd	S	–0.466	30Fru
	1 **M** KI/H_2O	4.8 % Cd	S	–0.653	30Fru
Calcium (Ca)	not reported		n.a.	–1.7	53Vas
Carbon (C)	1 **N** H_2SO_4/H_2O		CA	0.07	43Chu
	1 **N** KCl/H_2O		CA	0.07	43Chu
	1 **N** KBr/H_2O		CA	0.0 [18])	71Mor
	1 **N** KBr/H_2O		CA	0.07	43Chu
	4 **N** KBr/H_2O		CA	0.0 [18])	71Mor
	1 **N** KI/H_2O		CA	–0.13	43Chu
	1 **N** K_2SO_4/H_2O	pH=7	S	–0.26	69And
	0.5 **N** $NaCl/H_2O$		H	0.25	49Ven
	1 **N** $NaOH/H_2O$		CA	0.07	43Chu
	0.05 **N** Na_2SO_4/H_2O	pH=7	IA	0.05	40Kuc
	1 **N** Na_2SO_4/H_2O		CA	0.22	43Chu
	1 **N** Na_2SO_4+0.005 **M** H_2SO_4/H_2O		IA	0.07	40Kuc
	1 **M** $NaClO_4/H_2O$	2 mol% In	T	–0.359	95Koe
Cerium (Ce)	not reported		n.a.	–1.61	53Vas
Cesium (Cs)	not reported		n.a.	–2.59	53Vas

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Chromium (Cr)	0.05 **M** H_2SO_4/H_2O		CA	–0.45	55Uks1
	0.1 **M** $NaOH/H_2O$		CA	–0.45	55Uks1
Cobalt (Co)	0.01 **N** Na_2SO_4/H_2O	pH=5.6	S	–0.43	78Laz
	0.02 **N** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=1	T	–0.32	57Khe
	0.02 **N** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=1	T	–0.32	57Khe
	0.02 **N** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=2	T	–0.405	57Khe
	0.02 **N** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=3	T	–0.43	57Khe
	0.02 **N** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=3	T	–0.43	57Khe
	0.02 **N** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=4	T	–0.465	57Khe
	0.02 **N** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=5	T	–0.475	57Khe
	0.02 **N** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=5	T	–0.46	57Khe
	0.02 **N** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=6	T	–0.48	57Khe
Copper (Cu)	0.03 **M** Cu_2SO_4+0.05 **M** $Na_4P_2O_7/H_2O$		ER	0.11	53Lev1
	0.1 **M** H_2SO_4/H_2O		OA	–0.052	73Lou
	0.5 **M** H_2SO_4/H_2O		OA	–0.052	73Lou
	0.5 **M** H_2SO_4/H_2O		T	0.101	01Ma
	1 **M** H_2SO_4/H_2O		OA	–0.052	73Lou
	0.0001 **m** KBr/H_2O		IM	0.073 [19])	68Koz
	1 **M** KBr/H_2O		S	–0.74	74Cla2
	0.0001 **M** KCl/H_2O		IM	0.073	62Jak1
	0.0001**M** KCl/H_2O		IM	0.078	64Jak1
	0.001 **M** KCl/H_2O		IM	0.073	62Jak1
	0.001**M** KCl/H_2O		IM	0.075	64Jak1
	0.01**M** KCl/H_2O		IM	0.073	62Jak1
	0.01**M** KCl/H_2O		IM	0.073	64Jak1
	0.1**M** KCl/H_2O		IM	–0.012	62Jak1
	0.1**M** KCl/H_2O		IM	–0.012	64Jak1
	0.1**M** KCl/H_2O	pH=7	S	–0.35	64Per
	0.3 **M** KCl/H_2O		IM	–0.049	64Jak1
	1 **M** KCl/H_2O		IM	–0.122	64Jak1
	1 **M** KCl/H_2O	pH=2.7...11	S	–0.51	74Cla2
	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=5.7	T	–0.73	04Luk
	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=4.2	T	–0.72	04Luk
	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=3.7	T	–0.70	04Luk
	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=3.2	T	–0.64	04Luk
	0.0025 **N** K_2SO_4/H_2O		T	–0.04	70Ego2
	0.01 **M** Cl^-/H_2O		IM	0.075	65Jak1
	0.01**M** LiCl/methanol		IM	–0.072	64Jak2
	0.01**M** LiCl/ethanol		IM	–0.147	64Jak2
	0.01 **N** LiCl/1-butanol		IM	–0.236	68Jak1
	0.01 **N** LiCl/2-butanol		IM	–0.2	65Jak1
	0.01 **N** LiCl/2-propanol		IM	–0.214	65Jak1
	0.01 **N** LiCl/1-pentanol		IM	–0.21	68Jak1
	0.01 **N** LiCl/quinoline		IM	–0.764	68Jak1

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Cu	0.003 **N** $NaCl/H_2O$		T	–0.03	70Ego2
	0.025 **M** $NaClO_4/H_2O$		T	0.025	72Nov2
	0.0001 **M** NaF/H_2O		IM	0.073	62Jak1
	0.0001**M** NaF/H_2O		IM	0.079	64Jak1
	0.001 **M** NaF/H_2O		T	0.05	91Agl
	0.001 **M** NaF/H_2O		IM	0.073	62Jak1
	0.001**M** NaF/H_2O		IM	0.075	64Jak1
	0.001 **M** NaF/H_2O		T	0.09	70Ego1
	0.01 **M** NaF/H_2O		IM	0.073	62Jak1
	0.01**M** NaF/H_2O		IM	0.072	64Jak1
	0.1 **M** NaF/H_2O		IM	0.073	62Jak1
	0.1**M** NaF/H_2O		IM	0.068	64Jak1
	0.3**M** NaF/H_2O		IM	0.073	64Jak1
	1 **M** NaF/H_2O		IM	0.073	62Jak1
	1 **M** NaF/H_2O		IM	0.077	64Jak1
	0.0001 **M** F^-/H_2O		IM	0.075	65Jak1
	0.1 **M** $NaOH/H_2O$		CA	0.05	55Uks1
	0.05 **M** $Na_4P_2O_7/H_2O$		CA	0.15	55Uks1
	0.05 **M** $Na_4P_2O_7/H_2O$		ER	0.11	53Lev1
	0.01 **N** Na_2SO_4/H_2O		T	–0.433	86Dag
	0.01 **N** Na_2SO_4/H_2O		IM	–0.035	95Tur
	0.02 **N** $(Na_2SO_4 + H_2SO_4)/H_2O$	1< pH <7	T	0.025	57Khe
	0.02 **N** $(Na_2SO_4 + NaOH)/H_2O$	7< pH <10	T	0.025	57Khe
	0.02 **N** $(Na_2SO_4 + NaOH)/H_2O$	pH=11	T	0.015	57Khe
	0.02 **N** $(Na_2SO_4 + NaOH)/H_2O$	pH=12	T	0.05	57Khe
	0.01 **M** Na_2SO_4/H_2O		IM	0.189 [20])	90Sok
	0.01 **M** Na_2SO_4/H_2O		IM	–0.03	97Tur
	0.02 **M** Na_2SO_4/H_2O		T	0.03	57Khe
	0.05 **N** Na_2SO_4/H_2O		T	–0.45	86Dag
	0.5 **M** Na_2SO_4/H_2O		OA	–0.053	73Lou
	1 **N** Na_2SO_4/H_2O		F	–0.26	61Sta
	1 **N** Na_2SO_4/H_2O		CA	–0.04	65Bon
	1 **M** Na_2SO_4/H_2O		OA	–0.053	73Lou
	1 **N** Na_2SO_4/H_2O		T	–0.059	59McM
	Na_2SO_4/H_2O		EC	–0.049	65Bon
	0.001 **M** $Na_2SO_4+K_2S_2O_7/H_2O$		ER	0.0	55Fru
	0.01 **M** Na_2SO_4/H_2O+1-propanol	3 %wt	IM	0.12	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	6.3 %wt	IM	0.16	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	9.4 %wt	IM	0.18	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	11 %wt	IM	0.19	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	15.3 %wt	IM	0.21	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	25.1 %wt	IM	0.197	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	32.1 %wt	IM	0.14	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	43.4 %wt	IM	0.165	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	56 %wt	IM	0.155	95Tur

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	0.01 **M** Na_2SO_4/H_2O+1-propanol	63.8 %wt	IM	0.15	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	77 %wt	IM	0.14	95Tur
	n.a.		EC	–0.040 [21])	32Pro
	n.a.		EE	0.03	67Gri2
Cu (100)	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=5.7	T	–0.73	04Luk
	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=4.2	T	–0.72	04Luk
	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=3.7	T	–0.71	04Luk
	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=3.2	T	–0.62	04Luk
	0.01 **M** $NaClO_4/H_2O$		T	–0.723	06Luk
	0.04 **M** $NaClO_4/H_2O$		T	–0.54	85Lec
	0.007 **M** NaF/H_2O		T	–0.049	72Nov1
	0.007 **M** NaF/H_2O		T	–0.050 [22])	73Nov
	0.007 **M** NaF/H_2O		T	–0.075 [23])	73Nov
	0.01 **M** NaF/H_2O		T	–0.05 [23])	73Nov
	0.01 **M** NaF/H_2O		T	–0.05 [24])	73Nov
	0.01 **M** NaF/H_2O		T	–0. 075 [23])	73Nov
	0.05 **M** NaF/H_2O		T	–0.05 [23])	73Nov
	0.05 **M** NaF/H_2O		T	–0.05 [23])	73Nov
	0.05 **M** Na_2SO_4/H_2O		T	–0.459	79Hen
Cu (110)	0.005 **M** $KClO_4/H_2O$		T	–0.689	97For
	0.005 **M** $KClO_4/H_2O$		C	–0.689	97For
	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=5.7	T	–0.75	04Luk
	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=4.2	T	–0.73	04Luk
	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=3.7	T	–0.72	04Luk
	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=3.2	T	–0.66	04Luk
	0.01 **M** $NaClO_4/H_2O$		T	–0.748	06Luk
	0.007 **M** NaF/H_2O		T	–0.074	72Nov1
	0.007 **M** NaF/H_2O		T	–0.124	72Nov1
	0.007 **M** NaF/H_2O		T	–0.075 [22])	73Nov
	0.007 **M** NaF/H_2O		T	–0.125 [23])	73Nov
	0.01 **M** NaF/H_2O		T	–0.075 [23])	73Nov
	0.01 **M** NaF/H_2O		T	–0.125 [23])	73Nov
	0.05 **M** NaF/H_2O		T	–0.025 [24])	73Nov
	0.05 **M** NaF/H_2O		T	–0.125 [23])	73Nov
	0.007 **M** NaF/H_2O		T	–0.124	75Max
	0.01 **M** NaF/H_2O		T	–0.124	78Max
	0.005 **M** Na_2SO_4/H_2O		T	–0.629	79Hen
Cu (111)	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=5.7	T	–0.70	04Luk
	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=4.2	T	–0.68	04Luk
	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=3.7	T	–0.65	04Luk
	0.1 **M** $KClO_4$+ $HClO_4/H_2O$	pH=3.2	T	–0.55	04Luk
	0.01 **M** $NaClO_4/H_2O$		T	–0.698	06Luk
	0.04 **M** $NaClO_4/H_2O$		T	–0.2	85Lec
	0.007 **M** NaF/H_2O		T	–0.014	72Nov1
	0.007 **M** NaF/H_2O		T	–0.125 [23])	73Nov

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Cu (111)	0.007 **M** NaF/H_2O		T	–0.024	75Max
	0.05 **M** NaF/H_2O		T	–0.025 [22])	73Nov
	0.01 **M** NaF/H_2O		T	–0.025 [22])	73Nov
	0.01 **M** NaF/H_2O		T	–0.125 [23])	73Nov
	0.01 **M** NaF/H_2O		T	0.076	78Max
	0.050 **M** Na_2SO_4/H_2O		T	–0.329	79Hen
Gallium [25]) **(Ga)**	0.006 **N** HCl/H_2O		T	–0.700	68Bag
	$HClO_4$, HCl/H_2O		EC	–0.69	66Mor
	$HClO_4$, HCl/H_2O		T	–0.69	66Fru1, 68Bag
	0.1 **M** $HClO_4/H_2O$		TC	–0.645	66But2
	0.006 **N** H_2SO_4/H_2O		T	–0.680 [26])	72Gri4
	0.9 **N** KBr+0.1 **N** HCl/H_2O		S	–0.73	67Per
	0.9 **M** KBr+0.1 **M** HCl/H_2O		S	–0.73	74Cla2
	1 **N** KBr+HCl/H_2O		EC	–0.799	64Fru1
	1 **N** KBr+HCl/H_2O		T	–0.789	64Fru2
	1 **N** KBr+HCl/H_2O		T	–1.03	65Fru1
	1 **N** KBr+HCl/H_2O		EC	–1.04	65Fru1
	1 **M** KCl+HCl/H_2O		EC	–0.62	36Mur
	0.01 **N** KCl/H_2O		EC	–0.96	66Mor
	0.9 **N** KCl+0.1 **N** HCl/H_2O		S	–0.7	67Per
	1 **N** KCl+0.1 **N** HCl/H_2O		EC	–0.6	55Fru
	1 **N** KCl+HCl/H_2O		EC	–0.769	64Fru1
	1 **N** KCl+HCl/H_2O		T	–0.759	64Fru1
	1 **N** KCl+HCl/H_2O		EC	–1.01	65Fru1
	1 **N** KCl+HCl/H_2O		T	–1.0	65Fru1
	1 **N** KCl+0.01 **N** HCl/H_2O		CA	–0.42	34Gor
	1 **M** KCl+HCl/H_2O	pH=1	S	–0.71	74Cla2
	1 **N** KI+HCl/H_2O		EC	–0.869	64Fru1
	1 **N** KI+HCl/H_2O		T	–0.859	64Fru2
	1 **N** KI+HCl/H_2O		EC	–1.11	65Fru1
	1 **N** KI+HCl/H_2O		T	–1.1	65Fru1
	0.1 **M** KOH/H_2O		EC	–0.58	67Pol
	1 **N** K_2SO_4/H_2O	pH=7	S	–0.63	69And
	0.06...0.63 **N** $NaClO_4/H_2O$		T; C	–0.64	93Inn
	0.1 **M** $NaClO_4/H_2O$+ 0.1 **M** n-$C_5H_{11}OH$		T	–0.94	71Fru
	0.1 **M** $NaClO_4/H_2O$		EC	–0.96	66Mor
	0.1 **M** $NaClO_4$/*N*-methylformamide		T	–0.669	96Eme
	0.1 **M** $NaClO_4$/dimethylsulphoxide		T	–1.025	78Bag
	1 **N** $NaClO_4/H_2O$	pH=1	S	–0.63	69And
	1 **N** $NaClO_4$+$HClO_4/H_2O$		EC	–0.659	64Fru1
	1 **N** $NaClO_4$+$HClO_4/H_2O$		T	–0.649	64Fru2
	1 **N** $NaClO_4$+$HClO_4/H_2O$		EC	–0.9	65Fru1
	1 **N** $NaClO_4$+ClO_4/H_2O		T	–0.89	65Fru1

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	1 **N** $NaClO_4$+1 **N** $HClO_4/H_2O$		T	–0.61	65Fru1
	0.01...1 **N** ($NaClO_4$+$HClO_4$)/H_2O		T	–0.64	90Pez
	0.01...0.63 **N** Na_2SO_4/H_2O		T; C	–0.68	93Inn
	0.9 **N** Na_2SO_4+0.1 **N** H_2SO_4/H_2O		S	–0.63	67Per
	0.45 **M** Na_2SO_4+0.05 **M** H_2SO_4/H_2O		S	–0.63	74Cla2
	1 **N** Na_2SO_4+H_2SO_4/H_2O		EC	–0.689	64Fru1
	1 **N** Na_2SO_4+H_2SO_4/H_2O		T	–0.684	64Fru2
	1 **N** Na_2SO_4+H_2SO_4/H_2O		EC	–0.93	65Fru1
	1 **N** Na_2SO_4+H_2SO_4/H_2O		T	–0.925	65Fru1
	0.5 **M** Na_2SO_4+H_2SO_4/H_2O		PC	–0.684	85Pez
	1 **N** Na_2SO_4+0.1 **N** $HClO_4/H_2O$		T	–0.645	65Fru1
Ga amalgam	0.1 **M** KOH/H_2O	0.001 at% Ga	EC	–1.3	67Pol
	0.1 **M** KOH/H_2O	0.004 at% Ga	EC	–1.3	67Pol
	0.1 **M** KOH/H_2O	0.03 at% Ga	EC	–1.0	67Pol
	0.1 **M** KOH/H_2O	0.18 at% Ga	EC	–0.96	67Pol
	0.1 **M** KOH/H_2O	0.36 at% Ga	EC	–0.91	67Pol
	0.1 **M** KOH/H_2O	0.7at % Ga	EC	–0.88	67Pol
	0.1 **M** $HClO_4/H_2O$	89.36 at% Ga	TC	–0.589	66But2
Ga/In alloy	0.001 **M** $HClO_4/H_2O$	83.3 % Ga/ 16.7 % In	T	–0.68	72Gri1
	0.1 **M** $HClO_4/H_2O$	19.2 at% In	TC	–0.666	66But2
	0.1 **M** $NaClO_4/H_2O$+ 0.1 **M** n-$C_5H_{11}OH$	83.3 % Ga/ 16.7 % In	T	–1.02	71Fru
Ga/Tl alloy [27])	1 **N** $NaBr/H_2O$		EC	–0.859	78Bul
	1 **N** $NaBr/H_2O$		T	–0.849	78Bul
	1 **N** $NaCl/H_2O$		EC	–0.779	78Bul
	1 **N** $NaCl/H_2O$		T	–0.769	78Bul
	1 **N** $NaClO_4/H_2O$		EC,T	–0.759	78Bul
	0.1 **M** $NaClO_4$/*N*-methylformamide		T	–0.639 [28])	96Eme
	1 **N** NaI/H_2O		EC	–0.959	78Bul
	1 **N** NaI/H_2O		T	–0.959	78Bul
	1 **N** Na_2SO_4/H_2O		EC,T	–0.709	78Bul
Germanium (Ge)	0.1 **N** HCl/H_2O	p- and n-type	T	–0.66	59Efi
	not reported		n.a.	–0.34	53Vas
Gold (Au)	0.001 **N** HCl/H_2O		S	–0.01	73Non
	0.0749 **N** HCl/H_2O		EB	–0.00	74Fre
	0.1 **N** HCl/H_2O		S	–0.07	67Bod
	0.1 **N** HCl/H_2O		G	0.14	88Mur
	0.1 **N** HCl/H_2O		S	–0.06	64And
	1 **M** HCl/H_2O		OA	0.086	52Hil
	$HClO_4/H_2O$		T	0.199	77Cla
	0.001 **M** $HClO_4/H_2O$		IM	0.34	73Kim
	0.001 **M** $HClO_4/H_2O$		T	0.17	69Boc1
	0.001 **M** $HClO_4/H_2O$		OA	0.3	63Gre
	0.001 **M** $HClO_4/H_2O$		PE	0.207	96Jin

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Au	0.001 **M** $HClO_4$+1 **N** $NaClO_4/H_2O$		OA	0.3	63Gre
	0.005 **M** $HClO_4/H_2O$		PE	0.177	96Jin
	0.01 **N** $HClO_4/H_2O$		F	0.15	69Boc1
	0.01 **M** $HClO_4/H_2O$		PE	0.155	96Jin
	0.1 **M** $HClO_4/H_2O$		T	0.24	63Sch1
	0.1 **M** $HClO_4/H_2O$		F	0.1	69Boc1
	0.001 **M** H_2SO_4/H_2O		S	0.12	73Non
	0.0087 **M** H_2SO_4/H_2O		T	0.330	72Car
	0.1 **M** H_2SO_4/H_2O		G	0.3	89Mur
	0.1 **N** H_2SO_4+Na_2SO_4/H_2O		S	0.18	67Bod
	0.001 **M** $HClO_4/H_2O$	pH=3	T	0.15	69Boc1
	0.001 **M** $HClO_4/H_2O$	pH=4	T	0.18	69Boc1
	0.02 **N** H_2SO_4/H_2O		T	0.23	64Kra
	0.1 **M** H_2SO_4/H_2O		G	0.28	88Mur
	0.001 **N** KBr/H_2O		T	–0.179	75Huo
	0.01 **N** KBr/H_2O	4< pH <12	S	–0.22	73Non
	0.01 **N** KBr/H_2O		T	–0.18	73Cla
	0.1 **N** KBr/H_2O		S	–0.13	73Non
	0.1 **N** KBr/H_2O		S	–0.26	65Per
	0.1 **N** KBr/H_2O		S	–0.24	64And, 67Bod
	0.1 **N** KBr/H_2O		R	–0.46	76Lin
	0.001 **N** KCl/H_2O		S	–0.01	73Non
	0.001 **N** KCl/H_2O		DR	–0.05	62Vor
	0.001 **N** KCl/H_2O		T	–0.001	75Huo
	0.0025 **N** KCl/H_2O		T	0.061	68Cla1
	0.005 **N** KCl/H_2O		T	0.031	68Cla1
	0.01 **N** KCl/H_2O		T	0.001	68Cla1
	0.02 **N** KCl/H_2O		T	–0.029	68Cla1
	0.05 **N** KCl/H_2O		T	–0.069	68Cla1
	0.1 **N** KCl/H_2O		T	–0.089	68Cla1
	0.01 **N** KCl/H_2O		T	0.06	73Cla
	0.01 **N** KCl/H_2O	0< pH <6	S	–0.07	73Non
	0.1 **N** KCl/H_2O		ED	0.0	71Fre
	0.1 **N** KCl/H_2O		R	–0.21	76Lin
	0.1 **N** KCl/H_2O		S	–0.11	65Per
	0.1 **N** KCl/H_2O	pH=7	S	–0.09	64Per
	0.1 **N** KCl/H_2O		S	–0.06	64And
	1 **N** KCl/H_2O		H	0.15	56Ven
	$KClO_4/H_2O$		T	0.199	77Cla
	0.1 **N** $KClO_4/H_2O$		T	0.24	86Ric1
	0.1 **N** $KClO_4/H_2O$		S	–0.180	65Per
	0.1 **N** $KClO_4/H_2O$		R	–0.160	76Lin
	0.1 **N** $KClO_4/H_2O$		EB	0.336	96Efi
	0.005 **N** KF/H_2O		PC	0.181	80Brz

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	0.01 **N** KF/H_2O		T	0.180	73Cla
	0.01 **N** KF/methanol		PC	0.26	83Brz
	0.1 **N** KF/H_2O		S	–0.060	65Per
	0.1 **N** KF/H_2O		CA	0.1	71Mor
	0.001 **N** KI/H_2O		S	–0.34	73Non
	0.001 **N** KI/H_2O		T	–0.459	75Huo
	0.01 **N** KI/H_2O	5< pH <12	T	–0.46	73Cla
	0.01 **N** KI/H_2O		S	–0.5	73Non
	0.1 **N** KI/H_2O		S	–0.47	65Per, 64And, 67Bod
	0.1 **N** KI/H_2O		R	–0.74	76Lin
	0.005 **M** KNO_3/H_2O		T	0.171	68Cla1
	0.02 **M** KNO_3/H_2O		PC	0.041	80Brz
	0.035 **M** KNO_3/H_2O		PC	0.011	80Brz
	0.05 **M** KNO_3/H_2O		PC	–0.029	80Brz
	0.1 **N** KNO_3/H_2O		S	–0.04	65Per
	0.1 **N** KNO_3/H_2O		R	–0.16	76Lin
	1 **N** KNO_3/H_2O		FL	0.43	51Pfü
	0.001 **N** KOH/H_2O		S	–0.06	73Non
	0.1 **N** KOH/H_2O		R	–0.16	76Lin
	0.1 **N** KOH/H_2O		CA	0.1	71Mor
	0.1 **N** $KSCN/H_2O$		S	–0.44	65Per
	0.01 **N** K_2SO_4/H_2O	0< pH <6	S	0.08	73Non
	0.01 **N** K_2SO_4/H_2O		EB	0.241	95Lan
	0.1 **N** K_2SO_4/H_2O		R	–0.06	76Lin
	0.1 **N** K_2SO_4/H_2O		S	0.06	65Per
	0.017 **N** K_2SO_4/H_2O		T	0.1	67Pet1
	0.0005 **N** $La(NO)_3/H_2O$		DR	–0.05	62Vor
	0.0005 **N** $La(NO)_3/H_2O$		DR	–0.05	63Vor
	0.003 **M** $LiClO_4$/ dimethylsulphoxide		T	0.454	87Jar
	0.005 **M** $LiClO_4$/ dimethylsulphoxide		T	0.249	88Bor2
	0.0055 **M** $LiClO_4$/ dimethylsulphoxide		T	0.444	87Jar
	0.01 **M** $LiClO_4$/dimethylsulphoxide		T	0.464	87Jar
	0.03 **M** $LiClO_4$/dimethylsulphoxide		T	0.454	87Jar
	0.001 **M** $LiClO_4$/acetonitrile		T	0.896	87Ham2
	0.02 **M** $LiNO_3/H_2O$		PC	0.041	83Brz
	0.035 **M** $LiNO_3/H_2O$		PC	0.011	83Brz
	0.05 **M** $LiNO_3/H_2O$		PC	–0.029	83Brz
	0.1 **M** $LiNO_3/H_2O$		PC	–0.059	83Brz
	0.02 **M** $LiNO_3$/methanol		PC	0.297	83Brz
	0.035 **M** $LiNO_3$/methanol		PC	0.320	83Brz
	0.05 **M** $LiNO_3$/methanol		PC	0.170	83Brz

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Au	0.1 **M** $LiNO_3$/methanol		PC	0.155	83Brz
	0.1 **M** $LiNO_3$/methanol		PC	0.095	83Brz
	0.00025 **M** $MgSO_4/H_2O$		T	0.181	74Cla1
	0.1 **N** $NaBr/H_2O$		S	–0.24	67Bod, 64And
	0.001 **N** $NaCl/H_2O$		S	–0.01	73Non
	0.1 **N** $NaCl/H_2O$		S	–0.07	67Bod
	0.01 **N** $NaClO_4/H_2O$		OA	0.37	69Boc1
	0.1 **N** $NaClO_4/H_2O$		S	0.14	67Bod
	0.1 **N** $NaClO_4/H_2O$		CA	0.241	71Shi
	0.1 **N** $NaClO_4/H_2O$		PE	0.18	99Uen
	0.1 **M** $NaClO_4$/propylene carbonate		T	0.4	85Ngu
	0.1 **N** $NaClO_4$+ $HClO_4/H_2O$	pH=2.3	PE	0.22	95Rai
	0.1 **N** $NaClO_4$+$HClO_4/H_2O$	pH=12	PE	–0.29	95Rai
	1 **N** $NaClO_4/H_2O$	pH=7	S	0.12	69And
	1 **N** $NaClO_4/H_2O$	pH=11	S	0.05	69And
	1 **N** $NaClO_4/H_2O$		PE	0.08	96Seo
	0.00025 **M** NaF/H_2O		T	0.181	74Cla1
	0.001 **M** NaF/H_2O		PE	0.28	96Jin
	0.001 **M** NaF/H_2O		S	0.13	73Non
	0.0024 **M** NaF/H_2O		PE	0.098	92Dic
	0.0025 **M** NaF/H_2O		T	0.18	73Cla
	0.0048 **M** NaF/H_2O		PE	0.116	92Dic
	0.005 **M** NaF/H_2O		PE	0.273	96Jin
	0.0072 **M** NaF/H_2O		PE	0.087	92Dic
	0.0095 **M** NaF/H_2O		PE	0.119	92Dic
	0.01 **M** NaF/H_2O		T	0.2	78Zel2
	0.01 **M** NaF/H_2O		T	0.18	73Cla
	0.02 **M** NaF/H_2O		PE	0.138	92Dic
	0.02 **M** NaF/H_2O		T	0.181	69Cla
	0.04 **M** NaF/H_2O		T	0.189	73Cla
	0.05 **M** NaF/H_2O		PE	0.15	92Dic
	0.1 **M** NaF/H_2O		PE	0.15	92Dic
	0.1 **M** NaF/H_2O		S	0.19	67Bod
	0.1 **M** NaF/H_2O		G	0.17	87Mur
	0.1 **M** NaF/H_2O		S	0.1	73Non
	0.1 **M** NaF/H_2O		PE	0.278	96Jin
	0.1 **N** NaI/H_2O		S	–0.53	67Bod
	0.1 **N** $NaOH/H_2O$		S	–0.09	67Bod
	0.1 **N** $NaOH/H_2O$		S	–0.08	64And
	0.1 **N** $NaOH/H_2O$		G	0.11	88Mur
	0.0005 **M** Na_2SO_4/H_2O		PE	0.107	96Jin
	0.001 **M** Na_2SO_4/H_2O		S	0.11	73Non
	0.01 **M** Na_2SO_4/H_2O		T	0.09	75Non
	0.01 **M** Na_2SO_4/H_2O		S	0.09 [15])	75Non

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	0.01 **M** Na_2SO_4/H_2O		S	0.09 [14])	75Non
	0.01 **M** Na_2SO_4/H_2O		IM	0.284 [20])	90Sok
	0.01 **M** Na_2SO_4/H_2O		IM	0.06	95Tur
	0.01 **M** Na_2SO_4/H_2O		IM	–0.39	97Tur
	0.1 **M** Na_2SO_4/H_2O		PE	0.09	96Jin
	0.1 **M** Na_2SO_4+0.001 **M** H_2SO_4+ $5\cdot10^{-5}$ **M** Ag_2SO_4/H_2O		PE	0.9 [29])	94Seo
	0.01 **M** Na_2SO_4/H_2O+1-propanol	3 wt%	IM	0.155	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	6.3 wt%	IM	0.225	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	9.4 wt%	IM	0.23	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	11 wt%	IM	0.23	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	15.3 wt%	IM	0.19	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	25.1 wt%	IM	0.2	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	32.1 wt%	IM	0.225	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	43.4 wt%	IM	0.22	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	56 wt%	IM	0.207	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	63.8 wt%	IM	0.2	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	77 wt%	IM	0.195	95Tur
	0.1 **N** Na_2SO_4/H_2O		S	0.14	67Bod
	0.2 **N** Na_2SO_4/H_2O		T	0.23	64Kra
	0.4 **M** N LiBr+0.1 **M** LiCl/butan-1-ol		VIB	–0.29 [30])	96Pem
	0.4 **M** N LiBr+0.1 **M** LiCl/butan-2-ol		VIB	–0.05 [30])	96Pem
	0.4 **M** N LiBr+0.1 **M** LiCl/ 2-methylpropan-1-ol		VIB	–0.19 [30])	96Pem
Au (100)	0.001 **M** $HClO_4/H_2O$		PC	0.191	62Sch1
	0.001 **M** $HClO_4/H_2O$		T	0.291	94Ham
	0.01 **M** $HClO_4/H_2O$		T	0.341	85Kol
	0.01 **M** $HClO_4/H_2O$		T	0.29	96Sil
	0.1 **M** $KClO_4/H_2O$		T	0.341	90Ric
	0.005 **M** $LiClO_4$/ dimethylsulphoxide		T	0.229	88Bor2
	0.0025 **M** NaF/H_2O		T	0.381	71Ham1
	0.010 **M** NaF/H_2O		T	0.381	76Ham
	0.010 **M** NaF/H_2O		T	0.341	78Lec
	0.010 **M** NaF/H_2O	pH=5.6	T	0.341	82Lec
	0.020 **M** NaF/H_2O		T	0.301	80Kof
	0.01 **M** K_2SO_4/H_2O		T	0.02	69Ham2
	0.00096 **M** KPF_6/H_2O		T	0.388	87Ham1
	0.002 **M** TEAC [31])/acetonitrile		T	–0.05 [32])	95Pan1
Au (100)-(5x20)	0.01 **M** $HClO_4/H_2O$		T	0.541	86Kol
Au (100)-(1x1)	0.01 **M** $HClO_4/H_2O$		T	0.321	86Kol
	0.1 **M** $HClO_4$+5 m**M** KI/H_2O		T	0.559 [33])	94Gao
Au (110)	0.001 **M** $HClO_4/H_2O$		PC	0.241	62Sch1
	0.01 **M** $HClO_4/H_2O$	2(111)-(111)	T	0.2	96Sil

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Au (110)	0.1 **M** $HClO_4$+50 m**M** KI/H_2O		T	0.409	94Gao
	0.1 **M** $KClO_4$/H_2O		T	0.191	90Ric
	0.005 **M** $LiClO_4$/ dimethylsulphoxide		T	0.374	88Bor2
	0.0025 **M** NaF/H_2O		T	0.191	71Ham1
	0.01 **M** NaF/H_2O	pH=5.6	T	0.177	82Lec
	0.01 **M** NaF/H_2O		T	0.191	76Ham
	0.02 **M** NaF/H_2O	pH=5	T	0.151	80Kof
	0.01 **M** K_2SO_4/H_2O		T	0.1	69Ham2
	0.00096 **M** KPF_6/H_2O		T	0.201	87Ham1
	0.005 **M** K_2SO_4/H_2O		T	0.101	74Ham
Au(110)-(1x2)	0.01 **M** $HClO_4$/H_2O		T	0.201	86Kol
Au(110)-(1x1)	0.01 **M** $HClO_4$/H_2O		T	0.221	86Kol
Au (111)	0.01 **M** $HClO_4$/H_2O		CV	0.52	00Alv
	0.01 **M** $HClO_4$/H_2O		T	0.52	96Sil
	0.1 **M** $HClO_4$/H_2O		IM	0.53	96Ham1
	0.1 **M** KCl/H_2O		R	–0.25	98Bru
	0.1 **M** $KClO_4$/H_2O		T	0.516	90Ric
	0.1 **M** $KClO_4$/H_2O		T	0.521	01Li
	0.01 **M** KPF_6/H_2O		T	0.561	92Lec
	0.005 **M** $LiClO_4$/ dimethylsulphoxide		T	0.204	88Bor2
	0.0025 **M** NaF/H_2O		T	0.501	71Ham1
	0.002 **M** NaF/H_2O		T	0.571	74Lec
	0.01 **M** NaF/H_2O		T	0.571	75Lec1, 75Lec2, 76Ham
	0.01 **M** NaF/H_2O	pH=5.6	T	0.571	82Lec
	0.02 **M** NaF/H_2O	pH=5	T	0.481	80Kof
	0.01 **M** K_2SO_4/H_2O		T	0.45	69Ham2
	0.005 **M** K_2SO_4/H_2O		T	0.361	74Ham
	0.00096 **M** KPF_6/H_2O		T	0.541	87Ham1
	0.002 **M** TEAC [31])/acetonitrile		T	0.2 [32])	95Pan1
	0.002 **M** TEAC [31])/acetonitrile		T	0.2 [32])	95Pan1
	0.1 **M** $NaClO_4$/H_2O		EC	–0.239 [34])	04Iwa
	0.1 **M** $NaClO_4$/H_2O		EC	–0.209 [35])	04Iwa
	0.1 **M** $NaClO_4$/H_2O		EC	–0.019 [36])	04Iwa
	0.1 **M** $NaClO_4$/H_2O		EC	0.759 [37])	04Iwa
Au(111)-(1x23)	0.01 **M** $HClO_4$/H_2O		T	0.561	86Kol
	0.01 **M** $HClO_4$/H_2O		T	0.551	85Kol
Au(111)-(1x1)	0.01 **M** $HClO_4$/H_2O		T	0.471	86Kol
Au (210)	0.01 **M** $HClO_4$/H_2O	2(100)-(110)	T	0.12	96Sil
	0.01 **M** $HClO_4$/H_2O	2(110)-(101)	T	0.12	96Sil

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	0.1 **M** $KClO_4/H_2O$		T	0.141	90Ric
	0.005 **M** $LiClO_4$/ dimethylsulphoxide		T	0.414	88Bor2
	0.01 **M** $LiClO_4/H_2O$		T	0.161	82Ham2
	0.01 **M** NaF/H_2O	pH=6.5	C	0.121	82Ham1
	0.010 **M** NaF/H_2O	pH=5.6	T	0.104	82Lec
	0.01 **M** $NaF+HClO_4/H_2O$	pH=5.6	C	0.096	82Ham1
	0.01 **M** $NaBF_4/H_2O$		C	0.151	82Ham1
Au (221)	0.010 **M** NaF/H_2O	pH=5.6	T	0.204	82Lec
Au (310)	0.010 **M** NaF/H_2O	pH=5.6	T	0.122	82Lec
Au (311)	0.1 **M** $KClO_4/H_2O$		T	0.241	90Ric
	0.01 **M** NaF/H_2O	pH=5.6	T	0.259	82Lec
	0.01 **M** $HClO_4/H_2O$	2(100)-(111)	T	0.241	93Sil
	0.01 **M** $HClO_4/H_2O$	2(100)-(111)	T	0.24	96Sil
	0.01 **M** $HClO_4/H_2O$	2(111)-(100)	T	0.24	96Sil
Au (320)	0.01 **M** NaF/H_2O	pH=5.6	T	0.122	82Lec
	0.01 **M** $HClO_4/H_2O$	3(110)-(100)	T	0.13	96Sil
Au (331)	0.01 **M** NaF/H_2O	pH=5.6	T	0.286	82Lec
	0.01 **M** $HClO_4/H_2O$	3(111)-(111)	T	0.216	93Sil
	0.01 **M** $HClO_4/H_2O$	3(111)-(111)	T	0.22	96Sil
	0.01 **M** $HClO_4/H_2O$	2(110)-(111)	T	0.22	96Sil
Au (332)	0.01 **M** NaF/H_2O	pH=5.6	T	0.304	82Lec
	0.01 **M** $HClO_4/H_2O$	6(111)-(111)	T	0.22	96Sil
Au (410)	0.01 **M** $HClO_4/H_2O$	4(100)-(110)	T	0.2	96Sil
Au (511)	0.01 **M** NaF/H_2O	pH=5.6	T	0.241	82Lec
Au (533)	0.01 **M** NaF/H_2O	pH=5.6	T	0.307	82Lec
	0.01 **M** $HClO_4/H_2O$	2(100)-(111)	T	0.303	93Sil
	0.01 **M** $HClO_4/H_2O$	4(111)-(100)	T	0.31	96Sil
Au (551)	0.01 **M** NaF/H_2O	pH=5.6	T	0.186	82Lec
	0.01 **M** $HClO_4/H_2O$		T	0.217	93Sil
Au (554)	0.010 **M** NaF/H_2O	pH=5.6	T	0.413	82Lec
Au (755)	0.010 **M** NaF/H_2O	pH=5.6	T	0.377	82Lec
	0.01 **M** KPF_6/H_2O		T	0.401	92Lec
Au (771)	0.01 **M** $HClO_4/H_2O$	7/3(111)-(111)	T	0.18	96Sil
Au (10.1.1)	0.010 **M** NaF/H_2O	pH=5.6	T	0.268	82Lec
Au (11.9.9)	0.010 **M** NaF/H_2O	pH=5.6	T	0.459	82Lec
Au (19.1.1)	0.010 **M** NaF/H_2O	pH=5.6	T	0.313	82Lec
Indium	0.001 **M** HCl+0.01 **M** KCl/H_2O		P	–0.9	74Rot
(In)	2 **N** H_2SO_4/H_2O		ER	–0.8	70Rot2
	2 **N** H_2SO_4/H_2O		ER	–0.65	70Rot2 [38])
	1 **N** $NaClO_4/H_2O$	pH=7	S	–0.65	69And
	0.003 **M** NaF/H_2O		T	–0.65	72Gri3
	1 **N** K_2SO_4/H_2O	pH=7	S	–0.64	69And

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
In amalgam	0.1 **M** $HClO_4/H_2O$	4.7 10^{-3} mol% In	TC	–0.255	64But1 [39])
	0.1 **M** $HClO_4/H_2O$	9.4 10^{-3} mol% In	TC	–0.268	64But1
	1 **M** $HClO_4/H_2O$	0.015 mol% In	TC	–0.229	66But1
	0.1 **M** $HClO_4/H_2O$	0.094 mol% In	TC	–0.281	64But1
	1 **N** KBr+0.01 **N** H_2SO_4/H_2O	0.11 mol% In	EC	–0.449	65Pol
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	0.11 mol% In	EC	–0.33	65Pol
	1 **M** $HClO_4/H_2O$	0.138 mol% In	TC	–0.271	66But1
	1 **N** KBr+0.01 **N** H_2SO_4/H_2O	0.49 mol% In	EC	–0.471	65Pol
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	0.49 mol% In	EC	–0.33	65Pol
	0.1 **M** $HClO_4/H_2O$	0.92 mol% In	TC	–0.322	64But1
	0.1 **M** $HClO_4/H_2O$	0.9 at% In	T	–0.322	65But6
	1 **N** KBr+0.01 **N** H_2SO_4/H_2O	1.14 mol% In	EC	–0.525	65Pol
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	1.14 mol% In	EC	–0.38	65Pol
	1 **M** $HClO_4/H_2O$	1.23 mol% In	TC	–0.324	66But1
	0.1 **M** $HClO_4/H_2O$	1.71 mol% In	EC	–0.34	68Pol
	1 **M** $NaClO_4/H_2O$	2 mol% In	T	–0.359	95Koe
	1 **N** KBr+0.01 **N** H_2SO_4/H_2O	2.04 mol% In	EC	–0.525	65Pol
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	2.04 mol% In	EC	–0.39	65Pol
	0.1 **M** $HClO_4/H_2O$	3.9 mol% In	TC	–0.36	64But1
	0.1 **M** $HClO_4/H_2O$	3.9 at% In	T	–0.36	65But6
	1 **N** KBr+0.01 **N** H_2SO_4/H_2O	4.6 mol% In	EC	–0.568	65Pol
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	4.6 mol% In	EC	–0.41	65Pol
	1 **M** $HClO_4/H_2O$	5.0 mol% In	TC	–0.366	66But1
	0.1 **M** $HClO_4/H_2O$	5.95 mol% In	EC	–0.37	68Pol
	1 **M** $HClO_4/H_2O$	5.97 mol% In	TC	–0.37	66But1
	0.01 **M** $HClO_4/H_2O$	5.97 mol% In	TC	–0.365	66But1
	0.1 **M** $HClO_4/H_2O$	8.4 mol% In	TC	–0.382	64But1
	0.1 **M** $HClO_4/H_2O$	8.4 at% In	T	–0.387	65But6
	1 **M** $HClO_4/H_2O$	10.0 mol% In	TC	–0.403	66But1
	1 **M** $HClO_4/H_2O$	10.0 mol% In	TC	–0.416	66But1
	1 **M** $NaClO_4/H_2O$	10 mol% In	T	–0.414	95Koe
	0.01 **M** $HClO_4/H_2O$	10.02 mol% In	TC	–0.391	66But1
	0.1 **M** $HClO_4/H_2O$	10.02 mol% In	TC	–0.369	66But1
	1 **N** KBr+0.01 **N** H_2SO_4/H_2O	10.3 mol% In	EC	–0.622	65Pol
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	10.3 mol% In	EC	–0.46	65Pol
	0.1 **M** $HClO_4/H_2O$	16.3 mol% In	TC	–0.434	64But1
	0.1 **M** $HClO_4/H_2O$	20.0 mol% In	TC	–0.457	64But1
	1 **M** $NaClO_4/H_2O$	20 mol% In	T	–0.469	95Koe
	0.01 **M** $HClO_4/H_2O$	20.34 mol% In	TC	–0.445	66But1
	1 **M** $HClO_4/H_2O$	20 mol% In	TC	–0.471	66But1
	0.1 **M** $HClO_4/H_2O$	20.34 mol% In	TC	–0.457	66But1
	0.1 **M** $HClO_4/H_2O$	20.0 mol% In	EC	–0.457	68Pol
	1 **N** KBr+0.01 **N** H_2SO_4/H_2O	20.0 mol% In	EC	–0.667	65Pol

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	20.0 mol% In	EC	–0.54	65Pol
	0.1 **M** $HClO_4/H_2O$	30.0 mol% In	TC	–0.489	64But1
	1 **M** $NaClO_4/H_2O$	30 mol% In	T	–0.526	95Koe
	1 **N** KBr+0.01 **N** H_2SO_4/H_2O	30.2 mol% In	EC	–0.719	65Pol
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	30.2 mol% In	EC	–0.58	65Pol
	0.1 **M** $HClO_4/H_2O$	30.4 mol% In	TC	–0.515	64But1
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	38.9 mol% In	EC	–0.63	65Pol
	1 **M** $HClO_4/H_2O$	39.83 mol% In	TC	–0.570	66But1
	0.01 **M** $HClO_4/H_2O$	39.89 mol% In	TC	–0.524	66But1
	0.1 **M** $HClO_4/H_2O$	39.89 mol% In	TC	–0.552	66But1
	0.1 **M** $HClO_4/H_2O$	40 mol% In	TC	–0.525	64But1
	1 **M** $NaClO_4/H_2O$	40 mol% In	T	–0.571	95Koe
	0.1 **M** $HClO_4/H_2O$	42.7 mol% In	EC	–0.545	68Pol
	1 **N** KBr+0.01 **N** H_2SO_4/H_2O	48.5 mol% In	EC	–0.784	65Pol
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	48.5 mol% In	EC	–0.66	65Pol
	0.1 **M** $HClO_4/H_2O$	50.0 mol% In	TC	–0.585	64But1
	1 **M** $NaClO_4/H_2O$	50 mol% In	T	–0.610	95Koe
	0.1 **M** $HClO_4/H_2O$	52.2 mol% In	TC	–0.605	64But1
	1 **N** KBr+0.01 **N** H_2SO_4/H_2O	57.7 mol% In	EC	–0.816	65Pol
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	57.7 mol% In	EC	–0.68	65Pol
	1 **M** $NaClO_4/H_2O$	60 mol% In	T	–0.639	95Koe
	0.01 **M** $HClO_4/H_2O$	63.02 mol% In	TC	–0.592	66But1
	0.1 **M** $HClO_4/H_2O$	63.02 mol% In	TC	–0.622	66But1
	1 **M** $HClO_4/H_2O$	63.02 mol% In	TC	–0.646	66But1
	0.1 **M** $HClO_4/H_2O$	63.6 mol% In	TC	–0.621	64But1
	0.1 **M** $HClO_4/H_2O$	64.4 mol% In	TC	–0.625	64But1
	0.1 **M** $HClO_4/H_2O$	64.5 mol% In	EC	–0.646	68Pol
In/Ga (eutectic alloy)	1 **N** KBr/H_2O		T	–0.68	72Gri5
	1 **N** KCl/H_2O		T	–0.74	72Gri5
	1 **N** KI/H_2O		T	–0.89	72Gri5
	1 **N** K_2SO_4/H_2O		T	–0.68	72Gri5
	0.1 **M** $NaClO_4$/*N*-methylformamide		T	–0.629	96Eme
	1 **N** Na_2SO_4/H_2O		T	–0.68	72Gri5
	0.001 **M** $LiClO_4$/acetonitrile		T	–0.509	75Bag
Iridium (Ir)	0.01 **N** NaCl+0.001 **N** HCl		P	–0.63	74Bro1
	1 **N** $NaClO_4/H_2O$	pH=7	S	–0.04	69And
	1 **N** $NaClO_4/H_2O$	pH=11	S	–0.17	69And
	1 **N** K_2SO_4/H_2O	pH=7	S	–0.04	69And
	1 **N** K_2SO_4/H_2O	pH=11	S	–0.17	69And
Iron (Fe)	n.a./H_2O		n.a.	–0.02	58Ant
	0.003 **N** HCl/H_2O		DR	–0.4	63Vor
	0.01 **N** HCl/H_2O		DR	–0.4	63Vor
	0.3 **N** $HClO_4/H_2O$		T	–0.28	67Mur
	0.001 **N** H_2SO_4/H_2O		T	–0.37	55Aya, 55Fru

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Fe	0.1 **N** H_2SO_4/H_2O		T	–0.33	55Aya
	0.01 **N** KCl/H_2O [40])	pH=3.5	T	–0.7	75Ryb2
	0.1 **M** KF/H_2O	pH=3.5	T	–0.7	76Ryb
	0.3 **N** $NaBr/H_2O$		PE	–0.65	97Seo
	0.3 **N** $NaCl/H_2O$		PE	–0.65	97Seo
	0.3 **N** $NaClO_4/H_2O$		PE	–0.52	97Seo
	0.3 **N** NaI/H_2O		PE	–0.64	97Seo
	0.01 **N** Na_2SO_4/H_2O [40])	pH=3.5	T	–0.7	75Ryb2
	0.01 **N** Na_2SO_4/H_2O		S	–0.64	78Laz
	0.01 **M** Na_2SO_4/H_2O		IM	0.074 [20])	90Sok
	0.01 **M** Na_2SO_4/H_2O		IM	–0.16	95Tur
	0.01 **M** Na_2SO_4/H_2O		IM	–0.13	97Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	3 wt%	IM	–0.3	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	6.3 wt%	IM	0.075	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	9.4 wt%	IM	0.13	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	11 wt%	IM	0.142	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	15.3 wt%	IM	0.095	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	25.1 wt%	IM	0.025	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	32.1 wt%	IM	0.1	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	43.4 wt%	IM	0.105	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	56 wt%	IM	0.095	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	63.8 wt%	IM	0.08	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	77 wt%	IM	0.065	95Tur
	0.02 **N** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=1.5	T	–0.325	57Khe
	0.02 **N** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=3	T	–0.48	57Khe
	0.3 **N** Na_2SO_4/H_2O		PE	–0.59	97Seo
Fe (100)	0.005 **N** $NaClO_4$+3 m**M** $HClO_4/H_2O$		T	–0.72	79Bat
Fe (111)	0.01 **N** $NaClO_4$+3 m**M** $HClO_4/H_2O$		T	–0.69	79Bat
Fe-Ni alloy	0.02 **N** $Na_2SO_4+H_2SO_4/H_2O$	pH=3.2, 5 % Fe	T	–0.345	57Kra
	0.02 **N** $Na_2SO_4+H_2SO_4/H_2O$	pH=3.2, 8 % Fe	T	–0.353	57Kra
	0.02 **N** $Na_2SO_4+H_2SO_4/H_2O$	pH=3.2, 12 % Fe	T	–0.36	57Kra
	0.02 **N** $Na_2SO_4+H_2SO_4/H_2O$	pH=3.2, 20 % Fe	T	–0.375	57Kra
	0.02 **N** $Na_2SO_4+H_2SO_4/H_2O$	pH=3.2, 25 % Fe	T	–0.38	57Kra
	0.02 **N** $Na_2SO_4+H_2SO_4/H_2O$	pH=3.2, 40 % Fe	T	–0.383	57Kra
	0.02 **N** $Na_2SO_4+H_2SO_4/H_2O$	pH=3.2, 50 % Fe	T	–0.383	57Kra
	0.02 **N** $Na_2SO_4+H_2SO_4/H_2O$	pH=3.2, 60 % Fe	T	–0.382	57Kra

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	0.02 **N** $Na_2SO_4+H_2SO_4/H_2O$	pH=3.2, 70 % Fe	T	–0.384	57Kra
	0.02 **N** $Na_2SO_4+H_2SO_4/H_2O$	pH=3.2, 80 % Fe	T	–0.382	57Kra
	0.02 **N** $Na_2SO_4+H_2SO_4/H_2O$	pH=3.2, 90 % Fe	T	–0.383	57Kra
Lead	0.001 **M** HCl+0.01 **M** KCl/H_2O		P	–0.85	74Rot
(Pb)	0.0034 **N** HNO_3/H_2O		T	–0.59	71Car
	0.0056 **N** H_3PO_4/H_2O		T	–0.59	71Car
	0.001 **N** H_2SO_4/H_2O		T	–0.67	55Fru
	0.001 **N** H_2SO_4/H_2O		AE	–0.62	66Kuk2
	0.0027 **N** H_2SO_4/H_2O		T	–0.59	71Car
	0.001 **M** KCl/H_2O		DR	0.2	59Vor
	0.01 **N** KCl/H_2O		T	–0.69	48Bor
	0.01 **N** KCl/H_2O		H	–0.60	49Ven
	0.01 **N** KCl/H_2O		T	–0.75	50Bor
	0.001 **M** KCl/H_2O		DR	0.2	59Vor
	KCl/H_2O		H	–0.37	44Reb
	0.0027 **N** KNO_3/H_2O		T	–0.59	71Car
	0.25 **N** $NaCH_3COO$/glacial acetic acid		T	–0.6	72Pan3
	0.01 **N** NaCl+0.001 **N** HCl/H_2O		P	–0.6	74Bro1
	0.1 **N** $NaCl/H_2O$		H	–0.62	55Fru, 56Ven
	1 **N** $NaCl/H_2O$		H	–0.6	49Ven
	1 **N** $NaCl/H_2O$		H	–0.63	44Reb
	0.001 **N** NaF/H_2O		T	–0.56	67Ryb
	0.01 **N** NaF/H_2O		T	–0.56	67Ryb
	0.01 **N** NaF/H_2O		T	–0.6	89Saf
	0.026 **N** NaF/H_2O		T	–0.59	78Mis
	0.001 **N** NaF/methanol		T	0.921	76Ush
	1 **N** $NaOH/H_2O$		H	–0.58	44Reb
	1 **N** $NaOH/H_2O$		H	–0.55	49Ven
	0.001 **M** Na_2SO_4+0.001 N $K_2S_2O_8$/ H_2O		ER	–0.65	55Fru
	0.01 **N** Na_2SO_4/H_2O		T	–0.6	89Saf
	0.1 **N** Na_2SO_4/H_2O		T	–0.64	67Ryb
	0.1 **N** Na_2SO_4/H_2O		H	–0.6	52Ven
	1 **N** Na_2SO_4/H_2O		H	–0.59	44Reb
	1 **N** Na_2SO_4/H_2O		H	–0.56	49Ven
	1 **N** Na_2SO_4/H_2O		H	–0.59	56Ven
	1 **N** Na_2SO_4/H_2O		H	–0.6	66Kuk1
	$S_2O_8^{2-}/H_2O$		T	–0.65	57Khe
	KF/MeOH		T	–0.43	89Saf
Pb (s.c.) [41])	0.01 **N** H_2SO_4/H_2O		FL	–0.7	69Lik

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Pb (s.c.) [41])	0.01 **M** $NaBr/H_2O$		FL	–0.8	69Lik
	0.01 **M** $NaCl/H_2O$		FL	–0.71	69Lik
	0.01 **M** NaF/H_2O		FL	–0.68	69Lik
Pb (100)	0.001 **M** NaF/H_2O		T	–0.56	67Ryb
	0.01 **M** NaF/H_2O		T	–0.589	78Kme
	0.01 **M** NaF/H_2O		T	–0.59	89Saf
	0.005 **M** Na_2SO_4/H_2O		T	–0.609	78Kme
	0.005 **M** Na_2SO_4/H_2O		T	–0.59	89Saf
Pb (110)	0.01 **M** NaF/H_2O		T	–0.579	78Kme
	0.01 **M** NaF/H_2O		T	–0.58	89Saf
	0.005 **M** Na_2SO_4/H_2O		T	–0.609	78Kme
	0.005 **M** Na_2SO_4/H_2O		T	–0.58	89Saf
Pb (111)	0.01 **M** NaF/H_2O		T	–0.619	78Kme
	0.01 **M** NaF/H_2O		T	–0.62	89Saf
	0.005 **M** Na_2SO_4/H_2O		T	–0.649	78Kme
	0.005 **M** Na_2SO_4/H_2O		T	–0.62	89Saf
Pb (112)	0.010 **M** NaF/H_2O		T	–0.579	78Kme
	0.01 **M** NaF/H_2O		T	–0.58	89Saf
	0.005 **M** Na_2SO_4/H_2O		T	–0.609	78Kme
	0.005 **M** Na_2SO_4/H_2O		T	–0.58	89Saf
Lithium (Li)	not reported		n.a.	–2.24	53Vas
Magnesium (Mg)	not reported		n.a.	1.08	53Vas
Manganese (Mn)	not reported		n.a.	–1.01	53Vas
Mercury (Hg)	1 **N** $BaCl_2/H_2O$		CA	–0.28	55Uks2
	1 **N** $BaCl_2/H_2O$		EC	–0.28	67Arg
	1 **N** CH_3COONa/H_2O		CA	–0.24	55Uks2
	0.01 **N** CsF/H_2O		T	–0.193	58Dam
	0.1 **N** CsF/H_2O		T	–0.192	58Dam
	0.9 **N** CsF/H_2O		T	–0.189	58Dam
	0.1 **N** CsI/H_2O		T	–0.452	58Dam
	0.9 **N** CsI/H_2O		T	–0.532	58Dam
	0.01 **M** HCl/H_2O		S	0.133 [42])	65Boc
	0.01 **M** HCl/H_2O		T	0.138 [42])	65Boc
	0.03 **M** HCl/H_2O		S	0.142	65Boc
	0.03 **M** HCl/H_2O		T	0.147	65Boc
	0.1 **M** HCl/H_2O		H	0.19 [43])	53Boc
	0.1 **M** HCl/H_2O		S	0.168 [42])	65Boc
	0.1 **M** HCl/H_2O		T	0.182 [42])	65Boc
	0.1 **N** HCl/H_2O		CA	–0.5 [44])	71Mor
	0.1 **N** HCl/H_2O		CA	–0.259	68Mor
	0.1 **M** $(HCl+KCl/H_2O)$	pH=1	S	–0.639	86Hor

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	0.1 **M** (HCl+KCl/H_2O)	pH=1.3	S	–0.639	86Hor
	0.1 **M** (HCl+KCl/H_2O)	pH=2.1	S	–0.643	86Hor
	0.1 **M** (HCl+KCl/H_2O)	pH=3.2	S	–0.645	86Hor
	0.1 **M** (HCl+KCl/H_2O)	pH=3.8	S	–0.663	86Hor
	0.1 **M** (HCl+KCl/H_2O)	pH=4	S	–0.675	86Hor
	0.1 **M** (HCl+KCl/H_2O)	pH=4.1	S	–0.684	86Hor
	0.3 **M** HCl/H_2O		S	0.214 [42])	65Boc
	0.3 **M** HCl/H_2O		T	0.227 [42])	65Boc
	1 **N** HCl/H_2O		CA	–0.3	55Uks2
	1 **M** HCl/H_2O		S	0.290 [42])	65Boc
	1 **N** HCl/H_2O		EC	–0.3	67Arg
	1 **N** HCl/H_2O		T	–0.321	69Dek
	2.94 **M** HCl/H_2O		S	0.401 [42])	65Boc
	2.94 **M** HCl/H_2O		T	0.402 [42])	65Boc
	7.5 **N** HCl/H_2O		T	–0.562	69Dek
	0.01 **M** $HClO_4$/H_2O		TC	–0.21	66But1
	0.1 **M** $HClO_4$/H_2O		TC	–0.23	64But1
	0.1 **M** $HClO_4$/H_2O		T	–0.235	65But6
	0.1 **M** $HClO_4$/H_2O		TC	–0.229	66But1
	0.1 **M** $HClO_4$/H_2O		EC	–0.21	68Pol
	1 **M** $HClO_4$/H_2O		TC	–0.282	66But1
	5.2 **M** $HClO_4$/H_2O		T	–0.316	69Dek
	0.25 **M** HPF_6/H_2O		EC, T	–0.241	73Bau
	0.1 **N** H_2SO_4/H_2O		EC	–0.258	38Win
	1 **N** H_2SO_4/H_2O		CA	–0.23	55Uks2
	1 **N** H_2SO_4/H_2O		EC	–0.23	67Arg
	0.1 **N** KBr/H_2O		S	–0.277	62Jen
	0.1 **N** KBr/H_2O		EC	–0.295	52Gra
	1 **N** KBr+0.01 **N** H_2SO_4/H_2O		EC	–0.395	65Pol
	1 **N** KBr/H_2O		CA	–0.37	55Uks2
	1 **N** KBr/H_2O		EC	–0.37	67Arg
	0.01 **N** KCl/H_2O		S	–0.448	62Jen
	0.01 **N** KCl/H_2O		P	–0.199	69Ple
	0.01 **N** KCl/H_2O		S	–0.326	68Log
	0.01 **N** KCl/H_2O [45])		P	–0.199	72Bro
	0.01 **M** KCl/ H_2O+1,4-dioxane	5 wt%	VE	–0.301	83Mly
	0.01 **M** KCl/ H_2O+1,4-dioxane	10 wt%	VE	–0.271	83Mly
	0.01 **M** KCl/ H_2O+1,4-dioxane	20 wt%	VE	–0.251	83Mly
	0.1 **m** KCl/H_2O		S	–0.224	56Ran
	0.1 **N** KCl/H_2O		EC	–0.226	49Gra
	0.1 **N** KCl/H_2O		T	–0.190	47Gra
	0.1 **N** KCl/H_2O		EC	–0.226	52Gra
	0.1 **N** KCl/H_2O		S	–0.226	62Jen
	0.1 **N** KCl/H_2O		S	–0.291	68Log
	0.1 **N** KCl/H_2O		NS	–0.240	98Pal

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Hg	1 **N** KCl/H_2O		S	–0.275	62Jen
	0.01 **N** $KClO_4/H_2O$		CA	–0.5 [44])	71Mor
	0.01 **N** $KClO_4/H_2O$		CA	–0.259	68Mor
	0.1 **N** $KClO_4/H_2O$		EC	–0.228	52Gra
	0.1 **M** $KClO_4$/dimethylsulfoxide		S	–0.226	87Bor
	0.01 **N** KF/ H_2O		IM	–0.290	71Jak
	0.1 **N** $KCNS/H_2O$		EC	–0.348	52Gra
	1.1 **N** $KCNS/H_2O$		EC	–0.437	36Erd
	3 **N** $KCNS/H_2O$		CA	–0.49	55Uks2
	3 **N** $KCNS/H_2O$		EC	–0.49	67Arg
	0.1 **N** K_2CO_3/H_2O		EC	–0.193	52Gra
	1 **N** K_2CO_3/H_2O		CA	–0.2	55Uks2
	1 **N** K_2CO_3/H_2O		EC	–0.2	67Arg
	0.1 **N** $KC_2H_3O_2/H_2O$		EC	–0.209	52Gra
	0.001 **N** KF/ H_2O		IM	–0.193	71Jak
	0.01 **N** KF/H_2O		P	–0.19	74Bro1
	0.1 **N** KF/H_2O		CA	–0.209	68Mor
	0.1 **N** KF/H_2O		CA	–0.45 [44])	71Mor
	0.1 **N** KF/formamide		VE	–0.510	83Mly
	0.005...0.2 **M** KF/H_2O		T	–0.184	88Jap
	0.005...0.2 **M** KF/ethanediol		T	–0.089	88Jap
	0.005...0.2 **M** KF/1,2-propanediol		T	–0.084	88Jap
	0.005...0.2 **M** KF/1,3-propanediol		T	–0.154	88Jap
	0.005...0.2 **M** KF/1,2-butanediol		T	–0.094	88Jap
	0.005...0.2 **M** KF/1,3-butanediol		T	–0.154	88Jap
	0.005...0.2 **M** KF/1,4-butanediol		T	–0.129	88Jap
	0.005...0.2 **M** KF/2,3-butanediol		T	–0.089	88Jap
	1 **N** KF/H_2O		EC	–0.194	52Gra
	0.001 **M** $KF+N_2O/H_2O$		P	–0.43	74Rot
	0.1 **N** $KHCO_2/H_2O$		EC	–0.193	52Gra
	1 **N** K_2HPO_4/H_2O		CA	–0.21	55Uks2
	1 **N** K_2HPO_4/H_2O		EC	–0.21	67Arg
	0.001 **N** KI/H_2O		IM	–0.287	71Jak
	0.1 **N** KI/H_2O		EC	–0.453	52Gra
	0.1 **N** KI/H_2O		CA	–0.479	68Mor
	0.1 **N** KI/H_2O		CA	–0.72 [44])	71Mor
	1 **N** KI/H_2O		CA	–0.6 [44])	71Mor
	3 **N** KI/H_2O		EC	–0.59	67Arg
	0.01 **N** KNO_3/H_2O		EC	–0.246	38Win
	0.1 **N** KNO_3/H_2O		EC	–0.261	38Win
	0.1 **N** KNO_3/H_2O		EC	–0.236	52Gra
	1 **N** KNO_3/H_2O		EC	–0.322	38Win
	1 **N** KNO_3/H_2O		CA	–0.28	55Uks2
	1 **N** KNO_3/H_2O		EC	–0.28	67Arg
	3 **N** KNO_3/H_2O		CA	–0.59	55Uks2

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	KNO_3/H_2O		CA	–0.19	67Vet
	0.1 **N** KOH/H_2O		EC	–0.196	52Gra
	0.1 **M** KOH/H_2O		EC	–1.3	67Pol
	0.1 **N** KOH/H_2O		EC, T	–0.176	73Bau
	1 **N** KOH/H_2O		EC	–0.19	67Arg
	1 **N** KOH/H_2O		CA	–0.19	55Uks2
	0.01 **M** KPF_6/propylene carbonate		VE	–0.300	83Mly
	0.01 **M** KPF_6/nitromethan		EC	–0.204	81Jur
	0.1 **M** KPF_6/H_2O		T	–0.243	67Pay
	0.1 **M** KPF_6/methanol		VE	–0.120	83Mly
	0.1 **M** KPF_6/dimethyl sulfoxide		T	–0.073	67Pay
	0.1 **M** KPF_6/4-butyrolactone		T	–0.039	67Pay
	0.1 **M** KPF_6/4-valerolactone		T	–0.032	67Pay
	0.1 **M** KPF_6/propylene carbonate		T	–0.089	67Pay
	0.1 **M** KPF_6/propylene carbonate		VE	–0.333	83Mly
	0.1 **M** KPF_6/ethylene carbonate		T	–0.093	67Pay
	0.1 **M** KPF_6/acetonitrile		T	–0.0093	67Pay
	0.1 **M** KPF_6/dimethylformamide		T	–0.039	67Pay
	0.1 **M** $KClO_4$/dimethylsulfoxide		S	–0.226	87Bor
	0.2 **M** KPF_6/water		T	–0.218	77Bor
	0.2 **M** KPF_6/acetone		T	–0.058	77Bor
	0.1 **N** K_2SO_4/H_2O		EC	–0.190	52Gra
	0.1 **N** K_2SO_4/H_2O		T	–0.189	73Iva
	0.1 **m** K_2SO_4/H_2O		S	–0.851 [46])	56Ran
	1 **N** $LaCl_3/H_2O$		CA	–0.32	55Uks2
	1 **N** $LaCl_3/H_2O$		EC	–0.32	67Arg
	0.01 **M** LiCl/methanol		S	–0.392 [47])	67Gar
	0.1 **M** $LiCl/H_2O$		T	–0.279	70Tim
	1 **M** $LiCl/H_2O$		T	–0.229	70Tim
	3.1 **M** LiCl/glycerol		VE	–0.510	83Mly
	0.005...0.2 **M** $LiClO_4$/ethanediol		T	–0.089	88Jap
	0.005...0.2 **M** $LiClO_4$/1,2-propanediol		T	–0.084	88Jap
	0.005...0.2 **M** $LiClO_4$/1,3-propanediol		T	–0.154	88Jap
	0.005...0.2 **M** $LiClO_4$/1,3-butanediol		T	–0.154	88Jap
	0.005...0.2 **M** $LiClO_4$/1,4-butanediol		T	–0.129	88Jap
	0.005...0.2 **M** $LiClO_4$/1,2-butanediol		T	–0.094	88Jap
	0.005...0.2 **M** $LiClO_4$/2,3-butanediol		T	–0.089	88Jap
	0.001 **M** $LiClO_4$/acetonitrile		T	–0.049	87Ham2
	0.1 **M** $LiClO_4$/dimethylsulfoxide		S	–0.226	87Bor
	0.1 **M** $LiNO_3$/dimethylsulfoxide		S	–0.236	87Bor
	0.2 **M** $LiNO_3$/water		T	–0.253	77Bor
	0.2 **M** $LiNO_3$/acetone		T	–0.042	77Bor
	0.1 **N** $(NH_4)_2SO_4/H_2O$		T	–0.199	73Iva
	0.853 **N** $NaBr/H_2O$		EC	–0.348	36Erd
	0.1 **N** $NaBr/H_2O$		S	–0.220	62Jen

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Hg	1 **N** $NaCH_3COO/H_2O$		EC	–0.24	67Arg
	0.1 **N** $NaCl/H_2O$		EC	–0.223	91Pas
	0.1 **N** $NaCl/H_2O$		EC	–0.225	03Gou
	0.1 **N** $NaCl/H_2O$		T	–0.190	47Gra
	0.1 **N** $NaCl/H_2O$		EC	–0.235	49Gra
	0.1 **N** $NaCl/H_2O$		S	–0.226	62Jen
	1 **N** $NaCl/H_2O$		S	–0.275	62Jen
	1.18 **N** $NaCl/H_2O$		EC	–0.277	36Erd
	0.014 **M** $NaClO_4$/methanol		T	–0.326 [48])	81Bor
	0.05 **M** $NaClO_4$/methanol		T	–0.327 [48])	81Bor
	0.1 **M** $NaClO_4$/methanol		T	–0.330 [48])	81Bor
	0.1 **M** $NaClO_4$/methanol	*T*=–5 °C	T	–0.320 [48])	82Bor
	0.1 **M** $NaClO_4$/methanol	*T*=10 °C	T	–0.325 [48])	82Bor
	0.1 **M** $NaClO_4$/methanol	*T*=25 °C	T	–0.330 [48])	82Bor
	0.1 **M** $NaClO_4$/methanol	*T*=40 °C	T	–0.335 [48])	82Bor
	0.1 **M** $NaClO_4$/methanol		VE	–0.31	83Mly
	0.1 **M** $NaClO_4$/formamide		VE	–0.215	83Mly
	0.1 **M** $NaClO_4$/*N*-methylformamide	*T*=5 °C	EC	–0.398	84Bor
	0.1 **M** $NaClO_4$/*N*-methylformamide	*T*=25 °C	EC	–0.401	84Bor
	0.1 **M** $NaClO_4$/*N*-methylformamide	*T*=30 °C	EC	–0.400	84Bor
	0.1 **M** $NaClO_4$/*N*-methylformamide	*T*=40 °C	EC	–0.420	84Bor
	0.1 **M** $NaClO_4$/*N*-methylformamide		T	–0.079	96Eme
	0.1 **M** $NaClO_4$/dimethylsulfoxide		S	–0.226	87Bor
	0.2 **M** $NaClO_4$/methanol		T	–0.333 [48])	81Bor
	0.2 **M** $NaClO_4/H_2O$		T	–0.244	77Bor
	0.2 **M** $NaClO_4$/acetone		T	–0.028	77Bor
	1.08 **M** $NaClO_4/H_2O$		S	–0.250	36Erd
	1 **M** $NaClO_4/H_2O$		T	0.03	68Ham2
	1 **N** $NaClO_4/H_2O$		CA	–0.25 [44])	71Mor
	1 **M** $NaClO_4/H_2O$		T	–0.273	95Koe
	NaF/H_2O		T	–0.192	54Gra
	0.01 **N** NaF/H_2O		T	–0.193	58Dam
	0.1 **N** NaF/H_2O		EC	–0.194	52Gra
	0.1 **N** NaF/H_2O		T	–0.193	58Dam
	0.9 **N** NaF/H_2O		T	–0.192	58Dam
	1 **N** NaF/H_2O		EC	–0.192	52Gra
	0.25 **N** NaF/H_2O		S	–0.191	79Ama
	0.25 **N** NaF/H_2O		EC	–0.199	79Ama
	0.1 **N** NaI/H_2O		T	–0.446	58Dam
	0.9 **N** NaI/H_2O		T	–0.519	58Dam
	1.09 **N** NaI/H_2O		EC	–0.542	36Erd
	1.38 **N** $NaNO_3/H_2O$		EC	–0.279	36Erd
	0.1 **m** $NaOH/H_2O$		S	–0.567	56Ran
	0.1 **M** $NaOH/H_2O$		S	–0.524	62Jen
	0.514 **M** Na_2S/H_2O		EC	–0.791	36Erd

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	0.0066 **M** Na_2SO_4/H_2O		EC	–0.186	36Erd
	0.01 **N** Na_2SO_4/H_2O		EC	–0.232	38Win
	0.005 **M** Na_2SO_4/H_2O		EC	–0.194	75Pay
	0.01 **M** Na_2SO_4/H_2O		T	–0.1916	84Dag
	0.01 **M** Na_2SO_4/H_2O		IM	–0.19	95Tur
	0.01 **M** Na_2SO_4/H_2O		EC	–0.195	75Pay
	0.02 **M** Na_2SO_4/H_2O		T	–0.1914	84Dag
	0.02 **M** Na_2SO_4/H_2O		EC	–0.196	75Pay
	0.03 **M** Na_2SO_4/H_2O		T	–0.199	71Elk
	0.05 **M** Na_2SO_4/H_2O		T	–0.1912	84Dag
	0.05 **M** Na_2SO_4/H_2O		EC	–0.199	75Pay
	0.1 **N** Na_2SO_4/H_2O		EC	–0.248	38Win
	0.1 **M** Na_2SO_4/H_2O		T	–0.209	71Elk
	0.1 **M** Na_2SO_4/H_2O		EC	–0.201	75Pay
	0.1 **M** Na_2SO_4/H_2O		T	–0.1907	84Dag
	0.2 **N** Na_2SO_4/H_2O		S	–0.190	62Jen
	0.25 **N** Na_2SO_4/H_2O		S	–0.191	79Ama
	0.25 **N** Na_2SO_4/H_2O		EC	–0.199	79Ama
	0.2 **M** Na_2SO_4/H_2O		EC	–0.203	75Pay
	0.2 **M** Na_2SO_4/H_2O		T	–0.1900	84Dag
	0.3 **M** Na_2SO_4/H_2O		T	–0.229	71Elk
	1 **N** Na_2SO_4/H_2O		EC	–0.213	38Win
	0.5 **M** Na_2SO_4/H_2O		T	–0.1894	84Dag
	0.5 **M** Na_2SO_4/H_2O		EC	–0.206	75Pay
	1 **M** Na_2SO_4/H_2O		EC	–0.210	75Pay
	1 **M** Na_2SO_4/H_2O		T	–0.1892	84Dag
	1.01 **M** Na_2SO_4/H_2O		EC	–0.169	36Erd
	Na_2SO_4/H_2O		CA	–0.19	67Vet
	0.5 **M** Na_2SO_4+0.001 **M** H_2SO_4/ H_2O		T	–0.148	65Sat
	0.01 **M** Na_2SO_4/H_2O+1-propanol	3 wt%	IM	0.01	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	6.3 wt%	IM	0.039	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	15.3 wt%	IM	0.036	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	25.1 wt%	IM	0.023	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	32.1 wt%	IM	0.018	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	43.4 wt%	IM	0.006	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	56 wt%	IM	0.004	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	77 wt%	IM	–0.024	95Tur
	0.1 **M** Na_2SO_4/H_2O		EC	–0.24	60Del
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	1< pH <7	T	–0.26	57Khe
	0.02 **N** (Na_2SO_4+ NaOH)/H_2O	7< pH <10	T	–0.26	57Khe
	0.02 **N** (Na_2SO_4+NaOH)/H_2O	pH=11	T	–0.255	57Khe
	0.02 **N** (Na_2SO_4+NaOH)/H_2O	pH=12	T	–0.2	57Khe
	1 **N** Na_2SO_4/H_2O		EC	–0.2	28Fru2
	1 **N** Na_2SO_4/H_2O		CA	–0.12	34Gor

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Hg	1 **N** Na_2SO_4/H_2O		CA	–0.2	55Uks2
	1 **N** Na_2SO_4/H_2O		EC	–0.2	67Arg
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O		EC	–0.22	65Pol
	n.a.		EC	–0.22	32Pro
Molybdenum (Mo)	not reported		n.a.	–0.62	53Vas
Nickel (Ni)	0.0015 **N** HBr/H_2O		ER	–0.28	40Leg
	0.0015 **N** HCl/H_2O		ER	–0.28	40Leg
	0.001 **N** HCl/H_2O		ER	–0.06	39Luk
	1 **M** HCl/H_2O		OA	–0.264	52Hil
	0.0001 **M** $HClO_4/H_2O$		IM	0.03	73Kim
	0.01 **M** KCl/H_2O		IM	0.193	62Jak1
	0.0005 **M** $KClO_4/H_2O$	pH=10.3	T	–0.24	69Boc1
	0.01 **N** K_2SO_4/H_2O	pH=3	T	–0.33	70Vol
	0.01 **M** LiCl/methanol		IM	0.088	64Jak2
	0.01 **M** LiCl/ethanol		IM	–0.02	64Jak2
	0.01 **N** LiCl/2-butanol		IM	–0.105	65Jak1
	0.01 **N** LiCl/2-propanol		IM	–0.1	65Jak1
	0.01 **N** $NaClO_4/H_2O$		OA	–0.32	69Boc1
	0.1 **N** $NaClO_4/H_2O$	pH=12.5	F	–0.3	69Boc1
	0.0001 **M** $NaOH/H_2O$		T	–0.27	69Boc1
	0.0005 **M** Na_2SO_4/H_2O	pH=3.5	PC	–0.110	83Ohm
	0.01 **N** Na_2SO_4/H_2O		S	–0.37	78Laz
	0.01 **M** Na_2SO_4/H_2O		IM	0.034 [20])	90Sok
	0.01 **M** Na_2SO_4/H_2O		IM	–0.19	95Tur
	0.01 **M** Na_2SO_4/H_2O		IM	–0.18	97Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	3 wt%	IM	–0.12	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	6.3 wt%	IM	–0.087	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	9.4 wt%	IM	–0.11	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	11 wt%	IM	–0.125	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	15.3 wt%	IM	–0.12	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	25.1 wt%	IM	–0.17	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	32.1 wt%	IM	–0.185	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	43.4 wt%	IM	–0.18	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	56 wt%	IM	–0.19	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	63.8 wt%	IM	–0.2	95Tur
	0.01 **M** Na_2SO_4/H_2O+1-propanol	77 wt%	IM	–0.2	95Tur
	0.02 **N** Na_2SO_4/H_2O	pH=1	T	–0.21	53Khe
	0.02 **N** Na_2SO_4/H_2O	pH=3	T	–0.33	53Khe
	0.02 **N** Na_2SO_4+ H_2SO_4/H_2O	pH=3.2	T	–0.328	57Kra
	0.02 **N** Na_2SO_4/H_2O	pH=5	T	–0.37	53Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=1	T	–0.205	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=2	T	–0.27	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=3	T	–0.3	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=4	T	–0.325	57Khe

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	0.02 **N** (Na_2SO_4+ H_2SO_4)/H_2O	pH=5	T	–0.305	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=6	T	–0.355	57Khe
Ni (100)	3 m**M** $HClO_4$/H_2O		T	–0.39	89Aro
Ni (110)	3 m**M** $HClO_4$/H_2O		T	–0.53	89Aro
Ni-Zn alloy	0.01 **N** K_2SO_4/H_2O	1 % Zn, pH=3	T	–0.31	70Vol
	0.01 **N** K_2SO_4/H_2O	2 % Zn, pH=3	T	–0.32	70Vol
	0.01 **N** K_2SO_4/H_2O	2.8 % Zn, pH=3	T	–0.41	70Vol
	0.01 **N** K_2SO_4/H_2O	3 % Zn, pH=3	T	–0.43	70Vol
	0.01 **N** K_2SO_4/H_2O	3.5 % Zn, pH=3	T	–0.48	70Vol
	0.01 **N** K_2SO_4/H_2O	4.1 % Zn, pH=3	T	–0.54	70Vol
	0.01 **N** K_2SO_4/H_2O	5 % Zn, pH=3	T	–0.58	70Vol
	0.01 **N** K_2SO_4/H_2O	6.3 % Zn, pH=3	T	–0.62	70Vol
	0.01 **N** K_2SO_4/H_2O	7 % Zn, pH=3	T	–0.63	70Vol
	0.01 **N** K_2SO_4/H_2O	8 % Zn, pH=3	T	–0.63	70Vol
Niobium	1 **N** H_2SO_4/H_2O		T	–0.79 [49])	65Koz
(Nb)	1 **N** H_2SO_4/H_2O		T	–0.66 [50])	65Koz
Palladium	1 **N** K_2SO_4/H_2O	pH=7	S	–0.0	69And
(Pd)	0.001 **M** $LiClO_4$/acetonitrile		T	0.279	80Pet
	0.003 **M** $LiClO_4$/ dimethylsulphoxide		T	0.241	84Ale
	1 **N** $NaClO_4$/H_2O	pH=7	S	–0.0	69And
	5 m**M** NaF/H_2O [51])		T	0.031	06ElA
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=2	T	0.2	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=3	T	–0.06	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=4	T	–0.227	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=6	T	–0.31	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=7	T	–0.36	57Khe
	0.02 **N** (Na_2SO_4+NaOH)/H_2O	pH=8	T	–0.375	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=9	T	–0.385	57Khe
	0.02 **N** (Na_2SO_4+NaOH)/H_2O	pH=11	T	–0.39	57Khe
	0.01 **N** $HClO_4$/H_2O		CO	0.25	00Alv
Pd (111)	5 m**M** NaF/H_2O		T	0.117 [52])	02ElA
	5 m**M** NaF/H_2O		T	0.151 [53])	02ElA
	5 m**M** NaF/H_2O		T	0.141 [54])	02ElA
	5 m**M** NaF/H_2O		T	0.119 [55])	02ElA

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Platinum [56])	0.01 **N** $CdBr_2$+0.1 **N** H_2SO_4/H_2O	(Pt)	IA	0.45	66Man
(Pt)	0.01 **N** CdI_2+0.1 **N** H_2SO_4/H_2O	(Pt)	IA	0.01	66Man
	0.01 **N** $CdSO_4$+0.1 **N** H_2SO_4/H_2O	(Pt)	IA	0.65	66Man
	0.005 **N** Cs_2SO_4/H_2O	pH=2.7 (Pt)	IA	0.18	61Bal
	0.001...0.01 **N** Cs_2SO_4/H_2O	2.5<pH<6	RT	0.18	68Bal
	0.025 **M** Cs_2SO_4+0.005 **M** H_2SO_4/H_2O	(H)	IA	0.11	33Fru
	0.01 **N** Cs_2SO_4+0.01 **N** H_2SO_4/H_2O	(Pt)	RT	0.17	64Kaz
	0.01 **N** HCl + 0.1 **N** KCl/H_2O		ISS	0.132	72Pod
	1 **N** HCl/H_2O		T	0.47	53Rob
	0.015 **M** HCl+0.12 **M** KCl/H_2O		ISS	0.13	74Pod
	$HClO_4/H_2O$		T	0.1	69Bur
	0.1 **N** $HClO_4/H_2O$		F	0.5	69Boc1
	0.1 **N** $HClO_4/H_2O$		IM	1.1	96Ham1
	$HClO_4$,$NaClO_4$,$NaOH/H_2O$	2.5<pH<11.2, $c_{total}=3\cdot10^{-3}$ **M**	T	0.56–0.059 pH	66Gil
	0.3 **N** HF+0.1 **N** KF/H_2O		ISS	0.233	72Pod
	0.3 **M** HF+0.12 **M** KF/H_2O		ISS	0.235	74Pod
	0.000025 **N** H_2SO_4/H_2O	(H)	ED	0.12	38Bal
	0.001 **N** H_2SO_4/H_2O		T	0.18	63Bir
	0.001 **N** H_2SO_4/H_2O		DR	0.2	63Vor
	0.001 **N** H_2SO_4/H_2O		AE	0.21	66Kuk2
	0.005 **N** H_2SO_4/H_2O		T	0.35	53Bor
	0.005 **M** H_2SO_4+0.5 **M** Na_2SO_4/H_2O		ISS	0.2	74Pod
	0.005 **M** H_2SO_4+0.06 **M** K_2SO_4/H_2O		ISS	0.205	74Pod
	0.01 **N** H_2SO_4+0.1 **N** K_2SO_4/H_2O		ISS	0.205	72Pod
	0.01 **N** H_2SO_4/H_2O		T	0.35	53Bor
	0.01 **N** H_2SO_4/H_2O		T	0.18	63Bir
	0.1 **N** H_2SO_4/H_2O		F	0.3	50Bow
	H_2SO_4, $NaOH/H_2O$	0<pH<14	PE	0.32–0.04 pH	86Seo
	1 **M** KBr+0.01 **M** HBr/H_2O	(Pt)	ISS	–0.09	68Pet
	1 **M** KBr+0.01 **M** $NaOH/H_2O$	(Pt)	ISS	–0.39	68Pet
	0.001 **M** KCl/H_2O		DR	0.2	59Vor
	0.001 **M** KCl/H_2O		DR	0.2	62Vor
	0.001 **M** KCl/H_2O		DR	0.2	63Vor
	0.01 **M** KCl/H_2O		DR	0.2	63Vor
	0.1 **M** KCl/H_2O		ED	0.09	71Fre
	0.1 **M** KCl/H_2O	pH=12	S	–0.4	64Per
	1 **N** KCl+0.01 **N** HCl/H_2O	(Pt)	ISS	0.04	69Pet
	1 **M** KCl/H_2O		S	–0.08	74Cla2
	0.1 **N** KF/H_2O		CA	–0.06	71Mor
	0.1 **N** KI/H_2O		S	–0.51	69And
	1 **M** KI+0.01 **M** $NaOH/H_2O$	(Pt)	ISS	–0.52	68Pet
	1 **N** K_2SO_4/H_2O	pH=7	S	–0.02	69And

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	1 **N** K_2SO_4/H_2O	pH=11	S	–0.3	69And
	0.0005 **M** $LiClO_4$/acetonitrile		T	0.099	80Pet
	0.003 **M** $LiClO_4$/dimethylsulphoxide		T	0.091	84Ale
	0.00001 **N** $MgSO_4/H_2O$		DR	0.2	63Vor
	0.0003 **N** $MgSO_4/H_2O$		DR	0.2	63Vor
	0.001 **N** $MgSO_4/H_2O$		DR	0.2	63Vor
	0.001 **N** NaBr+001 **N** H_2SO_4/H_2O	(Pt)	IA	0.04	65Bal1
	1 **N** NaBr+0.01 **N** HBr/H_2O	(Pt)	IA	–0.02	36Shl
	1 **N** NaBr+0.05 **N** $NaOH/H_2O$	(Pt)	IA	–0.26	36Shl
	0.001 **N** NaCl+0.001 **N** HCl/H_2O	(Pt)	IA	0.14	65Bal1
	1 **N** NaCl+0.01 **N** HCl/H_2O	(Pt)	IA	0.05	36Shl
	$NaClO_4/H_2O$		T	0.1	69Bur
	0.01 **N** $NaClO_4/H_2O$	pH=3	OA	0.5	69Boc1
	0.003 **N** $NaClO_4$+$HClO_4/H_2O$	pH=2.5	T	0.47	69Boc1
	0.003 **N** $NaClO_4$+$HClO_4/H_2O$	pH=2.8	T	0.41	69Boc1
	0.003 **N** $NaClO_4$+$HClO_4/H_2O$	pH=3.3	T	0.37	69Boc1
	0.003 **N** $NaClO_4$+$HClO_4/H_2O$	pH=4.1	T	0.33	69Boc1
	0.003 **N** $NaClO_4$+$NaOH/H_2O$	pH=8.0	T	0.07	69Boc1
	0.003 **N** $NaClO_4$+$NaOH/H_2O$	pH=10.3	T	–0.083	69Boc1
	0.003 **N** $NaClO_4$+$NaOH/H_2O$	pH=11.2	T	–0.11	69Boc1
	0.1 **N** $NaClO_4/H_2O$	pH=12	OA	–0.07	69Boc1
	0.1 **N** $NaClO_4/H_2O$	pH=2.3	F	0.35	69Boc1
	0.1 **N** $NaClO_4/H_2O$	pH=12	F	–0.2	69Boc1
	1 **N** $NaClO_4/H_2O$	pH=7	S	–0.02	69And
	1 **N** $NaClO_4/H_2O$	pH=11	S	–0.3	69And
	0.001 **N** NaF+0.001 **N** H_2SO_4/H_2O	(Pt)	IA	0.18	65Bal1
	0.005 **N** NaI+005 **N** H_2SO_4/H_2O	(Pt)	IA	–0.5	65Bal1
	0.001 **N** Na_2SO_4+0.001 **N** H_2SO_4/H_2O	(Pt)	IA	0.18	65Bal1
	0.001...0.01 **N** Na_2SO_4/H_2O	2.5< pH <6	RT	0.18	68Bal
	0.025 **M** Na_2SO_4+0.005 **M** H_2SO_4/ H_2O	(H)	IA	0.11	36Shl
	0.02 **N** Na_2SO_4/H_2O	pH=2 (H)	T	0.48	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=2 (H)	T	0.48	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=3 (H)	T	0.3	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=4 (H)	T	0.18 [57])	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=4 (H)	T	0.13	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=5 (H)	T	0.0	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=5 (H)	T	0.12 [57])	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=6 (H)	T	–0.14	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=6 (H)	T	0.08	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=7 (H)	T	–0.21	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=7 (H)	T	0.06 [57])	57Khe
	0.02 **N** Na_2SO_4/H_2O	pH=7 (H)	T	–0.21	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=8 (H)	T	–0.28	57Khe

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Pt	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=8 (H)	T	0.03 [57])	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=9 (H)	T	–0.31	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=9 (H)	T	0.01 [57])	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=10 (H)	T	–0.35	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=10 (H)	T	0.01 [57])	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=11 (H)	T	–0.36	57Khe
	0.02 **N** (Na_2SO_4+H_2SO_4)/H_2O	pH=12 (H)	T	0.36	57Khe
	0.1 **N** Na_2SO_4/H_2O		T	0.22	59McM
	1 **N** Na_2SO_4/H_2O	acidic	CA	0.28	34Gor
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	(Pt)	ISS	0.16	69Pet
	1 **N** Na_2SO_4/H_2O		T	0.11	35Fru
	0.001 **N** $ZnCl_2/H_2O$		DR	0.2	63Vor
Pt (111)	0.1 **M** $HClO_4/H_2O$		CO	0.25	00Gom
	0.1 **M** $HClO_4/H_2O$		CO	0.25... –0.75 [58])	00Gom
	0.1 **M** $HClO_4/H_2O$		CO	0.19... 0.33 [58])	99Cli
	0.1 **M** $HClO_4/H_2O$		ER	0.277	95Att
	0.1 **M** $HClO_4/H_2O$		VIB	0.287	96Iwa
Potassium (K)	not reported		n.a.	–2.31	53Vas
Rhodium (Rh)	0.001 **N** Cs_2SO_4/H_2O		OA	0.18	64Kra
	0.01 **N** HCl+1 **N** KCl/H_2O		ISS	0.0 [59])	67Pet2
	0.01 **N** HCl+1 **N** KCl/H_2O		ISS	0.0	70Pet
	0.3 **M** HF+0.12 **M** KF/H_2O		ISS	0.085	74Pod
	0.02 **N** H_2SO_4/H_2O		T	0.28	64Kra
	0.005 **M** H_2SO_4+0.06 **M** K_2SO_4/H_2O		ISS	0.03	74Pod
	0.005 **M** H_2SO_4+0.5 **M** Na_2SO_4/H_2O		ISS	0.025	74Pod
	0.01 **N** H_2SO_4+1 **N** Na_2SO_4/H_2O		ISS	0.5	70Pet
	KCl+0.01 **N** HCl/H_2O		ISS	–0.12	68Fru [60])
	1 **N** $NaClO_4/H_2O$	pH=7	S	–0.02	69And
	1 **N** $NaClO_4/H_2O$	pH=11	S	–0.36	69And
	0.01 **N** NaOH+1 **N** Na_2SO_4/H_2O		ISS	–0.4	74Pod
	Na_2SO_4/H_2O		RT	–0.03	63Bir
	0.001 **N** Na_2SO_4/H_2O		OA	0.18	64Kra
	0.01 **N** Na_2SO_4/H_2O		ISS	0.15	82Laz
	1 **N** Na_2SO_4/H_2O	pH=2.5	ISS	0.02	82Laz
	Na_2SO_4+0.01 **N** H_2SO_4/H_2O		ISS	–0.04	68Fru [60])
	0.001 **N** Na_2SO_4+0.001 **N** H_2SO_4/H_2O		ISS	–0.07 [59])	67Pet2
	0.01 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O		RT	–0.03	67Bal
	1 **N** K_2SO_4/H_2O	pH=7	S	–0.02	69And
	1 **N** K_2SO_4/H_2O	pH=11	S	–0.36	69And
Rubidium (Rb)	not reported		n.a.	–2.42	53Vas

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Silicon (Si)	not reported		n.a.	–0.13	53Vas
Silver (Ag)	$AgNO_3/H_2O$		IA,IM	–0.51	31Pro
	$AgNO_3/H_2O$		S	–0.475	26Ben
	0.001 **M** $HClO_4/H_2O$		T	–0.45	69Boc1
	0.001 **M** $HClO_4/H_2O$		BAW	–0.683	97Che
	0.002 **M** $HClO_4/H_2O$		BAW	–0.683	97Che
	0.005 **M** $HClO_4/H_2O$		BAW	–0.666	97Che
	0.001 **M** KCl+0.199 **M** NaF/H_2O		T	–0.719	78Shl
	0.002 **M** $HClO_4/H_2O$		T	–0.41	69Boc1
	0.01 **N** KCl/H_2O		T	–0.86	67Dag
	0.01 **M** KCl+0.19 **M** NaF/H_2O		T	–0.704	78Shl
	0.05 **M** KCl+0.15 **M** NaF/H_2O		T	–0.679	78Shl
	0.1 **N** KCl/H_2O		S	–0.82	64And
	0.1 **N** KCl/H_2O	pH=7	S	–0.80	64Per
	0.1 **N** KCl/H_2O		S	–0.80	67Per
	0.1 **M** KCl+0.1 **M** NaF/H_2O		T	–0.669	78Shl
	0.1 **M** $KClO_4/H_2O$		SFG	–0.709	95Tad
	0.1 **M** $KClO_4$+0.01 **M** KCN/H_2O		SFG	–0.959	95Tad
	0.1 **N** KNO_3/H_2O		various methods	0.046	39Ves
	0.1 **N** KNO_3/H_2O		RT	0.05	55Fru
	0.01 **N** KOH/H_2O		T	–0.7	66Dag
	0.1 **N** KOH/H_2O		T	–0.7	74Zel
	1 **N** K_2SO_4/H_2O	pH=7	S	–0.64	69And
	1 **N** K_2SO_4/H_2O	pH=11	S	–0.64	69And
	0.1 **N** $NaBr/H_2O$		S	–0.78	67Bod
	0.1 **N** $NaCl/H_2O$		S	–0.62	67Bod
	0.005 **M** $NaClO_4/H_2O$		T	–0.66	02Lan[61])
	0.01 **N** $NaClO_4/H_2O$		T	–0.7	67Ham2
	0.1 **N** $NaClO_4/H_2O$		S	–0.48	67Bod
	0.1 **N** $NaClO_4/H_2O$		OA	0.08	69Boc1
	0.1 **N** $NaClO_4/H_2O$	pH=2	F	–0.4	69Boc1
	0.5 **M** $NaClO_4/H_2O$		T	–0.719	83Hup
	1 **N** $NaClO_4/H_2O$	pH=7	S	–0.64	69And
	1 **N** $NaClO_4/H_2O$	pH=11	S	–0.64	69And
	0.001 **M** NaF/H_2O		T	–0.7	91Agl
	0.001 **M** NaF/H_2O		BAW	–0.662	97Che
	0.005 **M** NaF/H_2O		T	–0.72	72Val
	0.005 **M** NaF/H_2O	pH=4	T	–0.7	78Nov
	0.005 **M** NaF/H_2O		T	–0.729	78Shl
	0.005 **M** NaF/H_2O		T	–0.7	78Zel1
	0.005 **M** NaF/H_2O		BAW	–0.664	97Che
	0.1 **M** NaF/H_2O		BAW	–0.673	97Che
	0.1 **M** $NaOH/H_2O$		CA	–0.46	69Mor

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Ag	0.1 **N** $NaOH/H_2O$		S	–0.56	67Bod
	0.0005 **M** Na_2SO_4/H_2O		BAW	–0.675	97Che
	0.001 **N** $NaSO_4/H_2O$		T	–0.7	66Dag
	0.001 **N** $NaSO_4/H_2O$		T	–0.7	66Dag
	0.005 **N** $NaSO_4/H_2O$		T	–0.7	66Lei
	0.01 **N** Na_2SO_4/H_2O		T	–0.7	60Lei
	0.01 **N** Na_2SO_4/H_2O		T	–0.7	66Dag
	0.01 **N** Na_2SO_4/H_2O	pH=3.4	T	–0.7	66Dag
	0.01 **N** Na_2SO_4/H_2O		T	–0.709 [5])	70And
	0.01 **N** Na_2SO_4/H_2O		T	–0.709 [62])	70And
	0.01 **N** Na_2SO_4/H_2O		T	–0.7	75Ale1
	0.01 **M** Na_2SO_4/H_2O		T	–0.66	75Non
	0.01 **M** Na_2SO_4/H_2O		S	–0.65 [14])	75Non
	0.2 **N** Na_2SO_4/H_2O		PE	–0.75	66Gok
	0.02 **N** (Na_2SO_4+ H_2SO_4)/H_2O	1< pH <6	T	0.05	57Khe
	0.02 **N** Na_2SO_4/H_2O		T	0.05	57Khe
	0.02 **N** Na_2SO_4/H_2O		T	–0.7	60Lei
	0.1 **M** Na_2SO_4/H_2O		CA	–0.58	69Mor
	0.1 **M** Na_2SO_4/H_2O		BAW	–0.668	97Che
	1 **N** Na_2SO_4/H_2O	acidic	CA	–0.02	34Gor
	0.05 **M** $KClO_4$/acetonitrile		T	–0.744	93Dou
	0.4 **M** N LiBr/ethanol		VIB	–1.0 [30])	92Sob
	0.4 **M** N LiBr/propanol		VIB	–1.0 [30])	92Sob
	0.4 **M** N LiBr/pentanol		VIB	–1.0 [30])	92Sob
	0.4 **M** N LiBr+0.1 **M** LiCl/butan-1-ol		VIB	–0.94 [30])	96Pem
	0.4 **M** N LiBr+0.1 **M** LiCl/butan-2-ol		VIB	–1.3 [30])	96Pem
	0.4 **M** N LiBr+0.1 **M** LiCl/ 2-methylpropan-1-ol		VIB	–1.45 [30])	96Pem
	n.a.		EC	–0.510 [21])	32Pro
$Ag_{Au(111)}$ [63])	0.1 **M** $KClO_4/H_2O$		T	–0.259	06Sol
$Ag_{Ir(111)}$ [63])	0.1 **M** $KClO_4/H_2O$		T	–0.559	06Sol
$Ag_{Pd(111)}$ [63])	0.1 **M** $KClO_4/H_2O$		T	–0.359	06Sol
$Ag_{Pt(111)}$ [63])	0.1 **M** $KClO_4/H_2O$		T	–0.239	06Sol
$Ag_{Rh(111)}$ [63])	0.1 **M** $KClO_4/H_2O$		T	–0.359	06Sol
Ag (100)	0.0025 **M** Cs_2SO_4/H_2O		T	–0.649	75Vit2
	1 **N** HCl/H_2O		H,F	–0.28	57Fle
	0.002 **M** KF/H_2O		T	–0.609	75Vit2, 76Vit2
	0.002 **M** KF/H_2O		T	–0.608	82Vit
	0.0025 **M** K_2SO_4/H_2O		T	–0.599	69Ham1
	0.005 **M** KPF_6/H_2O		T	–0.624	82Val
	0.1 **M** $LiClO_4/H_2O$		T	–0.619	02Jur
	0.0025 **M** Li_2SO_4/H_2O		T	–0.669	75Vit1
	0.005 **M** NaF/H_2O		T	–0.669	73Val
	0.005 **M** NaF/H_2O		T	–0.624	82Val

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	0.005 **M** NaF/H_2O		T	–0.616	95Tra
	0.1 **M** NaF/H_2O		T	–0.669	81Fle
	0.1 **M** NaF/H_2O		T	–0.639	02Jur
	0.002 **M** $NaClO_4/H_2O$		T	–0.647	82Vit
	0.005 **M** $NaClO_4/H_2O$		T	–0.632	82Val
	0.0005 **M** Na_2SO_4/H_2O		T	–0.669	74Vit1
	0.001 **N** Na_2SO_4/H_2O		T	–0.7	66Dag
	0.0025 **M** Na_2SO_4/H_2O		T	–0.669	75Vit1
	0.005 **N** Na_2SO_4/H_2O		T	–0.650	69Bud
	0.005 **M** Na_2SO_4/H_2O		T	–0.609	75Vit1, 76Vit1
	0.001 **M** $NaNO_3/H_2O$		T	–0.659	75Vit1
Ag (110)	0.005 **M** KBF_4/H_2O		T	–0.734	81Val
	0.05 **M** $KClO_4/H_2O$		T	–0.749	89Ham
	0.002 **M** KF/H_2O		T	–0.609	82Vit
	0.005 **M** KPF_6/H_2O		T	–0.734	89Ham
	0.005 **M** KPF_6/H_2O		T	–0.734	81Val
	0.0025 **M** K_2SO_4/H_2O		T	–0.659	69Ham1
	0.1 **M** $LiClO_4/H_2O$		T	–0.734	02Jur
	0.005 **M** NaF/H_2O		T	–0.769	71Val2, 73Val
	0.005 **M** NaF/H_2O		T	–0.734	81Val
	0.1 **M** NaF/H_2O		T	–0.759	81Fle
	0.1 **M** NaF/H_2O		T	–0.649	02Jur
	0.002 **M** $NaClO_4/H_2O$		T	–0.653	82Vit
	0.005 **M** $NaClO_4/H_2O$		T	–0.734	81Val
	0.005 **M** Na_2SO_4/H_2O		T	–0.719	74Vit1
Ag (111)	0.0025 **M** Cs_2SO_4/H_2O		T	–0.499	75Vit2
	1 **N** HCl/H_2O		H,F	–0.18	57Fle
	0.1 **M** $KClO_4/H_2O$		T	–0.459	06Sol
	0.001 **M** KF/H_2O		T	–0.459 [64])	72Sev
	0.002 **M** KF/H_2O		T	–0.459	82Vit
	(100–x) m**M** KF + x m**M** KCl/H_2O	(1< x <40)	T	–0.519	77Vit
	0.1 **M** KPF_6/H_2O		T	–0.454	89Val
	0.0025 **M** K_2SO_4/H_2O		T	–0.509	69Ham1
	0.0025 **M** Li_2SO_4/H_2O		T	–0.519	75Vit2
	0.1 **M** NaF/H_2O		T	–0.449	81Fle
	0.005 **M** NaF/H_2O		T	–0.449	73Val, 71Val1, 71Val2
	0.01 **M** NaF/H_2O		T	–0.469	87Ham2
	0.002 **M** $NaClO_4/H_2O$		T	–0.479	82Vit
	0.005 **M** $NaClO_4/H_2O$		T	–0.46	02Lan [65])
	0.005 **M** $NaClO_4/H_2O$		T	–0.24	02Lan [66])
	0.001 **M** $NaNO_3/H_2O$		T	–0.499	75Vit2

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Ag (111)	0.005 **N** Na_2SO_4/H_2O		T	–0.500	69Bud
	0.005 **M** Na_2SO_4/H_2O		T	–0.479	72Sev
	0.005 **M** Na_2SO_4/H_2O		T	–0.519	78Vit
	0.0005 **M** Na_2SO_4/H_2O		T	–0.519	74Vit1
	0.0025 **M** Na_2SO_4/H_2O		T	–0.519	75Vit2
	0.005 **M** Na_2SO_4/H_2O		T	–0.519	77Vit
Ag (311)	0.005 **M** KPF_6/H_2O		T	–0.666	90Ham
	0.05 **M** $KClO_4/H_2O$		T	–0.709	90Ham
Ag/Au alloy	2 **N** H_2SO_4/H_2O	100 % Ag/ 0 % Au	T	–0.700	72Lev
	2 **N** H_2SO_4/H_2O	15 % Ag/ 85 % Au	T	–0.600	72Lev
	2 **N** H_2SO_4/H_2O	12 % Ag/ 88 % Au	T	–0.500	72Lev
	2 **N** H_2SO_4/H_2O	9 % Ag/ 91 % Au	T	–0.370	72Lev
	2 **N** H_2SO_4/H_2O	6 % Ag/ 94 % Au	T	–0.300	72Lev
	2 **N** H_2SO_4/H_2O	3 % Ag/ 97 % Au	T	–0.250	72Lev
	2 **N** H_2SO_4/H_2O	0 % Ag/ 100 % Au	T	0.070	72Lev
Sodium (Na)	not reported		n.a.	–2.25	53Vas
Na amalgam	1 **N** $NaOH/H_2O$	$c_{Na\ in\ Hg}$= 0.000126	EC	–1.69	57Smi
	1 **N** $NaOH/H_2O$	$c_{Na\ in\ Hg}$= 0.000783	EC	–1.73	57Smi
	1 **N** $NaOH/H_2O$	$c_{Na\ in\ Hg}$= 0.000587	EC	–1.78	57Smi
	1 **N** $NaOH/H_2O$	$c_{Na\ in\ Hg}$= 0.00174	EC	–1.81	57Smi
	1 **N** $NaOH/H_2O$	$c_{Na\ in\ Hg}$= 0.0558	EC	–1.82	57Smi
Strontium (Sr)	not reported		n.a.	–2.32	53Vas
Tantalum	1 **N** H_2SO_4/H_2O		T	–0.85 [49])	65Koz
(Ta)	1 **N** H_2SO_4/H_2O		T	–0.73 [50])	65Koz
Tellurium	1 **N** H_2SO_4/H_2O		H	–0.6	56Ven
(Te)	1 **N** $NaCl/H_2O$		H	–0.02	44Reb
	1 **N** Na_2SO_4/H_2O		H	0.56	49Ven
	1 **N** Na_2SO_4/H_2O		H	0.6	45Reb, 52Ven, 55Fru
	1 **N** $NaCl/H_2O$		H	0.510	49Ven

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Thallium (Tl)	1 **N** KBr/H_2O		H	–0.85	44Reb
	1 **N** KBr/H_2O		H	–0.82	49Ven
	0.001 **N** KCl/H_2O		T	–0.82	55Fru, 50Bor
	0.01 **N** KCl/H_2O		T	–0.82	50Bor
	0.01 **N** KCl/H_2O		T	–0.71	67Dag
	0.1 **N** KCl/H_2O		T	–0.84	50Bor
	1 **N** Na_2CO_3/H_2O		H	–0.69	49Ven
	0.001 **N** NaF/H_2O		T	–0.67	67Dag
	0.003 **N** NaF/H_2O		T	–0.67	75Gri
	Na_2SO_4/H_2O		H	–0.69	55Fru
	0.1 **N** Na_2SO_4/H_2O		H	–0.69	49Ven
	1 **N** Na_2SO_4/H_2O		H	–0.69	49Ven
Tl amalgam [67])	1 **N** Na_2SO_4/H_2O	99.009 % Hg/ 0.0011 % Tl	S	–0.214	30Fru
	1 **M** $NaClO_4/H_2O$	1.4 mol% In	T	–0.434	95Koe
	1 **M** $NaClO_4/H_2O$	2 mol% In	T	–0.469	95Koe
	1 **M** $NaClO_4/H_2O$	7.9 mol% In	T	–0.544	95Koe
	1 **M** $NaClO_4/H_2O$	10 mol% In	T	–0.561	95Koe
	0.1 **N** $HClO_4/H_2O$	10.16 % Tl	S	–0.51	65But1
	0.1 **N** $HClO_4/H_2O$	19.88 % Tl	S	–0.52	65But1
	1 **M** $NaClO_4/H_2O$	20 mol% In	T	–0.622	95Koe
	1 **M** $NaClO_4/H_2O$	30 mol% In	T	–0.673	95Koe
	0.1 **N** $HClO_4/H_2O$	30.39 % Tl	S	–0.55	65But1
	1 **M** $NaClO_4/H_2O$	40 mol% In	T	–0.698	95Koe
	0.002 **N** NaF/H_2O	89.8 % Hg/ 1.2 % Tl	T	–0.503	60Bog
	0.002 **N** NaF/H_2O	95 % Hg/ 5 % Tl	T	–0.548	60Bog
	0.002 **N** NaF/H_2O	90.5 % Hg/ 9.5 % Tl	T	–0.575	60Bog
	1 **N** Na_2SO_4/H_2O	99 % Hg/ 1 % Tl	EC	–0.39	28Fru2
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	99 % Hg/ 1 % Tl	S	–0.388	30Fru
	1 **N** Na_2SO_4+0.1 **N** $NaOH/H_2O$	99 % Hg/ 1 % Tl	S	–0.388	30Fru
	1 **N** Na_2SO_4/H_2O	96.65 % Hg/ 3.35 % Tl	EC	–0.45	28Fru2
	1 **N** Na_2SO_4/H_2O	acidic, 59 % Hg/41 % Tl	CA	–0.67	34Gor
	0.1 **M** Na_2SO_4/H_2O	96.55 % Hg/ 3.45 % Tl	EC	–0.329	60Del
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	90 % Hg/ 10 % Tl	S	–0.512	30Fru

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Tl amalgam [67])	1 **N** Na_2SO_4+0.1 **N** NaOH/H_2O	99 % Hg/ 1 % Tl	S	–0.388	30Fru
	0.1 M Na_2SO_4/H_2O	90 % Hg/ 10 % Tl	EC	–0.439	60Del
	1 **N** Na_2SO_4/H_2O	89.65 % Hg/ 10.35 % Tl	EC	–0.52	28Fru2
	0.002 **N** NaF/H_2O	80.78 % Hg/ 19.22 % Tl	T	–0.596	60Bog
	0.1 M Na_2SO_4/H_2O	69 % Hg/ 31 % Tl	EC	–0.609	60Del
	1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O	67 % Hg/ 33 % Tl	S	–0.618	30Fru
	1 **N** Na_2SO_4+0.1 **N** NaOH/H_2O	67 % Hg/ 33 % Tl	S	–0.388	30Fru
	1 **N** Na_2SO_4/H_2O	66.1 % Hg/ 33.9 % Tl	EC	–0.64	28Fru2
	0.002 **N** NaF/H_2O	60 % Hg/ 40 % Tl	T	–0.645	60Bog
	1 **N** Na_2SO_4/H_2O	58.5 % Hg/ 41.5 % Tl	EC	–0.65	28Fru2
	Na_2SO_4/H_2O	58.5 % Hg/ 41.5 % Tl	T	–0.65	28Fru1
	1 **N** Na_2SO_4/H_2O	Hg/sat. Tl	EC	–0.65	55Fru
Tin	0.01 **M** KCl/H_2O		IM	–0.382	62Jak1
(Sn)	0.001 **N** $KClO_4$/H_2O		T	–0.46	65Fru2
	0.002 **N** $KClO_4$/H_2O		T	–0.38	71Erl
	1 **N** K_2SO_4/H_2O	pH=7	S	–0.02	69And
	0.01 **M** LiCl/methanol		IM	–0.532	64Jak2
	0.01 **M** LiCl/ethanol		IM	–0.573	64Jak2
	0.01 **N** LiCl/1-butanol		IM	–0.685	68Jak1
	0.01 **N** LiCl/2-butanol		IM	–0.655	65Jak1
	0.01 **N** LiCl/2-propanol		IM	–0.655	65Jak1
	0.01 **N** LiCl/1-pentanol		IM	–0.63	68Jak1
	0.01 **N** LiCl/quinoline		IM	–1.224	68Jak1
	0.0007 **N** $NaClO_4$/H_2O	pH=3.1	T	–0.43	68Ham1
	0.0025 **N** Na_2SO_4/H_2O		T	–0.42	70Bar2
	0.01 **N** Na_2SO_4/H_2O		T	–0.35	65Gri
	0.1 **N** Na_2SO_4+0.01 **N** H_2SO_4/H_2O		H	–0.24	52Ven
	Na_2SO_4/H_2O		T	–0.38	73Gri
	1 **N** $NaClO_4$/H_2O	pH=7	S	–0.3	69And
Sn (001)	0.0015 **M** Na_2SO_4/H_2O		T	–0.359	81Khm
Sn ($\bar{1}$10)	0.0015 **M** Na_2SO_4/H_2O		T	–0.359	81Khm
Sn (110)	0.0015 **M** Na_2SO_4/H_2O		T	0.851	81Khm
Sn/Cd alloy	0.01 **N** Na_2SO_4/H_2O	7 at%Cd	T	–0.63	65Gri
	0.01 **N** Na_2SO_4/H_2O	15 at%Cd	T	–0.61	65Gri
	0.01 **N** Na_2SO_4/H_2O	40 at%Cd	T	–0.68	65Gri

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
	0.01 **N** Na_2SO_4/H_2O	65 at%Cd	T	–0.69	65Gri
	0.01 **N** Na_2SO_4/H_2O	40 at%Cd	T	–0.68	65Gri
	0.01 **N** Na_2SO_4/H_2O	90 at%Cd	T	–0.695	65Gri
Titanium	1 **N** Na_2SO_4/H_2O		T	–1.05	67Che
(Ti)	0.1 **M** HCl/H_2O		T	–1.20	67Che
Tungsten (W)	not reported		n.a.	–0.32	53Vas
Zinc	0.01 **N** K_2SO_4/H_2O	pH=3	T	–0.63	70Vol
(Zn)	0.005 **M** $NaClO_4/H_2O$	pH=3.48	T	–0.92	69Cas
	0.01 **N** NaF/H_2O		T	–0.7	70Bar1
	0.01 **N** NaI/H_2O		T	–0.7	70Bar1
	0.01 **N** Na_2SO_4/H_2O		T	–0.7	70Bar1
	1 **N** Na_2SO_4/H_2O		H	–0.62	49Ven
	1 **N** Na_2SO_4/H_2O		H	–0.65	65Kuk, 66Kuk1
Zn (0001)	0.1 **M** KCl/H_2O		T	–0.649	60Chu
	0.1 **M** KCl/H_2O		T	–0.769	74Bat1
	0.3 **M** KCl/H_2O		SO	–0.769	75Bat
	0.1 **M** KCl/H_2O		SO	–0.759	74Fru2
	0.1 **M** KCl/H_2O		DE	–0.759	75Bat
	0.3 **M** NaF/H_2O		SO	–0.769	76Ipa1
	0.1 **M** KCl/H_2O		T	–0.779	76Ipa1
	0.1 **M** KCl/H_2O		DE	–0.759	76Ipa1
	0.005 **M** $NaClO_4/H_2O$	pH=3.48	T	–0.9	69Cas
	0.005 **M** $NaClO_4/H_2O$		T	–0.899	69Cas
	NaF/H_2O		T	–0.769	76Ipa2
	0.02 **N** $Na_2SO_4+H_2SO_4/H_2O$	pH=3.2	T	–0.6 [68])	57Kra
	0.02 **M** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=2	T	–0.61	57Khe
	0.02 **M** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=3	T	–0.61	57Khe
	0.02 **M** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=4	T	–0.61	57Khe
	0.02 **M** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=5	T	–0.61	57Khe
	0.02 **M** $(Na_2SO_4+H_2SO_4)/H_2O$	pH=6	T	–0.61	57Khe
	0.01 **M** Na_2SO_4/H_2O		T	–0.599	57Kra
Zn (0101)	0.02 **N** $Na_2SO_4+H_2SO_4/H_2O$	pH=3.2	T	–0.68 [68])	57Kra
	0.02 **N** Na_2SO_4/H_2O	pH=2.6	T	–0.6 [68])	57Kra
Zn $(10\bar{1}0)$	0.21 **M** NaF/H_2O		T	–0.919	74Bat1
	NaF/H_2O		T	–0.839	76Ipa2
	0.2 **M** NaF/H_2O		SO	–0.849	76Ipa1
	0.010 **M** Na_2SO_4/H_2O		T	–0.679	57Kra
	0.2 **M** KCl/H_2O		SO	–0.829	75Bat
	0.1 **M** KCl/H_2O		T	–0.844	76Ipa1
Zn $(11\bar{2}0)$	0.2 **M** NaF/H_2O		SO	–0.869	76Ipa1
	0.2 **M** KCl/H_2O		SO	–0.879	75Bat
	0.003 **M** KCl/H_2O		T	–0.859	75Bat

Electrode	Solution [1])	Comments [2])	Met. [3])	E_{pzc} [V vs. SHE] [4])	Ref.
Zn (11$\bar{2}$0)	0.003 **M** KCl/H_2O		T	–0.849	76Ipa1
	0.1 **M** KCl/H_2O	pH=3.7	T	–0.864	76Ipa1

[1]) A lack of an explicit concentration value indicates, that the value of E_{pzc} did not depend on the concentration.

[2]) All data refer to room temperature if not stated otherwise.

[3]) Met. = Method. AE = Acousto-electrochemical method, BAW = bulk acoustic wave method, C = Chronocoulometry, CA = Contact angle method, CO = CO charge displacement method, CV = Cyclic voltamogram, DE = Desorption potential, DR = Double layer repulsion method, EB = Film electrode bending method, EC = Electrocapillary method, ED = Electrode displacement method, EE = Theory of electron emission, ER = Electroreduction of ions, F = Friction method, FL = Flow measurement, G = Gravimetric method, H = Hardness measurements, IA = Ionic adsorption method, IM = Immersion method, ISS = Isoelectronic shift of potential, NS = Null solution method, OA = Organic adsorption method, P = Photoemission method, PC = Electrode capacitance measured with a pulse method, PE = Piezoelectric method, R = Surface stress measurement (ribbon extension method), RT = Radiotracer technique, S = Scratching or streaming electrode method, SFG = Sum frequency generation, SO = Salting out method, T = Tensammetry, TC = Modified charge determination method, VE = Vibrating electrode method, VIB = Vibrational spectroscopic method (for details and further acronyms see text). See also list of symbols.

[4]) All reported values are converted with respect to the standard hydrogen electrode for the solvent under investigation if not stated otherwise.

[5]) Electropolished surface.

[6]) Surface formed by crystal cleavage at liquid nitrogen temperature.

[7]) Electrochemically etched surface.

[8]) Chemically etched surface.

[8a]) Chemically treated surface.

[9]) Remelted solid drop electrode.

[10]) Mechanically polished and chemically etched surface.

[11]) Solid drop electrode obtained without remelting.

[12]) No significant temperature dependence was observed between 0 and 50 °C.

[13]) Electrodeposited electrode surface.

[14]) Without preelectrolysis of the electrolyte solution.

[15]) With preelectrolysis of the electrolyte solution.

[16]) This value was recorded in a separate experiment, no explanation of the reproducable difference is given.

[17]) Amalgam compositions are given as stated by the authors; conversions were made only for the organisation of the table, mol%≡at%, if not stated otherwise, values are wt%, values are listed according to amalgam composition.

[18]) Cleavage plane of high pressure stress annealed pyrolytic graphite.

[19]) Values for concentration up to 1 **m** were reported.

[20]) The reference electrode is ill defined in this report.

[21]) Electrode prepared by reduction of the metal oxide with hydrogen at 400 °C.

[22]) Copper purified by sublimation and recrystallization.

[23]) Copper purified according to Bridgeman.

[24]) Copper purified by zone melting.

[25]) Except where stated otherwise all measurements were done at $T > 30$ °C.

[26]) Solid gallium.

[27]) Gallium saturated with thallium (approx. 0.02...0.06 at%).

[28]) Gallium with thallium (0.02 at%).

[29]) Based on the experimental conditions the gold electrode is most likely covered with underpotential deposited (upd) silver. Consequently the value of E_{pzc} should be compared with the corresponding value for a silver electrode.

[30]) Measured vs. a reference electrode of Ag/Ag^+ (0.01 **M** $AgNO_3$) in methanol + 0.1 **M** Bu_4NClO_4.

[31]) Tetraethylammoniumperchlorate.

[32]) Versus a $Ag/AgNO_3$ (0.01 **M**) + 0.1 **M** tetraethylammoniumperchlorate in acetonitrile reference electrode.

[33]) The value refers to both the hexagonal and the (1x1) phase.

[34]) Electrode coated with 1-octadecanethiol.

[35]) Electrode coated with 1-undecanethiol.

[36]) Electrode coated with 1-propanethiol.

[37]) Electrode coated with 1H,1H,2H,2H-perfluorodecanethiol.

[38]) Electrolyticall deposited indium electrode.

[39]) Because of the many different electrolyte solutions employed values are grouped according to the indium concentration.

[40]) Acidified with H_2SO_4.

[41]) Single crystal. The crystallographic face was not specified.

[42]) Versus hydrogen electrode in the same solution.

[43]) An amalgamated copper surface was used.

[44]) Amalgamated gold electrode.

[45]) Solution saturated with N_2O.

[46]) *E* vs. a pool of mercury in the same solution.

[47]) *E* vs. a reference electrode of mercury in methanol + 0.01 **M** LiCl.

[48]) Measured vs. a saturated calomel electrode filled with methanol.

[49]) On an electrode not saturated with hydrogen.

[50]) On an electrode saturated with hydrogen.

[51]) Palladium overlayer on a Pt(111) electrode surface.

[52]) The value refers to palladium islands on a Au(111) electrode.

[53]) One monolayer of palladium deposited on a Au(111) electrode.

[54]) Two monolayers of palladium deposited on a Au(111) electrode.

[55]) Five monolayers of palladium deposited on a Au(111) electrode.

[56]) If a coverage of the platinum with hydrogen was specified this is indicated by (H), if the electrode was platinized this is indicated by (Pt).

[57]) Value measured after prolonged cathodic polarisation.

[58]) Values vary as a function of step density on Pt[*n*(111)x(111)] surfaces.

[59]) The value is given with respect to the potential of a hydrogen electrode in this electrolyte solution.

[60]) A rhodium black electrode was used.

[61]) Polycrystalline silver layer electrodeposited on a polycrystalline platinum electrode.

[62]) Mechanical surface treatment of the electrode.

[63]) Silver monolayer deposited on foreign metal single crystal.

[64]) Silver purified by sublimation and recrystallization.

[65]) Silver layer epitaxially electrodeposited on a Pt(111) electrode.

[66]) Incomplete upd-silver layer electrodeposited on a Pt(111) electrode.

[67]) Because of the many different electrolyte solutions employed values are grouped according to the thallium concentration.

[68]) Zinc purified by sublimation and recrystallization.

References

91Pas Paschen, F.: Ann. Phys. (Leipzig) **43** (1891) 568.
98Pal Palmaer, W.: Z. Phys. Chem. **25** (1898) 257.
03Gou Gouy, G.: Ann. Chim. Phys. **29** (1903) 145.
26Ben Bennewitz, K.: Z. Phys. Chem. **124** (1926) 115.
28Fru1 Frumkin, A., Gorodetzkaja, A.: Z. Phys. Chem. **136** (1928) 215.
28Fru2 Frumkin, A., Gorodetzkaja, A.: Z. Phys. Chem. **136** (1928) 451.
30Fru Frumkin, A., Cirves, F.J.: J. Phys. Chem. **34** (1930) 74.
31Pro Proskurnin, M., Frumkin, A.: Z. Phys. Chem. Abt. A **155** (1931) 29.
32Pro Proskurnin, M.A.: Zh. Fiz. Khim. **3** (1932) 91.
33Fru Frumkin, A.: Phys. Z. Sowjetunion **4** (1933) 239.
34Gor Gorodetzkaja, A., Kabanov, B.: Phys. Z. Sowjetunion **5** (1934) 418.
35Fru Frumkin, A.: Actual. Sci. Ind. (1935) 373.
36Erd Erdey-Grúz, T., Szarvas, P.: Z. Phys. Chem. Abt. A **177** (1936) 277.
36Mur Murtasaev, A., Gorodetskaya, A.V.: Acta Physicochim. URSS **4** (1936) 75.
36Shl Shlygin, A., Frumkin, A., Medvedovskii, V.: Acta Physicochim. URSS **4** (1936) 911.
38Bal Balashova, N., Frumkin, A.: C. R. Acad. Sci. URSS **20** (1938) 449.
38Win Winkel, A., Siebert, H.: Z. Elektrochem. **44** (1938) 127.
39Luk Lukovtsev, P., Levina, S., Frumkin, A.N.: Acta Physicochim. URSS **11** (1939) 21.
39Ves Veselovsky, V.I.: Acta Physicochim. URSS **11** (1939) 815.
40Kuc Kuchinskii, E., Burshtein, R., Frumkin, A.: Zh. Fiz. Khim. **14** (1940) 441.
40Leg Legran, A., Levina, S.: Acta Physicochim. URSS **12 (**1940) 243.
43Chu Chugunov, P., Bruns, B.: Acta Physicochim. URSS **18** (1943) 351.
44Reb Rebinder, P.A., Venstrem, E.K.: Acta Physicochim. URSS **19** (1944) 36.
45Reb Rebinder, P.A., Venstrem, E.K.: Zh. Fiz. Khim. **19** (1945) 1.
47Gra Grahame, D.C.: Chem. Rev. **41** (1947) 441.
48Bor Borisova, T.I., Ershler, B., Frumkin, A.: Zh. Fiz. Khim. **22** (1948) 925.
49Gra Grahame, D., Larsen, R.P., Poth, M.A.: J. Am. Chem. Soc. **71** (1949) 2978.
49Ven Venstrem, E.K., Rebinder, P.A.: Dokl. Akad. Nauk SSSR **68** (1949) 329.
50Bor Borisova, T.I., Ershler, B.V.: Zh. Fiz. Khim. **24** (1950) 337.
50Bow Bowden, F.P., Young, L.: Research (London) **3** (1950) 235.
51Pfü Pfützenreuther, A., Masing, G.: Z. Metallkd. **42** (1951) 361.
52Gra Grahame, D.C., Coffin, E.M., Cummings, J.I., Poth, M.A.: J. Am. Chem. Soc. **74** (1952) 1207.
52Hil Hillson, P.J.: Trans. Faraday Soc. **48** (1952) 462.
52Nik Nikolaeva, N.V., Shapiro, N.S., Frumkin, A.N.: Dokl. Akad. Nauk SSSR **36** (1952) 581.
52Ven Venstrem, E.K., Rebinder, P.A.: Zh. Fiz. Khim. **26** (1952) 1847.
53Boc Bockris, J.O'M., Jones, R.P.: Nature (London) **171** (1953) 930.
53Bor Borisova, T.I., Veselovskii, V.J., Karpov, L.Ya.: Zh. Fiz. Khim. **27** (1953) 1195.
53Khe Kheifets, V.L., Reishakhrit, L.S.: Uch. Zap. Leningr. Gos. Univ. im. A. A. Zhdanova Ser. Khim. Nauk. **13** (1953) 173.
53Lev1 Levin, A.I., Ukshe, E.A., Brilina, N.S.: Dokl. Akad. Nauk SSSR **88** (1953) 697.
53Rob Robertson, W.D.: J. Electrochem. Soc. **100** (1953) 194.
53Vas Vasenin, R.M.: Zh. Fiz. Khim. **27** (1953) 878.
54Gra Grahame, D.: J. Am. Chem. Soc. **76** (1954) 4819.
55Aya Ayazyan, E.O.: Dokl. Akad. Nauk SSSR **100** (1955) 473.
55Fru Frumkin, A.: Z. Elektrochem. **59** (1955) 807.
55Uks1 Ukshe, E.A., Levin, A.I.: Dokl. Akad. Nauk SSSR **105** (1955) 119.
55Uks2 Ukshe, E.A., Levin, A.I.: Zh. Fiz. Khim. **29** (1955) 219.
56Ran Randles, J.E.B., Whiteley, K.S.: Trans. Faraday Soc. **52** (1956) 1509.

56Ven Venstrem, E.K., Lichtman, V.I., Rebinder, P.A.: Dokl. Akad. Nauk SSSR **107** (1956) 105.
57Fle Fleischmann, M., Sowerby, J., Thirsk, H.R.: Trans. Faraday Soc. **53** (1957) 91.
57Khe Kheifets, V.L., Krasikov, B.S.: Zh. Fiz. Khim. **31** (1957) 1992.
57Kra Krasikov, B., Sysoeva, V.: Dokl. Akad. Nauk SSSR **114** (1957) 826.
57Smi Smirnov, V.A., Antropov, L.I: Dokl. Akad. Nauk SSSR **113** (1957) 1098.
58Ant Antropov, L.I., Banerjee, S.N.: J. Indian Chem. Soc. **35** (1958) 531.
58Dam Damaskin, B.B., Nikolaeva-Fedorovic, N.V., Frumkin, A.N.: Dokl. Akad. Nauk SSSR **121** (1958) 129.
59Efi Efimov, E.A., Erusalimchik, I.G.: Zh. Fiz. Khim. **33** (1959) 441.
59McM McMullen, J.J., Hackerman, N.: J. Electrochem. Soc. **106** (1959) 341.
59Vor Voropaeva, T.N., Deryagin, B.V., Kabanov, B.N.: Dokl. Akad. Nauk SSSR **128** (1959) 981.
60Bog Boguslavskii, L.I., Damskin, B.B.: Zh. Fiz. Khim. **34** (1960) 2099.
60Chu Chuan-Sin, T., Iofa, Z.A.: Dokl. Akad. Nauk SSSR **131** (1960) 137.
60Del Delahay, P., Kleinerman, M.: J. Am. Chem. Soc. **82** (1960) 4509.
60Lei Leikis, D.I.: Dokl. Akad. Nauk SSSR **135** (1960) 1429.
61Bal Balashova, N.A., Merkulova, N.: Soviet Electrochemistry (Proceedings of the Fourth Conference on Electrochemistry), Vol. 1, New York, 1961, p. 23.
61Sta Staicopoulos, D.N.: J. Electrochem. Soc. **108** (1961) 900.
62Jak1 Jakuszewski, B., Kozlowski, Z.: Rocz. Chem. **36** (1962) 1873.
62Jen Jenkins, D., Newcombe, R.: Electrochim. Acta **6** (1962) 685.
62Sch1 Schmid, G.M., Hackerman, N.: J. Electrochem. Soc. **109** (1962) 243.
62Vor Voropaeva, T.N., Deryagin, B.V., Kabanov, B.N.: Kolloidn. Zh. **24** (1962) 396.
63Bir Birintseva, T.P., Kabanov, B.N.: Zh. Fiz. Khim. **37** (1963) 2600.
63Gre Green, M., Dahms, H.: J. Electrochem. Soc. **110** (1963) 466.
63Sch1 Schmid, G.M., Hackerman, N.: J. Electrochem. Soc. **110** (1963) 440.
63Vor Voropaeva, T.N., Deryagin, B.V., Kabanov, B.N.: Izv. Akad. Nauk SSSR, Otd. Khim. Nauk (1963) 257.
63Wu Wu, H., Lin, C.: Hua Hsueh Hsueh Pao **29** (1963) 95.
64And Andersen, T.N., Perkins, R.S., Eyring, H.: J. Am. Chem. Soc. **86** (1964) 4496.
64But1 Butler, J.N., Makrides, A.C.: Trans. Faraday Soc. **60** (1964) 1664.
64Fru1 Frumkin, A.N., Polianovskaya, N., Grigor'ev, N.B.: Dokl. Akad. Nauk SSSR **157** (1964) 1455.
64Fru2 Frumkin, A.N., Grigor'ev, N.B., Bagotskaya, I.A.: Dokl. Akad. Nauk SSSR **157** (1964) 957.
64Jak1 Jakuszewski, B., Kozlowski, Z.: Lodz. Tow. Nauk. Wydz. 3, Acta Chim. **9** (1964) 25.
64Jak2 Jakuszewski, B., Kozlowski, Z.: Rocz. Chem. **38** (1964) 93.
64Kaz Kazarinov, V.E., Balashova, N.A.: Dokl. Akad. Nauk SSSR **157** (1964) 1174.
64Kra Krasikov, B.S.: Zh. Prikl. Khim. **37** (1964) 2420.
64Per Perkins, R.S., Eyring, H.: J. Am. Chem. Soc. **86** (1964) 4496.
65Bal1 Balashova, N.A., Kazarinov, V.E.: Elektrokhimiya **1** (1965) 512.
65Boc Bockris, J.O'M., Müller, K., Wroblowa, H., Kovac, Z.: J. Electroanal. Chem. **10** (1965) 416.
65Bon Bonnemay, M., Bronoël, G., Jonville, P., Levart, E.: C. R. Acad. Sci. (Paris) **260** (1965) 5262.
65But1 Butler, J.N., Meehan, M.L.: J. Phys. Chem. **69** (1965) 4051.
65But6 Butler, J.N., Meehan, M.L., Makrides, A.C.: J. Electroanal. Chem. **9** (1965) 237.
65Fru1 Frumkin, A., Polianovskaya, N., Grigoriev, N.B., Bagotskaya, I.: Electrochim. Acta **10** (1965) 793.
65Fru2 Frumkin, A.: Sven. Kem. Tidskr. **77** (1965) 300.
65Gri Grilikhes, M.S., Krasikov, B.S.: Vestn. Leningr. Univ. Ser. 4, Fiz. Khim. **4** (1965) 131.
65Jak1 Jakuszewski, B., Kozlowski, Z.: Soc. Sci. Lodz. Acta Chim. **10** (1965) 5.
65Koz Kozhevnikova, N.M., Rotinyan, A.L.: Elektrokhimiya **1** (1965) 664.
65Kuk Kukoz, F.I., Semenchenko, S.A.: Elektrokhimiya **1** (1965) 1454.
65Per Perkins, R.S., Livingston, R.C., Andersen, T.N. , Eyring, H.: J. Phys. Chem. **69** (1965) 3329.

65Pol Polyanovskaya, N.S., Frumkin, A.N.: Elektrokhimiya **1** (1965) 538.
65Sat Sathyanarayana, S.: J. Electroanal. Chem. **10** (1965) 119.
66But1 Butler, J.N.: J. Phys. Chem. **70** (1966) 2312.
66But2 Butler, J.N., Meehan, M.L.: J. Phys. Chem. **70** (1966) 3582.
66Dag Dagaeva, I.G., Leikis, D.I., Sevastyanov, E.S.: Elektrokhimiya **2** (1966) 820.
66Fru1 Frumkin, A., Grigoryev, N., Bagotskaya, I.: Elektrokhimiya **2** (1966) 329.
66Gil Gileadi, E., Argade, S.D., Bockris, J.O'M.: J. Phys. Chem. **70** (1966) 2044.
66Gok Gokhstein, A.Y.: Elektrokhimiya **2** (1966) 1318.
66Kuk1 Kukoz, F.I., Semenchenko, S.A.: Elektrokhimiya **2** (1966) 74.
66Kuk2 Kukoz, F.I., Semenchenko, S.A., Bogdanov, V.I.: Issled. Obl. Khim. Istochnikov Toka (1966) 201.
66Lei Leikis, D.I., Sevastyanov, E.S., Dagaeva, I.G.: J. Electrochem. Soc. **113** (1966) 1341.
66Man Mansorov, G.N., Kazarinov, V.E., Balashova, N.A.: Elektrokhimiya **2** (1966) 1438.
66Mor Morozov, A., Grigoryev, N., Bagotskaya, I.: Elektrokhimiya **2** (1966) 1235.
66Pal Palm, U., Past, V., Pullerits, R.: Elektrokhimiya **2** (1966) 604.
67Arg Argade, S.G., Gileadi, E., in: Electrosorption, Gileadi, E. (ed.), New York: Plenum Press, 1967, p. 87.
67Bal Balashova, N.A., Kossaya, A.M., Gorokhova, N.T.: Elektrokhimiya **3** (1967) 656.
67Bod Bodé, D.D., Andersen, T.N., Eyring, H.: J. Phys. Chem. **71** (1967) 792.
67Che Chervyakov, P.I., Parfenov, G.S.: Uch. Zap. Omsk. Gos. Pedagog. Inst. **26** (1967) 58.
67Dag Dagaeva, I.G., Leikis, D.I., Sevastyanov, E.S.: Elektrokhimiya **3** (1967) 891.
67Gar Garnish, J.D., Parsons, R.: Trans. Faraday Soc. **63** (1967) 1754.
67Gri2 Grilikhes, M.S., Krasikov, B.S.: Zh. Prikl. Khim. **40** (1967) 921.
67Ham1 Hampson, N.A., Larkin, D.: J. Electrochem. Soc. **114** (1967) 933.
67Ham2 Hampson, N.A., Larkin, D., Morley, J.R.: J. Electrochem. Soc. **114** (1967) 817.
67Mur Murakawa, T., Kato, T., Nagaura, S., Hackerman, N.: Corros. Sci. **7** (1967) 657.
67Pay Payne, R.: J. Phys. Chem. **71** (1967) 1548.
67Per Perkins, R.S.: Ph.D. Thesis, University of Utah, 1967.
67Pet1 Petit, M., Clavilier, J.: C. R. Acad. Sci. (Paris), Ser. C **265** (1967) 145.
67Pet2 Petrii, O.A., Kossaya-Tsybulevskay, A.M., Tyurin, Yu.M.: Elektrokhimiya **3** (1967) 617.
67Pol Polianovskaya, N., Frumkin, A.: Elektrokhimiya **3** (1967) 1129.
67Ryb Rybalka, L.E., Leikis, D.I.: Elektrokhimiya **3** (1967) 383.
67Vet Vetter, K.J.: Electrochemical Kinetics, New York: Academic Press, 1967, p. 73.
68Bag Bagotskaya, I., Morozov, A., Grigoryev, N.: Electrochim. Acta **13** (1968) 873.
68Bal Balashova, N.A., Gorokhova, N.T., Kulezneva, M.I.: Elektrokhimiya **4** (1968) 871.
68Bar Bartenev, V.J., Sevastjanov, E.S., Leikis, D.I.: Elektrokhimiya **4** (1968) 745.
68Cla1 Clavilier, J., van Huong, N.: C. R. Acad. Sci. (Paris), Ser. C **267** (1968) 207.
68Fru Frumkin, A., Petry, O., Kossaya, A., Entina, V., Topolev, V.: J. Electroanal. Chem. **16** (1968) 175.
68Ham1 Hampson, N.A., Larkin, D.: J. Electrochem. Soc. **115** (1968) 612.
68Ham2 Hampson, N.A., Larkin, D.: J. Electroanal. Chem. **18** (1968) 401.
68Jak1 Jakuszewski, B., Kozlowski, Z., Granosik, W.: Soc. Sci. Lodz. Acta Chim. **13** (1968) 15.
68Koz Kozlowski, Z., Nast, A.M.: Lodz. Tow. Nauk. Wydz. 3, Acta Chim. **13** (1968) 21.
68Log Loge, A., Lorenz, W.: Z. Phys. Chem. (Leipzig) **238** (1968) 299.
68Mor Morcos, I., Fisher, H.: J. Electroanal. Chem. **17** (1968) 7.
68Pet Petrii, O.A., Kotlov, Yu.G.: Elektrokhimiya **4** (1968) 774.
68Pol Polianovskaya, N.: Elektrokhimiya **4** (1968) 549.
69And Andersen, T.N., Anderson, J.L., Eyring, H.: J. Phys. Chem. **73** (1969) 3562.
69Bar1 Bartenev, V.Ya., Sevastyanov, E.S., Leikis, D.I.: Elektrokhimiya **5** (1969) 1502.
69Boc1 Bockris, J.O'M., Argade, S.E., Gileadi, E.: Electrochim. Acta **14** (1969) 1259.
69Bud Budevski, E., Vitanov, T., Sevastyanov, E.S., Popov, A.: Elektrokhimiya **5** (1969) 90.
69Bur Burshstein, R.K., Pshenichnikov, A.G.: Elektrokhimiya **5** (1969) 332.

69Cas Caswell, P., Hampson, N.A., Larkin, D.: J. Electroanal. Chem. **20** (1969) 335.
69Cla Clavilier, J., van Huong, N.: C. R. Acad. Sci. (Paris), Ser. C **269** (1969) 736.
69Dek Dekker, B.G., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **21** (1969) 137.
69Hag Haga, M., Past, V.: Trans. Tartu State Univ. **235** (1969) 47.
69Ham1 Hamelin, A., Valette, G.: C. R. Acad. Sci. (Paris), Ser. C **269** (1969) 1020.
69Ham2 Hamelin, A., Sotto, M., Valette, G.: C. R. Acad. Sci. (Paris), Ser. C **268** (1969) 213.
69Kha1 Khaga, M., Past, V.: Elektrokhimiya **5** (1969) 618.
69Kha2 Khaga, M., Past, V.: Uch. Zap. Tartu. Gos. Univ. **235** (1969) 47.
69Lik Likhtman, V., Kochanova, L., Leikis, D., Shchukin, E.: Elektrokhimiya **5** (1969) 729.
69Mor Morcos, I.: J. Electroanal. Chem. **20** (1969) 479.
69Pal Palts, K., Palm, U., Past, V., Pullerits, R.: Uch. Zap. Tartu. Gos. Univ. **235** (1969) 57.
69Pet Petrii, O.A., Frumkin, A.N., Kotlov, Yu.G.: Elektrokhimiya **5** (1969) 476.
69Ple Pleskov, Yu.V., Rotenberg, Z.A.: J. Electroanal. Chem. **20** (1969) 1.
70And Andrusev, M.M., Ershler, A.B., Tedoradze, G.A.: Elektrokhimiya **6** (1970) 1163.
70Bar1 Bartenev, V.J., Sevastjanov, E.S., Leikis, D.I.: Elektrokhimiya **6** (1970) 1197.
70Bar2 Bartenev, V.J., Sevastjanov, E.S., Leikis, D.I.: Elektrokhimiya **6** (1970) 1868.
70Ego1 Egorov, L., Novosel'sky, I.M.: Elektrokhimiya **6** (1970) 521.
70Ego2 Egorov, L.Y., Novoselskij, I.M.: Elektrokhimiya **6** (1970) 869.
70Pet Petrii, O.A., Frumkin, A.N., Kotlov, Yu.G.: Elektrokhimiya **6** (1970) 725.
70Rot2 Rotinyan, A.L., Levin, E.D.: Elektrokhimiya **6** (1970) 328.
70Tim Timmer, B., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **24** (1970) 287.
70Vol Volkov, L.V., Ponomarjev, A.F., Yurev, B.P.: Tr. Leningr. Politekh. Inst. im. M. I. Kalinina **304** (1970) 94.
71Car Carr, J.P., Hampson, N.A., Holley, S.N., Taylor, R.: J. Electroanal. Chem. **32** (1971) 345.
71Elk Elkin, V.V., Alekseyev, V.N., Knots, L.L., Leikis, D.I.: Dokl. Akad. Nauk SSSR **199** (1971) 638.
71Erl Erlikh, G., Kukk, Yu., Past, V.: Uch. Zap. Tartu. Gos. Univ. **289** (1971) 9.
71Fre Fredlein, R.A., Damjanovic, A., Bockris, J.O'M.: Surf. Sci. **25** (1971) 261.
71Fru Frumkin, A., Polianovskaya, N., Bagotskaya, I., Grigoryev, N.: J. Electroanal. Chem. **33** (1971) 319.
71Ham1 Hamelin, A., Lecoeur, J.: Collect. Czech. Chem. Commun. **36** (1971) 714.
71Jak Jakuszewsi, B., Kozlowski, Z., Partyka, S., Przasnyski, M., Romanovski, S.: Elektrokhimiya **7** (1971) 804.
71Mor Morcos, I.: J. Colloid Interface Sci. **37** (1971) 410.
71Pet Petyarv, E., Kolk, K., Palm, U.: Tartu Riikliku Ulik. Toim. **289** (1971) 22.
71Shi Shimokawa, M., Takamura, T.: J. Electroanal. Chem. **32** (1971) 314.
71Val1 Valette, G., Hamelin, A.: C. R. Acad. Sci. (Paris), Ser. C **272** (1971) 602.
71Val2 Valette, G.: C. R. Acad. Sci. (Paris), Ser. C **273** (1971) 320.
72Bro Brodsky, A., Pleskov, Yu.: Progress in Surface Science 2, Davidson, S. (ed.), Oxford: Pergamon, 1972, p. 2.
72Car Carr, J.P., Hampson, N.A.: J. Electrochem. Soc. **119** (1972) 325.
72Gri1 Grigor'ev, N.B., Gedvillo, I., Bardina, N.: Elektrokhimiya **8** (1972) N3.
72Gri3 Grigor'ev, N.B., Gedvillo, I., Bardina, N.: Elektrokhimiya **8** (1972) 409.
72Gri4 Grigor'ev, N.B., Fateev, S.L., Bagotskaya, I.A.: Elektrokhimiya **8** (1972) 311.
72Gri5 Grigor'ev, N.B., Fateev, S.L., Bagotskaya, I.A.: Elektrokhimiya **8** (1972) 1525.
72Ham Hampson, N.A., Latham, R.J.: J. Electroanal. Chem. **34** (1972) 247.
72Lei Leikis, E.D., Panin, V., Rybalka, K.: J. Electroanal. Chem. **40** (1972) 6.
72Lev Levin, E.D., Rotinyan, A.L.: Elektrokhimiya **8** (1972) 242.
72Nov1 Novoselskiy, I.M., Konevskih, N.I., Egorov, L.Ya.: Double Layer and Adsorption on Solid Electrodes, Vol. III, Tartu University, 1972, p. 195.
72Nov2 Novoselskii, I.M., Konevskikh, N.I., Egorov, L.Ya.: Elektrokhimiya **8** (1972) 1480.
72Pan1 Panin, V.A., Rybalka, K.V., Leikis, D.I.: Elektrokhimiya **8** (1972) 390.

72Pan3 Panin, V.A., Rybalka, K.V., Leikis, D.I.: Elektrokhimiya **8** (1972) 1507.
72Pet1 Petyarv, E.K., Kolk, K.A., Palm, U.V.: Elektrokhimiya **8** (1972) 100.
72Pod Podlovchenko, B.I., Epstein, N.A.: Elektrokhimiya **8** (1972) 1522.
72Sev Sevastyanov, E.S., Vitanov, T., Popov, A.: Elektrokhimiya **8** (1972) 412.
72Val Valette, G.: C. R. Acad. Sci. (Paris), Ser. C **274** (1972) 2046.
73Bau Baugh, L.M., Parsons, R.: J. Electroanal. Chem. **41** (1973) 311.
73Cla Clavilier, J., van Huong, C.N.: J. Electroanal. Chem. **41** (1973) 193.
73Gri Grigor'ev, N.B., Kuprin, V.P., Loshkarev, Yu.M.: Elektrokhimiya **9** (1973) 1842.
73Iva Ivanov, V.F., Damaskin, B.B., Segel'man, B.S., Melekhova, N.I.: Elektrokhimiya **9** (1973) 389.
73Kim Kim, S.H.: J. Phys. Chem. **77** (1973) 2787.
73Lou Loutfy, R.O.: Electrochim. Acta **18** (1973) 227.
73Non Noninski, K.I., Lazarova, E.M.: Elektrokhimiya **9** (1973) 673.
73Nov Novoselskiy, I.M., Maksimyuk, N.I., Egorov, L.Ya.: Elektrokhimiya **9** (1973) 1518.
73Ryb3 Rybalka, K.V., Panin, V.A.: Elektrokhimiya **9** (1973) 172.
73Val Valette, G., Hamelin, A.: J. Electroanal. Chem. **45** (1973) 301.
74Bat1 Batrakov, V.V., Damaskin, B.B., Ipatov, Y.P.: Elektrokhimiya **10** (1974) 144.
74Bro1 Brodsky, A., Gurevich, V.Ya., Pleskov, Yu., Rotenberg, Z.: Sovremennaya Foto-elektrokhimiya, Nauka, Moscow, 1974.
74Cla1 Clavilier, J., van Huong, N.: C. R. Acad. Sci. (Paris), Ser. C **279** (1974) 993.
74Cla2 Clark, G.J., Andersen, T.N., Valentine, R.S., Eyring, H.: J. Electrochem. Soc. **121** (1974) 618.
74Fre Fredlein, R.A., Bockris, J.O'M.: Surf. Sci. **46** (1974) 641.
74Fru1 Frumkin, A.N., Pyarnoya, M.P., Grigoriev, N.B., Pal'm, U.V.: Elektrokhimiya **10** (1974) 1130.
74Fru2 Frumkin, A.N., Damaskin, B.B., Grigoryev, N.B., Bagotskaya, I.A.: Electrochim. Acta **19** (1974) 69.
74Fru3 Frumkin, A.N., Grigoryev, N.B., Pyarnoya, M., Palm, U.: Elektrokhimiya **10** (1974) N5.
74Ham Hamelin, A., Bellier, J.P.: C. R. Acad. Sci. (Paris), Ser. C **279** (1974) 371.
74Lec Lecoeur, J., Hamelin, A.: C. R. Acad. Sci. (Paris), Ser. C **279** (1974) 1081.
74Pod Podlovchenko, B.I., Epstein, N.A., Frumkin, A.N.: J. Electroanal. Chem. **53** (1974) 95.
74Rot Rotenberg, Z.A., Prishchepa, Yu.A., Pleskov, Yu.V.: J. Electroanal. Chem. **56** (1974) 345.
74Vit1 Vitanov, T., Popov, A., Sevastyanov, E.S.: Elektrokhimiya **10** (1974) 346.
74Zel Zelinsky, A.G., Bek, R.Yu.: Elektrokhimiya **10** (1974) 1212.
75Ale1 Aleksandrova, D.P., Sevastyanov, E.S., Andrusev, M.M., Leikis, D.I.: Elektrokhimiya **11** (1975) 648.
75Bag Bagotskaya, I.A., Grigor'ev, N.B., Doubova, L., Kalyuzknaya, A., Fateev, S.: Double Layer and Adsorption on Solid Electrodes, Vol. IV, Tartu University, 1975, p. 14.
75Bat Batrakov, V.V., Damaskin, B.B.: J. Electroanal. Chem. **65** (1975) 361.
75Gri Grigor'ev, N.B., Bulavka, V.A., Loshkarev, Yu.M.: Elektrokhimiya **11** (1975) 1404.
75Huo van Huong, N., Clavilier, J., Bonnemay, M.: J. Electroanal. Chem. **65** (1975) 531.
75Lec1 Lecoeur, J., Sella, C., Tertian, L., Hamelin, A.: C. R. Acad. Sci. (Paris), Ser. C **280** (1975) 247.
75Lec2 Lecoeur, J., Sella, C., Martin, K., Tertian, L., Deschamps, J.: C. R. Acad. Sci. (Paris), Ser. C **281** (1975) 71.
75Max Maximuk, N.I., Egorov, L.Ya.: Double Layer and Adsorption on Solid Electrodes, Vol. IV, Tartu University, 1975, p. 177.
75Mis Mishuk, V.Ya., Solomatin, E.A., Elkin, V.V., Knots, L.L.: Elektrokhimiya **11** (1975) 1897.
75Non Noninski, K.I., Lazarova, E.M.: Elektrokhimiya **11** (1975) 1103.
75Pal Palm, U.V., Vaartnou, M.G., Petyarv, E.K.: Elektrokhimiya **11** (1975) 1849.
75Pay Payne, R.: J. Electroanal. Chem. **60** (1975) 183.
75Pul3 Pullerits, R., Moldau, M., Past V.: Double Layer and Adsorption on Solid Electrodes, Vol. IV, Tartu University, 1975, p. 257.
75Ryb2 Rybalka, L.E., Leikis, D.I.: Elektrokhimiya **11** (1975) 1619.
75Vit1 Vitanov, T., Popov, A.: Trans. SAEST **10** (1975) 5.

75Vit2 Vitanov, T., Popov, A., Sevastyanov, E.S.: Elektrokhimiya **11** (1975) 170.
76Ham Hamelin, A., Lecoeur, J.: Surf. Sci. **57** (1976) 771.
76Ipa1 Ipatov, Y.P., Batrakov, V.V.: Elektrokhimiya **12** (1976) 1174.
76Ipa2 Ipatov, Y.P., Batrakov, V.V., Shalaginov, V.V.: Elektrokhimiya **12** (1976) 286.
76Lin Lin, K.-F., Beck, T.R.: J. Electrochem. Soc. **123** (1976) 1145.
76Pet Petyarv, V., Palm, U.W.: Elektrokhimiya **12** (1976) 806.
76Ryb Rybalka, L.E., Leikis, D.I., Zelinski, A.G.: Elektrokhimiya **12** (1976) 1340.
76Ush Ushakova, Z.N., Ivanov, V.F.: Elektrokhimiya **12** (1976) 485.
76Vit1 Vitanov, T., Popov, A.: Dokl. Akad. Nauk SSSR **226** (1976) 373.
76Vit2 Vitanov, T., Popov, A., Sevastyanov, E.S.: Elektrokhimiya **12** (1976) 582.
77Bor Borkowska, Z.: J. Electroanal. Chem. **79** (1977) 206.
77Cla Clavilier, J., van Huong, N.: J. Electroanal. Chem. **80** (1977) 101.
77Pal Pal'm, U.V., Pyarnoya, M.P., Sal've, M.A.: Elektrokhimiya **13** (1977) 873.
77Vit Vitanov, T., Popov, A.: 28th Meeting of ISE (International Society of Electrochemistry), Varna; Ext. Abstract, 1977, p. 230.
78Bag Bagotskaya, I.A., Dubova, L.M.: Elektrokhimiya **14** (1978) 1383.
78Bul Bulavka, V.A., Grigor'ev, N.B., Loshkarev, Yu.M.: Elektrokhimiya **14** (1978) 1037.
78Kme Kmelevaya, L.P., Chizhov, A.V., Damaskin, B.B.: Elektrokhimiya **14** (1978) 1304.
78Laz Lazarova, E.M.: Elektrokhimiya **14** (1978) 1300.
78Lec Lecoeur, J., Sella, C., Martin, J.C.: C. R. Acad. Sci. (Paris), Ser. C **287** (1978) 447.
78Max Maximuk, N.I., Egorov, L.Ya.: Double Layer and Adsorption on Solid Electrodes, Vol. V, Tartu University, 1978, p. 148.
78Mis Mishuk, V.Ya., Solomatin, E.A., Elkin, V.V.: Elektrokhimiya **14** (1978) 1135.
78Nov Novitskii, S.P., Burenkov, I.I., Kenzin, V.I., Zelinsky, A.G., Leikis, D.I.: Elektrokhimiya **14** (1978) 270.
78Shl Shlepakov, A.V., Sevastyanov, E.S.: Elektrokhimiya **14** (1978) 287.
78Vaa Vaartnou, M., Palm, U.: Elektrokhimiya **14** (1978) 311.
78Vit Vitanov, T., Popov, A.: 29th Meeting of ISE (International Society of Electrochemistry), Budapest; Ext. Abstracts, Part I, 1978, p. 539.
78Zel1 Zelinsky, A.G., Bek, R.Yu.: Elektrokhimiya **14** (1978) 1825.
78Zel2 Zelinsky, A.G., Bek, R.Yu., Makurin, A.L., Abdulov, S.D.: Elektrokhimiya **14** (1978) 1740.
79Ama Amadelli, R., Daghetti, A., Vergano, L., De Battisti, A., Trasatti, S.: J. Electroanal. Chem. **100** (1979) 379.
79Bat Batrakov, B.B., Naumova, N.I.: Elektrokhimiya **15** (1979) 551.
79Hen Hennig, H., Batrakov, V.V.: Elektrokhimiya **15** (1979) 1833.
79Vaa1 Vaartnou, M.G., Palm, U.V.: Elektrokhimiya **15** (1979) 591.
80Brz Brzostowska-Smolska, M., Minc, S.: Pol. J. Chem. **54** (1980) 2391.
80Kar Karbasov, B.G., Tikhonov, K.I., Vyacheslavov, P.M.: Elektrokhimiya **16** (1980) 1273.
80Kof Kofman, R., Garrigos, R., Cheyssac, P.: Surf. Sci. **101** (1980) 231.
80Pan Pankina, G.W., Leikis, D.I., Sevastyanov, E.S.: Elektrokhimiya **16** (1980) 213.
80Pet Petrii, O.A., Khomchenko, I.G.: J. Electroanal. Chem. **106** (1980) 277.
80Vae Vaertnou, M., Palm, U.: Elektrokhimiya **16** (1980) 1877.
81Ann Anni, K.L., Erlich, Y.O., Pal'm, U.V.: Double Layer and Adsorption on Solid Electrodes, Vol. VI, Tartu University, 1981, p. 5.
81Bor Borkowska, Z., Fawcett, W.R.: Can. J. Chem. **59** (1981) 710.
81Fle Fleischmann, M., Robinson, J., Waser, R.: J. Electroanal. Chem. **117** (1981) 257.
81Jur Jurkiewicz-Herbich, M.: J. Electroanal. Chem. **119** (1981) 275.
81Khm Khmelevaya, L.P., Damaskin, B.B.: Elektrokhimiya **17** (1981) 1721.
81Val Valette, G.: J. Electroanal. Chem. **122** (1981) 285.
82Bor Borkowska, Z., Fawcett, W.R.: Can. J. Chem. **60** (1982) 1787.
82Ham1 Hamelin, A.: J. Electroanal. Chem. **138** (1982) 395.
82Ham2 Hamelin, A., Le Lan, A.: C. R. Acad. Sci. (Paris), Ser. 2 **295** (1982) 161.

82Laz Lazarova, E.M.: Elektrokhimiya **18** (1982) 1654.
82Lec Lecoeur, J., Andro, J., Parsons, R.: Surf. Sci. **114** (1982) 320.
82Val Valette, G.: J. Electroanal. Chem. **138** (1982) 37.
82Vit Vitanov, T., Popov, A., Sevastyanov, E.S.: J. Electroanal. Chem. **142** (1982) 289.
83Brz Brzostowska-Smolska, M., Minc, S.: Pol. J. Chem. **57** (1983) 1005.
83Hup Hupp, J.T., Larkin, D., Weaver, M.J.: Surf. Sci. **125** (1983) 429.
83Mly Mlynarczyk, R., Figaszewski, Z., Koczorowski, Z.: Pol. J. Chem. **57** (1983) 1011.
83Ohm Ohmori, T.: J. Electroanal. Chem. **157** (1983) 159.
84Ale Alekseeva, E.Yu., Safonov, V.A., Petrii, O.A.: Elektrokhimiya **20** (1984) 945.
84Bor Borkowska, Z., Stafiej, J.: J. Electroanal. Chem. **170** (1984) 289.
84Dag Daghetti, A., Trasatti, S.: J. Electroanal. Chem. **162** (1984) 327.
85Kol Kolb, D.M., Schneider, J.: Surf. Sci. **162** (1985) 764.
85Lec Lecoeur, J., Bellier, J.P.: Electrochim. Acta **30** (1985) 1027.
85Lus1 Lust, E.I., Palm, Yu.V.: Elektrokhimiya **21** (1985) 1256.
85Lus2 Lust, E.I., Palm, Yu.V.: Elektrokhimiya **21** (1985) 1381.
85Ngu Nguyen Van Huong, C.: J. Electroanal. Chem. **194** (1985) 131.
85Pez Pezzatini, G., Moncelli, M.R., Foresti, M.L., Pergola, F., Guidelli, R.: J. Electroanal. Chem. **196** (1985) 429.
86Ann Anni, K.L., Vaertnou, M., Palm, J.V.: Elektrokhimiya **22** (1986) 992.
86Dag Dagaeva, I., Nikolov, Z.: Elektrokhimiya **22** (1986) 1217.
86Hor Horanyi, T.S., Takács, M.: J. Electroanal. Chem. **215** (1986) 83.
86Kol Kolb, D.M., Schneider, J.: Electrochim. Acta **31** (1986) 929.
86Lus1 Lust, E.I., Palm, Yu.V.: Elektrokhimiya **22** (1986) 407.
86Lus2 Lust, E.I., Erlich, Yu.I., Palm, Yu.V.: Elektrokhimiya **22** (1986) 695.
86Pas Past, V., Pullerits, R., Moldau, M.: Trans.Tartu State Univ. **755** (1986) 140.
86Ric1 Richer, J., Lipkowski, J.: J. Electrochem. Soc. **133** (1986) 121.
86Seo Seo, M., Makino, T., Sato, N.: J. Electrochem. Soc. **133** (1986) 1138.
87Bor Borkowska, Z., Stafiej, J.: J. Electroanal. Chem. **226** (1987) 283.
87Ham1 Hamelin, A., Stoicoviciu, L.: J. Electroanal. Chem. **234** (1987) 93.
87Ham2 Hamelin, A., Doubova, L., Wagner, D., Schirmer, H.: J. Electroanal. Chem. **220** (1987) 155.
87Jar Jarzabek, G., Borkowska, Z.: J. Electroanal. Chem. **226** (1987) 295.
87Mur Murphy, O.J., Wainright, J.S.: J. Electrochem. Soc. **134** (1987) 267.
88Ann Anni, K.L., Väärtnou, M.G., Palm, U.V.: Elektrokhimiya **24** (1988) 846.
88Bor2 Borkowska, Z., Hamelin, A.: J. Electroanal. Chem. **241** (1988) 373.
88Jap Japaridze, J.I., Khutzishvili, Zh.A., Chagelishvili, V.A., Borghesani, G., De Battisti, A., Locatelli, C., Trasatti, S.: J. Electroanal. Chem. **257** (1988) 123.
88Mur Murphy, O.J., Wainright, J.S.: J. Electrochem. Soc. **135** (1988) 138.
88Vae Vaertnou, M.G., Palm, U.V.: Elektrokhimiya **24** (1988) 553.
89Aro Arold, J., Tamm, J.: Elektrokhimiya **25** (1989) 1417.
89Ham Hamelin, A., Morin, S., Richer, J., Lipkowski, J.: J. Electroanal. Chem. **272** (1989) 241.
89Mur Murphy, O.J., Wainright, J.S.: Langmuir **5** (1989) 519.
89Nan Naneva, R., Bostanov, V., Popov, O., Vitanov, T.: J. Electroanal. Chem. **274** (1989) 179.
89Saf Safonov, V.A., Damaskin, B.B., Choba, M.A.: Elektrokhimiya **25** (1989) 1432.
89Val Valette, G.: J. Electroanal. Chem. **269** (1989) 191.
90Ham Hamelin, A., Morin, S., Richer, J., Lipkowski, J.: J. Electroanal. Chem. **285** (1990) 249.
90Lus Lust, E.J., Lust, K.K., Jänes, A.A.-J.: Elektrokhimiya **26** (1990) 1627.
90Pez Pezzatini, G., Foresti, M.L., Innocenti, M., Guidelli, R.: J. Electroanal. Chem. **295** (1990) 265.
90Ric Richer, J.: Ph.D. Thesis, University of Guelph, 1990.
90Sok Sokolowski, J., Czajkowski, J.M., Turowska, M.: Electrochim. Acta **35** (1990) 1393.
91Agl Agladze, T., Podobayev, A.: Electrochim. Acta **36** (1991) 859.
91Lus1 Lust, E., Anni, K.: Double Layer and Adsorption on Solid Electrodes, Vol. IX, Tartu University, 1991, p. 109.

91Lus2 Lust, E.: Sov. Electrochem. (English Transl.) **27** (1991) 389; Elektrokhimiya **27** (1991) 424.
91Obr Obraztsov, V.B., Parfenov, Yu.A., Danilov, F.I.: Elektrokhimiya **27** (1991) 980.
91Saf Safonov, V.A., Sokolov, S.A.: Elektrokhimiya **27** (1991) 1317.
92Dic Dickinson, K.M., Hanson, K.E., Fredlein, R.A.: Electrochim. Acta **37** (1992) 139.
92Lec Lecoeur, J., Bellier, J.P., Koehler, C.: J. Electroanal. Chem. **337** (1992) 197.
92Lus Lust, E.J., Jänes, A.A.-J.: Elektrokhimiya **28** (1992) 802.
92Nan Naneva, R., Vitanov, T., Dimitrov, N., Bostanov, V., Popov, A.: J. Electroanal. Chem. **328** (1992) 287.
92Sob Sobocinski, R.L., Pemberton, J.E.: Langmuir **8** (1992) 2049.
93Dou Doubova, L.M., Trasatti, S., Valcher, S.: J. Electroanal. Chem. **349** (1993) 187.
93Inn Innocenti, M., Pezzatini, G., Foresti, M.L., Guidelli, R.: J. Electroanal. Chem. **349** (1993) 113.
93Sil Silva, F., Sottomayor, M.J., Martins, A.: J. Electroanal. Chem. **360** (1993) 199.
94Gao Gao, X.P., Edens, G.J., Weaver, M.J.: J. Electroanal. Chem. **376** (1994) 21.
94Ham Hamelin, A., Stoicoviciu, L., Edens, G.J., Gao, X.P., Weaver, M.J.: J. Electroanal. Chem. **365** (1994) 47.
94Lus Lust, E.J., Jänes, A.A.-Ja.: Elektrokhimiya **30** (1994) 357.
94Seo Seo, M., Aomi, M., Yoshida, K.: Electrochim. Acta **39** (1994) 1039.
95Att Attard, G.A., Ahmadi, A.: J. Electroanal. Chem. **389** (1995) 175.
95Koe Koene, L., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **396** (1995) 569.
95Lan Lang, G., Heusler, K.E.: J. Electroanal. Chem. **391** (1995) 169.
95Pan1 Panzram, E., Baumgärtel, H., Roelfs, B., Schröter, C., Solomun, T.: Ber. Bunsenges. Phys. Chem. **99** (1995) 827.
95Rai Raiteri, R., Butt, H.-J.: J. Phys. Chem. **99** (1995) 15728.
95Tad Tadjeddine, A., Peremans, A., Guyot-Sionnest, P.: Surf. Sci. **335** (1995) 210.
95Tra Trasatti, S., Dubova, L.: J. Chem. Soc. Faraday Trans. **91** (1995) 3311.
95Tur Turowska, M., Sokolowski, J.: Proc. 3rd Int. Symp. Electrochemistry in Practice and Theory, Lodz, Sept. 11-13, 1995.
96Efi Efimov, I.O., Heusler, K.E.: J. Electroanal. Chem. **414** (1996) 75.
96Eme Emets, V.V., Kazarinov, V.E., Bagotskaya, I.A.: Sov. Electrochem. (English Transl.) **32** (1996) 1069; Elektrokhimiya **32** (1996) 1157.
96Ham1 Hamm, U.W., Kramer, D., Zhai, R.S., Kolb, D.M.: J. Electroanal. Chem. **414** (1996) 85.
96Iwa Iwasita, T., Xia, X.H.: J. Electroanal. Chem. **411** (1996) 95.
96Jin Jin-Hua, C., Li-Hua, N., Zhuo, Y.S.: J. Electroanal. Chem. **414** (1996) 53.
96Lus1 Lust, E., Lust, K., Jänes, A.: J. Electroanal. Chem. **413** (1996) 111.
96Pem Pemberton, J.E., Joa, S.L., Shen, A.J., Woelfel, K.J.: J. Chem. Soc. Faraday Trans. **92** (1996) 3683.
96Seo Seo, M., Ueno, K.: J. Electrochem. Soc. **143** (1996) 899.
96Sil Silva, F., Sottomayor, M.J., Martins, A.: J. Chem. Soc. Faraday Trans. **92** (1996) 3693.
97Che Chen, J.-H., Si, S.-H., Nie, L.-H., Yao, S.-Z.: Electrochim. Acta **42** (1997) 689.
97For Foresti, M.L., Pezzatini, G., Innocenti, M.: J. Electroanal. Chem. **434** (1997) 191.
97Lus1 Lust, E., Jänes; A., Lust, K., Sammelselg, V., Miidla, P.: Electrochim. Acta **42** (1997) 2861.
97Lus2 Lust, E., Jänes, A., Lust, K., Miidla, P.: Electrochim. Acta **42** (1997) 771.
97Seo Seo, M., Ueno, K.: Electrochemical Society Meeting 97-2, Paris, France, 1.-5.9.1997, Ext. Abstr. p. 426.
97Tur Turowska, M., Baczynski, J., Sokolowski, J.: Elektrokhimiya **33** (1997) 1301.
98Bru Brunt, T.A., Chabala, E.D., Rayment, T., O'Shea, S.J., Welland, W.E.: Electrochemical Nanotechnology, Lorenz, W.J., Plieth, W. (eds.), Weinheim: Wiley-VCH, 1998, p. 87.
99Cli Climent, V., Gomez, R., Feliu, J.M.: Electrochim. Acta **45** (1999) 629.
99Uen Ueno, K., Seo, M.: J. Electrochem. Soc. **146** (1999) 1496.
00Alv Álvarez, B., Climent, V., Feliu, J.M., Aldaz, A.: Electrochem. Commun. **2** (2000) 427.
00Gom Gómez, R., Climent, V., Feliu, J.M., Weaver, M.J.: J. Phys. Chem. B **104** (2000) 597.
01Li Li, H.-Q., Chen, A., Roscoe, S.G., Lipkowski, J.: J. Electroanal. Chem. **500** (2001) 299.

02ElA El-Aziz, A.M., Kibler, L.A., Kolb, D.M.: Electrochem. Commun. **4** (2002) 535.
02Jur Jurkiewicz-Herbich, M., Milkowska, M., Slojkowska, R.: Colloids Surf. A **197** (2002) 235.
02Lan Langkau, T., Baltruschat, H.: Electrochim. Acta **47** (2002) 1595.
04Iwa Iwami, Y., Hobara, D., Yamamoto, M., Kakiuchi, T.: J. Electroanal. Chem. **564** (2004) 77.
04Luk Łukomska, A., Sobkowski, J.: J. Electroanal. Chem. **567** (2004) 95.
06ElA El-Aziz, A.M., Hoyer, R., Kibler, L.A., Kolb, D.M.: Electrochim. Acta **51** (2006) 2518.
06Luk Łukomska, A., Sobkowski, J.: J. Electroanal. Chem. **592** (2006) 68.
06Sol Soliman, K., Kibler, L.A., Kolb, D.M.: 57th Annual Meeting of the International Society of Electrochemistry, Edinburg, 27.8.-1.9.2006, Abstr. S8-O-16.

Table 3.2. Electrode potentials of zero charge of metal oxide electrodes in contact with electrolyte solutions. [1])

Electrode	Solution [2])	Method [3])	E_{pzc} [mV vs. SHE] [4])	Ref.
Iron (passivated) **(Fe)**	0.01 **N** $NaOH/H_2O$	DR	0.34	63Vor
PbO_2	0.001 **N** Na_2SO_4/H_2O	T	1.8	54Kab
	0.005 **M** H_2SO_4/H_2O	T	1.8	54Kab
	0.005 **M** H_2SO_4/H_2O	T	1.8	60Lei
	0.005 **M** H_2SO_4/H_2O	H	1.8	57Lei
	0.01 **N** $HClO_4/H_2O$	T	1.8	54Kab
	0.1 **N** H_2SO_4/H_2O	H	1.9	57Lei
	8 **N** H_2SO_4/H_2O	H	1.7	57Lei
	2 **N** H_2SO_4/H_2O	H	1.85	65Kuk
	2 **N** Na_2SO_4/H_2O	H	1.85	66Kuk1
	10 **N** Na_2SO_4/H_2O	H	1.9	66Kuk1
Silver oxide (AgO_2)	0.001 **N** $NaOH/H_2O$	T	0.6	53Bor

[1]) All data refer to room temperature if not stated otherwise.

[2]) A lack of an explicit concentration value indicates, that the value of E_{pzc} did not depend upon the concentration.

[3]) DR = Double layer repulsion method, H = Hardness measurements, T = Tensammetry (for details see text). See also list of symbols.

[4]) All reported values are converted with respect to the standard hydrogen electrode if not stated otherwise.

References

53Bor Borisova, T.I., Veselovskii, V.J., Karpov, L.Ya.: Zh. Fiz. Khim. **27** (1953) 1195.

54Kab Kabanov, B.N., Kiseieva, I.G., Leikis, D.I.: Dokl. Akad. Nauk SSSR **99** (1954) 805.

57Lei Leikis, D.I., Venstrem, E.K.: Dokl. Akad. Nauk SSSR **112** (1957) 97.

60Lei Leikis, D.I.: Dokl. Akad. Nauk SSSR **135** (1960) 1429.

63Vor Voropaeva, T.N., Deryagin, B.V., Kabanov, B.N.: Izv. Akad. Nauk SSSR, Otd. Khim. Nauk (1963) 257.

65Kuk Kukoz, F.I., Semenchenko, S.A.: Elektrokhimiya **1** (1965) 1454.

66Kuk1 Kukoz, F.I., Semenchenko, S.A.: Elektrokhimiya **2** (1966) 74.

Table 3.3. Electrode potentials of zero charge of metal electrodes in contact with molten electrolyte salts. [1])

Electrode	Electrolyte melt [2])	Reference electrode REF	T [°C]	Met. [3])	E_{pzc} [4]) [mV vs. REF]	Ref.
Aluminum (Al)	KCl-LiCl eutectic	Pb	750	EC	–0.3	42Kar, 44Kar
Antimony (Sb)	KCl-LiCl eutectic	Pb	750	EC	0.2	42Kar
	KCl-LiCl eutectic	Pb	750	EC	0.1	44Kar
	KCl-LiCl eutectic	Pb	1050	EC	0.2	44Kar
	KCl-LiCl 1:1	Pb-2.5 mol% $PbCl_2$	1050	EC	–0.1	60Uks
	KCl-LiCl 1:1	Pb-2.5 mol% $PbCl_2$	750	EC	0.0	60Uks
	KCl-LiCl 1:1	Pb-2.5 mol% $PbCl_2$	500	T	–0.92	60Uks
	KCl-LiCl 1:1	Pb-2.5 mol% $PbCl_2$	700	T	–0.17	60Uks
Bismuth (Bi)	KCl-LiCl eutectic	Pb	450	EC	–0.23	44Kar
	KCl-LiCl eutectic	Pb	450	EC	–0.3	48Kar
	KCl-LiCl eutectic	Pb	450	EC	–0.3	49Kar
	KCl-LiCl eutectic	Pb	450	EC	–0.3	42Kar
	KCl-LiCl eutectic	Pt/Pt^{2+} (1 **M**)	450	PC	–1.52	60Lai1
	KCl-LiCl 1:1	Pb-2.5 mol% $PbCl_2$	420	EC	–0.48	60Uks
	KCl-LiCl 1:1	Pb-2.5 mol% $PbCl_2$	700	T	–0.32	60Uks
Bi amalgams						
5.2 at% Bi	KCl-LiCl eutectic	Pb	420	EC	–0.38	48Kar
15.25 at% Bi	KCl-LiCl eutectic	Pb	420	EC	–0.32	48Kar
29 at% Bi	KCl-LiCl eutectic	Pb	420	EC	–0.3	48Kar
30.4 at% Bi	KCl-LiCl eutectic	Pb	420	EC	–0.3	48Kar
50.5 at% Bi	KCl-LiCl eutectic	Pb	420	EC	–0.3	48Kar
61.5 at% Bi	KCl-LiCl eutectic	Pb	420	EC	–0.29	48Kar
73 at% Bi	KCl-LiCl eutectic	Pb	420	EC	–0.3	48Kar
Bi-Te alloy						
10...20 at% Te	KCl–LiCl eutectic	Pb	450	EC	0.15	49Kar
Cadmium (Cd)	KCl-LiCl eutectic	Pb	400	EC	–0.79	36Kar
	KCl-LiCl eutectic	Pb	450	EC	–0.63	42Kar, 44Kar
	KCl-LiCl eutectic	Pb	450	EC	–0.6	44Str, 53Kuz
	KCl-LiCl 1:1	Pb-2.5 mol% $PbCl_2$	450	EC	–0.82	60Uks
	KCl-LiCl 1:1	Pb-2.5 mol% $PbCl_2$	700	T	–0.8	60Uks
Cd-Sn alloys						
10 at% Cd	KCl-LiCl eutectic	Pb	450	EC	–0.31	53Kuz
20 at% Cd	KCl-LiCl eutectic	Pb	450	EC	–0.35	53Kuz
40 at% Cd	KCl-LiCl eutectic	Pb	450	EC	–0.41	53Kuz
60 at% Cd	KCl-LiCl eutectic	Pb	450	EC	–0.42	53Kuz
80 at% Cd	KCl-LiCl eutectic	Pb	450	EC	–0.5	53Kuz
Chromium (Cr)	KCl-LiCl 1:1	Ag/AgCl (0.05 **M**)	700	T	–1.16	65Del1

Electrode	Electrolyte melt [2])	Reference electrode	T [°C]	Met. [3])	E_{pzc} [4]) [mV vs. REF]	Ref.
Cobalt (Co)	KCl-LiCl 1:1	Ag/AgCl (0.05 **M**)	700	T	–0.9	65Del1
Gallium (Ga)	KCl-LiCl eutectic	Pb	450	EC	–0.4	42Kar, 44Kar
	KCl-LiCl eutectic	Pb-10 wt.% $PbCl_2$	450	T	–0.63	63Buk
Indium (In)	KCl-LiCl eutectic	Pb	450	EC	–0.52	61Kuz
Iron (Fe)	KCl-LiCl 1:1	Ag/AgCl (0.05 **M**)	700	T	–0.98	65Del1
Lead (Pb)	KCl-LiCl 1:1	Pb-2.5 mol% $PbCl_2$	450	EC	–0.58	60Uks
	KCl-LiCl 1:1	Pb-2.5 mol% $PbCl_2$	700	T	–0.55	60Uks
	KBr	Pb-2.5 mol% $PbCl_2$	820	EC	–0.595	63Uks
	KCl	Pb-2.5 mol% $PbCl_2$	820	EC	–0.6	63Uks
	KCl	Graphite in melt	850	EC	1.4	10Hev
	KCl-LiCl eutectic	Pb	400	EC	–0.47	36Kar
	KCl-LiCl eutectic	Pb	450	EC	–0.47	42Kar, 44Kar
	KCl-LiCl eutectic	Pb	1050	EC	–0.38	44Kar
	KCl-LiCl eutectic	Pb-10 wt% $PbCl_2$	450	T	–0.7	63Buk
	KCl-LiCl eutectic	Pb	1050	EC	–0.38	42Kar
	KCl-KI eutectic	Graphite in melt	700	EC	1.36	10Hev
	NaBr	Pb-2.5 mol% $PbCl_2$	820	EC	–0.5	63Uks
	NaCl	Pb-2.5 mol% $PbCl_2$	820	EC	–0.47	63Uks
	NaCl-NaI eutectic	Graphite in melt	650	EC	1.44	10Hev
Manganese (Mn)	KCl-LiCl 1:1	Ag/AgCl (0.05 **M**)	700	T	–1.7	65Del1
Mercury (Hg)	KCl-LiCl eutectic	Pb	420	EC	–0.1	42Kar, 44Kar, 48Kar
	$LiNO_3$-KNO_3 (34%+66%)	$Hg/HgSO_4$	140	T	–1.12	55Ran
	$LiNO_3$-$NaNO_3$-KNO_3 (27%+18%+55%)	$Hg/HgSO_4$	140	T	–1.12	55Ran
	$LiNO_3$-$NaNO_3$ eutectic	Graphite in melt	145	EC	0.68	10Hev
Nickel (Ni)	KCl-LiCl 1:1	Ag/AgCl (0.05 **M**)	700	T	–0.72	65Del1
Platinum (Pt)	KCl-LiCl eutectic	Pt/Pt^{2+} (1 **M**)	450	T	–0.83	65Gra
	KCl-LiCl eutectic	Pt/Pt^{2+} (1 **M**)	450	PC	–1.91	60Lai1
Silver (Ag)	KCl-LiCl eutectic	Pb	1050	EC	–0.3	42Kar, 44Kar
	KCl-LiCl 1:1	Ag/AgCl (0.05 **M**)	700	T	–0.8	65Del1

Electrode	Electrolyte melt [2])	Reference electrode	T [°C]	Met. [3])	E_{pzc} [4]) [mV vs. REF]	Ref.
Tellurium (Te)	KCl-LiCl eutectic	Pb	550	EC	0.60	42Kar, 44Kar
	KCl-LiCl eutectic	Pb	475	EC	0.60	59Kuz2
	KCl-LiCl eutectic	Pb-10 wt% $PbCl_2$	550	T	0.4	63Buk
Te-Tl alloys						
2.83 at% Tl	KCl-LiCl eutectic	Pb	475	EC	0.55	59Kuz2
10.9 at% Tl	KCl-LiCl eutectic	Pb	475	EC	0.48	59Kuz2
22.1 at% Tl	KCl-LiCl eutectic	Pb	475	EC	0.35	59Kuz2
42.3 at% Tl	KCl-LiCl eutectic	Pb	475	EC	0.15	59Kuz2
53.7 at% Tl	KCl-LiCl eutectic	Pb	475	EC	0.09	59Kuz2
65.2 at% Tl	KCl-LiCl eutectic	Pb	475	EC	–0.05	59Kuz2
Thallium (Tl)	KCl-LiCl eutectic	Pb	400	EC	–1.04	36Kar
	KCl-LiCl eutectic	Pb	420	EC	–0.65	42Kar, 44Kar
	KCl-LiCl eutectic	Pb	475	EC	–0.6	59Kuz2
	KCl-LiCl 1:1	Pb-2.5 mol% $PbCl_2$	420	EC	–0.85	60Uks
	KCl-LiCl 1:1	Pb-2.5 mol% $PbCl_2$	700	T	–0.87	60Uks
Tin (Sn)	KCl-LiCl eutectic	Pb	400	EC	–0.25	36Kar
	KCl-LiCl eutectic	Pb	400	EC	–0.31	37Kar
	KCl-LiCl eutectic	Pb	400	EC	–0.18	44Str
	KCl-LiCl eutectic	Pb	450	EC	–0.23	42Kar, 44Kar
	KCl-LiCl eutectic	Pb	420	EC	–0.23	48Kar
	KCl-LiCl eutectic	Pb	450	EC	–0.2	53Kuz
	KCl-LiCl eutectic	Graphite in melt	600	EC	1.04	10Hev
	KCl-LiCl 1:1	Pb-2.5 mol% $PbCl_2$	420	EC	–0.43	60Uks
	KCl-LiCl 1:1	Pb-2.5 mol% $PbCl_2$	700	T	–0.32	60Uks
	KI-LiI (70 mol% LiI)	Pb	400	EC	–0.3	36Kar
Sn amalgams						
2.5 at% Sn	KCl-LiCl eutectic	Pb	420	EC	–0.21	48Kar
10.7 at% Sn	KCl-LiCl eutectic	Pb	420	EC	–0.32	48Kar
15.9 at% Sn	KCl-LiCl eutectic	Pb	420	EC	–0.3	48Kar
23.5 at% Sn	KCl-LiCl eutectic	Pb	420	EC	–0.27	48Kar
27.9 at% Sn	KCl-LiCl eutectic	Pb	420	EC	–0.26	48Kar
32.1 at% Sn	KCl-LiCl eutectic	Pb	420	EC	–0.25	48Kar
42.7 at% Sn	KCl-LiCl eutectic	Pb	420	EC	–0.25	48Kar
50.8 at% Sn	KCl-LiCl eutectic	Pb	420	EC	–0.24	48Kar
61.8 at% Sn	KCl-LiCl eutectic	Pb	420	EC	–0.24	48Kar
75.9 at% Sn	KCl-LiCl eutectic	Pb	420	EC	–0.23	48Kar
Sn-Au alloy						
2...69 at% Au	KCl-LiCl eutectic	Pb	450	EC	0.2	49Kar

Electrode	Electrolyte melt [2])	Reference electrode	T [°C]	Met. [3])	E_{pzc} [4]) [mV vs. REF]	Ref.
Sn-Zn alloys						
5 mol% Zn	KCl-LiCl eutectic	Pb	400	EC	–0.2	37Kar
10 mol% Zn	KCl-LiCl eutectic	Pb	400	EC	–0.3	37Kar
25 mol% Zn	KCl-LiCl eutectic	Pb	400	EC	–0.35	37Kar
50 mol% Zn	KCl-LiCl eutectic	Pb	400	EC	–0.4	37Kar
75 mol% Zn	KCl-LiCl eutectic	Pb	400	EC	–0.45	37Kar
90 mol% Zn	KCl-LiCl eutectic	Pb	400	EC	–0.5	37Kar
95 mol% Zn	KCl-LiCl eutectic	Pb	400	EC	–0.5	37Kar
5 mol% Zn	KI-LiI (70mol%LiI)	Pb	400	EC	–0.4	36Kar
10 mol% Zn	KI-LiI (70mol%LiI)	Pb	400	EC	–0.3	36Kar
25 mol% Zn	KI-LiI (70mol%LiI)	Pb	400	EC	–0.35	36Kar
50 mol% Zn	KI-LiI (70mol%LiI)	Pb	400	EC	–0.4	36Kar
75 mol% Zn	KI-LiI (70mol%LiI)	Pb	400	EC	–0.45	36Kar
90 mol% Zn	KI-LiI (70mol%LiI)	Pb	400	EC	–0.5	36Kar
95 mol% Zn	KI-LiI (70mol%LiI)	Pb	400	EC	–0.5	36Kar
Titanium (Ti)	KCl-LiCl 1:1	Ag/AgCl (0.05 **M**)	700	T	–1.35	65Del1
Zinc (Zn)	KCl-LiCl eutectic	Pb	400	EC	–0.55	37Kar
	KCl-LiCl eutectic	Pb	450	EC	–0.55	42Kar, 44Kar
	KCl-LiCl eutectic	Pb	450	EC	–0.47	44Str
	KCl-LiCl 1:1	Ag/AgCl (0.05 **M**)	700	T	–1.16	65Del1
	KI-LiI (70mol%LiI)	Pb	400	EC	–0.6	36Kar

[1]) All data refer to room temperature if not stated otherwise.

[2]) A lack of an explicitly stated composition indicates, that the value of E_{pzc} does not depend on the concentration of the constituents.

[3]) Met. = Method. EC = electrocapillary method, PC = electrode capacitance measured with a pulse method, T = tensammetry. See also list of symbols.

[4]) All reported values are given with respect to the stated reference electrode.

References

10Hev von Hevesy, G., Lorenz, R.: Z. Phys. Chem. **74** (1910) 443.

36Kar Karpatschoff, S., Stromberg, A.: Z. Phys. Chem. Abt. A **176** (1936) 182.

37Kar Karpachev, S., Stromberg, A.: Zh. Fiz. Khim. **10** (1937) 739.

42Kar Karpachev, S., Stromberg, A.G.: Acta Physicochim. URSS **16 (**1942) 331.

44Kar Karpachev, S., Stromberg, A.G.: Zh. Fiz. Khim. **18** (1944) 47.

44Str Stromberg, A., Chukina, T.: Zh. Fiz. Khim. **18** (1944) 234.

48Kar Karpachev, S., Kochergin, W.P., Iordan, E.F.: Zh. Fiz. Khim. **22** (1948) 521.

49Kar Karpachev, S.V., Rodigina, E.: Zh. Fiz. Khim. **23** (1949) 953.

53Kuz Kuznetsov, V.A., Kotschergin, W.P., Titschenko, M.W., Posdnischewa, E.G.: Dokl. Akad. Nauk SSSR **92** (1953) 1197.

55Ran Randles, J.E.B., White, W.: Z. Elektrochem. **59** (1955) 666.

59Kuz2 Kuznetsov, V.A., Akzenov, V.I., Klevzova, M.P.: Dokl. Akad. Nauk SSSR **128** (1959) 763.

60Lai1 Laitinen, H.A., Roe, D.K.: Collect. Czech. Chem. Commun. **25** (1960) 3065.

60Uks Ukshe, E.A., Bukun, N.G., Leikis, D.P.: Dokl. Akad. Nauk SSSR **135** (1960) 1183.

61Kuz Kuznetsov, V.A., Zagainova, L.S.: Zh. Fiz. Khim. **35** (1961) 1640.

63Buk Bukun, N.G., Ukshe, E.A.: Zh. Fiz. Khim. **37** (1963) 1401.

63Uks Ukshe, E.A., Tomskikh, I.V.: Dokl. Akad. Nauk SSSR **150** (1963) 347.

65Del1 Delimarski, Yu.K., Kikhno, V.S.: Ukr. Khim. Zh. **31** (1965) 116.

65Gra Graves, A.D., Inman, D.: Nature (London) **208** (1965) 481.

Table 3.4. Electrode potentials of zero charge of miscellaneous materials in contact with aqueous electrolyte solutions. [1])

Electrode	Solution	Method [2])	E_{pzc} [mV vs. SHE]	Ref.
Polypyrrole	$LiClO_4/H_2O$	CA	0.256	99Lee

[1]) All data refer to room temperature if not stated otherwise.

[2]) CA = Contact angle method. See also list of symbols.

References

99Lee Lee, M.-H., Kim, S.S., Kang, Y.: Synth. Met. **101** (1999) 442.

4 Free enthalpies of adsorption

4.1 Thermodynamics of the adsorption at electrodes

The adsorption of any particle (molecule, ion) at the electrode [1]) surface is essentially a competitive process. In the presence of only the solvent of the electrolyte solution and the electrolyte dissolved therein solely molecules of the solvent and ions formed by the dissociation of the electrolyte in the solvent are interacting with the surface. This results in case of e.g. platinum interacting with an aqueous solution in a situation wherein about 80 % of all platinum atoms are directly interacting with water molecules. The ratio of the number of metal atoms interacting with a species of the adjacent phase versus the total number of metal atoms in the topmost layer of the electrode can be defined as the degree of coverage $\theta = N_{cov}/N_{tot}$. Alternatively the ratio of the number of occupied adsorbate places versus the total number of adsorbate places can be defined as θ. If an additional kind of species (subsequently called adsorbate molecule) is dissolved in the solution phase, a further partner will participate in the competitive adsorption process. In a simplified case wherein the electrolyte ions are not interacting with the electrode to an appreciable extent (i.e. specific adsorption is absent; this can be accomplished by using e.g. salts of alkali cations and anions like e.g. F^- or ClO_4^-) only solvent and adsorbate molecules are present on the electrode surface. The degree of coverage θ of the electrode with species of both kinds will depend primarily on the relative strengths of the interaction of both species. Quantitatively this can be expressed by the free enthalpy of adsorption [2]) for the adsorbate molecule G_{ad}^0. A closer inspection has to take into account all conceivable interactions. Thus the free enthalpy of adsorption is easily defined for the case of adsorption at the gas/solid interface according to $A \rightarrow A_{ad}$ with $G_{ad,\,g}$ representing solely the free enthalpy associated with the adsorption of A on the solid surface. The corresponding formal reaction equation for adsorption from the solution phase is more complicated:

$$A_{sol} + n\ S_{ad} \rightarrow A_{ad} + n\ S_{sol} \tag{4.1}$$

with n being the number of solvent molecules S desorbing in the process. Assuming $\Delta G_{ad,S}$ being the free enthalpy associated with the solvent adsorption and $\Delta G_{ad,A}$ being the free enthalpy associated with the adsorption of A the free enthalpy of the substitution process is

$$\Delta G_{ad} = \Delta G_{ad,A} - \Delta G_{ad,S} \tag{4.2}$$

[1]) The term „electrode“ is used primarily to designate an electronically (electron or hole) conducting material. According to W. Nernst the term „electrode“ should be used to name an electronically conducting materal in combination with a specific ionically conducting material. Although this definition is convincing it is not generally accepted. Thus within this chapter „electrode“ will be used in the generally accepted meaning.

[2])The term free enthalpy of adsorption is synonymously used for free energy of adsorption and Gibbs energy of adsorption.

Obviously ΔG_{ad} depends on the strength of the solvent-surface and the adsorbate-surface interaction; in addition S-S-, A-A- and S-A-interactions contribute. Various simplifications are possible [74Tra]. As a further complication the electrical field being always present (except at E_{pzc}) at the electrode/solution-interface has to be considered. Consequently the free enthalpy of adsorption can be split into a chemical part and an electrical part:

$$\Delta G_{ad} = \Delta G_{chem} + \Delta G_{el} \tag{4.3}$$

This separation assumes, that the type of interaction causing ΔG_{chem} is not influenced by the electric field. This is certainly only a simplification. Nevertheless a comparison of the adsorption at the solid/gas- and the solid/solution-interface implies as a simple consequence

$$\Delta G_{ad} < \Delta G_{ad,g} \tag{4.4}$$

Proper definition of the reference state is essential for the compatibility of values of the free enthalpy of adsorption results in values of ΔG_{ad}^0 obtained by various authors and with different methods [87Jas, 88Gon].

The strength of the adsorptive interaction as expressed in values of ΔG_{ad}^0 depends on the mode of interaction between the adsorbate molecule and the electrode surface. Weak adsorption (physisorption) is based mostly on van-der-Waals-interactions. It shows typically values of $\Delta G_{ad}^0 > -20$ kJ mol^{-1}. Stronger interactions involve some contributions from covalent bonding, typical values are $\Delta G_{ad}^0 < -20$ kJ mol^{-1}.

Further details are collected in numerous reviews [92Tra1, 84Con, 67Gil, 67Pie, 92Lip]. Attempts to model the interface based on quantum chemistry have been reported [95Rom]. General patterns in ionic and molecular adsorption have been reviewed elsewhere [75Ans].

4.2 Determination of the free enthalpy of the adsorption at electrodes

Various methods have been employed for the determination of ΔG_{ad}^0 for electrochemical systems. A number of reviews providing further details is available [72Con, 72Pay].

Electrocapillary curves

For liquid electrode metals (mercury, gallium) the determination of the surface tension can be applied. From changes of the surface tension as a function of dissolved adsorbable species the surface coverage and finally ΔG_{ad}^0 can be derived. With solid electrodes this method is not applicable directly. (Data obtained with this method are labelled EC.)

Cyclic voltammetry

When cyclic voltammograms of an electrode partially or completely covered with an adsorbate in contact with an electrolyte solution are recorded, various characteristic features of the obtained voltammogram can be used to deduce the amount of adsorbed material. This procedure can be repeated at various concentrations of the species to be adsorbed in the solution. From the obtained relationship between coverage and concentration an adsorption isotherm can be constructed; finally a value of ΔG_{ad}^0 can be

obtained. The features of the voltammogram to be evaluated are: The amount of electric charge needed to reduce or oxidize the adsorbed species; changes in the charge corresponding to some other adsorbed (preferably hydrogen on platinum, oxygen on gold etc.), which can be related to the coverage with the species of interest, changes in the voltammogram caused by a shift of the electrochemical reaction from a pathway involving only adsorbed species to a pathway involving solution phase species. (Data obtained with this method are labelled CV.) Details of the various possible procedures can be found in the literature (see below).
Cyclic voltammetry as described above can be performed in an electrochemical cell, wherein the solution volume is confined to a space of a few micrometers thickness between the flat working electrode and an inert plate mounted adjacent to the working electrode. Connection to the counter and reference electrode is made by inserting the assembly into an electrolyte solution contained in a standard electrochemical cell. Because of the thin liquid film enclosed between the working electrode and the inert plate diffusional transport of solution species from within the bulk solution is impeded. The amount of species converted electrochemically within the time frame of a potential sweep is thus limited to the amount of liquid confined in the mentioned volume. When these species can be adsorbed on the electrode surface, the amount of species adsorbed in a single run, i.e. before the described gap is opened again in order to allow an exchange of electrolyte solution, can be calculated easily from the composition of the solution. In order to detect the number of fillings necessary to saturate the electrode surface, a check of the solution in the gap for species not adsorbed has to be made. In most cases this can be done by running a cyclic voltammogramm, which shows distinctly different features for adsorbed and dissolved species. As soon as the latter features appear, the surface is presumed to be completely covered. For further information and a review see [69Hub, 70Hub, 72Lai].

In the case of the reactive chemisorption the electrode redox potentials assigned to the chemisorption step represent the thermodynamic free energy of adsorption according to $\Delta G_{ad} = - n F E_{1/2}$. This can be visualized by considering the example of the reactive adsorption of an n-alkanethiolate on a silver electrode surface. The reaction is

$$Ag + C_nH_{2n}S^- + H^+ (a = 1) \rightarrow Ag - SC_nH_{2n} + \tfrac{1}{2} H_2 (a = 1) \text{ [98Hat]} \quad (4.5)$$

For this experimental approach the standard states as used in the evaluation of adsorption isotherms do not apply, care has thus to be taken when comparing data. (Data obtained with this method are labelled RCV.)

Cathode overpotential method

During metal deposition processes the addition of adsorbable species has been found to cause an increase in the deposition overpotential [71Lou]. Evaluation of the data results in the calculation of an adsorption isotherm. (Data obtained with this method are labelled CT.)

Tensammetry

Measurements of the interfacial capacitance (the differential double layer capacity C_{DL}) have been used widely, the method has been labelled tensammetry [46Bre, 52Bre, 51Dos1, 52Dos1, 63Bre]. Various experimental setups based on arrangements for AC polarography, lock-in-amplifier, impedance measurement etc. have been employed. In all reports evaluated in the lists of data below the authors have apparently taken precautions in order to measure only the value of C_{DL}.

Since replacement of an adsorbed solvent molecule by an adsorbate molecule generally results in a change of the composition of the first layer of the electrolyte solution being in contact with the electrode, the different values of the dielectric properties of the solvent and adsorbate molecules as well as their different size result in a change of the value of C_{DL}. Evaluation of the changes of C_{DL} as a function of the concentration of the adsorbate in the solution phase results in adsorption isotherms. Various isotherms (Frumkin, Langmuir etc.) have been used in the further evaluation [75Dam]. Because of interferences of the kinetics of the adsorption process in particular by slow transport and slow adsorption tensammetry

has been shown to provide unreliable results in some cases. (Data obtained with this method are labelled T.)

Chronocoulometry

In order to overcome certain shortcomings of tensammetry outlined above as an additional method chronocoulometry has been established [86Sto, 86Ric1, 86Ric2, 87Sto, 88Ric]. By means of electrode potential jumps between a reference value E_r, where complete desorption is established, to another value E_m, where adsorption with a certain coverage θ is established , the charge σ_M involved during this process is measured. A plot of σ_M vs. E_M allows the determination of the adsorption isotherm. (Data obtained with this method are labelled C.)

Corrosion Current Measurements

The coverage θ of a metallic electrode surface with an adsorbed species, which inhibits corrosion of the metal, can be related to the corrosion current according to

$$\theta = 1 - \frac{j_{corr,i}}{j_{corr,0}} = \frac{j_{corr,0} - j_{corr,i}}{j_{corr,0}} \tag{4.6}$$

Details of the method have been reported elsewhere [73Kam, 61Fis, 67Amm, 78Yah, 99Dur]. (Data obtained with this method are labelled CC.)

Radiotracer technique

The relationship between the amount of adsorbed particles of a certain type (ion, organic molecule etc.) and the concentration of the adsorbable species in the solution as a function of the electrode potential can be used to calculate adsorption isotherms and the free enthalpy of adsorption. Provided the species to be adsorbed is radioactive the radiotracer technique can be used. In particular species emitting β-radiation are employed. The amount of radiation emitted from the electrode is directly proportional to the amount of adsorbed species. In order to increase the surface area accessible for adsorption, electrodes to be investigated can be prepared as rough or even porous electrodes with a considerably large ratio of true to geometric surface area. Since species both in the solution phase as well as in the adsorbed state are both radioactive, the radiation emitted from the dissolved species has to be removed or at least decreased somehow. This can be achieved preferably by using a thin layer technique, wherein the electrolyte solution is squeezed away from the electrode as soon as the adsorption equilibrium is attained. The main advantage of this method is the fact, that a perfect stationary equilibrium is attained. Thus the influence of slow adsorption or diffusion of the species to be adsorbed can be excluded entirely rather easily in contrast to the electrochemical methods. Numerous reviews describing details of this method are available [95Hor, 92Boc, 86Zel, 85Hor, 84Kaz, 64Boc1, 60Blo], for general information see [65Erw]. (Data obtained with this method are labelled RT.)

Thin film resistance method

The electrical resistance of a conducting material depends according to the Drude equation on the concentration, the specific charge and the mobility of a charge carrier in the conducting medium. A more thorough examination of the mobility takes into account the way the charged species (electron, hole) is moving in the conductor. Any distortion of the material in particular in the solid phase (defects in crystals etc.) will affect the conductivity. In the case of very thin conducting films the effects of particles sitting on the surface of the thin film will be particularly strong because of the fact, that their effect will extend into almost the complete conducting layer. Consequently the phenomenon is called surface resistance. Species adsorbed on the surface will introduce additional 'defect-like' changes in the film resulting in

changes of the film conductivity [98Win]. More recently the change of the DC resistance was explained by Schumacher et al. [91Sch] by considering the adsorbate-induced density of states at the Fermi level and the half-width of the Newns-Anderson resonance corresponding to the lowest unoccupied orbital of the adsorbate molecule. Since the adsorption of neutral species depends on the electrode potential and reaches a maximum around the potential of maximum adsorption, measurements of the thin film resistance (or conductance) as a function of adsorbate concentration in the solution phase and electrode potential will provide the necessary information to construct an adsorption isotherm (see also [87Tuc, 91Kör]). As the mode and intensity of adsorbate-surface interaction may be different for different surface places (differing in local energy, topography, atomic arrangement) results may differ from those of other methods not sensitive to local differences in surface properties. A method which allows the contact resistance measurement has been proposed for the determination of electrode coverage with adsorbates, but no thermodynamic data determined this way have been reported so far [99Mar]. (Data obtained with the former method are labelled TF.)

Ellipsometric method

When linearly polarized light is reflected at a surface, the perpendicular and the parallel components undergo different changes both in amplitude and phase. Whereas both components are in phase and of the same amplitude upon incidence this may have changed considerably after reflection. The resulting optical vector of the reflected light is not any more at a given fixed angle as during incidence, instead it describes a spiral. Its projection is an ellipse, consequently measurements of the shifts outlined above upon reflection are called ellipsometry. From the obtained changes in phase shift and amplitude optical data of layers on the electrode surface can be calculated. Based on a model suggested by Paik and Bockris [71Pai] the coverage of the electrode surface with an adsorbate can be derived [92Boc]. (Data obtained with this method are labelled EL.)

Infrared spectroscopic method

The coverage of an electrode surface with adsorbed species can be calculated from infrared spectra obtained in the reflection mode. A beam of infrared radiation is directed towards the electrode surface *in situ*, i.e. it passes an infrared-transparent window and the electrolyte solution film between the window and the electrode surface. The reflected beam passes again the solution film and window and is directed towards the infrared sensitive detector. In order to remove unwanted spectral information related to the window, that the solution and the constituents of the atmosphere the infrared beam passes through outside the electrochemical cell, a modulation (or so-called differential) method has to be used. Either the electrode potential or the plane of polarization of the incident light beam is modulated. With both methods two sets of spectra are initially collected. In the former method the sets of spectra correspond to two different electrode potentials, in the latter case they correspond to infrared absorption of light polarized parallelly and perpendicularly to the plane of incidence. Calculating the spectral difference of both results in a spectrum showing only infrared absorption of species adsorbed on the electrode surface. With electrode potential modulation an additional prerequisite for bands to show up in the differential spectrum is their change in band position and/or intensity as a function of electrode potential. In the latter case no potential modulation is involved; consequently so-called absolute bands showing the infrared absorption at a given electrode potential are obtained. Because of surface selection rules only vibrational modes with a change of their dipole moment perpendicular to the surface will interact absorptively with the infrared radiation. Further details of both spectroscopic techniques have been reported in detail in numerous reports [07Hol]. Spectral absorption calculated from the infrared spectra can be related directly to the coverage of the electrode surface as detailed elsewhere [92Boc, 87Cha]. (Data obtained with this method are labelled IR.)

Weight loss method

The coverage of a metal surface with an adsorbate affecting (inhibiting or accelerating) corrosion can be deduced from the corrosion rate of the metal in the absence/presence of the adsorbate as determined by measuring the weight loss of the metal [06Bou]. (Data obtained with this method are labelled WL.)

General remarks on the tables

The following table contains values of the free enthalpy of adsorption ΔG^0_{ad} taken from literature sources. If not stated otherwise the values refer to the maximum absolute value calculated assuming standard reference states unit mole fraction for the adsorbed species in solution, unit coverage of adsorbate on the electrode surface (for further details see [74Tra], [84Nik], [91Nik], [90Tor], [93Tor]). The method employed in determining the value is stated. If a specific variant of a method is used, only the general type of method is stated, further details on the experimental approach can be found in the original report. In some cases additional information on the employed adsorption isotherm, further details on the thermodynamics as well as kinetics of the adsorption have been reported which are included in this table as footnotes.
Data are listed according to the alphabetical order of the electrode material. Polycristalline materials are listed first followed by single crystal samples and amalgam electrodes for every element. Within a given electrode material data are arranged according to the alphabetically listed adsorbate, inorganic adsorbates are listed first in this sequence. Within a given adsorbate data are listed with increasing concentration of the supporting electrolyte solution. At the same electrolyte concentration data obtained with water as a solvent are listed first, subsequently listed organic solvents are arranged according to Hill's system [00Hil] with the ordering of the element symbol and trivial names kept in the usual manner.

References

00Hil Hill, E.A.: J. Am. Chem. Soc. **22** (1900) 478.

46Bre Breyer, B., Gutmann, F.: Trans. Faraday Soc. **42** (1946) 645.

51Dos1 Doss, K.S.G., Kalyanasundaram, A.: Current Sci. **20** (1951) 199.

52Bre Breyer, B., Hacobian, S.: Aust. J. Sci. Res. Ser. A **5** (1952) 500.

52Dos1 Doss, K.S.G., Kalyanasundaram, A.: Proc. Indian Acad. Sci. Sect. A **35** (1952) 27.

60Blo Blomgren, E.A., Bockris, J.O'M.: Nature (London) **186** (1960) 305.

61Fis Fisher, H.: Ann. Univ. Ferrara, Sez. 5, 1960, suppl. No. 3 (1961) 1.

63Bre Breyer, B., Bauer, H.H.: Alternating Current Polarography and Tensammetry, New York: Interscience, 1963.

64Boc1 Bockris, J.O'M., Green, M., Swinkels, D.A.J.: J. Electrochem. Soc. **111** (1964) 743.

65Erw Erwall, L.G., Forsberg, H.G., Ljunggren, K.: Radioaktive Isotope in der Technik, Braunschweig: Friedr. Vieweg & Sohn, 1965.

67Amm Ammar, I.A., Darwish, S.: Corros. Sci. **7** (1967) 579.

67Gil Gileadi, E., in: Electrosorption, Gileadi, E. (ed.), New York: Plenum Press, 1967, p. 1.

67Pie Piersma, B.J., in: Electrosorption, Gileadi, E. (ed.), New York: Plenum Press, 1967, p. 19.

69Hub Hubbard, A.T.: J. Electroanal. Chem. **22** (1969) 165.

70Hub Hubbard, A.T., Anson, F.C.: Electroanalytical Chemistry, Vol. 4, Bard, A.J. (ed.), New York: Marcel Dekker, 1970.

71Lou Loutfy, R.O., Sukava, A.J.: J. Electrochem. Soc. **118** (1971) 216.

71Pai Paik, W.-K., Bockris, J.O'M.: Surf. Sci. **28** (1971) 61.

72Con Conway, B.E.: Techniques of Electrochemistry, Vol. 1, Yeager, E., Salkind, A.J. (eds.), New York: Wiley-Interscience, 1972, p. 389.

72Lai Lai, C.-H., Hubbard, A.T.: Inorg. Chem. **11** (1972) 2081.

72Pay Payne, R.: Techniques of Electrochemistry, Vol. 1, Yeager, E., Salkind, A.J. (eds.), New York: Wiley-Interscience, 1972, p. 43.

73Kam Kaminski, M., Szklarska-Smialowska, Z.: Corros. Sci. **13** (1973) 557.

74Tra Trasatti, S.: J. Electroanal. Chem. **53** (1974) 335.

75Ans Anson, F.C.: Acc. Chem. Res. **8** (1975) 400.

75Dam Damaskin, B.B., Petrii, O.A., Batrakov, V.V.: Adsorption organischer Verbindungen an Elektroden, Berlin: Akademie-Verlag, 1975, p. 70.

78Yah Yahalom, J., Poznansky, A.: Passivity of Metals, Frankenthal, R.P., Kruger, J. (eds.), Princeton, NJ: The Electrochemical Society, 1978, p. 328.

84Con Conway, B.E.: Prog. Surf. Sci. **16** (1984) 1.

84Kaz Kazarinov, V.E., Andreev, V.N.: Comprehensive Treatise of Electrochemistry, Vol. 9, Yeager, E., Bockris, J.O'M., Conway, B.E., Sarangapani, S. (eds.), New York: Plenum Press, 1984, p. 393.

84Nik Nikitas, P.: J. Electroanal. Chem. **170** (1984) 333.

85Hor Horanyi, G.: Wiss. Z. Humboldt-Univ. Berlin, Math. Naturwiss. Reihe **34** (1985) 37.

86Ric1 Richer, J., Lipkowski, J.: J. Electrochem. Soc. **133** (1986) 121.

86Ric2 Richer, J., Lipkowski, J.: Langmuir **2** (1986) 630.

86Sto Stolberg, L., Richer, J., Irish, D.E., Lipkowski, J.: J. Electroanal. Chem. **207** (1986) 213.

86Zel Zelenay, P., Habib, M.A., Bockris, J.O'M.: Langmuir **2** (1986) 393.

87Cha Chandrasekaran, K., Bockris, J.O'M.: Surf. Sci. **185** (1987) 495.

87Jas Jastrzebska, J., Jurkiewicz-Herbich, M., Trasatti, S.: J. Electroanal. Chem. **216** (1987) 21.

87Sto Stolberg, D.E., Irish, J., Lipkowski, J.: Electroanal. Chem. **238** (1987) 333.

87Tuc Tucceri, R.I., Posadas, D.: Electrochim. Acta **32** (1987) 27.

88Gon Gonzalez, R., Torrent, J., Sanz, F.: J. Electroanal. Chem. **244** (1988) 53.

88Ric Richer, J., Lipkowski, J.: J. Electroanal. Chem. **251** (1988) 217.

90Tor Torrent, J., Sanz, F.: J. Electroanal. Chem. **286** (1990) 207.

91Kör Körwer, D., Schumacher, D., Otto, A.: Ber. Bunsenges. Phys. Chem. **95** (1991) 1484.

91Nik Nikitas, P.: J. Electroanal. Chem. **309** (1991) 1.

91Sch Schumacher, D., Otto, A., Persson, B.N.J.: Chem. Phys. Lett. **178** (1991) 204.

92Boc Bockris, J.O'M, Jeng, K.T.: J. Electroanal. Chem. **330** (1992) 541.

92Lip Lipkowski, J., Stolberg, L.: Adsorption of Molecules at Metal Electrodes, Lipkowski, J., Ross, P.N. (eds.), New York: VCH, 1992.

92Tra1 Trasatti, S.: Electrochim. Acta **37** (1992) 2137.

93Tor Torrent, J., Sanz, F.: J. Electroanal. Chem. **359** (1993) 273.

95Hor Horanyi, G., Kolics, A.: Sov. Electrochem. (English Transl.) **31** (1995) 941.

95Rom Romanowski, S., Maksymiuk, K., Galus, Z.: J. Electroanal. Chem. **385** (1995) 95.

98Hat Hatchett, D.W., Uibel, R.H., Stevenson, K.J., Harris, J.M., White, H.S.: J. Am. Chem. Soc. **120** (1998) 1062.

98Win Winkes, H., Schumacher, D., Otto, A.: Surf. Sci. **400** (1998) 44.

99Dur Durnie, W., De Marco, R., Kinsella, B., Jefferson, A.: J. Electrochem. Soc. **146** (1999) 1751.

99Mar Marichev, V.A.: Russ. J. Electrochem. **35** (1999) 417.

06Bou Bouklah, M., Hammouti, B., Lagrenée, M., Bentiss, F.: Corros. Sci. **48** (2006) 2831.

07Hol Holze, R.: Surface and Interface Analysis: An Electrochemists Toolbox, Heidelberg: Springer-Verlag, 2007.

Table 4.1. Free enthalpies of adsorption ΔG^0_{ad} of ions and organic molecules at metal electrodes.

Electrode	Solution [1])	Adsorbate [2])	Met.[3])	ΔG^0_{ad} [4])[5]) [kJ mol^{-1}]	Ref.
Antimony (Sb)	0.1 **M** NaF/H_2O	cyclohexanol	T	–19.8	94Lus
Sb (001)	0.001 **M** NaF/H_2O	cyclohexanol	T	–20.1	97Lus2
	0.1 **M** NaF/H_2O	cyclohexanol	T	–20.1	94Lus
Sb (111)	0.001 **M** NaF/H_2O	cyclohexanol	T	–18.9	97Lus2
	0.1 **M** NaF/H_2O	cyclohexanol	T	–18.9	94Lus
Sb ($2\bar{1}\bar{1}$)	0.001 **M** NaF/H_2O	cyclohexanol	T	–20.7	97Lus2
	0.1 **M** NaF/ H_2O	cyclohexanol	T	–20.7	94Lus
Bismuth (Bi)	0.1 **M** KBr+KF/methanol	Br^-	T	–94.1	72Pet1
	0.1 **M** KCl+KF/methanol	Cl^-	T	–91.6	72Pet1
	0.1 **M** LiCl+$LiClO_4$/ethanol	Cl^-	T	–70.7 [6])	79Vaa2
	0.1 **M** LiCl+$LiClO_4$/ethanol	Cl^-	T	–75.7 [7])	79Vaa2
	0.1 **M** LiCl+$LiClO_4$/ethanol	Cl^-	T	–79.5 [8])	79Vaa2
	0.1 **M** LiCl+$LiClO_4$/ethanol	Cl^-	T	–82.8 [9])	79Vaa2
	0.1 **M** LiCl+$LiClO_4$/ethanol	Cl^-	T	–86.2 [10])	79Vaa2
	0.1 **M** LiCl+$LiClO_4$/ethanol	Cl^-	T	–87.9 [11])	79Vaa2
	0.126 **M** LiCl+CsCl/methanol	Cs^+	T	–16.8	72Pet2
	0.1 **M** KI+KF/H_2O	I^-	T	–114.0	80Pal
	0.1 **M** KI+KF/methanol	I^-	T	–107.1	72Pet1
	0.126 **M** LiCl+KCl/methanol	K^+	T	–14.0	72Pet2
	0.126 **M** LiCl+RbCl/methanol	Rb^+	T	–15.5	72Pet2
	0.1 **N** K_2SO_4/H_2O	methanol	T	–5.23	75Pul2
	0.1 **N** K_2SO_4+H_2SO_4/H_2O	ethanol	T	–7.82	69Pul
	0.1 **N** K_2SO_4/H_2O	ethanol	T	–7.8	75Pul2
	0.1 **N** K_2SO_4+H_2SO_4/H_2O	*n*-propanol	T	–11.09	69Pul
	0.1 **N** KF+H_2SO_4/H_2O	*n*-propanol	T	–11.0	72Pul
	0.1 **N** K_2SO_4/H_2O	*n*-propanol	T	–11.1	75Pul2
	0.1 **N** KF+H_2SO_4/H_2O	*n*-propanol	T	–11.07	78Pya
	0.1 **N** K_2SO_4+H_2SO_4/H_2O	*n*-butanol	T	–14.52	69Pul
	0.1 **N** K_2SO_4/H_2O	*n*-butanol	T	–14.7	75Pul2
	0.1 **N** K_2SO_4+H_2SO_4/H_2O	*n*-pentanol	T	–17.83	69Pul
	0.1 **N** K_2SO_4/H_2O	*n*-pentanol	T	–18.5	75Pul2
	0.1 **N** K_2SO_4+H_2SO_4/H_2O	*n*-hexanol	T	–21.31	69Pul
	0.1 **N** K_2SO_4+H_2SO_4/H_2O	*n*-hexanol	T	–21.31	72Alu
	0.1 **N** K_2SO_4/H_2O	*n*-hexanol	T	–22.1	75Pul2
	0.1 **N** K_2SO_4+H_2SO_4/H_2O	*n*-heptanol	T	–24.78	69Pul
	0.1 **N** K_2SO_4+H_2SO_4/H_2O	*n*-heptanol	T	–24.78	72Alu
	0.1 **N** K_2SO_4/H_2O	*n*-heptanol	T	–25.6	75Pul2
	0.1 **N** KF+H_2SO_4/H_2O	cyclohexanol	T	–17.42	72Alu

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
	0.1 **N** $KF+H_2SO_4/H_2O$	cyclohexanol	T	–17.38	78Pya
	0.05 **M** Na_2SO_4/H_2O	cyclohexanol	T	–17.4	85Lus2
	0.12 **N** KF/H_2O	phenol	T	–23.87	70Alu
	0.1 **N** $KF+H_2SO_4/H_2O$	phenol	T	–21.69	72Alu
	1 **N** $Na_2SO_4+H_2SO_4/H_2O$	acetone	T	–12.3	74Erl
	1 **N** $Na_2SO_4+H_2SO_4/H_2O$	methylethylketone	T	–14.7	74Erl
	1 **N** $Na_2SO_4+H_2SO_4/H_2O$	methylpropylketone	T	–17.9	74Erl
	1 **N** $Na_2SO_4+H_2SO_4/H_2O$	methylbutylketone	T	–20.9	74Erl
	1 **N** $Na_2SO_4+H_2SO_4/H_2O$	diethylketone	T	–18.7	74Erl
	1 **N** $Na_2SO_4+H_2SO_4/H_2O$	cyclohexanone	T	–19.2	74Erl
	0.99 **N** Na_2SO_4+ 0.01 **N** H_2SO_4/H_2O	acetic acid	T	–12.1	73Pal
	0.99 **N** Na_2SO_4+ 0.01 **N** H_2SO_4/H_2O	propionic acid	T	–13.8	73Pal
	0.99 **N** Na_2SO_4+ 0.01 **N** H_2SO_4/H_2O	butyric acid	T	–15.5	73Pal
	0.99 **N** Na_2SO_4+ 0.01 **N** H_2SO_4/H_2O	pentanoic acid	T	–18.4	73Pal
	0.99 **N** Na_2SO_4+ 0.01 **N** H_2SO_4/H_2O	hexanoic acid	T	–21.3	73Pal
	0.99 **N** Na_2SO_4+ 0.01 **N** H_2SO_4/H_2O	heptanoic acid	T	–24.7	73Pal
	0.99 **N** Na_2SO_4+ 0.01 **N** H_2SO_4/H_2O	β-pentanoic acid	T	–17.6	73Pal
	0.1 **N** $NaF+HClO_4/H_2O$	camphor	T	–22.3	79Pal
	0.1 **N** $KF+H_2SO_4/H_2O$	aniline	T	–19.8	71Alu
Bi (001)	0.1 **M** $LiCl+LiClO_4/H_2O$	Cl^-	T; C	–79.5	98Vää
	0.1 **M** $KBr+KF/H_2O$	Br^-	T; C	–84.0	02Lus
	0.1 **M** $LiBr+LiClO_4/H_2O$	Br^-	T; C	–85.8	98Vää
	0.1 **M** $LiBr+LiClO_4$/methanol	Br^-	T; C	–86.0	02Lus
	0.1 **M** $LiBr+LiClO_4$/ethanol	Br^-	T; C	–89.0	02Lus
	0.1 **M** $LiBr+Li\ ClO_4$/propanol	Br^-	T; C	–92.0	02Lus
	0.1 **M** $KJ+KF/H_2O$	I^-	T; C	–93.0	02Lus
	0.1 **M** $LiI+LiClO_4/H_2O$	I^-	T; C	–94.2	99Vää2 [12])
	0.1 **M** $LiI+LiClO_4$/methanol	I^-	T; C	–94.6	01Vää2 [13])
	0.1 **M** $LiI+LiClO_4$/ethanol	I^-	T; C	–99.8	01Vää2 [12])
	0.1 **M** $LiI+LiClO_4$/ethanol	I^-	T; C	–100.7	01Vää2 [13])
	0.1 **M** $LiI+LiClO_4$/ethanol	I^-	T; C	–102.0	02Lus
	0.1 **N** Na_2SO_4/H_2O	n-propanol	T; C	–11.5	97Lus4
	0.1 **N** Na_2SO_4/H_2O	n-butanol	T	–14.2	96Lus2
	0.1 **N** Na_2SO_4/H_2O	i-butanol	T	–12.7	96Lus2
	0.1 **N** Na_2SO_4/H_2O	2-butanol	T	–13.2	96Lus2
	0.05 **N** Na_2SO_4/H_2O	1-pentanol	T; C	–17.4	99Jän2
	0.05 **M** Na_2SO_4/H_2O	t-butanol	T	–12.6	86Lus1
	0.1 **N** Na_2SO_4/H_2O	t-butanol	T	–12.6	96Lus2
	0.1 **N** Na_2SO_4/H_2O	t-pentanol	T	–15.5	01Nur

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
Bi (001)	0.1 **N** Na_2SO_4/H_2O	heptanol	T	–25.7	03Jän
	0.05 **M** Na_2SO_4/H_2O	cyclohexanol	T	–18.5	85Lus2
	0.05 **M** Na_2SO_4/H_2O	butylacetate	T	–20.5	86Lus2
	0.1 **M** NaF/H_2O	pyridine	T	–24.7	97Lus3
	0.1 **M** NaF/H_2O	pyridine	C	–25.1	97Lus3
	0.1 **N** Na_2SO_4/H_2O	adenosine	T	–34.2 [14])	98Lus
	0.1 **N** Na_2SO_4/H_2O	adenosine	T	–34.4 [15])	98Lus
	0.1 **N** Na_2SO_4/H_2O	adenosine	T	–34.2	01Jän
	0.1 **N** Na_2SO_4/H_2O	d-ribose	T	–16.66	99Jän1
	0.1 **N** Na_2SO_4/H_2O	d-ribose	T	–16.52	01Jän
	0.1 **M** KBr+KF/H_2O	Br^-	T; C	–89.0	02Lus
	0.1 **M** LiBr+$LiClO_4/H_2O$	Br^-	T; C	–88.3	98Vää
	0.1 **M** LiBr+$LiClO_4$/ethanol	Br^-	T; C	–90.0	02Lus
	0.1 **M** LiBr+$LiClO_4$/methanol	Br^-	T; C	–85.8	99Vää1 [13])
	0.1 **M** LiBr+$LiClO_4$/methanol	Br^-	T; C	–86.0	99Vää1 [12])
	0.1 **M** LiBr+$LiClO_4$/methanol	Br^-	T; C	–86.0	98Vää
	0.1 **M** LiBr+Li ClO_4/methanol	Br^-	T; C	–86.0	99Vää2 [12])
	0.1 **M** LiBr+$LiClO_4$/2-propanol	Br^-	T; C	–90.8	98Vää
	0.1 **M** LiBr+$LiClO_4$/methanol	Br^-	T; C	–86.0	01Vää1
	0.1 **M** LiBr+Li ClO_4/ethanol	Br^-	T; C	–89.2	01Vää1
	0.1 **M** LiBr+$LiClO_4$/ethanol	Br^-	T; C	–89.2	01Vää1 [12])
	0.1 **M** LiBr+$LiClO_4$/2-propanol	Br^-	T; C	–91.4	01Vää1
	0.1 **M** LiBr+$LiClO_4$/2-propanol	Br^-	T	–90.7	00Vää [12]), 96Vää
	0.1 **M** LiBr+$LiClO_4$/2-propanol	Br^-	T	–91.4	00Vää [13]), 96Vää
	0.1 **M** LiBr+$LiClO_4$/methanol	Br^-	T; C	–86.8	01Vää1
	0.1 **M** LiBr+$LiClO_4$/ethanol	Br^-	T; C	–89.8	01Vää1
	0.1 **M** LiBr+$LiClO_4$/methanol	Br	T; C	–86.2	99Vää1 [13])
	0.1 **M** LiBr+$LiClO_4$/methanol	Br^-	T; C	–86.8	99Vää2 [12])
	0.1 **M** LiBr+$LiClO_4$/methanol	Br^-	T; C	–86.8	99Vää1 [12])
	0.1 **M** LiBr+$LiClO_4$/ methanol	Br^-	T; C	–86.8	98Vää
	0.1 **M** LiCl+$LiClO_4/H_2O$	Cl^-	T; C	–80.5	98Vää
	0.1 **M** LiCl+$LiClO_4$/methanol	Cl^-	T; C	–81.7	99Vää1 [13])
	0.1 **M** LiCl+$LiClO_4$/methanol	Cl^-	T; C	–81.3	99Vää2 [12])
	0.1 **M** LiCl+$LiClO_4$/methanol	Cl^-	T; C	–81.3	99Vää1 [12])
	0.1 **M** LiCl+$LiClO_4$/methanol	Cl^-	T; C	–81.3	98Vää
	0.1 **M** LiCl+$LiClO_4$/methanol	Cl^-	T; C	–81.4	99Vää1 [13])
	0.1 **M** LiCl+$LiClO_4$/methanol	Cl^-	T; C	–81.3	98Vää
	0.1 **M** LiCl+$LiClO_4$/methanol	Cl^-	T; C	–81.3	99Vää2 [12])
	0.1 **M** LiCl+$LiClO_4$/methanol	Cl^-	T; C	–81.3	99Vää1 [12])
	0.1 **M** LiCl+$LiClO_4$/ethanol	Cl^-	T; C	–82.3	98Vää
	0.1 **M** LiCl+Li ClO_4/ethanol	Cl^-	T; C	–82.8	98Vää
	0.1 **M** LiCl+$LiClO_4$/2-propanol	Cl^-	T; C	–84.8	98Vää

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
	0.1 **M** LiCl+$LiClO_4$/2-propanol	Cl^-	T	–84.8	00Vää [12]), 96Vää
	0.1 **M** LiCl+$LiClO_4$/2-propanol	Cl^-	T	–85.3	00Vää [13]), 96Vää
	0.1 **M** LiI+$LiClO_4$/methanol	I^-	T; C	–97.5	99Vää1 [12])
	0.1 **M** LiI+$LiClO_4$/methanol	I^-	T; C	–96.3	99Vää1 [13])
	0.1 **M** LiI+$LiClO_4$/methanol	I^-	T; C	–97.5	99Vää2 [12])
	0.1 **M** LiI+$LiClO_4$/ethanol	I^-	T; C	–98.7	99Vää2 [12])
	0.1 **M** LiI+$LiClO_4$/acetonitrile	I^-	T; C	–110.0	02Lus
	0.1 **M** LiI+Li ClO_4/methanol	I^-	T; C	–94.6	99Vää1 [13])
	0.1 **M** LiI+$LiClO_4$/ethanol	I^-	T; C	–98.7	99Vää2 [12])
	0.1 **M** LiI+$LiClO_4$/methanol	I^-	T; C	–95.2	99Vää1 [12])
	0.1 **M** LiI+$LiClO_4$/methanol	I^-	T; C	–95.2	99Vää2 [12])
	0.1 **M** KI+KF/H_2O	I^-	T; C	–87.0	02Lus
	0.1 **M** LiI+$LiClO_4$/H_2O	I^-	T; C	–100.2	99Vää2 [12])
	0.1 **M** KI+KF/H_2O	I^-	T	–116.0	80Pal
	0.1 **M** LiI+Li ClO_4/methanol	I^-	T; C	–96.3	01Vää2 [13])
	0.1 **M** LiI+Li ClO_4/ethanol	I^-	T; C	–101.2	01Vää2 [12])
	0.1 **M** LiI+Li ClO_4/ethanol	I^-	T; C	–102.4	01Vää2 [13])
	0.1 **M** LiI+Li ClO_4/ethanol	I^-	T; C	–103.0	02Lus
	0.1 **N** Na_2SO_4/H_2O	n-propanol	T; C	–12.3	97Lus4
	0.1 **N** Na_2SO_4/H_2O	n-butanol	T	–15.1	96Lus2
	0.1 **N** Na_2SO_4/H_2O	2-butanol	T	–14.5	96Lus2
	0.1 **N** Na_2SO_4/H_2O	i-butanol	T	–14.5	96Lus2
	0.05 **M** Na_2SO_4/H_2O	t-butanol	T	–13.6	86Lus1
	0.1 **N** Na_2SO_4/H_2O	t-butanol	T	–13.6	96Lus2
	0.1 **N** Na_2SO_4/H_2O	t-pentanol	T	–17.3	01Nur
	0.1 **N** KF+H_2SO_4/H_2O	cyclohexanol	T	–19.77	78Pal
	0.001 **M** NaF/H_2O	cyclohexanol	T	–20.1	97Lus2
	0.05 **M** Na_2SO_4/H_2O	cyclohexanol	T	–20.1	85Lus2
	0.1 **N** Na_2SO_4/H_2O	heptanol	T	–27.3	03Jän
	0.05 **M** Na_2SO_4/H_2O	butylacetate	T	–22.1	86Lus2
	0.1 **N** KF+H_2SO_4/H_2O	aniline	T	–20.0	82Pal
	0.1 **M** NaF/H_2O	pyridine	T	–25.4	97Lus3
	0.1 **M** NaF/H_2O	pyridine	C	–26.2	97Lus3
	0.05 **N** Na_2SO_4/H_2O	1-pentanol	T; C	–19.2	99Jän2
	0.1 **N** Na_2SO_4/H_2O	d-ribose	T	–30.76	99Jän1
	0.1 **N** Na_2SO_4/H_2O	d-ribose	T	–30.66	01Jän
	0.05 **M** Na_2SO_4/H_2O	uracil	T	–17.1	05Kas
Bi $(\bar{1}0\bar{1})$	0.001 **M** NaF/H_2O	cyclohexanol	T	–19.0	97Lus2
	0.1 **N** Na_2SO_4/H_2O	n-butanol	T	–14.4	96Lus2
	0.1 **N** Na_2SO_4/H_2O	2-butanol	T	–14.0	96Lus2
	0.05 **M** Na_2SO_4/H_2O	t-butanol	T	–13.0	86Lus1
	0.1 **N** Na_2SO_4/H_2O	t-butanol	T	–13.0	96Lus2
	0.05 **M** Na_2SO_4/H_2O	cyclohexanol	T	–19.04	85Lus2

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
Bi ($\bar{1}0\bar{1}$)	0.05 **M** Na_2SO_4/H_2O	butylacetate	T	–21.3	86Lus2
Bi (111)	0.1 **M** KBr+KF/ H_2O	Br^-	T; C	–83.0	02Lus
	0.1 **M** KBr+KF/H_2O	Br^-	T	–84	00Lus
	0.1 **M** LiBr+$LiClO_4/H_2O$	Br^-	T; C	–85.3	98Vää
	0.1 **M** LiBr+$LiClO_4$/methanol	Br^-	T; C	–85.8	99Vää1 [13])
	0.1 **M** LiBr+$LiClO_4$/methanol	Br^-	T; C	–85.3	99Vää2 [12])
	0.1 **M** LiBr+$LiClO_4$/ethanol	Br^-	T; C	–88.0	02Lus
	0.1 **M** LiBr+$LiClO_4$/methanol	Br^-	T; C	–85.3	99Vää1 [12])
	0.1 **M** KBr+KF/methanol	Br^-	T	–88	00Lus
	0.1 **M** LiCl+$LiClO_4$/methanol	Br^-	T; C	–85.3	98Vää
	0.1 **M** LiBr+$LiClO_4$/methanol	Br^-	T; C	–85.8	01Vää1
	0.1 **M** LiBr+$LiClO_4$/ethanol	Br^-	T; C	–88.8	01Vää1
	0.1 **M** LiBr+$LiClO_4$/ethanol	Br^-	T; C	–88.7	01Vää1 [12])
	0.1 **M** KBr+KF/2-propanol	Br^-	T	–93	00Lus
	0.1 **M** LiBr+$LiClO_4$/2-propanol	Br^-	T	–90.6	00Vää [13]), 96Vää
	0.1 **M** LiBr+$LiClO_4$/2-propanol	Br^-	T	–89.9	00Vää [12]), 96Vää
	0.1 **M** LiBr+$LiClO_4$/2-propanol	Br^-	T; C	–89.8	98Vää
	0.1 **M** LiBr+$LiClO_4$/2-propanol	Br^-	T; C	–90.6	01Vää1
	0.1 **M** KCl+KF/H_2O	Cl^-	T	–79	00Lus
	0.1 **M** KCl+KF/methanol	Cl^-	T	–82	00Lus
	0.1 **M** LiCl+$LiClO_4/H_2O$	Cl^-	T; C	–80.8	98Vää
	0.1 **M** LiCl+$LiClO_4$/methanol	Cl^-	T; C	–79.8	99Vää1 [13])
	0.1 **M** LiCl+$LiClO_4$/methanol	Cl^-	T; C	–79.8	99Vää2 [12])
	0.1 **M** LiCl+$LiClO_4$/methanol	Cl^-	T; C	–79.8	99Vää1 [12])
	0.1 **M** LiCl+$LiClO_4$/methanol	Cl^-	T; C	–79.8	98Vää
	0.1 **M** LiCl+$LiClO_4$/ethanol	Cl^-	T; C	–79.8	98Vää
	0.1 **M** KCl+KF/2-propanol	Cl^-	T	–83	00Lus
	0.1 **M** LiCl+$LiClO_4$/2-propanol	Cl^-	T	–83.0	00Vää [12]), 96Vää
	0.1 **M** LiCl+$LiClO_4$/2-propanol	Cl^-	T	–83.3	00Vää [13]), 96Vää
	0.1 **M** LiCl+$LiClO_4$/2-propanol	Cl^-	T; C	–83.0	98Vää
	0.1 **M** KI+KF/H_2O	I^-	T	–112.0	80Pal
	0.1 **M** LiI+$LiClO_4/H_2O$	I^-	T; C	–92.2	99Vää2 [12])
	0.1 **M** KI+KF/H_2O	I^-	T	–96	00Lus
	0.1 **M** KI+KF/H_2O	I^-	T; C	–96.0	02Lus
	0.1 **M** LiI+$LiClO_4$/methanol	I^-	T; C	–94.2	01Vää2 [13])
	0.1 **M** LiI+$LiClO_4$/ethanol	I^-	T; C	–98.5	01Vää2 [12])
	0.1 **M** LiI+$LiClO_4$/ethanol	I^-	T; C	–99.2	01Vää2 [13])
	0.1 **M** LiI+$LiClO_4$/methanol	I^-	T; C	–96.0	02Lus
	0.1 **M** LiI+$LiClO_4$/ethanol	I^-	T; C	–99.0	02Lus
	0.1 **M** LiI+$LiClO_4$/methanol	I^-	T; C	–95.2	99Vää1 [13])
	0.1 **M** LiI+$LiClO_4$/methanol	I^-	T; C	–94.2	99Vää2 [12])

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
	0.1 **M** LiI+$LiClO_4$/methanol	I^-	T; C	–94.2	99Vää1 [12])
	0.1 **M** KI+KF/methanol	I^-	T	–97	00Lus
	0.1 **M** LiI+$LiClO_4$/ethanol	I^-	T; C	–97.0	99Vää2 [12])
	0.1 **M** KI+KF/2-propanol	I^-	T	–96	00Lus
	0.1 **M** LiI+$LiClO_4$/acetonitrile	I^-	T; C	–108.0	02Lus
	0.02 (1–x) **M** $LiClO_4$+ 0.02x **M** NEt_4ClO_4/ethanol	NEt_4^+	T	–81.0	06Vää
	0.1 **N** Na_2SO_4/H_2O	n-butanol	T	–13.9	96Lus2
	0.1 **N** Na_2SO_4/H_2O	i-butanol	T	–12.7	96Lus2
	0.1 **N** Na_2SO_4/H_2O	2-butanol	T	–13.1	96Lus2
	0.05 **M** Na_2SO_4/H_2O	t-butanol	T	–12.2	86Lus1
	0.1 **N** Na_2SO_4/H_2O	t-butanol	T	–12.2	96Lus2
	0.1 **N** Na_2SO_4/H_2O	t-butanol	C	–12.0	96Lus2
	0.1 **N** Na_2SO_4/H_2O	n-propanol	T; C	–11.2	97Lus4
	0.05 **N** Na_2SO_4/H_2O	1-pentanol	T; C	–17.0	99Jän2
	0.1 **N** Na_2SO_4/H_2O	t-pentanol	T	–17.0	01Nur
	0.1 **N** KF+H_2SO_4/H_2O	cyclohexanol	T	–17.01	78Pal
	0.001 **M** NaF/H_2O	cyclohexanol	T	–17.1	97Lus2
	0.001 **M** NaF/H_2O	cyclohexanol	C	–16.8	97Lus2
	0.05 **M** Na_2SO_4/H_2O	cyclohexanol	T	–17.1	85Lus2
	0.1 **N** Na_2SO_4/H_2O	heptanol	T	–24.5	03Jän
	0.05 **M** Na_2SO_4/H_2O	butylacetate	T	–18.9	86Lus2
	0.1 **N** Na_2SO_4/H_2O	d-ribose	T	–16.52	99Jän1
	0.1 **N** Na_2SO_4/H_2O	d-ribose	T	–16.66	01Jän
	0.1 **N** NaF+ $HClO_4/H_2O$	camphor	T	–21.6	79Pal
	0.1 **N** KF+H_2SO_4/H_2O	aniline	T	–18.5	82Pal
	0.1 **M** NaF/ H_2O	pyridine	T	–22.1	97Lus3
	0.1 **M** NaF/ H_2O	pyridine	C	–24.3	97Lus3
	0.1 **N** Na_2SO_4/H_2O	adenosine	T	–32.7 [14])	98Lus
	0.1 **N** Na_2SO_4/H_2O	adenosine	T	–31.5 [15])	98Lus
	0.1 **N** Na_2SO_4/H_2O	adenosine	T	–32.7	01Jän
	0.05 **M** Na_2SO_4/H_2O	uracil	T	–16.4	05Kas
Bi ($2\bar{1}\bar{1}$)	0.1 **M** KI+KF/H_2O	I^-	T	–114.0	80Pal
	0.1 **N** KF+H_2SO_4/H_2O	n-propanol	T	–13.2	78Pya
	0.1 **N** Na_2SO_4/H_2O	i-butanol	T	–14.2	96Lus2
	0.1 **N** Na_2SO_4/H_2O	2-butanol	T	–14.4	96Lus2
	0.1 **N** Na_2SO_4/H_2O	t-butanol	T	–13.4	96Lus2
	0.1 **N** KF+H_2SO_4/H_2O	cyclohexanol	T	–18.55	78Pal
	0.1 **N** KF+H_2SO_4/H_2O	cyclohexanol	T	–18.55	78Pya
	0.001 **M** NaF/H_2O	cyclohexanol	T	–18.6	97Lus2
	0.1 **N** NaF+ $HClO_4/H_2O$	camphor	T	–23.2	79Pal
	0.1 **N** KF+H_2SO_4/H_2O	aniline	T	–19.3	82Pal
Cadmium (Cd)	0.2 **N** KI+KF/H_2O	I^-	T	–92.0	72Pan2
	0.1 **N** KF/H_2O	ethanol	T	–7.46	75Ryb1
	0.1 **N** KF/H_2O	n-propanol	T	–9.53	75Ryb1

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
Cd	0.1 **N** KF/H_2O	n-propanol	T	–9.61	73Ryb1
	0.1 **N** NaF/H_2O	n-propanol	T	–10.5	91Obr
	1 **N** Na2SO_4+0.05 **N** H_2SO_4/H_2O	n-propanol	T	–10.4	91Obr
	0.1 **N** KF/H_2O	n-butanol	T	–12.85	75Ryb1
	0.1 **N** KF/H_2O	n-pentanol	T	–16.00	75Ryb1
	0.1 **N** KF/H_2O	cyclohexanol	T	–8.29	73Ryb2
Cd (000)	0.001 **M** NaF/H_2O	cyclohexanol	C	–15.8	97Lus2
Cd (0001)	0.001 **M** NaF/H_2O	cyclohexanol	T	–16.2	97Lus2
Cd (10$\bar{1}$0)	0.001 **M** NaF/H_2O	cyclohexanol	T	–17.4	97Lus2
Cd (10$\bar{2}$0)	0.001 **M** NaF/H_2O	cyclohexanol	T	–17.3	97Lus2
Carbon (C)	0.1 **N** $NaClO_4$+0.1 **N** $HClO_4$/H_2O	*N*,*N*,*N*',*N*'-tetra-methyl-1,1'-naphthidine	CV	–24.68 [16])	93Ote
Copper (Cu)	1 **M** H_2SO_4+0.5 **M** $CuSO_4$/H_2O	propanoic acid	CT	–12.55 [17])	71Lou
	1 **M** H_2SO_4+0.5 **M** $CuSO_4$/H_2O	butanoic acid	CT	–15.5 [17])	71Lou
	1 **M** H_2SO_4+0.5 **M** $CuSO_4$/H_2O	pentanoic acid	CT	–18.4 [17])	71Lou
	1 **M** H_2SO_4+0.5 **M** $CuSO_4$/H_2O	hexanoic acid	CT	–21.3 [17])	71Lou
	1 **M** H_2SO_4+0.5 **M** $CuSO_4$/H_2O	heptanoic acid	CT	–24.3 [17])	71Lou
	1 **M** H_2SO_4+0.5 **M** $CuSO_4$/H_2O	octanoic acid	CT	–27.2 [17])	71Lou
	0.9 **N** $NaClO_4$+0.1 **N** NaOH/H_2O	n-decylamine	RT	–30.51	64Boc2
	1 **N** $NaClO_4$/H_2O	naphthalene	RT	–29.3	64Boc1
	0.1 **M** $HClO_4$/H_2O	benzoic acid	RT	–18.6 [18])	95Was
	0.1 **M** $HClO_4$/H_2O	benzoic acid	RT	–26.4 [19])	95Was
	0.1 **M** $HClO_4$/H_2O	benzoic acid	RT	–26.4	95Sob
	0.5 **M** HCl/H_2O	imidazole	CC	–14.5	00Gas
	0.5 **M** HCl/H_2O	4-methylimidazole	CC	–13.9	00Gas
	0.5 **M** HCl/H_2O	4-methyl-5-hydroxy-methylimidazole	CC	–14.2	00Gas
	0.5 **M** HCl/H_2O	1-phyenyl-4-methyl-imidazole	CC	–15.2	00Gas
	0.5 **M** HCl/H_2O	1-(p-tolyl)-4-methyl-imidazole	CC	–16.3	00Gas
Cu (100)	0.1 **M** $HClO_4$/H_2O	thiourea	RT	–22.3	01Luk
	0.01 **M** $NaClO_4$/H_2O	thiourea	RT	–33.1	06Luk
Cu (110)	0.1 **M** $HClO_4$/H_2O	thiourea	RT	–22.4	01Luk
	0.01 **M** $NaClO_4$/H_2O	thiourea	RT	–31.9	06Luk
Cu (111)	0.1 **M** $HClO_4$/H_2O	thiourea	RT	–22.3	01Luk
	0.01 **M** $NaClO_4$/H_2O	thiourea	RT	–33.4	06Luk
Gallium (Ga)	0.5 **M** Na_2SO_4/H_2O	2-propanol	C	–9.2	91Pez
Ga-Bi [20])	0.05 **M** Na_2SO_4/H_2O	n-butanol	T	–14.55	05Eme

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
Ga-Pb [21])	0.05 **M** Na_2SO_4/H_2O	n-butanol	T	–14.55	05Eme
Ga-Sn [22])	0.05 **M** Na_2SO_4/H_2O	n-butanol	T	–14.55	05Eme
	0.5 **M** Na_2SO_4/H_2O	t-pentanol	C	–16.3	91Pez
Gold	1 **N** H_2SO_4/H_2O	cyclohexane	RT	–24.3	63Dah
(Au)	0.02 **M** $NaClO_4/H_2O$	1,2-ethanediol	T	–6.9	93Luc
	0.1 **M** $KClO_4/H_2O$	1,2-butanediol	T	–12.8	93Luc
	0.02 **M** $NaClO_4/H_2O$	1-propanol	T	–5.0	86Hol
	0.02 **M** NaF/H_2O	2-propanol	T	–8.2	98Bel
	0.02 **M** $NaClO_4/H_2O$	2-propen-1-ol	T	–31.8	95Luc1
	0.02 **M** $NaClO_4/H_2O$	2-propin-2-ol	T	–34.0	95Luc1
	0.02 **M** $NaClO_4/H_2O$	1-butanol	T	–8.5	85Hol
	0.02 **M** $NaClO_4/H_2O$	2-butanol	T	–11.1	86Hol
	0.04 **M** $KClO_4/H_2O$	t-butanol	T	–9.63	96Sta
	0.02 **M** $NaClO_4/H_2O$	1-pentanol	T	–13.4	87Tuc
	0.02 **M** $NaClO_4/H_2O$	1-pentanol	T	–13.7	85Hol
	0.02 **M** $NaClO_4/H_2O$	2-pentanol	T	–14.5	85Hol
	0.02 **M** $NaClO_4/H_2O$	1-hexanol	T	–17.9	85Hol
	0.02 **M** $NaClO_4/H_2O$	2-hexanol	T	–16.5	86Hol
	0.04 **M** $KClO_4/H_2O$	1,2-cyclohexanediol	T	–19.0	95Luc2
	0.04 **M** $KClO_4/H_2O$	1,2-cyclohexanediol	T	–18.9	89Hol
	0.04 **M** $KClO_4/H_2O$	trans-1,2-cyclohexanediol	T	–17.5	94Luc
	0.04 **M** $KClO_4/H_2O$	1,3-cyclohexanediol	T	–19.6	89Hol
	0.04 **M** $KClO_4/H_2O$	1,4-cyclohexanediol	T	–16.1	89Hol
	0.02 **M** $NaClO_4/H_2O$	1-heptanol	T	–21.2	86Hol
	0.02 **M** $NaClO_4/H_2O$	2-heptanol	T	–19.5	86Hol
	1 **N** H_2SO_4/H_2O	benzene	RT	–25.5	63Dah
	1 **N** H_2SO_4/H_2O	naphthalene	RT	–2.9	63Dah
	0.1 **M** $KClO_4/H_2O$	p-toluenesulphonate [23])	T	–40.00	92Sko1
	0.01 **M** $NaClO_4/H_2O$	β-cyclodextrine	T	–34.70	90Jar
	0.1 **M** $HClO_4/H_2O$	benzoic acid	RT	–16.4 [24])	94Zel
	0.1 **M** $HClO_4/H_2O$	benzoic acid	RT	–23.1 [25])	94Zel
	0.1 **M** $HClO_4/H_2O$	benzoic acid	RT	–16.4	95Sob
	0.1 **M** $HClO_4/H_2O$	benzoic acid	T	–43.7	71Kat
	0.04 **M** $KClO_4/H_2O$	urea	T	–24.0	90Hol
	0.04 **M** $KClO_4/H_2O$	aniline	T	–42.45	90Pap
	0.1 **M** $KClO_4/H_2O$	aniline	T	–47.00	89Ngu
	0.1 **M** $KClO_4/H_2O$	aniline	T	–45.08	90Pap
	0.04 **M** $KClO_4/H_2O$	o-toluidine	T	–34.05	89Sch
	0.04 **M** $KClO_4/H_2O$	m-toluidine	T	–34.43	89Sch
	0.04 **M** $KClO_4/H_2O$	p-toluidine	T	–34.35	89Sch
	0.04 **M** $KClO_4/H_2O$	*N*,*N*-dimethylaniline	T	–34.05	89Sch
	0.04 **M** $KClO_4/H_2O$	2,5-dimethylaniline	T	–35.42	89Sch
	0.1 **M** $KClO_4/H_2O$	pyridine	C	–38.0	86Sto
	0.1 **M** $KClO_4/H_2O$	pyrazine	C	–32.5	97Ian

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
Au	0.1 **M** $LiClO_4/H_2O$	DL-serine	C	–37.0	01Slo
	0.1 **M** $LiClO_4/H_2O$	L-tyrosine	C	–46.2	01Slo
	0.1 **M** $LiClO_4/H_2O$, pH=9.8	L-histidine	C	–44.6 [26])	01Slo
	0.1 **M** $LiClO_4/H_2O$, pH=9.8	L-histidine	C	–42.4 [27])	01Slo
	0.1 **M** $LiClO_4/H_2O$, pH=1.4...7	L-histidine	C	–42.0 [26])	01Slo
	0.1 **M** $LiClO_4/H_2O$, pH=1.4...7	L-histidine	C	–41.4 [27])	01Slo
	0.1 **M** $LiClO_4/H_2O$, pH=1.4...7	L-histidine	C	–42.3 [28])	01Slo
	0.04 **M** $KClO_4/H_2O$	thiourea	T	–43.3	90Hol
	0.1 **M** $NaClO_4/H_2O$	thiourea	T	–26.5	94Brz
	0.05 **M** KF/methanol	thiourea	T	–21.5	94Brz
Au (100)	0.1 **M** $KClO_4/H_2O$	t-pentanol	C	–17.4	88Ric
	0.04 **M** $KClO_4/H_2O$	2-butene-1,4-diol	T	–29.8	93Luc
	0.04 **M** $KClO_4/H_2O$	2-butyne-1,4-diol	T	–29.0	93Luc
	0.01 **M** NaF/H_2O	diethylether	C	–13.5	83Lip
	0.1 **M** $KClO_4/H_2O$	benzonitrile	C	–27.9	90Ric
	0.1 **M** $KClO_4/H_2O$	pyridine	C	–36.0	87Sto
Au (110)	0.05 **M** $KClO_4/H_2O$	t-pentanol	C	–18.5	86Ric2
	0.01 **M** NaF/H_2O	diethylether	C	–10.0	83Lip
	0.1 **M** $KClO_4/H_2O$	pyridine	C	–42.0	90Sto
Au (111)	0.1 **M** $HClO_4/H_2O$	SO_4^{2-}	C	–110.0 [13])	95Shi
	$NaClO_4/H_2O$	n-hexanol	C	–22.3	97Sil
	$NaClO_4/H_2O$	n-hexanol	T	–22.3	97Sil
	0.01 **M** NaF/H_2O	diethylether	C	–15.5	83Lip
	0.02 **M** $HClO_4/H_2O$	sorbitol	T	–17.3	94Sot
	0.02 **M** $HClO_4/H_2O$	mannitol	T	–12.4	94Sot
	0.02 **M** $NaClO_4/H_2O$	sorbitol	T	–17.7	94Sot
	0.02 **M** $NaClO_4/H_2O$	mannitol	T	–15.2	94Sot
	0.1 **M** $KClO_4/H_2O$	benzonitrile	C	–27.0	92Ric
	0.1 **M** $KClO_4/H_2O$	pyridine	C	–37.0	91Sto
	0.1 **M** $KClO_4/H_2O$	pyrazine	C	–26.0	94Ian
	0.1 **M** $KClO_4/H_2O$	pyrazine	C	–26.2	97Ian
	0.1 **M** $KClO_4/H_2O$	benzonitrile	C	–28.1	90Ric
	0.1 **M** $KClO_4/H_2O$	4-cyanopyridine	C	–37.0	95Yan
	0.05 **M** $KClO_4/H_2O$	1,3-dimethyluracil	C	–35.4	96Wan1
	0.05 **M** $KClO_4/H_2O$	uracil	C	–14.85	96Wan2
	0.05 **M** $KClO_4/H_2O$	5-methyluracil	C	–17.95	96Wan1
	0.05 **M** $KClO_4/H_2O$	5,6-dimethyluracil	C	–22.1	96Wan2
	0.1 **M** $KClO_4/H_2O$	L-phenylalanine	C	–27.0 [13])	01Li
	0.1 **M** $NaClO_4/H_2O$ [29])	benzenesulfonic acid	T	–35.0	96Dut
	0.1 **M** $KClO_4/H_2O$	bromide ions	T	–140.0 [13])	96Shi
Au (210)	0.1 **M** $KClO_4/H_2O$	pyrazine	C	–39.0	97Ian
Au (211)	0.1 **M** $KClO_4/H_2O$	benzonitrile	C	–26.4	90Ric
Au (311)	0.1 **M** $NaClO_4/H_2O$	*p*-toluenesulphonate	C	–35.0	92Sko2
	0.1 **M** $KClO_4/H_2O$	benzonitrile	C	–26.2	90Ric

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
	0.1 **M** $KClO_4/H_2O$	pyridine	C	–44.0	92Sto
Au (322)	0.1 **M** $KClO_4/H_2O$	benzonitrile	C	–25.8	90Ric
Au (511)	0.1 **M** $KClO_4/H_2O$	benzonitrile	C	–25.6	90Ric
Indium (In)					
In-Ga eutectic alloy (16.4 at.% In)	0.1 **N** Na_2SO_4/H_2O	ethanol	T	–6.13	74Gri
	0.1 **N** Na_2SO_4/H_2O	propanol	T	–11.31	74Gri
	0.1 **N** Na_2SO_4/H_2O	n-butanol	T	–14.46	74Gri
	0.1 **N** Na_2SO_4/H_2O	n-pentanol	T	–17.03	74Gri
Iron (Fe)	1 **N** $NaClO_4/H_2O$	naphthalene	RT	–29.3	64Boc1
	0.9 **N** $NaClO_4$+0.1 **N** $NaOH/H_2O$	n-decylamine	RT	–27.58	64Boc2
mild steel	3% W/V NaCl+100 mg Na_2CO_3 100 ml H_2O	hexadecenylsuccinic anhydride	CC	–37.00	99Dur
	3% W/V NaCl+100 mg Na_2CO_3/ 100 ml H_2O	hexadecylsuccinic anhydride	CC	–31.00	99Dur
	3% W/V NaCl+100 mg Na_2CO_3 100 ml H_2O	octenylsuccinic anhydride	CC	–27.00	99Dur
	3% W/V NaCl+100 mg Na_2CO_3 100 ml H_2O	benzyldimethylhexadecyl-ammonium chloride	CC	–37.00	99Dur
	3% W/V NaCl+100 mg Na_2CO_3 100 ml H_2O	trimethylexadecyl-ammonium chloride	CC	–32.00	99Dur
	3% W/V NaCl+100 mg Na_2CO_3 100 ml H_2O	lauric acid	CC	–34.00	99Dur
	3% W/V NaCl+100 mg Na_2CO_3 100 ml H_2O	sebacic acid	CC	–25.00	99Dur
	3% W/V NaCl+100 mg Na_2CO_3 100 ml H_2O	cetylpyridinium chloride	CC	–35.00	99Dur
	3% W/V NaCl+100 mg Na_2CO_3 100 ml H_2O	cetylamine	CC	–36.00	99Dur
	3% W/V NaCl+100 mg Na_2CO_3 100 ml H_2O	stearylamine	CC	–34.00	99Dur
	3% W/V NaCl+100 mg Na_2CO_3 100 ml H_2O	dimethyldodecyl-amine	CC	–23.00	99Dur
	3% W/V NaCl+100 mg Na_2CO_3 100 ml H_2O	cetyl cyanide	CC	–29.00	99Dur
	3% W/V NaCl+100 mg Na_2CO_3 100 ml H_2O	cetylmercaptan	CC	–31.00	99Dur
	3% W/V NaCl+100 mg Na_2CO_3 100 ml H_2O	dimethyldipalmityl-ammonium chloride	CC	–31.00	99Dur
	3% W/V NaCl+100 mg Na_2CO_3 100 ml H_2O	folic acid	CC	–33.00	99Dur
steel	0.5 **M** H_2SO_4/H_2O	2,5-bis(4-methoxy-phenyl)-1,3,4-oxadiazole	WL	–37.61	06Bou
Lead (Pb)	0.1 **N** Na_2SO_4/H_2O	n-pentanol	T	–18.00	69Gri1
	0.1 **N** Na_2SO_4/H_2O	cyclohexanol	T	–18.74	72Gri2

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
Pb	0.1 **N** Na_2SO_4/H_2O	cyclohexanol	T	–18.76	72Alu
	0.1 **N** Na_2SO_4/H_2O	cyclohexanol	T	–18.8	80Khm
	1 **N** Na_2SO_4/H_2O	phenol	T	–22.59	72Gri2
	0.1 **N** Na_2SO_4/H_2O	aniline	T	–22.2	72Gri2
	0.9 **N** $NaClO_4$+0.1 **N** NaOH/ H_2O	n-decylamine	RT	–25.9	64Boc2
	0.12 **N** KF/H_2O	phenol	T	–22.62	70Alu
Pb (100)	0.1 **N** Na_2SO_4/H_2O	cyclohexanol	T	–18.6	80Khm
Pb (110)	0.1 **N** Na_2SO_4/H_2O	cyclohexanol	T	–18.9	80Khm
Pb (111)	0.1 **N** Na_2SO_4/H_2O	cyclohexanol	T	–18.1	80Khm
Pb (112)	0.1 **N** Na_2SO_4/H_2O	cyclohexanol	T	–18.4	80Khm
Mercury (Hg)	1 **M** KCl+KF/H_2O [30])	Cl^-	T	–81.2	68Pay
	$KNCS/H_2O$	NCS^-	EC	–90.4 [31])	68Min
	$KNCS/H_2O$	NCS^-	EC	–97.1 [32])	68Min
	0.1 **N** NaF/H_2O	toluene	T	–5.34	85Dam
	0.2 **N** NH_4NO_3/methanol	toluene	T	–19.2	85Dam
	0.1 **N** $NaClO_4$/DMF [33])	naphthalene	T	–4.69	70Kag
	0.1 **N** $NaClO_4$/DMF [33])	naphthalene	T	–8.93 [34])	70Kag
	0.1 **N** $NaClO_4$/DMF [33])	diphenyl	T	–6.46 [35])	70Kag
	0.1 **N** $NaClO_4$/DMF [33])	diphenyl	T	–8.8 [34])	70Kag
	0.1 **M** Na_2SO_4/H_2O	methanol	T	–13.8	68Mur3
	0.1 **N** NaF/H_2O	ethanol	T	–7.78	67Dam
	0.4 **N** NaF/H_2O	2-(methylthio)-ethanol	EC	–17.9	79Ike
	0.1 **N** NaF/H_2O	n-propanol	T	–12.76	67Dam
	1 **N** KCl/H_2O	n-propanol	T	–14.6 [36])	65Bar2
	0.1 **N** KF/H_2O	n-propanol	T	–12.75	73Ryb1
	0.1 **M** NaF/H_2O	n-propanol	C	–19.9	93For
	0.5 **M** Na_2SO_4/H_2O	2-propanol	C	–12.5	91Pez
	0.05 **M** Na_2SO_4/H_2O	n-propanol	C	–19.7	93For
	0.5 **M** Na_2SO_4/ H_2O	hexafluoro-2-propanol	EC	–10.7	74Ped
	0.1 **N** NaF/H_2O	n-butanol	T	–14.6	97Dam1
	0.1 **N** NaF/H_2O	n-butanol	T	–14.0 [37])	97Dam1
	0.1 **N** NaF/H_2O	n-butanol	T	–13.1 [38])	97Dam1
	0.05 **M** Na_2SO_4/H_2O	n-butanol	T	–14.7	97Dam1
	0.5 **M** Na_2SO_4/H_2O	n-butanol	T	–15.4	97Dam1
	0.1 **N** NaF/H_2O	n-butanol	T	–16.03	67Dam
	0.1 **N** $HClO_4/H_2O$	n-butanol	EC	–14.5	76Bro
	0.1 **N** $HClO_4/H_2O$	n-butanol	T	–14.0	76Bro
	0.5 **M** Na_2SO_4/H_2O	n-butanol	C	–17.6	90Mon2
	0.1 **M** NaF/H_2O	n-butanol	C	–16.4	93For
	0.05 **M** Na_2SO_4/H_2O	n-butanol	C	–15.9	93For
	1 **N** HCl/H_2O	n-butanol	EC	–10.47	61Blo
	0.5 **M** Na_2SO_4/H_2O	t-butanol	T	–14.0	75Pul1

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
	0.25 **N** NaF/H_2O	t-butanol	EC	–13.0	76Pul
	0.1 **M** Na_2SO_4/H_2O	t-butanol	T	–14.0	76deB1
	0.1 **N** NaF/H_2O	n-pentanol	T	–19.17	67Dam
	0.1 **M** $HClO_4$+0.001 **M** HCl/ H_2O	1-pentanol	T	–11.31	57Han
	0.5 **M** Na_2SO_4/H_2O	n-pentanol	C	–20.4	90Mon2
	0.1 **N** NaF/H_2O	n-pentanol	T	–18.1	97Dam1
	0.1 **N** NaF/H_2O	n-pentanol	T	–17.6 [37])	97Dam1
	0.1 **N** NaF/H_2O	n-pentanol	T	–15.6 [38])	97Dam1
	0.1 **N** Na_2SO_4/H_2O	i-pentanol	T	–17.98	70Dob
	0.1 **N** $HClO_4/H_2O$	i-pentanol	EC	–14.0	76Bro
	0.1 **N** $HClO_4/H_2O$	i-pentanol	T	–13.9	76Bro
	0.5 **M** NaF/H_2O	t-pentanol	T	–18.0	62Dam
	0.5 **M** Na_2SO_4/H_2O	t-pentanol	C	–19.6	91Pez
	0.1 **N** NaF/H_2O	n-hexanol	T	–22.56	67Dam
	0.1 **M** NaF/H_2O	n-hexanol	C	–23.8	93For
	0.05 **M** Na_2SO_4/H_2O	n-hexanol	C	–23.6	93For
	0.5 **M** Na_2SO_4/H_2O	n-hexanol	C	–24.5	90Mon2
	0.1 **N** NaF/H_2O	n-pentanol	T	–21.7	97Dam1
	0.1 **M** $HClO_4/H_2O$	3-pentanol	EC	–18.0	69Bro1
	0.5 **M** Na_2SO_4/H_2O	n-heptanol	C	–28.5	90Mon2
	0.1 **M** NaF/H_2O	n-heptanol	C	–27.3	93For
	0.05 **M** Na_2SO_4/H_2O	n-heptanol	C	–27.0	93For
	0.5 **M** Na_2SO_4/H_2O	n-octanol	C	–31.8	90Mon2
	0.1 **M** NaF/H_2O	n-octanol	C	–30.8	93For
	0.05 **M** Na_2SO_4/H_2O	n-octanol	C	–31.5	93For
	0.1 **N** Na_2SO_4/H_2O	cyclohexanol	T	–19.45	70Dob
	0.1 **N** Na_2SO_4/H_2O	cyclohexanol	T	–19.35	72Alu
	0.1 **M** $HClO_4/H_2O$	cyclohexanol	EC	–20.3	69Bro1
	0.1 **M** KBr/H_2O	1,2-ethanediol	EC	–8.37	70Tra
	0.1 **M** KCl/H_2O	1,2-ethanediol	EC	–8.45	70Tra
	0.1 **M** KJ/H_2O	1,2-ethanediol	EC	–8.03	70Tra
	0.1 **M** NaF/H_2O	1,2-ethanediol	EC	–8.7	70Tra
	0.5 **M** Na_2SO_4/H_2O	1,2-ethanediol	C	–8.8	90Mon1
	0.5 **M** Na_2SO_4/H_2O	1,3-propanediol	C	–15.9	90Mon1
	0.1 **M** Na_2SO_4/H_2O	2-propen-1-ol	T	–24.06	68Mur3
	0.25 **M** Na_2SO_4/H_2O	1,4-butanediol	EC	–15.5	68Dut
	0.25 **M** NaF/H_2O	1,4-butanediol	EC	–15.5	78Pul
	0.25 **M** Na_2SO_4/H_2O	1,4-butanediol	EC	–15.5	78Pul
	0.5 **M** Na_2SO_4/H_2O	1,4-butanediol	C	–17.6	90Mon1
	0.5 **M** Na_2SO_4/H_2O	1,5-pentanediol	C	–21.3	90Mon1
	0.5 **M** Na_2SO_4/H_2O	1,6-hexanediol	EC; C	–23.1	85Gol
	0.5 **M** Na_2SO_4/H_2O	1,6-hexanediol	C	–23.7	90Mon1
	0.5 **M** Na_2SO_4/H_2O	1,7-heptanediol	C	–26.4	90Mon1
	0.1 **M** Na_2SO_4/H_2O	β-cyclodextrine	EC	–34.7	88Bor1

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
Hg	0.7953 **M** NaF/H_2O	sucrose	T	–22.1	81Par
	1 **N** KCl/H_2O	allyl alcohol	T	–13.0 [36])	65Bar2
	1 **N** KCl/H_2O	propargyl alcohol	T	–12.3 [36])	65Bar2
	0.12 **N** KF/H_2O	phenol	T	–28.4	70Alu
	1 **N** Na_2SO_4/H_2O	phenol	C	–28.37	64Dam2
	1 **N** HCl/H_2O	phenol	EC	–10.89	61Blo
	0.1 **M** $HClO_4$+0.001 **M** HCl/H_2O	phenol	T	–21.82	57Han
	0.5 **M** $NaClO_4$/H_2O	phenol	T	–26.0 [39])	86Rol
	0.5 **M** $NaClO_4$/H_2O	phenol	T	–18.5 [40])	86Rol
	0.5 **M** $NaClO_4$/H_2O	phenol	T; EC	–6.3 [41])	83Sar
	0.1 **M** Na_2SO_4/H_2O	phenol	T	–14.64	68Mur3
	0.1 **N** Na_2SO_4/H_2O	pyrocatechol	T	–31.07	69Dya
	0.5 **M** $NaClO_4$/H_2O	pyrocatechol	T; EC	–6.3 [41])	83Sar
	0.1 **N** Na_2SO_4/H_2O	resorcinol	T	–30.8	69Dya
	0.1 **N** Na_2SO_4/H_2O	hydroquinone	T	–28.34	69Dya
	1 **N** HCl/H_2O	α-naphthol	EC	–27.22	61Blo
	0.5 **M** $NaClO_4$	phloroglucinol	EC	–25.5 [41])	79Sar
	0.5 **M** $NaClO_4$/H_2O	phloroglucinol	T; EC	–20.1 [41])	83Sar
	0.1 **M** $HClO_4$+0.001 **M** HCl/H_2O	3-pentanone	T	–12.10	57Han
	1 **N** HCl/H_2O	n-pentanone	EC	–27.6	61Blo
	0.1 **M** $HClO_4$+0.001 **M** HCl/H_2O	2,4-pentanedione	T	–12.19	57Han
	0.1 **M** $HClO_4$/H_2O	cyclohexanone	EC	–21.10	69Bro1
	1 **N** Na_2SO_4/H_2O	methylpropylketone	T	–19.6	73Gus
	1 **N** Na_2SO_4/H_2O	methylbutylketone	T	–22.3	73Gus
	1 **N** HCl/H_2O	di-n-butylketone	EC	–25.55	61Blo
	1 **N** Na_2SO_4/H_2O	methylpentylketone	T	–25.1	73Gus
	1 **N** Na_2SO_4/H_2O	methylhexylketone	T	–27.8	73Gus
	1 **N** KBr/H_2O	dioxane	T	–13.8	63Wat
	1 **N** KCl/H_2O	dioxane	T	–12.5 [36])	65Bar2
	1 **N** KF/H_2O	dioxane	T	–13.8	63Wat
	1 **N** KI/H_2O	dioxane	T	–10.9	63Wat
	1 **N** NaCl/H_2O	dioxane	T	–12.6	63Wat
	0.1 **M** Na_2SO_4/H_2O	dioxane	T	–15.5	68Mur3
	1 **N** HCl/H_2O	benzaldehyde	EC	–18.85	61Blo
	1 **N** HCl/H_2O	α-naphthaldehyde	EC	–28.90	61Blo
	1 **M** KCl/H_2O	cinnamaldehyde	T	–34.7	84Gon
	0.8 **N** Na_2SO_4+0.2 **N** H_2SO_4/H_2O	acetic acid	T; EC	–12.6	70Dam
	0.4 **M** Na_2SO_4+0.1 **M** H_2SO_4/H_2O	acetic acid	T	–13.2	97Dam2
	0.8 **N** Na_2SO_4+0.2 **N** H_2SO_4/H_2O	propionic acid	T; EC	–15.9	70Dam
	0.4 **M** Na_2SO_4+0.1 **M** H_2SO_4/H_2O	propionic acid	T	–14.8	97Dam2

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
	0.8 **N** Na_2SO_4+0.2 **N** H_2SO_4/ H_2O	n-butyric acid	EC	–15.21	69Chi
	0.8 **N** Na_2SO_4+0.2 **N** H_2SO_4/ H_2O	n-butyric acid	T	–15.21	69Chi
	0.8 **N** Na_2SO_4+0.2 **N** H_2SO_4/ H_2O	n-butyric acid	T; EC	–19.1	70Dam
	0.4 **M** Na_2SO_4+0.1 **M** H_2SO_4/ H_2O	n-butyric acid	T	–18.5	97Dam2
	0.1 **M** $HClO_4$+0.001 **M** HCl/ H_2O	n-pentanoic acid	T	–10.80	57Han
	1 **N** HCl/H_2O	n-pentanoic acid	EC	–12.14	61Blo
	0.8 **N** Na_2SO_4+0.2 **N** H_2SO_4/ H_2O	n-pentanoic acid	T; EC	–23.1	70Dam
	0.8 **N** Na_2SO_4+0.2 **N** H_2SO_4/ H_2O	α-iso-valeric acid	T; EC	–18.0	70Dam
	0.8 **N** Na_2SO_4+0.2 **N** H_2SO_4/ H_2O	β-iso-valeric acid	T; EC	–18.6	70Dam
	0.8 **N** Na_2SO_4+0.2 **N** H_2SO_4/ H_2O	trimethyl acetic acid	T; EC	–17.0	70Dam
	0.1 **N** $HClO_4$/H_2O	n-pentanoic acid	EC	–15.3	76Bro
	0.1 **N** $HClO_4$/H_2O	n-pentanoic acid	T	–15.3	76Bro
	0.4 **M** Na_2SO_4+0.1 **M** H_2SO_4/ H_2O	n-pentanoic acid	T	–21.8	97Dam2
	0.8 **N** Na_2SO_4+0.2 **N** H_2SO_4/ H_2O	n-hexanoic acid	T; EC	–26.6	70Dam
	0.4 **M** Na_2SO_4+0.1 **M** H_2SO_4	n-hexanoic acid	T	–23.5	97Dam2
	0.1 **M** $HClO_4$+0.001 **M** HCl/ H_2O	octanoic acid	T	–12.14	57Han
	0.1 **M** $NaClO_4$+0.05 **M** $HClO_4$/ H_2O	maleic acid	EC	–12.8	69Pos
	1 **N** HCl/H_2O	benzoic acid	EC	–19.27	61Blo
	1 **N** H_2SO_4/H_2O	benzoic acid	EC; T	–26.0	77Tan
	0.1 **m** Na_2SO_4/H_2O	o-hydroxybenzoic acid	EC	–29.9	72Jos
	0.1 **m** Na_2SO_4/H_2O	p-hydroxybenzoic acid	EC	–26.0	72Jos
	0.05 **N** Na_2SO_4+0.05 **N** H_2SO_4/ H_2O	o-hydroxybenzoic acid	EC; T	–38.62	81Dob
	0.05 **N** Na_2SO_4+0.05 **N** H_2SO_4/ H_2O	p-hydroxybenzoic acid	EC; T	–39.08	81Dob
	1 **N** HCl/H_2O	α-naphthoic acid	EC	–32.25	61Blo
	0.5 **M** NaF/H_2O+0.01 **M** Na_2HPO_4	D-ribose	EC	–6.93	78Bra1
	0.5 **M** NaF/H_2O+0.01 **M** Na_2HPO_4	2'-deoxy-D-ribose	EC	–7.51	78Bra1
	0.5 **M** NaF/H_2O+0.01 **M** Na_2HPO_4	D-ribose 5'-phosphate	EC	–3.32	78Bra1

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
Hg	1 **N** KNO_3/H_2O	benzofurazan	T	–27.2	71Tsv
	1 **N** Na_2SO_4/H_2O	butylamine	EC	–20.5	65Zwi
	0.5 **M** Na_2SO_4/H_2O	t-butylamine	EC	–16.5	83Bor
	1 **N** Na_2SO_4/H_2O	*N*-methylaniline	EC	–23.84	65Zwi
	0.25 **M** NaF/H_2O	acetonitrile	EC; T	–9.1	73DeB
	0.25 **M** NaF/H_2O	propionitrile	EC; T	–10.6	75Abd
	0.25 **M** NaF/H_2O	butyronitrile	EC; T	–13.8	79Ama
	0.1 **M** Na_2SO_4/H_2O	butyronitrile	EC; T	–13.8	79Ama
	0.1 **M** $HClO_4$+0.001 **M** HCl/ H_2O	pentane nitrile	T	–10.38	57Han
	0.25 **M** NaF/H_2O	succinonitrile	EC; T	–12.6	76deB2
	1 **N** HCl/H_2O	benzonitrile	EC	–19.68	61Blo
	1 **N** HCl/H_2O	n-pentanoic acid	EC	–12.14	61Blo
	1 **N** HCl/H_2O	n-butylammonium chloride	EC	–4.2... –10.05	61Blo
	1 **N** HCl/H_2O	benzylammonium chloride	EC	–4.2... –10.05	61Blo
	1 **N** HCl/H_2O	α-naphthyl-ammonium chloride	EC	–4.2... –10.05	61Blo
	1 **N** HCl/H_2O	n-thiobutanol	EC	–23.04	61Blo
	1 **N** HCl/H_2O	thiophenol	EC	–25.55	61Blo
	0.1 **N** NaF/H_2O	thiourea	T	–95.72	61Par
	0.1 **N** KNO_3/H_2O	thiourea	T	–94.13	68Par
	0.1 **N** $NaClO_4$/DMF [33])	thiourea	T	–90.9	70Gan
	1 **N** KNO_3/H_2O	piazthiole	T	–33.5	71Tsv
	1 **N** HCl/H_2O	di-n-butylsulfide	EC	–30.99	61Blo
	1 **N** HCl/H_2O	dibenzylsulfide	EC	–36.44	61Blo
	1 **N** HCl/H_2O	n-butyl sulfonic acid	EC	–4.2... –10.05	61Blo
	1 **N** HCl/H_2O	benzylsulfonic acid	EC	–4.2... –10.05	61Blo
	1 **N** HCl/H_2O	α-naphthyl sulfonic acid	EC	–6.70	61Blo
	1 **N** KCl/H_2O	aniline	C	–22.22	64Dam1
	1 **N** KI/H_2O	aniline	C	–18.6	64Dam1
	0.1 **N** NaF/H_2O	aniline	T	–23.54 [42])	85Dam
	1 **N** HCl/H_2O	aniline	EC	–17.15 [36])	61Bar2
	1 **N** KCl/H_2O	aniline	EC	–18.82 [36])	61Con
	0.1 **M** NaH_2PO_4/H_2O	acetanilide	EC	–102.1	66Par1
	1 **M** NaH_2PO_4/H_2O	acetanilide	EC	–103.8	66Par1
	1 **N** HCl/H_2O	pyridine	EC	–10.46 [36])	61Bar2
	1 **N** KCl/H_2O	pyridine	EC	–13.8 [36])	61Con
	1 **N** KCl/H_2O	pyridine	EC	–17.7	65Con
	1 **N** KCl/H_2O	pyridine	T	–18.6 [36])	65Bar2
	0.03 **N** $NaClO_4/H_2O$	pyridine	EC	–13.8	74Con1

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
	0.03 **N** $NaClO_4$+0.01 **M** NaOH/ H_2O	pyridine	EC	–13.8	74Con1
	1 **N** HCl/H_2O	1,2,3,6-tetrahydro-pyridine	EC	–15.48 [36])	61Bar2
	1 **N** KCl/H_2O	1,2,3,6-tetrahydro-pyridine	EC	–18.82 [36])	61Con
	1 **N** HCl/H_2O	2-Cl-pyridine	EC	–14.64 [36])	61Bar2
	1 **N** KCl/H_2O	2-Cl-pyridine	EC	–22.17 [36])	61Con
	1 **N** HCl/H_2O	piperidine	EC	–15.06 [36])	61Bar2
	1 **N** KCl/H_2O	piperidine	EC	–19.24 [36])	61Con
	1 **N** KCl/H_2O	piperidine	T	–18.8 [36])	65Bar2
	0.03 **N** $NaClO_4$+0.01 **M** NaOH/ H_2O	pyrazine	EC	–17.99 [43])	73Con
	0.03 **N** $NaClO_4/H_2O$	pyrazine	EC	–11.4	74Con1
	0.03 **N** $NaClO_4$+0.01 **M** NaOH/ H_2O	pyrazine	EC	–11.4	74Con1
	0.1 **N** KF/H_2O	nicotinic acid	T	–100.8	92Jur
	1 **N** HCl/H_2O	2-NH_2-pyridine	EC	–18.82 [36])	61Bar2
	1 **N** KCl/H_2O	2-NH_2-pyridine	EC	–16.73 [36])	61Con
	1 **N** KCl/H_2O	morpholine	T	–14.6 [36])	65Bar2
	1 **N** KNO_3/H_2O	piazselenole	T	–36.1	71Tsv
	0.916 **M** NaF/H_2O	sodium tetraphenyl-borate	T	–41.38	72Gio
	1 **M** $LiClO_4/H_2O$	cis-1,2-dibromo-ethylene	T	–21.5	69Mar
	1 **M** $LiClO_4/H_2O$	trans-1,2-dibromo-ethylene	T	–21.03	69Mar
	0.1 **M** NaF/H_2O	tyrosine	T	–31.3	97Jur
	0.5 **M** NaF/H_2O+0.01 **M** Na_2HPO_4	adenine	EC	–18.46	78Bra2
	0.5 **M** borate buffer/H_2O	adenine	T	–18.5	77Kin
	0.5 **M** NaF/H_2O+0.01 **M** Na_2HPO_4	adenosine	EC	–24.4	78Bra2
	0.5 **M** borate buffer/H_2O	deoxyadenosine	T	–23.8	77Kin
	0.5 **M** borate buffer/H_2O	deoxyadenosine-5'-monophosphate	T	–20.7	77Kin
	0.5 **M** borate buffer/H_2O	deoxyadenosine-5'-diphosphate	T	–20.6	77Kin
	0.5 **M** borate buffer/H_2O	deoxyadenosine-5'-triphosphate	T	–23.0	77Kin
	0.5 **M** NaF/H_2O+0.01 **M** Na_2HPO_4	adenosine 5'-monophosphate	EC	–20.44	78Bra2
	0.5 **M** NaF/H_2O+0.01 **M** Na_2HPO_4	uracil	EC	–12.53	78Bra2
	0.5 **M** NaF/H_2O+0.01 **M** Na_2HPO_4	uridine	EC	–17.47	78Bra2

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG^0_{ad} [4])[5]) [kJ mol^{-1}]	Ref.
Hg	0.5 **M** NaF/H_2O+0.01 **M** Na_2HPO_4	uridine 5'-monophosphate	EC	–16.16	78Bra2
	0.5 **M** NaF/H_2O, pH = 7	uracil	T	–23.7	91Wan
	0.5 **M** NaF/H_2O, pH = 7	2-thiouracil	T	–32.8	91Wan
	0.5 **M** NaF/H_2O, pH = 7	5-ethyl-2-thiouracil	T	–40.6	91Wan
	0.1 **M** $HClO_4/H_2O$	chloroform	EC	–14.2	69Bro1
Nickel (Ni)	0.9 **N** $NaClO_4$+0.1 **N** NaOH/H_2O	n-decylamine	RT	–28.42	64Boc2
	1 **N** $NaClO_4/H_2O$	naphthalene	RT	–25.1	64Boc1
Palladium (Pd)	0.9 **N** $NaClO_4$+0.1 **N** NaOH/H_2O	n-decylamine	RT	–25.91	64Boc2
Platinum (Pt)	1 **N** H_2SO_4/H_2O	ethylene	RT	–20.48	65Gil
	1 **N** H_3PO_4/H_2O	benzene	RT	–23.87	66Hei
	1 **N** H_3PO_4/H_2O	benzene	RT	–32.18 [44])	66Hei
	1 **N** H_2SO_4/H_2O	benzene	RT	–27.22	66Hei
	1 **N** H_2SO_4/H_2O	benzene	RT	–36.36 [44])	66Hei
	1 **N** $NaClO_4/H_2O$	naphthalene	RT	–35.1	64Boc1
	0.01 **M** HCl/H_2O	n-butyl alcohol	EL; IR	–34.7	92Boc
	0.01 **M** HCl/H_2O	phenol	EL; IR	–40.2	92Boc
	0.01 **M** HCl/H_2O	1-naphthol	EL; IR	–56.1	92Boc
	0.01 **M** HCl/H_2O	1-naphthaldehyde	EL; IR	–54.0	92Boc
	0.01 **M** HCl/H_2O	n-valeraldehyde	EL; IR	–32.6	92Boc
	0.01 **M** HCl/H_2O	n-valeric alcohol	EL; IR	–35.6	92Boc
	0.01 **M** HCl/H_2O	benzoylchloride	EL; IR	–39.3	92Boc
	0.01 **M** HCl/H_2O	n-valeroylchloride	EL; IR	–33.9	92Boc
	0.01 **M** HCl/H_2O	1-naphthoyl chloride	EL; IR	–59.4	92Boc
	0.01 **M** HCl/H_2O	n-butylamine	EL; IR	–33.5	92Boc
	0.9 **N** $NaClO_4$+0.1 **N** NaOH/H_2O	n-decylamine	RT	–30.93	64Boc2
	0.01 **M** HCl/H_2O	aniline	EL; IR	–42.3	92Boc
	0.01 **M** HCl/H_2O	1-naphthamine	EL; IR	–54.8	92Boc
	0.01 **M** HCl/H_2O	benzaldehyde	EL; IR	–38.5	92Boc
	0.01 **M** HCl/H_2O	benzoic acid	EL; IR	–43.5	92Boc
	0.1 **M** $HClO_4/H_2O$	benzoic acid	RT	–49.5	95Sob
	0.01 **M** HCl/H_2O	1-naphthoic acid	EL; IR	–61.1	92Boc
	0.01 **M** HCl/H_2O	benzonitrile	EL; IR	–43.9	92Boc
	0.01 **M** HCl/H_2O	n-valeronitrile	EL; IR	–38.1	92Boc
	0.01 **M** HCl/H_2O	1-butylmercaptane	EL; IR	–36.0	92Boc
	0.01 **M** HCl/H_2O	thiophenol	EL; IR	–41.0	92Boc
Silver (Ag)	0.1 **M** NaF/H_2O	Cl^-	T	–86.0	82Lar
	0.5 **M** NaF/H_2O	Cl^-	T	–83.0	82Lar
	0.5 **M** $NaClO_4/H_2O$	Cl^-	T	–95.0	83Hup
	0.5 **M** NaF/H_2O	Br^-	T	–96.0	82Lar
	0.5 **M** $NaClO_4/H_2O$	Br^-	T	–101.0	83Hup
	0.5 **M** $NaClO_4/H_2O$	I^-	T	–95.0	83Hup

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
Ag	0.5 **M** NaF/H_2O	N_3^-	T	–93.0	82Lar
	0.5 **M** Na_2SO_4/H_2O	CN^-	T	–23.0 [45])	97Rog
	0.1 **M** $LiClO_4/H_2O$	4-phenylpyridine	T	–40.0	02Buk
	0.1 **M** KF/H_2O	thiourea	T	–85.6	88Mil
Ag (100)	0.04 **M** $(NaF+NaCl)/H_2O$	Cl^-	T	–100.0	78Val
	0.1 **N** NaF/H_2O	1-butanol	T	–21.4	02Jur
	0.1 **N** NaF/H_2O	2-butanol	T	–23.0	02Jur
	0.1 **N** Na_2SO_4/H_2O	isobutanol	T	–20.5	76Vit3
	0.1 **M** KF/H_2O	amyl alcohol	T	–15.8	88Pop
	0.05 **M** KPF_6/H_2O	*n*-pentanol	T; C	–14.1	01Dou
Ag (110)	0.04 **M** $(NaF+NaCl)/H_2O$	Cl^-	T	–93.0	78Val
	0.1 **N** NaF/H_2O	1-butanol	T	–20.6	02Jur
	0.1 **N** NaF/H_2O	2-butanol	T	–22.7	02Jur
	0.1 **N** $LiClO_4F/H_2O$	1-butanol	T	–21.0	02Jur
	0.1 **N** $LiClO_4/H_2O$	2-butanol	T	–18.4	02Jur
	0.1 **N** Na_2SO_4/H_2O	n-hexanol	T	–23.0	74Vit3
	0.05 **M** KPF_6/H_2O	*n*-pentanol	T; C	–12.2	01Dou
	0.1 **M** $LiClO_4/H_2O$	tyrosine	T	–28.1	98Mil
Ag (111)	0.04 **M** $(NaF+NaCl)/H_2O$	Cl^-	T	–100.0	78Val
	0.1 **N** Na_2SO_4/H_2O	isobutanol	T	–20.9	76Vit3
	0.1 **M** KF/H_2O	amyl alcohol	T	–15.3	88Pop
	0.05 **M** $KClO_4/H_2O$	amyl alcohol	T	–16.4	94Dou
	0.1 **N** Na_2SO_4/H_2O	*n*-hexanol	T	–24.7	74Vit3
	0.05 **M** $KClO_4/H_2O$	uracil	C	–11.38	96Wan2
	0.05 **M** $KClO_4/H_2O$	5-methyluracil	C	–14.65	96Wan2
	0.05 **M** $KClO_4/H_2O$	5,6-dimethyluracil	C	–21.1	96Wan2
	0.1 **M** $LiClO_4/H_2O$	tyrosine	T	–28.1	98Mil
	0.5 **M** $NaOH/H_2O$	thioethanol	CV	–78.0	98Hat [46])
	0.5 **M** $NaOH/H_2O$	*n*-thiopropanol	CV	–82.3	98Hat [46])
	0.5 **M** $NaOH/H_2O$	*n*-thiobutanol	CV	–86.5	98Hat [46])
	0.5 **M** $NaOH/H_2O$	*n*-thiopentanol	CV	–90.8	98Hat [46])
	0.5 **M** $NaOH/H_2O$	*n*-thiohexanol	CV	–95.0	98Hat [46])
	0.5 **M** $NaOH/H_2O$	*n*-thioheptanol	CV	–99.3	98Hat [46])
	0.5 **M** $NaOH/H_2O$	*n*-thiooctanol	CV	–103.6	98Hat [46])
	0.5 **M** $NaOH/H_2O$	*n*-thiodecanol	CV	–112.0	98Hat [46])
	0.5 **M** NaOH/methanol+5% H_2O	thioethanol	CV	–87.0	98Hat [46])
	0.5 **M** NaOH/methanol+5% H_2O	*n*-thiopropanol	CV	–89.1	98Hat [46])
	0.5 **M** NaOH/methanol+5% H_2O	*n*-thiobutanol	CV	–91.2	98Hat [46])
	0.5 **M** NaOH/methanol+5% H_2O	*n*-thiopentanol	CV	–94.5	98Hat [46])
	0.5 **M** NaOH/methanol+5% H_2O	*n*-thiohexanol	CV	–95.0	98Hat [46])
	0.5 **M** NaOH/methanol+5% H_2O	*n*-thioheptanol	CV	–96.6	98Hat [46])
	0.5 **M** NaOH/methanol+5% H_2O	*n*-thiodecanol	CV	–102.5	98Hat [46])
	0.5 **M** NaOH/methanol+5% H_2O	*n*-thiododecanol	CV	–106.7	98Hat [46])
	0.5 **M** NaOH/methanol+5% H_2O	*n*-thiohexadecanol	CV	–110.0	98Hat [46])
Tin	0.1 **N** Na_2SO_4/H_2O	ethanol	C	–7.36	73Gri
(Sn)	0.1 **N** Na_2SO_4/H_2O	*n*-propanol	C	–11.88	73Gri

Electrode	Solution [1])	Adsorbate [2])	Met. [3])	ΔG_{ad}^{0} [4])[5]) [kJ mol^{-1}]	Ref.
Sn	0.1 **N** Na_2SO_4/H_2O	*n*-butanol	C	–15.4	73Gri
	0.1 **N** Na_2SO_4/H_2O	*n*-pentanol	C	–17.86	73Gri
Zinc (Zn)					
Zn (0001)	0.1 **N** KCl+0.0002 **N** H_2SO_4/H_2O	*n*-propanol	T	–9.94	76Ipa2
	0.1 **N** KCl+0.0002 **N** H_2SO_4/H_2O	*n*-butanol	T	–12.51	76Ipa2
	0.1 **N** KCl+0.0002 **N** H_2SO_4/H_2O	*n*-pentanol	T	–15.42	76Ipa2
	0.1 **N** KCl+0.0002 **N** H_2SO_4/H_2O	cyclohexanol	T	–15.9	75Ipa1
	0.1 **N** KCl+0.0002 **N** H_2SO_4/H_2O	cyclohexanol	T	–15.9	75Ipa2
Zn (10$\bar{1}$0)	0.1 **N** KCl+0.0002 **N** H_2SO_4/H_2O	cyclohexanol	T	–15.1	75Ipa2
Zn (10$\bar{2}$0)	0.1 **N** KCl+0.0002 **N** H_2SO_4/H_2O	cyclohexanol	T	–14.7	75Ipa2

[1]) If no concentration is given the reported value of ΔG_{ad}^{0} did not depend on the concentration of the supporting electrolyte; concentrations are given as molar values (**M**) or molal (**m**) values. If the original report contains concentration numbers in units of **N** this is maintained. In case of mixed electrolytes (this is mostly used to maintain constant ionic strength) the electrolyte containing the species to be asdorbed is mentioned first.

[2]) Adsorbates are arranged alphabetically starting with inorganic adsorbates; within a given adsorbate data are arranged with increasing concentration of the supporting electrolyte.

[3]) Met. = Method. C = chronocoulometry, CC = corrosion current method, EC = Electrocapillary method, EL = Ellipsometric method, IR = Infrared spectroscopic method, CV = cyclic voltammetry, RT = Radiotracer method, T = Tensammetry, TF = Thin film resistance, WL = weight loss method. When a combination of methods was used all methods are listed separated by slashes. See also list of symbols.

[4]) If values of ΔG_{ad}^{0} as a function of electrode potential are reported only maximum values are listed except for data reported for specific adsorbate orientations.

[5]) Standard state: unit mole fraction in solution and unit coverage in monolayer.

[6]) T = –15 °C.

[7]) T = –0 °C.

[8]) T = 12 °C.

[9]) T = 25 °C.

[10]) T = 40 °C.

[11]) T = 50 °C.

[12]) Obtained at σ = 0 C cm^{-2}.

[13]) Obtained at $E = E_{pzc}$.

[14]) Obtained with assuming a Frumkin isotherm.

[15]) Obtained with assuming a Henry isotherm.

[16]) The actual value varies slightly depending on the way the glassy carbon electrode was pretreated.

[17]) Extrapolated to zero coverage.

[18]) At E_{RHE} = –0.2 V.

[19]) At E_{RHE} = –0.2 V.

[20]) Gallium saturated with bismuth.

[21]) Gallium saturated with lead.

[22]) Gallium saturated with tin.

[23]) The concentration of the supporting electrolyte and of the adsorbate were adjusted to keep constant ionic strength.

[24]) At E_{RHE} = 0.3 V.

[25]) At E_{RHE} = 1.3 V.
[26]) Value calculated using Frumkin's isotherm.
[27]) Value calculated using Henry's isotherm.
[28]) Value calculated using Langmuir's isotherm.
[29]) Concentration of $NaClO_4$ and benzensulfonic acid kept constant at 0.1 **M**.
[30]) Variable concentration of chloride, total concentration kept at constant ionic strength.
[31]) T = 293 K.
[32]) T = 313 K.
[33]) DMF = dimethylformamide.
[34]) Flat adsorbate orientation.
[35]) Inclined adsorbate orientation.
[36]) θ = 0.5.
[37]) T = 50 °C.
[38]) T = 75 °C.
[39]) Flat orientation.
[40]) Perpendicular orientation.
[41]) At E_{pzc}.
[42]) E_{NHE} = 0.241 V.
[43]) θ = 0.17.
[44]) T = 70 °C.
[45]) At E_{SHE} = –0.7 V.
[46]) θ = 1.0, $c_{calc.}$ = 5 m**M.**

References

57Han Hansen, R.S., Minturn, R.E., Hickson, D.A.: J. Phys. Chem. **61** (1957) 953.
61Bar2 Barradas, R.G., Conway, B.E.: Electrochim. Acta **5** (1961) 349.
61Blo Blomgren, E., Bockris, J.O'M., Jesch, C.: J. Phys. Chem. **65** (1961) 2000.
61Con Conway, B.E., Barradas, R.G.: Electrochim. Acta **5** (1961) 319.
61Par Parsons, R.: Proc. R. Soc. London, Ser. A **261** (1961) 79.
62Dam Damaskin, B.B., Grigorjev, N.B.: Dokl. Akad. Nauk SSSR **147** (1962) 135.
63Dah Dahms, H., Green, M.: J. Electrochem. Soc. **110** (1963) 1075.
63Wat Watanabe, A., Tsuji, F., Ueda, S.: Kolloid Z. Z. Polym. **191** (1963) 147.
64Boc1 Bockris, J.O'M., Green, M., Swinkels, D.A.J.: J. Electrochem. Soc. **111** (1964) 743.
64Boc2 Bockris, J.O'M., Swinkels, D.A.J.: J. Electrochem. Soc. **111** (1964) 736.
64Dam1 Damaskin, B.B., Mishutushkina, I.P., Gerovich, V.M., Kaganovich, R.I.: Zh. Fiz. Khim. **38** (1964) 1797.
64Dam2 Damaskin, B.B., Gerovich, V.M., Gladkikh, I.P., Kaganovich, R.I.: Zh. Fiz. Khim. **38** (1964) 2495.
65Bar2 Barradas, R.G., Hamilton, P.G.: Can. J. Chem. **43** (1965) 2468.
65Con Conway, B.E., Barradas, R.G., Hamilton, P.G., Parry, J.M.: J. Electroanal. Chem. **10** (1965) 485.
65Gil Gileadi, E., Rubin, B.T., Bockris, J.O'M.: J. Phys. Chem. **69** (1965) 3335.
65Zwi Zwierzykowska, I.: Rocz. Chem. **39** (1965) 101.
66Hei Heiland, W., Gileadi, E., Bockris, J.O'M.: J. Phys. Chem. **70** (1966) 1207.
66Par1 Parsons, R., Zobel, F.G.R.: Trans. Faraday Soc. **62** (1966) 3511.
67Dam Damaskin, B.B., Survila, A.A., Rybalka, L.E.: Elektrokhimiya **3** (1967) 146.
68Dut Dutkiewicz, E., Garnish, J.T., Parsons, R.: J. Electroanal. Chem. **16** (1968) 505.
68Min Minc, S., Andrejczak, J.: J. Electroanal. Chem. **17** (1968) 101.
68Mur3 Murti, M.V.R., Singh, N.: Allg. Prakt. Chem. **19** (1968) 365.
68Par Parsons, R.: Trans. Faraday Soc. **64** (1968) 1077.
68Pay Payne, R.: Trans. Faraday Soc. **64** (1968) 1638.
69Bro1 Broadhead, D.E., Hansen, R.S., Potter jr., G.W.: J. Colloid Interface Sci. **31** (1969) 61.
69Chi Chizhov, A.V., Damaskin, B.B.: Elektrokhimiya **5** (1969) 1012.
69Dya Dyatkina, S.L., Damaskin, B.B.: Elektrokhimiya **5** (1969) 438.
69Gri1 Grigor'ev, N.B., Macavariani, D.N.: Elektrokhimiya **5** (1969) 87.
69Mar Markova, I.G., Feoktistov, L.G.: Elektrokhimiya **5** (1969) 1095.
69Pos Pospisil, L., Kuta, J.: Collect. Czech. Chem. Commun. **34** (1969) 3047.
69Pul Pullerits, R.J., Palm, U.V., Past, V.E.: Elektrokhimiya **5** (1969) 886.
70Alu Alumaa, A.R., Palm, U.V.: Elektrokhimiya **6** (1970) 580.
70Dam Damaskin, B., Frumkin, A., Chizhov, A.: J. Electroanal. Chem. **28** (1970) 93.
70Dob Dobren'kov, G.A., Schilotkach, G.D.: Elektrokhimiya **6** (1970) 1416.
70Gan Ganzhina, I.M., Damaskin, B.B.: Elektrokhimiya **6** (1970) 1715.
70Kag Kaganovich, R.I., Damaskin, B.B., Kaisheva, M.K.: Elektrokhimiya **6** (1970) 1359.
70Tra Trasatti, S.: J. Electroanal. Chem. **28** (1970) 257.
71Alu Alumaa, A., Palm, U.: Uch. Zap. Tartu. Gos. Univ. **289** (1971) 41.
71Kat Katoh, K., Schmid, G.M.: Bull. Chem. Soc. Jpn. **44** (1971) 2007.
71Lou Loutfy, R.O., Sukava, A.J.: J. Electrochem. Soc. **118** (1971) 216.
71Tsv Tsveniashvili, V.Sh., Zhdanov, S.I., Todres, Z.V.: Elektrokhimiya **7** (1971) 28.
72Alu Alumaa, A.R., Palm, U.V.: Elektrokhimiya **8** (1972) 790.
72Gio Giomini, C., Rampazzo, L.: J. Phys. Chem. **76** (1972) 707.
72Gri2 Grigor'ev, N.B., Machavariani, D.N.: Elektrokhimiya **8** (1972) 406.
72Jos Joshi, K.M., Rajagopalan, S.: Electrochim. Acta **17** (1972) 479.
72Pan2 Panin, V.A., Rybalka, K.V.: Elektrokhimiya **8** (1972) 1202.
72Pet1 Petyarv, E.K., Kolk, K.A., Palm, U.V.: Elektrokhimiya **8** (1972) 100.

72Pet2 Petyarv, E.K., Kolk, K.A., Palm, U.V.: Elektrokhimiya **8** (1972) 104.
72Pul Pullerits, R., Moldau, M., Palm, U., Past, B.: Uch. Zap. Tartu. Gos. Univ. **302** (1972) 41.
73Con Conway, B.E., Dhar, H.P., Gottesfeld, S.: J. Colloid Interface Sci. **43** (1973) 303.
73DeB De Battisti, A., Trasatti, S.: J. Electroanal. Chem. **48** (1973) 213.
73Gri Grigor'ev, N.B., Kuprin, V.P., Loshkarev, Yu.M.: Elektrokhimiya **9** (1973) 1842.
73Gus Gusakova, O.Yu., Damaskin, B.B., Fedorovitch, N.V.: Elektrokhimiya **9** (1973) 1134.
73Pal Palm, Yu.V., Palm, V.E., Erlich, Yu.I., Erlich, T.E.: Elektrokhimiya **9** (1973) 1399.
73Ryb1 Rybalka, L.E., Damaskin, B.B., Leikis, D.I.: Elektrokhimiya **9** (1973) 414.
73Ryb2 Rybalka, L.E., Damaskin, B.B.: Elektrokhimiya **9** (1973) 1562.
74Con1 Conway, B.E., Mathieson, J.G., Dhar, H.P.: J. Phys. Chem. **78** (1974) 1226.
74Erl Erlich, Yu.I., Pal'm, U.V.: Elektrokhimiya **10** (1974) 1866.
74Gri Grigor'ev, N.B., Kalyushnaja, A.M.: Elektrokhimiya **10** (1974) 1287.
74Ped Pedriali, R., Kuta, J.: J. Electroanal. Chem. **51** (1974) 395.
74Vit3 Vitanov, T., Popov, A.: Elektrokhimiya **10** (1974) 1373.
75Abd Abd-El-Nabey, B.A., Trasatti, S.: J. Chem. Soc. Faraday Trans. 1 **71** (1975) 1230.
75Ipa1 Ipatov, Yu.P., Batrakov, B.B.: Elektrokhimiya **11** (1975) 1282.
75Ipa2 Ipatov, Yu.P., Batrakov, B.B.: Elektrokhimiya **11** (1975) 1717.
75Pul1 Pulidori, F., Borghesani, G., Pedriali, R., Bighi, C.: Ann. Chim. (Rome) **65** (1975) 581.
75Pul2 Pullerits, R.Ya., Moldau, M.E., Palm, U.V.: Elektrokhimiya **11** (1975) 487.
75Ryb1 Rybalka, D.E., Damaskin, B.B., Leijkis, D.I.: Elektrokhimiya **11** (1975) 9.
76Bro Broadhead, D.E., Balkerikar, K.G., Hansen, R.S.: J. Phys. Chem. **80** (1976) 370.
76Ipa2 Ipatov, Y.P., Batrakov, V.V., Shalaginov, V.V.: Elektrokhimiya **12** (1976) 286.
76Pul Pulidori, F., Borghesani, G., Pedriali, R., Bighi, C.: J. Electroanal. Chem. **72** (1976) 65.
76Vit3 Vitanov, T., Popov, A.: Elektrokhimiya **12** (1976) 319.
76deB1 de Battisti, A., Abd-El-Nabey, B., Trasatti, S.: J. Chem. Soc. Faraday Trans. 1 **72** (1976) 2076.
76deB2 de Battisti, A., Trasatti, S.: J. Electroanal. Chem. **73** (1976) 327.
77Kin Kinoshita, H., Christian, S.D., Dryhurst, G.: J. Electroanal. Chem. **83** (1977) 151.
77Tan Taniguchi, I., Machida, K., Sato, N., Sekine, T.: Bull. Chem. Soc. Jpn. **50** (1977) 1245.
78Bra1 Brabec, V., Christian, S.D., Dryhurst, G.: J. Electrochem. Soc. **125** (1978) 1236.
78Bra2 Brabec, V., Christian, S.D., Dryhurst, G.: Biophys. Chem. **7** (1978) 253.
78Pal Pal'm, U.V., Pyarnoya, M.P.: Elektrokhimiya **14** (1978) 1070.
78Pul Pulidori, F., Borghesani, G., Pedriali, R., De Battisti, A., Trasatti, S.: J. Chem. Soc. Faraday Trans. 1 **74** (1978) 79.
78Pya Pyarnoya, M.P., Pal'm, U.V.: Elektrokhimiya **14** (1978) 1229.
78Val Valette, G., Hamelin, A., Parsons, R.: Z. Phys. Chem. N. F. **113** (1978) 71.
79Ama Amadelli, R., Daghetti, A., Vergano, L., De Battisti, A., Trasatti, S.: J. Electroanal. Chem. **100** (1979) 379.
79Ike Ikeda, O., Kogo, K., Tamura, H.: J. Electroanal. Chem. **95** (1979) 177.
79Pal Paltusova, N., Alumaa, A., Palm, U.: Elektrokhimiya **15** (1979) 1870.
79Sar Sarangapani, S., Venkatesan, V.K.: Electrochim. Acta **24** (1979) 975.
79Vaa2 Vaartnou, M.G., Palm, U.V.: Elektrokhimiya **15** (1979) 1568.
80Khm Khmelevaya, L.P., Chizhov, A.V., Damaskin, B.B., Vainblat, T.I.: Elektrokhimiya **16** (1980) 257.
80Pal Palm, U.V., Pyarnoya, M.P.: Elektrokhimiya **16** (1980) 1599.
81Dob Dobrenkov, G.A., Gucewa, L.T.: Elektrokhimiya **17** (1981) 1272.
81Par Parsons, R., Peat, R.: J. Electroanal. Chem. **122** (1981) 299.
82Lar Larkin, D., Guyer, K.L., Hupp, J.T., Weaver, M.J.: J. Electroanal. Chem. **138** (1982) 401.
82Pal Paltusova, N.A., Alumaa, A.R., Palm, U.V.: Elektrokhimiya **18** (1982) 475.
83Bor Borghesani, G.: Electrochim. Acta **28** (1983) 483.
83Hup Hupp, J.T., Larkin, D., Weaver, M.J.: Surf. Sci. **125** (1983) 429.

83Lip Lipkowski, J., van Huong, C.N., Hinnen, C., Parsons, R., Chevalet, J.: J. Electroanal. Chem. **143** (1983) 375.

83Sar Sarangapani, S., Venkatesan, V.K.: Proc. Indian Natl. Sci. Acad. Part A **49** (1983) 124.

84Gon Gonzalez-Arjona, D., Andreu, R., Rueda, M.: J. Electroanal. Chem. **178** (1984) 305.

85Dam Damaskin, B.B., Dyatkina, S.L., Gerovich, V.M.: Elektrokhimiya **21** (1985) 716.

85Gol Goledzinowski, M., Dojlido, J., Lipkowski, J.: J. Electroanal. Chem. **185** (1985) 131.

85Hol Holze, R., Bełtowska-Brzezinska, M.: Electrochim. Acta **30** (1985) 937.

85Lus2 Lust, E.I., Palm, Yu.V.: Elektrokhimiya **21** (1985) 1381.

86Hol Holze, R., Bełtowska-Brzezinska, M.: J. Electroanal. Chem. **201** (1986) 387.

86Lus1 Lust, E.I., Palm, Yu.V.: Elektrokhimiya **22** (1986) 407.

86Lus2 Lust, E.I., Erlich, Yu.I., Palm, Yu.V.: Elektrokhimiya **22** (1986) 695.

86Ric2 Richer, J., Lipkowski, J.: Langmuir **2** (1986) 630.

86Rol Rolle, D., Schultze, J.W.: Electrochim. Acta **31** (1986) 991.

86Sto Stolberg, L., Richer, J., Irish, D.E., Lipkowski, J.: J. Electroanal. Chem. **207** (1986) 213.

87Sto Stolberg, D.E., Irish, J., Lipkowski, J.: Electroanal. Chem. **238** (1987) 333.

87Tuc Tucceri, R.I., Posadas, D.: Electrochim. Acta **32** (1987) 27.

88Bor1 Borkowska, Z.: J. Electroanal. Chem. **246** (1988) 423.

88Mil Milkowska, M.: Electrochim. Acta **33** (1988) 161.

88Pop Popov, A., Velev, O., Vitanov, T.: J. Electroanal. Chem. **256** (1988) 405.

88Ric Richer, J., Lipkowski, J.: J. Electroanal. Chem. **251** (1988) 217.

89Hol Holze, R., Fischer, G.: DECHEMA-Monogr. 117, Weinheim: VCH, 1989, p. 331.

89Ngu Nguyen Van Huong, C.: J. Electroanal. Chem. **264** (1989) 247.

89Sch Schomaker, S.: Diplomarbeit, Universität Oldenburg, 1989.

90Hol Holze, R., Schomaker, S.: Electrochim. Acta **35** (1990) 613.

90Jar Jarzabek, G.: J. Electroanal. Chem. **294** (1990) 253.

90Mon1 Moncelli, M.R., Guidelli, R.: J. Electroanal. Chem. **295** (1990) 239.

90Mon2 Moncelli, M.R., Foresti, M.L., Guidelli, R.: J. Electroanal. Chem. **295** (1990) 225.

90Pap Pape, U.: unpublished results, Oldenburg, 1990.

90Ric Richer, J.: Ph.D. Thesis, University of Guelph, 1990.

90Sto Stolberg, L., Lipkowski, J., Irish, D.E.: J. Electroanal. Chem. **296** (1990) 171.

91Obr Obraztsov, V.B., Parfenov, Yu.A., Danilov, F.I.: Elektrokhimiya **27** (1991) 980.

91Pez Pezzatini, G., Foresti, M.L., Moncelli, M.R., Guidelli, R.: J. Colloid Interface Sci. **146** (1991) 452.

91Sto Stolberg, L., Morin, S., Lipkowski, J., Irish, D.E.: J. Electroanal. Chem. **307** (1991) 241.

91Wan Wandlowski, T.: J. Electroanal. Chem. **312** (1991) 245.

92Boc Bockris, J.O'M, Jeng, K.T.: J. Electroanal. Chem. **330** (1992) 541.

92Jur Jurkiewicz-Herbich, M.: J. Electroanal. Chem. **332** (1992) 265.

92Ric Richer, J., Iannelli, A., Lipkowski, J.: J. Electroanal. Chem. **324** (1992) 339.

92Sko1 Skoluda, P., Dutkiewicz, E.: Electrochim. Acta **37** (1992) 75.

92Sko2 Skoluda, P., Dutkiewicz, E.: J. Electroanal. Chem. **329** (1992) 279.

92Sto Stolberg, L., Lipkowski, J., Irish, D.E.: J. Electroanal. Chem. **322** (1992) 357.

93For Foresti, M.L., Mocelli, M.R.: J. Electroanal. Chem. **348** (1993) 429.

93Luc Łuczak, T., Bełtowska-Brzezinska, M., Holze, R.: Electrochim. Acta **38** (1993) 717.

93Ote Otero, L., Vettorazzi, N., Barbero, C., Miras, M.C., Silber, J.J., Sereno, L.: J. Electroanal. Chem. **350** (1993) 251.

94Brz Brzostowska-Smolska, M.: Pol. J. Chem. **68** (1994) 1379.

94Dou Doubova, L.M., Valcher, S. Trasatti, S.: J. Electroanal. Chem. **376** (1994) 73.

94Ian Iannelli, A., Merza, J., Lipkowski, J.: J. Electroanal. Chem. **376** (1994) 49.

94Luc Łuczak, T., Holze, R., Bełtowska-Brzezinska, M.: Electroanalysis **6** (1994) 773.

94Lus Lust, E.J., Jänes, A.A.-Ja.: Elektrokhimiya **30** (1994) 357.

94Sot Sottomayor, M.J., Silva, F.: J. Electroanal. Chem. **376** (1994) 59.

94Zel Zelenay, P., Waszczuk, P., Dobrowolska, K., Sobkwoski, J.: Electrochim. Acta **39** (1994) 655.

95Luc1 Łuczak, T., Bełtowska-Brzezinska, M., Holze, R.: Electroanalysis **7** (1995) 251.
95Luc2 Łuczak, T.: Thesis, University of Poznan, 1995.
95Shi Shi, Z.H., Lipkowski, J., Mirwald, S., Pettinger, B.: J. Electroanal. Chem. **396** (1995) 115.
95Sob Sobkowski, J., Waszczuk, P., Zelenay, P.: Elektrokhimiya **31** (1995) 920.
95Was Waszczuk, P., Zelenay, P., Sobkowski, J.: Electrochim. Acta **40** (1995) 1717.
95Yan Yang, D., Lipkowski, J.: Elektrokhimiya **31** (1995) 836.
96Dut Dutkiewicz, E., Skoluda, P.: J. Chem. Soc. Faraday Trans. **92** (1996) 3763.
96Lus2 Lust, E., Jänes, A., Lust, K., Miidla, P.: J. Electroanal. Chem. **413** (1996) 175.
96Shi Shi, Z., Lipkowski, J., Mirwald, S., Pettinger, B.: J. Chem. Soc. Faraday Trans. **92** (1996) 3737.
96Sta Standke, C.: Diploma Thesis, Technische Universität Chemnitz-Zwickau, 1996.
96Vää Väärtnõu, M., Pärsimägi, P., Lust, E.: J. Electroanal. Chem. **407** (1996) 227.
96Wan1 Wandlowski, T., Hölzle, M.H.: Langmuir **12** (1996) 6597.
96Wan2 Wandlowski, T., Hölzle, M.H.: Langmuir **12** (1996) 6604.
97Dam1 Damaskin, B.B., Safonov, V.A., Baturina, O.A.: Elektrokhimiya **33** (1997) 117.
97Dam2 Damaskin, B.B., Baturina, O.A.: Elektrokhimiya **33** (1997) 1141.
97Ian Iannelli, A., Brolo, A.G., Irish, D.E., Lipkowski, J.: Can. J. Chem. **75** (1997) 1694.
97Jur Jurkiewicz-Herbich, M., Muszalska, A., Slojkowska, R.: Electrochemical Society Meeting 97-2, Paris, France, 1.-5.9.1997, Ext. Abstr. p. 1068.
97Lus2 Lust, E., Jänes, A., Lust, K., Miidla, P.: Electrochim. Acta **42** (1997) 771.
97Lus3 Lust, E., Jänes, A., Miidla, P., Lust, K.: J. Electroanal. Chem. **425** (1997) 25.
97Lus4 Lust, E., Jänes, A., Lust, K.: J. Electroanal. Chem. **436** (1997) 141.
97Rog Rogozhnikov, N.A., Beck, R.Y.: J. Electroanal. Chem. **434** (1997) 19.
97Sil Silva, F., Sottomayor, M.J., Coelho, V.: Electrochemical Society Meeting 97-2, Paris, France, 1.-5.9.1997, Ext. Abstr. p. 956.
98Bel Bełtowska-Brzezinska, M., Łuczak, T., Holze, R.: Surf. Sci. **418** (1998) 281.
98Hat Hatchett, D.W., Uibel, R.H., Stevenson, K.J., Harris, J.M., White, H.S.: J. Am. Chem. Soc. **120** (1998) 1062.
98Lus Lust, E., Jänes, A., Lust, K.: J. Electroanal. Chem. **449** (1998) 153.
98Mil Milkowska, M., Jurkiewicz-Herbich, M.: Pol. J. Chem. **72** (1998) 1096.
98Vää Väärtnõu, M., Lust, E.: J. Electroanal. Chem. **445** (1998) 165.
99Dur Durnie, W., De Marco, R., Kinsella, B., Jefferson, A.: J. Electrochem. Soc. **146** (1999) 1751.
99Jän1 Jänes, A., Vagel, A., Miidla, P., Lust, E.: Ext. Abstracts of 50th ISE-Meeting, Pavia, Italy, September 5th-10th, 1999, Abstract 2-466.
99Jän2 Jänes, A., Miidla, P., Lust, E.: J. Solid State Electrochem. **3** (1999) 277.
99Vää1 Väärtnõu, M., Lust, E.: J. Electroanal. Chem. **469** (1999) 182.
99Vää2 Väärtnõu, M., Lust, E.: Electrochim. Acta **44** (1999) 2437.
00Gas Gasparac, R., Martin, C.R., Stupnisek-Lisac, E.: J. Electrochem. Soc. **147** (2000) 548.
00Lus Lust, K., Väärtnou, M., Lust, E.: Electrochim. Acta **45** (2000) 3543.
00Vää Väärtnõu, M., Lust, E.: Electrochim. Acta **45** (2000) 1623.
01Dou Doubova, L.M., De Battista, A., Daolio, S., Trasatti, S.: J. Electroanal. Chem. **500** (2001) 134.
01Jän Jänes, A., Lust, E.: Electrochim. Acta **47** (2001) 967.
01Li Li, H.-Q., Chen, A., Roscoe, S.G., Lipkowski, J.: J. Electroanal. Chem. **500** (2001) 299.
01Luk Łukomska, A., Smolinski, S., Sobkwoski, J.: Electrochim. Acta **46** (2001) 3111.
01Nur Nurk, G., Jänes, A., Miidla, P., Lust, K., Lust, E.: J. Electroanal. Chem. **515** (2001) 33.
01Slo Slojkowska, R., Jurkiewicz-Herbich, M.: Colloids Surf. **178** (2001) 325.
01Vää1 Väärtnõu, M., Lust, E.: J. Electroanal. Chem. **499** (2001) 136.
01Vää2 Väärtnõu, M., Lust, E.: Electrochim. Acta **47** (2001) 997.
02Buk Bukowska, J., Calvo, M., Jurkiewicz-Herbich, M., Jackowska, K.: Pol. J. Chem. **76** (2002) 1151.
02Jur Jurkiewicz-Herbich, M., Milkowska, M., Slojkowska, R.: Colloids Surf. A **197** (2002) 235.
02Lus Lust, K., Väärtnõu, M., Lust, E.: J. Electroanal. Chem. **532** (2002) 303

03Jän Jänes, A., Miidla, P., Lust, E.: J. Solid State Electrochem. **7** (2003) 189.

05Eme Emets, V.V., Damaskin, B.B.: J. Electroanal. Chem. **582** (2005) 97.

05Kas Kasuk, H., Nurk, G., Lust, K., Lust, E.: J. Electroanal. Chem. **580** (2005) 128.

06Bou Bouklah, M., Hammouti, B., Lagrenée, M., Bentiss, F.: Corros. Sci. **48** (2006) 2831.

06Luk Łukomska, A., Sobkowski, J.: J. Electroanal. Chem. **592** (2006) 68.

06Vää Väärtnõu, M., Lust, E.: J. Electroanal. Chem. **587** (2006) 220.

5 Exchange current densities

5.1 Kinetics of electrode processes at an electrode in contact with an electrolyte solution [1])

In a simplified electrode reaction of the form

$$\mathrm{Ox} + n\,\mathrm{e}^- \rightarrow \mathrm{Red} \tag{5.1}$$

an oxidized species Ox is reduced by a transfer of n electrons from the electrode onto the species forming the reduced species Red. The number n of electrons transferred is sometimes called reaction valency. An electrode reaction is always a heterogeneous reaction; its rate is assumed to be proportional to the surface area A. With ξ designating the extent of the reaction the rate divided by the surface area is

$$v=\frac{1}{A}\frac{\mathrm{d}\xi}{\mathrm{d}t}=v_\mathrm{a}-v_\mathrm{c} \tag{5.2}$$

with v_a and v_c denoting the rate of the oxidation (anodic) and the reduction (cathodic) reaction. In the following text as well as in the tables it is assumed (if not explicitly stated otherwise), that with a single rate determining step a number of n electrons is transferred and a stochiometric number υ of generated/consumed species of the transition state is passed in the process of the electrode reaction. The total Faradaic current density (given in A cm^{-2}) is composed of the anodic current density j_a and the cathodic current density j_c according to

$$j=j_\mathrm{a}+j_\mathrm{c} \tag{5.3}$$

with the partial current densities related to the corresponding reaction rates

$$j_\mathrm{a}=nFv_\mathrm{a} \tag{5.4}$$

and

$$j_\mathrm{c}=-nFv_\mathrm{c} \tag{5.5}$$

By definition a current of positive charge (an oxidation reaction) is taken as a positive current, a current of negative charge (a reduction reaction) is taken as a negative current. When the electrode is at equilibrium, the net current density j equals zero. This implies, that

$$j_\mathrm{a}=|j_\mathrm{c}|=j_0 \tag{5.6}$$

[1]) As outlined in chapter 1 the term „electrode“ is used - contrary to the suggestion of W. Nernst - to designate the electronically conducting phase only, the term „electrolyte solution“ covers all types of ionically conducting phases (solutions, melts, solids) being in contact with the former phase.

This exchange current j_0 density shows a characteristic value for a given reacting species, electrolyte solution composition and electrode material, in addition it depends on the concentration of the reacting species. Other exchange current densities related to surface properties like e.g. surface step or dislocation densities have been reported [74Vit2], they are not included in the following tables. The relationship between these various currents and the electrode potential is given by the Butler-Volmer-equation, a complete treatment of this equation has been provided by various authors elsewhere [80Bar2, 01Bar]. In order to obtain values of j_0 which can be compared directly without regard to different reactant concentrations a standard exchange current density j_{00} has been defined, which refers to a reactant concentration of 1 mol L^{-1} (for the definition see below). This current density has also been called conditional exchange current density j^0 (see e.g. [75Tam]). Unfortunately there are numerous examples, where the values of j_{00} calculated from values of j_0 obtained at various reactant concentrations do not agree. Consequently in the table reported below values of j_{00} are listed only when reported by the authors. When the partial current density is found to be proportional to the concentration of a reactant c_i or the product of concentrations of several reactants $\prod_i c_i^{\nu i}$ with ν_i being the reaction order with respect to c_i, an electrochemical rate constant for the anodic reaction

$$k_{ox} = j_a / nF \prod_i c_i^{\nu i} \tag{5.7}$$

and for the cathodic reaction

$$k_{red} = j_c / nF \prod_i c_i^{\nu i} \tag{5.8}$$

can be defined. With Σ being the overall reaction order (the sum of all ν_i) the electrochemical rate constant has the dimension $mol^{(1-\Sigma)}cm^{(3\Sigma-2)}s^{-1}$. For a first order reaction the dimensions are cm s^{-1}.

In addition to the exchange current density the transfer coefficient α is needed to describe the relationship between the electrode potential and the current flowing across the electrode/solution interface. From a formal point of view α can be obtained by calculating the partial current densities with respect to the electrode potential for the anodic reaction:

$$\alpha_a / \nu = (RT/nF)(\partial \ln j_a / \partial E)_{T,p,c_i} = (RT/nF)(\partial \ln k_{ox} / \partial E)_{T,p} \tag{5.9}$$

and for the cathodic reaction

$$\alpha_c / \nu = -(RT/nF)(\partial \ln j_c / \partial E)_{T,p,c_i} = (RT/nF)(\partial \ln k_{red} / \partial E)_{T,p} \tag{5.10}$$

From a kinetic point of view α describes the influence of a change of the electrode potential on the energy of activation for the charge transfer reaction which in turn influences the partial current density. The transfer coefficients α_a for the anodic charge transfer reaction and α_c for the cathodic reaction add up according to

$$\alpha_a + \alpha_c = 1 \tag{5.11}$$

Assuming $\alpha_a + \alpha_c$ to be independent of the electrode potential integration of (5.9) and (5.10) yields

$$k_{ox} = k_{ox}^0 \exp(\alpha_a nFE/RT) \tag{5.12}$$

and

$$k_{red} = k_{red}^0 \exp(-\alpha_c nFE/RT) \tag{5.13}$$

with k_{ox}^0 and k_{red}^0 being the rate constant at E = 0 V with respect to a reference electrode. These values are sometimes reported for electrode reactions with unknown equilibrium potentials. A closely related rate constant also reported sometimes is the apparent rate constant k_{app}. This constant is calculated from the observed current I at a selected electrode potential, the surface area of the electrode A and the surface concentrations c_s of the reacting species according to

$$k_{app} = I/FAc_s \tag{5.14}$$

The surface concentration c_s can be obtained in various ways. With cyclic voltammetry c_s can be determined according to

$$c_s = c_b\,[(I_1 - I)/I_1] \tag{5.15}$$

with c_b being the bulk reactant concentration; I_1 being the peak current in the voltammogram and I being the current in the rising part of the current wave [80Bar1]. This rate constant varies with electrode potential, thus a precise comparison of values is possible only when the electrode potential is known.

A conditional electrode reaction rate constant [2]) k_c^θ which is defined as the common value of the anodic and cathodic rate constants at the conditional electrode potential $E^{0'}$ is given according to

$$k_c^\theta = k_{ox}/\exp[\alpha_a nF(E - E^{0'})/RT] \tag{5.16}$$

and

$$k_c^\theta = k_{red}/\exp[-\alpha_c nF(E - E^{0'})/RT] \tag{5.17}$$

It should be kept in mind, that these rate constants are defined based on the volume concentrations of the reacting species. Another standard rate constant k^0 can be defined with regard to the rate of the reaction at the standard electrode potential of the electrode reaction. This rate constant refers consequently to standard activities instead of concentrations.

As a result of polarographic investigations using e.g. a dropping mercury electrode electrochemical rate constants at the half wave potential $E_{1/2}$ are reported:

$$k_{1/2} = k_{red}/\exp[-\alpha_c nf(E - E_{1/2})/RT] \tag{5.18}$$

If the reaction is diffusion controlled, i.e. the charge transfer is fast compared to reactant transport, the value of $k_{1/2}$ is equal to k_c^θ.

With the electrochemical rate constant a relationship between the exchange current density, the electrochemical rate constant and the concentration of the reacting species can be derived:

$$j_0 = nFc_{red}k_{ox}/\exp(\alpha_a nF\eta/RT) \tag{5.19}$$

and

$$j_0 = nFc_{ox}k_{red}/\exp(-\alpha_c nF\eta/RT) \tag{5.20}$$

with the overpotential $\eta = E - E_0$. With (5.11) the conditional electrode reaction rate constant (5.17) and (5.18) can be fused

$$j_0 = nFk_c^\theta c_{ox}^{\alpha_a} c_{red}^{\alpha_c} \tag{5.21}$$

The standard (conditional) exchange current density can be rewritten accordingly:

$$j_{00} = nFk_c^\theta (c_{ox}^\theta)^{\alpha_a} (c_{red}^\theta)^{\alpha_c} \tag{5.22}$$

wherein c_i^θ is the standard concentration (usually 1 mol dm^{-1}). In order to obtain an electrochemical quantity not depending on the actual reactant concentration (as the rate constant k_0) the exchange current density j_0 can be converted into the standard exchange current density j_{00} according to

$$j_0 = j_{00} c_{ox}^{\alpha_a} c_{red}^{\alpha_c} (c^\#)^{-1} \tag{5.23}$$

Unfortunately there are numerous examples in the literature where conversion of values of j_0 obtained at various concentrations did not result in the same value of the standard exchange current density j_{00} (e.g. 83Kaw]).

Further relationships have been discussed elsewhere [64Tan, 70Mat]. Compilations of exchange current densities j_0 or the corresponding kinetic parameter k^0 etc. have been published only sparingly [64Tan, 75Tam]. Relationships between the rate of an electrochemical reaction and other properties of the participating materials and species like e.g. the work function or electronic properties of the metal acting as an electrode have been discussed elsewhere [57Con1, 72Tra, 77Tra], the influence of the energies of the hydrogen-platinum bond on the exchange current density has been treated using density functional theory [05Nor]. Relationships between electrode kinetics (including j_0) and E_{pzc} have been discussed [70Khe]. The way the electron is actually passing the phase boundary, the influence of interactions between the electrode and the reacting species (inner-sphere versus outer-sphere reaction [76Wea, 96Tor, 01Wea]) and the mechanism of processes closely associated with the electron transfer itself (i.e. the electrode reaction mechanism) will not be considered in this text.

5.2 Methods for the determination of exchange current densities

Current density versus electrode potential curves (current-potential curves)

At an electrode immersed in an electrolyte solution a rest electrode potential will be established provided that a potential determining electrode reaction can take place (e.g. at the interface between a copper wire and a copper ion-containing solution or at the interface between a platinum wire and an electrolyte solution containing both iron(II)- and iron(III)-ions). If this electrode is connected to a suitable electronic device (i.e. a potentiostat) a change of this electrode potential can be effected by forcing an electric current of sign and magnitude appropriate to change the interfacial concentration of those species through the electrode/solution interface. Various steps of the overall electrode process may limit this current and may thus interfere with the establishment of the desired electrode potential. Provided that the charge transfer is the rate determining step the relationship between current density and electrode potential is given by the Butler-Volmer-equation [80Bar2, 01Bar]. At deviations from the rest potential, small enough to avoid transport to become the rate controlling step and large enough to neglect the corresponding back reaction, mathematic simplification results in a linearized form of this equation (Tafel approximation). From respective plots of current density and electrode potential the exchange current density and the transfer coefficient can be derived. An experimental procedure for measurements with a dropping mercury electrode involving photographic determination of drop size and surface area has been described [64But1]. (Data obtained with this method are labelled CP.)

Cyclic voltammetry

Beyond applying a fixed electrode potential different from the rest potential various other time dependent potential changes are feasible. In an experimental approach widely used the electrode potential is changed linearly with time. Upper and lower potential limits are fixed, thus the change of the electrode potential plotted as a function of time yields a triangular plot. Thus this methods is frequently called triangular voltage sweep method. Since the potential is cycled between these limits and since the potential (in volts) and the current (in ampere) are recorded, the method is also called cyclic voltammetry. In the presence of a redox couple in the electrolyte solution typical voltammograms with current peaks are obtained which provide the information needed for the calculation of electrode reaction rate constants. An introduction to the method was provided by Nicholson [65Nic]. For reviews see [84Hei2, 89Bon]. High speed cyclic voltammetry [88Wip] has been reviewed by Wightman and Wipf [90Wig]. The suitability of microelectrodes in cyclic voltammetry in particular at low temperatures has been discussed [92Saf]. As the term voltammetry itself can be applied to a broad variety of methods, where the fundamental principle (i.e. re-

cording the flowing current as a function of the electrode potential) is combined with different methodological variations (i.e. modulation of the electrode potential with an additional signal), there are numerous „voltammetries". They are treated in the appropriate context of the additional modification below (e.g. AC voltammetry). (Data obtained with the basic method are labelled CV.)

Thin-layer voltammetry

Cyclic voltammetry as described above can be performed in an electrochemical cell wherein the solution volume is confined to a space of a few micrometers thickness between the flat working electrode and an inert plate mounted adjacent to the working electrode. Connection to the counter and reference electrode is made by inserting the assembly into the electrolyte solution contained in a standard electrochemical cell. Because of the thin liquid film enclosed between working electrode and inert plate diffusional transport of solution species from within the bulk solution is impeded. At sufficiently low scan rates ($dE/dt <$ 2 mV s^{-1}) and electrolyte layer thickness ($d < 4 \cdot 10^{-3}$ cm) the reactant concentration at the electrode surface will differ only negligibly from the bulk concentration. The amount of species converted electrochemically within the time frame of a potential sweep is thus limited to the amount of liquid confined in the said volume. When the species under investigation undergoes an electrode reaction wherein the species converted from a given state of oxidation into another one can be converted back into its initial state (this criterion has been called sometimes misleadingly „reversibility") the potential splitting between the oxidation and reduction peak can be used to calculate the rate of reaction. For further information and reviews see [66Hub, 69Hub, 70Hub, 72Lai]. (Data obtained with this method are labelled TLV.)

Coulostatic method

In the absence of an external current flowing across the solution/electrode interface the electrochemical interface can be treated in good approximation as a capacity (i.e. the double layer capacity) connected in parallel to a resistor (i.e. the charge transfer resistance). Any electrical charge brought onto the capacity will be discharged into the resistance. The rate of discharge will depend on the magnitude of the resistance; the latter being directly related to the exchange current j_0. The charging of the double layer is done in various ways, the duration of the charging process has to be sufficiently short in order to avoid any appreciable concentration perturbation because of the electrolysis taking place. The potential-time curves recorded after the charging are evaluated resulting in the desired kinetic parameters [62Del1, 62Del2, 62Del3, 62Rei1, 62Rei2]. A consideration of the limitations of this method has been reported [67Koo2]. The method has also been called coulostatic relaxation method. (Data obtained with this method are labelled CS.)

Potential step method

A sudden change of the electrode potential (i.e. a potential step, sometimes less precisely called voltage step) applied by e.g. a potentiostat results in a current flow needed to change the concentration of the species involved in the electrode potential determining reaction to the new values determined by the new value of the electrode potential. An additional current flow is needed to charge the electrochemical double layer capacitance to the new value. The latter process is accomplished within usually 10^{-3} s. During a brief period of time following immediately the charging time the potential-time relationship is controlled predominantly by the charge transfer process. A plot of the measured current versus $t^{-1/2}$ extrapolated for $t = 0$ s for various values of ΔE yields the relationship of charge transfer overpotential versus charge transfer current, which in turn provides the desired exchange current density j_0 [54Lor, 55Ger1, 55Ger2, 57Vie]. A treatment taking into account double layer effects has been reported [69Tan1], the method has also been simply called potentiostatic method. An experimental approach employing two matching electrochemical cells (only one containing the species dissolved in the supporting electrolyte solution being present in both cells), wherein the difference in response to the potential step is recorded (thus the method has been termed differential potentiostatic technique) has been described [64Bon]. (Data obtained with this method are labelled PS.)

Cyclic potential step method

Based on an evaluation of the drawbacks and limitations of the established potential step methods Smit and Wijnen [60Smi, 60Wij2] have developed a method wherein a cyclic potential-step is applied to a polarizable electrode. Mathematical treatment of the current response predicts a rapidly established constant pattern. This prediction is free from limitations pertaining to reversibility of the electrochemical system [54Kam, 56Kou] or amplitude of the potential perturbation (compare with e.g. AC and square wave polarography [58Bar2, 58Mat]). Exchange current densities and transfer coefficients can be obtained. (Data obtained with this method are labelled CPS.)

Impedance method

The measurement of the electrode impedance has also been called Faradaic impedance method. Since measurements are possible by applying either an electrode potential modulated by an AC voltage of discrete frequency (which is varied subsequently) or by applying a mix of frequencies (pink noise, white noise) followed by Fourier transform analysis, the former method is sometimes called AC impedance method. The optimization of this method for the use with ultramicroelectrodes has been described [91Bar1]. (Data obtained with these methods are labelled IP.)

Faradaic rectification

When the electrode potential of the working electrode is modulated with a sinusoidal alternating current, the mean potential is shifted by a small increment in many cases. This effect has been named Faradaic rectification, it is caused by the non-linearity of the proceeding electrode processes, in particular their variation of rate with the electrode potential [51Dos2, 54Hil, 57Old, 71Aga, 94Gal]. A theoretical treatment of an electrode in contact with a solution containing a redox system has been provided [57Old, 51Dos3, 51Dos4, 52Dos2]. It was extended to reactions where one reactant is present in its elementary form dispersed throughout a metallic liquid phase (e.g. $Cd^{2+} \rightarrow Cd(Hg)$) [58Bar1]. An improved evaluation technique has been proposed [67deL], some inherent problems have been reviewed [60Bau]. A variant of this method applied to polarography has been described by van der Pol et al. [73Pol]. The measurement of the second harmonic in AC polarography (see below) has been described, the approach closely resembles Faradaic rectification [66Bau]. (Data obtained with all variants of the Faradaic rectification method are labelled FR.)

Faradaic distortion method

When a pure sinusoidal AC current passes across the electrode/solution interface, the cell voltage (a two electrode arrangement is used) shows a sinusoidal perturbation. It contains multiples of the fundamental frequency of the modulation, the first harmonic dominates. The magnitude of the effect is comparable to Faradaic rectification, but experiments may be easier to perform. Measurement and evaluation have been described in detail [60Old, 72Hil2]. (Data obtained with this method are labelled FD.)

Galvanostatic method

The potential of an electrode being initially at equilibrium (i.e. no net current is flowing across the electrochemical interface) is shifted when a constant current is applied. Recording the change of the electrode potential as a function of time after turning-on of the current (i.e. applying a current step) results in curves, which yield the desired kinetic data. Since only the initial part of the curve is of interest (in this time domain transport etc. does not influence the current-overpotential relationship) a fast electronic current switch is required. Instead of only switching on the current a current pulse of defined duration can be applied, in this case the method is called galvanostatic single pulse method [55Ber1, 55Ber2], for details

see below. A consideration of the limitations of this method has been reported [67Koo2]. (Data obtained with this method are labelled GM.)

Galvanostatic single pulse method

An electrochemical system (i.e. an electrode) can be displaced from its equilibrium situation by applying a current pulse or step. From the electrode potential response kinetic data pertaining to the redox system present at the electrode/solution interface (j_0 and α) can be derived [55Ger1, 55Ger2, 55Ber1, 55Ber2, 55Ber3]. In addition to the change of the electrode potential this pulse or step also charges the double layer capacitance. The actual value of the differential capacitance of the double layer can be extracted from the charging curve. From the subsequent relaxation of the electrode potential the charge transfer resistance can be calculated taking into account the already obtained value of the double layer capacitance. By limiting the pulse duration to very short values (ns to μs) the supplied electrical charge can be kept very low, consequently mass-transport contributions need not be taken into account. Infrequently this method has been called 'current impulse relaxation technique' [67Wei1, 67Wei2]. Further details can be found elsewhere [67Wei1, 69Dau]. (Data obtained with this method are labelled GS.)

Galvanostatic double pulse method

The current flowing across the interface between the electronically conducting phase (electrode) and the ionically conducting phase (electrolyte solution etc.) causes a deviation of the electrode potential from its rest value in the absence of a net current to a new value. The difference is called overpotential; it is attributed to the limited rate of the involved steps of the overall electrode process. The aim of many experimental methods used for the measurement of the rate of the electrochemical charge transfer is to exclude the unwanted influence of all other steps than the central charge transfer step. A sudden change (e.g. from zero to a given value) of the current flowing across the interface effected by a suitable instrument (a galvanostat) will result in a current immediately after the step which is only limited by the rate of the charge transfer because all other steps of the electrode reaction are still almost in equilibrium. Evaluation of the potential-time transients should yield the charge transfer overvoltage from the relationship between applied current and observed initial overpotential, which can be combined with the applied current to yield the charge transfer resistance and finally the exchange current density

$$R_{ct} = \frac{RT}{nFj_0} \tag{5.24}$$

Unfortunately the sudden change in current and potential causes a substantial charging current necessary to adapt the electrochemical double layer to the new electrode potential. This distorts the potential-time transients in the crucial initial period. Gerischer and Krause have developed a method which provides a short pulse of a large current used predominantly to charge the double layer practically without affecting the concentration of the reactant followed by a second much smaller current step [57Ger2]. At the beginning of the second step the slope of the potential-time transient should be zero, from the relationship between the measured potential and the applied current the charge transfer resistance etc. can be calculated. A proper match of time and duration of the first pulse has to be found by trial and error. (Data obtained with this method are labelled GD.)

Galvanostatic current interruption

Interruption of the galvanostatically controlled current flowing across the electrode/solution interface results in potential-time curves which can be evaluated in order to obtain values of j_0 [70New]. An experimental setup has been described [87Wru]. (Data obtained with this method are labelled GCI.)
A method based on the oscillographic observation of the response after current interruption or turning-on suitable for its interpretation taking into account the effect of side reactions on potential-time curves has been described [59Khe, 61Zin]. (Data obtained with this method are labelled OCS.)

Cyclic current step method

The inherent drawbacks and limitations of current step and pulse methods (e.g. need of photographic recording or transients recorders, influence of double layer capacitance) have caused further developments including the application of a cyclic current square wave. The observed cell response reaches a steady state very soon and can be inspected easily with an oscilloscope. Exchange current density and double layer capacitance can be obtained. A mathematical treatment for a symmetrical square wave and a cyclic state of the cell response has been provided by Berzins and Delahay [55Ber3]. An extended treatment including asymmetric square waves applicable even when the cell response has not yet attained a steady state has been reported by Wijnen and Smit [60Wij1, 60Wij2, 60Wij3]. This method (together with the cyclic potential step method) has been infrequently called oscilloscopic square wave method [66Tep1, 66Tep2, 68Tep]. (Data obtained with this method are labelled CCS.)

Dropping mercury electrode

The dropping electrode, in particular used with mercury and with amalgams, has been frequently employed in kinetic studies [66But3]. Data obtained with this electrode/method are labelled DME, if no specific technique is applied that is more helpful in qualifying the obtained data.

Open-circuit method

When the area A of the electrode/solution interface with a redox system in the solution varies (e.g. when using a streaming mercury electrode), the double layer capacity which is proportional to A, varies too. The corresponding double layer charging current has to be supplied at open circuit conditions by the Faradaic current of the redox reaction. The associated overpotential can be measured with respect to a reference electrode. By measuring the overpotential at different capacitive current densities (i.e. Faradaic current densities) the current density vs. electrode potential relationship can be determined, subsequently kinetic data can be obtained [65Del3]. (Data obtained with this method are labelled OC.)

Chronopotentiometry

A constant current flowing across the electrolyte solution/electrode interface causes a potential shift because of the changing concentrations of educts and products, which are consumed and generated respectively. The change of the electrode potential as a function of time is recorded in a chronopotentiometric experiment. Depending on the rate of the electrode reaction various mathematical treatments are possible providing access to rate constants; for details see e.g. [01Bar]. (Data obtained with this method are labelled CH.)

Chronocoulometric method

When the potential of an electrode is switched from a value, where an electrode reaction under investigation proceeds at a negligible rate, to a value, where the rate is measurable, a charge-time behavior can be observed, which contains besides numerous other electrochemical parameters the rate constant of the electrode reaction. Details and a complete derivation have been given elsewhere [73Rod, 75Wea, 76Wea] (Data obtained with this method are labelled CM.)

Direct current polarography

Kinetic data especially of slow electrode reactions can be extracted from polarographic curves obtained with direct current polarography (DC polarography) and a dropping mercury electrode (DME, see entry above). A complete review of the fundamentals of polarography including the use of DC polarography for the derivation of values of i_0 has been provided [60Sta, 65Hey, 53Del]. Equations describing the current potential relationship useful for the derivation of i_0 have been reported [65Del2]. The analysis of irreversible waves has been described by Koutecky and Weber [53Kou, 55Web].
The term polarography basically refers to a method, where the current flowing across the electrochemical interface is recorded as a function of the applied electrode potential, historically in most cases a mercury electrode is involved. Thus polarography might be called also voltammetry. This sometimes results in confusing terms like e.g. AC voltammetry, which is obviously equivalent to AC polarography (see following entry). (Data obtained with this method are labelled DCP.)

Alternating current polarography

Superposition of a small amplitude AC voltage on the DC ramp in a polarographic or voltammetric experiment results in an alternating current added to the direct current. Analysis of the AC part as a function of the applied electrode (DC) potentials yields kinetic information about the proceeding electrode process [58Mat, 58Bau]. An advanced technique, wherein an excitation signal composed of several discrete, non-coherent frequencies is supplied to the electrochemical cell and the response is measured at every frequency simultaneously using appropriately tuned electronic circuits (the method is termed AC polarography in the non-coherent wave frequency mode), has been reported [72Hue]. A combination of the Faradaic rectification method with polarography has been described by van der Pol et al.; for theoretical and experimental details see [73Pol]; results obtained with this method are designated FR. For alternating current harmonics methods, i.e. the measurement of the second harmonic in AC polarography [66Bau] see above because of the similarities with Faradaic rectification. (All data obtained with AC polarographies and voltammetry are labelled ACP.)

Oscillographic polarography

Application of a suitable electrode potential to a small electrode immersed in a solution containing a low concentration of an electrochemically active species with a large concentration of an indifferent supporting electrolyte can result in a diffusion controlled flow. Sudden application of this potential results in an initially large current which decreases with time according to the growing diffusion-layer thickness. Using a sufficiently fast recording device (e.g. a cathode ray oscilloscope (i.e. an oscillograph)) the time-dependent current-potential relationship can be recorded. Essentially the same information obtained from a polarogram recorded within about ten minutes can be obtained here within a few seconds. The form of the obtained curves and their evaluation have been described [48Ran, 48Sev, 55Mat]. (Data obtained with this method are labelled OP.)

Pulse polarography

Theoretical treatment of polarographic curves for the calculation of values of j_0 has been described [65Hey, 66Hey], for an overview see [94Gal], a further evaluation procedure has been described [68Old]. Experimental details, in particular of solid electrodes in combination with a rotating disk electrode have been reported elsewhere [84Guy]. (Data obtained with this method are labelled PP.)

Miscellaneous polarographic methods

Beyond specific polarographic techniques variations of these methods have been developed aiming at simplified and fast evaluation using modified measurement procedures and data processing. Oldham and

Parry have described procedures to determine the kinetic parameters of totally irreversible electrode processes using various polarographic techniques [68Old]. (Data obtained with these methods are collectively labelled POL.)

Rotating disc electrode

The current flowing across an electrochemical interface is in most cases controlled or at least influenced by transport limitations. A determination of a pure charge transfer overpotential or charge transfer current is thus hampered. This problem may be solved either by providing unlimited reactant supply (which is impossible) or by providing a reactant supply which can be mathematically treated. This treatment can be used to remove the transport contribution from the experimental data. So far the reactant flow towards a rotating disc electrode and the turbulent flow within a tube or a channel (see below) have been described mathematically with sufficient precision. In case of a rotating disc electrode the active electrode area is disc shaped and embedded in the front end of a rod of insulating material. Rotation of the rod drags the fluid along the rod surface, the effective centrifugal force flings the solution outwards, and fresh solution from the bulk is flows normal to the disk surface. The hydrodynamic and the convective-diffusion equations have been solved rigorously for this system [62Lev, 71Alb, 76Ple, 84Fil]. By extrapolating current versus rate of rotation data obtained at various overpotentials towards infinite rates the transport contribution can be eliminated and pure charge transfer controlled data become available resulting finally in values of j_0. (Data obtained with this method are labelled RDE.)

Controlled hydrodynamic convection

Alternatively controlled hydrodynamic convection can be achieved by using a rotating bucket (electrochemical cell) or a controlled electrolyte flow and a stationary electrode [55Jor, 62Los1, 62Jor]. The evaluation of cyclic voltammograms obtained with a stationary electrode exposed to a flowing electrolyte solution moved by a rotating disc has been treated in detail [60Oza1, 60Oza2, 68Suz, 69Suz, 70Suz]. The method has been called hydrodynamic voltammetry; the electrode has been named 'convection electrode'. A setup suitable for liqud metal electrodes has been described [64Suz]. In order to minimize transport control limiting the charge transfer reaction forced convection can be achieved by using a rotating metal wire as a working electrode [50Ger]. (Data obtained with this method are labelled CF).

Turbulent pipe flow

Beside laminar flow created by e.g. a rotating disc electrode turbulent flow provides a means of artificially enhanced transport. A consistent mathematical description and analytical treatment of this mode of transportation is not possible. Various approximations have been proposed and tested for correctness [84Bar1], an experimental setup has been described [78Ber, 83Her, 83Iwa]. From comparisons of measured and calculated current density vs. electrode potential relationships exchange current densities are available. (Data obtained with this method are labelled TPF.)

Stirred solution

Steady state polarization curves showing the dependence of the electrode potential vs. the current flowing across the solution or electrolyte melt/electrode interface can be recorded in a simple three-electrode arrangement in either the galvanostatic or the potentiostatic mode. In order to minimize transport control limiting the charge transfer reaction, stirring the electrolyte solution can be used to enhance convection artificially [68Boc2]. (Data obtained with this method are labelled SS.)

Radiotracer (radiochemical) method

The electron transfer between an electrode (e.g. a metal) and a corresponding species in solution (e.g. the metal ion) is a dynamic equilibrium process at $E = E_0$ with $j_a = |j_c| = j_0$. This implies a continuous exchange of species between solution and electrode. Such an exchange can be detected and measured with radioactive species (radiotracers) being present initially only in one phase. As a function of time the other phase will become radioactive too. The rate of the increase of radioactivity corresponds to the value of j_0. Because of the limited speed of measurements of radioactivity the measurable values of j_0 and k_c^θ are low ($k_c^\theta \leq 10^{-2}$ cm s^{-1}). In case of redox systems like e.g. Ce(III)/Ce(IV) the different solubilities of the involved species in a separate liquid phase can be used for separation. Experimental setups have been described [56Fro1, 63Ric, 63Bom, 63Fio]. (Data obtained with this method are labelled RT.)

Derivative cyclic voltabsorptometry

Using optically transparent electrodes (OTE) the concentration of light absorbing species formed or consumed during electrochemical reactions can be monitored [84Hei1]. Usually this is done under stationary or quasistationary conditions, i.e. at a constant or at an only slowly changing electrode potential. Results are displayed as absorption A vs. electrode potential E. Instead of the absorption the derivative $\partial A/\partial E$ can be registered. The obtained plots closely resemble cyclic voltammograms. Digital simulation of these plots can be used to obtain the rate constant of the involved electrode reaction [81Ban]. (Data obtained with this method are labelled DCVA.)

Circular dichroic spectroelectrochemical method

Optically active molecules show circular dichroism. Their extinction coefficients ε_L and ε_R are different and change as a function of wavelength. Using a suitable spectroelectrochemical cell, $\Delta\varepsilon = \varepsilon_L - \varepsilon_R$, which is usually small compared to conventional extinction coefficients, can be measured. Combined with the special properties of a thin layer cell kinetic data can be extracted from CD-data [01Liu]. (Data obtained with this method are labelled CD.)

Mixed potential method

A titration procedure essentially equivalent to the redox titration known from analytical chemistry can be employed to obtain kinetic parameters including rate constants of processes. In a two-compartment cell with an ionically conducting connection (porous glass frit, salt bridge) wire electrodes of the desired electrode material and reference electrodes are mounted in both cells. A supporting electrolyte solution purged with inert gas is filled into both compartments. The reduced (or oxidized) form of the redox system under study is filled into one compartment at a concentration large enough to avoid diffusion limitation. Into the other compartment small amounts of the oxidized (or reduced) form of the opposite couple are added stepwise. Measurements of the current flowing between both electrodes under short-circuit conditions and of the electrode potential of the electrode in the first compartment enables the determination of titration curves, which can be further evaluated in order to obtain the desired kinetic data [70Her]. (Data obtained with this method are labelled MP.)

General remarks

Because of the outlined difficulties in a simple transformation of the exchange current density j_0 into the standard exchange current density j_{00} or the rate constant k_c^θ all values are listed as reported in the original literature. The exchange current density is usually reported in units of A cm^{-2}, the rate constant in units of cm s^{-1}. Sometimes the exchange current density or the standard exchange current density are reported in units of A cm mol^{-1}. All values reported below refer to the units outlined initially referring to

molar concentrations. Apparent rate constants k_{app} are given in parentheses. Since the conditions for calculating k_{app} are rather complex and detailed, in many cases the reader is referred to the original reports for exact descriptions. Further data treatment of obtained kinetic data (e.g. conversion of the initially calculated rate constant to the value the constant is assumed to reach at a certain different electrode potential (e.g. [72Lai])) is not listed. In case a specific type of electrode (e.g. dropping mercury electrode) is used, nevertheless the employed method is stated preferably.

Data are listed according to the alphabetical order of the electrode material. Polycrystalline materials are listed first followed by single crystal samples and amalgam electrodes for every element. Within a given electrode material data are arranged according to the alphabetically listed reactants. Within a given reactant data are listed with increasing concentration of the supporting electrolyte solution. Electrode reactions proceeding at a mercury electrode may be listed at the electrode material (mercury) or at the reduced form of the element undergoing the studied reaction with giving its state as amalgam. In case a reporting author states explicitly the use of an amalgam electrode the latter assignment is used.

References

48Ran Randles, J.E.B.: Trans. Faraday Soc. **44** (1948) 327.
48Sev Sevcik, A.: Collect. Czech. Chem. Commun. **13** (1948) 349.
50Ger Gerischer, H.: Z. Elektrochem. **54** (1950) 366.
51Dos2 Doss, K.S.G., Agarwal, H.P.: Proc. Indian Acad. Sci. Sect. A **34** (1951) 229.
51Dos3 Doss, K.S.G., Agarwal, H.P.: Proc. Indian Acad. Sci. Sect. A **34** (1951) 263.
51Dos4 Doss, K.S.G., Agarwal, H.P.: Proc. Indian Acad. Sci. Sect. A **33** (1951) 66.
52Dos2 Doss, K.S., Agarwal, H.P.: Proc. Indian Acad. Sci. Sect. A **35** (1952) 45.
53Del Delahay, P.: J. Am. Chem. Soc. **75** (1953) 1430.
53Kou Koutecky, J.: Collect. Czech. Chem. Commun. **18** (1953) 597.
54Hil Hillson, P.J.: Trans. Faraday Soc. **50** (1954) 385.
54Kam Kambara, T.: Bull. Chem. Soc. Jpn. **27** (1954) 529.
54Lor Lorenz, W.: Z. Elektrochem. **58** (1954) 912.
55Ber1 Berzins, T., Delahay, P.: J. Am. Chem. Soc. **77** (1955) 6448.
55Ber2 Berzins, T., Delahay, P.: J. Chem. Phys. **23** (1955) 972.
55Ber3 Berzins, T., Delahay, P.: Z. Elektrochem. **59** (1955) 792.
55Ger1 Gerischer, H.: Z. Elektrochem. **59** (1955) 604.
55Ger2 Gerischer, H., Vielstich, W.: Z. Phys. Chem. (Frankfurt/Main) **3** (1955) 16.
55Jor Jordan, J.: Anal. Chem. **27** (1955) 1708.
55Mat Matsuda, H., Ayabe, Y.: Z. Elektrochem. **59** (1955) 494.
55Web Weber, J., Koutecky, J.: Collect. Czech. Chem. Commun. **20** (1955) 980.
56Fro1 Fronaeus, S., Östman, C.O.: Acta Chem. Scand. **10** (1956) 769.
56Kou Koutecky, J.: Collect. Czech. Chem. Commun. **21** (1956) 433.
57Con1 Conway, B.E., Bockris, J.O'M.: J. Chem. Phys. **26** (1957) 532.
57Ger2 Gerischer, H., Krause, M.: Z. Phys. Chem. N. F. **10** (1957) 264.
57Old Oldham, K.B.: Trans. Faraday Soc. **53** (1957) 80.
57Vie Vielstich, W., Delahay, P.: J. Am. Chem. Soc. **79** (1957) 1874.
58Bar1 Barker, G.C., Faircloth, R.L., Gardner, A.W.: Nature (London) **181** (1958) 247.
58Bar2 Barker, G.C.: Anal. Chim. Acta **18** (1958) 118.
58Bau Bauer, H.H., Elving, P.: Anal. Chem. **30** (1958) 341.
58Mat Matsuda, H.: Z. Elektrochem. **62** (1958) 977.
59Khe Kheifets, V.L., Sheinin, A.B.: Zh. Fiz. Khim. **33** (1959) 1945.
60Bau Bauer, H.H., Smith, D.L., Elving, P.J.: J. Am. Chem. Soc. **82** (1960) 2094.
60Old Oldham, K.B.: J. Electrochem. Soc. **107** (1960) 766.
60Oza1 Ozaki, T., Nakayama, T.: J. Chem. Soc. Jpn. Pure Chem. Sec. (Nippon Kagaku Zasshi) **81** (1960) 90.
60Oza2 Ozaki, T., Nakayama, T.: J. Chem. Soc. Jpn. Pure Chem. Sec. (Nippon Kagaku Zasshi) **81** (1960) 98.
60Smi Smit, W.M., Wijnen, M.D.: Recl. Trav. Chim. Pays Bas **79** (1960) 5.
60Sta von Stackelberg , M.: Polarographische Arbeitsmethoden, Berlin: Walter de Gruyter & Co., 1960.
60Wij1 Wijnen, M.D., Smit, W.M.: Recl. Trav. Chim. Pays Bas **79** (1960) 22.
60Wij2 Wijnen, M.D., Smit, W.M.: Recl. Trav. Chim. Pays Bas **79** (1960) 203.
60Wij3 Wijnen, M.D., Smit, W.M.: Recl. Trav. Chim. Pays Bas **79** (1960) 289.
61Zin Zinov'ev, V.A., Sheinin, A.B., Kheifets, V.L.: Zh. Fiz. Khim. **35** (1961) 98.
62Del1 Delahay, P.: J. Phys. Chem. **66** (1962) 2204.
62Del2 Delahay, P.: Anal. Chem. **34** (1962) 1161.
62Del3 Delahay, P., Aramata, A.: J. Phys. Chem. **66** (1962) 2208.
62Jor Jordan, J., Javick, R.A.: Electrochim. Acta **6** (1962) 23.

62Lev Levich, V.G.: Physicochemical Hydrodynamics, Englewood Cliffs: Prentice-Hall Inc., 1962, p. 107; 314.

62Los1 Loshkarev, M.A., Tomilov, B.I.: Russ. J. Phys. Chem. (English Transl.) **36** (1962) 66; Zh. Fiz. Khim. **36** (1962) 132.

62Rei1 Reinmuth, W.H., Wilson, C.E.: Anal. Chem. **34** (1962) 1159.

62Rei2 Reinmuth, W.H.: Anal. Chem. **34** (1962) 1272.

63Bom Bombi, G.G., Fiorani, M.: Chim. Ind. (Milan) **45** (1963) 793.

63Fio Fiorani, M., Bombi, G.G., Marchetto, M.T.: Chim. Ind. (Milan) **45** (1963) 796.

63Ric Riccoboni, L., Bombi, G.G., Fiorani, M.: Chim. Ind. (Milan) **45** (1963) 786.

64Bon Bonnemay, M., Costa, M.: C. R. Acad. Sci. (Paris) **258** (1964) 2054.

64But1 Butler, J.N., Makrides, A.C.: Trans. Faraday Soc. **60** (1964) 1664.

64Suz Suzuki, J., Ozaki, T.: Bull. Chem. Soc. Jpn. **37** (1964) 230.

64Tan Tanaka, N., Tamamushi, R.: Electrochim. Acta **9** (1964) 963.

65Del2 Delahay, P.: Double Layer and Electrode Kinetics, New York: John Wiley&Sons, 1965, p. 171.

65Del3 Delahay, P.: J. Electroanal. Chem. **10** (1965) 1.

65Hey Heyrovsky, J., Kuta, J.: Grundlagen der Polarographie, Berlin: Akademie-Verlag, 1965.

65Nic Nicholson, R.S.: Anal. Chem. **37** (1965) 1351.

66Bau Bauer, H.H., Foo, D.C.S.: Aust. J. Chem. **19** (1966) 1103.

66But3 Butler, J.N., Meehan,M.L.: Trans. Faraday Soc. **62** (1966) 3524.

66Hey Heyrovsky, J., Kuta, J.: Principles of Polarography, New York: Academic Press, 1966.

66Hub Hubbard, A.T., Anson, F.C.: Anal. Chem. **38** (1966) 58.

66Tep1 Teppema, P., Sluyters-Rehbach, M., Sluyters, J.H.: Recl. Trav. Chim. Pays Bas **85** (1966) 303.

66Tep2 Teppema, P., Sluyters-Rehbach, M., Sluyters, J.H.: Recl. Trav. Chim. Pays Bas **85** (1966) 1161.

67Koo2 Kooijman, D.J., Sluyters, J.H.: Electrochim. Acta **12** (1967) 1579.

67Wei1 Weir, W.D., Enke, C.G.: J. Phys. Chem. **71** (1967) 275.

67Wei2 Weir, W.D., Enke, C.G.: J. Phys. Chem. **71** (1967) 280.

67deL de Leuwe , R., Sluyters-Rehbach, M., Sluyters, J.H.: Electrochim. Acta **12** (1967) 1593.

68Boc2 Bockris, J.O'M., Mannan, R.J., Damjanovich, A.: J. Chem. Phys. **48** (1968) 1898.

68Old Oldham, K.B., Parry, E.P.: Anal. Chem. **40** (1968) 65.

68Suz Suzuki, J.: Rev. Polarogr. **15** (1968) 21.

68Tep Teppema, P., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **16** (1968) 165.

69Dau Daum, P.H., Enke, C.G.: Anal. Chem. **41** (1969) 653.

69Hub Hubbard, A.T.: J. Electroanal. Chem. **22** (1969) 165.

69Suz Suzuki, J.: Bull. Chem. Soc. Jpn. **42** (1969) 3487.

69Tan1 Tanaka, N., Yamada, A.: Electrochim. Acta **14** (1969) 491.

70Her Herbelin, J.M., Andersen, T.N., Eyring, H.: Electrochim. Acta **15** (1970) 1455.

70Hub Hubbard, A.T., Anson, F.C.: Electroanalytical Chemistry, Vol. 4, Bard, A.J. (ed.), New York: Marcel Dekker, 1970.

70Khe Kheifets, V.L., Krasikov, B.S., Rotinyan, A.L.: Elektrokhimiya **6** (1970) 916.

70Mat Matsuda, A., Notoya, R.: J. Res. Inst. Catal. Hokkaido Univ. **18** (1970) 59.

70New Newman, J.: J. Electrochem. Soc. **117** (1970) 507.

70Suz Suzuki, J.: Bull. Chem. Soc. Jpn. **43** (1970) 755.

71Aga Agarwal, H.P.: Electrochim. Acta **16** (1971) 1395.

71Alb Albery, W.J., Hitchman, M.L: Ring-Disc Electrodes, Oxford: Clarendon Press, 1971.

72Hil2 Hill, D.L., Fung, K.W., Lee, H.L.: J. Electroanal. Chem. **38** (1972) 67.

72Hue Huebert, B.J., Smith, D.E.: Anal. Chem. **44** (1972) 1179.

72Lai Lai, C.-H., Hubbard, A.T.: Inorg. Chem. **11** (1972) 2081.

72Tra Trasatti, S.: J. Electroanal. Chem. **39** (1972) 163.

73Pol van der Pol, F., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **45** (1973) 377.

73Rod Rodgers, R.S., Anson, F.C.: J. Electroanal. Chem. **42** (1973) 381.

74Vit2 Vitanov, T., Popov, A., Budevski, E.: J. Electrochem. Soc. **121** (1974) 207.

75Tam Tamamushi, R.: Kinetic Parameters of Electrode Reactions of Metallic Compounds, London: Butterworth, 1975.

75Wea Weaver, M.J., Anson, F.C.: J. Electroanal. Chem. **65** (1975) 711.

76Ple Pleskov, Yu.V., Filinovskii V.Yu.: The Rotating Disc Electrode, New York: Consultants Bureau, 1976.

76Wea Weaver, M.J., Anson, F.C.: Inorg. Chem. **15** (1976) 1871.

77Tra Trasatti, S.: Advances in Electrochemistry and Electrochemical Engineering, Vol. 10, Gerischer, H., Tobias, C.W. (eds.), New York: John Wiley, 1977, p. 213.

78Ber Bernstein, C., Heindrichs, A., Vielstich, W.: J. Electroanal. Chem. **87** (1978) 81.

80Bar1 Barr, S., Guyer, K., Weaver, M.J.: J. Electroanal. Chem. **111** (1980) 41.

80Bar2 Bard, A.J., Faulkner L.R.: Electrochemical Methods, New York: Wiley, 1980.

81Ban Bancroft, E.E., Sidwell, J.S., Blount, H.N.: Anal. Chem. **53** (1981) 1390.

83Her Herrmann, J.: Ph.D. Dissertation, Bonn University, 1983.

83Iwa Iwasita, T., Schmickler, W., Herrmann, J., Vielstich, W.: J. Electrochem. Soc. **130** (1983) 2026.

83Kaw Kawiak, J., Jedral, T., Galus, Z: J. Electroanal. Chem. **145** (1983) 163.

84Bar1 Barz, F., Bernstein, Ch., Vielstich W.: Advances in Electrochemistry and Electrochemical Engineering, Vol. 13, Gerischer, H. (ed.), New York: John Wiley, 1984.

84Fil Filinovski, V.Yu., Pleskov, Yu.V.: Comprehensive Treatise of Electrochemistry, Yeager, E., Bockris, J.O'M., Conway, B.E., Sarangapani, S. (eds.), New York: Plenum Press, 1984, p. 293.

84Guy Guyer, K.L., Weaver, M.J.: Inorg. Chem. **23** (1984) 1664.

84Hei1 Heinemann, W.R., Hawkridge, F.M., Blount, H.N.: Electroanalytical Chemistry, Vol. 13, Bard, A.J. (ed.), New York: Marcel Dekker, 1984, p. 1.

84Hei2 Heinze, J.: Angew. Chem. **96** (1984) 823.

87Wru Wruck, W.J., Machaldo, R.M., Chapman, T.W.: J. Electrochem. Soc. **134** (1987) 539.

88Wip Wipf, D.O., Kristensen, E.W., Deakin, M.R., Wightman, R.M.: Anal. Chem. **60** (1988) 306.

89Bon Bontempelli, G., Magno, F., Mazzochin, G.A., Seeber, R.: Ann. Chim. (Rome) **79** (1989) 103.

90Wig Wightman, R.M., Wipf, D.O.: Acc. Chem. Res. **23** (1990) 64.

91Bar1 Baranski, A.S.: J. Electroanal. Chem. **300** (1991) 309.

92Saf Safford, L.K., Weaver, M.J.: J. Electroanal. Chem. **331** (1992) 857.

94Gal Galus, Z.: Fundamentals of Electrochemical Analysis, 2nd edition, Chichester: Ellis Horwood, 1994.

96Tor Torres, L.M., Gil, A.F., Galicia, L., Gonzalez, I.: J. Chem. Educ. **73** (1996) 808.

01Bar Bard, A.J., Faulkner, L.R.: Electrochemical Methods, 2nd ed., New York: Wiley, 2001.

01Liu Liu, X., Liu, K., Cheng, G., Dong, S.: J. Electroanal. Chem. **510** (2001) 103.

01Wea Weaver, M.J.: J. Electroanal. Chem. **498** (2001) 105.

05Nor Norskov, J.K., Bligaard, T., Logadottir, A., Kitchin, J.R., Chen, J.G., Pandelov, S., Stimming, U.: J. Electrochem. Soc. **152** (2005) J23.

Table 5.1. Exchange current densities and rate constants in aqueous systems.

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Aluminum (Al)	1 **N** CH_3COOH/H_2O	H^+/H_2	1	$10^{-9.28}$	CP	68Vij
	1 **N** CH_3COOH/H_2O	H^+/H_2	12.5	$10^{-8.86}$	CP	68Vij
	1 **N** CH_3COOH/H_2O	H^+/H_2	30.5	$10^{-8.82}$	CP	68Vij
	1 **N** CH_3COOH/H_2O	H^+/H_2	38.5	$10^{-8.6}$	CP	68Vij
	1 **N** CH_3COOH/H_2O	H^+/H_2	55.5	$10^{-8.15}$	CP	68Vij
Antimony (Sb)	7 **M** HCl/H_2O	$Sb^{3+/0}$, 0.1 **M** Sb^{3+}	25	*$1.98\cdot10^{-7}$*	CP	69Nor
	7 **M** HCl/H_2O	$Sb^{3+/0}$, 0.25 **M** Sb^{3+}	25	*$1.05\cdot10^{-7}$*	CP	69Nor
	7 **M** HCl/H_2O	$Sb^{3+/0}$, 0.5 **M** Sb^{3+}	25	*$0.74\cdot10^{-7}$*	CP	69Nor
	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-8.1}$	CP	69Pun
Sb_{Hg}	0.5 **M** H_2SO_4/H_2O +0.001 % gelatin	$Sb^{3+/0}$	25	$8.7\cdot10^{-5}$	DCP	63Fuz
Bismuth (Bi)	2 **M** HCl/H_2O	$Bi^{3+/0}$, $4.6\cdot10^{-5}$ **M**	20	$6.2\cdot10^{-6}$	CP	69Gor
	2 **M** HCl/H_2O	$Bi^{3+/0}$, $2\cdot10^{-2}$ **M**	20	$1.0\cdot10^{-4}$	CP	69Gor
	2 **M** HCl/H_2O	$Bi^{3+/2+}$, $4.6\cdot10^{-5}$ **M**	20	$3.7\cdot10^{-6}$	CP	69Gor
	2 **M** HCl/H_2O	$Bi^{3+/2+}$, $2\cdot10^{-2}$ **M**	20	$5.1\cdot10^{-4}$	CP	69Gor
	2 **M** HCl/H_2O	$Bi^{2+/0}$, $4.6\cdot10^{-5}$ **M**	20	$5.3\cdot10^{-6}$	CP	69Gor
	2 **M** HCl/H_2O	$Bi^{2+/0}$, $2\cdot10^{-2}$ **M**	20	$9.3\cdot10^{-5}$	CP	69Gor
	2 **M** HCl/H_2O	$Bi^{1+/0}$, $4.6\cdot10^{-5}$ **M**	20	$4.4\cdot10^{-5}$	CP	69Gor
	2 **M** HCl/H_2O	$Bi^{1+/0}$, $2\cdot10^{-2}$ **M**	20	$6.1\cdot10^{-5}$	CP	69Gor
	2 **M** HCl/H_2O	$Bi^{3+/0}$	20	$7.2\cdot10^{-4}$	CP	69Gor
	2 **M** HCl/H_2O	$Bi^{3+/0}$	30	$2.5\cdot10^{-3}$	CP	69Gor
	2 **M** HCl/H_2O	$Bi^{3+/0}$	40	$9.4\cdot10^{-3}$	CP	69Gor
$Bi_{(Hg)}$	1 **M** HCl/H_2O	$Bi^{3+/0}$	20	*>1*	IP	52Ran1
$Bi_{(Hg)}$	1 **M** HCl/H_2O	$Bi^{3+/0}$	26	*1.2..3.6*	IP	60Mou
$Bi_{(Hg)}$	1 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	20	*0.0003*	IP	52Ran1
$Bi_{(Hg)}$	1 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	26	*0.00037*	IP	60Mou
$Bi_{(Hg)}$	1 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, +1 m**M** Cl^-	26	*0.0028*	IP	60Mou
$Bi_{(Hg)}$	1 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, 2 m**M**	25	$1.61\cdot10^{-4}$	IP	64Joh
$Bi_{(Hg)}$	1 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, 4 m**M**	25	$1.98\cdot10^{-4}$	IP	64Joh
$Bi_{(Hg)}$	1 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, 8 m**M**	25	$2.55\cdot10^{-4}$	IP	64Joh
$Bi_{(Hg)}$	1 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, 17.7 m**M**	25	$3.51\cdot10^{-4}$	IP	64Joh
$Bi_{(Hg)}$	1 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, 38.4 m**M**	25	$4.55\cdot10^{-4}$	IP	64Joh [6])
$Bi_{(Hg)}$	1 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	25	*$3.0\cdot10^{-4}$*	CS	64Ham1
$Bi_{(Hg)}$	1 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	25	*$3.0\cdot10^{-4}$*	ACP	61Kam
$Bi_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	20	*$3.5\cdot10^{-5}$*	SS	66Gor
$Bi_{(Hg)}$ [10])	2 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, $4.3\cdot10^{-6}$ **M**	20	$2.8\cdot10^{-6}$	CP	68Gor1
$Bi_{(Hg)}$ [7])	2 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, 1 m**M**	–8	$4.3\cdot10^{-6}$	CP	71Gor
$Bi_{(Hg)}$ [7])	2 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, 1 m**M**	20	$3.2\cdot10^{-5}$	CP	71Gor
$Bi_{(Hg)}$ [7])	2 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, 1 m**M**	59	$1.9\cdot10^{-4}$	CP	71Gor
$Bi_{(Hg)}$ [7])	2 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, 1 m**M**	78	$3.8\cdot10^{-4}$	CP	71Gor
$Bi_{(Hg)}$ [8])	2 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, 0.1 **M**	20	$8.5\cdot10^{-5}$	RT	67Gor
$Bi_{(Hg)}$ [10])	2 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, 0.001 **M**	20	$2.6\cdot10^{-5}$	RT	66Gor
$Bi_{(Hg)}$ [9])	2 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, 0.0028 **M**	20	$1.0\cdot10^{-2}$	CP	63Gor

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Bi$_{(Hg)}$ [10])	12.5 **M** $HClO_4/H_2O$	$Bi^{3+/0}$, 0.0001 **M**	20	$7.2\cdot10^{-5}$	CP	68Gor2
Bi$_{(Hg)}$	2 **M** $HClO_4$ +0.25 m**M** $NaCl/H_2O$	$Bi^{3+/0}$, 0.001 m**M**	20	*1.0·10^{-5}*	CP	67Los
Bi$_{(Hg)}$	2 **M** $HClO_4$ +0.1 m**M** $NaCl/H_2O$	$Bi^{3+/0}$, 0.001 m**M**	20	$6.6\cdot10^{-5}$	RT	66Gor
Bi$_{(Hg)}$	2 **M** $HClO_4$ +1 m**M** $NaCl/H_2O$	$Bi^{3+/0}$	20	*1.4·10^{-3}*	RT	66Gor
Bi$_{(Hg)}$	2 **M** $HClO_4$ +1 **M** $NaClO_4/H_2O$	$Bi^{3+/0}$	25	*1.8·10^{-4}*	RT	63Fio
Bi$_{(Hg)}$	2 **M** $HClO_4$ +1 **M** $NaClO_4/H_2O$	$Bi^{3+/0}$	35	*2.5·10^{-4}*	RT	63Fio
Bi$_{(Hg)}$	2 **M** $HClO_4$ +1 **M** $NaClO_4/H_2O$	$Bi^{3+/0}$	45	*3.59·10^{-4}*	RT	63Fio
Bi$_{(Hg)}$	2 **M** $HClO_4$ +10^{-7} **M** NaI/H_2O	$Bi^{3+/0}$, 0.001 m**M**	20	$8.4\cdot10^{-5}$	RT	66Gor
Bi$_{(Hg)}$	2 **M** $HClO_4$ +10^{-8} **M** NaI/H_2O	$Bi^{3+/0}$, 0.001 m**M**	20	$4.7\cdot10^{-4}$	RT	66Gor
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	20	*3.5·10^{-5}*	RT	63Gor
Bi$_{(Hg)}$ [11])	2 **M** $HClO_4/H_2O$	$Bi^{2+/1+}$, 0.00078 **M**	20	$2.0\cdot10^{-5}$	CP	67Gor
Bi$_{(Hg)}$ [11])	2 **M** $HClO_4/H_2O$	$Bi^{2+/1+}$, 0.00078 **M**	20	$5.6\cdot10^{-5}$	CP	68Gor1
Bi$_{(Hg)}$ [10])	2 **M** $HClO_4/H_2O$	$Bi^{2+/1+}$, $4.3\cdot10^{-6}$ **M**	20	$2.0\cdot10^{-6}$...$5.9\cdot10^{-6}$	CP	68Gor1
Bi$_{(Hg)}$ [12])	2 **M** $HClO_4/H_2O$	$Bi^{2+/1+}$, 0.1 **M**	20	$4.3\cdot10^{-5}$	CP	67Gor
Bi$_{(Hg)}$ [13])	2 **M** $HClO_4/H_2O$	$Bi^{2+/0}$, 0.1 **M**	20	$6.9\cdot10^{-6}$	CP	67Gor
Bi$_{(Hg)}$ [13])	2 **M** $HClO_4/H_2O$	$Bi^{2+/0}$, 0.1 **M**	20	$3.5\cdot10^{-5}$	CP	68Gor1
Bi$_{(Hg)}$ [12])	2 **M** $HClO_4/H_2O$	$Bi^{3+/2+}$, 0.1 **M**	20	$2.1\cdot10^{-4}$	CP	67Gor
Bi$_{(Hg)}$ [13])	2 **M** $HClO_4/H_2O$	$Bi^{3+/2+}$, 0.1 **M**	20	$9.6\cdot10^{-5}$	CP	67Gor
Bi$_{(Hg)}$ [13])	2 **M** $HClO_4/H_2O$	$Bi^{3+/2+}$, 0.1 **M**	20	$8.6\cdot10^{-5}$	CP	68Gor1
Bi$_{(Hg)}$ [11])	2 **M** $HClO_4/H_2O$	$Bi^{3+/2+}$, 0.00078 **M**	20	$3.6\cdot10^{-5}$	RT	67Gor
Bi$_{(Hg)}$ [11])	2 **M** $HClO_4/H_2O$	$Bi^{3+/2+}$, 0.00078 **M**	20	$6.0\cdot10^{-5}$	CP	68Gor1
Bi$_{(Hg)}$ [10])	2 **M** $HClO_4/H_2O$	$Bi^{3+/2+}$, $4.3\cdot10^{-6}$ **M**	20	$6.8\cdot10^{-7}$...$2.0\cdot10^{-6}$	CP	68Gor1
Bi$_{(Hg)}$ [12])	2 **M** $HClO_4/H_2O$	$Bi^{3+/2+}$, 0.1 **M**	20	$9.6\cdot10^{-5}$	CP	68Gor1
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{3+/2+}$, 18/50 m**M** [14])	25	1.52	PS	66Ham
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{3+/2+}$, 9/50 m**M** [14])	25	0.9	PS	66Ham
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{3+/2+}$, 1.8/- m**M** [14])	25	0.26	PS	66Ham
Bi$_{(Hg)}$ [11])	2 **M** $HClO_4/H_2O$	$Bi^{1+/0}$, $7.8\cdot10^{-4}$ **M**	20	$1.4\cdot10^{-4}$	CP	68Gor1
Bi$_{(Hg)}$ [11])	2 **M** $HClO_4/H_2O$	$Bi^{1+/0}$, $7.8\cdot10^{-4}$ **M**	20	$1.3\cdot10^{-4}$	CP	67Gor
Bi$_{(Hg)}$ [10])	2 **M** $HClO_4/H_2O$	$Bi^{1+/0}$, $4.3\cdot10^{-6}$ **M**	20	$6.4\cdot10^{-4}$	CP	67Gor
Bi$_{(Hg)}$ [10])	2 **M** $HClO_4/H_2O$	$Bi^{1+/0}$, $4.3\cdot10^{-6}$ **M**	20	$5.6\cdot10^{-4}$	CP	68Gor1
Bi$_{(Hg)}$ [12])	2 **M** $HClO_4/H_2O$	$Bi^{1+/0}$, 0.1 **M**	20	$1.2\cdot10^{-6}$	CP	67Gor
Bi$_{(Hg)}$ [12])	2 **M** $HClO_4/H_2O$	$Bi^{1+/0}$, 0.1 **M**	20	$1.2\cdot10^{-6}$	CP	68Gor1
Bi$_{(Hg)}$ [12])	2 **M** $HClO_4/H_2O$	$Bi^{1+/0}$, 0.1 **M**	20	$3.2\cdot10^{-5}$	CP	67Gor

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{0/2+}$, -/1 m**M** [14])	25	0.1	PS	66Ham
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{0/2+}$, -/5 m**M** [14])	25	0.04	PS	66Ham
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{0/3+}$, 0.5/100 m**M** [14])	20	$8.5\cdot10^{-5}$	RT	67Gor
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{0/3+}$, 120/7.8 m**M** [14])	20	$7.2\cdot10^{-5}$	CP,RT	67Gor
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{3+/2+}$, $5\cdot10^{-4}$/0.1 **M** [14])	20	$2.1\cdot10^{-4}$	CP	67Gor
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{3+/2+}$, 0.12/$7.8\cdot10^{-4}$ **M** [14])	20	$3.6\cdot10^{-5}$	CP,RT	67Gor
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{3+/2+}$, $5\cdot10^{-6}$/0.1 **M** [14])	20	$9.6\cdot10^{-5}$	CP	67Gor
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{2+/0}$, $5\cdot10^{-4}$/0.1 **M** [14])	20	$4.3\cdot10^{-5}$	CP,RT	67Gor
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{2+/0}$, 0.12/$7.8\cdot10^{-4}$ **M** [14])	20	$2.0\cdot10^{-5}$	CP,RT	67Gor
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{2+/0}$, $5\cdot10^{-6}$/0.1 **M** [14])	20	$6.9\cdot10^{-6}$	CP	67Gor
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{+/0}$, $5\cdot10^{-4}$/0.1 **M** [14])	20	$1.2\cdot10^{-6}$	CP	67Gor
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{+/0}$, $5\cdot10^{-4}$/0.1 **M** [14])	20	$3.2\cdot10^{-5}$	CP	67Gor
Bi$_{(Hg)}$	2 **M** $HClO_4/H_2O$	$Bi^{+/0}$, 0.12/$7.8\cdot10^{-4}$ **M** [14])	20	$1.3\cdot10^{-4}$	CP	67Gor
Bi$_{(Hg)}$	1 **M** HNO_3/H_2O	$Bi^{3+/0}$	26	*0.0037*	IP	60Mou
Bi$_{(Hg)}$	0.5 **M** HNO_3/H_2O	$Bi^{3+/0}$	25	*$5.0\cdot10^{-3}$*	CS	64Ham1
Bi$_{(Hg)}$	0.5 **M** H_2SO_4/H_2O	$Bi^{3+/0}$	26	*0.0018*	IP	60Mou
Bi$_{(Pt)}$	2 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	20	$7.2\cdot10^{-4}$	CP	69Gri2
Bi$_{(Pt)}$	2 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	30	$2.5\cdot10^{-3}$	CP	69Gri2
Bi$_{(Pt)}$	2 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	40	$9.4\cdot10^{-3}$	CP	69Gri2
Bi	0.1 **M** $HClO_4/H_2O$	H^+/H_2	-	$10^{-8.1}$	CP	61Bic
Bi	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-10.8}$	CP	70Ten
Cadmium (Cd)	0.72 **m** $CdSO_4/H_2O$	Cd^{2+}	RT	$1.4\cdot10^{-3}$	IP	61Bro
(Cd)$_{(Hg)}$	1 **M** $HClO_4/H_2O$	Cd^{2+}	22	68	IP	70Sal1
Cd$_{(Hg)}$	1 **M** KBr/H_2O	Cd^{2+}	27	*0.91*	FR	71Aga
	1 **M** KCl/H_2O	Cd^{2+}	22	*2.9*	FR	58Bar1
Cd$_{(Hg)}$	1 **M** KCl/H_2O	Cd^{2+}	0	>12.0	IP	63Slu
Cd$_{(Hg)}$	1 **M** KCl/H_2O	Cd^{2+}	25	4 .. 8	IP	63Slu
Cd$_{(Hg)}$	1 **M** KCl/H_2O	Cd^{2+}	25	>5.0	GS	67Koo1
Cd$_{(Hg)}$	1 **M** KCl/H_2O	Cd^{2+}	25	*2.5*	FR	69deL[15])
Cd$_{(Hg)}$	1 **M** KCl/H_2O	$Cd^{2+/0}$, $1.5\cdot10^{-3}$	27	*1.25*	FR	71Aga
Cd$_{(Hg)}$	1 **M** KCl/H_2O	$Cd^{2+/0}$	0	*≈4...8*	IP	63Slu
Cd$_{(Hg)}$	1 **M** KCl/H_2O	$Cd^{2+/0}$	25	*>12*	IP	63Slu
Cd$_{(Hg)}$	1 **M** KCl/H_2O	$Cd^{2+/0}$	-	*≈5*	IP	61Ran
Cd$_{(Hg)}$	1 **M** KCl/H_2O	Cd^{2+}	0	700	IP	70Sal1
Cd$_{(Hg)}$	1 **M** KCl/H_2O	Cd^{2+}	-	700	IP	70Sal2

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Cd$_{(Hg)}$	1 **M** KCl/H_2O, pH=2	Cd^{2+}	27	*1.25*	FR	71Aga
Cd$_{(Hg)}$	1 **M** KCl/H_2O, pH=4	$Cd^{2+/0}$	25	*>5*	GM	67Koo1
Cd$_{(Hg)}$	1 **M** KF/H_2O	Cd^{2+}, 1.5 m**M**	27	*2.0*	FR	71Aga
Cd$_{(Hg)}$	1 **M** KF/H_2O	Cd^{2+}, 3 m**M**	27	*6.1*	FR	71Aga
Cd$_{(Hg)}$	1 **M** KI/H_2O	$Cd^{2+/0}$, $1.5\cdot10^{-3}$	27	*6.5...13.8*	FR	71Aga
Cd$_{(Hg)}$	1 **M** KI/H_2O	$Cd^{2+/0}$, $3\cdot10^{-3}$	27	*8.5...13.6*	FR	71Aga
Cd$_{(Hg)}$	1 **M** KNO_3/H_2O	Cd^{2+}	20	*~0.6*	IP	52Ran1
Cd$_{(Hg)}$	1 **M** KNO_3/H_2O, pH=2	Cd^{2+}	22	*6.3*	FR	58Bar1
Cd$_{(Hg)}$	1 **M** KNO_3/H_2O	Cd^{2+}	25	*~1.0*	ACP	60Bau
Cd$_{(Hg)}$	1 **M** KNO_3/H_2O	Cd^{2+}	25	*0.6*	ACP	61Kam
Cd$_{(Hg)}$	1 **M** KNO_3/H_2O	$Cd^{2+/0}$, $1.5\cdot10^{-3}$	28	*1.5*	FR	71Aga
Cd$_{(Hg)}$	1 **M** KNO_3/H_2O	$Cd^{2+/0}$, $3\cdot10^{-3}$	28	*3.1*	IP	71Aga
Cd$_{(Hg)}$	0.4 **M** K_2SO_4/H_2O	Cd^{2+}, 0.01 **M**	20	$5\cdot10^{-4}$	PS	54Lor
Cd$_{(Hg)}$	0.5 **M** K_2SO_4/H_2O	Cd^{2+}	0	5.5	IP	61Mül
Cd$_{(Hg)}$	1 **M** K_2SO_4/H_2O	Cd^{2+}	0	*0.2*	IP	57Del
Cd$_{(Hg)}$	1 **M** K_2SO_4/H_2O	Cd^{2+}	0	*0.2*	IP	58Del
Cd$_{(Hg)}$	1 **M** NaBr/H_2O	Cd^{2+}	-	30...100	IP	70Sal2
Cd$_{(Hg)}$	1 **M** $NaClO_4$/H_2O	Cd^{2+}	22	*2.3*	FR	58Bar1
Cd$_{(Hg)}$	1 **M** $NaClO_4$/H_2O	Cd^{2+}	0	40	IP	61Mül
Cd$_{(Hg)}$	1 **M** $NaClO_4$/H_2O	Cd^{2+}	25	*0.5*	GS	67Koo1
Cd$_{(Hg)}$	1 **M** $NaClO_4$/H_2O	Cd^{2+}	36	*1.2*	FR	71Aga
Cd$_{(Hg)}$	1 **M** $NaClO_4$/H_2O, pH=4	Cd^{2+}	25	*0.46*	GS	67Koo2
Cd$_{(Hg)}$	1 **M** $NaClO_4$/H_2O, pH=2.9	Cd^{2+}	23	*0.4*	IP	68Ham2
Cd$_{(Hg)}$	1 **M** $NaClO_4$/H_2O, pH=3.2	Cd^{2+}	23	*0.38*	IP	68Ham2
Cd$_{(Hg)}$	1 **M** $NaClO_4$/H_2O, pH=3.8	Cd^{2+}	23	*0.48*	IP	68Ham2
Cd$_{(Hg)}$	1 **M** $NaClO_4$/H_2O, pH=5.4...6.9	Cd^{2+}	23	*0.38*	IP	68Ham2
Cd$_{(Hg)}$	1 **M** $NaClO_4$/H_2O, pH=3	Cd^{2+}	25	*0.46*	FR	69deL[16])
Cd$_{(Hg)}$	1 **M** $NaClO_4$/H_2O, $c(Cd_{(Hg)})$=1.3... $5.7\cdot10^{-3}$ **M**	Cd^{2+}, $3.6\cdot10^{-5}$ **M**	23	0.16	IP	70Ham
Cd$_{(Hg)}$	0.5 **M** Na_2SO_4/H_2O	Cd^{2+}	20	*$4.2\cdot10^{-2}$*	IP	53Ger1[17])
Cd$_{(Hg)}$	0.5 **M** Na_2SO_4/H_2O	Cd^{2+}	20	*$4.2\cdot10^{-2}$*	IP	57Vie
Cd$_{(Hg)}$	0.5 **M** Na_2SO_4/H_2O	Cd^{2+}	20	*$4.5\cdot10^{-2}$*	GS	57Vie
Cd$_{(Hg)}$	0.5 **M** Na_2SO_4/H_2O	Cd^{2+}	20	*$2.6\cdot10^{-2}$*	PS	57Vie
Cd$_{(Hg)}$	1 **M** Na_2SO_4/H_2O	Cd^{2+}	27	4.31	GM	55Ber1
Cd$_{(Hg)}$	1 **M** Na_2SO_4/H_2O	Cd^{2+}	25	*0.55*	GS	67Koo1
Cd$_{(Hg)}$	1 **M** Na_2SO_4/H_2O, pH=4	Cd^{2+}	25	*$6.3\cdot10^{-2}$*	GS	67Koo1
Cd$_{(Hg)}$	1 **M** Na_2SO_4/H_2O	Cd^{2+}, 5 m**M**	25	0.065	GS	69van

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Cd$_{(Hg)}$	1 **M** Na_2SO_4/H_2O	Cd^{2+}, 5 m**M**	25	0.058	CS	69van
Cd	0.5 **N** H_2SO_4/H_2O	H^+/H_2	-	$10^{-10.77}$	CP	64Boc3
	1 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-11.6}$	CP	70Rot1
Carbon (C) [19])	1 **M** $CaCl_2/H_2O$	$Co(CN)_6^{3-/4-}$	-	$0.09\cdot10^{-3}$	IP	93Gao [18])
	1 **M** H_2SO_4/H_2O	$Ce^{4+/3+}$	-	*$5.3\cdot10^{-5}$*	CV	84Gro
	1 **M** H_2SO_4/H_2O	$Ce^{4+/3+}$	RT	*10^{-6}*	IP	99Evs [20])
	1 **M** $HClO_4/H_2O$	$Ce^{4+/3+}$	RT	*$4.8\cdot10^{-6}$*	CP, IP	02Fer1 [21])
	1 **M** $HClO_4/H_2O$	$Ce^{4+/3+}$	RT	*$1.48\cdot10^{-5}$*	CP, IP	02Fer2 [22])
	1 **M** H_2SO_4/H_2O	$Ce^{4+/3+}$	25	*$3.8\cdot10^{-4}$*	CP	63Ada
	1 **M** H_2SO_4/H_2O	$Ce^{4+/3+}$	-	*$3.2\cdot10^{-4}$*	CF	84Gro
	1 **M** H_2SO_4/H_2O	$Ce^{4+/3+}$	25	*$3.8\cdot10^{-4}$*	RDE	63Gal [46])
	1 **M** H_2SO_4/H_2O	$Ce^{4+/3+}$	RT	*$5.3\cdot10^{-8}$*	IP	98Evs2 [23])
	1 **M** H_2SO_4/H_2O	$Ce^{4+/3+}$	RT	*$1.7\cdot10^{-5}$*	IP	98Evs2 [24])
	1 **M** H_2SO_4/H_2O	$Ce^{4+/3+}$	RT	*$5.3\cdot10^{-8}$*	IP	98Ple [23])
	1 **M** H_2SO_4/H_2O	$Ce^{4+/3+}$	RT	*$6.5\cdot10^{-7}$*	IP	98Ple [25])
	1 **M** $LiCl/H_2O$	$Co(CN)_6^{3-/4-}$	-	$0.13\cdot10^{-3}$	IP	93Gao [18])
	1 **M** $MgCl_2/H_2O$	$Co(CN)_6^{3-/4-}$	-	$0.1\cdot10^{-3}$	IP	93Gao [18])
	0.01 **M** $NaCl/H_2O$	$Co(CN)_6^{3-/4-}$	-	$0.64\cdot10^{-3}$	IP	93Gao [18])
	0.1 **M** $NaCl/H_2O$	$Co(CN)_6^{3-/4-}$	-	$2.1\cdot10^{-3}$	IP	93Gao [18])
	1 **M** $NaCl/H_2O$	$Co(CN)_6^{3-/4-}$	-	$5.2\cdot10^{-3}$	IP	93Gao [18])
	1 **M** KCl/H_2O	$Cr^{3+/2+}$	-	*$1.27\cdot10^{-4}$*	CV	84Gro
	0.5 **M** KCl +0.1 **M** HCl/H_2O	$Cu^{2+/+}$	-	*$1.6...9.5\cdot10^{-4}$*	CF	84Gro
	0.2 **M** $HClO_4/H_2O$	$Eu^{2+/3+}$	-	*$8.4\cdot10^{-4}$*	CV	95Che [26])
	0.01 **M** HCl +0.09 **M** $HClO_4$ +3.9 **M** $NaClO_4/H_2O$	$Fe^{3+/2+}$	25	*$1.37\cdot10^{-5}$*	CP	63Gal [27])
	0.1 **M** HCl/H_2O	$Fe^{3+/2+}$	25	*$7.4...10\ 10^{-5}$*	CP	63Gal
	0.1 **M** HCl +3.8 **M** KCl/H_2O	$Fe^{3+/2+}$	25	*$2.29\cdot10^{-5}$*	CP	63Gal
	0.1 **M** HCl+3.9 **M** $LiCl/H_2O$	$Fe^{3+/2+}$	25	*$2.19\cdot10^{-5}$*	CP	63Gal
	0.1 **M** HCl +0.5 **M** NaCl +3.4 **M** $NaClO_4/H_2O$	$Fe^{3+/2+}$	25	*$1.41\cdot10^{-5}$*	CP	63Gal
	0.1 **M** HCl +3.9 **M** $NaCl/H_2O$	$Fe^{3+/2+}$	-	*$9.4\cdot10^{-4}$*	CP	63Ada [28])
	0.1 **M** HCl +3.9 **M** $NaCl/H_2O$	$Fe^{3+/2+}$	-	*$2.1\cdot10^{-4}$*	CP	63Ada [29])
	0.1 **M** HCl +3.9 **M** $NaCl/H_2O$	$Fe^{3+/2+}$	-	*$3.1\cdot10^{-4}$*	CP	63Ada [30])

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.1 **M** HCl +3.9 **M** $NaCl/H_2O$	$Fe^{3+/2+}$	-	*5.4·10^{-4}*	CP	63Ada [31])
	0.1 **M** HCl +3.9 **M** $NaCl/H_2O$	$Fe^{3+/2+}$	-	*5.8·10^{-4}*	CP	63Ada [32])
	0.1 **M** HCl +3.9 **M** $NaCl/H_2O$	$Fe^{3+/2+}$	-	*6.0·10^{-4}*	CP	63Ada [33])
	0.1 **M** HCl +3.9 **M** $NaCl/H_2O$	$Fe^{3+/2+}$	-	*7.3·10^{-4}*	CP	63Ada [34])
	0.1 **M** HCl +3.9 **M** $NaCl/H_2O$	$Fe^{3+/2+}$	-	*8.2·10^{-4}*	CP	63Ada [35])
	0.1 **M** HCl +0.1 **M** $NaClO_4/H_2O$	$Fe^{3+/2+}$	25	*8.5·10^{-5}*	CP	63Gal
	0.1 **M** HCl +3.9 **M** $NaClO_4/H_2O$	$Fe^{3+/2+}$	25	*1.41·10^{-5}*	CP	63Gal
	1 **M** HCl/H_2O	$Fe^{3+/2+}$	21	*1.2·10^{-4}*	DCP	70Sha1
	1 **M** $HClO_4/H_2O$	$Fe^{3+/2+}$	23	*(4.5·10^{-5})*	CV	00Gra [36])
	1 **M** $HClO_4/H_2O$	$Fe^{3+/2+}$	23	*(1.0·10^{-5})*	CV	00Gra [37])
	0.2 **M** $HClO_4/H_2O$	$Fe^{2+/3+}$	-	*2.5·10^{-3}*	CV	95Che [26])
	0.2 **M** $HClO_4/H_2O$	$Fe^{2+/3+}$	-	*2.1·10^{-6}*	CV	95Che [38])
	0.2 **M** $HClO_4/H_2O$	$Fe^{2+/3+}$	-	*1.6·10^{-5}*	CV	95Che [39])
	0.2 **M** $HClO_4/H_2O$	$Fe^{2+/3+}$	-	*1.1·10^{-3}*	CV	95Che [40])
	0.2 **M** $HClO_4/H_2O$	$Fe^{2+/3+}$	-	*7.4·10^{-8}*	CV	95Che [41])
	0.2 **M** $HClO_4/H_2O$	$Fe^{2+/3+}$	-	*3.8·10^{-4}*	CV	95Che [42])
	0.2 **M** $HClO_4/H_2O$	$Fe^{2+/3+}$	-	*9.1·10^{-5}*	CV	95Che [43])
	1 **M** $HClO_4/H_2O$	$Fe^{3+/2+}$	RT	*(1.0·10^{-6})*	CP, IP	02Fer1 [44])
	1 **M** $HClO_4/H_2O$	$Fe^{3+/2+}$	RT	(*0.6·10^{-5}*)	CP, IP	02Fer2 [45])
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*5.4·10^{-5}*	CP	63Ada
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*5.4·10^{-5}*	CP	63Gal
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*4.5·10^{-5}*	OC	63Gal
	0.1 **M** $NaClO_4$ +0.1 **M** HCl/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*8.5·10^{-5}*	RDE	63Gal [46])
	3.9 **M** $NaClO_4$ +0.01 **M** HCl +0.09 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	25	*1.37·10^{-4}*	RDE	63Gal [46])
	0.09 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	-	*1.8·10^{-3}*	CV	84Gro
	0.9 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	-	*1.34·10^{-3}*	CV	84Gro
	0.1 **M** $HClO_4$ +0.09 **M** HCl /H_2O	$Fe(H_2O)_6^{2+/3+}$	-	*5.7·10^{-3}*	CV	84Gro
	3.9 **M** $HClO_4$ +0.1 **M** HCl/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*1.41·10^{-4}*	RDE	63Gal [46])

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
C	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*5.4·10^{-5}*	RDE	63Gal [46])
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	-	*9.0·10^{-4}*	RDE	84Gro
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	-	*1.2·10^{-3}*	CV	84Gro
	0.45 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	-	*8.6·10^{-4}*	CV	84Gro
	3.4 **M** $NaClO_4$ +0.5 **M** NaCl +0.1 **M** HCl/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*1.41·10^{-4}*	RDE	63Gal [46])
	3.9 **M** $NaClO_4$ +0.09 **M** HCl/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*1.2·10^{-3}*	RDE	63Gal [46])
	3.8 **M** KCl +0.1 **M** HCl/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*2.29·10^{-4}*	RDE	63Gal [46])
	0.5 **M** potassium oxalate/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*1.4·10^{-3}*	CV	84Gro
	3.9 **M** LiCl +0.1 **M** HCl/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*2.19·10^{-4}*	RDE	63Gal [46])
	-/H_2O	$Fe(CN)_6^{3-/4-}$	20	0.58	RDE	61Reg
	0.1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	*1.48·10^{-3}*	RDE	63Gal [46])
	0.3 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	*2.65·10^{-3}*	RDE	63Gal [46])
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	*7.08·10^{-3}*	RDE	63Gal [46])
	1 **M** $LiCl/H_2O$	$Fe(CN)_6^{3-/4-}$	25	*1.06·10^{-3}*	RDE	63Gal [46])
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	-	*5.3·10^{-2}*	ACP	84Gro
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	-	*7.0·10^{-2}*	CV	84Gro
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	*0.1*	CV	95Che [26])
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	*0.037*	CV	95Che [38])
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	*0.026*	CV	95Che [39])
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	*0.084*	CV	95Che [40])
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	*0.047*	CV	95Che [41])
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	*0.078*	CV	95Che [42])
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	*0.076*	CV	95Che [43])
	0.1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	23	*9.0·10^{-4}*	CV	95Ale [47])
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	23	*(0.017)*	CV	00Gra [36])
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	23	*(0.019)*	CV	00Gra [37])
	0.5 **M** K_2SO_4/H_2O	$Fe(CN)_6^{3-/4-}$	25	*0.025*	CV	84Gro
	0.5 **M** K_2SO_4/H_2O	$Fe(CN)_6^{3-/4-}$	22	*0.015*	CV	84Gro [48])

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	phosphate buffer/ H_2O, pH=7.5	$Fe(CN)_6^{3-/4-}$	-	*0.02*	TPF	84Gro
	0.1 **M** THAP [49]) /H_2O+ethanol	ferrocene/ ferrocinium	-	*1.3·10^{-2}*	CV	84Gro
	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	1·10^{-10}	CP	96Mar [44])
	1 **M** H_2SO_4/H_2O	$Hg^{2+/+}$	-	*10^{-7}*	CF	84Gro
	1 **M** NaOH/H_2O	IO_3^-/I^-	-	(*2.1·10^{-19}*)	CV	84Gro
	1 **M** KCl/H_2O	$IrCl_6^{3-/4-}$	25	*0.5*	CV	95Che [26])
	1 **M** KCl/H_2O	$IrCl_6^{3-/4-}$	25	*0.11*	CV	95Che [38])
	1 **M** KCl/H_2O	$IrCl_6^{3-/4-}$	25	*0.21*	CV	95Che [39])
	1 **M** KCl/H_2O	$IrCl_6^{3-/4-}$	25	*0.25*	CV	95Che [41])
	1 **M** KCl/H_2O	$IrCl_6^{3-/4-}$	25	*0.22*	CV	95Che [42])
	1 **M** KCl/H_2O	$IrCl_6^{3-/4-}$	25	*0.2*	CV	95Che [43])
	0.1 **M** KCl/H_2O	$IrCl_6^{2-/3-}$	23	*0.004*	CV	95Ale [47])
	1 **M** KCl/H_2O	$IrCl_6^{2-/3-}$	23	*(0.01)*	CV	00Gra [36])
	1 **M** KCl/H_2O	$IrCl_6^{2-/3-}$	23	*(0.308)*	CV	00Gra [37])
	1 **M** KOH/H_2O	$MnO_4^{-/2-}$	20	2.1·10^{-2}	CP,GS	67Thi1
	0.9 **M** LiOH/H_2O	$MnO_4^{-/2-}$	-	*6.3·10^{-3}*	RDE	62Gal [46])
	0.9 **M** LiOH/H_2O	$MnO_4^{-/2-}$	25	*6.92·10^{-3}*	RDE	63Gal [46])
	0.9 **M** NaOH/H_2O	$MnO_4^{-/2-}$	-	*1.4·10^{-2}*	RDE	62Gal [46])
	0.9 **M** NaOH/H_2O	$MnO_4^{-/2-}$	25	*1.45·10^{-2}*	RDE	63Gal [46])
	0.9 **M** LiOH/H_2O	$MnO_4^{-/2-}$, 0.1 m**M**	-	*4.45·10^{-3}*	CP	70Sha1
	1.33 **M** NaOH +1.95 **M** $NaNO_3/H_2O$	NO_3^-/NO	25	3.6·10^{-12}	CV	96Boc [50])
	0.5 **M** H_2SO_4/H_2O	H_2O/O_2	-	5·10^{-11}	CP	96Mar [44])
	0.1 **M** DMSO [51]) /TBAP [52])	O_2/O_2^-	25	*9.3·10^{-2}*	CV	03Ort
	1 **M** KCl/H_2O	$Ru(NH_3)_6^{3+/2+}$	25	*0.24*	CV	95Che [26])
	1 **M** KCl/H_2O	$Ru(NH_3)_6^{3+/2+}$	25	*0.039*	CV	95Che [38])
	1 **M** KCl/H_2O	$Ru(NH_3)_6^{3+/2+}$	25	*0.13*	CV	95Che [39])
	1 **M** KCl/H_2O	$Ru(NH_3)_6^{3+/2+}$	25	*0.15*	CV	95Che [40])
	1 **M** KCl/H_2O	$Ru(NH_3)_6^{3+/2+}$	25	*0.15*	CV	95Che [41])
	1 **M** KCl/H_2O	$Ru(NH_3)_6^{3+/2+}$	25	*0.093*	CV	95Che [42])
	1 **M** KCl/H_2O	$Ru(NH_3)_6^{3+/2+}$	25	*0.13*	CV	95Che [43])

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
C	0.1 **M** KCl/H_2O	$Ru(NH_3)_6^{3+/2+}$	23	*0.003*	CV	95Ale [47])
	1 **M** KCl/H_2O	$Ru(NH_3)_6^{3+/2+}$	23	*(0.012)*	CV	00Gra [36])
	1 **M** KCl/H_2O	$Ru(NH_3)_6^{3+/2+}$	23	*(0.017)*	CV	00Gra [37])
	0.5 **M** $TBAPF_6$ /acetonitrile [52])	$(Ru(Cl_2dpy)_2(NO_2)(NO))^{+/2+}$	-	*0.06*	CV	84Gro
	0.5 **M** $TBAPF_6$ /acetonitrile [52])	$(Ru(dpy)_2(NO_2)(NO))^{+/2+}$	-	*0.19*	CF	84Gro
	0.5 **M** $TBAPF_6$ /acetonitrile [52])	$(Ru((C_6H_5)_{2)}dpy)_2(NO_2)(NO))^{+/2+}$	-	*0.09*	CF	84Gro
	0.5 **M** $TBAPF_6$ /acetonitrile [52])	$(Ru((C_4H_9)_{2)}dpy)_2(NO_2)(NO))^{+/2+}$	-	*0.07*	CF	84Gro
	0.5 **M** $TBAPF_6$ /acetonitrile [52])	$(Ru((CH_3)_{2)}dpy)_2(NO_2)(NO))^{+/2+}$	-	*0.15*	CF	84Gro
	0.5 **M** $TBAPF_6$ /acetonitrile [52])	$(Ru((OCH_3)_{2)}dpy)_2(NO_2)(NO))^{+/2+}$	-	*0.12*	CF	84Gro
	9.5 **M** HCl/H_2O	$Sb^{5+/3+}$	-	*$4.0\cdot10^{-4}$*	CF	84Gro
	6 **M** HCl/H_2O	$SbCl_6^-$(r)	-	*$2.0\cdot10^{-3}$*	CF	84Gro
	6 **M** HCl/H_2O	$SbCl_6^-$(r)	-	*$2.0\cdot10^{-3}$*	CV	84Gro
	12 **M** HCl/H_2O	$SbCl_6^-$(r)	-	*$2.0\cdot10^{-3}$*	CF	84Gro
	12 **M** HCl/H_2O	$SbCl_6^-$(r)	-	*$1.0\cdot10^{-3}$*	CF	84Gro
	4 **M** HCl+1 **M** Hbr +9.5 **M** LiBr/H_2O	$Sn^{4+/2+}$(ox)	-	*$2.0\cdot10^{-4}$*	CF	84Gro
	4 **M** HCl+1 **M** Hbr +9.5 **M** LiBr/H_2O	$Sn^{4+/2+}$(r)	-	*$5.6\cdot10^{-4}$*	CF	84Gro
	1 **M** KCl/H_2O	$Tl^{3+/1+}$	21	*$8.7\cdot10^{-6}$*	DCP	70Sha1
	1 **M** NaOH/H_2O	$Tl^{3+/1+}$	21	*$5.75\cdot10^{-6}$*	DCP	70Sha1
	0.2 **M** $HClO_4/H_2O$	$V_{aq}^{2+/3+}$	25	*$6.8\cdot10^{-4}$*	CV	95Che [26])
	0.2 **M** $HClO_4/H_2O$	$V_{aq}^{2+/3+}$	25	*$2.0\cdot10^{-5}$*	CV	95Che [38])
	0.2 **M** $HClO_4/H_2O$	$V_{aq}^{2+/3+}$	25	*$1.0\cdot10^{-6}$*	CV	95Che [39])
	0.2 **M** $HClO_4/H_2O$	$V_{aq}^{2+/3+}$	25	*$7.8\cdot10^{-5}$*	CV	95Che [40])
	0.2 **M** $HClO_4/H_2O$	$V_{aq}^{2+/3+}$	25	*$2.1\cdot10^{-3}$*	CV	95Che [41])
	0.2 **M** $HClO_4/H_2O$	$V_{aq}^{2+/3+}$	25	*$2.9\cdot10^{-6}$*	CV	95Che [42])
	0.2 **M** $HClO_4/H_2O$	$V_{aq}^{2+/3+}$	25	*$1.5\cdot10^{-6}$*	CV	95Che [43])
	1 **M** H_2SO_4/H_2O	$V^{5+/4+}$	-	($1.5\cdot10^{-3}$)	CV	91Kan
	2 **M** H_2SO_4/H_2O	$V^{5+/4+}$	-	($1.9\cdot10^{-3}$)	CV	91Kan
	3 **M** H_2SO_4/H_2O	$V^{5+/4+}$	-	*($5.3\cdot10^{-3}$)*	CV	91Kan
	4 **M** H_2SO_4/H_2O	$V^{5+/4+}$	-	*($6.1\cdot10^{-3}$)*	CV	91Kan
	1 **M** H_2SO_4/H_2O	$V^{4+/5+}$	-	*($0.91\cdot10^{-3}$)*	CV	91Kan

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	2 **M** H_2SO_4/H_2O	$V^{4+/5+}$	-	*(2.1·10^{-3})*	CV	91Kan
	3 **M** H_2SO_4/H_2O	$V^{4+/5+}$	-	*(6.1·10^{-3})*	CV	91Kan
	4 **M** H_2SO_4/H_2O	$V^{4+/5+}$	-	*(7.7·10^{-3})*	CV	91Kan
	1 **M** H_2SO_4/H_2O	$V^{2+/3+}$	-	*(3.2·10^{-3})*	CV	91Kan
	2 **M** H_2SO_4/H_2O	$V^{2+/3+}$	-	*(4.7·10^{-3})*	CV	91Kan
	3 **M** H_2SO_4/H_2O	$V^{2+/3+}$	-	*(7.7·10^{-3})*	CV	91Kan
	1 **M** H_2SO_4/H_2O	$V^{3+/2+}$	-	*(4.8·10^{-3})*	CV	91Kan
	2 **M** H_2SO_4/H_2O	$V^{3+/2+}$	-	*(5.6·10^{-3})*	CV	91Kan
	3 **M** H_2SO_4/H_2O	$V^{3+/2+}$	-	*(9.7·10^{-3})*	CV	91Kan
	0.1 **M** TEAP /acetonitrile	1,1'dimethyl2,2'-dibenzylhydrazine (ox, r)	-	*1.8·10^{-3}*	CV	84Gro
	0.05 **M** $LiClO_4$ /ethanol	vitamin $D_2^{0/+}$	20	*4.3·10^{-4}*	CD	01Liu
	0.05 **M** acetate buffer /H_2O, pH=4.6	cyano-cobalamin$^{0/-}$	-	*6.0·10^{-4}*	ACP	84Gro
	0.05 **M** acetate buffer /H_2O, pH=4.6	cyano-cobalamin$^{-/2-}$	-	*6.0·10^{-3}*	ACP	84Gro
	0.05 **M** carbonate buffer/H_2O, pH=10	cyano-cobalamin$^{0/-}$	-	*<·10^{-4}*	ACP	84Gro
	0.05 **M** carbonate buffer/H_2O, pH=10	cyano-cobalamin$^{-/2-}$	-	*4.0·10^{-3}*	ACP	84Gro
	0.05 **M** acetate buffer/ H_2O, pH=4.6	aquo-cobalamin$^{0/-}$	-	*5.0·10^{-3}*	ACP	84Gro
	0.05 **M** acetate buffer/ H_2O, pH=4.6	aquo-cobalamin$^{-/2-}$	-	*1.0·10^{-2}*	ACP	84Gro
	0.05 **M** carbonate buffer/H_2O, pH=10	hydroxy-cobalamin$^{0/-}$	-	*<·10^{-3}*	ACP	84Gro
	0.05 **M** carbonate buffer/H_2O, pH=10	hydroxy-cobalamin$^{-/2-}$	-	*6.0·10^{-3}*	ACP	84Gro
	0.1 **M** KCl/H_2O	dopamine	23	<10^{-6}	CV	95Ale [47])
	1 **M** $HClO_4/H_2O$	dopamine	23	*(7.3·10^{-5})*	CV	00Gra [36])
	1 **M** $HClO_4/H_2O$	dopamine	23	*(3.2·10^{-4})*	CV	00Gra [37])
	1 **M** $HClO_4/H_2O$	methylviologen$^{2+/+}$	23	*(0.247)*	CV	00Gra [37])
	1 **M** $HClO_4/H_2O$	chlorpromazine$^{0/+}$	23	*(0.023)*	CV	00Gra [37])
	0.1 **M** KCl/H_2O	4-methylcatechol	23	<10^{-6}	CV	95Ale [47])
	1 **M** $HClO_4/H_2O$	4-methylcatechol	23	*(5.3·10^{-4})*	CV	00Gra [37])
	0.1 **M** KCl/H_2O	hydroquinone	23	<10^{-6}	CV	95Ale [47])
	1 **M** H_2SO_4/H_2O	quinone /hydro-quinone	RT	*0.8·10^{-3}*	IP	98Evs1 [53])
	Britton&Robinson buffer/H_2O	*p*-phenylenediamine	-	*1.9·10^{-3}*	CV	84Bal

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Cesium (Cs)						
Cs$_{(Hg)}$	0.1 **M** NEt_4PF_6 /DMF [54])	Cs^{1+}	25	*1.1*	IP/PP	80Bar3
Chromium (Cr)	1 **N** H_2SO_4/H_2O	H^+/H_2	25	$0.35\cdot10^{-6}$	CP	69Wil [55])
	1 **N** H_2SO_4/H_2O	H^+/H_2	38	$5.0\cdot10^{-6}$	CP	69Wil [55])
	1 **N** H_2SO_4/H_2O	H^+/H_2	45	$22.0\cdot10^{-6}$	CP	69Wil [55])
	1 **N** H_2SO_4/H_2O	H^+/H_2	25	$5.3\cdot10^{-9}$	CP	69Wil [56])
	1 **N** H_2SO_4/H_2O	H^+/H_2	38	$1.0\cdot10^{-8}$	CP	69Wil [56])
	1 **N** H_2SO_4/H_2O	H^+/H_2	45	$1.7\cdot10^{-8}$	CP	69Wil [56])
Cobalt (Co)	0.1 **N** $CoCl_2/H_2O$	$Co^{2+/0}$	25	$2.8\cdot10^{-7}$	CS	61She
	0.5 **N** $CoCl_2/H_2O$	$Co^{2+/0}$	25	$4.2\cdot10^{-7}$	CS	61She
	1 **N** $CoCl_2/H_2O$	$Co^{2+/0}$	25	$6.8\cdot10^{-7}$	CS	61She
	2 **N** $CoCl_2/H_2O$	$Co^{2+/0}$	25	$8.4\cdot10^{-7}$	CS	61She
	$CoSO_4/H_2O$	$Co^{2+/0}$	25	$3\cdot10^{-5}$	OCS	61Zin
	2 **N** $CoSO_4/H_2O$	$Co^{2+/0}$	25	$1.3\cdot10^{-5}$	CS	61She
	2 **N** $CoSO_4$+0.002 **N** $CoCl_2/H_2O$	$Co^{2+/0}$	25	$5.2\cdot10^{-6}$	CS	61She
	2 **N** $CoSO_4$+0.02 **N** $CoCl_2/H_2O$	$Co^{2+/0}$	25	$1.1\cdot10^{-6}$	CS	61She
	2 **N** $CoSO_4$ +0.2 **N** $CoCl_2/H_2O$	$Co^{2+/0}$	25	$4.8\cdot10^{-7}$	CS	61She
	2 **N** $CoSO_4$ +0.02 **N** KI/H_2O	$Co^{2+/0}$	25	$1.0\cdot10^{-7}$	CS	61She
Copper (Cu)	0.7 **M** LiCl+0.8 **M** $AlCl_3$/propylene-carbonate	$Cu^{2+/0}$ [57])	25	$4.5\cdot10^{-5}$	CP	70Eis
	0.001 **M** $Cu(NO_3)_2/H_2O$	$Cu^{2+/0}$	20	$1.0\cdot10^{-9}$	CP	43Phi
	0.001 **M** $Cu(NO_3)_2/H_2O$	$Cu^{2+/0}$	20	$1.0\cdot10^{-9}$	CP	52Con
	0.01 **M** $Cu(NO_3)_2/H_2O$	$Cu^{2+/0}$	20	$1.0\cdot10^{-11}$	CP	43Phi
	0.01 **M** $Cu(NO_3)_2/H_2O$	$Cu^{2+/0}$	20	$1.0\cdot10^{-11}$	CP	52Con
	0.1 **M** $Cu(NO_3)_2/H_2O$	$Cu^{2+/0}$	20	$1.0\cdot10^{-11}$	CP	43Phi
	0.1 **M** $Cu(NO_3)_2/H_2O$	$Cu^{2+/0}$	20	$1.0\cdot10^{-11}$	CP	52Con
	1 **M** $Cu(NO_3)_2/H_2O$	$Cu^{2+/0}$	20	$1.0\cdot10^{-10}$	CP	43Phi
	1 **M** $Cu(NO_3)_2/H_2O$	$Cu^{2+/0}$	20	$1.0\cdot10^{-10}$	CP	52Con
	0.15 **N** $CuSO_4$ +1 **N** H_2SO_4/H_2O	$Cu^{2+/0}$	30	0.136 [58])	GM	59Mat1
	0.15 **N** $CuSO_4$ +1 **N** H_2SO_4/H_2O	$Cu^{0/2+}$	30	0.052 [58])	GM	59Mat1
	0.8 **M** $CuSO_4$+1.02 **M** H_2SO_4/H_2O	$Cu^{0/2+}$	30	$7.08\cdot10^{-3}$	CH	68Che

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.05 **M** $CuSO_4$ +1 **M** KCl +2 **M** NH_3/H_2O	$Cu^{2+/0}$	-	0.06	IP	59Lor
	0.05 **M** $CuSO_4$ +1 **M** H_2SO_4/H_2O	$Cu^{2+/0}$	-	0.03	IP	59Lor
	0.001 **M** $CuSO_4$ +1 **M** KNO_3/H_2O	$Cu^{0/2+}$, 0.001 **M**	-	2.3	FR	54Hil
	0.025 **M** $CuSO_4$ +1 **M** KNO_3/H_2O	$Cu^{0/2+}$, 0.025 **M**	-	3.2	FR	54Hil
	0.05 **M** $CuSO_4$ +1 **M** KNO_3/H_2O	$Cu^{0/2+}$, 0.05 **M**	-	3.9	FR	54Hil
	0.025 **M** $CuSO_4$ +0.05 **M** H_2SO_4 +1 **M** KNO_3/H_2O	$Cu^{2+/0}$	-	3.9	IP	54Hil
	0.025 **M** $CuSO_4$ +0.05 **M** H_2SO_4 +1 **M** KNO_3 +0.0025 % cetyl trimethyl ammonium bromide/H_2O	$Cu^{2+/0}$	-	5.2	IP	54Hil
	0.025 **M** $CuSO_4$ +0.05 **M** H_2SO_4 +1 **M** KNO_3 +0.0125 % cetyl trimethyl ammonium bromide/H_2O	$Cu^{2+/0}$	-	7.6	IP	54Hil
	0.025 **M** $CuSO_4$ +0.05 **M** H_2SO_4 +1 **M** KNO_3 +0.05 % cetyl trimethyl ammonium bromide/H_2O	$Cu^{2+/0}$	-	3.5	IP	54Hil
	0.025 **M** $CuSO_4$ +0.05 **M** H_2SO_4 +1 **M** KNO_3 +0.0025 % sodium dodecyl sulphate/H_2O	$Cu^{2+/0}$	-	10.0	IP	54Hil
	0.025 **M** $CuSO_4$ +0.05 **M** H_2SO_4 +1 **M** KNO_3 +0.0125 % sodium dodecyl sulphate/H_2O	$Cu^{2+/0}$	-	23.0	IP	54Hil

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Cu	0.01 **M** $CuSO_4$ +0.05 **M** H_2SO_4 +1 **M** KNO_3 +0.0025 % gelatin/H_2O	$Cu^{2+/0}$	-	2.1	IP	54Hil
	0.01 **M** $CuSO_4$ +0.05 **M** H_2SO_4 +1 **M** KNO_3 +0.0125 % gelatin/H_2O	$Cu^{2+/0}$	-	2.0	IP	54Hil
	0.01 **M** $CuSO_4$ +0.05 **M** H_2SO_4 +1 **M** KNO_3 +0.05 % gelatin/H_2O	$Cu^{2+/0}$	-	2.0	IP	54Hil
	0.025 **M** $CuSO_4$ +0.05 **M** H_2SO_4 +1 **M** KNO_3 +0.0125 % gelatin/H_2O	$Cu^{2+/0}$	-	3.8	IP	54Hil
	0.025 **M** $CuSO_4$+ 0.05 **M** H_2SO_4 +1 **M** KNO_3+0.025 % gelatin/H_2O	$Cu^{2+/0}$	-	3.7	IP	54Hil
	0.025 **M** $CuSO_4$+0.05 **M** H_2SO_4+1 **M** KNO_3+0.05 % gelatin/H_2O	$Cu^{2+/0}$	-	4.6	IP	54Hil
	0.05 **M** $CuSO_4$+0.05 **M** H_2SO_4 +1 **M** KNO_3 +0.0125 % gelatin/H_2O	$Cu^{2+/0}$	-	3.8	IP	54Hil
	0.05 **M** $CuSO_4$ +0.05 **M** H_2SO_4 +1 **M** KNO_3 +0.025 % gelatin/H_2O	$Cu^{2+/0}$	-	3.2	IP	54Hil
	0.05 **M** $CuSO_4$ +0.05 **M** H_2SO_4 +1 **M** KNO_3 +0.05 % gelatin/H_2O	$Cu^{2+/0}$	-	3.2	IP	54Hil
	0.05 **M** H_2SO_4/H_2O	$Cu^{2+/0}$	0	$2.3 \cdot 10^{-3}$	GM	61Hur1, 61Hur2
	0.05 **M** H_2SO_4/H_2O	$Cu^{2+/0}$	15	$1.2 \cdot 10^{-2}$	GM	61Hur1, 61Hur2

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.05 **M** H_2SO_4/H_2O	$Cu^{2+/0}$	20	$2.4\cdot10^{-2}$	GM	61Hur1, 61Hur2
	0.05 **M** H_2SO_4/H_2O	$Cu^{2+/0}$	25	$8.1\cdot10^{-2}$	GM	61Hur1, 61Hur2
	0.05 **M** H_2SO_4/H_2O	$Cu^{2+/0}$	35	$5.4\cdot10^{-2}$	GM	61Hur1, 61Hur2
	0.05 **M** H_2SO_4/H_2O	$Cu^{2+/0}$	45	$1.4\cdot10^{-1}$	GM	61Hur1, 61Hur2
	0.7 **M** LiCl +0.8 **M** $AlCl_3$ /propylene-carbonate	$Cu^{2+/0}$ [59])	25	$4.5\cdot10^{-5}$	CP	70Eis
	0.08 **M** CN^-/H_2O, (I=2 **M**)	$Cu^{1+/0}$, 0.01 **M**	31	$6.5\cdot10^{-3}$	CP	62Bon
	0.10 **M** CN^-/H_2O, (I=2 **M**)	$Cu^{1+/0}$, 0.01 **M**	31	$5.5\cdot10^{-3}$	CP	62Bon
	0.12 **M** CN^-/H_2O, (I=2 **M**)	$Cu^{1+/0}$, 0.01 **M**	31	$2.9\cdot10^{-3}$	CP	62Bon
	0.001 **N** HCl/H_2O	H^+/H_2	-	$10^{-7.56}$	CP	57Pen
	0.01 **N** HCl/H_2O	H^+/H_2	-	$10^{-7.25}$	CP	57Pen
	0.1 **N** HCl/H_2O	H^+/H_2	-	$10^{-6.84}$	CP	57Pen
	1 **M** KOH/H_2O	$MnO_4^{-/2-}$	20	$2.6\cdot10^{-2}$	CP,GS	67Thi1
Cu$_{(Hg)}$ [60])	5.0 **M** $HClO_4$ +10^{-5} **M** $Cu(ClO_4)_2/H_2O$	$Cu^{2+/0}$	20	*$2.5\cdot10^{-6}$*	CP	66Los
Cu$_{(Hg)}$ [61])	5.0 **M** $HClO_4$ +$2\cdot10^{-3}$ **M** $Cu(ClO_4)_2/H_2O$	$Cu^{2+/0}$	20	*$1.5\cdot10^{-6}$*	CP	66Los
Cu$_{(Hg)}$ [62])	5.0 **M** $HClO_4$ +10^{-5}...0.1 **M** $Cu(ClO_4)_2/H_2O$	$Cu^{2+/1+}$	20	*$5.7\cdot10^{-4}$*	CP	66Los
Cu$_{(Hg)}$ [62])	5.0 **M** $HClO_4$ +10^{-5}...0.1 **M** $Cu(ClO_4)_2/H_2O$	$Cu^{1+/0}$	20	*$4\cdot10^{-4}$*	CP	66Los
	0.01 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-7.2}$	CP	68Gos
	0.1 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-7.25}$	CP	68Gos
	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	$2.3\cdot10^{-8}$	CP	65Iof
	1 **M** H_2SO_4/H_2O	H^+/H_2	RT	$1.6\cdot10^{-7}$	CP	63Gos
	1 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-7.25}$	CP	68Gos
	1.5 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-6.97}$	CP	68Gos
	3 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-7.0}$	CP	68Gos
	5 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-6.75}$	CP	68Gos
	0.01 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-6.09}$	CP	57Pen
	1 **M** KF/H_2O	$[Ru(NH_3)_6]^{3+/2+}$	20	113	TPF	85Iwa
Cu$_{(Hg)}$	1.0 **M** KNO_3/H_2O	$Cu^{2+/0}$	20	*0.045*	IP	52Ran1

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Cu$_{(Hg)}$	1.0 **M** en [63]) HCl-en/H_2O	$Cuen_2^{2+}$	20	0.037	IP	52Ran1
Cu$_{(Hg)}$	1.0 **M** NH_4OH+NH_3/H_2O	$Cu(NH_3)_2^{2+}$	20	*0.011*	IP	52Ran1
	0.1 **N** HCl/H_2O	H^+/H_2	10	$4.26\cdot10^{-7}$	CP	59Boc [64])
	0.1 **N** HCl/H_2O	H^+/H_2	20	$5.49\cdot10^{-7}$	CP	59Boc [64])
	0.1 **N** HCl/H_2O	H^+/H_2	30	$1.58\cdot10^{-6}$	CP	59Boc [64])
	0.1 **N** HCl/H_2O	H^+/H_2	40	$2.13\cdot10^{-6}$	CP	59Boc [64])
	0.01 **N** HCl/H_2O	H^+/H_2	10	$1.44\cdot10^{-6}$	CP	59Boc [64])
	0.01 **N** HCl/H_2O	H^+/H_2	20	$1.94\cdot10^{-6}$	CP	59Boc [64])
	0.01 **N** HCl/H_2O	H^+/H_2	30	$2.95\cdot10^{-6}$	CP	59Boc [64])
	0.01 **N** HCl/H_2O	H^+/H_2	40	$1.23\cdot10^{-5}$	CP	59Boc [64])
	0.001 **N** HCl/H_2O	H^+/H_2	10	$4.36\cdot10^{-6}$	CP	59Boc [64])
	0.001 **N** HCl/H_2O	H^+/H_2	20	$7.76\cdot10^{-6}$	CP	59Boc [64])
	0.001 **N** HCl/H_2O	H^+/H_2	30	$1.65\cdot10^{-5}$	CP	59Boc [64])
	0.001 **N** HCl/H_2O	H^+/H_2	40	$2.34\cdot10^{-5}$	CP	59Boc [64])
	0.15 **N** NaOH/H_2O	H^+/H_2	16	$1.02\cdot10^{-6}$	CP	59Boc [64])
	0.02 **N** NaOH/H_2O	H^+/H_2	16	$5.12\cdot10^{-7}$	CP	59Boc [64])
	0.005 **N** NaOH/H_2O	H^+/H_2	16	$3.98\cdot10^{-7}$	CP	59Boc [64])
Cu(111)	0.5 **M** H_2SO_4 +0.5 **M** $CuSO_4/H_2O$	$Cu^{2+/0}$	-	$4\cdot10^{-4}$	IP	69Tka
Gallium (Ga)	0.1 **M** $HClO_4/H_2O$	H^+/H_2, 0.1 **M**	32	$1.2\cdot10^{-10}$	CP	66But3
	0.1 **M** $HClO_4/H_2O$	H^+/H_2, 0.1 **M**	58	$2.08\cdot10^{-10}$	CP	66But3
+19.2mol%In	0.1 **M** $HClO_4/H_2O$	H^+/H_2, 0.1 **M**	37.5	$1.15\cdot10^{-11}$	CP	66But3
	0.1 **M** $HClO_4/H_2O$	H^+/H_2, 0.1 **M**	55.5	$5.75\cdot10^{-12}$	CP	66But3
+0.122mol%Hg	0.1 **M** $HClO_4/H_2O$	H^+/H_2, 0.1 **M**	31.9	$3.38\cdot10^{-11}$	CP	66But3
+0.36mol%Hg	0.1 **M** $HClO_4/H_2O$	H^+/H_2, 0.1 **M**	43.6	$8.7\cdot10^{-11}$	CP	66But3
+1.64mol%Hg	0.1 **M** $HClO_4/H_2O$	H^+/H_2, 0.1 **M**	33.3	$1.69\cdot10^{-13}$	CP	66But3
+99.65mol%Hg	0.1 **M** $HClO_4/H_2O$	H^+/H_2, 0.1 **M**	25	$2.39\cdot10^{-12}$	CP	66But3
	0.5 **M** H_2SO_4/H_2O	H^+/H_2, 0.1 **M**	-	$10^{-10.1}$	CP	65Rai
Gold (Au)	0.54 **M** KCN +0.0052 **M** $KAu(CN)_2/H_2O$	$Au^{+/0}$	60	$0.82\cdot10^{-3}$	CP	71Che
	1 **M** H_2SO_4/H_2O	$Ce^{3+/4+}$	-	$(4.0\cdot10^{-4})$	CP	70Bon
	1 **M** $HClO_4/H_2O$	$Ce^{3+/4+}$	-	$(3.0\cdot10^{-4})$	CP	70Bon
	1 **M** $HClO_4/H_2O$	$Ce^{4+/3+}$	RT	$3.6\cdot10^{-5}$	CP	02Fer1 [65])
	1 **M** $HClO_4/H_2O$	$Ce^{4+/3+}$	RT	$5.8\cdot10^{-5}$	IP	02Fer1 [65])
	0.025 **M** TBAP [52]) /DMF [54])	$[Co(cp)_2]$ (o)	-	*2.63*	ACP	93Win
	0.05 **M** TBAP [52]) /DMF [54])	$[Co(cp)_2]$ (o)	-	*1.75*	ACP	93Win
	0.1 **M** TBAP [52]) /DMF [54])	$[Co(cp)_2]$ (o)	-	*0.94*	ACP	93Win

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.25 **M** TBAP [52]) /DMF [54])	$[Co(cp)_2]$ (o)	-	*0.71*	ACP	93Win
	0.1 **M** KPF_6/H_2O	$[Co(NH_3)_6]^{3+/2+}$	25	*(1.0·10^{-2})*	RDE	80Bar1
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	25	*(2.0·10^{-3})*	RDE	80Bar1
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	23	*(4.2·10^{-4})*	CV	86Ham
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5]$ $F^{2+/1+}$	23	*(30·10^{-4})*	CV	86Ham
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5]OH_2^{3+/2+}$	25	*(1.0·10^{-3})*	RDE	80Bar1
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	$[Co(NH_3)_5]Cl^{2+/1+}$	24	*(0.1)*	PP	84Bar2
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$ +5 m**M** NaI/H_2O	$[Co(NH_3)_5]Cl^{2+/1+}$	24	*(1.0·10^{-7})*	PP	84Bar2
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	$[Co(NH_3)_5]Br^{2+/1+}$	24	*(0.15)*	PP	84Bar2
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$ +5 m**M** NaI/H_2O	$[Co(NH_3)_5]Br^{2+/1+}$	24	*(1.0·10^{-7})*	PP	84Bar2
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	$[Co(NH_3)_5]NCS^{2+/1+}$	24	*(5.0·10^{-2})*	PP	84Bar2
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$ +5 m**M** NaI/H_2O	$[Co(NH_3)_5]NCS^{2+/1+}$	24	*(2.5·10^{-2})*	PP	84Bar2
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	*cis*-$([Co(NH_3)_4]Cl_2)^{2+/1+}$	24	*(4.0·10^{-3})*	PP	84Bar2
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$ +5 m**M** NaI/H_2O	*cis*-$([Co(NH_3)_4]Cl_2)^{2+/1+}$	24	*(1.0·10^{-5})*	PP	84Bar2
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	*cis*-$([Co(en)_2]Cl_2)^{2+/1+}$	24	*(5.0·10^{-3})*	PP	84Bar2 [63])
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$ +5 m**M** NaI/H_2O	*cis*-$([Co(en)_2]Cl_2)^{2+/1+}$	24	*(7.0·10^{-6})*	PP	84Bar2 [63])
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	*cis*-$([Co(en)_2](NCS)_2)^{2+/1+}$	24	*(6.5·10^{-4})*	PP	84Bar2 [63])
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$ +5 m**M** NaI/H_2O	*cis*-$([Co(en)_2](NCS)_2)^{2+/1+}$	24	*(1.0·10^{-4})*	PP	84Bar2 [63])
	0.1 **M** $NaClO_4/H_2O$	$([Co(NH_3)_5]HCOO)^{2+/1+}$	24	*(6.5·10^{-5})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	$([Co(NH_3)_5]CH_3COO)^{2+/1+}$	24	*(3.1·10^{-5})*	PP	85Li

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Au	0.1 **M** $NaClO_4/H_2O$	$([Co(NH_3)_5]CCl_3COO)^{2+/1+}$	24	*(6.8·10^{3})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	$([Co(NH_3)_5]CF_3COO)^{2+/1+}$	24	*(3.2·10^{-4})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	$([Co(NH_3)_5](CH_3)_3COO)^{2+/1+}$	24	*(3.0·10^{-5})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	$([Co(NH_3)_5]CH_3C(O)COO)^{2+/1+}$	24	*(1.2·10^{-4})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	$([Co(NH_3)_5]HC(O)COO)^{2+/1+}$	24	*(1.0·10^{-4})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	$([Co(NH_3)_5]CH_2OHCOO)^{2+/1+}$	24	*(4.8·10^{-5})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	$([Co(NH_3)_5]CH_3C(OH)HCOO)^{2+/1+}$	24	*(6.0·10^{-5})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	$([Co(NH_3)_5]CH_2FCOO)^{2+/1+}$	24	*(1.8·10^{-4})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	([Co(NH$_3$)$_5$] cyclopentane carboxylate)$^{2+/1+}$	24	*(4.5·10^{-4})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	([Co(NH$_3$)$_5$] cyclopentane acetate)$^{2+/1+}$	24	*(5.0·10^{-4})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	([Co(NH$_3$)$_5$] cyclohexane carboxylate)$^{2+/1+}$	24	*(4.6·10^{-4})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	([Co(NH$_3$)$_5$] cyclohexane acetate)$^{2+/1+}$	24	*(1.6·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	([Co(NH$_3$)$_5$] 3-cyclohexane propionate)$^{2+/1+}$	24	*(3.1·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	([Co(NH$_3$)$_5$] benzoate)$^{2+/1+}$	24	*(1.2·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	([Co(NH$_3$)$_5$] phenylacetate)$^{2+/1+}$	24	*(6.5·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	([Co(NH$_3$)$_5$]3-phenyl propionate)$^{2+/1+}$	24	*(9.7·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	([Co(NH$_3$)$_5$]2-phenyl acrylate)$^{2+/1+}$	24	*(2.0·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	([Co(NH$_3$)$_5$]furan-3-carboxylate)$^{2+/1+}$	24	*(1.1·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	([Co(NH$_3$)$_5$]furan-2-carboxylate)$^{2+/1+}$	24	*(1.5·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4/H_2O$	([Co(NH$_3$)$_5$]furan-2-acrylate)$^{2+/1+}$	24	*(1.6·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([Co(NH$_3$)$_5$]pyridine-2-carboxylate)$^{2+/1+}$	24	*(1.5·10^{-3})*	PP	85Li

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$]pyridine-3-carboxylate)$^{2+/1+}$	24	*(1.3·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$/H_2O	([$Co(NH_3)_5$] pyridine-2-acetate)$^{2+/1+}$	24	*(8.9·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$/H_2O	([$Co(NH_3)_5$] 3-pyridyl acrylate)$^{2+/1+}$	24	*(7.2·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$/H_2O	([$Co(NH_3)_5$] thiophene-2-carboxylate)$^{2+/1+}$	24	*(2.3·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$/H_2O	([$Co(NH_3)_5$] thiophene-2-acetate)$^{2+/1+}$	24	*(7.3·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$/H_2O	([$Co(NH_3)_5$] thiophene-2-propionate)$^{2+/1+}$	24	*(7.0·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$/H_2O	([$Co(NH_3)_5$] thiophene-2-butyrate)$^{2+/1+}$	24	*(3.0·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$/H_2O	([$Co(NH_3)_5$] thiophene-3-carboxylate)$^{2+/1+}$	24	*(2.8·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$/H_2O	([$Co(NH_3)_5$] thiophene-2-acetate)$^{2+/1+}$	24	*(7.5·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$/H_2O	([$Co(NH_3)_5$] thiophene-2-acrylate)$^{2+/1+}$	24	*(3.2·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$/H_2O	([$Co(NH_3)_5$] thiophene-2-carboxylate)$^{2+/1+}$	24	*(7.0·10^{-5})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$/H_2O	([$Co(NH_3)_5$] thiophene-2-acetate)$^{2+/1+}$	24	*(1.2·10^{-5})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$/H_2O	([$Co(NH_3)_5$] thiophene-2-propionate)$^{2+/1+}$	24	*(1.8·10^{-5})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$/H_2O	([$Co(NH_3)_5$] thiophene-2-butyrate)$^{2+/1+}$	24	*(1.3·10^{-5})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$/H_2O	([$Co(NH_3)_5$] thiophene-2-acrylate)$^{2+/1+}$	24	*(2.0·10^{-5})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$/H_2O	([$Co(NH_3)_5$] thiophene-3-carboxylate)$^{2+/1+}$	24	*(7.5·10^{-5})*	PP	84Li

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Au	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	($[Co(NH_3)_5]$ thiophene-3-acetate)$^{2+/1+}$	24	*(3.0·10^{-5})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	($[Co(NH_3)_5]$ benzothiophene-3-carboxylate)$^{2+/1+}$	24	*(2.0·10^{-5})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	($[Co(NH_3)_5]$ benzothiophene-3-malonate)$^{2+/1+}$	24	*(4.0·10^{-5})*	PP	84Li
	0.1 **M** $NaClO_4/H_2O$	$Cr_{aq}^{2+/3+}$	25	*(7.0·10^{-4})*	RDE	80Bar1
	0.5 **M** $NaClO_4/H_2O$	$Cr_{aq}^{2+/3+}$	25	*(1.5·10^{-3})*	RDE	80Bar1
	0.1 **M** H_2SO_4+0.1 **M** $NaCl/H_2O$ [66])	$Cr_{aq}^{2+/3+}$	-	*1.5·10^{-3}*	CV	89Rod
	0.1 **M** H_2SO_4+0.2 **M** $NaCl/H_2O$ [66])	$Cr_{aq}^{2+/3+}$	-	*2.7·10^{-3}*	CV	89Rod
	0.1 **M** H_2SO_4+0.5 **M** $NaCl/H_2O$ [66])	$Cr_{aq}^{2+/3+}$	-	*3.3·10^{-3}*	CV	89Rod
	0.1 **M** H_2SO_4 +1 **M** $NaCl/H_2O$ [66])	$Cr_{aq}^{2+/3+}$	-	*5.8·10^{-3}*	CV	89Rod
	0.1 **M** H_2SO_4 +2 **M** $NaCl/H_2O$ [66])	$Cr_{aq}^{2+/3+}$	-	*7.3·10^{-3}*	CV	89Rod
	0.2 **M** HCl/H_2O [67])	$Cr_{aq}^{2+/3+}$	-	*7.3·10^{-3}*	CV	89Rod
	0.2 **M** HCl/H_2O [68])	$Cr_{aq}^{2+/3+}$	-	*4.8·10^{-3}*	CV	89Rod
	0.2 **M** HCl/H_2O [69])	$Cr_{aq}^{2+/3+}$	-	*9.4·10^{-3}*	CV	89Rod
	0.2 **M** HCl/H_2O [70])	$Cr_{aq}^{2+/3+}$	-	*6.0·10^{-3}*	CV	89Rod
	0.4 **M** K_2SO_4 +0.01 **N** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	19.0	TPF	88Boc
	0.4 **M** K_2SO_4 +0.01 **N** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	18.0	CV	88Boc
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$, 10 m**M**	25	2.5·10^{-4}	SS	68Boc2 [71])
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$, 0.1 **M**	25	7.2·10^{-4}	SS	68Boc2 [72])
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$, 0.004 **M**	-	6.0·10^{-6}	RDE	96Tor
	1 **N** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	*10·10^{-3}*	RDE	72Ang
	0.1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	22	*7.8·10^{-3}*	CV	89Chu

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.5 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	20	*(5·10^{-5})*	RDE	78Web
	0.5 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	RT	*2.2·10^{-5}*	PS	87Hun [73])
	0.5 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	25	*8.0·10^{-5}*	CP	91Cur
	1 **N** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	5.0	RDE	83Vie
	1 **M** KF/H_2O	$Fe(CN)_6^{3-/4-}$	20	*37.0*	TPF	83Iwa
	0.5 **M** K_2SO_4/H_2O	$Fe(CN)_6^{3-/4-}$	20	*26.0*	TPF	83Iwa
	1 **M** $NaNO_3/H_2O$	$Fe(CN)_6^{3-/4-}$, 0.004 **M**	-	$2.01\cdot10^{-4}$	RDE	96Tor
	0.5 **M** K-oxalate/H_2O	$Fe(ox)_3^{4-/3-}$	25	*5.0·10^{-3}*	IP	52Ran2
	0.3 **M** TEAH/acetone	$FeC^{+/0}$	–75	*0.083*	CV	92Saf
	0.3 **M** TEAH/acetone	$FeC^{+/0}$	–55	*0.38*	CV	92Saf
	0.3 **M** TEAH/acetone	$FeC^{+/0}$	–35	*0.92*	CV	92Saf
	0.3 **M** TEAH/acetone	$FeC^{+/0}$	–15	*1.4*	CV	92Saf
	0.3 **M** TEAH/acetone	$FeC^{+/0}$	5	*2.5*	CV	92Saf
	0.3 **M** TEAH/acetone	$FeC^{+/0}$	25	*5.5*	CV	92Saf
	0.3 **M** TEAH/propionitrile	$FeC^{+/0}$	–75	*0.07*	CV	92Saf
	0.3 **M** TEAH/propionitrile	$FeC^{+/0}$	–55	*0.15*	CV	92Saf
	0.3 **M** TEAH/propionitrile	$FeC^{+/0}$	–35	*0.43*	CV	92Saf
	0.3 **M** TEAH/propionitrile	$FeC^{+/0}$	–15	0.68	CV	92Saf
	0.3 **M** TEAH/propionitrile	$FeC^{+/0}$	5	*2.1*	CV	92Saf
	0.3 **M** TEAH/propionitrile	$FeC^{+/0}$	25	*3.0*	CV	92Saf
	0.3 **M** TEAH/butyronitrile	$FeC^{+/0}$	–75	*0.02*	CV	92Saf
	0.3 **M** TEAH/butyronitrile	$FeC^{+/0}$	–55	*0.04*	CV	92Saf
	0.3 **M** TEAH/butyronitrile	$FeC^{+/0}$	–35	*0.17*	CV	92Saf
	0.3 **M** TEAH/butyronitrile	$FeC^{+/0}$	–15	0.36	CV	92Saf
	0.3 **M** TEAH/butyronitrile	$FeC^{+/0}$	5	*0.48*	CV	92Saf
	0.3 **M** TEAH/butyronitrile	$FeC^{+/0}$	25	*0.95*	CV	92Saf
	1 **M** KOH/H_2O	$MnO_4^{-/2-}$	20	$0.8\cdot10^{-2}$	CP, GS	67Thi1
	0.001 **N** HCl/H_2O	H^+/H_2	-	$10^{-7.32}$	CP	57Pen
	0.01 **N** HCl/H_2O	H^+/H_2	-	$10^{-6.63}$	CP	57Pen

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Au	0.1 **N** HCl/H_2O	H^+/H_2	-	$10^{-5.64}$	CP	57Pen
	0.1 **M** HCl/H_2O	H^+/H_2	-	$10^{-6.3}$	CP	59Ive
	0.1 **M** $HClO_4/H_2O$	H^+/H_2	-	$1.4\cdot10^{-7}$	CP	98Per
	0.01 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-5.83}$	CP	68Gos
	0.1 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-5.75}$	CP	68Gos
	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-5.9}$	CP	67Man
	1 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-5.86}$	CP	68Gos
	2 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-5.81}$	CP	68Gos
	3 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-5.78}$	CP	68Gos
	5 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-5.71}$	CP	68Gos
	8 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-5.61}$	CP	68Gos
	12 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-5.48}$	CP	68Gos
	0.001 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-5.64}$	CP	57Pen
	0.01 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-7.05}$	CP	57Pen
	0.1 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-7.04}$	CP	57Pen
	0.6 **M** TEAP [74])/AN	ferrocene	-	*3.1*	CV	88Wip
	0.2 **M** $NaClO_4/H_2O$	$V_{aq}^{2+/3+}$	25	*(>0.1)*	CV	80Bar1
	0.2 **M** $NaClO_4$ +1 m**M** Cr^{2+}/H_2O	$V_{aq}^{2+/3+}$	25	*(5.5·10^{-5})*	CV	80Bar1
	0.2 **M** KPF_6/H_2O	$Ru_{aq}^{2+/3+}$	25	*(>5.0·10^{-2})*	CV	80Bar1
	0.2 **M** KPF_6 +1 m**M** Cr^{2+}/H_2O	$Ru_{aq}^{2+/3+}$	25	*(5·10^{-5})*	CV	80Bar1
	1 **M** KF/H_2O	$[Ru(NH_3)_6]^{3+/2+}$	20	98	TPF	85Iwa
	0.1 **M** TEAP [74]) /DMF [54])	benzoquinone (r)	25	*0.0053*	CV	73Cap
	0.3 **M** TEAH/propionitrile	*o*-nitrotoluene$^{0/-}$	–55	*0.014*	CV	92Saf
	0.3 **M** TEAH/propionitrile	*o*-nitrotoluene$^{0/-}$	–35	*0.071*	CV	92Saf
	0.3 **M** TEAH/propionitrile	*o*-nitrotoluene$^{0/-}$	–15	*0.16*	CV	92Saf
	0.3 **M** TEAH/propionitrile	*o*-nitrotoluene$^{0/-}$	5	*0.36*	CV	92Saf
	0.3 **M** TEAH/propionitrile	*o*-nitrotoluene$^{0/-}$	25	*0.75*	CV	92Saf
	0.3 **M** TEAH/butyronitrile	*o*-nitrotoluene$^{0/-}$	–55	*0.014*	CV	92Saf
	0.3 **M** TEAH/butyronitrile	*o*-nitrotoluene$^{0/-}$	–35	*0.042*	CV	92Saf
	0.3 **M** TEAH/butyronitrile	*o*-nitrotoluene$^{0/-}$	–15	*0.12*	CV	92Saf
	0.3 **M** TEAH/butyronitrile	*o*-nitrotoluene$^{0/-}$	5	*0.28*	CV	92Saf
	0.3 **M** TEAH/butyronitrile	*o*-nitrotoluene$^{0/-}$	25	*0.58*	CV	92Saf

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.3 **M** TEAH/propionitrile	nitromesitylene$^{0/-}$	–35	*0.029*	CV	92Saf
	0.3 **M** TEAH/propionitrile	nitromesitylene$^{0/-}$	–15	*0.075*	CV	92Saf
	0.3 **M** TEAH/propionitrile	nitromesitylene$^{0/-}$	5	*0.19*	CV	92Saf
	0.3 **M** TEAH/propionitrile	nitromesitylene$^{0/-}$	25	*0.36*	CV	92Saf
	0.3 **M** TEAH/butyronitrile	nitromesitylene$^{0/-}$	–35	*0.025*	CV	92Saf
	0.3 **M** TEAH/butyronitrile	nitromesitylene$^{0/-}$	–15	*0.082*	CV	92Saf
	0.3 **M** TEAH/butyronitrile	nitromesitylene$^{0/-}$	5	*0.16*	CV	92Saf
	0.3 **M** TEAH/butyronitrile	nitromesitylene$^{0/-}$	25	*0.31*	CV	92Saf
Au(100)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	23	*(1.8·10^{-4})*	CV	86Ham
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5]F^{2+/1+}$	23	*(5·10^{-4})*	CV	86Ham
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5OSO_3]^{2+/1+}$	23	*(2.5·10^{-4})*	CV	86Ham
	0.1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{3+}$ (r)	-	*(1.0)*	CV	97Faw
	0.1 **M** $HClO_4/H_2O$	H^+/H_2	-	0.5·10^{-7}	CP	98Per
Au(110)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	23	*(1.5·10^{-4})*	CV	86Ham
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5]F^{2+/1+}$	23	*(1.5·10^{-4})*	CV	86Ham
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5OSO_3]^{2+/1+}$	23	*(50·10^{-4})*	CV	86Ham
	0.1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{3+}$ (r)	-	*(1.0)*	CV	97Faw
	0.1 **M** $HClO_4/H_2O$	H^+/H_2	-	0.3·10^{-7}	CP	98Per
Au(111)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	23	*(3·10^{-4})*	CV	86Ham
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5]F^{2+/1+}$	23	*(10·10^{-4})*	CV	86Ham
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5OSO_3]^{2+/1+}$	23	*(40·10^{-4})*	CV	86Ham
	0.1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{3+}$ (r)	-	*(1.0)*	CV	97Faw
	0.1 **M** $HClO_4/H_2O$	H^+/H_2	-	2.5·10^{-7}	CP	98Per
Gold/Palladium alloy 25 at%	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$, 0.1 **M**	25	0.0018	SS	68Boc2
Gold/Palladium alloy 50 at%	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$, 0.1 **M**	25	0.0046	SS	68Boc2
Gold/Palladium alloy 75 at%	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$, 0.1 **M**	25	0.0063	SS	68Boc2
Gold/Platinum alloy 20 at%	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$, 0.1 **M**	25	6.9·10^{-4}	SS	68Boc2

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Indium (In)	0.1 **M** $HClO_4/H_2O$	H^+/H_2	30	$10^{-10.83}$	CP	65But5 [75])
	0.1 **M** $HClO_4/H_2O$	H^+/H_2	25	$1.4\cdot10^{-12}$	CP	64But1
	10^{-2} **M** $HClO_4$ $+10^{-2}$ **M** $In(ClO_4)_3/H_2O$	In^{3+}/In	20	$6.7\cdot10^{-5}$	CF	68Los
	10^{-2} **M** $HClO_4$ $+10^{-2}$ **M** $In(ClO_4)_3$ $+5\cdot10^{-2}$ **M** $NaCl/H_2O$	In^{3+}/In	20	$2.4\cdot10^{-4}$	CF	68Los
	10^{-2} **M** $HClO_4$ $+10^{-2}$ **M** $In(ClO_4)_3$ +0.16 **M** $NaCl/H_2O$	In^{3+}/In	20	$4.2\cdot10^{-4}$	CF	68Los
	3 m**M** $HClO_4/H_2O$	In^{3+}/In	20	$2.88\cdot10^{-3}$	RT	68Los
In$_{(Pt)}$ [76])	3 m**M** $HClO_4$ $+NaClO_4/H_2O$ (I=3 **M**)	In^{3+}/In	20	$2.9\cdot10^{-3}$	RT	70Los
In$_{(Pt)}$	10 m**M** $HClO_4$ $+NaClO_4/H_2O$ (I =3 **M**)	In^{3+}/In	20	$6.7\cdot10^{-5}$	RT	70Los
In$_{(Pt)}$	10 m**M** $HClO_4$ $+NaClO_4$+ 0.05 **M** $NaCl/H_2O$ (I =3 **M**)	In^{3+}/In	20	$2.4\cdot10^{-4}$	RT	70Los
In$_{(Pt)}$	10 m**M** $HClO_4$ $+NaClO_4$+0.16 **M** $NaCl/H_2O$ (I =3 **M**)	In^{3+}/In	20	$4.2\cdot10^{-4}$	RT	70Los
In$_{(Hg)}$0.97 **M** In	1 **M** $HClO_4/H_2O$	In^{3+}/In, 5.3 m**M** In^{3+}	20	$4\cdot10^{-6}$	CP	65Mol3
In$_{(Hg)}$ 0.002...0.9 **M** In	0.01 **M** $HClO_4$ $+NaClO_4/H_2O$ (I =3 **M**)	In^{3+}/In	-	0.0067	RT	63Bud
In$_{(Hg)}$ 0.1**M** In	0.0077 **M** $HClO_4$ +2.9923 **M** $NaClO_4/H_2O$	In^{3+}/In, 10 m**M** In^{3+}	20	$3.5\cdot10^{-4}$	CP, RT	60Los2, 62Los2
In$_{(Hg)}$ 0.1**M** In	0.016 **M** $HClO_4$ +2.984 **M** $NaClO_4/H_2O$	In^{3+}/In, 10 m**M** In^{3+}	20	$2.2\cdot10^{-4}$	CP, RT	60Los2, 62Los2
In$_{(Hg)}$ 0.1**M** In	0.05 **M** $HClO_4$ +2.95 **M** $NaClO_4/H_2O$	In^{3+}/In, 10 m**M** In^{3+}	20	$5.2\cdot10^{-5}$	CP, RT	60Los2, 62Los2
In$_{(Hg)}$ 0.1**M** In	0.01 **M** $HClO_4$ +1 **M** $NaClO_4/H_2O$	In^{3+}/In, 10 m**M** In^{3+}	20	$6.9\cdot10^{-4}$	RT	64Mol, 65Mol2
In$_{(Hg)}$ 0.1**M** In	0.01 **M** $HClO_4$ +3 **M** $NaClO_4/H_2O$	In^{3+}/In, 10 m**M** In^{3+}	20	$1.52\cdot10^{-4}$	RT	64Mol, 65Mol2
In$_{(Hg)}$ 0.1**M** In	0.01 **M** $HClO_4$ +6.1 **M** $NaClO_4/H_2O$	In^{3+}/In, 10 m**M** In^{3+}	20	$2.5\cdot10^{-4}$	RT	64Mol, 65Mol2

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
In$_{(Hg)}$ 0.1**M** In	0.01 **M** $HClO_4$ +2.9 **M** $NaClO_4$ +0.1 **M** Na_2SO_4/H_2O	In^{3+}/In, 10 m**M** In^{3+}	20	$3.0\cdot10^{-4}$	CP, RT	60Los3, 63Los
In$_{(Hg)}$ 0.1**M** In	0.01 **M** $HClO_4$ +2.4 **M** $NaClO_4$ +0.6 **M** Na_2SO_4/H_2O	In^{3+}/In, 10 m**M** In^{3+}	20	$1.7\cdot10^{-4}$	CP, RT	60Los3, 63Los
In$_{(Hg)}$ 0.1**M** In	0.01 **M** $HClO_4$ +2.9 **M** $NaClO_4$ +0.1 **M** $NaBr/H_2O$	In^{3+}/In, 10 m**M** In^{3+}	20	$1.1\cdot10^{-3}$	CP, RT	63Los
In$_{(Hg)}$ 0.1**M** In	0.01 **M** $HClO_4$ +2.5 **M** $NaClO_4$ +0.5 **M** $NaBr/H_2O$	In^{3+}/In, 10 m**M** In^{3+}	20	$2.9\cdot10^{-3}$	CP, RT	63Los
In$_{(Hg)}$ 0.1**M** In	0.01 **M** $HClO_4$ +2.9 **M** $NaClO_4$ +0.1 **M** NaI/H_2O	In^{3+}/In, 10 m**M** In^{3+}	20	$4.7\cdot10^{-3}$	CP, RT	63Los
In$_{(Hg)}$ 0.1**M** In	0.01 **M** $HClO_4$ +2.5 **M** $NaClO_4$ +0.5 **M** NaI/H_2O	In^{3+}/In, 10 m**M** In^{3+}	20	$8.7\cdot10^{-3}$	CP, RT	63Los
In$_{(Hg)}$ 0.1**M** In	0.01 **M** $HClO_4$ +2.5 **M** $NaClO_4$ +0.5 **M** $NaCl/H_2O$	In^{3+}/In, 10 m**M** In^{3+}	20	$1.7\cdot10^{-4}$	CP, RT	63Los
In$_{(Hg)}$ 0.1**M** In	0.01 **M** $HClO_4$ +2.9 **M** $NaClO_4$ +0.1 **M** $NaCl/H_2O$	In^{3+}/In, 10 m**M** In^{3+}	20	$1.7\cdot10^{-4}$	CP, RT	63Los
In$_{(Hg)}$ 0.1**M** In	0.1 **M** $HClO_4$ +0.3 **M** NaF/H_2O	In^{3+}/In, $4\cdot10^{-6}$ m**M**In^{3+}	20	$1.55\cdot10^{-5}$	CP	67Los
In$_{(Hg)}$	0.86 **M** H_3PO_4/H_2O	$In^{3+/0}$	25	*(6.6·10^{-11})*	DCP	58Ima
In$_{(Hg)}$ 0.1**M** In	2.98 **M** $NaClO_4$ +0.02 **M** NaF +$HClO_4/H_2O$ (pH=1.3)	In^{3+}/In, 0.01 **M** In^{3+}	20	$9.1\cdot10^{-5}$	CP, RT	63Los
In$_{(Hg)}$ 0.1**M** In	2.9 **M** $NaClO_4$ +0.1 **M** NaF +$HClO_4/H_2O$ (pH=1.3)	In^{3+}/In, 0.01 **M** In^{3+}	20	$1.6\cdot10^{-5}$	CP, RT	63Los
In$_{(Hg)}$						
8.4% In	0.1 **M** $HClO_4$	H^+/H_2	25	$2.9\cdot10^{-13}$	CP	64But1
10.0% In	0.1 **M** $HClO_4$	H^+/H_2	25	$1.9\cdot10^{-14}$	CP	65But4
10.0% In	0.1 **M** $HClO_4$	H^+/H_2	30	$3.3\cdot10^{-14}$	CP	65But4
10.0% In	0.1 **M** $HClO_4$	H^+/H_2	40	$1.0\cdot10^{-13}$	CP	65But4
10.0% In	0.1 **M** $HClO_4$	H^+/H_2	50	$2.7\cdot10^{-13}$	CP	65But4
10.0% In	0.1 **M** $HClO_4$	H^+/H_2	61	$8.9\cdot10^{-13}$	CP	65But4
16.3% In	0.1 **M** $HClO_4$	H^+/H_2	25	$3.8\cdot10^{-13}$	CP	64But1
20.0% In	0.1 **M** $HClO_4$	H^+/H_2	14	$4.2\cdot10^{-15}$	CP	65But4
20.0% In	0.1 **M** $HClO_4$	H^+/H_2	25	$1.3\cdot10^{-15}$	CP	65But4

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
20.0% In	0.1 **M** $HClO_4$	H^+/H_2	35	$3.3\cdot10^{-14}$	CP	65But4
20.0% In	0.1 **M** $HClO_4$	H^+/H_2	44	$5.2\cdot10^{-13}$	CP	65But4
20.0% In	0.1 **M** $HClO_4$	H^+/H_2	55	$1.2\cdot10^{-13}$	CP	65But4
20.0% In	0.1 **M** $HClO_4$	H^+/H_2	66	$1.8\cdot10^{-12}$	CP	65But4
30.4% In	0.1 **M** $HClO_4$	H^+/H_2	25	$5.5\cdot10^{-13}$	CP	64But1
40.0% In	0.1 **M** $HClO_4$	H^+/H_2	14	$4.9\cdot10^{-15}$	CP	65But4
40.0% In	0.1 **M** $HClO_4$	H^+/H_2	25	$2.9\cdot10^{-14}$	CP	65But4
40.0% In	0.1 **M** $HClO_4$	H^+/H_2	35	$9.5\cdot10^{-14}$	CP	65But4
40.0% In	0.1 **M** $HClO_4$	H^+/H_2	51	$1.2\cdot10^{-12}$	CP	65But4
40.0% In	0.1 **M** $HClO_4$	H^+/H_2	65	$3.3\cdot10^{-12}$	CP	65But4
52.2% In	0.1 **M** $HClO_4$	H^+/H_2	25	$8.1\cdot10^{-13}$	CP	64But1
63.0% In	0.1 **M** $HClO_4$	H^+/H_2	16	$8.1\cdot10^{-14}$	CP	65But4
63.0% In	0.1 **M** $HClO_4$	H^+/H_2	25	$1.0\cdot10^{-13}$	CP	65But4
63.0% In	0.1 **M** $HClO_4$	H^+/H_2	36	$4.3\cdot10^{-13}$	CP	65But4
63.0% In	0.1 **M** $HClO_4$	H^+/H_2	46	$5.2\cdot10^{-13}$	CP	65But4
63.0% In	0.1 **M** $HClO_4$	H^+/H_2	56	$1.5\cdot10^{-12}$	CP	65But4
63.0% In	0.1 **M** $HClO_4$	H^+/H_2	66	$9.7\cdot10^{-12}$	CP	65But4
64.4% In	0.1 **M** $HClO_4$	H^+/H_2	25	$1.1\cdot10^{-12}$	CP	64But1
Iridium (Ir)	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$ 0.1 **M**	25	0.0052	SS	68Boc2
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$ 0.01 **M**	25	0.0015	SS	68Boc2
	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-3.7}$	CP	65Mat
	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-3.4}$	CP	67Man
	0.1 **M** TEAP [74]) /DMF [54])	benzoquinone (r)	25	*6.2·10*$^{-3}$	CV	73Cap
	HCl-buffer/H_2O, pH=1.1	hydroquinone/ quinone, 5.5 m**M**	25	$6.3\cdot10^{-6}$	CP	62Tom [77])
	HCl-buffer/H_2O, pH=1.1	hydroquinone/ quinone, 5.5 m**M**	25	$3.4\cdot10^{-5}$	CF	62Tom
	acetate buffer/H_2O, pH=3.0	hydroquinone/ quinone, 5.5 m**M**	25	$2.6\cdot10^{-5}$	CF	62Tom
	acetate buffer/H_2O, pH=5.0	hydroquinone/ quinone, 5.5 m**M**	25	$4.9\cdot10^{-5}$	CF	62Tom
Iron (Fe)	1 **M** H_2SO_4 +0.1 **M** Na_2SO_4/H_2O	$Fe^{2+/3+}$, 0.05 **M**	25	$1\cdot10^{-5}$	CP	64Mak [78])
	0.001 **N** HCl/H_2O	H^+/H_2	-	$10^{-6.19}$	CP	57Pen
	0.01 **N** HCl/H_2O	H^+/H_2	-	$10^{-6.29}$	CP	57Pen
	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-5.6}$	CP	67Man
	0.01 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-6.62}$	CP	57Pen
	0.1 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-6.06}$	CP	57Pen
	1.33 **M** NaOH +1.95 **M** $NaNO_3/H_2O$	NO_3^-/NO	25	$4.19\cdot10^{-13}$	CV	96Boc [79])

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Lead (Pb)	0.01...1 **M** HCl/H_2O	H^+/H_2	-	$10^{-11.5}$	CP	70Avd
	0.5 **N** H_2SO_4/H_2O	H^+/H_2	-	$10^{-11.33}$	CP	64Boc3
	H_2SO_4/H_2O, ρ=1.225 g·cm^{-3}	H^+/H_2	30	$1.8\cdot10^{-13}$	CS	57Rüe [80])
	H_2SO_4/H_2O, ρ=1.225 g·cm^{-3}	H^+/H_2	30	$7.5\cdot10^{-13}$	CS	57Rüe [81])
	H_2SO_4/H_2O, ρ=1.225 g·cm^{-3}	H^+/H_2	30	$1.4\cdot10^{-12}$	CS	57Rüe [82])
	H_2SO_4/H_2O, ρ=1.225 g·cm^{-3}	H^+/H_2	30	$4.5\cdot10^{-12}$	CS	57Rüe
	0.5 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-6.47}$	CP	64Boc3
	H_2SO_4/H_2O, ρ=1.225 g·cm^{-3}	H_2O/O_2	30	$1.4\cdot10^{-10}$	CS	57Rüe [83])
	H_2SO_4/H_2O, ρ=1.225 g·cm^{-3}	H_2O/O_2	30	$8.9\cdot10^{-11}$	CS	57Rüe [84])
	H_2SO_4/H_2O, ρ=1.225 g·cm^{-3}	H_2O/O_2	30	$2.1\cdot10^{-10}$	CS	57Rüe [82])
	H_2SO_4/H_2O, ρ=1.225 g·cm^{-3}	H_2O/O_2	30	$4.4\cdot10^{-10}$	CS	57Rüe
	1 **M** $HClO_4/H_2O$	$Pb^{2+/0}$	16	0.3	IP	67Har
	0.9946 **M** KNO_3 +2.7 m**M** $Pb(NO_3)_2/H_2O$	$Pb^{2+/0}$	23	0.0262, 0.0356	IP, GD	69Ham3
	0.986 **M** KNO_3 +7 m**M** $Pb(NO_3)_2/H_2O$	$Pb^{2+/0}$	23	0.0349, 0.0395	IP, GD	69Ham3
	0.9 **M** KNO_3 +50 m**M** $Pb(NO_3)_2/H_2O$	$Pb^{2+/0}$	23	0.0468, 0.064	IP, GD	69Ham3
	0.8 **M** KNO_3 +0.1 **M** $Pb(NO_3)_2/H_2O$	$Pb^{2+/0}$	23	0.0498, 0.067	IP, GD	69Ham3
	0.5 **M** $Pb(NO_3)_2/H_2O$	$Pb^{2+/0}$	23	0.0714, 0.1	IP, GD	69Ham3
	0.05 **M** $PbCl_2$ +1 **M** KNO_3/H_2O	$Pb^{2+/0}$	-	0.1	IP	59Lor
	1.33 **M** NaOH +1.95 **M** $NaNO_3/H_2O$	NO_3^-/NO	25	$6.6\cdot10^{-10}$	CV	96Boc [79])
Pb$_{(Hg)}$	1 **M** $HClO_4/H_2O$	$Pb^{2+/0}$	-	2.0	IP	61Ran
Pb$_{(Hg)}$	1 **M** $HClO_4/H_2O$	$Pb^{2+/0}$	25	*1*	IP	62Ran
Pb$_{(Hg)}$	1 **M** KBr/H_2O, pH=2	$Pb^{2+/0}$, 1.5 m**M**	28	*0.96*	FR	71Aga
Pb$_{(Hg)}$	1 **M** KCl/H_2O	$Pb^{2+/0}$, 1.2 m**M**	32	*5.2*	FR	71Aga
Pb$_{(Hg)}$	1 **M** KCl/H_2O	$Pb^{2+/0}$, 2.4 m**M**	32	*7.7*	FR	71Aga
Pb$_{(Hg)}$	1 **M** KNO_3/H_2O	$Pb^{2+/0}$	20	*>1*	IP	52Ran1
Pb$_{(Hg)}$	1 **M** KNO_3/H_2O	$Pb^{2+/0}$	22	*0.8*	FR	58Bar1

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Pb$_{(Hg)}$	1 **M** KCl/H_2O	$Pb^{2+/0}$	22	*0.2*	FR	58Bar1
Pb$_{(Hg)}$	1 **M** KCl/H_2O	$Pb^{2+/0}$	-	250	IP	70Sal2
Pb$_{(Hg)}$	1 **M** KF/H_2O	$Pb^{2+/0}$, $3.7\cdot10^{-5}$ **M**	28	*$2.1\cdot10^{-3}$*	FR	71Aga
Pb$_{(Hg)}$	1 **M** KNO_3/H_2O	$Pb^{2+/0}$, 1.5 m**M** Pb^{2+}	28	*0.59*	FR	71Aga
Pb$_{(Hg)}$	1 **M** KNO_3/H_2O	$Pb^{2+/0}$, 3 m**M** Pb^{2+}	28	*1.0*	FR	71Aga
Pb$_{(Hg)}$	$K_4P_2O_7$+$NaNO_3/H_2O$	$Pb^{2+/0}$	-	*$1.66\cdot10^{-4}$*	DCP	68Pop
Pb$_{(Hg)}$	1 **M** $NaBr/H_2O$	$Pb^{2+/0}$	-	3...50	IP	70Sal2
Pb$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Pb^{2+/0}$	22	*3.3*	FR	58Bar1
Pb$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Pb^{2+/0}$	25	220	IP	70Sal1
Pb$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Pb^{2+/0}$, 1.2 m**M** Pb^{2+}	37	*10.2*	FR	71Aga
Pb$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Pb^{2+/0}$, 2.4 m**M** Pb^{2+}	37	*13.5*	FR	71Aga
Lithium (Li)	0.1 **M** NEt_4PF_6 /DMF [54])	$Li^{1+/0}$	25	*$1.1\cdot10^{-4}$*	IP/PP	80Bar3
	0.1 **M** NPr_4PF_6 [86]) /DMF [54])	$Li^{1+/0}$	25	*$2.8\cdot10^{-4}$*	IP/PP	80Bar3
	0.1 **M** NBu_4PF_6 [86]) /DMF [54])	$Li^{1+/0}$	25	*$4.7\cdot10^{-4}$*	IP/PP	80Bar3
Mercury (Hg)	0.1 **M** $LiCl/H_2O$	$Ba^{2+/0}$	25	*$1.6\cdot10^{-2}$*	IP	70Tim
	1 **M** $LiCl/H_2O$	$Ba^{2+/0}$	25	*$9\cdot10^{-3}$*	IP	70Tim
	0.1 **M** $(C_2H_5)_4NI/H_2O$	$Ba^{2+/0}$	25	*$1\cdot10^{-2}$*	IP	70Tim
	1 **M** HCl/H_2O	$Bi^{3+/0}$	26	*1.2...3.6*	IP	60Mou
	1 m**M** HCl +1 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	26	*$2.7\cdot10^{-3}$*	IP	60Mou
	1 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	25	*$3\cdot10^{-4}$*	ACP	61Kam
	1 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	26	*$3.7\cdot10^{-4}$*	IP	60Mou
	1 **M** HNO_3/H_2O	$Bi^{3+/0}$	26	*$2.8...4.2\cdot10^{-4}$*	IP	60Mou
	0.2 **M** HNO_3 +0.8 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	26	*$1.15\cdot10^{-3}$*	IP	60Mou
	0.4 **M** HNO_3 +0.6 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	26	*$1.75\cdot10^{-3}$*	IP	60Mou
	0.6 **M** HNO_3 +0.4 **M** $HClO_4/H_2O$	$Bi^{3+/0}$	26	*$2.5\cdot10^{-3}$*	IP	60Mou
	2 **M** $HClO_4$ +0.25 m**M** $NaCl/H_2O$	$Bi^{0/3+}$, $4\cdot10^{-6}$ **M**	20	10^{-5}	CP	67Los
	0.5 **M** H_2SO_4/H_2O	$Bi^{3+/0}$	26	*$1.85\cdot10^{-3}$*	IP	60Mou
	0.5 **M** HCl/H_2O	$Cd^{2+/0}$	23.7	*0.165*	ACP	58Bau
	0.5 **M** HCl/H_2O	$Cd^{2+/0}$	25	*≈0.5*	ACP	60Bau
	0.5 **M** HCl/H_2O	$Cd^{2+/0}$	25	*≈10*	IP	65Bie
	1 **M** HCl/H_2O	$Cd^{2+/0}$	25	*≈0.6*	ACP	60Bau
	1 **N** HCl/H_2O	$Cd^{2+/0}$	25	*2.6*	IP	65Bie
	1 **M** HCl/H_2O	$Cd^{2+/0}$	24	*0.94*	IP	67Fri

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	1 **M** $HClO_4/H_2O$	$Cd^{2+/0}$	25	*0.35*	IP	65Bie
	1 **M** $HClO_4/H_2O$	$Cd^{2+/0}$	24	*0.25*	IP	67Fri
	0.1 **M** $HClO_4$ +0.9 **M** $NaClO_4/H_2O$	$Cd^{2+/0}$	25	*0.46*	IP	65Bie
	1 **N** H_2SO_4/H_2O	$Cd^{2+/0}$	24	*0.14*	IP	67Fri
	1 **M** H_2SO_4/H_2O	$Cd^{2+/0}$	24	*0.12*	IP	67Fri
	0.5 **M** H_2SO_4/H_2O	$Cd^{2+/0}$, $1\cdot10^{-3}$	24	*0.17*	ACP	68Moo
	1 **M** KBr/H_2O, pH=2	$Cd^{2+/0}$, $2\cdot10^{-5}$	22	*>0.2*	FR	61Bar1
	1 **M** KBr/H_2O	$Cd^{2+/0}$, $1.5\cdot10^{-3}$	27	*0.91*	FR	71Aga
	0.5 **M** KCl/H_2O	$Cd^{2+/0}$	25	*≈1.5*	IP	60Bau
	1 **M** KCl/H_2O, pH=2	$Cd^{2+/0}$	25	*≈2.9*	FR	60Bau
	1 **M** KCl/H_2O, pH=3	$Cd^{2+/0}$	25	*≈0.7*	FR	69deL
	1 **M** KCl/H_2O, pH=2	$Cd^{2+/0}$	25	*(≈2.9)*	FR	61Bar1
	1 **M** KI/H_2O, pH=2	$Cd^{2+/0}$, $2\cdot10^{-5}$	22	*>0.2*	FR	61Bar1
	1 **M** KI/H_2O	$Cd^{2+/0}$, $2\cdot10^{-5}$	22	*>1*	ACP	61Kam
	1 **M** KCl/H_2O	$Cd^{2+/0}$	24	*1.2*	IP	67Fri
	1 **N** KF/H_2O	$Cd^{2+/0}$, $2\cdot10^{-5}$	22	*(0.25)*	FR	61Bar1
	1 **N** KI/H_2O	$Cd^{2+/0}$, $2\cdot10^{-5}$	22	*>0.2*	FR	61Bar1
	1 **N** KI/H_2O	$Cd^{2+/0}$, $2\cdot10^{-5}$	22	*>1*	ACP	61Kam
	1 **M** KNO_3/H_2O	$Cd^{2+/0}$	25	*0.6*	ACP	61Kam
	1 **M** KNO_3/H_2O	$Cd^{2+/0}$	24	*0.063*	IP	67Fri
	1 **M** KNO_3/H_2O	$Cd^{2+/0}$	25	*≈0.4*	ACP	60Bau
	1 **M** KNO_3/H_2O	$Cd^{2+/0}$	25	*2.6*	IP	65Bie
	1 **M** KNO_3/H_2O	$Cd^{2+/0}$	24	*0.63*	ACP	67Fri
	1 **N** K_2SO_4/H_2O	$Cd^{2+/0}$, $2\cdot10^{-5}$	22	*0.08*	FR	61Bar1
	1 **N** K_2SO_4/H_2O, pH=2	Cd^{2+}	22	*(0.08)*	FR	61Bar1
	1 **M** $NaBr/H_2O$	$Cd^{2+/0}$	25	*$7.7\cdot10^{-3}$*	IP	70Mou
	1 **M** $NaCl/H_2O$	$Cd^{2+/0}$	24	*1.2*	IP	67Fri
	1 **M** $NaCl/H_2O$	$Cd^{2+/0}$	25	*$3.9\cdot10^{-3}$*	IP	70Mou
	1 **M** $NaClO_4/H_2O$, pH=2	$Cd^{2+/0}$	22	*2.3*	FR	58Bar1
	1 **M** $NaClO_4/H_2O$, pH=2	$Cd^{2+/0}$	22	*(0.43)*	FR	61Bar1
	1 **M** $NaClO_4/H_2O$, pH=3	$Cd^{2+/0}$	25	*0.45*	FR	69deL
	1 **M** $NaClO_4/H_2O$	$Cd^{2+/0}$	24	*0.34*	IP	67Fri
	1 **M** $NaClO_4/H_2O$	$Cd^{2+/0}$, $1\cdot10^{-3}$	24	*0.33*	ACP	68Moo
	1 **M** $NaClO_4/H_2O$	$Cd^{2+/0}$, $3\cdot10^{-3}$	27	*1.2*	FR	71Aga
	1 **M** NaN_3/H_2O	$Cd^{2+/0}$	25	*$7.5\cdot10^{-3}$*	IP	70Mou
	1 **N** Na_2SO_4/H_2O	$Cd^{2+/0}$	20	*$2.6\cdot10^{-2}$*	PS	57Vie
	1 **N** Na_2SO_4/H_2O	$Cd^{2+/0}$	25	*0.25*	ACP	60Bau
	1 **N** Na_2SO_4/H_2O	$Cd^{2+/0}$	-	*$8.5\cdot10^{-2}$*	IP	61Ran
	1 **N** Na_2SO_4/H_2O	$Cd^{2+/0}$	25	*$3.6\cdot10^{-2}$*	PS	63Tos
	1 **N** Na_2SO_4/H_2O	$Cd^{2+/0}$	25	*$3.8\cdot10^{-2}$*	GS	65Tos

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	1 **N** Na_2SO_4/H_2O	$Cd^{2+/0}$	24	*0.076*	IP	67Fri
	1 **M** Na_2SO_4/H_2O	$Cd^{2+/0}$	24	*0.063*	IP	67Fri
	1 **M** Na_2SO_4/H_2O	$Cd^{2+/0}$	24	*0.24*	CP	65Nic
	1 **M** Na_2SO_4/H_2O	$Cd^{2+/0}$	25	*0.15*	ACP	72Hue
	1 **M** Na_2SO_4+ $5\cdot10^{-3}$ **M** *n*-hexanol/H_2O	$Cd^{2+/0}$	30	*$2.2\cdot10^{-4}$*	ACP	58Del
	1 **M** Na_2SO_4+10^{-2} **M** *n*-hexanol/H_2O	$Cd^{2+/0}$	30	*$1.0\cdot10^{-4}$*	ACP	58Del
	0.9 **M** $NaClO_4/H_2O$	$Cd^{2+/0}$	25	*0.45*	IP	71Bie
	0.9 **M** $NaClO_4$ /methanol	$Cd^{2+/0}$	25	*0.013*	IP	71Bie
	0.9 **M** $NaClO_4$ /propylene carbonate	$Cd^{2+/0}$	25	*~0.1*	IP	71Bie
	0.9 **M** $NaClO_4$ /DMF [54])	$Cd^{2+/}$	25	*0.15*	IP	71Bie
	0.9 **M** $NaClO_4$ /NMF [88])	$Cd^{2+/0}$	25	*0.12*	IP	71Bie
	1 **M** $NaClO_4/H_2O$	$Cd^{2+/0}$, $2\cdot10^{-5}$	22	*(0.43)*	FR	61Bar1
	1 **M** $NaClO_4/H_2O$ +CH_3CN 0.5 mol +0.5 mol	$Cd^{2+/0}$	25	*0.0145*	IP	71Bie
	1 **M** $NaClO_4/CH_3CN$	$Cd^{2+/0}$	25	*0.19*	IP	71Bie
	1 **M** $NaClO_4/CH_3CN$ +CH_3OH 0.8 mol +0.2 mol	$Cd^{2+/0}$	25	*0.08*	IP	71Bie
	1 **M** $NaClO_4/CH_3CN$ +CH_3OH 0.5 mol +0.5 mol	$Cd^{2+/0}$	25	*0.028*	IP	71Bie
	1 **M** $NaClO_4/CH_3OH$	$Cd^{2+/0}$	25	*0.013*	IP	71Bie
	1 **M** $NaClO_4$/DMF [54]) +NMF [88]) 0.8 mol +0.2 mol	$Cd^{2+/0}$	25	*0.081*	IP	71Bie
	1 **M** $NaClO_4$/DMF [54]) +NMF [88]) 0.5 mol +0.5 mol	$Cd^{2+/0}$	25	*0.089*	IP	71Bie
	1 **M** $NaClO_4$/DMF [54]) +NMF [88]) 0.2 mol +0.8 mol	$Cd^{2+/0}$	25	*0.108*	IP	71Bie
	1 **M** KCl/H_2O	$Co^{2+/0}$	25	*$1.99\cdot10^{-10}$*	OP	55Mat
	0.1 **M** $NaCl/H_2O$	$Co^{2+/0}$	30	*$8.0\cdot10^{-17}$*	DCP	70Gau
	0.1 **M** $NaClO_4/H_2O$	$Co^{2+/0}$	30	*$1.5\cdot10^{-16}$*	DCP	70Gau
	0.1 **M** $NaSO_4/H_2O$	$Co^{2+/0}$	30	*$5.6\cdot10^{-16}$*	DCP	70Gau
	0.1 **M** NMe_4Br/H_2O	$Co^{2+/0}$	30	*$4.0\cdot10^{-16}$*	DCP	70Gau
	0.1 **M** $NaClO_4$ /formamide	$Co^{2+/0}$	30	*$2.8\cdot10^{-15}$*	DCP	70Gau

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.1 **M** $NaClO_4$ /formamide (20vol%)+H_2O	$Co^{2+/0}$	30	*1.0·10^{-14}*	DCP	70Gau
	0.1 **M** $NaClO_4$ /formamide (40vol%)+H_2O	$Co^{2+/0}$	30	*5.6·10^{-15}*	DCP	70Gau
	0.1 **M** $NaClO_4$ /formamide (60vol%)+H_2O	$Co^{2+/0}$	30	*4.5·10^{-15}*	DCP	70Gau
	0.1 **M** $NaClO_4$ /formamide (80vol%)+H_2O	$Co^{2+/0}$	30	*1.4·10^{-14}*	DCP	70Gau
	0.1 **M** $NaClO_4$ /DMF [54])	$Co^{2+/0}$	30	*5.0·10^{-13}*	DCP	70Gau
	0.1 **M** $NaClO_4$ /DMF(20vol%) +H_2O	$Co^{2+/0}$	30	*1.0·10^{-14}*	DCP	70Gau
	0.1 **M** $NaClO_4$ /DMF(40vol%) +H_2O	$Co^{2+/0}$	30	*3.5·10^{-17}*	DCP	70Gau
	0.1 **M** $NaClO_4$ /DMF(60vol%) +H_2O	$Co^{2+/0}$	30	*3.2·10^{-16}*	DCP	70Gau
	0.1 **M** $NaClO_4$ /DMF(80vol%) +H_2O	$Co^{2+/0}$	30	*3.5·10^{-15}*	DCP	70Gau
	0.1 **M** $NaClO_4$ /acetonitrile	$Co^{2+/0}$	30	*2.8·10^{-6}*	DCP	70Gau
	0.1 **M** $NaClO_4$ /acetonitrile (40vol%)+H_2O	$Co^{2+/0}$	30	*1.4·10^{-10}*	DCP	70Gau
	0.1 **M** $NaClO_4$ /acetonitrile (60vol%)+H_2O	$Co^{2+/0}$	30	*4.5·10^{-9}*	DCP	70Gau
	0.1 **M** $NaClO_4$ /acetonitrile (80vol%)+H_2O	$Co^{2+/0}$	30	*7.0·10^{-13}*	DCP	70Gau
	0.1 **M** KPF_6/H_2O	$[Co(NH_3)_6]^{3+/2+}$	25	*(2.7·10^{-4})*	CV	80Bar1
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	25	*(3.2·10^{-3})*	CV	80Bar1
	1 **M** KF/H_2O	$[Co(NH_3)_5]F^{2+/+}$	25	*(1.5·10^{-5})*	DCP	78Sat
	0.1 **M** KPF_6/H_2O	$[Co(NH_3)_5]F^{2+/+}$	25	*(4.5·10^{-4})*	CV	80Bar1
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5]F^{2+/+}$	25	*(2.0·10^{-3})*	CV	80Bar1
	1 **M** NH_4F/H_2O	$[Co(NH_3)_5]F^{2+/+}$	25	*(1.5·10^{-5})*	DCP	78Sat
	0.2 **M** NaF/H_2O	$[Co(NH_3)_5]F^{2+/+}$	25	*(8.0·10^{-6})*	DCP	78Sat
	0.2 **M** KF/H_2O	$[Co(NH_3)_5]F^{2+/+}$	25	*(1.0·10^{-5})*	DCP	78Sat
	0.905 **M** NaF/H_2O	$[Co(NH_3)_5]F^{2+/+}$	25	*(1.7·10^{-5})*	DCP	78Sat
	1 **M** KF/H_2O	$[Co(NH_3)_5]OH^{2+/+}$	25	*(1.0·10^{-5})*	DCP	78Sat

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	1 **M** KF/H_2O	([$Co(NH_3)_5$] NO_3)$^{2+/+}$	25	*(3.2·10^{-2})*	DCP	78Sat
	1 **M** KF/H_2O	([$Co(NH_3)_5$]N_3)$F^{2+/+}$	25	*(1.7·10^{-2})*	DCP	78Sat
	1 **M** KF/H_2O	[$Co(NH_3)_5$]$NCS^{2+/+}$	25	*(1.1·10^{-3})*	DCP	78Sat
	1 **M** KF/H_2O	[$Co(NH_3)_6$]$^{3+/2+}$	25	*(8.0·10^{-7})*	DCP	78Sat
	1 **M** NH_4F/H_2O	[$Co(NH_3)_6$]$^{3+/2+}$	25	*(7.5·10^{-7})*	DCP	78Sat
	1 **M** KF/H_2O	[$Co(en)_3$]$F^{2+/+}$	25	*(1.5·10^{-5})*	DCP	78Sat [63])
	1 **M** KF/H_2O	cis-([$Co(en)_2$]$(N_3)_2$)$^{+/0}$	25	*(5.0·10^{-3})*	DCP	78Sat [63])
	1 **M** KF/H_2O	trans-([$Co(en)_2$]$(N_3)_2$)$^{+/0}$	25	*(8.0·10^{-4})*	DCP	78Sat [63])
	1 **M** KF/H_2O	[$Co(NH_3)_5OSO_3$]$^{+/0}$	25	*(4.1·10^{-4})*	DCP	78Sat
	1 **M** NH_4F/H_2O	[$Co(NH_3)_5OSO_3$]$^{+/0}$	25	*(4.5·10^{-4})*	DCP	78Sat
	1 **M** KF/H_2O	[$Co(NH_3)_5Cl$]$^{2+/+}$	25	*(>4.0·10^{-2})*	DCP	78Sat
	0.1 **M** KPF_6/H_2O	[$Co(NH_3)_5$] $OH_2^{3+/2+}$	25	*(1.6·10^{-4})*	CV	80Bar1
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4$/H_2O	[$Co(NH_3)_5$] $NCS^{2+/1+}$	24	*(1.0·10^{-3})*	PP	84Bar2
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] $HCOO)^{2+/1+}$	24	*(3.2·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] $CH_3COO)^{2+/1+}$	24	*(8.5·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] $CCl_3COO)^{2+/1+}$	24	*(4.2·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] $CF_3COO)^{2+/1+}$	24	*(6.5·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] $(CH_3)_3COO)^{2+/1+}$	24	*(4.0·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$]CH_3C $(O)COO)^{2+/1+}$	24	*(4.0·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] $HC(O)COO)^{2+/1+}$	24	*(1.5·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] $CH_2OHCOO)^{2+/1+}$	24	*(2.4·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] $CH_3C(OH)$ $HCOO)^{2+/1+}$	24	*(1.1·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] $CH_2FCOO)^{2+/1+}$	24	*(2.0·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] cyclopentane carboxylate)$^{2+/1+}$	24	*(1.1·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$]cyclopentane acetate)$^{2+/1+}$	24	*(2.5·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] cyclohexane carboxylate)$^{2+/1+}$	24	*(8.5·10^{-4})*	PP	85Li

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] cyclohexane acetate)$^{2+/1+}$	24	*(2.0·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] cyclohexane propionate)$^{2+/1+}$	24	*(6.1·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] benzoate)$^{2+/1+}$	24	*(1.0·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] phenylacetate)$^{2+/1+}$	24	*(1.1·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] 3-phenyl-propionate)$^{2+/1+}$	24	*(8.7·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] 2-phenyl-acrylate)$^{2+/1+}$	24	*(1.2·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] benzoate)$^{2+/1+}$	24	*(1.2·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] furan-3-carboxylate)$^{2+/1+}$	24	*(1.1·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] furan-2-carboxylate)$^{2+/1+}$	24	*(1.5·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$]furan-2-acrylate)$^{2+/1+}$	24	*(1.6·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] pyridine-2-carboxylate)$^{2+/1+}$	24	*(1.5·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] pyridine-3-carboxylate)$^{2+/1+}$	24	*(1.3·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] pyridine-2-acetate)$^{2+/1+}$	24	*(8.9·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] 3-pyridyl-acrylate)$^{2+/1+}$	24	*(7.2·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] thiophene-2-carboxylate)$^{2+/1+}$	24	*(2.3·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] thiophene-2-acetate)$^{2+/1+}$	24	*(7.3·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] thiophene-2-propionate)$^{2+/1+}$	24	*(7.0·10^{-3})*	PP	85Li

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] thiophene-2-butyrate)$^{2+/1+}$	24	*(3.0·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] thiophene-3-carboxylate)$^{2+/1+}$	24	*(2.8·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] thiophene-2-acetate)$^{2+/1+}$	24	*(7.5·10^{-3})*	PP	85Li
	0.1 **M** $NaClO_4$+/H_2O	([$Co(NH_3)_5$] thiophene-2-acrylate)$^{2+/1+}$	24	*(3.2·10^{-2})*	PP	85Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	([$Co(NH_3)_5$] thiophene-2-carboxylate)$^{2+/1+}$	24	*(1.8·10^{-3})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	([$Co(NH_3)_5$] thiophene-2-acetat)$^{2+/1+}$	24	*(4.5·10^{-4})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	([$Co(NH_3)_5$] thiophene-2-propionate)$^{2+/1+}$	24	*(6.3·10^{-4})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	([$Co(NH_3)_5$] thiophene-2-butyrate)$^{2+/1+}$	24	*(3.2·10^{-4})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	([$Co(NH_3)_5$] thiophene-2-acrylate)$^{2+/1+}$	24	*(2.8·10^{-3})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	([$Co(NH_3)_5$] thiophene-3-carboxylate)$^{2+/1+}$	24	*(1.7·10^{-3})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	([$Co(NH_3)_5$] thiophene-3-acetate)$^{2+/1+}$	24	*(4.9·10^{-4})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	([$Co(NH_3)_5$] benzothiophene-3-carboxylate)$^{2+/1+}$	24	*(2.7·10^{-4})*	PP	84Li
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	([$Co(NH_3)_5$] benzothiophene-3-malonate)$^{2+/1+}$	24	*(2.6·10^{-4})*	PP	84Li
	0.025 **M** TBAP [52]) /DMF [54])	[$Co(Cp)_2$] (o) [87])	-	*2.75*	ACP	93Win
	0.05 **M** TBAP [52]) /DMF [54])	[$Co(Cp)_2$] (o)	-	*1.55*	ACP	93Win
	0.1 **M** TBAP [52]) /DMF [54])	[$Co(Cp)_2$] (o)	-	*0.89*	ACP	93Win
	0.25 **M** TBAP [52]) /DMF [54])	[$Co(Cp)_2$] (o)	-	*0.62*	ACP	93Win

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.1 **M** TBAP [52]) /acetonitrile	$Co(Cp)_2^{+/0}$	23	*3.0*	ACP	85Gen
	0.1 **M** TBAP/acetone	$Co(Cp)_2^{+/0}$	23	*2.5*	ACP	85Gen
	0.1 **M** TBAP /formamide	$Co(Cp)_2^{+/0}$	23	*0.55*	ACP	85Gen
	0.1 **M** TBAP /NMF [88])	$Co(Cp)_2^{+/0}$	23	*0.6*	ACP	85Gen
	0.1 **M** TBAP /DMF [54])	$Co(Cp)_2^{+/0}$	23	*1.0*	ACP	85Gen
	0.1 **M** TBAP /dimethyl sulfoxide	$Co(Cp)_2^{+/0}$	23	*0.55*	ACP	85Gen
	0.1 **M** TBAP /benzonitrile	$Co(Cp)_2^{+/0}$	23	*0.25*	ACP	85Gen
	0.1 **M** TEAP [74]) /acetonitrile	$Co(Cp)_2^{+/0}$	23	*5.0*	ACP	86McM
	0.1 **M** TEAP [74]) /CH_2Cl_2	$Co(Cp)_2^{+/0}$	23	*5.0*	ACP	86McM
	0.1 **M** TEAP [74]) /formamide	$Co(Cp)_2^{+/0}$	23	*0.6*	ACP	86McM
	0.1 **M** TEAP [74]) /NMF [88])	$Co(Cp)_2^{+/0}$	23	*0.45*	ACP	86McM
	0.1 **M** TEAP [74]) /DMF [54])	$Co(Cp)_2^{+/0}$	23	*2.0*	ACP	86McM
	0.1 **M** TEAP [74]) /dimethylacetamide	$Co(Cp)_2^{+/0}$	23	*2.0*	ACP	86McM
	0.1 **M** TEAP [74]) /tetramethylurea	$Co(Cp)_2^{+/0}$	23	*0.35*	ACP	86McM
	0.1 **M** TEAP [74]) /dimethyl sulfoxide	$Co(Cp)_2^{+/0}$	23	*0.75*	ACP	86McM
	0.1 **M** TEAP [74]) /benzonitrile	$Co(Cp)_2^{+/0}$	23	*0.6*	ACP	86McM
	0.1 **M** TEAP [74])/PC [113])	$Co(Cp)_2^{+/0}$	23	*1.0*	ACP	86McM
	0.1 **M** TBAFB [89]) /methanol	$Co(Cp)_2^{+/0}$	23	*6.0*	ACP	86McM
	0.1 **M** TBAFB /ethanol	$Co(Cp)_2^{+/0}$	23	*4.0*	ACP	86McM
	0.1 **M** TBAFB /1-propanol	$Co(Cp)_2^{+/0}$	23	*1.0*	ACP	86McM
	0.1 **M** TBAP [52]) /acetonitrile	$Co(Cp)_2^{0/-}$	22	*0.112*	IP	92Faw1
	0.1 **M** TBAP [52]) /butyronitrile	$Co(Cp)_2^{0/-}$	22	*0.079*	IP	92Faw1
	0.1 **M** TBAP [52]) /dimethylacetamide	$Co(Cp)_2^{0/-}$	22	*0.126*	IP	92Faw1
	0.1 **M** TBAP /dimethyl sulfoxide	$Co(Cp)_2^{0/-}$	22	*0.04*	IP	92Faw1

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	0.1 **M** TBAP [52]) /PC [113])	$Co(Cp)_2^{0/-}$	22	*0.034*	IP	92Faw1
	0.1 **M** TBAP /acetonitrile	$Co(Cp')_2^{+/0}$ [90])	23	*3.5*	ACP	85Gen
	0.1 **M** TBAP/acetone	$Co(Cp')_2^{+/0}$	23	*3.0*	ACP	85Gen
	0.1 **M** TBAP/NMF [88])	$Co(Cp')_2^{+/0}$	23	*0.8*	ACP	85Gen
	0.1 **M** TBAP/DMF [54])	$Co(Cp')_2^{+/0}$	23	*1.5*	ACP	85Gen
	0.1 **M** TBAP /dimethyl sulfoxide	$Co(Cp')_2^{+/0}$	23	*0.75*	ACP	85Gen
	0.5 **M** Na_2SO_4/H_2O	*trans*-$([Co(SO_3)_2(CN)_4)^{5-/7-}$	15	*$(2.67\cdot10^{-13})$*	DCP	70Mak
	0.5 **M** Na_2SO_4/H_2O	*trans*-$([Co(SO_3)_2(CN)_4)^{5-/7-}$	20	*$(3.32\cdot10^{-13})$*	DCP	70Mak
	0.5 **M** Na_2SO_4/H_2O	*trans*-$([Co(SO_3)_2(CN)_4)^{5-/7-}$	25	*$(5.29\cdot10^{-13})$*	DCP	70Mak
	0.5 **M** Na_2SO_4/H_2O	*trans*-$([Co(SO_3)_2(CN)_4)^{5-/7-}$	30	*$(6.59\cdot10^{-13})$*	DCP	70Mak
	0.5 **M** Na_2SO_4/H_2O	*trans*-$([Co(SO_3)_2(CN)_4)^{5-/7-}$	35	*$(9.12\cdot10^{-13})$*	DCP	70Mak
	0.5 **M** Na_2SO_4/H_2O	$([Co(SO_3)(CN)_5)^{4-/6-}$	10	*$(1.66\cdot10^{-17})$*	DCP	70Mak
	0.5 **M** Na_2SO_4/H_2O	$([Co(SO_3)(CN)_5)^{4-/6-}$	15	*$(3.06\cdot10^{-17})$*	DCP	70Mak
	0.5 **M** Na_2SO_4/H_2O	$([Co(SO_3)(CN)_5)^{4-/6-}$	20	*$(6.26\cdot10^{-17})$*	DCP	70Mak
	0.5 **M** Na_2SO_4/H_2O	$([Co(SO_3)(CN)_5)^{4-/6-}$	30	*$(1.5\cdot10^{-17})$*	DCP	70Mak
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$	25	*$1.0\cdot10^{-5}$*	IP	52Ran2
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$	25	*$7.9\cdot10^{-8}$*	OP	55Mat
	0.4 **M** KPF_6/H_2O	$Cr_{aq}^{2+/3+}$	25	*$(2.1\cdot10^{-3})$*	DME	80Bar1
	0.4 **M** $NaClO_4/H_2O$	$Cr_{aq}^{2+/3+}$	25	*$(2.0\cdot10^{-2})$*	CV	80Bar1
	1 **M** $NaClO_4/H_2O$	$Cr_{aq}^{3+/2+}$	25	*$(0.75\cdot10^{5})$*	CM	76Wea
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$, 0.018; 0.082 **M**	24	*$0.95\cdot10^{-4}$*	OC	65Sri
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$, 0.035; 0.065 **M**	24	*$1.4\cdot10^{-4}$*	OC	65Sri
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$, 0.076; 0.024 **M**	24	*$2.2\cdot10^{-4}$*	OC	65Sri
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$, 0.08; 0.02 **M**	24	*$1.6\cdot10^{-4}$*	OC	65Sri

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$, 0.085; 0.015 **M**	24	*1.8·10^{-4}*	OC	65Sri
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$, 0.003; 0.037 **M**	24	*2.1·10^{-4}*	OC	65Sri
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$, 0.008; 0.032 **M**	24	*1.7·10^{-4}*	OC	65Sri
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$, 0.012; 0.028 **M**	24	*1.8·10^{-4}*	OC	65Sri
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$, 0.017; 0.023 **M**	24	*1.7·10^{-4}*	OC	65Sri
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$, 0.023; 0.017 **M**	24	*1.9·10^{-4}*	OC	65Sri
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$, 0.027; 0.013 **M**	24	*2.2·10^{-4}*	OC	65Sri
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$, 0.031; 0.009 **M**	24	*2.0·10^{-4}*	OC	65Sri
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$, 0.036; 0.004**M**	24	*2.3·10^{-4}*	OC	65Sri
	KCl/H_2O	Cr(VI)/Cr(IV)	-	*2.7·10^{-6}*	DCP	61Pam
	KCl/H_2O	Cr(IV)/Cr(III)	-	*2.7·10^{-6}*	DCP	61Pam
	KCl/H_2O	Cr(III)/Cr(II)	-	*8.0·10^{-7}*	DCP	61Pam
	0.08 **M** $CsClO_4/H_2O$	Cr(III)/Cr(II)	25	*2.1·10^{-5}*	DCP	70Ans
	0.5 **M** KBr/H_2O	Cr(III)/Cr(II)	25	*1.7·10^{-5}*	DCP	70Ans
	1 **M** KBr/H_2O	Cr(III)/Cr(II)	25	*1.4·10^{-5}*	DCP	70Ans
	0.5 **M** KCl/H_2O	Cr(III)/Cr(II)	25	*8.5·10^{-6}*	DCP	70Ans
	0.5 **M** KCl/H_2O	Cr(III)/Cr(II)	25	*8.5·10^{-6}*	DCP	71deK
	0.5 **M** KI/H_2O	Cr(III)/Cr(II)	25	*9.8·10^{-5}*	DCP	70Ans
	0.0883 **M** $LaCl_3/H_2O$	Cr(III)/Cr(II)	25	*6.5·10^{-6}*	DCP	70Ans
	0.1666 **M** $LaCl_3/H_2O$	Cr(III)/Cr(II)	25	*2.6·10^{-6}*	DCP	70Ans
	0.5 **M** $LiClO_4/H_2O$	Cr(III)/Cr(II)	25	*6.5·10^{-6}*	DCP	70Ans
	0.5 **M** $LiClO_4/H_2O$	Cr(III)/Cr(II)	25	*6.5·10^{-6}*	DCP	71deK
	1 **M** $LiClO_4/H_2O$	Cr(III)/Cr(II)	25	*3.6·10^{-6}*	DCP	70Ans
	1 **M** $LiClO_4/H_2O$	Cr(III)/Cr(II)	25	*3.6·10^{-6}*	DCP	71deK
	0.5 **M** $NaClO_4/H_2O$	Cr(III)/Cr(II)	25	*8.1·10^{-6}*	DCP	70Ans
	0.5 **M** $NaClO_4/H_2O$	Cr(III)/Cr(II)	25	*8.1·10^{-6}*	DCP	71deK
	0.5 **M** $NaClO_4/H_2O$, pH=3.4	Cr(III)/Cr(II)	25	*7.1·10^{-6}*	DCP	66Par2
	0.5 **M** $NaClO_4/H_2O$, pH=3.4	Cr(III)/Cr(II)	35	*1.4·10^{-5}*	DCP	66Par2
	0.5 **M** $NaClO_4/H_2O$, pH=3.4	Cr(III)/Cr(II)	45	*2.5·10^{-5}*	DCP	66Par2

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	0.9 **M** $NaClO_4$ /+0.1 **M** $HClO_4/H_2O$	Cr(III)/Cr(II)	-	*1.0·10^{-5}*	DCP	62Ran
	1 **M** $NaClO_4/H_2O$	Cr(III)/Cr(II)	25	*7.9·10^{-6}*	DCP	70Ans
	1 **M** $NaClO_4/H_2O$	Cr(III)/Cr(II)	25	*7.9·10^{-6}*	DCP	71deK
	0.1 **M** $NaOH/H_2O$	CrO_4^{2-} (r)	25	*(1.2·10^{-9})*	DCP	61Mei
	0.1 **M** $NaOH/H_2O$	CrO_4^{2-} (r)	-	*(1.8·10^{-9})*	DCP	71Sad
	0.1 **M** NaOH /10 vol% methanol/H_2O	CrO_4^{2-} (r)	-	*(1.9·10^{-11})*	DCP	71Sad
	0.1 **M** NaOH /20 vol% methanol/H_2O	CrO_4^{2-} (r)	-	*(1.6·10^{-11})*	DCP	71Sad
	0.1 **M** NaOH /30 vol% methanol/H_2O	CrO_4^{2-} (r)	-	*(1.1·10^{-11})*	DCP	71Sad
	0.1 **M** NaOH /40 vol% methanol/H_2O	CrO_4^{2-} (r)	-	*(2.1·10^{-11})*	DCP	71Sad
	0.1 **M** NaOH /50 vol% methanol/H_2O	CrO_4^{2-} (r)	-	*(4.1·10^{-11})*	DCP	71Sad
	0.1 **M** NaOH /60 vol% methanol/H_2O	CrO_4^{2-} (r)	-	*(4.1·10^{-11})*	DCP	71Sad
	0.1 **M** NaOH /70 vol% methanol/H_2O	CrO_4^{2-} (r)	-	*(1.1·10^{-11})*	DCP	71Sad
	0.1 **M** NaOH /10 vol% ethanol/H_2O	CrO_4^{2-} (r)	-	*(3.5·10^{-13})*	DCP	71Sad
	0.1 **M** NaOH /20 vol% ethanol/H_2O	CrO_4^{2-} (r)	-	*(2.4·10^{-15})*	DCP	71Sad
	0.1 **M** NaOH /30 vol% ethanol/H_2O	CrO_4^{2-} (r)	-	*(8.5·10^{-16})*	DCP	71Sad
	0.1 **M** NaOH /40 vol% ethanol/H_2O	CrO_4^{2-} (r)	-	*(7.4·10^{-17})*	DCP	71Sad
	0.1 **M** NaOH /50 vol% ethanol/H_2O	CrO_4^{2-} (r)	-	*(2.2·10^{-15})*	DCP	71Sad

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.1 **M** NaOH /60 vol% ethanol/H_2O	CrO_4^{2-} (r)	-	*(6.9·10^{-15})*	DCP	71Sad
	0.1 **M** NaOH /70 vol% ethanol/H_2O	CrO_4^{2-} (r)	-	*(7.2·10^{-16})*	DCP	71Sad
	0.1 **M** NaOH /10 vol% propanol/H_2O	CrO_4^{2-} (r)	-	*(1.1·10^{-16})*	DCP	71Sad
	0.1 **M** NaOH /20 vol% propanol/H_2O	CrO_4^{2-} (r)	-	*(1.1·10^{-18})*	DCP	71Sad
	0.1 **M** NaOH /30 vol% propanol/H_2O	CrO_4^{2-} (r)	-	*(1.6·10^{-17})*	DCP	71Sad
	0.1 **M** NaOH /40 vol% propanol/H_2O	CrO_4^{2-} (r)	-	*(3.0·10^{-15})*	DCP	71Sad
	0.1 **M** NaOH /50 vol% propanol/H_2O	CrO_4^{2-} (r)	-	*(2.7·10^{-14})*	DCP	71Sad
	0.1 **M** NaOH /60 vol% propanol/H_2O	CrO_4^{2-} (r)	-	*(5.8·10^{-16})*	DCP	71Sad
	0.1 **M** NaOH/70 vol% propanol/H_2O	CrO_4^{2-} (r)	-	*(1.1·10^{-15})*	DCP	71Sad
	0.1 **M** NaOH/10 vol% dioxane/H_2O	CrO_4^{2-} (r)	-	*(5.3·10^{-18})*	DCP	71Sad
	0.1 **M** NaOH/20 vol% dioxane/H_2O	CrO_4^{2-} (r)	-	*(2.2·10^{-16})*	DCP	71Sad
	0.1 **M** NaOH/30 vol% dioxane/H_2O	CrO_4^{2-} (r)	-	*(8.4·10^{-18})*	DCP	71Sad
	0.1 **M** NaOH/40 vol% dioxane/H_2O	CrO_4^{2-} (r)	-	*(8.4·10^{-18})*	DCP	71Sad
	0.1 **M** NaOH/50 vol% dioxane/H_2O	CrO_4^{2-} (r)	-	*(1.8·10^{-17})*	DCP	71Sad
	0.1 **M** NaOH/60 vol% dioxane/H_2O	CrO_4^{2-} (r)	-	*(4.6·10^{-17})*	DCP	71Sad

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	0.1 **M** KCN/H_2O	$Cr(CN)_6^{3-/4-}$	30	*0.018*	IP	62Del4
	0.1 **M** KCN +0.1 **M** KCl/H_2O	$Cr(CN)_6^{3-/4-}$	25	*2.2·10^{-2}*	PS	70Yam
	0.1 **M** KCN +0.1 **M** KCl/H_2O	$Cr(CN)_6^{3-/4-}$	25	*0.052*	IP	62Del4
	0.1 **M** KCN +0.3 **M** KCl/H_2O	$Cr(CN)_6^{3-/4-}$	25	*0.14*	IP	62Del4
	0.1 **M** KCN +0.5 **M** KCl/H_2O	$Cr(CN)_6^{3-/4-}$	25	*0.31*	IP	62Del4
	0.1 **M** KCN +0.7 **M** KCl/H_2O	$Cr(CN)_6^{3-/4-}$	25	*0.91*	IP	62Del4
	0.5 **M** KCN +3 **M** KCl/H_2O	$Cr(CN)_6^{3-/4-}$	25	*1.25*	FR	62Ima1
	1 **M** KCN/H_2O	$Cr(CN)_6^{3-/4-}$	20	*(0.25)*	IP	52Ran2
	1 **M** KCN/H_2O	$Cr(CN)_6^{3-/4-}$	30	*0.22*	ACP	61Smi
	0.1 **M** KSCN/H_2O	$Cr(SCN)_6^{3-/4-}$	-	*4.0·10^{-3}*	DCP	62Kis
	0.3 **M** KSCN/H_2O	$Cr(SCN)_6^{3-/4-}$	-	*4.0·10^{-3}*	DCP	62Kis
	1 **M** KSCN/H_2O	$Cr(SCN)_6^{3-/4-}$	-	*2.0·10^{-3}*	DCP	62Kis
	3 **M** KSCN/H_2O	$Cr(SCN)_6^{3-/4-}$	-	*7.0·10^{-3}*	DCP	62Kis
	7 **M** KSCN/H_2O	$Cr(SCN)_6^{3-/4-}$	-	*1.0·10^{-2}*	DCP	62Kis
	1 **M** $NaClO_4$/H_2O	$[Cr(OH_2)F^-]^{2+/1+}$	25	*(0.6·10^5)*	CM	76Wea
	1 **M** $NaClO_4$/H_2O	$[Cr(OH_2)SO_4^{2-}]^{1+/0}$	25	*(0.46·10^5)*	CM	76Wea
	1 **M** $NaClO_4$/H_2O	$[Cr(OH_2)Cl^-]^{2+/1+}$	25	*(1.2·10^3)*	CM	76Wea
	1 **M** $NaClO_4$/H_2O	$[Cr(OH_2)Br^-]^{2+/1+}$	25	*(7.·10^4)*	CM	76Wea
	1 **M** $NaClO_4$/H_2O	$[Cr(OH)_2NCS^-]^{2+/1+}$	25	*(4.2·10^2)*	CM	76Wea
	1 **M** $NaClO_4$/H_2O	$[Cr(OH_2)N_3^-]^{2+/1+}$	25	*(12)*	CM	76Wea
	1 **M** $NaClO_4$/H_2O	$[Cr(OH_2)NO_3^-]^{2+/1+}$	25	*(35)*	CM	76Wea
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$/H_2O	$[Cr(NH_3)_5]NCS^{2+/1+}$	24	*(1.0·10^{-5})*	PP	84Guy
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$/H_2O	$[Cr(OH_2)_5]NCS^{2+/1+}$	24	*(7.5·10^{-4})*	PP	84Guy
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$/H_2O	$([Cr(NH_3)_5]N_3)^{2+/1+}$	24	*(5.0·10^{-7})*	PP	84Guy
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$/H_2O	$([Cr(OH_2)_5]N_3)^{2+/1+}$	24	*(2.0·10^{-5})*	PP	84Guy

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	$[Cr(NH_3)_5]Cl^{2+/1+}$	24	*(1.0·10^{-5})*	PP	84Guy
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	$[Cr(OH_2)_5]Cl^{2+/1+}$	24	*(1.5·10^{-3})*	PP	84Guy
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	$[Cr(NH_3)_5]Br^{2+/1+}$	24	*(6.0·10^{-3})*	PP	84Guy
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4/H_2O$	$[Cr(OH_2)_5]Br^{2+/1+}$	24	*(>2.5)*	PP	84Guy
	0.1 **M** TBAP [52]) /acetonitrile	$Cr(C_6H_6)_2^{+/0}$	23	*4.0*	ACP	85Gen
	0.1 **M** TBAP/acetone	$Cr(C_6H_6)_2^{+/0}$	23	*2.5*	ACP	85Gen
	0.1 **M** TBAP /formamide	$Cr(C_6H_6)_2^{+/0}$	23	*0.55*	ACP	85Gen
	0.1 **M** TBAP /NMF [91])	$Cr(C_6H_6)_2^{+/0}$	23	*0.65*	ACP	85Gen
	0.1 **M** TBAP /DMF [54])	$Cr(C_6H_6)_2^{+/0}$	23	*1.2*	ACP	85Gen
	0.1 **M** TBAP /dimethyl sulfoxide	$Cr(C_6H_6)_2^{+/0}$	23	*0.5*	ACP	85Gen
	0.1 **M** TBAP /benzonitrile	$Cr(C_6H_6)_2^{+/0}$	23	*0.3*	ACP	85Gen
	0.1 **M** TBAP /acetonitrile	$Cr(acac)_3^{0/-}$ [92])	23	*1.37·10^{-2}*	PP/CV	95Urb
	0.1 **M** TBAP/PC [113])	$Cr(acac)_3^{0/-}$	23	*0.56·10^{-2}*	PP/CV	95Urb
	0.1 **M** TBAP /butyrolactone	$Cr(acac)_3^{0/-}$	23	*0.57·10^{-2}*	PP/CV	95Urb
	0.1 **M** TBAP/DMF	$Cr(acac)_3^{0/-}$	23	*0.96·10^{-2}*	PP/CV	95Urb
	0.1 **M** TBAP /dimetylacetamide	$Cr(acac)_3^{0/-}$	23	*0.75·10^{-2}*	PP/CV	95Urb
	0.1 **M** TBAP /dimethylsulfoxide	$Cr(acac)_3^{0/-}$	23	*0.53·10^{-2}*	PP/CV	95Urb
	0.1 **M** TBAP/DEF [93])	$Cr(acac)_3^{0/-}$	23	*0.83·10^{-2}*	PP/CV	95Urb
	0.1 **M** TBAP /HMPA [124])	$Cr(acac)_3^{0/-}$	23	*0.37·10^{-2}*	PP/CV	95Urb
	0.1 **M** TBAP /formamide	$Cr(acac)_3^{0/-}$	23	*0.89·10^{-2}*	PP/CV	95Urb
	0.1 **M** TBAP /NMF [88])	$Cr(acac)_3^{0/-}$	23	*0.59·10^{-2}*	PP/CV	95Urb

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	0.1 **M** acetate buffer +0.4 **M** NaCl/H_2O	$[CrEDTA]^{-/2-}$ [94])	25	*2.1·10^{-1}*	PS	69Tan1
	0.1 **M** acetate buffer +0.4 **M** NaCl/H_2O, pH=5	$[CrEDTA]^{-/2-}$ [94])	25	*2.1·10^{-1}*	PS	69Yam
	0.1 **M** acetate buffer +0.4 **M** $NaNO_3/H_2O$	$[CrEDTA]^{-/2-}$ [94])	25	*2.9·10^{-2}*	PS	69Tan1
	0.1 **M** TEAP [74]) /acetonitrile	$[CrEDTA]^{-/2-}$ [94])	25	*3.4·10^{-3}*	CV	95Hec1
	0.3 **M** Na_2SO_4/H_2O	$[CrEDTA]^{-/2-}$ [94])	22	*0.45*	IP	87Zha [95])
	0.1 **M** TEAP [74]) /propionitrile	$[CrEDTA]^{-/2-}$ [94])	25	*1.1·10^{-2}*	CV	95Hec1
	0.1 **M** TEAP [74]) /DMA [96])	$[CrEDTA]^{-/2-}$ [94])	25	*4.3·10^{-3}*	CV	95Hec1
	0.1 **M** TEAP [74]) /DMF [54])	$[CrEDTA]^{-/2-}$ [94])	25	*3.7·10^{-3}*	CV	95Hec1
	0.1 **M** TEAP [74]) /DMSO [51])	$[CrEDTA]^{-/2-}$ [94])	25	*3.2·10^{-4}*	CV	95Hec1
	0.1 **M** TEAP [74]) /ethylene glycol	$[CrEDTA]^{-/2-}$ [94])	25	*3.4·10^{-3}*	CV	95Hec1
	0.1 **M** TEAP [74]) /PC [113])	$[CrEDTA]^{-/2-}$ [94])	25	*3.7·10^{-3}*	CV	95Hec1
	0.1 **M** TEAP [74]) /acetonitrile	$[CrEDTA]^{-/2-}$ [94])	25	*3.4·10^{-3}*	CV	95Hec2
	0.1 **M** TEAP [74])/ propionitrile	$[CrEDTA]^{-/2-}$ [94])	25	*1.1·10^{-2}*	CV	95Hec2
	0.1 **M** TEAP [74]) /DMA [96])	$[CrEDTA]^{-/2-}$ [94])	25	*4.3·10^{-3}*	CV	95Hec2
	0.1 **M** TEAP [74]) /DMF [54])	$[CrEDTA]^{-/2-}$ [94])	25	*3.7·10^{-3}*	CV	95Hec2
	0.1 **M** TEAP [74]) /DMSO [51])	$[CrEDTA]^{-/2-}$ [94])	25	*3.2·10^{-4}*	CV	95Hec2
	0.1 **M** TEAP [74]) /ethylene glycol	$[CrEDTA]^{-/2-}$ [94])	25	*3.4·10^{-3}*	CV	95Hec2
	0.1 **M** TEAP [74]) /PC [113])	$[CrEDTA]^{-/2-}$ [94])	25	*3.1·10^{-3}*	CV	95Hec2
	0.02 **M** K-acetate buffer+0.08 **M** KCl +2 µ**M** LEO [98]) /H_2O, pH=4.5	$[CrCYDTA]^{-/2-}$ [97])	25	*8.5·10^{-3}*	PS, DCP	72Tan
	0.04 **M** K-acetate buffer+0.16 **M** KCl +2 µ**M** LEO [98]) /H_2O, pH=4.5	$[CrCYDTA]^{-/2-}$ [97])	25	*17.0·10^{-3}*	PS, DCP	72Tan

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.1 **M** K-acetate buffer+0.4 **M** KCl +2 µ**M** LEO [98]) /H_2O, pH=4.5	[CrCYDTA]$^{-/2-}$ [97])	25	*34.0·10^{-3}*	PS, DCP	72Tan
	0.2 **M** K-acetate buffer+0.8 **M** KCl +2 µ**M** LEO [98]) /H_2O, pH=4.5	[CrCYDTA]$^{-/2-}$ [97])	25	*58.0·10^{-3}*	PS, DCP	72Tan
	0.02 **M** Li-acetate buffer+0.08 **M** LiCl +2 µ**M** LEO [98]) /H_2O, pH=4.5	[CrCYDTA]$^{-/2-}$ [97])	25	*6.6·10^{-3}*	PS, DCP	72Tan
	0.04 **M** Li-acetate buffer+0.16 **M** LiCl +2 µ**M** LEO [98]) /H_2O, pH=4.5	[CrCYDTA]$^{-/2-}$ [97])	25	*9.6·10^{-3}*	PS, DCP	72Tan
	0.1 **M** Li-acetate buffer+0.4 **M** LiCl +2 µ**M** LEO [98]) /H_2O, pH=4.5	[CrCYDTA]$^{-/2-}$ [97])	25	*18.0·10^{-3}*	PS, DCP	72Tan
	0.2 **M** Li-acetate buffer+0.8 **M** LiCl +2 µ**M** LEO [98]) /H_2O, pH=4.5	[CrCYDTA]$^{-/2-}$ [97])	25	*25.0·10^{-3}*	PS, DCP	72Tan
	0.02 **M** Na-acetate buffer+0.08 **M** NaCl +2 µ**M** LEO [98]) /H_2O, pH=4.5	[CrCYDTA]$^{-/2-}$ [97])	25	*5.9·10^{-3}*	PS, DCP	72Tan
	0.04 **M** Na-acetate buffer+0.16 **M** NaCl +2 µ**M** LEO [98]) /H_2O, pH=4.5	[CrCYDTA]$^{-/2-}$ [97])	25	*11.0·10^{-3}*	PS, DCP	72Tan
	0.1 **M** Na-acetate buffer+0.4 **M** NaCl +2 µ**M** LEO [98]) /H_2O, pH=4.5	[CrCYDTA]$^{-/2-}$ [97])	25	*29.0·10^{-3}*	PS, DCP	72Tan
	0.2 **M** Na-acetate buffer+0.8 **M** NaCl +2 µ**M** LEO [98]) /H_2O, pH=4.5	[CrCYDTA]$^{-/2-}$ [97])	25	*37.0·10^{-3}*	PS, DCP	72Tan
	0.1 **M** acetate buffer +0.4 **M** NaCl +2 µ**M** LEO [98]) /H_2O, pH=5	[CrCYDTA]$^{-/2-}$ [97])	0	*4.5·10^{-3}*	PS, DCP	72Tan

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	0.1 **M** acetate buffer +0.4 **M** NaCl +2 µ**M** LEO [98]) /H_2O, pH=5	$[CrCYDTA]^{-/2-}$ [97])	10	*11.0·10^{-3}*	PS, DCP	72Tan
	0.1 **M** acetate buffer +0.4 **M** NaCl +2 µ**M** LEO [98]) /H_2O, pH=5	$[CrCYDTA]^{-/2-}$ [97])	20	*18.0·10^{-3}*	PS, DCP	72Tan
	0.1 **M** acetate buffer +0.4 **M** NaCl +2 µ**M** LEO [98]) /H_2O, pH=5	$[CrCYDTA]^{-/2-}$ [97])	25	*29.0·10^{-3}*	PS, DCP	72Tan
	0.1 **M** acetate buffer +0.4 **M** NaCl +2 µ**M** LEO [98]) /H_2O, pH=5	$[CrCYDTA]^{-/2-}$ [97])	35	*51.0·10^{-3}*	PS, DCP	72Tan
	0.1 **M** TEAP [74]) /acetonitrile	$[CrCYDTA]^{-/2-}$ [97])	25	*2.1·10^{-5}*	CV	95Hec1
	0.1 **M** TEAP [74]) /benzonitrile	$[CrCYDTA]^{-/2-}$ [97])	25	*2.7·10^{-5}*	CV	95Hec1
	0.1 **M** TEAP [74]) /propionitrile	$[CrCYDTA]^{-/2-}$ [97])	25	*8.5·10^{-2}*	CV	95Hec1
	0.1 **M** TEAP [74]) /DMA [96])	$[CrCYDTA]^{-/2-}$ [97])	25	*2.2·10^{-4}*	CV	95Hec1
	0.1 **M** TEAP [74]) /DMF [54])	$[CrCYDTA]^{-/2-}$ [97])	25	*7·10^{-3}*	CV	95Hec1
	0.1 **M** TEAP [74]) /DMSO [51])	$[CrCYDTA]^{-/2-}$ [97])	25	*3.9·10^{-6}*	CV	95Hec1
	0.1 **M** TEAP [74]) /PC [113])	$[CrCYDTA]^{-/2-}$ [97])	25	*1.6·10^{-3}*	CV	95Hec1
	0.1 **M** TEAP [74]) /DMA [96])	$[CrHEDTRA]^{0/-}$ [99])	25	*7.6·10^{-4}*	CV	95Hec1
	0.1 **M** TEAP [74]) /DMF [54])	$[CrHEDTRA]^{0/-}$	25	*8.5·10^{-3}*	CV	95Hec1
	0.1 **M** TEAP [74]) /DMSO [51])	$[CrHEDTRA]^{0/-}$	25	*3.1·10^{-3}*	CV	95Hec1
	0.1 **M** TEAP [74]) /HMPA [100])	$[CrHEDTRA]^{0/-}$	25	*5.1·10^{-3}*	CV	95Hec1
	0.1 **M** TEAP [74]) /propionitrile	$[CrHEDTRA]^{0/-}$	25	*4.0·10^{-2}*	CV	95Hec1
	0.1 **M** TEAP [74]) /tetramethylurea	$[CrHEDTRA]^{0/-}$	25	*1.1·10^{-2}*	CV	95Hec1
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Cr(bipy)_3^{2+/+}$	25	*>0.4*	GD	75Saj1
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Cr(bipy)_3^{+/0}$	25	*1.3*	GD	75Saj1

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Cr(bipy)_3^{0/-}$	25	*0.23*	GD	75Saj1
	1 **M** KCl/H_2O	$Cr_{aq}^{2+/3+}$	25	*1.0·10^{-5}*	IP	52Ran2
	0.2 **M** NaCN/H_2O	$Cr(CN)_6^{3-/4-}$, 2.5 m**M**	30	3.9·10^{-3}	IP	60Del
	1 **M** KCN/H_2O	$Cr(CN)_6^{3-/4-}$	25	*0.25*	IP	52Ran2
	1 **M** KCN/H_2O	$Cr(CN)_6^{3-/4-}$	25	*0.27*	ACP	72Hue
	1 **M** NMe_4OH/H_2O	Cs^{1+}	20	*~0.2*	IP	52Ran1
	0.1 **M** CsCl +0.01 **M** EDTA +0.01 **M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*1.3·10^{-3}*	DCP	71Bar4
	0.15 **M** CsCl +0.01 **M** EDTA +0.01 **M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*1.3·10^{-3}*	DCP	71Bar4
	0.2 **M** CsCl +0.01 **M** EDTA +0.01 **M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*1.1·10^{-3}*	DCP	71Bar4
	0.3 **M** CsCl +0.01 **M** EDTA+ 0.01 **M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*1.6·10^{-3}*	DCP	71Bar4
	0.5 **M** CsCl +0.01 **M** EDTA +0.01 **M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*1.6·10^{-3}*	DCP	71Bar4
	1 **M** CsCl +0.01 **M** EDTA +0.01 **M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*1.6·10^{-3}*	DCP	71Bar4
	0.1 **M** $HClO_4$/ +0.05% TritonX 100/H_2O	$Cu^{2+/0}$	25	*4 .4·10^{-9}*	DCP	59Kol
	0.1 **M** KCl+0.1 **M** EDTA/H_2O; saturated with camphor	$Cu^{2+/0}$, 0.046 **M** $CuSO_4$	25	*1.43·10^{-9}*	DCP	58Lai
	0.1 **M** $KClO_4/H_2O$	$Cu^{2+/0}$	25	*2.2·10^{-9}*	CF	69Suz
	1 **M** KNO_3/H_2O	$Cu^{2+/0}$	25	*4.5·10^{-2}*	ACP	61Kam
	1 **M** KNO_3/H_2O	$Cu^{2+/0}$, 2 m**M** Cu^{2+}	25	*1.9·10^{-2}*	PS	63Tos
	1 **M** KNO_3/H_2O	$Cu^{2+/0}$	25	*1.9·10^{-2}*	GS	63Tos

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	1 **M** KNO_3+1 **M** en /H_2O; saturated with camphor [63])	$Cu^{2+/0}$, 0.047 **M** $CuSO_4$	25	*1.12·10^{-5}*	DCP	58Lai
	1 **M** KNO_3+0.95 **M** propylene-diamine /H_2O; saturated with camphor [63])	$Cu^{2+/0}$, 0.046 **M** $CuSO_4$	25	*9.44·10^{-6}*	DCP	58Lai
	0.1 **M** LiCl +0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*1.0·10^{-3}*	DCP	71Bar4
	0.15 **M** LiCl +0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*1.0·10^{-3}*	DCP	71Bar4
	0.2 **M** LiCl +0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*1.3·10^{-3}*	DCP	71Bar4
	0.3 **M** LiCl +0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*0.9·10^{-3}*	DCP	71Bar4
	0.5 **M** LiCl +0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*1.3·10^{-3}*	DCP	71Bar4
	1 **M** LiCl +0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*1.4·10^{-3}*	DCP	71Bar4
	1 **M** $LiNO_3$/H_2O, pH=5	$Cu^{2+/0}$	25	*0.48*	ACP	72Haw
	0.5 **M** NaCl +0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*2.1·10^{-3}*	DCP	71Ruz
	NaCl+NaI +0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*1.0·10^{-2}*	DCP	71Bar4
	0.05 **M** NaI +0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M**$CuSO_4$	25	*1.1·10^{-6}*	DCP	71Bar4

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.05 **M** NaI +0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*7.7·10^{-7}*	CV	71Bar4
	0.1 **M** NaI +0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*6.9·10^{-7}*	DCP	71Bar4
	0.25 **M** NaI +0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*4.1·10^{-7}*	DCP	71Bar4
	0.25 **M** NaI +0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*1.9·10^{-7}*	CV	71Bar4
	0.3 **M** NaI +0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*3.1·10^{-7}*	DCP	71Bar4
	1 **M** NaI+0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*3.2·10^{-7}*	DCP	71Bar4
	1 **M** NaI+0.01 **M** EDTA +0.01**M** borate buffer, pH=8	$Cu^{2+/0}$, 0.5 m**M** $CuSO_4$	25	*1.2·10^{-7}*	CV	71Bar4
	1 **M** $NaNO_3$ +0.1 **M** EDTA/H_2O	$Cu^{2+/0}$, 2 m**M** $Cu(NO_3)_2$	25	*1.07·10^{-3}*	DCP	70Aya
	1 **M** $NaNO_3$+0.1 **M** EDTA+0.01% polyvinylalcohol /H_2O	$Cu^{2+/0}$, 2 m**M** $Cu(NO_3)_2$	25	*(2.5·10^{-11})*	DCP	70Aya
	0.5 **M** $NaSO_4$/+1 m**M** H_2SO_4+ n-butanol/H_2O	$Cu^{2+/0}$	25	*0.065*·(1-θ)$^{3.97}$ [101])	DCP	65Sat
	1 **M** NH_4OH +1 **M** NH_4NO_3 +0.71 **M** EDTA /H_2O; saturated with camphor	$Cu^{2+/0}$, 0.025 **M** $CuSO_4$	25	*8.75·10^{-3}*	DCP	58Lai
	1 **M** NH_4OH +1 **M** NH_4NO_3 /H_2O; saturated with camphor	$Cu^{2+/1+}$,0.035 **M** $CuCl_2$	25	*3.64·10^{-4}*	DCP	58Lai

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	1 **M** NH_4OH +1 **M** NH_4NO_3 /H_2O; saturated with camphor	$Cu^{2+/1+}$, 0.025 **M** $CuSO_4$	25	*2.28·10^{-6}*	DCP	58Lai
	1 **M** NH_4OH +1 **M** NH_4Cl/H_2O; saturated with camphor	$Cu^{2+/1+}$, 0.01 **M** $CuSO_4$	25	*1.03·10^{-6}*	DCP	58Lai
	1 **M** NH_4OH +1 **M** NH_4NO_3 /H_2O; saturated with camphor	$Cu^{1+/0}$, 0.026 **M** $CuCl_2$	25	*9.75·10^{-5}*	DCP	58Lai
	1 **M** KCl/H_2O	$Eu_{aq}^{2+/3+}$	25	*2.1·10^{-4}*	IP	52Ran2
	1 **M** KCl/H_2O	$Eu_{aq}^{2+/3+}$	25	*(0.32...0.87)·10^{-3}*	DCP	70deK
	1 **M** KI/H_2O	$Eu_{aq}^{2+/3+}$	25	*0.0016*	IP	52Ran2
	1 **M** $KCNS/H_2O$	$Eu_{aq}^{2+/3+}$	25	*0.008*	IP	52Ran2
	0.1 **M** KPF_6/H_2O	$Eu_{aq}^{2+/3+}$	25	*(3.0·10^{-3})*	CV	80Bar1
	0.4 **M** KPF_6/H_2O	$Eu_{aq}^{2+/3+}$	25	*(7.0·10^{-3})*	CV	80Bar1
	0.2 **M** $LiCl/H_2O$	$Eu_{aq}^{2+/3+}$	25	*4...9.9·10^{-3}*	OP	68Kov
	$NaClO_4/H_2O$	$Eu_{aq}^{2+/3+}$	25	*2.0·10^{-3}*	DCP	60Gie
	$NaClO_4/H_2O$	$Eu_{aq}^{3+/2+}$	25	*1.6·10^{-5}*	DCP	60Gie
	0.03...1 **M** $NaClO_4$ +1m**M** $HClO_4/H_2O$	$Eu_{aq}^{3+/2+}$	25	*2·10^{-5}*	DCP	61Gie
	1 **M** $NaClO_4/H_2O$	$Eu_{aq}^{3+/2+}$	25	*2.7·10^{-4}*	ACP	67Tim
	1 **M** $NaClO_4/H_2O$	$Eu_{aq}^{3+/2+}$	25	*3·10^{-5}*	DCP	70Kin
	1 **M** $KSCN/H_2O$	$Eu_{aq}^{3+/2+}$	25	*(2.5...3.3)·10^{-2}*	DCP	71deK
	0.5 **M** NaSCN+ 0.5 **M** $NaClO_4/H_2O$	$Eu_{aq}^{3+/2+}$	25	*(2.5...3.3)·10^{-2}*	CM	73Rod
	KCl (I=1 **M**)/H_2O	Fe^{2+}(r)	-	*(2.5·10^{-22})*	DCP	61deM
	KCl (I=1 **M**) +0.1 **M** thiourea/H_2O	Fe^{2+}(r)	-	*(1.3·10^{-20})*	DCP	61deM
	KCl (I=1 **M**)+0.2 **M** thiourea/H_2O	Fe^{2+}(r)	-	*(1.2·10^{-20})*	DCP	61deM
	KCl (I=1 **M**)+0.4 **M** thiourea/H_2O	Fe^{2+}(r)	-	*(1.3·10^{-20})*	DCP	61deM
	0.5 **M** K-oxalate/H_2O	$Fe(ox)_3^{3-/4-}$	25	*>1.0*	IP	52Ran2 [102])
	0.5 **M** K-oxalate/H_2O	$Fe(ox)_3^{3-/4-}$	25	*1.2*	ACP	72Hue

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	1 **M** K-oxalate/H_2O	$Fe(ox)_3^{3-/4-}$	25	*1.44*	GD	72Roh
	1 **M** K-oxalate +0.05 **M** oxalicacid/H_2O	$Fe(ox)^{2+/3+}$	25	*0.86*	FR	69deL [103])
	1 **M** K-oxalate +0.05 **M** oxalicacid/H_2O	$Fe(ox)^{2+/3+}$	-	*1.29*	FR	73Pol [102])
	0.6 **M** TEAP [74])/AN	Ferrocene	-	*1.1*	CV	88Wip
	0.1 **M** TBAP [52]) /acetone	$Fe(Cp')_2^{+/0}$ [90])	23	*3.5*	ACP	85Gen
	0.1 **M** TBAP /NMF [88])	$Fe(Cp')_2^{+/0}$	23	*0.9*	ACP	85Gen
	0.1 **M** TBAP /DMF [54])	$Fe(Cp')_2^{+/0}$	23	*2.0*	ACP	85Gen
	0.1 **M** TBAP /dimethyl sulfoxide	$Fe(Cp')_2^{+/0}$	23	*0.85*	ACP	85Gen
	0.1 **M** TBAP /benzonitrile	$Fe(Cp')_2^{+/0}$	23	*0.4*	ACP	85Gen
	0.1 **M** TEAP [74]) /acetonitrile	$(COT)Fe(CO)_3^{0/-}$ [104])	23	*2.0*	ACP	86McM
	0.1 **M** TBAP [52]) /acetonitrile	$(COT)Fe(CO)_3^{0/-}$	23	*1.1*	ACP	86McM
	0.1 **M** TEAP [74]) /NMF [88])	$(COT)Fe(CO)_3^{0/-}$	23	*1.1*	ACP	86McM
	0.1 **M** TBAP [52]) /NMF [88])	$(COT)Fe(CO)_3^{0/-}$	23	*0.19*	ACP	86McM
	0.1 **M** TEAP [74]) /DMF [54])	$(COT)Fe(CO)_3^{0/-}$	23	*0.7*	ACP	86McM
	0.1 **M** TBAP [52]) /DMF [54])	$(COT)Fe(CO)_3^{0/-}$	23	*0.16*	ACP	86McM
	0.1 **M** TEAP [74]) /tetramethylurea	$(COT)Fe(CO)_3^{0/-}$	23	*0.5*	ACP	86McM
	0.1 **M** TBAP [52]) /tetramethylurea	$(COT)Fe(CO)_3^{0/-}$	23	*0.07*	ACP	86McM
	0.1 **M** TEAP [74]) /dimethyl sulfoxide	$(COT)Fe(CO)_3^{0/-}$	23	*0.45*	ACP	86McM
	0.1 **M** TBAP [52]) /dimethyl sulfoxide	$(COT)Fe(CO)_3^{0/-}$	23	*0.07*	ACP	86McM
	0.1 **M** TEAP [74]) /PC [113])	$(COT)Fe(CO)_3^{0/-}$	23	*0.35*	ACP	86McM
	0.1 **M** TBAP [52]) /PC [113])	$(COT)Fe(CO)_3^{0/-}$	23	*0.1*	ACP	86McM

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	0.1 **M** TBAFB [89]) /ethanol	$(COT)Fe(CO)_3^{0/-}$	23	*3.0*	ACP	86McM
	0.1 **M** TBAFB /1-propanol	$(COT)Fe(CO)_3^{0/-}$	23	*2.0*	ACP	86McM
	0.1 **M** TBAP [52]) /acetone	$Fe(acac)_3^{0/-}$ [105])	22	*0.42*	IP	92Faw2
	0.1 **M** TBAP /dimethylacetamide	$Fe(acac)_3^{0/-}$	22	*0.324*	IP	92Faw2
	0.1 **M** TBAP /dimethyl sulfoxide	$Fe(acac)_3^{0/-}$	22	*0.175*	IP	92Faw2
	0.1 **M** TBAP [52]) /PC [113])	$Fe(acac)_3^{0/-}$	22	*0.154*	IP	92Faw2
	0.1 **M** TBAP [52]) /HMPA [124])	$Fe(acac)_3^{0/-}$	22	*0.031*	IP	92Faw2
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Fe(bipy)_3^{2+/+}$	25	*0.14*	GD	75Saj1
	6 **M** $NaClO_4$+0.05 **M** NaSCN/H_2O	Ga^{3+}/Ga, $1\cdot10^{-3}$	24	*0.016*	ACP	68Moo
	6 **M** $NaClO_4$+0.1 **M** NaSCN/H_2O	Ga^{3+}/Ga, $1\cdot10^{-3}$	30	*0.026*	ACP	68Moo
	0.1 **M** HCl/H_2O +methanol	H^+/H_2	19	$10^{-13}...10^{-12}$	CP	49Boc [106])
	1 **M** HCl/H_2O	H^+/H_2	0	$10^{-13.65}$	IP	69Dek
	1 **M** HCl/H_2O	H^+/H_2	15	$10^{-13.08}$	IP	69Dek
	1 **M** HCl/H_2O	H^+/H_2	25	$10^{-11.98}$	IP	69Dek
	1 **M** HCl/H_2O	H^+/H_2	45	$10^{-11.28}$	IP	69Dek
	1 **M** HCl/H_2O	H^+/H_2	0	$10^{-13.65}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	–35	$10^{-11.66}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	–25	$10^{-11.23}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	–15	$10^{-10.70}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	–5	$10^{-10.7}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	5	$10^{-10.38}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	10	$10^{-10.0}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	15	$10^{-9.88}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	20	$10^{-9.77}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	25	$10^{-9.8}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	35	$10^{-9.33}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	45	$10^{-9.16}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	50	$10^{-8.85}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	55	$10^{-8.88}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	65	$10^{-8.52}$	IP	69Dek
	7.5 **M** HCl/H_2O	H^+/H_2	72	$10^{-8.4}$	IP	69Dek
	0.1 **M** $HClO_4/H_2O$	H^+/H_2	25	$2.2\cdot10^{-13}$	CP	64But1
	0.1 **M** $HClO_4/H_2O$	H^+/H_2	25	$5.88\cdot10^{-14}$	CP	66But3
	5.2 **M** $HClO_4/H_2O$	H^+/H_2	–39	$10^{-16.52}$	IP	69Dek

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	5.2 **M** $HClO_4/H_2O$	H^+/H_2	–35	$10^{-15.56}$	IP	69Dek
	5.2 **M** $HClO_4/H_2O$	H^+/H_2	–25	$10^{-15.64}$	IP	69Dek
	5.2 **M** $HClO_4/H_2O$	H^+/H_2	–5	$10^{-14.26}$	IP	69Dek
	5.2 **M** $HClO_4/H_2O$	H^+/H_2	0	$10^{-13.07}$	IP	69Dek
	5.2 **M** $HClO_4/H_2O$	H^+/H_2	5	$10^{-13.03}$	IP	69Dek
	5.2 **M** $HClO_4/H_2O$	H^+/H_2	10	$10^{-13.58}$	IP	69Dek
	5.2 **M** $HClO_4/H_2O$	H^+/H_2	15	$10^{-13.82}$	IP	69Dek
	5.2 **M** $HClO_4/H_2O$	H^+/H_2	25	$10^{-12.17}$	IP	69Dek
	0.05 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-12.6}$	CP	65Mat
	0.1 **M** $HClO_4/H_2O$	$Hg^{+/0}$	24	0.28	FR	62Ima2
	0.2 **M** $HClO_4/H_2O$	$Hg^{+/0}$	24	0.35	FR	62Ima2
	1.1 **M** $HClO_4/H_2O$	$Hg^{+/0}$	24	1.3	FR	62Ima2
	0.98 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 1 m**M**	25	0.248	GD	59Mat3
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 1 m**M**	0	*>1.5*	PS	56Ger
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 1 m**M**	0	*14*	GD	58Ger
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 1 m**M**	25	*16.5*	GD	58Ger
	1 **M** $HClO_4/H_2O$ [107])	Hg_2^{2+}/Hg	–38.8	10	GD	58Ger
	1 **M** $HClO_4/H_2O$ [108])	Hg_2^{2+}/Hg	–38.9	8.5	GD	58Ger
	1 **M** $HClO_4/H_2O$	Hg/Hg_2^{2+} , 1 m**M**	-	>0.46	IP	64Slu
	1 **M** $HClO_4/H_2O$	Hg/Hg_2^{2+} , 0.88 m**M**	25	$0.6<j_0<2.2$	GM	67Koo2
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 0.0206 m**M**	25	0.0212	GS	67Wei2 [109])
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 0.103 m**M**	25	0.0678	GS	67Wei2 [109])
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 0.206 m**M**	25	0.103	GS	67Wei2 [109])
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 0.515 m**M**	25	0.190	GS	67Wei2 [109])
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 1.03 m**M**	25	0.304	GS	67Wei2 [109])
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 2.06 m**M**	25	0.662	GS	67Wei2 [109])
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 5.15 m**M**	25	2.74	GS	67Wei2 [109])
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 0.0206 m**M**	25	0.0197	GS	67Wei2 [110])
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 0.103 m**M**	25	0.0717	GS	67Wei2 [110])
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 0.206 m**M**	25	0.104	GS	67Wei2 [110])
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 0.515 m**M**	25	0.182	GS	67Wei2[110])
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 1.03 m**M**	25	0.274	GS	67Wei2[110])
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 2.06 m**M**	25	0.447	GS	67Wei2[110])
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 5.15 m**M**	25	1.21	GS	67Wei2[110])

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 0.25 m**M**	25	0.040	FR,GS, IP,CS	70deL
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 0.5 m**M**	25	0.075	FR,GS, IP,CS	70deL
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 1 m**M**	25	0.140	FR,GS, IP,CS	70deL
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 2 m**M**	25	0.265	FR,GS, IP,CS	70deL
	1 **M** $HClO_4/H_2O$	Hg_2^{2+}/Hg , 4 m**M**	25	0.500	FR,GS, IP,CS	70deL
	0.5 **M** ClO_4^-/H_2O	$In^{3+/0}$	20	*(2.2·10^{-12})*	DCP	60Ino
	0.5 **M** ClO_4^-/H_2O	$In^{3+/0}$	26.5	*(4.7·10^{-12})*	DCP	60Ino
	0.5 **M** ClO_4^-/H_2O	$In^{3+/0}$	30	*(7.3·10^{-12})*	DCP	60Ino
	1 **M** KCl/H_2O (pH=3 with HNO_3)	$In^{3+/0}$	25	*6.0·10^{-4}*	IP	68Tim
	1 **M** KNO_3+0.012 **M** KCl/H_2O	$In^{3+/0}$	25	*3.1·10^{-3}*	IP	68Tim
	0.1 **M** $LiCl/H_2O$ [111])	$In^{3+/0}$	30	*3.147·10^{-4}*	DCP	67Jai
	0.1 **M** $LiCl/H_2O$ [111])	$In^{3+/0}$	40	*3.147·10^{-4}*	DCP	67Jai
	0.5 **M** $LiCl/H_2O$ [111])	$In^{3+/0}$	30	*4.414·10^{-4}*	DCP	67Jai
	0.5 **M** $LiCl/H_2O$ [111])	$In^{3+/0}$	40	*6.0·10^{-4}*	DCP	67Jai
	1 **M** $LiCl/H_2O$ [111])	$In^{3+/0}$	30	*6.65·10^{-4}*	DCP	67Jai
	1 **M** $LiCl/H_2O$ [111])	$In^{3+/0}$	40	*8.683·10^{-4}*	DCP	67Jai
	1 **M** $NaBr/H_2O$	$In^{3+/0}$	25	*0.65*	IP	70Mou
	1 **M** $NaCl/H_2O$	$In^{3+/0}$	25	*0.034*	IP	70Mou
	0.1 **M** $NaClO_4/H_2O$	$In^{3+/0}$	30	*1.282·10^{-3}*	DCP	66Gau
	0.1 **M** $NaClO_4/H_2O$	$In^{3+/0}$	40	*1.423·10^{-3}*	DCP	66Gau
	0.48 **M** $NaClO_4/H_2O$	$In^{3+/0}$	25	*(4.4·10^{-9})*	DCP	58Ima
	0.5 **M** $NaClO_4/H_2O$	$In^{3+/0}$	30	*1.251·10^{-3}*	DCP	66Gau
	0.5 **M** $NaClO_4$ +0.04 **M** acetyl-acetonate/H_2O, pH=1.35	$In^{3+/0}$	25	*1.4·10^{-3}*	DCP	66Cos
	0.5 **M** $NaClO_4$ +0.04 **M** acetyl-acetonate/H_2O, pH=1.58	$In^{3+/0}$	25	*1.57·10^{-3}*	DCP	66Cos
	0.5 **M** $NaClO_4$ +0.04 **M** acetyl-acetonate/H_2O, pH=1.82	$In^{3+/0}$	25	*1.27·10^{-3}*	DCP	66Cos
	0.5 **M** $NaClO_4$ +0.04 **M** acetyl-acetonate/H_2O, pH=2.59	$In^{3+/0}$	25	*5.42·10^{-3}*	DCP	66Cos

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.5 **M** $NaClO_4$ +0.04 **M** acetyl-acetonate/H_2O, pH=4.45	$In^{3+/0}$	25	*2.92·10^{-3}*	DCP	66Cos
	0.5 **M** $NaClO_4$ +0.04 **M** acetyl-acetonate/H_2O, pH=5.11	$In^{3+/0}$	25	*3.32·10^{-3}*	DCP	66Cos
	0.5 **M** $NaClO_4$ +0.04 **M** acetyl-acetonate/H_2O, pH=5.45	$In^{3+/0}$	25	*2.92·10^{-3}*	DCP	66Cos
	0.5 **M** $NaClO_4$ +0.04 **M** acetyl-acetonate /H_2O, pH=5.78	$In^{3+/0}$	25	*5.32·10^{-3}*	DCP	66Cos
	0.5 **M** $NaClO_4$ +0.04 **M** acetyl-acetonate/H_2O, pH=6.7	$In^{3+/0}$	25	*5.32·10^{-3}*	DCP	66Cos
	0.5 **M** $NaClO_4$ +0.04 **M** acetyl-acetonate/H_2O, pH=7.25	$In^{3+/0}$	25	*11.3·10^{-3}*	DCP	66Cos
	1 **M** $NaClO_4/H_2O$	$In^{3+/0}$	30	*1.213·10^{-3}*	DCP	66Gau
	0.3 **M** NaF +0.1 **M** HCl/H_2O	$In^{0/3+}$, 4·10^{-6} **M**	20	1.55·10^{-5}	CP	67Los
	1 **M** NaN_3/H_2O	$In^{3+/0}$	25	*0.05*	IP	70Mou
	0.1 **M** $NaNO_3/H_2O$	$In^{3+/0}$	30	*1.285·10^{-3}*	DCP	66Gau
	0.1 **M** $NaNO_3/H_2O$	$In^{3+/0}$	40	*1.89·10^{-3}*	DCP	66Gau
	0.5 **M** $NaNO_3/H_2O$	$In^{3+/0}$	30	*1.243·10^{-3}*	DCP	66Gau
	0.5 **M** $NaNO_3/H_2O$	$In^{3+/0}$	40	*1.875·10^{-3}*	DCP	66Gau
	1 **M** $NaNO_3/H_2O$	$In^{3+/0}$	30	*1.232·10^{-3}*	DCP	66Gau
	1 **M** KCl/H_2O	$K^{+/0}$	25	*(0.7)*	IP	66Slu
	1 **M** $LiCl/H_2O$	$Li^{+/0}$	25	*(0.14)*	IP,DC, P	72Ree
	1 **M** KCl/H_2O	Mn^{2+}(r)	25	*4·10^{-3}*	OP	55Mat
	1 **M** H_2SO_4+0.021 % sodiumdodecyl-sulfonate/H_2O	$Mn^{3+/2+}$	-	*(8.6·10^{-3})*	DCP	72Iss3
	1 **M** H_2SO_4+0.019 % TritonX-100/H_2O	$Mn^{3+/2+}$	-	*(5.1·10^{-3})*	DCP	72Iss3
	1 **M** H_2SO_4+0.059 % TritonX-100/H_2O	$Mn^{3+/2+}$	-	*(2.3·10^{-3})*	DCP	72Iss3
	1 **M** H_2SO_4+0.19 % TritonX-100/H_2O	$Mn^{3+/2+}$	-	*(1.6·10^{-3})*	DCP	72Iss3

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	1 **M** H_3PO_4+0.042 % sodiumdodecyl sulfonate/H_2O	$Mn^{3+/2+}$	-	*(2.0·10^{-2})*	DCP	72Iss3
	1 **M** H_3PO_4+0.145 % sodiumdodecyl-sulfonate/H_2O	$Mn^{3+/2+}$	-	*(1.4·10^{-2})*	DCP	72Iss3
	1 **M** H_3PO_4+0.01 % TritonX-100/H_2O	$Mn^{3+/2+}$	-	*(2.8·10^{-3})*	DCP	72Iss3
	1 **M** H_3PO_4+0.0136 % TritonX-100/H_2O	$Mn^{3+/2+}$	-	*(4.8·10^{-3})*	DCP	72Iss3
	1 **M** H_3PO_4+0.19 % TritonX-100/H_2O	$Mn^{3+/2+}$	-	*(6.6·10^{-3})*	DCP	72Iss3
	1 **M** H_3PO_4+0.536 % TritonX-100/H_2O	$Mn^{3+/2+}$	-	*(9.4·10^{-3})*	DCP	72Iss3
	0.08 **M** NaF+0.0046 % TritonX-100/H_2O (pH=4.5 with H_2SO_4)	MnF_4^-/Mn^{2+}	-	*(3.82·10^{-5})*	DCP	71Iss
	0.08 **M** NaF+0.018 % TritonX-100/H_2O (pH=4.5 with H_2SO_4)	MnF_4^-/Mn^{2+}	-	*(4.60·10^{-6})*	DCP	71Iss
	0.08 **M** NaF+0.052 % TritonX-100/H_2O (pH=4.5 with H_2SO_4)	MnF_4^-/Mn^{2+}	-	*(3.65·10^{-6})*	DCP	71Iss
	0.08 **M** NaF+0.0036 % TritonX-100/H_2O (pH=4.5 with H_2SO_4)	MnF_4^-/Mn^{2+}	-	*(2.01·10^{-3})*	DCP	71Iss
	0.08 **M** NaF+0.0108 % TritonX-100/H_2O (pH=4.5 with H_2SO_4)	MnF_4^-/Mn^{2+}	-	*(7.46·10^{-4})*	DCP	71Iss
	0.08 **M** NaF+0.068 % TritonX-100/H_2O (pH=4.5 with H_2SO_4)	MnF_4^-/Mn^{2+}	-	*(3.8·10^{-4})*	DCP	71Iss
	0.1 **M** TBAP [52]) /acetonitrile	$Mn(Cp')_2^{+/0}$ [90])	23	*4.0*	ACP	85Gen
	0.1 **M** TBAP/acetone	$Mn(Cp')_2^{+/0}$	23	*4.0*	ACP	85Gen
	0.1 **M** TBAP/DMF[54])	$Mn(Cp')_2^{+/0}$	23	*2.0*	ACP	85Gen
	0.1 **M** TBAP /dimethyl sulfoxide	$Mn(Cp')_2^{+/0}$	23	*0.9*	ACP	85Gen

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.1 **M** TBAP /acetonitrile	$Mn(acac)_3^{0/-}$ [105])	22	*0.448*	IP	92Faw2
	0.1 **M** TBAP/acetone	$Mn(acac)_3^{0/-}$	22	*0.189*	IP	92Faw2
	0.1 **M** TBAP /butyronitrile	$Mn(acac)_3^{0/-}$ [105])	22	*0.132*	IP	92Faw2
	0.1 **M** TBAP/PC [113])	$Mn(acac)_3^{0/-}$	22	*0.074*	IP	92Faw2
	0.1 **M** TBAP /acetonitrile	Mn(TPP)Cl (r) [112])	25	*0.18*	CV	93Mu
	0.1 **M** TBAP /propionitrile	Mn(TPP)Cl (r)	25	*0.068*	CV	93Mu
	0.1 **M** TBAP /butyronitrile	Mn(TPP)Cl (r)	25	*0.028*	CV	93Mu
	0.1 **M** TBAP /nitromethane	Mn(TPP)Cl (r)	25	*0.116*	CV	93Mu
	0.1 **M** TBAP /nitrobenzene	Mn(TPP)Cl (r)	25	*0.018*	CV	93Mu
	0.1 **M** TBAP/PC [113])	Mn(TPP)Cl (r)	25	*0.012*	CV	93Mu
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Mo(bipy)_3^{+/0}$	25	*1.3*	GD	75Saj1
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Mo(bipy)_3^{0/-}$	25	*0.22*	GD	75Saj1
	KCl/H_2O, I=2 **M**	$Ni^{2+/0}$	-	*(6.3·10^{-18})*	DCP	61deM
	0.1 **M** KCl/H_2O	$Ni^{2+/0}$	25	*(1.2·10^{-2})*	DCP	70Nis [114])
	0.1 **M** KCl+0.025 **M** acetyl-acetonate +0.025 **M** acetate buffer+0.005 carboxymethyl cellulose/H_2O, pH=3.5...4.5	$Ni^{2+/0}$	25	*(2.9·10^{-10})*	DCP	70Nis [115])
	0.1 **M** KCl+0.025 **M** acetyl-acetonate +0.025 **M** acetate buffer+0.005 carboxymethyl cellulose/H_2O, pH=3.5...4.5	$Ni^{2+/0}$	25	*(1.5·10^{-7})*	DCP	70Nis [116])
	0.1 **M** KCl+0.025 **M** acetyl-acetonate +0.025 **M** acetate buffer+0.005 carboxymethyl cellulose/H_2O, pH=3.5...4.5	$Ni^{2+/0}$	25	*(5.7·10^{-3})*	DCP	70Nis [114])

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	0.1 **M** KCl+0.025 **M** acetyl-acetonate +0.025 **M** phospate or borate buffer +0.005 carboxymethyl cellulose/H_2O, pH=7...8.5	$Ni^{2+/0}$	25	*(8.7·10^{-15})*	DCP	70Nis
	KCl+0.1 **M** thiourea /H_2O, *I*=2 **M**	$Ni^{2+/0}$	-	*(3.2·10^{-12})*	DCP	61deM
	KCl+0.2 **M** thiourea /H_2O, *I*=2 **M**	$Ni^{2+/0}$	-	*(3.1·10^{-12})*	DCP	61deM
	KCl+0.4 **M** thiourea /H_2O, *I*=2 **M**	$Ni^{2+/0}$	-	*(3.0·10^{-12})*	DCP	61deM
	KCl+0.7 **M** thiourea /H_2O, *I*=2 **M**	$Ni^{2+/0}$	-	*(3.2·10^{-12})*	DCP	61deM
	0.1 **M** KNO_3/H_2O	$Ni^{2+/0}$	25	*(2.0·10^{-14})*	DCP	56Mor
	0.1 **M** KNO_3 +1 **M** NH_3/H_2O	$Ni^{2+/0}$	25	*(1.6·10^{-17})*	DCP	56Mor
	0.1 **M** KNO_3 +1 **M** en/H_2O [63]) +0.01 % gelatin	$Ni^{2+/0}$	25	*(2.5·10^{-18})*	DCP	56Mor
	0.1 **M** KNO_3 /H_2O+0.01 % gelatin	$Ni^{2+/0}$	25	*(2.0·10^{-14})*	DCP	55Mor
	0.1 **M** KNO_3 /H_2O+0.01 % gelatin	$Ni^{2+/0}$	25	*(2.0·10^{-14})*	DCP	56Mat
	0.1 **M** KNO_3 +0.2...3 **M** HN_3 /H_2O+0.01 % gelatin	$Ni^{2+/0}$	25	*(1.6·10^{-17})*	DCP	55Mor
	0.1 **M** KNO_3+0.2...3 **M** HN_3/H_2O +0.01 % gelatin	$Ni^{2+/0}$	25	*(1.6·10^{-17})*	DCP	56Mat
	0.2 **M** KSCN/ H_2O	$Ni^{2+/0}$	25	*2.9·10^{-4}*	DCP	67Gal
	0.1 **M** KSCN/ H_2O	$Ni^{2+/0}$	25	*2.5·10^{-4}*	CV	67Gal
	0.2 **M** KSCN/ H_2O	$Ni^{2+/0}$	25	*1.5·10^{-4}*	CV	67Gal
	0.6 **M** KSCN/ H_2O	$Ni^{2+/0}$	25	*3.2·10^{-4}*	CV	67Gal
	0.6 **M** KSCN/ H_2O	$Ni^{2+/0}$	25	*3.4·10^{-4}*	DCP	67Gal
	1.8 **M** KSCN/ H_2O	$Ni^{2+/0}$	25	*3.8·10^{-4}*	CV	67Gal
	1.8 **M** KSCN/ H_2O	$Ni^{2+/0}$	25	*3.1·10^{-4}*	DCP	67Gal
	5.5 **M** KSCN/ H_2O	$Ni^{2+/0}$	25	*4.2·10^{-4}*	CV	67Gal
	5.5 **M** KSCN/ H_2O	$Ni^{2+/0}$	25	*2.5·10^{-4}*	DCP	67Gal
	0.12 **M** NaBr/H_2O +0.005 % methyl cellulose	$Ni^{2+/0}$	25	*(1.0·10^{-13})*	DCP	58Ima

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.18 **M** $NaBr/H_2O$ +0.005 % methyl cellulose	$Ni^{2+/0}$	25	*(4.0·10^{-13})*	DCP	58Ima
	0.24 **M** $NaBr/H_2O$ +0.005 % methyl cellulose	$Ni^{2+/0}$	25	*(2.7·10^{-12})*	DCP	58Ima
	0.18 **M** $NaClO_4$ +0.02 **M** $KSCN/H_2O$	$Ni^{2+/0}$	25	*0.25·10^{-4}*	DCP	67Gal
	0.16 **M** $NaClO_4$ +0.04 **M** $KSCN/H_2O$	$Ni^{2+/0}$	25	*0.63·10^{-4}*	DCP	67Gal
	0.13 **M** $NaClO_4$ +0.07 **M** $KSCN/H_2O$	$Ni^{2+/0}$	25	*2.3·10^{-4}*	DCP	67Gal
	0.18 **M** NaI/H_2O, +0.005 % methyl cellulose	$Ni^{2+/0}$	25	*(7.0·10^{-13})*	DCP	58Ima
	0.24 **M** NaI/H_2O, +0.005 % methyl cellulose	$Ni^{2+/0}$	25	*(8.0·10^{-12})*	DCP	58Ima
	0.12 **M** $NaCl/H_2O$, +0.005 % methyl cellulose	$Ni^{2+/0}$	25	*(1.0·10^{-13})*	DCP	58Ima
	0.18 **M** $NaCl/H_2O$, +0.005 % methyl cellulose	$Ni^{2+/0}$	25	*(5.0·10^{-13})*	DCP	58Ima
	0.24 **M** $NaCl/H_2O$, +0.005 % methyl cellulose	$Ni^{2+/0}$	25	*(2.7·10^{-12})*	DCP	58Ima
	0.1 **M** $NaClO_4/H_2O$	$Ni^{2+/0}$	25	*(1.22·10^{-9})*	DCP	58Lai
	0.1 **M** $NaClO_4/H_2O$, saturated with β-naphthol	$Ni^{2+/0}$	25	*(7.17·10^{-12})*	DCP	58Lai
	0.1 **M** $NaClO_4/H_2O$, saturated with camphor	$Ni^{2+/0}$	25	*(1.09·10^{-16})*	DCP	58Lai
	0.5 **M** $NaClO_4/H_2O$	$Ni^{2+/0}$	25	*(5.14·10^{-9})*	DCP	58Lai
	0.5 **M** $NaClO_4/H_2O$ +2 m**M** *n*-octyl alcohol	$Ni^{2+/0}$	25	*(6.06·10^{-16})*	DCP	58Lai
	1 **M** $NaClO_4/H_2O$	$Ni^{2+/0}$	25	*(8.05·10^{-9})*	DCP	58Lai
	2 **M** $NaClO_4/H_2O$	$Ni^{2+/0}$	25	*(3.53·10^{-8})*	DCP	58Lai
	3 **M** $NaClO_4/H_2O$	$Ni^{2+/0}$	25	*(1.69·10^{-9})*	DCP	58Lai
	0.1 **M** $NaClO_4/H_2O$ +0.01 % polyacrylamide	$Ni^{2+/0}$	25	*3.3·10^{-9}*	DCP	60Oki

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	0.1 **M** TBAP /acetonitrile	$[Ni(cp)_2]^{0/+}$	22	*0.71*	ACP	96Win
	0.1 **M** TBAP [52]) /acetone	$[Ni(cp)_2]^{0/+}$	22	*0.55*	ACP	96Win
	0.1 **M** TBAP [52]) /dichloromethane	$[Ni(cp)_2]^{0/+}$	22	*0.4*	ACP	96Win
	0.1 **M** TBAP [52]) /dichloroethane	$[Ni(cp)_2]^{0/+}$	22	*0.25*	ACP	96Win
	0.1 **M** TBAP [52]) /DMF [54])	$[Ni(cp)_2]^{0/+}$	22	*0.14*	ACP	96Win
	0.1 **M** TBAP [52]) /dimethyl sulfoxide	$[Ni(cp)_2]^{0/+}$	22	*0.1*	ACP	96Win
	0.1 **M** TBAP [52]) /PC [113])	$[Ni(cp)_2]^{0/+}$	22	*0.055*	ACP	96Win
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Os(bipy)_3^{2+/+}$	25	*0.15*	GD	75Saj1
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Os(bipy)_3^{+/0}$	25	*0.25*	GD	75Saj1
	1 **M** $HClO_4/H_2O$	$Pb^{2+/0}$	25	*0.9*	IP	65Bie
	1 **M** KCl/H_2O, pH=2 with $HClO_4$	$Pb^{2+/0}$	22	*0.2*	FR	58Bar1
	1 **M** KCl/H_2O, pH=2	$Pb^{2+/0}$, $2\cdot10^{-5}$ **M**	22	*(>0.2)*	FR	61Bar1
	1 **M** KF/H_2O, saturated with PbF_2	$Pb^{2+/0}$, $2\cdot10^{-5}$ **M**	22	*(>5)*	FR	61Bar1
	1 **M** KNO_3/H_2O	$Pb^{2+/0}$	25	*>1*	ACP	61Kam
	1 **M** KNO_3/H_2O	$Pb^{2+/0}$	25	*>1*	ACP	61Kam
	1 **M** KBr/H_2O	$Pb^{2+/0}$, $1.5\cdot10^{-3}$ **M**	27	*0.96*	FR	71Aga
	1 **M** KBr/H_2O, pH=2	$Pb^{2+/0}$, $2\cdot10^{-5}$ **M**	22	*(>0.2)*	IP	61Bar1
	1 **M** KCl/H_2O	$Pb^{2+/0}$, $1.2\cdot10^{-3}$ **M**	32	*5.2*	FR	71Aga
	1 **M** KCl/H_2O	$Pb^{2+/0}$, $2.4\cdot10^{-3}$ **M**	32	*7.7*	FR	71Aga
	1 **M** KI/H_2O, pH=2	$Pb^{2+/0}$, $2\cdot10^{-5}$ **M**	22	*(>0.2)*	FR	61Bar1
	1 **M** KNO_3/H_2O	$Pb^{2+/0}$	25	*>1*	ACP	61Kam
	1 **M** KNO_3/H_2O, pH=2 with $HClO_4$	$Pb^{2+/0}$	22	*0.8*	FR	58Bar1
	1 **M** KNO_3/H_2O, pH=2	$Pb^{2+/0}$	22	*(>0.8)*	FR	61Bar1
	1 **M** KNO_3/H_2O	$Pb^{2+/0}$	25	*>1*	ACP	61Kam
	1 **M** KNO_3/H_2O	$Pb^{2+/0}$, $1.5\cdot10^{-3}$	28	*0.59*	FR	71Aga
	1 **M** KNO_3/H_2O	$Pb^{2+/0}$, $3\cdot10^{-3}$	28	*1.0*	FR	71Aga
	1 **M** $NaBr/H_2O$	$Pb^{2+/0}$	25	*0.48*	IP	70Mou
	1 **M** $NaCl/H_2O$	$Pb^{2+/0}$	25	*0.17*	IP	70Mou
	0.9 **M** $NaClO_4$+0.1 **M** $NaOH/H_2O$	$Pb^{2+/0}$	25	*0.06*	IP	70Mou
	1 **M** $NaClO_4/H_2O$	$Pb^{2+/0}$, $1.2\cdot10^{-3}$	37	*10.2*	FR	71Aga
	1 **M** $NaClO_4/H_2O$	$Pb^{2+/0}$, $2.4\cdot10^{-3}$	37	*13.5*	FR	71Aga

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	1 **M** $NaClO_4/H_2O$, pH=2 with $HClO_4$	$Pb^{2+/0}$	22	*3.35*	FR	58Bar1
	1 **M** $NaClO_4/H_2O$, pH=2	$Pb^{2+/0}$	22	*(3.3)*	FR	61Bar1
	1 **M** NaN_3/H_2O	$Pb^{2+/0}$	25	*0.23*	IP	70Mou
	$Na_6P_6O_{18}/H_2O$ [117])	$Pb^{2+/0}$	27	*$2.0\cdot10^{-9}$*	DCP	61Pam
	$Na_6P_6O_{18}/H_2O$	$Pb^{2+/0}$	32	*$1.3\cdot10^{-8}$*	DCP	61Pam
	$Na_6P_6O_{18}/H_2O$	$Pb^{2+/0}$	42	*$1.7\cdot10^{-8}$*	DCP	61Pam
	$Na_6P_6O_{18}/H_2O$	$Pb^{2+/0}$	52	*$2.2\cdot10^{-7}$*	DCP	61Pam
	1 **M** Na_2CO_3/H_2O	Pu(IV)/Pu(III),(r)	25	*$10^{-4.3}$*	DCP	70Cas
	1 **M** Na_2CO_3/H_2O	Pu(IV)/Pu(III),(r)	25	*$10^{-4.6}$*	OC	70Cas
	0.2 **M** KPF_6/H_2O	$Ru_{aq}^{2+/3+}$	25	*$(5.0\cdot10^{-2})$*	CV	80Bar1
	0.1 **M** CF_3COONa/H_2O	$Ru(NH_3)_6^{3+/2+}$	-	*0.45*	CV	88Wip
	1 **M** KF/H_2O	$[Ru(NH_3)_6]^{3+/2+}$	20	50...79	TPF	85Iwa
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Ru(bipy)_3^{2+/+}$	25	*0.1*	GD	75Saj1
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Ru(bipy)_3^{+/0}$	25	*0.25*	GD	75Saj1
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Ru(bipy)_3^{0/-}$	25	*0.22*	GD	75Saj1
	0.5 **M** H_2SO_4/H_2O +0.001 % gelatin	$Sb^{3+/0}$	25	$8.7\cdot10^{-5}$	DCP	63Fuz
	0.1 **M** $NaOH/H_2O$	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*$(5.0\cdot10^{-15})$*	DCP	71Sad
	0.1 **M** $NaOH/H_2O$ +10 vol% methanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*$(1.2\cdot10^{-14})$*	DCP	71Sad
	0.1 **M** $NaOH/H_2O$ +20 vol% methanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*$(6.9\cdot10^{-17})$*	DCP	71Sad
	0.1 **M** $NaOH/H_2O$ +30 vol% methanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*$(6.3\cdot10^{-21})$*	DCP	71Sad
	0.1 **M** $NaOH/H_2O$ +40 vol% methanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*$(1.25\cdot10^{-21})$*	DCP	71Sad
	0.1 **M** $NaOH/H_2O$ +50 vol% methanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*$(7.9\cdot10^{-24})$*	DCP	71Sad
	0.1 **M** $NaOH/H_2O$ +10 vol% ethanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*$(6.2\cdot10^{-17})$*	DCP	71Sad
	0.1 **M** $NaOH/H_2O$ +20 vol% ethanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*$(3.2\cdot10^{-18})$*	DCP	71Sad
	0.1 **M** $NaOH/H_2O$ +30 vol% ethanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*$(1.15\cdot10^{-20})$*	DCP	71Sad

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	0.1 **M** NaOH/H_2O +40 vol% ethanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*(3.16·10^{-23})*	DCP	71Sad
	0.1 **M** NaOH/H_2O +50 vol% ethanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*(1.8·10^{-25})*	DCP	71Sad
	0.1 **M** NaOH/H_2O +10 vol% *n*-propanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*(1.6·10^{-17})*	DCP	71Sad
	0.1 **M** NaOH/H_2O +20 vol% *n*-propanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*(4.0·10^{-22})*	DCP	71Sad
	0.1 **M** NaOH/H_2O +30 vol% *n*-propanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*(2.0·10^{-24})*	DCP	71Sad
	0.1 **M** NaOH/H_2O +40 vol% *n*-propanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*(8.7·10^{-29})*	DCP	71Sad
	0.1 **M** NaOH/H_2O +50 vol% *n*-propanol	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*(4.5·10^{-31})*	DCP	71Sad
	0.1 **M** NaOH/H_2O +10 vol% dioxane	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*(5.5·10^{-17})*	DCP	71Sad
	0.1 **M** NaOH/H_2O +20 vol% dioxane	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*(7.2·10^{-21})*	DCP	71Sad
	0.1 **M** NaOH/H_2O +30 vol% dioxane	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*(5.0·10^{-22})*	DCP	71Sad
	0.1 **M** NaOH/H_2O +40 vol% dioxane	TeO_6^{6-}, 0.5 m**M** H_6TeO_6	-	*(8.3·10^{-26})*	DCP	71Sad
	0.23 **M** HCl/H_2O +0.005 wt% gelatin	$Ti^{4+/3+}$	16-17	$7.8·10^{-4}$	DCP	60Str [118])
	0.8 **M** NaCl/H_2O +0.002 wt% Triton X-30S	$Ti^{4+/3+}$	25	*(1.24...3.2)*	DCP	63Wol
	0.2 **M** oxalic acid +0.04 **M** $KClO_4/H_2O$	$Ti^{4+/3+}$	25	*0.046*	ACP	63Smi
	1 **M** tartaric acid/H_2O	$Ti^{3+/4+}$	25	*0.009*	IP	52Ran2
	1 **M** tartaric acid +0.1 **M** NaCl/H_2O	$Ti^{4+/3+}$	30	*4.1...7.2·10^{-3}*	ACP	58Del
	1 **M** tartaric acid +0.2 **M** Na_2SO_4/H_2O	$Ti^{4+/3+}$	-	*1.7·10^{-3}*	IP	64Lai

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	1 **M** tartaric acid +0.2 **M** Na_2SO_4 +0.1 m**M** camphor/H_2O	$Ti^{4+/3+}$	-	*1.6·10^{-3}*	IP	64Lai
	1 **M** tartaric acid +0.2 **M** Na_2SO_4 +0.15 m**M** camphor/H_2O	$Ti^{4+/3+}$	-	*1.6·10^{-3}*	IP	64Lai
	1 **M** tartaric acid +0.2 **M** Na_2SO_4 +0.2 m**M** camphor/H_2O	$Ti^{4+/3+}$	-	*1.6·10^{-3}*	IP	64Lai
	1 **M** tartaric acid +0.2 **M** Na_2SO_4 +0.25 m**M** camphor/H_2O	$Ti^{4+/3+}$	-	*0.9·10^{-3}*	IP	64Lai
	1 **M** tartaric acid +0.2 **M** Na_2SO_4 +0.3 m**M** camphor/H_2O	$Ti^{4+/3+}$	-	*0.5·10^{-3}*	IP	64Lai
	1 **M** tartaric acid +0.2 **M** Na_2SO_4 +0.35 m**M** camphor/H_2O	$Ti^{4+/3+}$	-	*0.3·10^{-3}*	IP	64Lai
	1 **M** tartaric acid +0.2 **M** Na_2SO_4 +0.45 m**M** camphor/H_2O	$Ti^{4+/3+}$	-	*0.2·10^{-3}*	IP	64Lai
	1 **M** tartaric acid +0.2 **M** Na_2SO_4 +0.1 m**M** thymol/H_2O	$Ti^{4+/3+}$	-	*1.1·10^{-3}*	IP	64Lai
	1 **M** tartaric acid +0.2 **M** Na_2SO_4 +0.2 m**M** thymol/H_2O	$Ti^{4+/3+}$	-	*0.5·10^{-3}*	IP	64Lai
	1 **M** tartaric acid +0.2 **M** Na_2SO_4 +0.3 m**M** thymol/H_2O	$Ti^{4+/3+}$	-	*0.1·10^{-3}*	IP	64Lai
	1 **M** tartaric acid +0.2 **M** Na_2SO_4 +0.4 m**M** thymol/H_2O	$Ti^{4+/3+}$	-	*0.09·10^{-3}*	IP	64Lai
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Ti(bipy)_3^{0/-}$	25	*0.5*	GD	75Saj1
	1 **M** KCl/H_2O	$Tl^{+/0}$	36	*5.0*	FR	71Aga
	1 **M** KCl/H_2O, pH=2	$Tl^{+/0}$	22	*0.15*	FR	58Bar1

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	1 **M** KCl/H_2O, pH=2	$Tl^{+/0}$	22	*>0.15*	FR	61Bar1
	1 **M** KF/H_2O	$Tl^{+/0}$, 0.1 m**M**	22	*4.1*	FR	61Bar1
	1 **M** KF/H_2O	$Tl^{+/0}$, $1.05\cdot10^{-3}$	36	*3.5*	FR	71Aga
	1 **M** KF/H_2O	$Tl^{+/0}$, $2.1\cdot10^{-3}$	36	*2.5*	FR	71Aga
	1 **M** KNO_3/H_2O	$Tl^{+/0}$, $1.05\cdot10^{-3}$	36	*0.69*	FR	71Aga
	1 **M** KNO_3/H_2O, pH=2	$Tl^{+/0}$, 0.1 m**M**	22	*>0.3*	FR	61Bar1
	1 **M** KNO_3/H_2O, pH=2	$Tl^{+/0}$, 0.1 m**M**	22	*0.3*	FR	58Bar1
	0.5 **M** KSO_4/H_2O, pH=2	$Tl^{+/0}$, 0.1 m**M**	22	*>3.1*	FR	61Bar1
	0.5 **M** KSO_4/H_2O, pH=2	$Tl^{+/0}$	22	*3.1*	FR	58Bar1
	0.5 **M** KSO_4/H_2O	$Tl^{+/0}$, $1.05\cdot10^{-3}$	36	*3.5*	FR	71Aga
	0.5 **M** KSO_4/H_2O	$Tl^{+/0}$, $2.1\cdot10^{-3}$	36	*7.4*	FR	71Aga
	1 **M** $NaClO_4$/H_2O, pH=2	$Tl^{+/0}$, 0.1 m**M**	22	*>1.2*	FR	61Bar1
	1 **M** $NaClO_4$/H_2O, pH=2	$Tl^{+/0}$	22	*1.2*	FR	58Bar1
	1 **M** $NaClO_4$/H_2O	$Tl^{+/0}$, $1.05\cdot10^{-3}$	36	*3.4*	FR	71Aga
	1 **M** $NaClO_4$/H_2O	$Tl^{+/0}$, $2.1\cdot10^{-3}$	36	*8.3*	FR	71Aga
	3 **M** $NaClO_4$+0.05 **M** $HClO_4$/H_2O	$Tl^{0/+}$, $4.3\cdot10^{-5}$ **M**	20	$1.5\cdot10^{-4}$	CP	67Los
	1 **M** Na_2CO_3/H_2O	U(VI)/U(V), (r)	22	$1\cdot10^{-4}$	DCP	70Cas
	1 **M** Na_2CO_3/H_2O	U(V)/U(V), (r)	22	$8.9\cdot10^{-5}$	OC	70Cas
	1 **M** $HClO_4$/H_2O	UO_2^{2+} , (r)	25	*(2.14·10^{-2})*	OC	61Gok
	0.1 **M** H_3PO_4/H_2O	U(VI),(r)	25	*(9·10^{-7}...2.9·10^{-3})*	DCP	72Iss2
	1 **M** H_3PO_4/H_2O	U(VI),(r)	25	*(1.5·10^{-7}...5·10^{-3})*	DCP	72Iss2
	1.85 **M** H_3PO_4/H_2O	U(VI),(r)	25	*(2.4·10^{-2}...0.15)*	DCP	72Iss2
	5 **M** H_3PO_4/H_2O	U(VI),(r)	25	*(12...113)*	DCP	72Iss2
	10 **M** H_3PO_4/H_2O	U(VI),(r)	25	*(1.2...28180)*	DCP	72Iss2
	0.1 **M** H_2SO_4/H_2O	U(VI),(r)	25	*(1.71·10^{-3})*	DCP	72Iss1
	1 **M** H_2SO_4/H_2O	U(VI),(r)	25	*(4.8...7.47·10^{-3})*	DCP	72Iss1
	5 **M** H_2SO_4/H_2O	U(VI),(r)	25	*(3.16)*	DCP	72Iss1
	0.1 **M** $NaCO_3$ +$NaClO_4$/H_2O	$[UO_2(CO_3)_2(O_2)]^{4-}$ /$[UO_2(CO_3)_3]^{5-}$	25	*(2.5·10^{-8})*	DCP	68Zut
	0.05...1 **M** LiOH/H_2O	$[UO_2(OH)_3(H_2O)_3]^-$ /$[UO_2(OH)_3(H_2O)_3]^{2-}$	25	$3.5\cdot10^{-2}$	DCP	70Zut
	4 **M** $HClO_4$/H_2O	U(V)/U(IV)	25	*(4.0·10^{-8})*	DCP	71Sip
	1 **M** $HClO_4$/H_2O	$V_{aq}^{2+/3+}$	25	*0.004*	IP	52Ran2
	1 **M** $HClO_4$/H_2O	$V_{aq}^{2+/3+}$	14	*0.0036*	IP	59Ran
	1 **M** $HClO_4$/H_2O	$V_{aq}^{2+/3+}$	25	*0.0022*	IP	59Ran
	0.5 **M** H_2SO_4/H_2O	$V_{aq}^{2+/3+}$	25	*1.45·10^{-3}*	PS	69Nik

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.5 **M** H_2SO_4/H_2O	$V_{aq}^{2+/3+}$	25	*1.04·10^{-3}*	DCP	69Nik
	0.5 **M** H_2SO_4/+ 3.44 m**M** *n*-butylalcohol/H_2O	$V_{aq}^{2+/3+}$	25	*1.25·10^{-4}*	PS	69Nik
	0.5 **M** H_2SO_4/+ 6.9 m**M** *n*-butylalcohol/H_2O	$V_{aq}^{2+/3+}$	25	*1.03·10^{-3}*	PS	69Nik
	0.5 **M** H_2SO_4/+ 13.8 m**M** *n*-butylalcohol/H_2O	$V_{aq}^{2+/3+}$	25	*8.0·10^{-4}*	PS	69Nik
	0.5 **M** H_2SO_4/+1 m**M** *n*-amylalcohol/H_2O	$V_{aq}^{2+/3+}$	25	*1.23·10^{-3}*	PS	69Nik
	0.5 **M** H_2SO_4/+2 m**M** *n*-amylalcohol/H_2O	$V_{aq}^{2+/3+}$	25	*9.9·10^{-4}*	PS	69Nik
	0.5 **M** H_2SO_4/+4 m**M** *n*-amylalcohol/H_2O	$V_{aq}^{2+/3+}$	25	*6.1·10^{-4}*	PS	69Nik
	0.5 **M** H_2SO_4/+ 0.213 m**M** *n*-hexyl-alcohol/H_2O	$V_{aq}^{2+/3+}$	25	*1.25·10^{-3}*	PS	69Nik
	0.5 **M** H_2SO_4/+ 0.43 m**M** *n*-hexyl-alcohol/H_2O	$V_{aq}^{2+/3+}$	25	*1.0·10^{-4}*	PS	69Nik
	0.5 **M** H_2SO_4/+ 0.85 m**M** *n*-hexyl-alcohol/H_2O	$V_{aq}^{2+/3+}$	25	*5.4·10^{-4}*	PS	69Nik
	1 **M** H_2SO_4/H_2O	$V_{aq}^{2+/3+}$	25	*0.0014*	IP	59Ran
	0.5 **M** K-oxalate/H_2O	$V_{aq}^{2+/3+}$	25	*0.0014*	IP	52Ran2
	0.1 **M** KPF_6/H_2O	$V_{aq}^{2+/3+}$	25	*(3.0·10^{-2})*	CV	80Bar1
	0.4 **M** KPF_6/H_2O	$V_{aq}^{2+/3+}$	25	*(7.0·10^{-2})*	CV	80Bar1
	1 **M** $HClO_4/H_2O$	$V^{3+/2+}$	24	*0.0044*	OC	65Sri
	1 **M** $HClO_4/H_2O$	$V^{3+/2+}$	24	*0.0057*	CP	65Sri
	1 **M** $HClO_4/H_2O$	$V^{3+/2+}$	25	*0.0036*	IP	67Ell
	0.02 **M** $HClO_4$+ 0.2 **M** $NaClO_4/H_2O$	$V^{3+/2+}$	25	*0.00436*	IP	67Ell
	0.02 **M** $HClO_4$ +1 **M** $NaClO_4/H_2O$	$V^{3+/2+}$	25	*0.00469*	IP	67Ell
	1 **M** $HClO_4$/ H_2O	$V^{3+/2+}$	20	*0.004*	IP	52Ran2
	1 **M** $HClO_4$/ H_2O	$V^{3+/2+}$	14.2	*0.0022*	DCP,IP	59Ran
	1 **M** $HClO_4$/ H_2O	$V^{3+/2+}$	25	*0.00364*	DCP,IP	59Ran
	1 **M** $HClO_4$/ H_2O	$V^{3+/2+}$	20	*0.0032*	IP	61Jos
	1 **M** NH_2SO_3H/H_2O	$V_{aq}^{3+/2+}$	25	*0.00336*	IP	71Bie

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	1 **M** NH_2SO_3H /NMF [88])	$V_{aq}^{3+/2+}$	25	*0.00116*	IP	71Bie
	1 **M** $NaClO_4$ +0.2 **M** $HClO_4$ +0.01 **M** $(NH_4)_2SO_4/H_2O$	$V_{aq}^{3+/2+}$	25	*0.0048*	ACP	68Ran
	5 **M** $NaClO_4$ +0.2 **M** $HClO_4$ +0.01 **M** $(NH_4)_2SO_4/H_2O$	$V_{aq}^{3+/2+}$	25	*0.0077*	ACP	68Ran
	0.5 **M** H_2SO_4/ H_2O	$V_{aq}^{3+/2+}$	25	*0.00145*	PS	71Nik
	1 **M** H_2SO_4/ H_2O	$V_{aq}^{3+/2+}$	15.3	*0.00139*	DCP, IP	59Ran
	1 **M** H_2SO_4/ H_2O	$V_{aq}^{3+/2+}$	15	*0.00103*	DCP	62Ran
	0.2 **M** Bu_4NClO_2/DMF	$V(bipy)_3^{2+/+}$	25	*0.35*	GD	75Saj1
	0.2 **M** Bu_4NClO_2/DMF	$V(bipy)_3^{+/0}$	25	*1.2*	GD	75Saj1
	0.2 **M** Bu_4NClO_2/DMF	$V(bipy)_3^{0/-}$	25	*0.66*	GD	75Saj1
	0.25 **M** $CsSO_4/H_2O$	$Zn^{2+/0}$	22	$4.1\cdot10^{-4}$	DCP	62Kor
	1 **M** KBr/H_2O	$Zn^{2+/0}$	25	$8.0\cdot10^{-3}$	ACP	61Kam
	1 **M** KBr+2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	$10.5\cdot10^{-3}$	DCP	62Tam
	1 **M** KBr+2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	$10.5\cdot10^{-3}$	DCP	63Tam
	2.88 **M** KBr+2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	$4.8\cdot10^{-3}$	DCP	62Tam
	2.88 **M** KBr+2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	$4.8\cdot10^{-3}$	DCP	63Tam
	0.1 **M** KCl/H_2O	$Zn^{2+/0}$, 2 m**M** $ZnCl_2$	1	$5.2\cdot10^{-3}$	CP	61Riu
	0.1 **M** KCl/H_2O	$Zn^{2+/0}$, 2 m**M** $ZnCl_2$	23	$1.23\cdot10^{-2}$	CP	61Riu
	0.1 **M** KCl/H_2O	$Zn^{2+/0}$, 2 m**M** $ZnCl_2$	38.5	$1.55\cdot10^{-2}$	CP	61Riu
	0.1 **M** KCl/H_2O	$Zn^{2+/0}$, 2 m**M** $ZnCl_2$	50	$3.24\cdot10^{-2}$	CP	61Riu
	0.1 **M** KCl/H_2O	$Zn^{2+/0}$, 4 m**M** $ZnCl_2$	19	$1.58\cdot10^{-2}$	CP	61Riu
	0.1 **M** KCl/H_2O	$Zn^{2+/0}$, 4 m**M** $ZnCl_2$	34	$1.78\cdot10^{-2}$	CP	61Riu
	0.1 **M** KCl/H_2O	$Zn^{2+/0}$, 4 m**M** $ZnCl_2$	48	$3.31\cdot10^{-2}$	CP	61Riu
	1 **M** KCl/H_2O	$Zn^{2+/0}$	25	$3.1\cdot10^{-3}$	OP	55Mat
	1 **M** KCl/H_2O	$Zn^{2+/0}$	25	$4.0\cdot10^{-3}$	ACP	61Kam
	1 **M** KCl/H_2O	$Zn^{2+/0}$	15	$3.5\cdot10^{-3}$	CP	61Riu
	1 **M** KCl/H_2O	$Zn^{2+/0}$	27.5	$4.7\cdot10^{-3}$	CP	61Riu
	1 **M** KCl/H_2O	$Zn^{2+/0}$	37.5	$7.7\cdot10^{-3}$	CP	61Riu
	1 **M** KCl/H_2O	$Zn^{2+/0}$	49.5	$9.8\cdot10^{-3}$	CP	61Riu
	1 **M** KCl/H_2O, pH=3	$Zn^{2+/0}$	25	$4.0\cdot10^{-3}$	DCP	67Tim

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	1 **M** KCl+ 2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*7.0·10^{-3}*	DCP	62Tam
	1 **M** KCl+ 2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*7.0·10^{-3}*	DCP	63Tam
	1 **M** KCl+ 2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$	25	*3.4·10^{-3}*	IP	59Tam
	0.1...1 **M** ZnL +KCl/H_2O [119])	$Zn^{2+/0}$	-	*3.4·10^{-5}*	DCP	62Mat
	0.1...1 **M** ZnL+ KCl/H_2O [119])	$ZnL_2^{/0}$	-	*3.4·10^{-4}*	DCP	62Mat
	1 **M** KI/H_2O	$Zn^{2+/0}$	25	*7.0·10^{-2}*	ACP	61Kam
	0.1 **M** KNO_3+ 0.1...2 **M** ethylene-diamine+0.02 % gelatin/H_2O	$Zn^{2+/0}$	25	*1.0·10^{-4}*	DCP	56Mor
	1 **M** KNO_3/H_2O	$Zn^{2+/0}$	25	*3.5·10^{-3}*	ACP	61Kam
	0.1 **M** KNO_3/H_2O	$Zn^{2+/0}$, 2 m**M** $ZnCl_2$	10.5	*1.7·10^{-2}*	CP	61Riu
	0.1 **M** KNO_3/H_2O	$Zn^{2+/0}$, 2 m**M** $ZnCl_2$	24	*1.32·10^{-2}*	CP	61Riu
	0.1 **M** KNO_3/H_2O	$Zn^{2+/0}$, 2 m**M** $ZnCl_2$	36.5	*2.82·10^{-2}*	CP	61Riu
	0.1 **M** KNO_3/H_2O	$Zn^{2+/0}$, 2 m**M** $ZnCl_2$	50	*3.34·10^{-2}*	CP	61Riu
	1 **M** KNO_3/H_2O	$Zn^{2+/0}$, 2 m**M** $ZnCl_2$	21	*0.4·10^{-2}*	CP	61Riu
	1 **M** KNO_3/H_2O	$Zn^{2+/0}$, 2 m**M** $ZnCl_2$	36	*0.46·10^{-2}*	CP	61Riu
	1 **M** KNO_3/H_2O	$Zn^{2+/0}$, 2 m**M** $ZnCl_2$	54	*1.2·10^{-2}*	CP	61Riu
	1 **M** KNO_3/H_2O	$Zn^{2+/0}$	25	*3.9·10^{-3}*	PS	63Tos
	3 **M** KNO_3/H_2O	$Zn^{2+/0}$	25	*3.25·10^{-3}*	OC	58Ima
	3 **M** KNO_3+0.002 % gelatin/H_2O	$Zn^{2+/0}$	25	*6.0·10^{-3}*	OC	58Ima
	3 **M** KNO_3+0.004% gelatin/H_2O	$Zn^{2+/0}$	25	*7.5·10^{-3}*	OC	58Ima
	3 **M** KNO_3+0.006 % gelatin/H_2O	$Zn^{2+/0}$	25	*8.7·10^{-3}*	OC	58Ima
	3 **M** KNO_3+0.008 % gelatin/H_2O	$Zn^{2+/0}$	25	*5.1·10^{-3}*	OC	58Ima
	3 **M** KNO_3+0.01 % gelatin/H_2O	$Zn^{2+/0}$	25	*4.25·10^{-3}*	OC	58Ima
	1 **M** KNO_3+2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*5.4·10^{-3}*	DCP	62Tam
	1 **M** KNO_3+2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*5.4·10^{-3}*	DCP	63Tam
	2 **M** KOH+KCl/H_2O	$Zn^{2+/0}$	25	*5.1·10^{-4}*	DCP	59Mat2
	1 **M** KSCN/H_2O	$Zn^{2+/0}$	25	*1.7·10^{-2}*	ACP	61Kam
	0.25 **M** K_2SO_4/H_2O	$Zn^{2+/0}$	22	*1.3·10^{-3}*	DCP	62Kor
	0.1 **M** $LaCl_3$/H_2O	$Zn^{2+/0}$	22	*2.5·10^{-3}*	DCP	62Kor
	0.25 **M** Li_2SO_4/H_2O	$Zn^{2+/0}$	22	*2.4·10^{-3}*	DCP	62Kor
	0.25...2.8 **M** NH_4Cl +NH_3/H_2O	$Zn^{2+/0}$	25	*1.3·10^{-5}*	DCP	59Mat2

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	1 **M** $(NH_4)_2SO_4$ +NH_3/H_2O	$Zn^{2+/0}$	25	*6.3·10^{-5}*	OP	59Ima
	x **M** NH_4OH +x **M** NH_4/Cl +(3-x) **M** $NaCl/H_2O$	$Zn^{2+/0}$	-	*1 .. 2.9·10^{-5}*	DCP	68Pop
	1 **M** $NaBr/H_2O$	$Zn^{2+/0}$	25	*7.7·10^{-3}*	IP	70Mou
	0.55 **M** NaCl +1 m**M** HCl/H_2O	$Zn^{2+/0}$	25	*0.005*	DCP	71Ruz
	0.55 **M** NaCl+0.01 **M** borate buffer/H_2O	$Zn^{2+/0}$	25	*0.0086*	DCP	71Ruz
	1 **M** $NaCl/H_2O$	$Zn^{2+/0}$	25	*3.9·10^{-3}*	IP	70Mou
	3 **M** NaCl+KOH/H_2O	$Zn^{2+/0}$	-	*1.7...3.6·10^{-5}*	DCP	68Pop
	0.1 **M** $NaClO_4$ +0.1 m**M** $HClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*11.9·10^{-3}*	DCP	69Tan2
	0.5 **M** $NaClO_4$ +0.1 m**M** $HClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*5.3·10^{-3}*	DCP	69Tan2
	0.75 **M** $NaClO_4$ +0.1 m**M** $HClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*3.4·10^{-3}*	DCP	69Tan2
	0.9 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$	25	1.02	IP	53Ger2
	0.9 **M** $NaClO_4$ +0.1 **M** NaCl +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*3.3·10^{-3}*	DCP	62Tam
	0.9 **M** $NaClO_4$ +0.1 **M** NaCl +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*3.3·10^{-3}*	DCP	63Tam
	0.9 **M** $NaClO_4$ +0.1 **M** NaI +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*4.7·10^{-3}*	DCP	62Tam
	0.9 **M** $NaClO_4$ +0.1 **M** NaI +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*4.7·10^{-3}*	DCP	63Tam
	0.99 **M** $NaClO_4$ +0.01 **M** KCl +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*2.8·10^{-3}*	DCP	62Tam
	0.99 **M** $NaClO_4$ +0.01 **M** KCl +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*2.8·10^{-3}*	DCP	63Tam
	0.99 **M** $NaClO_4$ +0.01 **M** NaI +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*2.7·10^{-3}*	DCP	62Tam
	0.99 **M** $NaClO_4$ +0.01 **M** NaI +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*2.7·10^{-3}*	DCP	63Tam

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	1 **M** $NaClO_4$ /H_2O, acidifed	$Zn^{0/2+}$, 1 m**M**	20	$5.7\cdot10^{-4}$	CP	67Los
	1 **M** $NaClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	0	*3.3·10^{-4}*	DCP	62Tam
	1 **M** $NaClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	0	*3.3·10^{-4}*	DCP	63Tam
	1 **M** $NaClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	15	*1.54·10^{-3}*	DCP	62Tam
	1 **M** $NaClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	15	*1.54·10^{-3}*	DCP	63Tam
	1 **M** $NaClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*4.0·10^{-3}*	DCP	62Tam
	1 **M** $NaClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*4.0·10^{-3}*	DCP	63Tam
	1 **M** $NaClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	35	*5.2·10^{-3}*	DCP	62Tam
	1 **M** $NaClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	35	*5.2·10^{-3}*	DCP	63Tam
	1 **M** $NaClO_4$ +0.1 m**M** $HClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	0.3	*0.64·10^{-3}*	DCP	69Tan2
	1 **M** $NaClO_4$ +0.1 m**M** $HClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	15	*1.6·10^{-3}*	DCP	69Tan2
	1 **M** $NaClO_4$ +0.1 m**M** $HClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	25	*2.6·10^{-3}*	DCP	69Tan2
	1 **M** $NaClO_4$ +0.1 m**M** $HClO_4$ +2 µ**M** LEO/H_2O [98])	$Zn^{2+/0}$, 1 m**M** $Zn(NO_3)_2$	35	*4.29·10^{-3}*	DCP	69Tan2
	0.014 **M** $NaNO_3/H_2O$	$Zn^{2+/0}$	-	*5.0·10^{-2}*	PP	66Chr
	0.1 **M** $NaNO_3/H_2O$	$Zn^{2+/0}$	22	*5.0·10^{-2}*	DCP	62Kor
	0.1 **M** $NaNO_3$+0.05 **M** $(CH_3)_4NCl/H_2O$	$Zn^{2+/0}$	22	*1.5·10^{-3}*	DCP	62Kor
	0.1 **M** $NaNO_3$+0.05 **M** $(C_2H_5)_4NCl/H_2O$	$Zn^{2+/0}$	22	*1.3·10^{-3}*	DCP	62Kor
	0.1 **M** $NaNO_3$+0.05 **M** $(C_4H_9)_4NCl/H_2O$	$Zn^{2+/0}$	22	*1.3·10^{-3}*	DCP	62Kor
	0.105 **M** $NaNO_3/H_2O$	$Zn^{2+/0}$	-	*2.0·10^{-2}*	PP	66Chr
	0.2 **M** $NaNO_3/H_2O$	$Zn^{2+/0}$	25	*1.5·10^{-2}*	DCP	71Ruz
	0.25 **M** $NaNO_3/H_2O$	$Zn^{2+/0}$	22	*2.0·10^{-2}*	DCP	62Kor
	0.35 **M** $NaNO_3/H_2O$	$Zn^{2+/0}$	-	*6.5·10^{-3}*	PP	66Chr
	0.49 **M** $NaNO_3/H_2O$	$Zn^{2+/0}$	-	*4.1·10^{-3}*	PP	66Chr

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Hg	0.5 **M** $NaNO_3/H_2O$	$Zn^{2+/0}$	22	*5.4·10^{-3}*	DCP	62Kor
	0.5 **M** $NaNO_3$ +0.5 m**M** HNO_3/H_2O	$Zn^{2+/0}$	25	*4.4·10^{-3}*	PS	69Tan1
	0.75 **M** $NaNO_3/H_2O$	$Zn^{2+/0}$	22	*3.3·10^{-3}*	DCP	62Kor
	1 **M** $NaNO_3/H_2O$	$Zn^{2+/0}$	22	*2.8·10^{-3}*	DCP	62Kor
	1 **M** $NaNO_3/H_2O$	$Zn^{2+/0}$	25	*0.31·10^{-2}*	DCP	71Ruz
	1.05 **M** $NaNO_3/H_2O$	$Zn^{2+/0}$	-	*3.2·10^{-3}*	PP	66Chr
	2.45 **M** $NaNO_3/H_2O$	$Zn^{2+/0}$	-	*2.4·10^{-3}*	PP	66Chr
	4 **M** $NaNO_3$ +NaL/H_2O [120])	$Zn^{2+/0}$	25	*1.8·10^{-4}*	DCP	63Mat
	4 **M** $NaNO_3$ +NaL/H_2O [120])	$ZnL^{+/0}$	25	*4.4·10^{-3}*	DCP	63Mat
	4 **M** $NaNO_3$ +NaL/H_2O [120])	$ZnL_3^{-/0}$	25	*6.2·10^{-4}*	DCP	63Mat
	0.05...0.2 **M** $NaNO_3$ +NaL/H_2O [121])	$Zn^{2+/0}$	25	*1.15·10^{-5}*	DCP	71Aya
	0.05...0.2 **M** $NaNO_3$ +NaL/H_2O [121])	$ZnL^{+/0}$	25	*0.99·10^{-3}*	DCP	71Aya
	0.05...0.2 **M** $NaNO_3$ +NaL/H_2O [121])	$ZnL_2^{/0}$	25	*1.0·10^{-5}*	DCP	71Aya
	1 **M** NaN_3/H_2O	$Zn^{2+/0}$	25	*7.4·10^{-3}*	IP	70Mou
	1 **M** $NaOH/H_2O$	$Zn^{2+/0}$	25	*7.4·10^{-5}*	ACP	61Kam
	1 **M** $NaOH/H_2O$	$Zn^{2+/0}$	20	*1.8·10^{-4}*	DCP	62Beh
	0.25 **M** Na_2SO_4/H_2O	$Zn^{2+/0}$	22	*2.4·10^{-3}*	DCP	62Kor
	0.5 **M** Na_2SO_4 +1 m**M** H_2SO_4	$Zn^{2+/0}$	22	*2.08·10^{-3}*	CP	62Bat
	0.5 **M** Na_2SO_4 +1 m**M** H_2SO_4 +n-butanol/H_2O	$Zn^{2+/0}$	22	*9.0·10^{-4}*	DCP	65Sat
	0.1 **M** TEAP [74]) /DMF [54])	benzoquinone (r)	25	*0.0052*	CV	73Cap
	acetate buffer/H_2O, pH=3.9	benzoquinone (r)	-	*(1.4·10^{-3})*	CP	63Hal
	0.1 **M** TEAP [74]) /acetone	benzophenone (r)	22	*2.0*	IP	93Faw
	0.1 **M** TEAP [74]) /acetonitrile	benzophenone (r)	22	*2.7*	IP	93Faw
	0.1 **M** TPAP [74]) /acetonitrile	benzophenone (r)	22	*1.03*	IP	93Faw
	0.1 **M** TBAP [52]) /acetonitrile	benzophenone (r)	22	*0.42*	IP	93Faw
	0.1 **M** THxAP [122]) /acetonitrile	benzophenone (r)	22	*0. 21*	IP	93Faw
	0.1 **M** TOAP [123]) /acetonitrile	benzophenone (r)	22	*0. 11*	IP	93Faw

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.1 **M** TEAP [74]) /propionitrile	benzophenone (r)	22	*1.82*	IP	93Faw
	0.1 **M** TEAP [74]) /dimethylacetamide	benzophenone (r)	22	*0.51*	IP	93Faw
	0.1 **M** TEAP [74]) /dimethylsulfoxide	benzophenone (r)	22	*0.31*	IP	93Faw
	0.1 **M** TEAP [74]) /HMPA [124])	benzophenone (r)	22	*0.11*	IP	93Faw
	0.1 **M** TEAP [74]) /PC [113])	benzophenone (r)	22	*0.18*	IP	93Faw
	acetate buffer /H_2O, pH=3.9	1,4-naphtha-quinone(r)	-	*(1.6·10^{-3})*	CP	63Hal
	acetate buffer /methanol, pH=3.9	benzoquinone (r)	-	*(1.1·10^{-3})*	CP	63Hal
	acetate buffer /methanol, pH=3.9	1,4-naphtha-quinone(r)	-	*(1.4·10^{-3})*	CP	63Hal
	acetate buffer /methanol, pH=3.9	1,2-benz-anthraquinone (r)	-	*(2.2·10^{-3})*	CP	63Hal
	acetate buffer /methanol, pH=3.9	anthraquinone (r)	-	*(2.8·10^{-3})*	CP	63Hal
	acetate buffer /methanol, pH=3.9	9,10-phen-anthraquinone (r)	-	*(0.5·10^{-3})*	CP	63Hal
	acetate buffer /ethanol, pH=3.9	benzoquinone (r)	-	*(0.7·10^{-3})*	CP	63Hal
	acetate buffer /ethanol, pH=3.9	1,4-naphtha-quinone(r)	-	*(1.5·10^{-3})*	CP	63Hal
	acetate buffer/DMF +H_2O, pH=3.9 [54])	anthraquinone (r)	-	*(0.2·10^{-3})*	CP	63Hal
	phosphate buffer /H_2O, pH=6.75	1-oxo-10-methyl-1.10-dihydro-phenazine (r)	25	*0.5*	IP	52Ran2
	phosphate buffer /H_2O, pH=6.75	tetramethyl-paraphenylene diamine (o)	25	*0.5*	IP	52Ran2
	0.05 **M** TBAP /acetonitrile	*p*-dicyanobenzene	22	*0.95*	ACP	96Win
	0.05 **M** TBAP [52]) /acetone	*p*-dicyanobenzene	22	*0.75*	ACP	96Win
	0.05 **M** TBAP [52]) /dichloromethane	*p*-dicyanobenzene	22	*0.54*	ACP	96Win
	0.05 **M** TBAP [52]) /DMF [54])	*p*-dicyanobenzene	22	*0.27*	ACP	96Win
	0.05 **M** TBAP [52]) /PC [113])	*p*-dicyanobenzene	22	*0.14*	ACP	96Win

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Molybdenum	0.001 **N** HCl/H_2O	H^+/H_2	-	$10^{-7.12}$	CP	57Pen
(Mo)	0.01 **N** HCl/H_2O	H^+/H_2	-	$10^{-7.19}$	CP	57Pen
	0.1 **N** HCl/H_2O	H^+/H_2	-	$10^{-6.45}$	CP	57Pen
	0.001 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-7.27}$	CP	57Pen
	0.01 **N** HCl/H_2O	H^+/H_2	-	$10^{-6.42}$	CP	57Pen
	0.1 **N** HCl/H_2O	H^+/H_2	-	$10^{-6.37}$	CP	57Pen
	1 **M** $NaOH/H_2O$	hydrogen evolution	25	$5.4\cdot10^{-6}$	CP	00Jak
MoNi	1 **M** NaOH/ H_2O	hydrogen evolution	25	$79\cdot10^{-6}$	CP	00Jak
$MoNi_2$	1 **M** NaOH/ H_2O	hydrogen evolution	25	$70\cdot10^{-6}$	CP	00Jak
$MoNi_3$	1 **M** NaOH/ H_2O	hydrogen evolution	25	$32\cdot10^{-6}$	CP	00Jak
$MoNi_4$	1 **M** NaOH/ H_2O	hydrogen evolution	25	$13\cdot10^{-6}$	CP	00Jak
Nickel (Ni)	1 **M** NaOH/ H_2O	hydrogen evolution	25	$3.3\cdot10^{-6}$	CP	00Jak
	1 **M** H_2SO_4+0.01 **M** $MgSO_4/H_2O$	$Fe^{2+/3+}$, 0.05 **M**	25	$7.5\cdot10^{-5}$	CP	64Mak [125])
	0.5 **N** H_2SO_4/H_2O	H^+/H_2	-	$10^{-5.22}$	CP	64Boc3
	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-5.3}$	CP	67Man
	0.5 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-6.1}$	CP	64Boc3
	0.2 **M** HCl/H_2O, pH=0.67	$Ni^{2+/0}$, 10 m**M**	-	$1.7\cdot10^{-9}$	CP	59Hol
	2 **M** HCl/H_2O, pH=0.67	$Ni^{2+/0}$, 10 m**M**	-	$1.1\cdot10^{-8}$	CP	59Hol
	1 **M** HCl+1 **M** KCl /H_2O, pH=0	$Ni^{2+/0}$, 10 m**M**	-	$6.9\cdot10^{-9}$	CP	59Hol
	0.1 **M** H_2SO_4/H_2O, pH=0.8	$Ni^{2+/0}$, 5 m**M**	-	$2.4\cdot10^{-9}$	CP	59Hol
	1 **M** H_2SO_4/H_2O, pH=0	$Ni^{2+/0}$, 5 m**M**	-	$8.3\cdot10^{-10}$	CP	59Hol
	1.5 **M** KCl+0.2 **M** HCl/H_2O, pH=0.6	$Ni^{2+/0}$, 10 m**M**	-	$1.7\cdot10^{-9}$	CP	59Hol
	2 **M** KCl/H_2O, pH=6	$Ni^{2+/0}$, 10 m**M**	-	$1.0\cdot10^{-8}$	CP	59Hol
$Ni_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Ni^{2+/0}$	25	*$(2.74\cdot10^{-4})$*	CP	63Mac
	1 **M** Na_2SO_4/H_2O, pH=5.9	$Ni^{2+/0}$, 5 m**M**	-	$4.2\cdot10^{-8}$	CP	59Hol
	0.5 **M** $NiSO_4$+acetate buffer/H_2O, pH=2.3	$Ni^{2+/0}$	35	$1.0\cdot10^{-6}$	GCI	47Sal
	0.5 **M** $NiSO_4$+acetate buffer/H_2O, pH=4.4	$Ni^{2+/0}$	35	$1.0\cdot10^{-6}$	GCI	47Sal
	0.5 **M** $NiSO_4$+acetate buffer/H_2O, pH=5.5	$Ni^{2+/0}$	35	$1.0\cdot10^{-7}$	GCI	47Sal
	0.5 **M** $NiSO_4$+acetate buffer/H_2O, pH=6.7	$Ni^{2+/0}$	35	$1.0\cdot10^{-5}$	GCI	47Sal
	0.5 **M** $NiSO_4$+acetate buffer/H_2O, pH=7.9	$Ni^{2+/0}$	35	$1.0\cdot10^{-5}$	GCI	47Sal
	1.33 **M** NaOH +1.95 **M** $NaNO_3/H_2O$	NO_3^-/NO	25	$4.17\cdot10^{-7}$	CV	96Boc [126])

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Niobium (Nb)	0.1 **N** H_2SO_4/H_2O	H^+/H_2, 0.1 **N**	25	$2.0\cdot10^{-7}$	CP	63Rot
	0.1 **N** H_2SO_4/H_2O	H^+/H_2, 0.1 **N**	40	$4.0\cdot10^{-7}$	CP	63Rot
	0.1 **N** H_2SO_4/H_2O	H^+/H_2, 0.1 **N**	55	$1.0\cdot10^{-6}$	CP	63Rot
	1 **N** H_2SO_4/H_2O	H^+/H_2, 0.1 **N**	25	$5.0\cdot10^{-7}$	CP	63Rot
	1 **N** H_2SO_4/H_2O	H^+/H_2, 0.1 **N**	40	$1.0\cdot10^{-6}$	CP	63Rot
	1 **N** H_2SO_4/H_2O	H^+/H_2, 0.1 **N**	55	$3.0\cdot10^{-6}$	CP	63Rot
	1 **N** H_2SO_4/H_2O	H^+/H_2, 1 **N**	25	$1.9\cdot10^{-7}$	GCI [127])	65Koz
Palladium (Pd)	0.4 **M** K_2SO_4+0.01 N H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	0.75	TPF	88Boc
	0.4 **M** K_2SO_4+0.01 N H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	1.23	TPF	88Boc [128])
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$; 0.1**M**	25	0.012	SS	68Boc2
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$; 0.01**M**	25	0.006	SS	68Boc2
	1 **N** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	0.5	RDE	83Vie
	0.01 **N** HCl/H_2O	H^+/H_2	-	$10^{-4.18}$	CP	57Pen
	0.1 **N** HCl/H_2O	H^+/H_2	-	$10^{-3.25}$	CP	57Pen
	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-3.0}$	CP	65Mat
	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-5.9}$	CP	67Man
	1 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-3.1}$	CP	61Shi
	0.001 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-5.88}$	CP	57Pen
	0.01 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-5.56}$	CP	57Pen
	0.1 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-5.01}$	CP	57Pen
	1 **M** HBr +3 **M** H_2SO_4/H_2O	$Pd^{2+/0}$, 0.05 **M**	25	$3.9\cdot10^{-5}$	GM	66Kra1
	1 **M** HCl +3 **M** H_2SO_4/H_2O	$Pd^{2+/0}$, 0.1 **M**	25	$4.8\cdot10^{-5}$	GM	66Kra2
	1 **M** KOH/H_2O	$MnO_4^{-/2-}$	20	$0.67\cdot10^{-2}$	CP,GS	67Thi1
	1 **M** KF/H_2O	$[Ru(NH_3)_6]^{3+/2+}$	20	97	TPF	85Iwa
	0.1 **M** TEAP [74]) /DMF [54])	benzoquinone (r)	25	*0.0057*	CV	73Cap
	HCl-buffer/H_2O, pH=1.1	hydroquinone /quinone, 5.5 m**M**	25	$1.9\cdot10^{-4}$	CF	62Tom [77])
	acetate buffer/H_2O, pH=3.0	hydroquinone /quinone, 5.5 m**M**	25	$2.5\cdot10^{-5}$	CF	62Tom [77])
	acetate buffer/H_2O, pH=5.0	hydroquinone /quinone, 5.5 m**M**	25	$4.8\cdot10^{-5}$	CF	62Tom [77])
Platinum (Pt)	3 **M** $HClO_4/H_2O$	$Ce^{4+/3+}$, 3 m**M**	0	*$0.72\cdot10^{-9}$*	RT	56Fro1
	3 **M** $HClO_4/H_2O$	$Ce^{4+/3+}$, 6 m**M**	0	*$2.9\cdot10^{-9}$*	RT	56Fro1
	3 **M** $HClO_4/H_2O$	$Ce^{4+/3+}$, 3 m**M**	0	*$0.25\cdot10^{-9}$*	IP	56Fro2
	3 **M** $HClO_4/H_2O$	$Ce^{4+/3+}$, 6 m**M**	0	*$1.0\cdot10^{-9}$*	IP	56Fro2
	3 **M** $HClO_4/H_2O$	$Ce^{4+/3+}$, 3 m**M**	20	*$0.5\cdot10^{-9}$*	IP	56Fro2

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Pt	3 **M** $HClO_4/H_2O$	$Ce^{4+/3+}$, 6 m**M**	20	*2.0·10^{-9}*	IP	56Fro2
	1 **M** H_2SO_4/H_2O	$Ce^{4+/3+}$	25	*3.7·10^{-4}*	RDE	63Gal
	1 **M** H_2SO_4/H_2O	$Ce^{4+/3+}$	25	*3.7·10^{-4}*	CP	63Ada
	1 **N** H_2SO_4/H_2O	$Ce^{4+/3+}$, 1 m**M**	25	43.5·10^{-6}	CP	51Vet
	0.1 **N** H_2SO_4/H_2O	$Ce^{4+/3+}$, 1 m**M**	25	40.5·10^{-6}	CP	51Vet
	0.1 **N** H_2SO_4 +0.2 **M** KNO_3/H_2O	$Ce^{4+/3+}$, 1 m**M**	25	17.4·10^{-6}	CP	51Vet
	0.1 **N** H_2SO_4 +1 **M** KNO_3/H_2O	$Ce^{4+/3+}$, 1 m**M**	25	15.8·10^{-6}	CP	51Vet
	0.1 **N** H_2SO_4+0.01 **M** $KAl(SO_4)_2/H_2O$	$Ce^{4+/3+}$, 1 m**M**	25	17.3·10^{-6}	CP	51Vet
	0.1 **N** H_2SO_4+0.1 **M** $KAl(SO_4)_2/H_2O$	$Ce^{4+/3+}$, 1 m**M**	25	11.7·10^{-6}	CP	51Vet
	1 **M** HNO_3/H_2O	$Ce^{4+/3+}$, 1 m**M**	25	28.1·10^{-6}	CP	51Vet
	2 **M** HNO_3/H_2O	$Ce^{4+/3+}$, 1 m**M**	25	25.9·10^{-6}	CP	51Vet
	4 **M** HNO_3/H_2O	$Ce^{4+/3+}$, 1 m**M**	25	15.5·10^{-6}	CP	51Vet
	7 **M** HNO_3/H_2O	$Ce^{4+/3+}$, 1 m**M**	25	10.6·10^{-6}	CP	51Vet
	10 **M** HNO_3/H_2O	$Ce^{4+/3+}$, 1 m**M**	25	5.5·10^{-6}	CP	51Vet
	1 **M** HNO_3/H_2O	$Ce^{4+/3+}$, 10 m**M**	25	56.23·10^{-6}	CP	51Vet
	0.025 **M** TBAP [52]) /DMF [54])	$[Co(cp)_2]$ (o)	-	*0.4*	ACP	93Win
	0.05 **M** TBAP [52]) /DMF [54])	$[Co(cp)_2]$ (o)	-	*0.37*	ACP	93Win
	0.1 **M** TBAP [52]) /DMF [54])	$[Co(cp)_2]$ (o)	-	*0.31*	ACP	93Win
	0.25 **M** TBAP [52]) /DMF [54])	$[Co(cp)_2]$ (o)	-	*0.3*	ACP	93Win
	1 **M** NH_4Cl +1 **M** NH_3/H_2O	$[Co(NH_3)_6]Cl_3$ (r)	25	*4.5·10^{-6}*	CF	70Suz
	0.1 **M** KPF_6/H_2O	$[Co(NH_3)_6]^{3+/2+}$	25	*(2.0·10^{-4})*	RDE	80Bar1
	0.1 **M** KPF_6 +5 m**M** NaI/H_2O	$[Co(NH_3)_6]^{3+/2+}$	25	*(2.0·10^{-4})*	RDE	80Bar1
	7 **M** NH_3+1 m**M** NH_4Cl/H_2O	$[Co(NH_3)_6]^{3+/2+}$	25	*(2.3·10^{-3})*	RDE	69Bar2
	7 **M** NH_3+1 m**M** NH_4Cl/H_2O	$[Co(NH_3)_6]^{3+/2+}$	25	*(1.33·10^{-4})*	RDE	71Bar1
	7 **M** NH_3+0.2 **M** NH_4Cl/H_2O	$[Co(NH_3)_6]^{3+/2+}$	25	*(2.3·10^{-3})*	RDE	69Bar2
	7 **M** NH_3+0.2 **M** NH_4Cl/H_2O	$[Co(NH_3)_6]^{3+/2+}$	25	*(2.3·10^{-3})*	RDE	71Bar1
	7 **M** NH_3 +1 **M** NH_4Cl/H_2O	$[Co(NH_3)_6]^{3+/2+}$	25	*(5.25·10^{-4})*	RDE	69Bar2
	7 **M** NH_3 +1 **M** NH_4Cl/H_2O	$[Co(NH_3)_6]^{3+/2+}$	25	*(5.25·10^{-4})*	RDE	71Bar1
	7 **M** NH_3+0.1 **M** $(NH_4)_2SO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	25	*(7.2·10^{-4})*	RDE	69Bar2

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	7 **M** NH_3+0.1 **M** $(NH_4)_2SO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	25	*(7.2·10^{-4})*	RDE	71Bar1
	7 **M** NH_3+0.5 **M** $(NH_4)_2SO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	25	*(2.1·10^{-4})*	RDE	69Bar2
	7 **M** NH_3+0.5 **M** $(NH_4)_2SO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	25	*(2.1·10^{-4})*	RDE	71Bar1
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	25	*(1.5·10^{-3})*	RDE	80Bar1
	0.1 **M** KPF_6/H_2O	$[Co(NH_3)_5]F^{2+/+}$	25	*(2.0·10^{-3})*	RDE	80Bar1
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5]F^{2+/+}$	25	*(5.0·10^{-4})*	RDE	80Bar1
	0.1 **M** KPF_6/H_2O	$[Co(NH_3)_5]OH_2^{3+/2+}$	25	*(~5.0·10^{-2})*	RDE	80Bar1
	0.1 **M** KPF_6 +5 m**M** NaI /H_2O	$[Co(NH_3)_5]OH_2^{3+/2+}$	25	*(~0.1)*	RDE	80Bar1
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5]OH_2^{3+/2+}$	25	*(~5.0·10^{-2})*	RDE	80Bar1
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	$[Co(NH_3)_5]Cl^{2+/1+}$	24	*(7.0·10^{-3})*	PP	84Bar2
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$ +5 m**M** NaI/H_2O	$[Co(NH_3)_5]Cl^{2+/1+}$	24	*(1.5·10^{-7})*	PP	84Bar2
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	$[Co(NH_3)_5]Br^{2+/1+}$	24	*(2.0·10^{-2})*	PP	84Bar2
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$ +5 m**M** NaI/H_2O	$[Co(NH_3)_5]Cl^{2+/1+}$	24	*(4.0·10^{-8})*	PP	84Bar2
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	$[Co(NH_3)_5]$ $NCS^{2+/1+}$	24	*(4.0·10^{-4})*	PP	84Bar2
	5 m**M** $HClO_4$ +5 m**M** NaI/H_2O	$[Co(NH_3)_5]$ $NCS^{2+/1+}$	24	*(2.0·10^{-4})*	PP	84Bar2
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	*cis*-($[Co(NH_3)_4]$ Cl_2)$^{2+/1+}$	24	*(2.5·10^{-2})*	PP	84Bar2
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$ +5 m**M** NaI/H_2O	*cis*-($[Co(NH_3)_4]$ Cl_2)$^{2+/1+}$	24	*(1.0·10^{-5})*	PP	84Bar2
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	*cis*-$[Co(en)_2]Cl_2)^{2+/1+}$	24	*(3.0·10^{-4})*	PP	84Bar2 [63])
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$ +5 m**M** NaI/H_2O	*cis*-$[Co(en)_2]Cl_2)^{2+/1+}$	24	*(4.0·10^{-6})*	PP	84Bar2 [63])
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	*cis*-($[Co(en)_2]$ $(NCS)_2$)$^{2+/1+}$	24	*(1.0·10^{-4})*	PP	84Bar2 [63])
	0.1 **M** $NaClO_4$ +5 m**M** $HClO_4$ +5 m**M** NaI/H_2O	*cis*-($[Co(en)_2]$ $(NCS)_2$)$^{2+/1+}$	24	*(1.5·10^{-4})*	PP	84Bar2 [63])
	1 **M** KBr +0.1 **M** L/H_2O L=ethylenediamine	$CoL_3^{3+/2+}$	25	*0.62·10^{-2}*	CP, RDE	69Bar3, 70Bar5

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Pt	1 **M** KCl +0.1 **M** L/H_2O L=ethylenediamine	$CoL_3^{3+/2+}$	25	*2.1·10^{-2}*	CP, RDE	69Bar3, 70Bar5
	1 **M** KCl +0.1 **M** L/H_2O L=ethylenediamine	$CoL_3^{3+/2+}$, 1 m**M**	25	2.6	RDE	70Bar4
	1 **M** $KClO_4$ +0.1 **M** L/H_2O L=ethylenediamine	$CoL_3^{3+/2+}$	25	*2.9·10^{-2}*	CP	69Bar3
				2.9·10^{-2}	RDE	70Bar5
	1 **N** K_2SO_4 +0.1 **M** L/H_2O L=ethylenediamine	$CoL_3^{3+/2+}$	25	*1.64·10^{-2}*	CP	69Bar3
			25	*1.64·10^{-2}*	RDE	70Bar5
	1 **M** $K_2S_2O_3$ +0.1 **M** L/H_2O L=ethylenediamine	$CoL_3^{3+/2+}$	25	*0.46·10^{-2}*	CP	69Bar3
			25	*0.46·10^{-2}*	RDE	70Bar5
	1 **M** KCl +0.1 **M** L/H_2O L=1,2-propandiamine	$CoL_3^{3+/2+}$	25	*1.4·10^{-2}*	CP	69Bar3
	1 **M** KCl +0.1 **M** L/H_2O L=1,2-propandiamine	$CoL_3^{3+/2+}$, 1 m**M**	25	1.35	RDE	70Bar4
	1 **M** KCl +0.1 **M** L/H_2O L=2,3-butandiamine	$CoL_3^{3+/2+}$	25	*1.04·10^{-2}*	CP	69Bar3
	1 **M** KCl +0.1 **M** L/H_2O L=2,3-butandiamine	$CoL_3^{3+/2+}$, 1 m**M**	25	*1.04·10^{-2}*	RDE	70Bar4
	1 **M** KCl+0.1 **M** L/H_2O L=1,2-cyclohexan-diamine	$CoL_3^{3+/2+}$	25	*0.65·10^{-2}*	CP	69Bar3;
			25	*7.5·10^{-4}*	RDE	70Bar3
	1 **M** KCl+4 m**M** L/H_2O L=1,10 phenanthroline	$CoL_3^{3+/2+}$	25	*4.8·10^{-2}*	CP	71Bar2
	1 **M** KCl +0.1 **M** L/H_2O L=diethylene-tramine	$CoL_3^{3+/2+}$	25	*7.7·10^{-2}*	CP	71Bar3
	0.1 **M** $NaClO_4/H_2O$	$Cr_{aq}^{2+/3+}$	25	*(1.0·10^{-3})*	RDE	80Bar1
	0.5 **M** $NaClO_4/H_2O$	$Cr_{aq}^{2+/3+}$	25	*(6.0·10^{-3})*	RDE	80Bar1
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Cr(bipy)_3^{2+/+}$	25	*>0.5*	GD	75Saj1
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Cr(bipy)_3^{+/0}$	25	*0.88*	GD	75Saj1
	1 **N** HCl/H_2O	$Fe(H_2O)_6^{2+/3+}$	-	*0.105*	FR	74Aga

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.1 **N** HCl/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*4.4·10^{-3}*	RDE	72Stu
	0.1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	25	*0.011*	CF	62Jor
	0.1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{3+/2+}$	25	*0.0076*	CF	70Suz
	0.1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$, 0.01 **M**	-	0.001	CP,CH	70Bre
	0.1 **M** $HClO_4$+3 m**M** $Th(ClO_4)_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$, 0.01 **M**	-	0.0005	CP,CH	70Bre
	0.1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	25	*7.6·10^{-3}*	CF	70Suz
	0.1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	22	*4.8·10^{-3}*	CV	89Chu
	0.1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	25	34.3	RDE	74Mom[129])
	0.5 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	20	(9·10^{-6})	RDE	78Web
	1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	-	*1·10^{-2}*	FR	74Aga
	1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	25	0.22	RDE	62Jah
	1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	25	*0.005*	IP	52Ran2
	1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	20	*9·10^{-3}*	RDE	72Ang
	0.1 **N** HNO_3/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*2.9·10^{-3}*	RDE	72Stu
	0.01 **M** H_2SO_4+0.5 **M** K_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	*5·10^{-3}*	RDE	72Ang
	0.1 **N** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*1.3·10^{-3}*	RDE	72Stu
	0.1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*4.3·10^{-3}*	RDE	63Gal
	0.5 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$, 2 m**M**	25	*3.3·10^{-2}*	FR	63Aga1, 63Aga2
	0.5 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$, 2 m**M**	30	*3.8·10^{-2}*	FR	63Aga1, 63Aga2
	0.5 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$, 2 m**M**	35	*4.4·10^{-2}*	FR	63Aga1, 63Aga2
	0.5 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$, 2 m**M**	40	*5.0·10^{-2}*	FR	63Aga1, 63Aga2
	0.5 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	35	*4.8·10^{-2}*	FR	63Aga3
	0.5 **M** H_2SO_4/H_2O [130])	$Fe(H_2O)_6^{2+/3+}$	35	*4.1·10^{-2}*	FR	62Aga
	0.5 **M** H_2SO_4/H_2O [131])	$Fe(H_2O)_6^{2+/3+}$	35	*4.1·10^{-2}*	FR	62Aga
	0.5 **M** H_2SO_4+0.4 **M** acetone/H_2O	$Fe(H_2O)_6^{2+/3+}$	35	*5.5·10^{-2}*	FR	63Aga3

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Pt	0.5 **M** H_2SO_4 +0.002% cetyl-pyridinium bromide/H_2O	$Fe(H_2O)_6^{2+/3+}$	35	*1.2·10^{-2}*	FR	63Aga3
	1 **N** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	*7·10^{-3}*	RDE	72Ang
	1 **N** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	-	*5·10^{-2}*	FR	74Aga
	1 **F** [132]) H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	*0.0053*	GM	61Ans [133])
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	0.29	CF	50Ger
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	4.63·10^2	CPS	60Wij3
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	5.2·10^2	CCS	60Wij3
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*4.3·10^{-3}*	CP	63Ada
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*4.3·10^{-3}*	RDE	63Gal
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$, 0.1 **M**	25	0.0011	SS	68Boc2
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$, 10 m**M**	25	0.0025	SS	68Boc2
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$, 0.004 **M**		1.0·10^{-4}	RDE	96Tor
	2 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	0.3	CF	50Ger
	4 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*3.1·10^{-3}*	RDE	63Gal
	0.4 **M** K_2SO_4 +0.01 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	1.0	TPF	88Boc
	0.4 **M** K_2SO_4 +0.01 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	1.0	CS	88Boc
	0.4 **M** K_2SO_4 +0.01 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	2.75	TPF	88Boc [134])
	2 **M** K_2SO_4 +1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*1.5·10^{-3}*	MP	70Her
	0.87 **M** $NaHSO_4/H_2O$	$Fe(H_2O)_6^{3+/2+}$	25	0.38	PS	70Bar5
	1 **F** [132]) $NaHSO_4/H_2O$	$Fe(H_2O)_6^{2+/3+}$	20	*0.0027*	GM	61Ans [133])
	1 **F** [132]) $NaHSO_4$ +1 **F** [132]) Na_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	20	*0.0014*	GM	61Ans [133])
	2.8 **M** Na_2SO_4 +1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*1.4·10^{-3}*	RDE	63Gal

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	1 **M** KCl/H_2O	$Fe(H_2O)_6^{2+/3+}$, 0.01 **M**	25	5.0	RDE	62Jah
	0.1 **N** acetic acid/H_2O	$Fe(H_2O)_6^{2+/3+}$	25	*$2.0 \cdot 10^{-3}$*	RDE	72Stu
	0.1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	-	*1.4910^{-5}*	CP	70Kap
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	*0.09*	IP	52Ran2
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	$9.2 \cdot 10^3$	CPS	60Wij3
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	$12.5 \cdot 10^3$	CCS	60Wij3
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	7.7	CF	55Jor
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	*0.08*	CF	62Jor
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	*0.24*	GS	69Dau [135])
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$, 2 m**M**	25	0.03	CP	70Sch
	1 **M** KCl/H_2O	$Fe(CN)_6^{3-/4-}$	25	*0.028*	GS	69Dau [135])
	1 **M** KF/H_2O	$Fe(CN)_6^{3-/4-}$	20	*40.0*	TPF	83Iwa
	0.2 **M** KF/H_2O	$Fe(CN)_6^{3-/4-}$, 2 m**M**	25	*0.18*	GD	75Saj2
	0.2 **M** KF/H_2O	$Fe(CN)_6^{3-/4-}$, 1 m**M**	25	*0.15*	GD	75Saj2
	1 **M** KNO_3/H_2O	$Fe(CN)_6^{3-/4-}$	35	*$6.6 \cdot 10^{-2}$*	FR	63Aga4
	1 **M** $LiCl/H_2O$	$Fe(CN)_6^{3-/4-}$, 2 m**M**	25	$6.4 \cdot 10^{-4}$	CP	70Sch
	0.5 **M** K_2SO_4/H_2O	$Fe(CN)_6^{3-/4-}$	20	*34.0*	TPF	83Iwa
	0.5 **M** K_2SO_4/H_2O	$Fe(CN)_6^{3-/4-}$	25	*0.13*	IP	52Ran2
	0.5 **M** K_2SO_4/H_2O	$Fe(CN)_6^{3-/4-}$, 0.01 **M**	25	7.0	RDE	62Jah
	0.5 **M** K_2SO_4/H_2O	$Fe(CN)_6^{3-/4-}$	25	*$2 \cdot 10^2$*	RDE	72Ang
	0.5 **M** K_2SO_4/H_2O	$Fe(CN)_6^{3-/4-}$	20	*0.12*	TPF	78Ber
	1 **M** $NaCl/H_2O$	$Fe(CN)_6^{3-/4-}$, 2 m**M**	25	$2.0 \cdot 10^{-3}$	CP	70Sch
	1 **M** $NaNO_3/H_2O$	$Fe(CN)_6^{3-/4-}$, 4 m**M**	-	$2.66 \cdot 10^{-4}$	RDE	96Tor
	1 **M** NH_4Cl/H_2O	$Fe(CN)_6^{3-/4-}$, 2 m**M**	25	0.032	CP	70Sch

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Pt	0.2 **M** phosphate buffer/H_2O, pH=7.2	$Fe(CN)_6^{3-/4-}$	-	*0.1*	IP	85Dur
	1 **M** KCl/H_2O	$[Fe(CN)_5NH_3]^{2-/3-}$, 2 m**M**	25	0.016	CP	70Sch
	1 **M** KNO_3/H_2O	$[Fe(CN)_5NH_3]^{2-/3-}$, 2 m**M**	25	0.0088	CP	70Sch
	1 **M** LiCl/H_2O	$[Fe(CN)_5NH_3]^{2-/3-}$, 2 m**M**	25	0.0048	CP	70Sch
	1 **M** NaCl/H_2O	$[Fe(CN)_5NH_3]^{2-/3-}$, 2 m**M**	25	0.0064	CP	70Sch
	1 **M** NH_4Cl/H_2O	$[Fe(CN)_5NH_3]^{2-/3-}$, 2 m**M**	25	0.030	CP	70Sch
	1 **M** KCl/H_2O	$[Fe(CN)_5 MeNH_2]^{2-/3-}$, 2 m**M**	25	0.0024	CP	70Sch
	1 **M** KCl/H_2O	$[Fe(CN)_5 EtNH_2]^{2-/3-}$, 2 m**M**	25	0.0008	CP	70Sch
	1 **M** KCl/H_2O	$[Fe(CN)_5 PrNH_2]^{2-/3-}$, 2 m**M**	25	0.0006	CP	70Sch [85])
	1 **M** KCl/H_2O	$[Fe(CN)_5 BuNH_2]^{2-/3-}$, 2 m**M**	25	0.0012	CP	70Sch [136])
	1 **M** KCl/H_2O	$[Fe(CN)_5H_2O]^{2-/3-}$, 2 m**M**	25	0.0012	CP	70Sch
	1 **M** LiCl/H_2O	$[Fe(CN)_5H_2O]^{2-/3-}$, 2 m**M**	25	0.00032	CP	70Sch
	0.5 **M** K-oxalate/H_2O	$Fe(ox)_3^{4-/3-}$	25	*0.009*	IP	52Ran2
	0.2 **M** KF/H_2O	$Fe(bipy)_2(CN)_2^+$ / $Fe(bipy)_2(CN)_2$	25	*0.63*	GD	75Saj2
	0.2 **M** TBAP [52]) /DMF [54])	$Fe(bipy)_2(CN)_2^+$ /$Fe(bipy)_2(CN)_2$	25	*0.41*	GD	75Saj2
	0.2 **M** KF/H_2O	$Fe(bipy)(CN)_4^-$ / $Fe(bipy)(CN)_4^{2-}$	25	*0.43*	GD	75Saj2
	1 **M** LiF/H_2O	$Fe(bipy)_3^{2+/3+}$	20	*0.9*	TPF	80Ber [137])
	0.2 **F** KF/H_2O	$Fe(bipy)_3^{3+/2+}$	25	*0.8*	GD	75Saj2 [133])
	0.2 m**F** TBAP [52]) /DMF [54])	$Fe(bipy)_3^{3+/2+}$	25	*1.1*	GD	75Saj2 [133])
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Fe(bipy)_3^{3+/2+}$	25	*0.8*	GD	75Saj1
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Fe(bipy)_3^{2+/+}$	25	*0.16*	GD	75Saj1
	0.6 **M** TEAP [74]) /acetonitrile	$[Fe(cp)_2]$ (o)	-	*0.95*	CV	88Wip

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.1 **M** TBAP [52]) /acetonitrile	$[Fe(cp)_2]$ (o)	12	*2.6*	CV/IP	91Bar2
	0.1 **M** TBAP [52]) /acetone	$[Fe(cp)_2]$ (o)	12	*2.0*	CV/IP	91Bar2
	0.1 **M** TBAP [52]) /butyronitrile	$[Fe(cp)_2]$ (o)	12	*0.51*	CV/IP	91Bar2
	0.1 **M** TBAP [52]) /dichloromethane	$[Fe(cp)_2]$ (o)	12	*0.62*	CV/IP	91Bar2
	0.1 **M** TBAP [52]) /1,2-dichloroethane	$[Fe(cp)_2]$ (o)	12	*0.34*	CV/IP	91Bar2
	0.025 **M** TBAP [52]) /DMF [54])	$[Fe(cp)_2]$ (o)	-	*0.47*	ACP	93Win
	0.05 **M** TBAP [52]) /DMF [54])	$[Fe(cp)_2]$ (o)	-	*0.45*	ACP	93Win
	0.1 **M** TBAP [52]) /DMF [54])	$[Fe(cp)_2]$ (o)	12	*0.4*	CV/IP	91Bar2
	0.1 **M** TBAP [52]) /DMF [54])	$[Fe(cp)_2]$ (o)	-	*0.47*	ACP	93Win
	0.25 **M** TBAP [52]) /DMF [54])	$[Fe(cp)_2]$ (o)	-	*0.43*	ACP	93Win
	0.1 **M** TBAP [52]) /dimethyl-sulphoxide	$[Fe(cp)_2]$ (o)	12	*0.2*	CV/IP	91Bar2
	0.1 **M** TBAP [52]) /propylene carbonate	$[Fe(cp)_2]$ (o)	12	*0.15*	CV/IP	91Bar2
	0.025 **M** TBAP [52]) /PC [113])	$[Fe(cp)_2]$ (o)	-	*0.2*	ACP	93Win
	0.05 **M** TBAP [52]) /PC [113])	$[Fe(cp)_2]$ (o)	-	*0.21*	ACP	93Win
	0.1 **M** TBAP [52]) /PC [113])	$[Fe(cp)_2]$ (o)	-	*0.19*	ACP	93Win
	0.25 **M** TBAP [52]) /PC[113])	$[Fe(cp)_2]$ (o)	-	*0.21*	ACP	93Win
	0.1 **M** TBAP [52]) /tetrahydrofurane	$[Fe(cp)_2]$ (o)	12	*0.14*	CV/IP	91Bar2
	0.1 **M** $LiClO_4$ /methanol	$[Fe(cp)_2]$ (o)	12	*1.55*	CV/IP	91Bar2
	0.1 **M** $LiClO_4$/ethanol	$[Fe(cp)_2]$ (o)	12	*0.55*	CV/IP	91Bar2
	0.1 **M** $LiClO_4$ /1-propanol	$[Fe(cp)_2]$ (o)	12	*0.28*	CV/IP	91Bar2
	0.1 **M** HCl/H_2O	H^+/H_2	23	$2.0\cdot10^{-4}$	CP	64Boc3 [138])
	0.2 **M** HCl/H_2O	H^+/H_2	23	$1.0\cdot10^{-3}$	CP	64Boc3 [139])
	1 **M** HCl/H_2O	H^+/H_2	23	$1.8\cdot10^{-4}$	CP	64Boc3 [138])
	1 **M** HCl/H_2O	H^+/H_2	23	$3.0\cdot10^{-5}$	CP	64Boc3 [140])
	1 **M** HCl/H_2O	H^+/H_2	23	$1.0\cdot10^{-3}$	CP	64Boc3 [139])
	1.5 **M** HCl/H_2O	H^+/H_2	23	$0.9\cdot10^{-3}$	CP	64Boc3 [139])

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Pt	1.5 **M** HCl/H_2O	H^+/H_2	23	$4.5\cdot10^{-4}$	CP	64Boc3 [138])
	1.5 **M** HCl/H_2O	H^+/H_2	23	$2.0\cdot10^{-5}$	CP	64Boc3 [140])
	2.5 **M** HCl/H_2O	H^+/H_2	23	$0.8\cdot10^{-3}$	CP	64Boc3 [139])
	0.1 **M** H_2SO_4/H_2O	H^+/H_2	23	$1.0\cdot10^{-3}$	CP	64Boc3 [139])
	0.1 **M** H_2SO_4+0.05 **M** K_2SO_4/H_2O	H^+/H_2	23	$0.7\cdot10^{-3}$	CP	64Boc3 [139])
	0.5 **N** H_2SO_4/H_2O	H^+/H_2	-	$10^{-3.53}$	CP	64Boc3
	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	10^{-3}	CP	67Man
	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-2.9}$	CP	69Gal
	1 **M** H_2SO_4/H_2O, $p_{H_2} = 1$ bar	H^+/H_2	25	$1.52\cdot10^{-3}$	CP	59Sch1
	1 **M** H_2SO_4/H_2O, $p_{H_2} = 0.46$ bar	H^+/H_2	25	$0.7\cdot10^{-3}$	CP	59Sch1
	1 **M** H_2SO_4/H_2O, $p_{H_2} = 0.31$ bar	H^+/H_2	25	$0.4\cdot10^{-3}$	CP	59Sch1
	1 **M** H_2SO_4/H_2O, $p_{H_2} = 0.24$ bar	H^+/H_2	25	$0.273\cdot10^{-3}$	CP	59Sch1
	1 **M** H_2SO_4/H_2O, $p_{H_2} = 0.14$ bar	H^+/H_2	25	$0.112\cdot10^{-3}$	CP	59Sch1
	1 **M** H_2SO_4/H_2O, $p_{H_2} = 0.05$ bar	H^+/H_2	25	$59.4\cdot10^{-6}$	CP	59Sch1
	0.5 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-4.06}$	CP	64Boc3
	1 **M** KCl/H_2O	$I^{0/-1}$	25	*0.015* [141])	CF	62Jor
	9.3 **M** KOH/H_2O	$MnO_4^{2-/3-}$	10	0.6	CP,GS	67Thi2
	10.0 **M** KOH/H_2O	$MnO_4^{2-/3-}$	10	0.7	CP,GS	67Thi2
	10.8 **M** KOH/H_2O	$MnO_4^{2-/3-}$	10	0.6	CP,GS	67Thi2
	11.8 **M** KOH/H_2O	$MnO_4^{2-/3-}$	10	0.5	CP,GS	67Thi2
	1 **M** KOH/H_2O	$MnO_4^{-/2-}$	20	$1.2\cdot10^{-2}$	CP,GS	67Thi1
	1 **M** KOD/D_2O (96 %)	$MnO_4^{-/2-}$, 6 m**M**	20	0.0072	CP,GS	67Thi1
	1 **M** KOH/H_2O	$MnO_4^{-/2-}$, 6 m**M**	20	0.0069	CP,GS	67Thi1
	1 **M** KOH +0.5 **M** K_2CO_3/H_2O	$MnO_4^{-/2-}$, 6 m**M**	20	0.0032	CP,GS	67Thi1
	1 **M** KOH +0.3 **M** $K_2C_2O_4/H_2O$	$MnO_4^{-/2-}$, 6 m**M**	20	0.0043	CP,GS	67Thi1
	1 **M** KOH +0.5 **M** KCl/H_2O	$MnO_4^{-/2-}$, 6 m**M**	20	0.0078	CP,GS	67Thi1
	1 **M** KOH +0.3 **M** KNO_3/H_2O	$MnO_4^{-/2-}$, 6 m**M**	20	0.0050	CP,GS	67Thi1

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	1 **M** KOH +0.2 **M** NaF/H_2O	$MnO_4^{-/2-}$, 6 m**M**	20	0.0056	CP,GS	67Thi1
	1 **M** KOH +1 m**M** H_6TeO_6/H_2O	$MnO_4^{-/2-}$, 6 m**M**	20	0.0027	CP,GS	67Thi1
	1 **M** KOH+1 m**M** $Zn(OH)_4^{2-}/H_2O$	$MnO_4^{-/2-}$, 6 m**M**	20	0.0019	CP,GS	67Thi1
	1 **M** $LiOH/H_2O$	$MnO_4^{-/2-}$, 6 m**M**	20	$5.5\cdot10^{-3}$	CP,GS	67Thi1
	1 **M** $NaOH/H_2O$	$MnO_4^{-/2-}$, 6 m**M**	20	$6.2\cdot10^{-3}$	CP,GS	67Thi1
	0.1 **M** TBAP /acetonitrile	$[Ni(cp)_2]$ (o)	22	*2.0*	ACP	96Win
	0.1 **M** TBAP /butyronitrile	$[Ni(cp)_2]$ (o)	22	*0.52*	ACP	96Win
	0.1 **M** TBAP [52]) /acetone	$[Ni(cp)_2]$ (o)	22	*1.8*	ACP	96Win
	0.1 **M** TBAP [52]) /dichloromethane	$[Ni(cp)_2]$ (o)	22	*1.4*	ACP	96Win
	0.1 **M** TBAP [52]) /dichloroethane	$[Ni(cp)_2]$ (o)	22	*0.8*	ACP	96Win
	0.1 **M** TBAP [52]) /DMF [54])	$[Ni(cp)_2]$ (o)	22	*0.54*	ACP	96Win
	0.1 **M** TBAP [54]) /dimethyl sulfoxide	$[Ni(cp)_2]$ (o)	22	*0.2*	ACP	96Win
	0.1 **M** TBAP [52]) /PC [113])	$[Ni(cp)_2]$ (o)	22	*0.12*	ACP	96Win
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Os(bipy)_3^{3+/2+}$	25	*0.83*	GD	75Saj1
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Os(bipy)_3^{2+/+}$	25	*0.26*	GD	75Saj1
	1 **M** HBr/H_2O	$PtBr_6^{2-}/PtBr_4^{2-}$, 1 m**M**	-	$6.3\cdot10^{-6}$	TLV	70Hub
	1 **M** NaBr +3 **M** H_2SO_4/H_2O	$PtBr_6^{2-}/PtBr_4^{2-}$, 5 m**M**	-	$1.9\cdot10^{-6}$	TLV	70Hub
	1 **M** HCl/H_2O	$PtCl_6^{2-}/PtCl_4^{2-}$, 1 m**M**	-	$2.3\cdot10^{-7}$	TLV	70Hub
	1 **M** NaCl +3 **M** H_2SO_4/H_2O	$PtCl_6^{2-}/PtCl_4^{2-}$, 5 m**M**	-	$2.1\cdot10^{-7}$	TLV	70Hub
	10 m**F** [132])$HClO_4$ +x $NaClO_4/H_2O$	$PtCl_6^{2-}/PtCl_4^{2-}$	24	$6.1\cdot10^{-6}$	TLV	70Lau [133])
	10 m**F** [132])$HClO_4$ +x $NaClO_4$ +10 m**F** [132]) $NaCl/H_2O$	$PtCl_6^{2-}/PtCl_4^{2-}$	24	$3.5\cdot10^{-6}$	TLV	70Lau [133])

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Pt	10 m**F** [132])$HClO_4$ +x $NaClO_4$+ 100 m**F** [132]) NaCl/H_2O	$PtCl_6^{2-}/PtCl_4^{2-}$	24	*1.5·10^{-6}*	TLV	70Lau [133])
	10 m**F** [132])$HClO_4$ +x $NaClO_4$ +300 m**F** [132]) NaCl/H_2O	$PtCl_6^{2-}/PtCl_4^{2-}$	24	*5.9·10^{-7}*	TLV	70Lau [133])
	10 m**F** [132])$HClO_4$ +x $NaClO_4$ +1 m**F** [132]) NaCl/H_2O	$PtCl_6^{2-}/PtCl_4^{2-}$	24	*5.7·10^{-7}*	TLV	70Lau [133])
	10 m**F** [132])$HClO_4$ +x $NaClO_4/H_2O$	$PtCl_4^{2-}/PtCl_6^{2-}$	24	*8.1·10^{-6}*	TLV	70Lau [133])
	10 m**F** [132])$HClO_4$ +x $NaClO_4$ +10 m**F** [132]) NaCl/H_2O	$PtCl_4^{2-}/PtCl_6^{2-}$	24	*3.4·10^{-6}*	TLV	70Lau [133])
	10 m**F** [132])$HClO_4$ +x $NaClO_4$ +100 m**F** [132]) NaCl/H_2O	$PtCl_4^{2-}/PtCl_6^{2-}$	24	*1.6·10^{-6}*	TLV	70Lau [133])
	10 m**F** [132])$HClO_4$ +x $NaClO_4$ +300 m**F** [132]) NaCl/H_2O	$PtCl_4^{2-}/PtCl_6^{2-}$	24	*≤10^{-9}*	TLV	70Lau [133])
	10 m**F** [132])$HClO_4$ +x $NaClO_4$ +1000 m**F** [132]) NaCl/H_2O	$PtCl_4^{2-}/PtCl_6^{2-}$	24	*≤10^{-9}*	TLV	70Lau [133])
Pt$_{mod}$ [142])	1 **F** [132])NaCl+10 m**M** $HClO_4/H_2O$	$PtCl_6^{2-}/PtCl_4^{2-}$	23	*(1.66)* [143])	TLV	71Lau
Pt$_{mod}$ [144])	1 **F** [132])NaCl+10 m**M** $HClO_4/H_2O$	$PtCl_6^{2-}/PtCl_4^{2-}$	23	*(21.11)* [143])	TLV	71Lau
Pt$_{mod}$ [145])	1 **F** [132])NaCl+10 m**M** $HClO_4/H_2O$	$PtCl_6^{2-}/PtCl_4^{2-}$	23	*(21.25)* [143])	TLV	71Lau
Pt$_{mod}$ [146])	1 **F** [132])NaCl+10 m**M** $HClO_4/H_2O$	$PtCl_6^{2-}/PtCl_4^{2-}$	23	*(20.86)* [143])	TLV	71Lau
Pt$_{mod}$ [147])	1 **F** [132])NaCl+10 m**M** $HClO_4/H_2O$	$PtCl_6^{2-}/PtCl_4^{2-}$	23	*(4.59)* [143])	TLV	71Lau
	1 **M** HI/H_2O	PtI_6^{2-}/PtI_4^{2-}, 1 m**M**	-	*7.2·10^{-5}*	TLV	70Hub
	1 **M** NaI +3 **M** H_2SO_4/H_2O	PtI_6^{2-}/PtI_4^{2-}, 5 m**M**	-	*4.5·10^{-4}*	TLV	70Hub

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	1 **F** $HClO_4$ +10 m**F** [132]) NaCl/H_2O	$Pt(CN)_4^{2-}$ /$Pt(CN)_4(H_2O)Cl$	24	*1.05·10^{-7}*	TLV	72Lai [133])
Pt$_{mod}$ [148])	1 **F** [132]) NaCl+10 m**M** $HClO_4$/H_2O	$Pt(NH_3)_5Cl^{3+}$ / $Pt(NH_3)_4^{2+}$	23	*(1.21)* [143])	TLV	71Lau
Pt$_{mod}$ [144])	1 **F** [132]) NaCl+10 m**M** $HClO_4$/H_2O	$Pt(NH_3)_5Cl^{3+}$ / $Pt(NH_3)_4^{2+}$	23	*(1.02)* [143])	TLV	71Lau
Pt$_{mod}$ [145])	1 **F** [132]) NaCl+10 m**M** $HClO_4$/H_2O	$Pt(NH_3)_5Cl^{3+}$ / $Pt(NH_3)_4^{2+}$	23	*(0.98)* [143])	TLV	71Lau
Pt$_{mod}$ [146])	1 **F** [132]) NaCl+10 m**M** $HClO_4$/H_2O	$Pt(NH_3)_5Cl^{3+}$ / $Pt(NH_3)_4^{2+}$	23	*(1.06)* [143])	TLV	71Lau
Pt$_{mod}$ [147])	1 **F** [132]) NaCl+10 m**M** $HClO_4$/H_2O	$Pt(NH_3)_5Cl^{3+}$ / $Pt(NH_3)_4^{2+}$	23	*(1.11)* [143])	TLV	71Lau
	1 **F** $HClO_4$ +10 m**F** [132]) NaCl/H_2O	*trans*-$Pt(NH_3)_2(CN)_2$ /*trans*-$Pt(NH_3)_2(CN)_2(H_2O)Cl^+$	24	*1.93·10^{-6}*	TLV	72Lai [133])
	1 **F** [132]) NaCl/H_2O	*trans*-$Pt(NH_3)_2(CN)_2$ /*trans*-$Pt(NH_3)_2(CN)_2Cl_2$	24	*7.6·10^{-9}*	TLV	72Lai [133])
Pt$_{mod}$ [148])	1 **F** [132]) NaCl+10 m**M** $HClO_4$/H_2O	$Pt(NH_3)_2(NO_2)_2Cl_2$ /$Pt(NH_3)_2(NO_2)_2$	23	*(1.03)* [143])	TLV	71Lau
Pt$_{mod}$ [144])	1 **F** [132]) NaCl+10 m**M** $HClO_4$/H_2O	$Pt(NH_3)_2(NO_2)_2Cl_2$ /$Pt(NH_3)_2(NO_2)_2$	23	*(1.13)* [143])	TLV	71Lau
Pt$_{mod}$ [145])	1 **F** [132]) NaCl+10 m**M** $HClO_4$/H_2O	$Pt(NH_3)_2(NO_2)_2Cl_2$ /$Pt(NH_3)_2(NO_2)_2$	23	*(1.54)* [143])	TLV	71Lau
Pt$_{mod}$ [146])	1 **F** [132]) NaCl+10 m**M** $HClO_4$/H_2O	$Pt(NH_3)_2(NO_2)_2Cl_2$ /$Pt(NH_3)_2(NO_2)_2$	23	*(1.34)* [143])	TLV	71Lau
Pt$_{mod}$ [147])	1 **F** [132]) NaCl+10 m**M** $HClO_4$/H_2O	$Pt(NH_3)_2(NO_2)_2Cl_2$/ $Pt(NH_3)_2(NO_2)_2$	23	*(1.01)* [143])	TLV	71Lau
	1 **F** [132]) $HClO_4$ +10 m**F** NaCl/H_2O	$Pt(en)(CN)_2$ /$Pt(en)(CN)_2(H_2O)Cl^+$	24	*2.04·10^{-7}*	TLV	72Lai [149])
	1 **F** [132]) NaCl/H_2O	$Pt(en)(CN)_2$ /$Pt(en)(CN)_2Cl_2$	24	*3.58·10^{-11}*	TLV	72Lai [133])
	1 **F** [132]) $HClO_4$ +10 m**F** NaCl/H_2O	*trans*-$Pt(NH_3)_2(SCN)_2$ /*trans*-$Pt(NH_3)_2(SCN)_2Cl_2$	24	*3.01·10^{-7}*	TLV	72Lai [133])
	1 **F** [132]) $HClO_4$ +10 m**F** NaCl/H_2O	$Pt(NO_2)_4^{2-}$ /$Pt(NO_2)_4(H_2O)Cl^-$	24	*4.87·10^{-7}*	TLV	72Lai [133])
	1 **F** [132]) $HClO_4$ +1 **F** NaCl/H_2O	*cis*-$Pt(NH_3)_2(NO_2)_2$ /$Pt(NH_3)_2(NO_2)_2$	23	*2.0·10^{-6}*	TLV	69Cus [133])
	1 **F** [132]) $HClO_4$ +10 m**F** NaCl/H_2O	*cis*-$Pt(NH_3)_2(NO_2)_2$ /*trans*-$Pt(NH_3)_2(NO_2)_2(H_2O)Cl^+$	24	*3.17·10^{-7}*	TLV	72Lai [133])

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Pt$_{mod}$ [147])	1 **F** [132]) $HClO_4$ +10 m**F** NaCl/H_2O	*cis*-$Pt(NH_3)_2(NO_2)_2$ /*trans*-$Pt(NH_3)_2(NO_2)_2Cl_2$	24	*3.17·10^{-7}*	TLV	72Lai [133])
	1 **F** [132]) $HClO_4$ +10 m**F** NaCl/H_2O	$Pt(C_2H_4)Cl_2(H_2O)$ /$Pt(C_2H_4)Cl_4(H_2O)$	24	*8.59·10^{-8}*	TLV	72Lai [133])
	1 **F** [132]) H_2SO_4 +10 m**F** NaCl/H_2O	$Pt(py)_4^{2+}$ / $Pt(py)_4Cl_2^{2+}$	24	*5.4·10^{-6}*	TLV	72Lai [133])
	1 **F** [132]) $HClO_4$ +10 m**F** NaCl/H_2O	$Pt(CN)_4Cl_2^{2-}$ / $Pt(CN)_4^{2-}$	24	*8.16·10^{-7}*	TLV	72Lai [133])
	1 **F** [132]) $HClO_4$	$Pt(SCN)_6^{2-}$ / $Pt(SCN)_4^{2-}$	24	*6.6·10^{-5}*	TLV	72Lai [133])
	1 **F** [132]) $HClO_4$ +10 m**F** NaCl/H_2O	$Pt(dien)(NO_2)^+$ /$Pt(dien)(NO_2)Cl_2^{2+}$	24	*2.0·10^{-5}*	TLV	69Cus [150])
	1 **F** [132]) $HClO_4$ +10 m**F** NaCl/H_2O	$Pt(en)(SCN)Cl^{2+}$ / $Pt(en)_2^{2+}$	24	*4.64·10^{-5}*	TLV	72Lai [133])
	1 **F** [132]) $HClO_4$ +2 m**F** NaCl/H_2O	$Pt(NO_2)_4Cl_2^{2-}$ / $Pt(NO_2)_4^{2-}$	24	*9.26·10^{-6}*	TLV	72Lai [133])
	1 **F** [132]) $HClO_4$ +10 m**F** NaCl/H_2O	$Pt(NH_3)_5Cl^{3+}$ / $Pt(NH_3)_4^{2+}$	24	*4.9·10^{-5}*	TLV	72Lai [133])
	0.2 **M** KPF_6/H_2O	$Ru_{aq}^{2+/3+}$	25	*(>5.0·10^{-2})*	RDE	80Bar1
	1 **M** KF/H_2O	$[Ru(NH_3)_6]^{3+/2+}$	20	116	TPF	85Iwa
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Ru(bipy)_3^{2+/+}$	25	*0.24*	GD	75Saj1
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Ru(bipy)_3^{+/0}$	25	*0.24*	GD	75Saj1
	0.2 **M** Bu_4NClO_2 /DMF [54])	$Ru(bipy)_3^{0/-}$	25	*0.26*	GD	75Saj1
	1 m**M** TBAP/CH_2Cl_2	bis(alkynyl) titanocene [151])	RT	*5.0·10^{-4}*	CV	02Ste
	0.2 **M** $NaClO_4/H_2O$	$V_{aq}^{2+/3+}$	25	*(>0.1)*	RDE	80Bar1
	0.2 **M** $NaClO_4$ +1 m**M** Cr^{2+}/H_2O	$V_{aq}^{2+/3+}$	25	*(1.5·10^{-5})*	RDE	80Bar1
	0.1 **M** TEAP [74]) /DMF [54])	benzoquinone (r)	25	*0.0052*	CV	73Cap
	HCl-buffer/H_2O, pH=1.1	hydroquinone /quinone, 5.5 m**M**	25	7.8·10^{-5}	CF	62Tom [77])
	HCl-buffer/H_2O, pH=1.1	hydroquinone /quinone, 5.5 m**M**	25	1.0·10^{-4}	CF	62Tom
	acetate buffer/H_2O, pH=3.0	hydroquinone /quinone, 5.5 m**M**	25	7.8·10^{-6}	CF	62Tom

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	acetate buffer/H_2O, pH=5.0	hydroquinone /quinone, 5.5 m**M**	25	$2.0\cdot10^{-5}$	CF	62Tom
	0.1 **M** HCl/H_2O	hydroquinone /quinone, 10 m**M**	5	$3.7\cdot10^{-7}$	CF	60Los1
	0.1 **M** HCl/H_2O	hydroquinone /quinone, 10 m**M**	15	$9.1\cdot10^{-7}$	CF	60Los1
	0.1 **M** HCl/H_2O	hydroquinone /quinone, 10 m**M**	25	$1.9\cdot10^{-6}$	CF	60Los1
	0.1 **M** HCl/H_2O	hydroquinone /quinone, 5 m**M**	25	$1.12\cdot10^{-6}$	CF	60Los1
	0.1 **M** HCl/H_2O	hydroquinone /quinone, 7.6 m**M**	25	1.3810^{-6}	CF	60Los1
	0.1 **M** HCl/H_2O	hydroquinone /quinone, 15.3 m**M**	25	$1.03\cdot10^{-6}$	CF	60Los1
	0.1 **M** HCl/H_2O	hydroquinone /quinone, 10 m**M**	35	$4.4\cdot10^{-6}$	CF	60Los1
	0.1 **M** HCl/H_2O	hydroquinone /quinone, 10 m**M**	45	$8.7\cdot10^{-6}$	CF	60Los1
	0.1 **M** HCl/H_2O	hydroquinone /quinone, 10 m**M**	60	$1.6\cdot10^{-6}$	CF	60Los1
	HCl/H_2O; pH=0.48	hydroquinone /quinone, 5.5 m**M**	25	$8.5\cdot10^{-5}$	CF	62Los1
	HCl/H_2O; pH=1.07	hydroquinone /quinone, 5.5 m**M**	25	$4.9\cdot10^{-5}$	CF	62Los1
	HCl/H_2O; pH=1.52	hydroquinone /quinone, 5.5 m**M**	25	$2.57\cdot10^{-5}$	CF	62Los1
	HCl/H_2O; pH=2.01	hydroquinone /quinone, 5.5 m**M**	25	$1.1\cdot10^{-5}$	CF	62Los1
	acetate buffer/H_2O, pH=2.03	hydroquinone /quinone, 5.5 m**M**	25	$1.0\cdot10^{-5}$	CF	62Los1
	acetate buffer/H_2O, pH=3.01	hydroquinone /quinone, 5.5 m**M**	25	$4.67\cdot10^{-6}$	CF	62Los1
	acetate buffer/H_2O, pH=3.92	hydroquinone /quinone, 5.5 m**M**	25	$5.58\cdot10^{-6}$	CF	62Los1
	acetate buffer/H_2O, pH=5.06	hydroquinone /quinone, 5.5 m**M**	25	$1.26\cdot10^{-5}$	CF	62Los1
	phosphate buffer/H_2O, pH=6.02	hydroquinone /quinone, 5.5 m**M**	25	$3.55\cdot10^{-5}$	CF	62Los1
	phosphate buffer /H_2O, pH=7.06	hydroquinone /quinone, 5.5 m**M**	25	$1.29\cdot10^{-4}$	CF	62Los1
	phosphate+borate buffer/H_2O, pH=7.92	hydroquinone /quinone, 5.5 m**M**	25	$4.26\cdot10^{-4}$	CF	62Los1
	borate buffer /H_2O, pH=8.6	hydroquinone /quinone, 5.5 m**M**	25	$3.1\cdot10^{-3}$	CF	62Los1

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Pt$_{mod}$ [147])	0.1 **M** HCl/H_2O	quinhydrone, 10 m**M**	25	$2.34\cdot10^{-6}$	CF	60Los1
	HCl/H_2O, pH=0.44	Na-β-hydroanthraquinonesulfonate /β-anthraquinonesulfonate, 1.8 m**M**	25	$1.51\cdot10^{-4}$	CF	62Los1
	H_2SO_4/H_2O, pH=2.05	Na-β-hydroanthraquinonesulfonate /β-anthraquinonesulfonate, 1.8 m**M**	25	$5.5\cdot10^{-5}$	CF	62Los1
	acetate buffer/H_2O, pH=3.06	Na-β-hydroanthraquinonesulfonate /β-anthraquinonesulfonate, 3.6 m**M**	25	$2.81\cdot10^{-5}$	CF	62Los1
	acetate buffer/H_2O, pH=4.1	Na-β-hydroanthraquinonesulfonate /β-anthraquinonesulfonate, 1.8 m**M**	25	$5.5\cdot10^{-5}$	CF	62Los1
	phosphate buffer /H_2O, pH=6.18	Na-β-hydroanthraquinonesulfonate /β-anthraquinonesulfonate, 1.8 m**M**	25	$1.21\cdot10^{-4}$	CF	62Los1
	phosphate buffer /H_2O, pH=8.12	Na-β-hydroanthraquinonesulfonate /β-anthraquinonesulfonate, 1.8 m**M**	25	$2.24\cdot10^{-4}$	CF	62Los1
	0.05...0.5 **M** $NaClO_4$ /acetonitrile	1,4-phenylenediamine (o)	25	*0.22*	CV	86Opa
	0.05...0.3 **M** $NaClO_4$ /DMF [54])	1,4-phenylenediamine (o)	25	*0.082*	CV	86Opa
	0.05...0.3 **M** $NaClO_4$ /DMSO [51])	1,4-phenylenediamine (o)	25	*0.074*	CV	86Opa
	0.05...0.5 **M** $NaClO_4$ /PC [113])	1,4-phenylenediamine (o)	25	*0.055*	CV	86Opa
	0.1...0.3 **M** $NaClO_4$ /THF [152])	1,4-phenylenediamine (o)	25	*0.074*	CV	86Opa
	0.1 **M** $NaClO_4$ /nitrobenzene	1,4-phenylenediamine (o)	25	*0.035*	CV	86Opa
	0.1...0.3 **M** $NaClO_4$ /sulfolane	1,4-phenylenediamine (o)	25	*0.022*	CV	86Opa
	0.1...0.3 **M** $NaClO_4$ /HMPA [124])	1,4-phenylenediamine (o)	25	*0.017*	CV	86Opa
	0.05...0.3 **M** $NaClO_4$ /NMF [88])	1,4-phenylenediamine (o)	25	*0.1*	CV	86Opa
	0.05...0.3 **M** $NaClO_4$ /methanol	1,4-phenylenediamine (o)	25	*0.18*	CV	86Opa

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.05...0.1 **M** $NaClO_4$ /2-propanol	1,4-phenylene-diamine (o)	25	*0.017*	CV	86Opa
	0.1 **M** $NaClO_4$ /1-butanol	1,4-phenylene-diamine (o)	25	*0.027*	CV	86Opa
	0.05...0.5 **M** KPF_6 /acetonitrile	phenothiazine (o)	25	*0.44*	CV	85Opa
	0.05...0.3 **M** KPF_6 /acetone	phenothiazine (o)	25	*0.22*	CV	85Opa
	0.1...0.5 **M** TBAP /CH_2Cl_2	phenothiazine (o)	25	*0.19*	CV	85Opa
	0.1...0.3 **M** KPF_6 /nitroethane	phenothiazine (o)	25	*0.19*	CV	85Opa
	0.05...0.5 **M** KPF_6 /DMF [54])	phenothiazine (o)	25	*0.1*	CV	85Opa
	0.05...0.3 **M** KPF_6 /DMSO [51])	phenothiazine (o)	25	*0.077*	CV	85Opa
	0.05...0.5 **M** KPF_6 /PC [113])	phenothiazine (o)	25	*0.44*	CV	85Opa
	0.1...0.5 **M** $NaClO_4$ /THF [152])	phenothiazine (o)	25	*0.046*	CV	85Opa
	0.1...0.5 **M** KPF_6 /sulfolane	phenothiazine (o)	25	*0.2*	CV	85Opa
	0.1...0.5 **M** KPF_6 /HMPT [124])	phenothiazine (o)	25	*0.014*	CV	85Opa
	0.05...0.3 **M** $NaClO_4$ /NMF [88])	phenothiazine (o)	25	*0.2*	CV	85Opa
	0.05...0.3 **M** $NaClO_4$ /methanol	phenothiazine (o)	25	*0.4*	CV	85Opa
	0.05...0.2 **M** $NaClO_4$ /ethanol	phenothiazine (o)	25	*0.22*	CV	85Opa
	0.05...0.3 **M** $NaClO_4$ /2-propanol	phenothiazine (o)	25	*0.045*	CV	85Opa
	0.1 **M** $NaClO_4$ /1-butanol	phenothiazine (o)	25	*0.062*	CV	85Opa
	0.1 **M** $NaClO_4$/acetone	tetrathiafulvalene$^{0/+}$	25	*1.6*	CV	90Gra
	0.1 **M** $NaClO_4$/aceto-nitrile	tetrathiafulvalene$^{0/+}$	25	*2.2*	CV	90Gra
	0.1 **M** $NaClO_4$/benzo-nitrile	tetrathiafulvalene$^{0/+}$	25	*0.03*	CV	90Gra
	0.1 **M** $NaClO_4$ /propionitrile	tetrathiafulvalene$^{0/+}$	25	*0.26*	CV	90Gra
	0.1 **M** $NaClO_4$/DMF [54])	tetrathiafulvalene$^{0/+}$	25	*0.31*	CV	90Gra
	0.1 **M** $NaClO_4$ /HMPA [124])	tetrathiafulvalene$^{0/+}$	25	*0.031*	CV	90Gra
	0.1 **M** $NaClO_4$/ NMP [153])	tetrathiafulvalene$^{0/+}$	25	*0.18*	CV	90Gra

Electrode [1]	Solution [2]	Educt, product, concentration [3]	T [4] [°C]	j_0 , j_{00} k_{app} or k_c^θ [5] [A cm^{-2} or cm s^{-1}]	Method [174]	Ref.
Pt$_{mod}$ [147]	0.1 **M** $NaClO_4$/PC [113]	tetrathiafulvalene$^{0/+}$	25	*0.087*	CV	90Gra
Pt(100)	0.1 **M** $KClO_4/H_2O$	$Fe(CN)_6^{3-/4-}$, 0.2 m**M**	25	*0.17*	IP	98Kit
	0.1 **M** $KClO_4/H_2O$	$Fe(CN)_6^{3-/4-}$, 0.5 m**M**	25	*0.04*	IP	98Kit
	0.1 **M** $KClO_4/H_2O$	$Fe(CN)_6^{3-/4-}$, 1 m**M**	25	*0.034*	IP	98Kit
	0.05 **M** H_2SO_4/H_2O	H^+/H_2	1	$0.36\cdot10^{-3}$	CP	97Mar
	0.05 **M** H_2SO_4/H_2O	H^+/H_2	30	$0.6\cdot10^{-3}$	CP	97Mar
	0.05 **M** H_2SO_4/H_2O	H^+/H_2	60	$0.76\cdot10^{-3}$	CP	97Mar
	0.1 **M** $HClO_4/H_2O$	O_2/H_2O	-	$2.0...4.9\cdot10^{-2}$	RDE	97Per
	0.05 **M** H_2SO_4/H_2O	O_2/H_2O	-	$1.4...3.3\cdot10^{-2}$	RDE	97Per
Pt(110)	0.1 **M** $KClO_4/H_2O$	$Fe(CN)_6^{3-/4-}$, 0.2 m**M**	25	*0.14*	IP	98Kit
	0.1 **M** $KClO_4/H_2O$	$Fe(CN)_6^{3-/4-}$, 0.5 m**M**	25	*0.032*	IP	98Kit
	0.1 **M** $KClO_4/H_2O$	$Fe(CN)_6^{3-/4-}$, 1 m**M**	25	*0.036*	IP	98Kit
	0.05 **M** H_2SO_4/H_2O	H^+/H_2	1	$0.36\cdot10^{-3}$	CP	97Mar
	0.05 **M** H_2SO_4/H_2O	H^+/H_2	30	$0.98\cdot10^{-3}$	CP	97Mar
	0.05 **M** H_2SO_4/H_2O	H^+/H_2	60	$1.35\cdot10^{-3}$	CP	97Mar
	0.1 **M** $HClO_4/H_2O$	O_2/H_2O	-	$6.6...19\cdot10^{-2}$	RDE	97Per
	0.05 **M** H_2SO_4/H_2O	O_2/H_2O	-	$3.2...8.7\cdot10^{-2}$	RDE	97Per
Pt(111)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	23	*$(1\cdot10^{-2})$*	CV	91Cha
	0.05 **M** H_2SO_4/H_2O	H^+/H_2	1	$0.21\cdot10^{-3}$	CP	97Mar
	0.05 **M** H_2SO_4/H_2O	H^+/H_2	30	$0.45\cdot10^{-3}$	CP	97Mar
	0.05 **M** H_2SO_4/H_2O	H^+/H_2	60	$0.83\cdot10^{-3}$	CP	97Mar
Pt(111)/(I√7) [154]	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	23	*$(3\cdot10^{-2})$*	CV	91Cha
Pt(111)/(I√7)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	23	*$(2\cdot10^{-4})$*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	23	*$(8\cdot10^{-4})$*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	23	*$(1\cdot10^{-4})$*	CV	91Cha
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5F]^{3+/2+}$	23	*$(4\cdot10^{-2})$*	CV	91Cha
Pt(111)/(I√7)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5F]^{3+/2+}$	23	*$(7\cdot10^{-4})$*	CV	91Cha
Pt(111)/(I√7)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_5F]^{3+/2+}$	23	*$(7\cdot10^{-4})$*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5F]^{3+/2+}$	23	*$(2\cdot10^{-4})$*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_5F]^{3+/2+}$	23	*$(2\cdot10^{-4})$*	CV	91Cha
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5SO_4]^{1+/0}$	23	*$(1.8\cdot10^{-2})$*	CV	91Cha
Pt(111)/(I√7)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5SO_4]^{1+/0}$	23	*$(3.5\cdot10^{-2})$*	CV	91Cha

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Pt(111)/(I√7)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_5SO_4]^{1+/0}$	23	*(9·10^{-3})*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5SO_4]^{1+/0}$	23	*(8·10^{-3})*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_5SO_4]^{1+/0}$	23	*(2·10^{-3})*	CV	91Cha
Pt(111)/(I√7)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5H_2O]^{3+/2+}$	23	*(0.15)*	CV	91Cha
Pt(111)/(I√7)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_5H_2O]^{3+/2+}$	23	*(2·10^{-2})*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5H_2O]^{3+/2+}$	23	*(3·10^{-2})*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_5H_2O]^{3+/2+}$	23	*(8·10^{-3})*	CV	91Cha
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5Ac]^{2+/1+}$	23	*(1.8·10^{-2})*	CV	91Cha [155])
Pt(111)/(I√7)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5Ac]^{2+/1+}$	23	*(3.5·10^{-2})*	CV	91Cha
Pt(111)/(I√7)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_5Ac]^{2+/1+}$	23	*(9·10^{-3})*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5Ac]^{2+/1+}$	23	*(8·10^{-3})*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_5Ac]^{2+/1}$	23	*(2·10^{-3})*	CV	91Cha
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5PAc]^{2+/1+}$	23	*(5.5·10^{-2})*	CV	91Cha [156])
Pt(111)/(I√7)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5PAc]^{2+/1+}$	23	*(4·10^{-3})*	CV	91Cha
Pt(111)/(I√7)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_5PAc]^{2+/1+}$	23	*(4·10^{-3})*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5PAc]^{2+/1+}$	23	*(4·10^{-3})*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_5PAc]^{2+/1+}$	23	*(3.5·10^{-3})*	CV	91Cha
Pt(111)/(I√7)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5Pyc]^{2+/1+}$	23	*(3·10^{-3})*	CV	91Cha [157])
Pt(111)/(I√7)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_5Pyc]^{2+/1+}$	23	*(4.5·10^{-3})*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5Pyc]^{2+/1+}$	23	*(1.5·10^{-3})*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_5Pyc]^{2+/1+}$	23	*(1.5·10^{-3})*	CV	91Cha
Pt(111)/(I√7)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5Cpc]^{2+/1+}$	23	*(2.5·10^{-4})*	CV	91Cha [158])
Pt(111)/(I√7)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_5Cpc]^{2+/1+}$	23	*(5.5·10^{-4})*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5Cpc]^{2+/1+}$	23	*(7·10^{-5})*	CV	91Cha
Pt(111)/(I√3x3)	0.1 **M** $NaClO_4$+50 m**M** $La(ClO_4)_3/H_2O$	$[Co(NH_3)_5Cpc]^{2+/1+}$	23	*(1·10^{-4})*	CV	91Cha
	0.1 **M** $KClO_4/H_2O$	$Fe(CN)_6^{3-/4-}$, 0.5 m**M**	25.0	*0.43*	IP	98Kit
	0.1 **M** $KClO_4/H_2O$	$Fe(CN)_6^{3-/4-}$, 1 m**M**	25.0	*0.19*	IP	98Kit
	0.1 **M** $HClO_4/H_2O$	O_2/H_2O	-	3.8...13·10^{-2}	RDE	97Per
	0.05 **M** H_2SO_4/H_2O	O_2/H_2O	-	0.5...1.1·10^{-2}	RDE	97Per

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Potassium (K)						
(K)$_{(Hg)}$	1.0 **M** NMe_4OH/H_2O	K^+	20	*~0.1*	IP	52Ran1
	0.1 **M** NEt_4PF_6 /DMF [54])	K^+	25	*0.65*	IP/PP	80Bar3
	0.1 **M** NPr_4PF_6 [85]) /DMF [54])	K^+	25	*0.81*	IP/PP	80Bar3
	0.1 **M** $NBut_4PF_6$ /DMF [54])	K^+	25	*1.33*	IP/PP	80Bar3
Rhodium (Rh)	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{3+/2+}$, 0.1**M**	25	0.0046	SS	68Boc2
	1 **M** H_2SO_4/H_2O	$Fe(H_2O)_6^{3+/2+}$, 0.01**M**	25	0.0022	SS	68Boc2
	1 **M** $HClO_4/H_2O$	$Fe(H_2O)_6^{3+/2+}$	-	*$4.6\cdot10^{-4}$*	CP	78Eve [159])
	0.01 **N** HCl/H_2O	H^+/H_2	-	$10^{-3.8}$	CP	57Pen
	0.5 **N** H_2SO_4/H_2O	H^+/H_2	-	$10^{-3.22}$	CP	64Boc3
	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-3.6}$	CP	65Mat
	0.5 **M** H_2SO_4/H_2O	H^+/H_2	-	$10^{-3.5}$	CP	67Man
	0.01 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-4.85}$	CP	57Pen
	0.1 **M** TEAP [74]) /DMF [54])	benzoquinone (r)	25	*0.0037*	CV	73Cap
	HCl-buffer/H_2O, pH=1.1	hydroquinone /quinone, 5.5 m**M**	25	$1.5\cdot10^{-6}$	CF	62Tom [77])
	HCl-buffer/H_2O, pH=1.1	hydroquinone /quinone, 5.5 m**M**	25	$3.2\cdot10^{-4}$	CF	62Tom
	acetate buffer/H_2O, pH=3.0	hydroquinone /quinone, 5.5 m**M**	25	$3.5\cdot10^{-5}$	CF	62Tom
	acetate buffer/H_2O, pH=3.0	hydroquinone /quinone, 5.5 m**M**	25	$4.0\cdot10^{-7}$	CF	62Tom [77])
	acetate buffer/H_2O, pH=5.0	hydroquinone /quinone, 5.5 m**M**	25	$2.3\cdot10^{-6}$	CF	62Tom [77])
	acetate buffer/H_2O, pH=5.0	hydroquinone /quinone, 5.5 m**M**	25	$1.0\cdot10^{-4}$	CF	62Tom
	1 **M** $HClO_4/H_2O$	hydroquinone /quinone	-	*$2.4\cdot10^{-4}$*	CP	78Eve [160])
Rubidium (Rb)	0.1 **M** NEt_4PF_6 /DMF [54])	Rb^{1+}	25	*0.59*	IP/PP	80Bar3
	0.1 **M** NPr_4PF_6 [85]) /DMF [54])	Rb^{1+}	25	*0.82*	IP/PP	80Bar3
Ruthenium (Ru)	1 **M** H_2SO_4/H_2O	$Fe^{2+/3+}$, 0.1 **M**	20	0.00653	CP	73Gal
	1 **M** H_2SO_4/H_2O	$Fe^{2+/3+}$, 0.1 **M**	25	0.00755	CP	73Gal
	1 **M** H_2SO_4/H_2O	$Fe^{2+/3+}$, 0.1 **M**	30	0.00867	CP	73Gal
	1 **M** H_2SO_4/H_2O	$Fe^{2+/3+}$, 0.1 **M**	35	0.00966	CP	73Gal
	1 **M** H_2SO_4/H_2O	$Fe^{2+/3+}$, 0.1 **M**	40	0.0108	CP	73Gal

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Silver (Ag)	0.58 **m** $AgNO_3$/ H_2O	Ag^+	RT	0.011	IP	61Bro
	0.14 **M** CsCN +2.1 **M** CsCl/H_2O	$Ag^{+/0}$, 0.33 **M**	25	$10.9\cdot10^{-3}$	CP	68Nec
	1 **N** $HClO_4$/H_2O	$Ag^{+/0}$	20	24	GD	57Ger1
	$HClO_4$/H_2O	$Ag^{+/0}$, 0.2 **M**	20	0.1	GS	57Meh
	0.14 **M** KCN +2.1 **M** KCl/H_2O	$Ag^{+/0}$, 0.33 **M**	25	$7.6\cdot10^{-3}$	CP	68Nec
	1 **M** KNO_3 +5 **M** NH_3/H_2O	$Ag^{+/0}$, 0.1 **M**	-	0.3	IP	58Lor
	1 **M** KCl +2...3 **M** NH_3/H_2O	$Ag^{+/0}$, 0.1 **M**	-	0.5	IP	59Lor
	35% KOH/H_2O	Ag/Ag_2O, (r)	RT	$2\cdot10^{-3}$	CP	67Dir
	0.14 **M** NaCN +2.1 **M** NaCl/H_2O	$Ag^{+/0}$, 0.33 **M**	25	$1.8\cdot10^{-3}$	CP	68Nec
	1 **M** NaOH/H_2O	$Ag^{0/+}$	-	*$8.09\cdot10^{-6}$* [161])	CV	68Sto
	0.28 **M** NaOH/H_2O	Ag/Ag_2O (r)	23	*$1.4\cdot10^{-3}$*	GM	68Cla2, 71Ham2
	1 **M** NaOH/H_2O	Ag/Ag_2O (r)	0	*$1.0\cdot10^{-3}$*	GM	68Cla2, 71Ham2
	1 **M** NaOH/H_2O	Ag/Ag_2O (r)	23	*$2.8\cdot10^{-3}$*	GM	68Cla2, 71Ham2
	1 **M** NaOH/H_2O	Ag/Ag_2O (r)	40	*$3.0\cdot10^{-3}$*	GM	68Cla2, 71Ham2
	1 **M** NaOH/H_2O	Ag/Ag_2O (r)	60	*$6.5\cdot10^{-3}$*	GM	68Cla2, 71Ham2
	4.86 **M** NaOH/H_2O	Ag/Ag_2O (r)	23	*$7.2\cdot10^{-3}$*	GM	68Cla2, 71Ham2
	1 **M** NaOH/H_2O	Ag/Ag_2O (r)	-	k_{ox}=*1.1* [161])	CV	68Sto
	1 **M** NaOH/H_2O	Ag/Ag_2O (r)	-	k_{red}=*1.1* [162])	CV	68Sto
	1 **M** NaOH/H_2O	Ag/AgO/Ag_2O (r)	23	*0.02*	GM	68Cla2, 71Ham2
	4 **M** NaSCN +0.1 **M** NaOH +0.019 **M** $AgNO_3$ /H_2O	$Ag^{+/0}$	20.3	0.055	IP	72Bek
	4 **M** NaSCN +0.1 **M** NaOH +0.019 **M** $AgNO_3$ /H_2O	$Ag^{+/0}$	26.5	0.068	IP	72Bek
	4 **M** NaSCN +0.1 **M** NaOH +0.019 **M** $AgNO_3$ /H_2O	$Ag^{+/0}$	33.5	0.080	IP	72Bek
	4 **M** NaSCN +0.1 **M** NaOH +0.019 **M** $AgNO_3$ /H_2O	$Ag^{+/0}$	42.4	0.120	IP	72Bek

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Ag	4 **M** NaSCN +0.1 **M** NaOH +0.031 **M** $AgNO_3/H_2O$	$Ag^{+/0}$	20.3	0.065	IP	72Bek
	4 **M** NaSCN +0.1 **M** NaOH +0.031 **M** $AgNO_3$ /H_2O	$Ag^{+/0}$	26.5	0.075	IP	72Bek
	4 **M** NaSCN +0.1 **M** NaOH +0.031 **M** $AgNO_3$ /H_2O	$Ag^{+/0}$	33.5	0.090	IP	72Bek
	4 **M** NaSCN +0.1 **M** NaOH +0.031 **M** $AgNO_3$ /H_2O	$Ag^{+/0}$	42.4	0.130	IP	72Bek
	4 **M** NaSCN +0.1 **M** NaOH +0.055 **M** $AgNO_3$ /H_2O	$Ag^{+/0}$	20.3	0.075	IP	72Bek
	4 **M** NaSCN +0.1 **M** NaOH +0.055 **M** $AgNO_3$ /H_2O	$Ag^{+/}$	26.5	0.09	IP	72Bek
	4 **M** NaSCN +0.1 **M** NaOH +0.055 **M** $AgNO_3$ /H_2O	$Ag^{+/0}$	33.5	0.110	IP	72Bek
	4 **M** NaSCN +0.1 **M** NaOH +0.055 **M** $AgNO_3$ /H_2O	$Ag^{+/0}$	42.4	0.160	IP	72Bek
	4 **M** NaSCN +0.1 **M** NaOH +0.097 **M** $AgNO_3$ /H_2O	$Ag^{+/0}$	26.5	0.11	IP	72Bek
	4 **M** NaSCN +0.1 **M** NaOH +0.097 **M** $AgNO_3$ /H_2O	$Ag^{+/0}$	20.3	0.1	IP	72Bek
	4 **M** NaSCN +0.1 **M** NaOH +0.097 **M** $AgNO_3$ /H_2O	$Ag^{+/0}$	33.5	0.16	IP	72Bek
	4 **M** NaSCN +0.1 **M** NaOH +0.097 **M** $AgNO_3$ /H_2O	$Ag^{+/0}$	42.4	0.22	IP	72Bek

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.14 **M** RbCN +2.1 **M** RbCl//H_2O	$Ag^{+/0}$, 0.33 **M**	25	$8.5\cdot10^{-3}$	CP	68Nec
	0.1 **M** KPF_6/H_2O	$[Co(NH_3)_6]^{3+/2+}$	25	*(3.0·10^{-4})*	RDE	80Bar1
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	25	*(7.0·10^{-4})*	RDE	80Bar1
	0.1 **M** KPF_6/H_2O	$[Co(NH_3)_5]F^{2+/+}$	25	*(4.5·10^{-4})*	RDE	80Bar1
	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_5]F^{2+/+}$	25	*(8.0·10^{-4})*	RDE	80Bar1
	0.1 **M** KPF_6/H_2O	$[Co(NH_3)_5]OH_2^{3+/2+}$	25	*(~10^{-4})*	RDE	80Bar1
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	$[Co(NH_3)_5]NCS^{2+/1+}$	24	*(2.0·10^{-2})*	PP	84Bar2
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	*cis*-$([Co(en)_2](NCS)_2)^{2+/1+}$	24	*(1.0·10^{-3})*	PP	84Bar2 [63])
	0.2 **M** KPF_6/H_2O	$Cr_{aq}^{2+/3+}$	25	*(>10^{-4})*	RDE	80Bar1
	0.2 **M** $NaClO_4/H_2O$	$Cr_{aq}^{2+/3+}$	25	*(>10^{-4})*	RDE	80Bar1
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	$[Cr(NH_3)_5]NCS^{2+/1+}$	24	*(6.0·10^{-4})*	PP	84Guy
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	$[Cr(OH_2)_5]NCS^{2+/1+}$	24	*(1.2·10^{-3})*	PP	84Guy
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	$([Cr(NH_3)_5]N_3)^{2+/1+}$	24	*(<·10^{-5})*	PP	84Guy
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	$([Cr(OH_2)_5]N_3)^{2+/1+}$	24	*(3.0·10^{-5})*	PP	84Guy
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	$[Cr(NH_3)_5]Cl^{2+/1+}$	24	*(1.3·10^{-2})*	PP	84Guy
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	$[Cr(OH_2)_5]Cl^{2+/1+}$	24	*(≈0.3)*	PP	84Guy
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	$[Cr(NH_3)_5]Br^{2+/1+}$	24	*(≈0.1)*	PP	84Guy
	0.1 **M** $NaClO_4$+5 m**M** $HClO_4/H_2O$	$[Cr(OH_2)_5]Br^{2+/1+}$	24	*(>0.5)*	PP	84Guy
	0.2 **M** KPF_6/H_2O	$Eu_{aq}^{2+/3+}$	25	*(1.0·10^{-2})*	RDE	80Bar1
	0.2 **M** KPF_6 +1 m**M** Cr^{2+}/H_2O	$Eu_{aq}^{2+/3+}$	25	*(~1·10^{-4})*	RDE	80Bar1
	0.5 **M** K-oxalate/H_2O	$Fe(ox)_3^{4-/3-}$	25	*0.001*	IP	52Ran2
	0.01 **M** HCl/H_2O	H^+/H_2	21	$10^{-8.94}$	CP	68Gos
	0.1 **M** HCl/H_2O	H^+/H_2	21	$10^{-9.47}$	CP	68Gos
	0.1 **M** HCl/H_2O	H^+/H_2	23	0.27; $8.4\cdot10^{-7}$	CP	64Boc3 [163])
	0.3 **M** HCl/H_2O	H^+/H_2	21	$10^{-8.94}$	CP	68Gos
	0.4 **M** HCl/H_2O	H^+/H_2	23	0.28; $4.4\cdot10^{-6}$	CP	64Boc3 [163])
	1 **M** HCl/H_2O	H^+/H_2	21	$10^{-9.81}$	CP	68Gos
	2 **M** HCl/H_2O	H^+/H_2	21	$10^{-8.14}$	CP	68Gos
	3 **M** HCl/H_2O	H^+/H_2	21	$10^{-8.35}$	CP	68Gos
	5 **M** HCl/H_2O	H^+/H_2	21	$10^{-8.33}$	CP	68Gos
	0.01 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-9.33}$	CP	68Gos
	0.1 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-10.02}$	CP	68Gos

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Ag	1 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-9.72}$	CP	68Gos
	3 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-9.73}$	CP	68Gos
	5 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-9.63}$	CP	68Gos
	7 **M** H_2SO_4/H_2O	H^+/H_2	21	$10^{-7.79}$	CP	68Gos
	1 **M** KF/H_2O	$[Ru(NH_3)_6]^{3+/2+}$	20	120	TPF	85Iwa
	0.3 **M** KPF_6/H_2O	$V_{aq}^{2+/3+}$	25	*(9.0·10^{-2})*	RDE	80Bar1
	0.3 **M** KPF_6 +1 m**M** Cr^{2+}/H_2O	$V_{aq}^{2+/3+}$	25	*(~10^{-4})*	RDE	80Bar1
	0.2 **M** KPF_6/H_2O	$Ru_{aq}^{2+/3+}$	25	*(>5.0·10^{-2})*	RDE	80Bar1
	0.2 **M** KPF_6 +1 m**M** Cr^{2+}/H_2O	$Ru_{aq}^{2+/3+}$	25	*(5·10^{-4})*	RDE	80Bar1
Ag(110)	0.1 **M** $NaClO_4/H_2O$	$[Co(NH_3)_6]^{3+/2+}$	23	*(8·10^{-4})*	CV	86Ham
Sodium (Na)	1.0 **M** NMe_4OH/H_2O	Na^+	20	*~0.4*	IP	52Ran1
(Na)$_{(Hg)}$	0.1 **M** NEt_4PF_6 /DMF [54])	Na^+	25	*9.0·10^{-2}*	IP/PP	80Bar3
	0.1 **M** NPr_4PF_6 [164]) /DMF [54])	Na^+	25	*0.2*	IP/PP	80Bar3
	0.1 **M** $NBut_4PF_6$ /DMF [54])	Na^+	25	*0.32*	IP/PP	80Bar3
Tantalum (Ta)	1 **M** $HClO_4/H_2O$	H^+/H_2	-	$10^{-8.6}$	CP	69Gal
	0.1 **N** H_2SO_4/H_2O	H^+/H_2, 0.1 **N**	25	$1.2\cdot10^{-8}$	CP	63Koz
	0.1 **N** H_2SO_4/H_2O	H^+/H_2, 0.1 **N**	35	$1.6\cdot10^{-8}$	CP	63Koz
	0.1 **N** H_2SO_4/H_2O	H^+/H_2, 0.1 **N**	45	$2.0\cdot10^{-8}$	CP	63Koz
	1 **N** H_2SO_4/H_2O	H^+/H_2, 1 **N**	25	$1.0\cdot10^{-8}$	CP	63Koz
	1 **N** H_2SO_4/H_2O	H^+/H_2, 1 **N**	35	$1.6\cdot10^{-8}$	CP	63Koz
	1 **N** H_2SO_4/H_2O	H^+/H_2, 1 **N**	45	$4.0\cdot10^{-8}$	CP	63Koz
	1 **N** H_2SO_4/H_2O	H^+/H_2, 1 **N**	55	$5.0\cdot10^{-8}$	CP	63Koz
	1 **N** H_2SO_4/H_2O	H^+/H_2, 1 **N**	65	$1.3\cdot10^{-7}$	CP	63Koz
	10 **N** H_2SO_4/H_2O	H^+/H_2, 1 **N**	25	$4.0\cdot10^{-8}$	CP	63Koz
	1 **N** H_2SO_4/H_2O	H^+/H_2, 10 **N**	25	$3.2\cdot10^{-8}$	GCI [165])	65Koz
	1 **M** H_2SO_4/H_2O	$Fe^{2+/3+}$, 0.1 **M**	20	$1.04\cdot10^{-6}$	CP [166])	73Gal
	1 **M** H_2SO_4/H_2O	$Fe^{2+/3+}$, 0.1 **M**	25	$1.22\cdot10^{-6}$	CP [166])	73Gal
	1 **M** H_2SO_4/H_2O	$Fe^{2+/3+}$, 0.1 **M**	30	$1.39\cdot10^{-6}$	CP [166])	73Gal
	1 **M** H_2SO_4/H_2O	$Fe^{2+/3+}$, 0.1 **M**	35	$1.63\cdot10^{-6}$	CP [166])	73Gal
	1 **M** H_2SO_4/H_2O	$Fe^{2+/3+}$, 0.1 **M**	40	$1.87\cdot10^{-6}$	CP [166])	73Gal
Thallium (Tl)						
(Tl)$_{(Hg)}$1% Tl	0.1 **M** $HClO_4/H_2O$	H^+(r)	15	$7.6\cdot10^{-14}$	CP	65But3
	0.1 **M** $HClO_4/H_2O$	H^+(r)	26	$2.5\cdot10^{-13}$	CP	65But3
	0.1 **M** $HClO_4/H_2O$	H^+(r)	36	$5\cdot10^{-13}$	CP	65But3
	0.1 **M** $HClO_4/H_2O$	H^+(r)	48	$6.8\cdot10^{-13}$	CP	65But3
	0.1 **M** $HClO_4/H_2O$	H^+(r)	58	$2.2\cdot10^{-12}$	CP	65But3
	0.1 **M** $HClO_4/H_2O$	H^+(r)	66	$2.2\cdot10^{-12}$	CP	65But3
(Tl)$_{(Hg)}$40.6% Tl	0.1 **M** $HClO_4/H_2O$	H^+(r)	14	$1.1\cdot10^{-14}$	CP	65But3
	0.1 **M** $HClO_4/H_2O$	H^+(r)	25	$4\cdot10^{-14}$	CP	65But3

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	0.1 **M** $HClO_4/H_2O$	H^+(r)	33	$1.1\cdot10^{-13}$	CP	65But3
	0.1 **M** $HClO_4/H_2O$	H^+(r)	42	$3.3\cdot10^{-13}$	CP	65But3
	0.1 **M** $HClO_4/H_2O$	H^+(r)	51	$1.9\cdot10^{-12}$	CP	65But3
	0.1 **M** $HClO_4/H_2O$	H^+(r)	62	$4.8\cdot10^{-12}$	CP	65But3
	0.1 **M** $HClO_4/H_2O$	H^+(r)	69	$3.1\cdot10^{-11}$	CP	65But3
10% Tl	0.2 **M** $NaCN/H_2O$	$Cr(CN)_6^{3-/4-}$, 2.5 m**M**	30	$3.9\cdot10^{-3}$	IP	60Del
31% Tl	0.2 **M** $NaCN/H_2O$	$Cr(CN)_6^{3-/4-}$, 2.5 m**M**	30	$4.8\cdot10^{-3}$	IP	60Del
(Tl)$_{(Hg)}$	1 **M** $HClO_4/H_2O$	$Tl^{+/0}$	36	*1.8*	IP	61Ran
(Tl)$_{(Hg)}$ 0.09% Tl	0.5 **M** H_2SO_4/H_2O	$Tl^{+/0}$, 1 m**M**	30	2.36	IP	67Bat
(Tl)$_{(Hg)}$ 0.0074% Tl	0.5 **M** H_2SO_4/H_2O	$Tl^{+/0}$, 0.97 m**M**	30	0.43	IP	67Bat
(Tl)$_{(Hg)}$ 0.0057% Tl	0.5 **M** H_2SO_4/H_2O	$Tl^{+/0}$, 10 m**M**	30	0.67	IP	67Bat
(Tl)$_{(Hg)}$	1 **M** KCl/H_2O	$Tl^{+/0}$	22	*0.15*	FR	58Bar1
(Tl)$_{(Hg)}$	1 **M** KCl/H_2O	$Tl^{+/0}$, 1.05 m**M**	36	*3.5*	FR	71Aga
(Tl)$_{(Hg)}$	1 **M** KCl/H_2O	$Tl^{+/0}$, 2.1 m**M**	36	*2.5*	FR	71Aga
(Tl)$_{(Hg)}$	1 **M** KNO_3/H_2O	$Tl^{+/0}$	22	*0.3*	FR	58Bar1
(Tl)$_{(Hg)}$	1 **M** KNO_3/H_2O	$Tl^{+/0}$	20	*>1*	IP	52Ran1
(Tl)$_{(Hg)}$	1 **M** KNO_3/H_2O	$Tl^{+/0}$, 1.05 m**M**	35	*0.69*	FR	71Aga
(Tl)$_{(Hg)}$	0.33 **M** KSO_4+5 m**M** H_2SO_4/H_2O	$Tl^{+/0}$	25	*1.25*	OC	61Gav
(Tl)$_{(Hg)}$	0.5 **M** K_2SO_4/H_2O	Tl^{1+}	22	*3.1*	FR	58Bar1
(Tl)$_{(Hg)}$	0.5 **M** K_2SO_4/H_2O	$Tl^{+/0}$, 2.1 m**M**	36	3.5 .. *7.4*	FR	71Aga
(Tl)$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	Tl^{1+}	22	*1.2*	FR	58Bar1
(Tl)$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Tl^{+/0}$	36	$> 1\cdot10^{2}$	IP	61Lor
(Tl)$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Tl^{+/0}$, 1.05 m**M**	36	*3.4*	FR	71Aga
(Tl)$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Tl^{+/0}$, 2.1 m**M**	36	*8.3*	FR	71Aga
Tin (Sn)	1 **M** HCl/H_2O	H^+/H_2	-	$10^{-10.2}$	CP	65Mat
	1 **M** H_2SO_4/H_2O	H^+/H_2	RT	$2.2\cdot10^{-9}$	CP	63Gos
Titanium (Ti)	1 **M** H_2SO_4+0.1 **M** Na_2SO_4/H_2O	$Fe^{2+/3+}$, 0.05**M**	25	$2.6\cdot10^{-7}$	CP	64Mak [167])
	H_2SO_4/H_2O, pH=0.25	H^+/H_2	RT	$15.4\cdot10^{-9}$	CP	70Tho
	H_2SO_4/H_2O, pH=0.5	H^+/H_2	RT	$13.5\cdot10^{-9}$	CP	70Tho
	H_2SO_4/H_2O, pH=0.9	H^+/H_2	RT	$11.5\cdot10^{-9}$	CP	70Tho
	H_2SO_4/H_2O, pH=1	H^+/H_2	RT	$10.5\cdot10^{-9}$	CP	70Tho
	H_2SO_4/H_2O, pH=1.35	H^+/H_2	RT	$8.8\cdot10^{-9}$	CP	70Tho
	H_2SO_4/H_2O, pH=1.65	H^+/H_2	RT	$7.5\cdot10^{-9}$	CP	70Tho
	H_2SO_4/H_2O, pH=1.90	H^+/H_2	RT	$6.5\cdot10^{-9}$	CP	70Tho
	H_2SO_4/H_2O, pH=2.25	H^+/H_2	RT	$5.5\cdot10^{-9}$	CP	70Tho
	1.33 **M** NaOH+1.95 **M** $NaNO_3/H_2O$	NO_3^-/NO	25	$1.5\cdot10^{-9}$	CV	96Boc [168])
Ti_2Ni	1 **M** $NaOH/H_2O$	hydrogen evolution	25	$46\cdot10^{-6}$	CP	97Krs

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Ti_3Ni	1 **M** $NaOH/H_2O$	hydrogen evolution	25	$12\cdot10^{-6}$	CP	97Krs
$TiNi_{0.7}$	1 **M** $NaOH/H_2O$	hydrogen evolution	25	$46\cdot10^{-6}$	CP	97Krs
TiNi	1 **M** $NaOH/H_2O$	hydrogen evolution	25	$13.2\cdot10^{-6}$	CP	97Krs
$TiNi_2$	1 **M** $NaOH/H_2O$	hydrogen evolution	25	$51\cdot10^{-6}$	CP	97Krs
$TiNi_3$	1 **M** $NaOH/H_2O$	hydrogen evolution	25	$19\cdot10^{-6}$	CP	97Krs
$TiNi_4$	1 **M** $NaOH/H_2O$	hydrogen evolution	25	$31.6\cdot10^{-6}$	CP	97Krs
Tungsten (W)	0.1 **M** HCl/H_2O	H^+/H_2	23	0.78; $2.8\cdot10^{-6}$	CP	64Boc3 [169])
	0.5 **N** H_2SO_4/H_2O	H^+/H_2	-	$10^{-6.25}$	CP	64Boc3
	0.5 **N** $NaOH/H_2O$	H^+/H_2	-	$10^{-7.05}$	CP	64Boc3
Vanadium (V)	1 **M** $HClO_4/H_2O$	V^{4+}(r)	25	$8.0\cdot10^{-14}$	PS	70Arm
	0.01 or 1 **M** $HClO_4$ $/H_2O$(adjusted with NaCl to ionic strength 1)	V^{4+}(r)	25	$8.0\cdot10^{-14}$	PS	70Arm
Zinc (Zn)	1 **M** H_2SO_4/H_2O	H^+/H_2	RT	$9.9\cdot10^{-11}$	CP	63Gos
	2 **M** KCl/H_2O	Zn^{2+}/Zn^0, 0.5 **M**	-	0.3	IP	59Lor
	7 **M** KOH/H_2O	Zn^{2+}/Zn^0, 0.04 **M**	25	0.241	GD	67Far
	7 **M** KOH/H_2O	Zn^{2+}/Zn^0, 0.08 **M**	25	0.231	GD	67Far
	7 **M** KOH/H_2O	Zn^{2+}/Zn^0, 0.64 **M**	25	0.224	GD	67Far
	7 **M** KOH/H_2O	Zn^{2+}/Zn^0, 0.016 **M**	25	0.238	GD	67Far
	7 **M** KOH/H_2O	Zn^{2+}/Zn^0, 0.04 **M**	25	0.241	GD	67Far
	7 **M** KOH/H_2O	Zn^{2+}/Zn^0, 0.08 **M**	25	0.231	GD	67Far
	7 **M** KOH/H_2O	Zn^{2+}/Zn^0, 0.64 **M**	25	0.224	GD	67Far
	4.95 **M** KOH/H_2O	Zn^{2+}/Zn^0, 0.056 **M**	25	0.167	GD	67Far
	4.95 **M** KOH/H_2O	Zn^{2+}/Zn^0, 0.14 **M**	25	0.175	GD	67Far
	4.95 **M** KOH/H_2O	Zn^{2+}/Zn^0, 0.28 **M**	25	0.166	GD	67Far
	$NaClO_4+Zn(ClO_4)_2$ $/H_2O$ [170])	Zn^{2+}/Zn^0, 0.0073 **M**	24	$2.84\cdot10^{-4}$	IP	70Cla
	$NaClO_4+Zn(ClO_4)_2$ $/H_2O$ [170])	Zn^{2+}/Zn^0, 0.022 **M**	24	$1.22\cdot10^{-4}$	IP	70Cla
	$NaClO_4+Zn(ClO_4)_2$ $/H_2O$ [170])	Zn^{2+}/Zn^0, 0.056 **M**	24	$3.2\cdot10^{-4}$	GM	70Cla
	$NaClO_4+Zn(ClO_4)_2$ $/H_2O$ [170])	Zn^{2+}/Zn^0, 0.057 **M**	24	$3.4\cdot10^{-4}$	GM	70Cla
	$NaClO_4+Zn(ClO_4)_2$ $/H_2O$ [170])	Zn^{2+}/Zn^0, 0.061 **M**	24	$1.97\cdot10^{-4}$	IP	70Cla
	$NaClO_4+Zn(ClO_4)_2$ $/H_2O$ [170])	Zn^{2+}/Zn^0, 0.087 **M**	24	$1.73\cdot10^{-4}$	IP	70Cla
	$NaClO_4+Zn(ClO_4)_2$ $/H_2O$ [170])	Zn^{2+}/Zn^0, 0.092 **M**	24	$6.68\cdot10^{-4}$	IP	70Cla
	$NaClO_4+Zn(ClO_4)_2$ $/H_2O$ [170])	Zn^{2+}/Zn^0, 0.17 **M**	24	$3.7\cdot10^{-4}$	GM	70Cla

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
	$NaClO_4$+$Zn(ClO_4)_2$ /H_2O [170])	Zn^{2+}/Zn^0, 0.27 **M**	24	$4.2\cdot10^{-4}$	IP	70Cla
	$NaClO_4$+$Zn(ClO_4)_2$ /H_2O [170])	Zn^{2+}/Zn^0, 0.46 **M**	24	$3.5\cdot10^{-4}$	GM	70Cla
	$NaClO_4$+$Zn(ClO_4)_2$ /H_2O [170])	Zn^{2+}/Zn^0, 0.51 **M**	24	$6.26\cdot10^{-4}$	IP	70Cla
	0.1 **M** NaOH+0.9 **M** $NaClO_4/H_2O$	Zn^{2+}/Zn^0, 0.00025 **M**	25	0.041	GD	67Far
	0.1 **M** NaOH+0.9 **M** $NaClO_4/H_2O$	Zn^{2+}/Zn^0, 0.0005 **M**	25	0.042	GD	67Far
	0.2 **M** NaOH+0.8 **M** $NaClO_4/H_2O$	Zn^{2+}/Zn^0, 0.00075 **M**	25	0.045	GD	67Far
	0.2 **M** NaOH+0.8 **M** $NaClO_4/H_2O$	Zn^{2+}/Zn^0, 0.0015 **M**	25	0.052	GD	67Far
	0.5 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.009 **M**	25	0.087	GD	67Far
	0.5 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.009 **M**	40	0.118	GD	67Far
	0.5 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.009 **M**	55	0.143	GD	67Far
	0.7 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.047 **M**	25	0.083	GD	67Far
	0.7 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.047 **M**	40	0.11	GD	67Far
	0.7 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.047 **M**	55	0.141	GD	67Far
	1.05 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.009 **M**	25	0.092	GD	67Far
	1.05 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.009 **M**	40	0.127	GD	67Far
	1.05 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.009 **M**	55	0.158	GD	67Far
	1.4 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.047 **M**	25	0.091	GD	67Far
	1.4 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.047 **M**	40	0.122	GD	67Far
	1.4 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.047 **M**	55	0.155	GD	67Far
	2.1 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.047 **M**	25	0.103	GD	67Far
	2.1 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.047 **M**	40	0.135	GD	67Far
	2.1 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.047 **M**	55	0.163	GD	67Far
	2.6 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.009 **M**	25	0.1	GD	67Far
	2.6 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.009 **M**	40	0.146	GD	67Far
	2.6 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.009 **M**	55	0.178	GD	67Far
	7 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.0014 **M**	25	0.16	GD	67Far
	7 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.0044 **M**	25	0.12	GD	67Far
	7 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.0094 **M**	25	0.13	GD	67Far
	7 **M** $NaOH/H_2O$	Zn^{2+}/Zn^0, 0.47 **M**	25	0.138	GD	67Far
	Na_2SO_4 +$ZnSO_4/H_2O$ [171])	Zn^{2+}/Zn^0, 0.011 **M**	24	$4.94\cdot10^{-4}$	IP	70Cla
	Na_2SO_4 +$ZnSO_4/H_2O$ [171])	Zn^{2+}/Zn^0, 0.14 **M**	24	$3.66\cdot10^{-4}$	IP	70Cla
	Na_2SO_4 +$ZnSO_4/H_2O$ [171])	Zn^{2+}/Zn^0, 0.5 **M**	24	$10.9\cdot10^{-4}$	IP	70Cla
	Na_2SO_4 +$ZnSO_4/H_2O$ [171])	Zn^{2+}/Zn^0, 0.88 **M**	24	$12.6\cdot10^{-4}$	IP	70Cla
	1 **M** $ZnSO_4/H_2O$	Zn^{2+}/Zn^0	RT	$2\cdot10^{-5}$	-	70Cla

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Zn$_{(Hg)}$	0.025 **M** $Ba(ClO_4)_2$ /H_2O	$Zn^{2+/0}$	26	0.0091	GS	64Ara
Zn$_{(Hg)}$	0.05 **M** $Ba(ClO_4)_2$ /H_2O	$Zn^{2+/0}$	26	0.0057	GS	64Ara
Zn$_{(Hg)}$	0.125 **M** $Ba(ClO_4)_2$ /H_2O	$Zn^{2+/0}$	26	0.0032	GS	64Ara
Zn$_{(Hg)}$	0.25 **M** $Ba(ClO_4)_2$ /H_2O	$Zn^{2+/0}$	26	0.0021	GS	64Ara
Zn$_{(Hg)}$	0.5 **M** $ZnCl_2$ +2 **M** KCl/H_2O	$Zn^{2+/0}$	-	0.3	IP	59Lor
Zn$_{(Hg)}$	1.0 **M** KBr/H_2O	$Zn^{2+/0}$	20	*0.008*	IP	52Ran1
Zn$_{(Hg)}$	1.0 **M** KBr/H_2O	$Zn^{2+/0}$	25	*0.005*	CS	64Ham2
Zn$_{(Hg)}$	1.0 **M** KBr/H_2O	$Zn^{2+/0}$	20	*0.008*	ACP	61Kam
Zn$_{(Hg)}$	1.0 **M** KBr/H_2O	$Zn^{2+/0}$, 0.01...0.1 **M** Zn^{2+}	27	*1.2·10^{-2}*	FR	69Bro2
Zn$_{(Hg)}$	1.0 **M** KCNS/H_2O	$Zn^{2+/0}$	25	1.64	IP	47Ran
Zn$_{(Hg)}$	0.2 **M** KCl/H_2O	$Zn^{2+/0}$	-	0.33	IP	62Beh
Zn$_{(Hg)}$	0.25 **M** KCl/H_2O	$Zn^{2+/0}$	25	*0.022*	FR	66Bau
Zn$_{(Hg)}$	0.25 **M** KCl/H_2O	$Zn^{2+/0}$	45	*0.038*	FR	66Bau
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	0	0.25	IP	61Mül
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	20	0.65	IP	61Mül
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	25	0.9	IP	61Mül
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	25	0.58	IP	47Ran
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	25	*0.004*	CS	62Del2
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	25	*0.004*	CS	64Ham2
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	25	*0.016*	OP	61Gav
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	0	0.25	GM	61Mül
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	20	0.65	GM	61Mül
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	25	0.9	GM	61Mül
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$, 0.5 m**M**	25	*0.0043*	CS	62Del3
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$, 1 m**M**	25	*0.0042*	CS	62Del3
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$, 5 m**M**	25	*0.0043*	CS	62Del3
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$, 10 m**M**	25	*0.0041*	CS	62Del3
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$, 50 m**M**	25	*0.0035*	CS	62Del3
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	20	*0.004*	IP	52Ran1
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	20	*0.004*	IP	62Beh
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	25	*0.00505*	IP	63Bla
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O, pH=3	$Zn^{2+/0}$	-	*0.0093*	IP	67Tim
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	25	1.1	IP	70Sal1
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	25	*0.0049*	IP	63Slu
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	25	*0.006*	FR	66Bau
Zn$_{(Hg)}$	1.0 **M** KCl/H_2O	$Zn^{2+/0}$	45	*0.014*	FR	66Bau
Zn$_{(Hg\ 172)}$)	1.0 **M** KCl/H_2O [172])	$Zn^{2+/0}$, 0.01...0.1 **M** Zn^{2+}	27	*0.0043*	FR	69Bro2
Zn$_{(Hg)}$	2.0 **M** KCl+1 **M** $K_2C_2O_4/H_2O$	$Zn^{2+/0}$,0.01 **M** $[Zn(C_2O_4)_3]^{4-}$	25	3.1·10^{-5}	CP	68Kra

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Zn$_{(Hg)}$	2.0 **M** KCl+1 **M** $K_2C_2O_4/H_2O$	$Zn^{2+/0}$,0.01 **M** $[Zn(C_2O_4)_3]^{4-}$	25	$1.6\cdot10^{-4}$	GM	66Kra1
Zn$_{(Hg)}$	1.0 **M** KCNS/H_2O	$Zn^{2+/0}$	20	*0.017*	IP	52Ran1
Zn$_{(Hg)}$	1.0 **M** KI/H_2O	$Zn^{2+/0}$	20	*0.07*	IP	52Ran1
Zn$_{(Hg)}$	1.0 **M** KI/H_2O	$Zn^{2+/0}$	25	*0.004*	CS	64Ham2
Zn$_{(Hg)}$	1.0 **M** KI+5 m**M** H_2SO_4/H_2O	$Zn^{2+/0}$	25	*0.12*	IP	63Slu
Zn$_{(Hg)}$	1.0 **M** KNO_3/H_2O	$Zn^{2+/0}$	20	*0.0035*	IP	52Ran1
Zn$_{(Hg)}$	0.025 **M** $Mg(ClO_4)_2$ /H_2O	$Zn^{2+/0}$, 2 m**M**	26	0.012	GS	64Ara
Zn$_{(Hg)}$	0.05 **M** $Mg(ClO_4)_2$ /H_2O	$Zn^{2+/0}$, 2 m**M**	26	0.009	GS	64Ara
Zn$_{(Hg)}$	0.05 **M** $Mg(ClO_4)_2$ /H_2O	$Zn^{2+/0}$, 2 m**M**	26	0.009	GS	64Ara
Zn$_{(Hg)}$	0.05 **M** $Mg(ClO_4)_2$ +0.02 **M** *n*-amyl alcohol/H_2O	$Zn^{2+/0}$, 2 m**M**	26	0.0077	GS	64Ara
Zn$_{(Hg)}$	0.05 **M** $Mg(ClO_4)_2$ +0.04 **M** *n*-amyl alcohol/H_2O	$Zn^{2+/0}$, 2 m**M**	26	0.0042	GS	64Ara
Zn$_{(Hg)}$	0.05 **M** $Mg(ClO_4)_2$ +0.08 **M** *n*-amyl alcohol/H_2O	$Zn^{2+/0}$, 2 m**M**	26	0.00065	GS	64Ara
Zn$_{(Hg)}$	0.05 **M** $Mg(ClO_4)_2$ +0.12 **M** *n*-amyl alcohol/H_2O	$Zn^{2+/0}$, 2 m**M**	26	0.00032	GS	64Ara
Zn$_{(Hg)}$	0.05 **M** $Mg(ClO_4)_2$ +0.1 **M** *n*-amyl alcohol/H_2O	$Zn^{2+/0}$, 2 m**M**	26	0.00032	GS	64Ara
Zn$_{(Hg)}$	0.125 **M** $Mg(ClO_4)_2/H_2O$	$Zn^{2+/0}$, 2 m**M**	26	0.0047	GS	64Ara
Zn$_{(Hg)}$	0.125 **M** $Mg(ClO_4)_2$ +0.1 **M** *n*-amyl alcohol/H_2O	$Zn^{2+/0}$, 2 m**M**	26	0.00020	GS	64Ara
Zn$_{(Hg)}$	0.2 **M** $Mg(ClO_4)_2$ +0.01 **M** $HClO_4$ /H_2O	$Zn^{2+/0}$, 1 m**M**	20	0.0006	RT	65Mol2
Zn$_{(Hg)}$	0.25 **M** $Mg(ClO_4)_2$ /H_2O	$Zn^{2+/0}$, 2 m**M**	26	0.0027	GS	64Ara
Zn$_{(Hg)}$	0.25 **M** $Mg(ClO_4)_2$ /H_2O	$Zn^{2+/0}$, 1 m**M**	26	0.00175	GS	64Ara
Zn$_{(Hg)}$	0.25 **M** $Mg(ClO_4)_2$ /H_2O	$Zn^{2+/0}$, 2 m**M**	26	0.00266	GS	64Ara
Zn$_{(Hg)}$	0.25 **M** $Mg(ClO_4)_2$ /H_2O	$Zn^{2+/0}$, 4 m**M**	26	0.00452	GS	64Ara
Zn$_{(Hg)}$	0.25 **M** $Mg(ClO_4)_2$ /H_2O	$Zn^{2+/0}$, 8 m**M**	26	0.00747	GS	64Ara

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^{θ} [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Zn$_{(Hg)}$	0.25 **M** $Mg(ClO_4)_2$ +0.1 **M** *n*-amyl alcohol/H_2O	$Zn^{2+/0}$, 2 m**M**	26	0.00014	GS	64Ara
Zn$_{(Hg)}$	2.8 **M** $Mg(ClO_4)_2$ +0.01 **M** $HClO_4$ /H_2O	$Zn^{2+/0}$, 1 m**M**	20	0.00015	RT	65Mol2
Zn$_{(Hg)}$	1.0 **M** NH_4SCN	$Zn^{2+/0}$	25	$17.5\cdot10^{-3}$	IP	63Bla
Zn$_{(Hg)}$ [173])	1.0 **M** NH_4OH +1 **M** NH_4Cl/H_2O	$Zn^{2+/0}$	20	$1.8\cdot10^{-2}$	CP	65Mol1
Zn$_{(Hg)}$	Mg_2SO_4+5 m**M** H_2SO_4/H_2O (I=2 **M**)	$Zn^{2+/0}$	25	0.11	RT	63Bud
Zn$_{(Hg)}$	1.0 **M** NaBr/H_2O	$Zn^{2+/0}$	25	2.29	IP	70Sal1
Zn$_{(Hg)}$	1.0 **M** NaBr/H_2O	$Zn^{2+/0}$	25	*$10.3\cdot10^{-3}$*	IP	63Bla
Zn$_{(Hg)}$	0.1 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$	45	*0.02*	FD	64Bau
Zn$_{(Hg)}$	0.1 **M** $NaClO_4$ +1 m**M** $HClO_4$ 2 µ**M** LEO [98])	$Zn^{2+/0}$, 1 m**M**	25	*$16.7\cdot10^{-3}$*	PS	69Tan2
Zn$_{(Hg)}$	0.3 **M** $NaClO_4$ +1 m**M** $HClO_4$ 2 µ**M** LEO [98])	$Zn^{2+/0}$, 1 m**M**	25	*$7.7\cdot10^{-3}$*	PS	69Tan2
Zn$_{(Hg)}$	0.5 **M** $NaClO_4$ +1 m**M** $HClO_4$ 2 µ**M** LEO [98])	$Zn^{2+/0}$, 1 m**M**	25	*$3.01\cdot10^{-3}$*	PS	69Tan2
Zn$_{(Hg)}$	0.1 **M** $NaClO_4$ +0.01 **M** $HClO_4$ /H_2O	$Zn^{2+/0}$, 1 m**M**	20	0.017	RT	65Mol2
Zn$_{(Hg)}$	0.2 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$	25	*0.00119*	IP	63Hus
Zn$_{(Hg)}$	0.25 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$	45	*0.013*	FD	64Bau
Zn$_{(Hg)}$	0.5 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$	45	*0.0075*	FD	64Bau
Zn$_{(Hg)}$	0.5 **M** $NaClO_4$ +0.01 **M** $HClO_4$ /H_2O	$Zn^{2+/0}$, 1 m**M**	20	0.025	RT	65Mol2
Zn$_{(Hg)}$	0.88 **M** $NaClO_4$ +1 m**M** $HClO_4$ /H_2O	$Zn^{2+/0}$	25	*$3.26\cdot10^{-3}$*	IP	60Slu
Zn$_{(Hg)}$	1 **M** $NaClO_4$ +0.01 **M** $HClO_4$ /H_2O	$Zn^{2+/0}$	-	*$3.6\cdot10^{-3}$*	IP	65Ijz
Zn$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$, 20 m**M**	0	0.0054	GM	55Ger1
Zn$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$, 20 m**M**	0	0.0058	PS	55Ger1
Zn$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$, 6 m**M**	1	0.0021	PS	55Vie
Zn$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$	0	0.1	IP	61Mül
Zn$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$	25	*$2.5\cdot10^{-3}$*	CM	63Ham
Zn$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$	45	*0.005*	FD	64Bau
Zn$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$	25	*0.002*	CS	64Ham2
Zn$_{(Hg)}$	1 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$	25	0.73	IP+FR	68Sal

Electrode [1])	Solution [2])	Educt, product, concentration [3])	T [4]) [°C]	j_0 , j_{00} k_{app} or k_c^θ [5]) [A cm^{-2} or cm s^{-1}]	Method [174])	Ref.
Zn$_{(Hg)}$	1 **M** $NaClO_4$ +0.01 **M** $HClO_4$ /H_2O	$Zn^{2+/0}$, 1 m**M**	20	0.00036	RT	65Mol2
Zn$_{(Hg)}$	2 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$	25	*0.00248*	IP	63Hus
Zn$_{(Hg)}$	2 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$	45	*0.0026*	CH	68Lar
Zn$_{(Hg)}$	2 **M** $NaClO_4$ +$NaCl/H_2O$	$Zn^{2+/0}$	25	*5.4*	IP	63Bla
Zn$_{(Hg)}$	2 **M** $NaClO_4$ +$NaBr/H_2O$	$Zn^{2+/0}$	25	*11.5*	IP	63Bla
Zn$_{(Hg)}$	2 **M** $NaClO_4$ +NaI/H_2O	$Zn^{2+/0}$	25	*66*	IP	63Bla
Zn$_{(Hg)}$	2 **M** $NaClO_4$ +$NaSCN/H_2O$	$Zn^{2+/0}$	25	*105*	IP	63Bla
Zn$_{(Hg)}$	3 **M** $NaClO_4/H_2O$	$Zn^{2+/0}$	25	*0.0028*	IP	63Hus
Zn$_{(Hg)}$	3 **M** $NaClO_4$+0.01 **M** $HClO_4/H_2O$	$Zn^{2+/0}$, 1 m**M**	20	$1.64\cdot10^{-4}$	RT	65Mol2
Zn$_{(Hg)}$	1.0 **M** NaI/H_2O	$Zn^{2+/0}$	25	*$131\cdot10^{-3}$*	IP	63Bla
Zn$_{(Hg)}$	1 **M** $NaOH/H_2O$	$Zn^{2+/0}$	20	*$2.5\cdot10^{-4}$*	IP	62Beh
Zn$_{(Hg)}$	0.5 **M** Na_2SO_4/H_2O	$Zn^{2+/0}$	25	*0.001*	CS	64Ham2

[1]) The subscript (Hg) denotes an amalgam electrode.

[2]) The concentration of substances added e.g. to maintain a certain pH is sometimes not stated exactly in the original report; this is indicated by an x. Concentrations are given in units as reported if not stated otherwise (**M** = molarity; **m** = molality, **F** equal to **M**).

[3]) The educt/reacting species for the charge transfer reaction to which the kinetic data are referring to is stated only if it is not already given by the composition of the electrolyte solution. If the concentrations of the reactants are given in addition the exchange current density j_0 will be stated subsequently. The arrangement of species identifies the reaction as a reduction or an oxidation in cases where no simple exchange reaction at rest potential was studied. In the absence of specific designations of the reacting species (e.g. reactions involving complex organometallic compounds) the type of reaction is designated (r) or (o) indicating a reduction or oxidation. Concentrations are given in units as reported if not stated otherwise (**M** = molarity; **m** = molality, **F** equal to **M**).

[4]) RT designates room temperature as specified in the original report.

[5]) If no concentration of the educt is given the standard exchange current density j_{00} is stated. Values of k_c^θ are printed in italics; values of the apparent rate constant k_{app} are printed in parentheses in italics. For electrode potentials where the latter rate constant was actually determined the reader is referred to the original literature.

[6]) The influence of the concentration of chloride ions was also studied, no values are listed here.

[7]) Bismuth concentration in amalgam was 3 m**M**.

[8]) Bismuth concentration in amalgam was $5\cdot10^{-6}$ **M**.

[9]) Bismuth concentration in amalgam was 0.11 **M**.

[10]) Bismuth concentration in amalgam was 0.3 **M**.

[11]) Bismuth concentration in amalgam was 0.12 **M**.

[12]) Bismuth concentration in amalgam was $5\cdot10^{-4}$ **M**.

[13]) Bismuth concentration in amalgam was $5\cdot10^{-6}$ **M**.

[14]) The concentration refers to Bi^{3+} and $Bi_{(Hg)}$.

[15]) The value was obtained using an extrapolation method; with a numerical method k_c^θ = 4.1 cm s^{-1} was obtained.

[16]) The value was obtained using an extrapolation method; with a numerical method k_c^θ = 0.44 cm s^{-1} was obtained.

[17]) The value of j_0 is related to the concentration of cadmium in the amalgam and Cd^{2+} in solution according to $j_0 = 8.4\ [Cd]^{0.17}\ [Cd^{2+}]^{0.83}$.

[18]) Because the reactant was present as a film on the electrode no reactant concentration is given.

[19]) Carbon refers to all kinds of carbonaceous, graphitic or similar electrode materials including glassy carbon.

[20]) A synthetic diamond electrode was used.

[21]) A boron-doped diamond electrode was used as prepared.

[22]) An electrochemically oxidized boron-doped diamond electrode was used. The observed rates were concentration-dependent.

[23]) A polycrystalline synthetic diamond electrode was used, values of k_0 range between $6 \cdot 10^{-9}$ and $2.4 \cdot 10^{-6}$ cm s^{-1}.

[24]) A single crystal synthetic diamond electrode was used, values of k_0 range between $4 \cdot 10^{-8}$ and $1.7 \cdot 10^{-5}$ cm s^{-1}.

[25]) A single crystal synthetic diamond electrode was used, values of k_0 range between $4 \cdot 10^{-8}$ and $6.5 \cdot 10^{-7}$ cm s^{-1}.

[26]) A polished glassy carbon electrode was used.

[27]) A carbon paste electrode was used.

[28]) A carbon paste made of Acheson graphite and bromonaphthalene was used.

[29]) A carbon paste made of Acheson graphite and nujol was used.

[30]) A carbon paste made of Acheson graphite and bromoform was used.

[31]) A carbon paste made of Acheson graphite and Silicon 1000 was used.

[32]) A carbon paste made of Acheson graphite and Silicon 350 was used.

[33]) A carbon paste made of Acheson graphite and Silicon 50 was used.

[34]) A carbon paste made of Acheson graphite and Silicon 20 was used.

[35]) A carbon paste made of Acheson graphite and Silicon 10 was used.

[36]) Highly boron doped diamond film supplied by Naval Research Laboratory, USA.

[37]) Highly boron doped diamond film supplied by Utah State University, USA.

[38]) An argon-sputtered glassy carbon electrode was used.

[39]) A dinitrophenylhydrazine-treated glassy carbon electrode was used.

[40]) A dinitrobenzoyl chloride-treated glassy carbon electrode was used.

[41]) A 2,6-anthraquinonedisulfonate modified glassy carbon electrode was used.

[42]) A methylene blue-modified sputtered glassy carbon electrode was used.

[43]) A 1,4-bis(2-methylstyryl)benzene coated glassy carbon electrode was used.

[44]) A boron-doped diamond electrode was used as prepared.

[45]) An electrochemically oxidized boron-doped diamond electrode was used. The observed rates were concentration-dependent.

[46]) A carbon paste electrode was used.

[47]) A boron-doped diamond electrode was used after wet etching.

[48]) Anodically activated glassy carbon.

[49]) THAP = Tetrahexylammoniumperchlorate.

[50]) Calculated with the transfer coefficient and the nitrate concentration only as reported.

[51]) DMSO = Dimethylsulphoxide

[52]) TBAP = Tetrabutylammoniumhexafluorophosphate.

[53]) A synthetic diamond electrode was used; values of k_0 range between $8.5 \cdot 10^{-7}$ and $1.2 \cdot 10^{-5}$ cm s^{-1}.

[54]) DMF = Dimethylformamide.

[55]) Chromium in its active state.

[56]) Chromium in its passive state.

[57]) An electrolytically deposited and a press-formed electrode prepared from equimolar ratios of Cu (32 wt%) and $CuCl_2$ (67 wt%) and 1 wt% polyethylene binder were used.

[58]) Taken from Tafel slope.

[59]) An electrode prepared from equimolar ratios of Cu (32 wt%) and $CuCl_2$ (67 wt%) and 1 wt% polyethylene binder was used.
[60]) 10^{-3} **M** Cu.
[61]) $3 \cdot 10^{-6}$ **M** Cu.
[62]) $3 \cdot 10^{-6}$ ... $2.5 \cdot 10^{-3}$ **M** Cu.
[63]) en = ethylene diamine.
[64]) Value calculated from j_0 assuming $\alpha = 0.5$ and the given acidity, activity of hydrogen set to 1.
[65]) The authors report a considerable scattering of the obtained data.
[66]) 0.1 m**M** $BiCl_3$ have been added to the electrolyte solution.
[67]) 0.07 m**M** $PbCl_2$ have been added to the electrolyte solution.
[68]) 0.2 m**M** $SnCl_2$ have been added to the electrolyte solution.
[69]) 0.05 m**M** TlCl have been added to the electrolyte solution.
[70]) The electrode was modified with an adsorbed layer of tin.
[71]) Measured with an educt concentration of 0.1 **M**.
[72]) Measured with an educt concentration of 0.01 **M**.
[73]) This study includes an extended evaluation of the catalytic effect of chloride anions.
[74]) TEAP = Tetraethyl ammonium perchlorate.
[75]) The actual values of j_0 depended slightly on the pretreatment of the solid electrode.
[76]) Indium-plated platinum electrode.
[77]) Electrode chemically cleaned with chromic acid.
[78]) Experimental conditions were adjusted to keep the working electrode in the passivated state.
[79]) Calculated with the transfer coefficient and the nitrate concentration only as reported.
[80]) Lead alloy containing 1.08 % antimony.
[81]) Lead alloy containing 2.16 % antimony.
[82]) Lead alloy containing 5.76 % antimony.
[83]) Lead alloy containing 1.08 % antimony.
[84]) Lead alloy containing 2.16 % antimony.
[85]) Pr = Propyl
[86]) Bu = Butyl
[87]) Cp = cyclopentadiene.
[88]) NMF = **N**-methylformamide.
[89]) TBAFB = *n*-Tetrabutylammonium fluoroborate.
[90]) Cp' = pentamethylcyclopentadiene.
[91]) NMF = *N*-methylformamide.
[92]) acac = acetylacetonato.
[93]) DEF = *N,N*-diethylformamide.
[94]) CrEDTA = ethylenediaminetetraacetatochromate.
[95]) As a function of viscosity values of k_0 between *0.45* and *0.021* cm s^{-1} were found.
[96]) DMA = *N,N*-dimethylacetamide.
[97]) CrCYDTA = cyclohexanediaminetetraacetatochromate
[98]) LEO = Polyoxyethylene lauryl ether.
[99]) CrHEDTRA = (2-hydroxyethyl)-ethylenediaminetriacetatochromium.
[100]) HMPA = hexamethylphosphoric triamide.
[101]) θ refers to the electrode coverage with n-butanol.
[102]) ox = oxalate ion.
[103]) The value was obtained using an extrapolation method; with a numerical method $k_c^{\theta} = 1.16$ cm s^{-1} was obtained.

[104]) $(COT)Fe(CO)_3^{0/-}$ = tricarbonyl(η^4-cyclooctatetraene(iron).

[105]) acac = acetylacetonato.

[106]) The value of j_0 depended on the water/methanol ratio.

[107]) Liquid mercury.

[108]) Solid mercury.

[109]) Calculated with capacitance value from charging experiment.

[110]) Calculated with capacitance value from relaxation experiment.

[111]) Solution contained Triton X-100.

[112]) Mn(TPP)Cl = Chloro(tetraphenylporphinato)manganese(III).

[113]) PC = Propylene carbonate

[114]) For the electrode reaction $NiOH^+ + 2\ e^- \rightarrow Ni + OH^-$

[115]) For the electrode reaction $Ni^{2+} + 2\ e^- \rightarrow Ni$

[116]) For the electrode reaction $Ni(AcAc)^+ + 2\ e^- \rightarrow Ni + AcAc$

[117]) Sodium hexaphosphate.

[118]) Standard concentration 1 m**M**.

[119]) L = tartrate.

[120]) L = acetate.

[121]) L = glycolate.

[122]) THxAP = Tetrahexyl ammonium perchlorate.

[123]) TOAP = Tetraoctyl ammonium perchlorate.

[124]) HMPA = hexamethylphosphoramide.

[125]) Experimental conditions were adjusted to keep the working electrode in the passivated state.

[126]) Calculated with the transfer coefficient and the nitrate concentration only as reported.

[127]) Value obtained after saturating the electrode with hydrogen for 73 hours.

[128]) With an underpotentially deposited coating of Thallium.

[129]) Contrary to theory the value of j_{00} depends on the concentration ration of the redox components with values ranging from 34. $3 < j_{00} < 154.2$ A cm^{-2}.

[130]) Solution saturated with benzene.

[131]) Solution saturated with toluene.

[132]) The activity of the solution was adjusted to 1.

[133]) The unit **F** as used throughout this report refers to formality and is equal to **M** for molarity.

[134]) With an underpotentially deposited coating of thallium.

[135]) The electrode was reduced electrochemically prior to the kinetic measurement.

[136]) Bu = Butyl.

[137]) bipy = bipyridyl complex ligand.

[138]) Electrode prepared in hydrogen.

[139]) Electrode prepared in hydrogen and anodically activated.

[140]) Electrode prepared in air.

[141]) Because the reaction is second order the units are cm^4 mol^{-1} s^{-1}.

[142]) The electrode surface has been modified with adsorbed CO.

[143]) The ratio of the rate constant as observed with the modified electrode at a selected electrode potential is divided by the rate constant observed under the same conditions with an unmodified electrode.

[144]) The electrode surface has been modified with adsorbed ethylene.

[145]) The electrode surface has been modified with adsorbed 2-butene.

[146]) The electrode surface has been modified with adsorbed 1,3-butadiene.

[147]) The electrode surface has been modified with adsorbed 1-octene.

[148]) The electrode surface has been modified with adsorbed CO.

[149]) en = ethylene diamine.

[150]) dien = diethylenetriamine.

[151]) $(\eta_5\text{-}C_5H_4SiMe_2)_2Ti(C{\equiv}CSiMe_3)_2$

[152]) THF = tetrahydrofurane.

[153]) NMP = *N*-methylpyrrolidone-(2).

[154]) The platinum electrode was modified with an iodine adlayer of the specified symmetry.

[155]) AC = acetate anion.

[156]) PAc = phenylacetic acid anion.

[157]) Pyc = 4-pyridine carboxylic acid anion.

[158]) Cpc = Cyclopentadienyl carboxylic acid anion.

[159]) Depending on the degree of coverage with oxygen the rate constant varied between k_0 range between $4.6{\cdot}10^{-4}$ and $2{\cdot}10^{-5}$ cm s^{-1}.

[160]) Depending on the degree of coverage with oxygen the rate constant varied between k_0 range between $2.4{\cdot}10^{-4}$ and $1.9{\cdot}10^{-5}$ cm s^{-1}.

[161]) Units: cm^2 mol^{-1} s^{-1}.

[162]) Units: s^{-1}.

[163]) Value of j_0 depends on selection of part of the current density vs. electrode potential curve.

[164]) Pr = Propyl.

[165]) Value obtained after saturating the electrode with hydrogen for 38 hours.

[166]) Average from cathodic and anodic Tafel lines.

[167]) Experimental conditions were adjusted to keep the working electrode in the passivated state.

[168]) Calculated with the transfer coefficient and the nitrate concentration only as reported.

[169]) Value of j_0 depends on selection of part of the current density vs. electrode potential curve.

[170]) $c_{perchlorate}$ = 3 **M**.

[171]) $c_{sulfate}$ = 1.5 **M**.

[172]) Mercury concentration in amalgam 0.1 %.

[173]) Mercury concentration in amalgam 0.32 m**M** ... 0.1 **M**.

[174]) See list of symbols.

References

43Phi Philbert, G.: J. Chim. Phys. **40** (1943) 157.
47Ran Randles, J.E.B.: Discuss. Faraday Soc. **1** (1947) 11.
47Sal Salt, F.W.: Discuss. Faraday Soc. **1** (1947) 169.
49Boc Bockris, J.O'M., Parsons, R.: Trans. Faraday Soc. **45** (1949) 916.
50Ger Gerischer, H.: Z. Elektrochem. **54** (1950) 366.
51Vet Vetter, K.J.: Z. Phys. Chem. (Leipzig) **196** (1951) 360.
52Con Conway, B.E.: Electrochemical Data, Amsterdam: Elsevier, 1952.
52Ran1 Randles, J.E.B., Somerton, K.W.: Trans. Faraday Soc. **48** (1952) 951.
52Ran2 Randles, J.E.B., Somerton, K.W.: Trans. Faraday Soc. **48** (1952) 937.
53Ger2 Gerischer, H.: Z. Phys. Chem. **202** (1953) 302.
54Hil Hillson, P.J.: Trans. Faraday Soc. **50** (1954) 385.
54Lor Lorenz, W.: Z. Elektrochem. **58** (1954) 912.
55Ber1 Berzins, T., Delahay, P.: J. Am. Chem. Soc. **77** (1955) 6448.
55Ger1 Gerischer, H.: Z. Elektrochem. **59** (1955) 604.
55Jor Jordan, J.: Anal. Chem. **27** (1955) 1708.
55Mat Matsuda, H., Ayabe, Y.: Z. Elektrochem. **59** (1955) 494.
55Mor Morinaga, K.: J. Chem. Soc. Jpn. Pure Chem. Sect. (Nippon Kagaku Zasshi) **76** (1955) 133.
55Vie Vielstich, W., Gerischer, H.: Z. Phys. Chem. N. F. **4** (1955) 10.
56Fro1 Fronaeus, S., Östman, C.O.: Acta Chem. Scand. **10** (1956) 769.
56Fro2 Fronaeus, S.: Acta Chem. Scand. **10** (1956) 1345.
56Ger Gerischer, H., Staubach, K.-E.: Z. Phys. Chem. N. F. **6** (1956) 118.
56Mat Matsuda, H., Ayabe, Y.: Bull. Chem. Soc. Jpn. **29** (1956) 134.
56Mor Morinaga, K.: Bull. Chem. Soc. Jpn. **29** (1956) 793.
57Del Delahay, P.: J. Chim. Phys. **54** (1957) 369.
57Ger1 Gerischer, H., Tischer, R.P.: Z. Elektrochem. **61** (1957) 1159.
57Meh Mehl, W., Bockris, J.O'M.: J. Chem. Phys. **27** (1957) 818.
57Pen Pentland, N., Bockris, J.O'M., Sheldon, E.: J. Electrochem. Soc. **104** (1957) 182.
57Rüe Rüetschi, P., Cahan, B.D.: J. Electrochem. Soc. **104** (1957) 406.
57Vie Vielstich, W., Delahay, P.: J. Am. Chem. Soc. **79** (1957) 1874.
58Bar1 Barker, G.C., Faircloth, R.L., Gardner, A.W.: Nature (London) **181** (1958) 247.
58Bau Bauer, H.H., Elving, P.: Anal. Chem. **30** (1958) 341.
58Del Delahay, P., Trachtenberg, I.: J. Am. Chem. Soc. **80** (1958) 2094.
58Ger Gerischer, H., Krause, M.: Z. Phys. Chem. N. F. **14** (1958) 184.
58Ima Imai, H.: J. Sci. Hiroshima Univ. Ser. A Math. Phys. Chem. **22** (1958) 291.
58Lai Laitinen, H.A., Subcasky, W.J.: J. Am. Chem. Soc. **80** (1958) 2623.
58Lor Lorenz, W.: Z. Phys. Chem. N. F. **17** (1958) 136.
59Boc Bockris, J.O'M., Pentland, N.: Trans. Faraday Soc. **55** (1959) 833.
59Hol Hollnagel, M., Landsberg, R.: Z. Phys. Chem. (Leipzig) **212** (1959) 94.
59Ima Imai, H.: J. Electrochem. Soc. Jpn. **27** (1959) 55.
59Ive Ives, D.J.G.: Can. J. Chem. **37** (1959) 213.
59Kol Kolthoff, I.M., Okinska, Y.: J. Am. Chem. Soc. **81** (1959) 2296.
59Lor Lorenz, W.: Z. Phys. Chem. N. F. **19** (1959) 377.
59Mat1 Mattsson, E., Bockris, J.O'M.: Trans. Faraday Soc. **55** (1959) 1586.
59Mat2 Matsuda, H., Ayabe, Y.: Z. Elektrochem. **63** (1959) 1164.
59Mat3 Matsuda, H., Oka, S., Delahay, P.: J. Am. Chem. Soc. **81** (1959) 5077.
59Ran Randles, J.E.B.: Can. J. Chem. **37** (1959) 238.
59Sch1 Schuldiner, S.: J. Electrochem. Soc. **106** (1959) 891.
59Tam Tamamushi, R., Tanaka, N.: Z. Phys. Chem. N. F. **21** (1959) 89.
60Bau Bauer, H.H., Smith, D.L., Elving, P.J.: J. Am. Chem. Soc. **82** (1960) 2094.
60Del Delahay, P., Kleinerman, M.: J. Am. Chem. Soc. **82** (1960) 4509.

60Gie Gierst, L., Cornelissen, R.: Collect. Czech. Chem. Commun. **25** (1960) 3004.
60Ino Inoue, S., Imai, H.: Bull. Chem. Soc. Jpn. **33** (1960) 149.
60Los1 Loshkarev, M.A., Tomilov, B.I.: Russ. J. Phys. Chem. (English Transl.) **34** (1960) 836; Zh. Fiz. Khim. **34** (1960) 1753.
60Los2 Losev, V.V., Molodov, A.I.: Dokl. Akad. Nauk SSSR **135** (1960) 1432.
60Los3 Losev, V.V., Molodov, A.I.: Dokl. Akad. Nauk SSSR **130** (1960) 111.
60Mou Moussa, A.A., Sammour, H.M.: J. Chem. Soc. (1960) 2151.
60Oki Okinaka, Y., Kolthoff, I.M.: J. Am. Chem. Soc. **82** (1960) 324.
60Slu Sluyters, J.H., Oomen, J.J.C.: Recl. Trav. Chim. Pays Bas **79** (1960) 1101.
60Str Stromberg, A.G., Kartushinskaya, A.I.: Zh. Fiz. Khim. **34** (1960) 1684.
60Wij3 Wijnen, M.D., Smit, W.M.: Recl. Trav. Chim. Pays Bas **79** (1960) 289.
61Ans Anson, F.C.: Anal. Chem. **33** (1961) 939.
61Bar1 Barker G.C.: Trans. Symp. Electrode Proc. Philadelphia 1959, New York: Wiley, 1961, p. 325.
61Bic Bicelli, L.P., La Vecchia, A., Graziano, M.: Wadd Tech. Rep. 61-63, Sept. 1961.
61Bro Brodd, R.J.: J. Res. Natl. Bur. Stand. Sect. A **65** (1961) 275.
61Gav Gavioli, G., Papoff, P.: Ric. Sci. Parte 2 Sez. A **1** (1961) 193.
61Gie Gierst, L.: Trans. Symp. Electrode Proc. Philadelphia 1959, New York: Wiley, 1961, p. 109.
61Gok Gokhshtein, Ya.P., Kao, T.-S.: Zh. Fiz. Khim. **35** (1961) 1699.
61Hur1 Hurlen, T.: Acta Chem. Scand. **15** (1961) 621.
61Hur2 Hurlen, T.: Acta Chem. Scand. **15** (1961) 630.
61Jos Joshi, K.M., Mehl, W., Parsons, R.: Trans. Symp. Electrode Proc. Philadelphia 1959, New York: Wiley, 1961, p. 249.
61Kam Kambara, T., Ishii, T.: Rev. Polarogr. **9** (1961) 30.
61Lor Lorenz, W., Salié, G.: Z. Phys. Chem. (Leipzig) **218** (1961) 259.
61Mei Meites, L., Israel, Y.: J. Am. Chem. Soc. **83** (1961) 4903.
61Mül Müller, W., Lorenz, W.: Z. Phys. Chem. N. F. **27** (1961) 23.
61Pam Pamfilov, A.V., Lopushanskaya, A.I.: Ukr. Khim. Zh. **27** (1961) 598.
61Ran Randles, J.E.B.: Trans. Symp. Electrode Proc. Philadelphia 1959, New York: Wiley, 1961, p. 209.
61Reg Regner, A., Balej, J.: Collect. Czech. Chem. Commun. **26** (1961) 237.
61Riu Rius, A., Tordesillas, I.M., Sacristan, A.: Electrochim. Acta **4** (1961) 62.
61She Sheinin, A.B., Zinov'ev, V.A., Kheifets, V.L.: Zh. Fiz. Khim. **35** (1961) 513.
61Shi Shiratori, H.: Denki Kagaku **29** (1961) 765.
61Smi Smith, D.E.: Ph.D. Thesis, Columbia University, 1961.
61Zin Zinov'ev, V.A., Sheinin, A.B., Kheifets, V.L.: Zh. Fiz. Khim. **35** (1961) 98.
61deM de Mars, R.D.: J. Electrochem. Soc. **108** (1961) 779.
62Aga Agarwal, H.P.: Aust. J. Chem. **15** (1962) 874.
62Bat Baticle, A.M.: C. R. Acad. Sci. (Paris) **254** (1962) 668.
62Beh Behr, B., Dojlido, J., Malyszko, J.: Rocz. Chem. **36** (1962) 725.
62Bon Bonnemay, M., Costa, M.: Electrochim. Acta **7** (1962) CITCE-meeting.
62Del2 Delahay, P.: Anal. Chem. **34** (1962) 1161.
62Del3 Delahay, P., Aramata, A.: J. Phys. Chem. **66** (1962) 2208.
62Del4 Delahay, P.: Progress in Polarography, Vol. 1, New York: Interscience, 1962.
62Gal Galus, Z., Adams, R.N.: Am. Chem. Soc. Meeting, 9.-14.09.1962, Atlantic City, NJ, USA.
62Ima1 Imai, H.: J. Phys. Chem. **66** (1962) 1744.
62Ima2 Imai, H., Delahay, P.: J. Phys. Chem. **66** (1962) 1108.
62Jah Jahn, D., Vielstich, W.: J. Electrochem. Soc. **109** (1962) 849.
62Jor Jordan, J., Javick, R.A.: Electrochim. Acta **6** (1962) 23.
62Kis Kisova, L., Jakesova, M., Fischer, O.: Collect. Czech. Chem. Commun. **27** (1962) 1962.
62Kor Koryta, J.: Electrochim. Acta **6** (1962) 67.
62Los1 Loshkarev, M.A., Tomilov, B.I.: Russ. J. Phys. Chem. (English Transl.) **36** (1962) 66; Zh. Fiz. Khim. **36** (1962) 132.
62Los2 Losev, V.V., Molodov, A.I.: Electrochim. Acta **6** (1962) 81.

62Mat Matsuda, H., Ayabe, Y.: Z. Elektrochem. **66** (1962) 469.
62Ran Randles, J.E.B.: Progress in Polarography 1, New York: Interscience, 1962, p.123.
62Tam Tamamushi, R., Ishibashi, K., Tanaka, N.: Z. Phys. Chem. N. F. **35** (1962) 209.
62Tom Tomilov, B.I., Loshkarev, M.A.: Russ. J. Phys. Chem. (English Transl.) **36** (1962) 1027; Zh. Fiz. Khim. **36** (1962) 1753.
63Ada Adams, R.N.: Rev. Polarogr. **11** (1963) 71.
63Aga1 Agarwal, H.P.: J. Electroanal. Chem. **5** (1963) 236.
63Aga2 Agarwal, H.P.: J. Electroanal. Chem. **5** (1963) 245.
63Aga3 Agarwal, H.P.: Agra Univ. J. Res. Sci. **12** (1963) 93.
63Aga4 Agarwal, H.P.: J. Electrochem. Soc. **110** (1963) 237.
63Bla Blackedge, J., Hush, N.S.: J. Electroanal. Chem. **5** (1963) 435.
63Bud Budov, G.M., Losev, V.V.: Zh. Fiz. Khim. **37** (1963) 1023.
63Fio Fiorani, M., Bombi, G.G., Marchetto, M.T.: Chim. Ind. (Milan) **45** (1963) 796.
63Fuz Fuzisawa, T.: Nippon Kagaku Zasshi (J. Chem. Soc. Jpn. Pure Chem. Sect.) **84** (1963) 629.
63Gal Galus, Z., Adams, R.N.: J. Phys. Chem. **67** (1963) 866.
63Gor Gorodetskii, V.V., Losev, V.V.: Dokl. Akad. Nauk SSSR **151** (1963) 361.
63Gos Gossner, K.: Z. Phys. Chem. N. F. **36** (1963) 392.
63Hal Hale, J.M., Parsons, R.: Trans. Faraday Soc. **59** (1963) 1429.
63Ham Hamelin, A.: C. R. Acad. Sci. (Paris) **257** (1963) 1709.
63Hus Hush, N.S., Blackledge, J.: J. Electroanal. Chem. **5** (1963) 420.
63Koz Kozhevnikova, N.M., Rotinyan, A.L.: Zh. Prikl. Khim. **36** (1963) 1950.
63Los Losev, V.V., Molodov, A.I.: Dokl. Akad. Nauk SSSR **148** (1963) 1114.
63Mac Macero, D.J., Anderson, L.B.: J. Electroanal. Chem. **6** (1963) 221.
63Mat Matsuda, H., Ayabe, Y., Adachi, K.: Ber. Bunsenges. Phys. Chem. **67** (1963) 593.
63Rot Rotinyan, A.L., Kozhevnikova, N.M.: Zh. Fiz. Khim. **37** (1963) 1818.
63Slu Sluyters-Rehbach, M., Sluyters, J.H.: Recl. Trav. Chim. Pays Bas **82** (1963) 535.
63Smi Smith, D.E.: Anal. Chem. **35** (1963) 610.
63Tam Tamamushi, R., Tanaka, N.: Z. Phys. Chem. N. F. **39** (1963) 117.
63Tos Toshima, S., Okinaka, Y., Okaniwa, H.: Denki Kagaku **31** (1963) 854.
63Wol Wolf, D.: J. Electroanal. Chem. **5** (1963) 186.
64Ara Aramata, A., Delahay, P.: J. Phys. Chem. **68** (1964) 880.
64Bau Bauer, H.H.: Aust. J. Chem. **17** (1964) 715.
64Boc3 Bockris, J.O'M., Srinivasan, S.: Electrochim. Acta **9** (1964) 31.
64But1 Butler, J.N., Makrides, A.C.: Trans. Faraday Soc. **60** (1964) 1664.
64Ham1 Hamelin, A.: C. R. Acad. Sci. (Paris), Ser. C **259** (1964) 362.
64Ham2 Hamelin, A.: Electrochim. Acta **9** (1964) 289.
64Joh Johansson, R.: Acta Chem. Scand. **18** (1964) 1809.
64Lai Laitinen, H.A., Eda, K., Nakanishi, M.: Talanta **11** (1964) 321.
64Mak Makrides, A.C.: J. Electrochem. Soc. **111** (1964) 392.
64Mol Molodov, A.I., Losev, V.V.: Zh. Fiz. Khim. **38** (1964) 1481.
64Slu Sluyters-Rehbach, M., Sluyters, J.H.: Recl. Trav. Chim. Pays Bas **83** (1964) 967.
65Bie Biegler, T., Laitinen, H.A.: Anal. Chem. **37** (1965) 572.
65But3 Butler, J.N., Barron-Apps, E.A., Meehan, M.L.: Trans. Faraday Soc. **61** (1965) 2509.
65But4 Butler, J.N., Meehan, M.L.: Trans. Faraday Soc. **61** (1965) 2521.
65But5 Butler, J.N., Dienst, M.: J. Electrochem. Soc. **112** (1965) 226.
65Ijz Ijzermans, A.B., Sluyters, J.H.: Recl. Trav. Chim. Pays Bas **84** (1965) 751.
65Iof Iofa, Z.A., Makorova, V.A.: Elektrokhimiya **1** (1965) 230.
65Koz Kozhevnikova, N.M., Rotinyan, A.L.: Elektrokhimiya **1** (1965) 664.
65Mat Matthews, D.B.: Ph.D. Thesis, University of Pennsylvania, 1965.
65Mol1 Molodov, A.I., Losev, V.V.: Elektrokhimiya **1** (1965) 149.
65Mol2 Molodov, A.I., Losev, V.V.: Elektrokhimiya **1** (1965) 53.
65Mol3 Molodov, A.I., Losev, V.V.: Elektrokhimiya **1** (1965) 651.
65Nic Nicholson, R.S.: Anal. Chem. **37** (1965) 1351.
65Rai Raicheva, S., Andreeva, L.: C. R. Acad. Bulg. Sci. **18** (1965) 1023.

65Sat Sathyanarayana, S.: J. Electroanal. Chem. **10** (1965) 119.
65Sri Srinivasan, V.S., Torsi, G., Delahay, P.: J. Electroanal. Chem. **10** (1965) 165.
65Tos Toshima, S., Okinaka, Y., Okaniwa, H.: Denki Kagaku **33** (1965) 19.
66Bau Bauer, H.H., Foo, D.C.S.: Aust. J. Chem. **19** (1966) 1103.
66But3 Butler, J.N., Meehan,M.L.: Trans. Faraday Soc. **62** (1966) 3524.
66Chr Christie, J.H., Parry, E.P., Osteryoung, R.A.: Electrochim. Acta **11** (1966) 1525.
66Cos Cosovic, B., Branica, M.: J. Polarogr. Soc. **12** (1966) 97.
66Gau Gaur, J.N., Jain, D.S.: Electrochim. Acta **11** (1966) 1661.
66Gor Gorodetskii, V.V., Losev, V.V.: Elektrokhimiya **2** (1966) 656.
66Ham Hamelin, A., Valette, G.: J. Chim. Phys. **63** (1966) 1285.
66Kra1 Kravtsov, V.I., Chamaev, V.N.: Vestn. Leningr. Univ. **4** (1966) 133.
66Kra2 Kravtsov, V.I., Zelensky, M.I.: Elektrokhimiya **2** (1966) 1138.
66Los Losev, V.V., Sribny, L.E., Molodov, A.I.: Elektrokhimiya **2** (1966) 1431.
66Par2 Parsons, R., Passeron, E.: J. Electroanal. Chem. **12** (1966) 524.
66Slu Sluyters-Rehbach, M., Sluyters, J.H.: Electrochim. Acta **11** (1966) 73.
67Bat Baticle, A.M., Perdu, F.: J. Electroanal. Chem. **13** (1967) 364.
67Dir Dirkse, T.P., DeWit, D., Shoemaker, R.: J. Electrochem. Soc. **114** (1967) 1196.
67Ell Elliot, D., Buchanan, G.S.: Anal. Chem. **39** (1967) 1245.
67Far Farr, J.P.G., Hampson, N.A.: J. Electroanal. Chem. **13** (1967) 433.
67Fri Frischmann, J.K., Timnick, A.: Anal. Chem. **39** (1967) 507.
67Gal Galus, Z., Jeftic, L.: J. Electroanal. Chem. **14** (1967) 415.
67Gor Gorodetskii, V.V., Losev, V.V.: Elektrokhimiya **3** (1967) 1192.
67Har Haruyama, S.: Electrochem. Soc. Jpn. **35** (1967) 62.
67Jai Jain, D.S., Gaur, J.N.: Electrochim. Acta **12** (1967) 413.
67Koo1 Kooijman, D.J., Sluyters, J.H.: Electrochim. Acta **12** (1967) 693.
67Koo2 Kooijman, D.J., Sluyters, J.H.: Electrochim. Acta **12** (1967) 1579.
67Los Losev, V.V., Molodov, A.I., Gorodetskii, V.V.: Electrochim. Acta **12** (1967) 475.
67Man Mannan, R.J.: Ph.D. Thesis, University of Pennsylvania, 1967.
67Thi1 Thiele, R., Landsberg, R.: Z. Phys. Chem. **236** (1967) 261.
67Thi2 Thiele, R., Landsberg, R.: Z. Phys. Chem. **236** (1967) 95.
67Tim Timmer, B., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **14** (1967) 181.
67Wei2 Weir, W.D., Enke, C.G.: J. Phys. Chem. **71** (1967) 280.
68Boc2 Bockris, J.O'M., Mannan, R.J., Damjanovich, A.: J. Chem. Phys. **48** (1968) 1898.
68Che Chernenko, V.I., Rysakov, A.A., Pristinskaya, Z.I.: Elektrokhimiya **4** (1968) 519.
68Cla2 Clarke, T.G., Hampson, N.A., Lee, J.B., Morley, J.R., Scanlon, B.: Can. J. Chem. **46** (1968) 3437.
68Gor1 Gorodetskii, V.V., Mishenina, K.A., Losev, V.V., Grimberg, A.M., Ostrovskii, G.M.: Elektrokhimiya **4** (1968) 46.
68Gor2 Gorodetskii, V.V., Fedortsov, L.I., Losev, V.V.: Elektrokhimiya **4** (1968) 967.
68Gos Gossner, K., Mansfeld, F.: Z. Phys. Chem. N. F. **58** (1968) 24.
68Ham2 Hampson, N.A., Larkin, D.: J. Electroanal. Chem. **18** (1968) 401.
68Kov Kovalenko, P.N., Tsyganov, E.M., Estifeev, M.M.: Elektrokhimiya **4** (1968) 828.
68Kra Kravtsov, V.I., Chamaev, V.N.: Vestn. Leningr. Univ. **10** (1968) 80.
68Lar Larionov, O.V., Durdin, Ya.V.: Elektrokhimiya **4** (1968) 630.
68Los Losev, V.V., Pchel'nikov, A.P.: Elektrokhimiya **4** (1968) 264.
68Moo Moorhead, E.D., Frame II, G.M.: J. Phys. Chem. **72** (1968) 3684.
68Nec Nechaev, E.A., Bek, R.Yu., Kudryavstev, N.T.: Elektrokhimiya **4** (1968) 545.
68Pop Popova, L.N., Stromberg, A.G.: Elektrokhimiya **4** (1968) 1147.
68Ran Randles, J.E.B., Whitehouse, D.R.: Trans. Faraday Soc. **64** (1968) 1376.
68Sal Salié, G.: Z. Phys. Chem. (Leipzig) **239** (1968) 411.
68Sto Stonehart, P.: Electrochim. Acta **13** (1968) 1789.
68Tim Timmer, B., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **19** (1968) 73.
68Vij Vijh, A.K.: J. Phys. Chem. **72** (1968) 1148.
68Zut Zutic, V., Branica, M.: J. Electroanal. Chem. **19** (1968) 259.

69Bar2 Bartelt, H., Landazury, S.: J. Electroanal. Chem. **22** (1969) 105.
69Bar3 Bartelt, H., Skilandat, H.: J. Electroanal. Chem. **23** (1969) 407.
69Bro2 Brodd, R.J.: J. Electrochem. Soc. Jpn. **37** (1969) 181.
69Cus Cushing, J.P., Hubbard, A.T.: J. Electroanal. Chem. **23** (1969) 183.
69Dau Daum, P.H., Enke, C.G.: Anal. Chem. **41** (1969) 653.
69Dek Dekker, B.G., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **21** (1969) 137.
69Gal Galizzioli, D.: Thesis, University of Milan, 1969.
69Gor Gorodetskii, V.V., Alenina, A.G., Losev, V.V.: Elektrokhimiya **5** (1969) 227.
69Gri2 Grilikhes, M.S., Krasikov, B.S., Solotskaya, N.S.: Elektrokhimiya **5** (1969) 859.
69Ham3 Hampson, N.A., Larkin, D.: Trans. Faraday Soc. **65** (1969) 1660.
69Nik Niki, K.K., Hackerman, N.: J. Phys. Chem. **73** (1969) 1023.
69Nor Norakidze, I.G., Kazakov, V.A., Vagramyan, A.T.: Elektrokhimiya **5** (1969) 970.
69Pun Punning, K., Past, V.: Uch. Zap. Tartu. Gos. Univ. **235** (1969) 35.
69Suz Suzuki, J.: Bull. Chem. Soc. Jpn. **42** (1969) 3487.
69Tan1 Tanaka, N., Yamada, A.: Electrochim. Acta **14** (1969) 491.
69Tan2 Tanaka, N., Aoki, Y., Yamada, A.: Electrochim. Acta **14** (1969) 1155.
69Tka Tkachik, Z.A., Gorbunova, K.M., Sevastyanov, E.S.: Elektrokhimiya **5** (1969) 1125.
69Wil Wilde, B.E., Hodge, F.G.: Electrochim. Acta **14** (1969) 619.
69Yam Yamada, A., Tanaka, N.: Bull. Chem. Soc. Jpn. **42** (1969) 1600.
69deL de Leeuwe, R., Sluyters-Rehbach, M., Sluyters, J.H.: Electrochim. Acta **14** (1969) 1183.
69van van Leeuwen, H.P., Kooijman, D.J., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **23** (1969) 475.
70Ans Anson, F.C., Rathjen, N., Frisbee, R.F.: J. Electrochem. Soc. **117** (1970) 477.
70Arm Armstrong, R.D., Henderson, M.: J. Electroanal. Chem. **26** (1970) 381.
70Avd Avdeyev, D.K., Aleksandrov, G.V., Durdin, Y.V.: Proc. 2nd Symposium on Double Layer and Adsorption on Solid Electrodes, Tartu, 1970, p. 58.
70Aya Ayabe, Y., Matsuda, H.: Abstracts of the 16th Meeting of the Polarographic Society of Japan, 1970, p. 51.
70Bar3 Bartelt, H., Skilandat, H.: J. Electroanal. Chem. **25** (1970) 79.
70Bar4 Bartelt, H., Skilandat, H.: J. Electroanal. Chem. **24** (1970) 207.
70Bar5 Bartelt, H., Skilandat, H.: J. Electroanal. Chem. **24** (1970) 407.
70Bon Bonewitz, R.A., Schmidt, G.M.: J. Electrochem. Soc. **117** (1970) 1367.
70Bre Breckenridge, J.H., Harris, W.E.: Can. J. Chem. **48** (1970) 1934.
70Cas Casadio, S., Orlandini, F.: J. Electroanal. Chem. **26** (1970) 91.
70Cla Clark, J.T., Hampson, N.A.: J. Electroanal. Chem. **26** (1970) 307.
70Eis Eisenberg, M., Kuppinger, R.E., Wong, K.M.: J. Electrochem. Soc. **117** (1970) 577.
70Gau Gaur, J.N., Goswami, N.K.: Electrochim. Acta **15** (1970) 519.
70Ham Hampson, N.A., Larkin, D.: Anal. Chem. **42** (1970) 1823.
70Her Herbelin, J.M., Andersen, T.N., Eyring, H.: Electrochim. Acta **15** (1970) 1455.
70Hub Hubbard, A.T., Anson, F.C.: Electroanalytical Chemistry, Vol. 4, Bard, A.J. (ed.), New York: Marcel Dekker, 1970.
70Kap Kapoor, R.C., Mathur, K.C.K.: Indian J. Chem. **8** (1970) 168.
70Kin Kinard, W.F., Philp jr., R.H.: J. Electroanal. Chem. **25** (1970) 373.
70Lau Lau, A.L.Y. Hubbard, A.T.: J. Electroanal. Chem. **24** (1970) 237.
70Los Losev, V.V., Pchel'nikov, A.P.: Elektrokhimiya **6** (1970) 41.
70Mak Maki, N., Yamamoto, K.: Bull. Chem. Soc. Jpn. **43** (1970) 2450.
70Mou Moussa, A.A., Abou Romia, M.M., El-Taib Heakal, F.: Electrochim. Acta **15** (1970) 1391.
70Nis Nishihara, C., Matsuda, H.: J. Electroanal. Chem. **28** (1970) 17.
70Rot1 Rotinyan, A.L., Kilimnin, A.B., Levin, E.D.: Proc. 2nd Symposium on Double Layer and Adsorption on Solid Electrodes, Tartu, 1970, p. 321.
70Sal1 Salie, G.: Z. Phys. Chem. (Leipzig) **244** (1970) 1.
70Sal2 Salie, G.: Z. Phys. Chem. (Leipzig) **240** (1970) 421.
70Sch Schleinitz, K.D., Landsberg, R., Lowis of Menar, G.V.: J. Electroanal. Chem. **28** (1970) 287.
70Sha1 Sharma, L.R., Dutt, J.: Indian J. Chem. **8** (1970) 170.

70Suz Suzuki, J.: Bull. Chem. Soc. Jpn. **43** (1970) 755.

70Ten Tenno, T.T., Slet, V.A., Palm, U.V.: Proc. 2nd Symposium on Double Layer and Adsorption on Solid Electrodes, Tartu, 1970, p. 361.

70Tho Thomas, N.T., Nobe, K.: J. Electrochem. Soc. **117** (1970) 622.

70Tim Timmer, B., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **24** (1970) 287.

70Yam Yamada, A., Tanaka, N.: Sci. Rep. Tohoku Univ. Ser.1 **53** (1970) 110.

70Zut Zutic, V., Branica, M.: J. Electroanal. Chem. **28** (1970) 187.

70deK de Kreuk, C.W., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **28** (1970) 391.

70deL de Leeuwe, R., Timmer, B., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **28** (1970) 105.

71Aga Agarwal, H.P.: Electrochim. Acta **16** (1971) 1395.

71Aya Ayabe, Y., Matsuda, H.: Denki Kagaku **39** (1971) 635.

71Bar1 Bartelt, H.: Electrochim. Acta **16** (1971) 307.

71Bar2 Bartelt, H.: Electrochim. Acta **16** (1971) 629.

71Bar3 Bartelt, H., Prügel, M.: Electrochim. Acta **16** (1971) 1815.

71Bar4 Baric, A.: Ph.D. Thesis, University of Zagreb, 1971.

71Bie Biegler, T., Gonzales, E.R., Parsons, R.: Collect. Czech. Chem. Commun. **36** (1971) 414.

71Che Cheh, H.Y., Sard, R.: J. Electrochem. Soc. **118** (1971) 1737.

71Gor Gorodetskii, V.V., Losev, V.V.: Elektrokhimiya **7** (1971) 631.

71Ham2 Hampson, N.A., Lee, J.B., Morley, J.R.: Electrochim. Acta **16** (1971) 637.

71Iss Issa, I.M., Ghoneim, M.M., El Samahy, A.A., Tharwat, M.: Electrochim. Acta **16** (1971) 1655.

71Lau Lau, A.L.Y., Hubbard, A.T.: J. Electroanal. Chem. **33** (1971) 77.

71Nik Niki, K., Okuda, Y., Tomonari, T., Buck, E., Hackerman, N.: Electrochim. Acta **16** (1971) 487.

71Ruz Ruzic, I., Baric, A., Branica, M.: J. Electroanal. Chem. **29** (1971) 411.

71Sad Sadek, H., Issa, R.M., Abd-El-Nabey, B.A.: Electrochim. Acta **16** (1971) 401.

71Sip Sipos, L., Jeftic, Lj., Branica, M., Galus, Z.: J. Electroanal. Chem. **32** (1971) 35.

71deK de Kreuk, C.W., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **33** (1971) 267.

72Ang Angell, D.H., Dickinson, T.: J. Electroanal. Chem. **35** (1972) 55.

72Bek Bek, R.Yu., Pautov, V.N., Lifshits, A.S.: Izv. Sib. Otd. Akad. Nauk SSSR Ser. Khim. Nauk (1972) 22.

72Haw Hawkridge jr., F.M., Bauer, H.H.: Anal. Chem. **44** (1972) 364.

72Hue Huebert, B.J., Smith, D.E.: Anal. Chem. **44** (1972) 1179.

72Iss1 Issa, I.M., Tharwat, M.: Electrochim. Acta **17** (1972) 343.

72Iss2 Issa, I.M., Tharwat, M., Elewady, Y.A.: Electrochim. Acta **17** (1972) 1065.

72Iss3 Issa, I.M., Ghoneim, M.M., El Samahy, A.A., Tharwat, M.: Electrochim. Acta **17** (1972) 1251.

72Lai Lai, C.-H., Hubbard, A.T.: Inorg. Chem. **11** (1972) 2081.

72Ree Reeves, R.M., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **34** (1972) 69.

72Roh Rohko, T., Kogoma, M., Aoyagui, S.: J. Electroanal. Chem. **38** (1972) 45.

72Stu Stulikova, M., Vydra, F.: J. Electroanal. Chem. **38** (1972) 349.

72Tan Tanaka, N., Tomita, T., Yamada, A.: Bull. Chem. Soc. Jpn. **45** (1972) 940.

73Cap Capon, A., Parsons, R.: J. Electroanal. Chem. **46** (1973) 215.

73Gal Galizzioli, D., Trasatti, S.: J. Electroanal. Chem. **44** (1973) 367.

73Pol van der Pol, F., Sluyters-Rehbach, M., Sluyters, J.H.: J. Electroanal. Chem. **45** (1973) 377.

73Rod Rodgers, R.S., Anson, F.C.: J. Electroanal. Chem. **42** (1973) 381.

74Aga Agarwal, H.P., Qureshi, S.: Electrochim. Acta **19** (1974) 607.

75Saj1 Saji, T., Aoyagui, S.: J. Electroanal. Chem. **63** (1975) 31.

75Saj2 Saji, T., Yamada, T., Aoyagui, S.: J. Electroanal. Chem. **61** (1975) 147.

76Wea Weaver, M.J., Anson, F.C.: Inorg. Chem. **15** (1976) 1871.

78Ber Bernstein, C., Heindrichs, A., Vielstich, W.: J. Electroanal. Chem. **87** (1978) 81.

78Eve Everett, K.G., Drew, L.A., Ericson, S.J., Schmid, G.M.: J. Electrochem. Soc. **125** (1978) 389.

78Sat Satterberg, T.L., Weaver, M.J.: J. Phys. Chem. **82** (1978) 1784.

78Web Weber, J., Samec, Z., Marecek, V.: J. Electroanal. Chem. **89** (1978) 271.

80Bar1 Barr, S., Guyer, K., Weaver, M.J.: J. Electroanal. Chem. **111** (1980) 41.

80Bar3 Baranski, A.S., Fawcett, W.R.: J. Chem. Soc. Faraday Trans. 1 **76** (1980) 1962.

80Ber Bernstein, C., Vielstich, W.: Proc. 3rd Symp. Electrode Processes, Boston 1979, The Electrochemical Society (ed.), Princeton: The Electrochemical Society, 1980, p. 350.

83Iwa Iwasita, T., Schmickler, W., Herrmann, J., Vielstich, W.: J. Electrochem. Soc. **130** (1983) 2026.

83Vie Vielstich, W.: Electrochemical Society Meeting, San Francisco, 1983, Ext. Abstr. 815.

84Bal Baldwin, R.P., Ravichandran, K., Johnson, R.K.: J. Chem. Educ. **61** (1984) 820.

84Bar2 Barr, S.W., Weaver, M.J.: Inorg. Chem. **23** (1984) 1657.

84Gro Gross, M., Jordan, J.: Pure Appl. Chem. **56** (1984) 1095.

84Guy Guyer, K.L., Weaver, M.J.: Inorg. Chem. **23** (1984) 1664.

84Li Li, T.-T., Liu, H.Y., Weaver, M.J.: J. Am. Chem. Soc. **106** (1984) 1233.

85Dur Durliat, H., Delorme, P., Comtat, M.: Electrochim. Acta **30** (1985) 1071.

85Gen Gennett, T., Milner, D.F., Weaver, M.J.: J. Phys. Chem. **89** (1985) 2787.

85Iwa Iwasita, T., Schmickler, W., Schultze, J.W.: Ber. Bunsenges. Phys. Chem. **89** (1985) 138.

85Li Li, T.T.-T., Weaver, M.J.: Inorg. Chem. **24** (1985) 1882.

85Opa Opallo, M., Kapturkiewicz, A.: Electrochim. Acta **30** (1985) 1301.

86Ham Hamelin, A., Weaver, M.J.: J. Electroanal. Chem. **209** (1986) 109.

86McM McManis, G.E., Golovin, M.N., Weaver, M.J.: J. Phys. Chem. **90** (1986) 6563.

86Opa Opallo, M.: J. Chem. Soc. Faraday Trans. 1 **82** (1986) 339.

87Hun Hung, N.C., Nagy, Z.: J. Electrochem. Soc. **134** (1987) 2215.

87Zha Zhang, X., Yang, H., Bard, A.J.: J. Am. Chem. Soc. **109** (1987) 1916.

88Boc Bochmann, H.G., Vielstich, W.: Electrochim. Acta **33** (1988) 805.

88Wip Wipf, D.O., Kristensen, E.W., Deakin, M.R., Wightman, R.M.: Anal. Chem. **60** (1988) 306.

89Chu Chu, D., Tryk, D., Gervasio, D., Yeager, E.B.: J. Electroanal. Chem. **272** (1989) 277.

89Rod Rodes, A., Feliu, J., Aldaz, A.: J. Electroanal. Chem. **271** (1989) 127.

90Gra Grampp, G., Kapturkiewicz, A., Jaenicke, W.: Ber. Bunsenges. Phys. Chem. **94** (1990) 439.

91Bar2 Baranski, A.S., Winkler, K., Fawcett, W.R.: J. Electroanal. Chem. **313** (1991) 367.

91Cha Chang, S.-C., Lau, S.-Y., Schardt, B.C., Weaver, M.J.: J. Phys. Chem. **95** (1991) 4787.

91Cur Curtiss, L.A., Halley, J.W., Hautmann, J., Hung, N.C., Nagy, Z., Rhee, Y.-J.: J. Electrochem. Soc. **138** (1991) 2032.

91Kan Kaneko, H., Nozaki, K., Wada, Y., Aoki, T., Negishi, A., Kamimoto, M.: Electrochim. Acta **36** (1991) 1191.

92Faw1 Fawcett, W.R., Opallo, M.: J. Phys. Chem. **96** (1992) 2920.

92Faw2 Fawcett, W.R., Opallo, M.: J. Electroanal. Chem. **331** (1992) 815.

92Saf Safford, L.K., Weaver, M.J.: J. Electroanal. Chem. **331** (1992) 857.

93Faw Fawcett, W.R., Fedurco, M.: J. Phys. Chem. **97** (1993) 7075.

93Gao Gao, Z.Q., Bobacka, J., Ivaska, A.: Electrochim. Acta **38** (1993) 379.

93Mu Mu, X.H., Schultz, F.A.: J. Electroanal. Chem. **353** (1993) 349.

93Win Winkler, K., Baranski, A.: J. Electroanal. Chem. **346** (1993) 197.

95Ale Alehashem, S., Chambers, F. J.W., Strojek, J.W., Swain, G.M., Ramesham, R.: Anal. Chem. **67** (1995) 2812.

95Che Chen, P., Fryling, M.A., McCreery, R.L.: Anal. Chem. **67** (1995) 3115.

95Hec1 Hecht, M., Fawcett, W.R.: J. Electroanal. Chem. **396** (1995) 473

95Hec2 Hecht, M., Fawcett, W.R.: J. Phys. Chem. **99** (1995) 1311.

95Urb Urbanczyk, A., Debek, E., Kalinowski, M.K.: J. Electroanal. Chem. **389** (1995) 141.

96Boc Bockris, J.O'M., Kim, J.: J. Electrochem. Soc. **143** (1996) 3801.

96Mar Martin, H.B., Argoitia, A., Landau, U., Anderson, A.B., Angus, J.C.: J. Electrochem. Soc. **143** (1996) L133.

96Tor Torres, L.M., Gil, A.F., Galicia, L., Gonzalez, I.: J. Chem. Educ. **73** (1996) 808.

96Win Winkler, K., Baranski, A.S., Fawcett, W.R.: J. Chem. Soc. Faraday Trans. **92** (1996) 3899.

97Faw Fawcett, W.R.: Electrochim. Acta **42** (1997) 833.

97Krs Krstajic, N.V., Grgur, B.N., Mladenovic, N.S., Vojnovic, M.V., Jaksic, M.M.: Electrochim. Acta **42** (1997) 323.

97Mar Markovic, N.M., Grgur, B.N., Ross, P.N.: J. Phys. Chem. B **101** (1997) 5405.
97Per Perez, J., Villullas, H.M., Gonzalez, E.R.: J. Electroanal. Chem. **435** (1997) 179.
98Evs2 Evstefeeva, Yu.E., Krotova, M.D., Pleskov, Yu.V., Mazin, V.M., Elkin, V.V., Mishuk, V.Y., Varnin, V.P., Teremetskaya, I.G.: Sov. Electrochem. (English Transl.) **34** (1998) 1352; Elektrokhimiya **34** (1998) 1171.
98Kit Kitamura, F., Nanbu, N., Ohsaka, T., Tokuda, K.: J. Electroanal. Chem. **456** (1998) 113.
98Per Perez, J., Gonzalez, E.R., Villullas, H.M.: J. Phys. Chem. B **102** (1998) 10931.
98Ple Pleskov, Y.V., Evstefeeva, Y.E., Krotova, M.D., Elkin, V.V., Mazin, V.M., Mishuk, V.Y., Varnin, V.P., Teremetskaya, I.G.: J. Electroanal. Chem. **455** (1998) 139.
99Evs Evstefeeva, Yu.E., Krotova, M.D., Pleskov, Y.V., Laptev, V.A.: Sov. Electrochem. (English Transl.) **35** (1999) 132; Elektrokhimiya **35** (1999) 138.
00Gra Granger, M.C., Witek, M., Xu, J.S., Wang, J., Hupert, M., Hanks, A., Koppang, M.D., Butler, J.E., Lucazeau, G.: Anal. Chem. **72** (2000) 3793.
00Jak Jaksic, J.M., Vojnovic, M.V., Krstajic, N.V.: Electrochim. Acta **45** (2000) 4151.
01Liu Liu, X., Liu, K., Cheng, G., Dong, S.: J. Electroanal. Chem. **510** (2001) 103.
02Fer1 Ferro, S., De Battisti, A.: Electrochim. Acta **47** (2002) 1641.
02Fer2 Ferro, S., De Battisti, A.: Phys. Chem. Chem. Phys. **4** (2002) 1915.
02Ste Stein, T., Lang, H., Holze, R.: J. Electroanal. Chem. **520** (2002) 163.
03Ort Ortiz, M.E., Núñez-Vergara, L.J., Squella, J.A.: J. Electroanal. Chem. **549** (2003) 157.

Table 5.2. Exchange current densities and rate constants in aqueous systems at miscellaneous surfaces.

Electrode	Solution	Educt, product, concentration [1])	T [°C]	j_0, j_{00} or k_c^θ [2]) [A cm^{-2} or cm^{-1}]	Method [3])	Ref.
Bronzes						
$Na_{0.9}WO_3$	0.5 **M** H_2SO_4/H_2O	$Fe(CN)_6^{3-/4-}$	RT	29	CP	69Voj
$Na_{0.7}WO_3$	0.5 **M** H_2SO_4/H_2O	$Fe(CN)_6^{3-/4-}$	RT	23	CP	69Voj
$Na_{0.5}WO_3$	0.5 **M** H_2SO_4/H_2O	$Fe(CN)_6^{3-/4-}$	RT	15	CP	69Voj
Lead oxide (PbO_2)	5 **M** H_2SO_4/H_2O	$Pb^{4+/2+}$	25	$3.2\cdot10^{-4}$	IP	68Agu
Tin oxide (SnO_2), In-doped	0.24 **M** cacodylic acid/H_2O, pH=7	horse heart cytochrome c (r), 38 µ**M**	22	*$4.9\cdot10^{-3}$*	CV	82Bow [4])
	0.24 **M** cacodylic acid/H_2O, pH=7	horse heart cytochrome c (r), 73 µ**M**	22	*$1.7\cdot10^{-3}$*	CV	82Bow
	0.24 **M** cacodylic acid/H_2O, pH=7	horse heart cytochrome c (r), 73 µ**M**	22	*$1\cdot10^{-4}$*	DCVA	82Bow
	0.24 **M** cacodylic acid/H_2O, pH=7	horse heart cytochrome c (r), 300 µ**M**	22	*$8.1\cdot10^{-4}$*	CV	82Bow
$LaSrMnO_3$	solid electrode	O_2 (r), p=1 atm	600	$4.2\cdot10^{-3}$	CV; IP	05Co
YSZ [5])	solid electrode	O_2 (r), p=1 atm	700	$3.7\cdot10^{-3}$	CV; IP	05Co
(ZrO_2+Y_2O_3)	solid electrode	O_2 (r), p=1 atm	800	$1.8\cdot10^{-3}$	CV; IP	05Co
	solid electrode	O_2 (r), p=1 atm	900	$4.4\cdot10^{-3}$	CV; IP	05Co

[1]) The educt/reacting species for the charge transfer reaction to which j_0 is referring to is stated only if it is not already given by the composition of the electrolyte solution. If the concentrations of the reactants are given in addition the exchange current density j_0 will be stated subsequently. The arrangement of species identifies the reaction as a reduction or an oxidation. In the absence of specific designations of the reaction product (e.g. reactions involving complex organometallic compounds) the type of reaction is designated (r) or (o) indicating a reduction or oxidation.

[2]) If no concentration of the educt is given the standard exchange current density j_{00} is stated. Values of k_c^θ are printed in italics.

[3]) CP = current-potential curves; CV = cyclic voltammetry; DCVA = derivative cyclic voltabsorptometry; IP = impedance method. See also list of symbols.

[4]) The unexpected dependence of the rate constant on the reactant concentration was tentatively assigned to traces of impurities.

[5]) YSZ = Y_2O_3-stabilized ZrO_2

References

68Agu Aguf, I.A.: Elektrokhimiya **4** (1968) 1130.

69Voj Vojnovic, M.V., Sepa, D.B.: J. Chem. Phys. **51** (1969) 5344.

82Bow Bowden, E.F., Bancroft, E.E., Blount, H.N., Chlebowski, J.F., Hawkridge, F.M., Thorpe, C.: J. Am. Chem. Soc. **104** (1982) 7641.

05Co Co, A.C., Xia, S.J., Birss, V.I.: J. Electrochem. Soc. **152** (2005) A570.

Table 5.3. Exchange current densities in molten salt electrolyte systems.

Electrode [1])	Electrolyte	Educt, product, concentration [2])	T [°C]	j_0, j_{00} or k_c^{θ} [3]) [A cm^{-2} or cm s^{-1}]	Method [9])	Ref.
Aluminum (Al)	66mol%$AlCl_3$ +20mol%NaCl +14mol%KCl	Al^{3+}, (r)	217	0.86		71Hof
Bismuth (Bi)	eutectic of LiCl+KCl	Bi^{2+}	450	8	PS	60Lai2
	eutectic of LiCl+KCl	$Bi^{0/3+}$	450	8	PS	61Lai
	eutectic of LiCl+KCl	$Bi^{3+/0}$	450	*$5.7\cdot10^{-3}$*	PS	70Bek
Cadmium (Cd)$_{(Hg)}$	eutectic of LiCl+KCl	Cd^{2+}	450	210	GD	60Lai2
	eutectic of $LiNO_3$, $NaNO_3$, KNO_3+0.02 mol% KCl	Cd^{2+}	156	*0.53*	FR	58Bar1
	eutectic of LiCl+KCl	Cd^{2+}	450	210	PS	60Lai2
	eutectic of LiCl+KCl	Cd^{2+}	450	210	GD, PS	61Lai
	eutectic of LiCl+KCl	Cd^{2+}	450	*$1.8\cdot10^{-3}$*	IP	70Bek
Carbon	eutectic of LiCl+KCl	ClO_2^+/ClO_2	450	*0.052*	CV	84Gro
(C)	NaCl-$AlCl_3$	Cl^-/Cl_2	175	$8.6\cdot10^{-6}$	CV	84Gro
Lead (Hg)	molten KCl+NaCl	$Pb^{2+/0}$	800	*$3.1\cdot10^{-4}$*	IP	70Bek
	molten KCl+NaCl	$Pb^{2+/0}$	450	*$1.2\cdot10^{-3}$*	IP	70Bek
	eutectic of LiCl+KCl	$Pb^{2+/0}$	450	*0.01*	GD	60Lai2
	eutectic of LiCl+KCl	$Pb^{2+/0}$, 1 **M**	450	30	GD	61Lai
	eutectic of $LiNO_3$, $NaNO_3$, KNO_3+0.02 mol% KCl	$Pb^{2+/0}$	156	*1.4*	FR	58Bar1
Manganese (Mn)	41.5% CaO+51.8% Al_2O_3 +6.7% MgO +0.22% MnO	$Mn^{2+/0}$	1450	0.298	IP	68Bul
	41.5% CaO+51.8% Al_2O_3 +6.7% MgO +0.31% MnO	$Mn^{2+/0}$	1450	0.334	IP	68Bul
	41.5% CaO+51.8% Al_2O_3 +6.7% MgO +0.52% MnO	$Mn^{2+/0}$	1450	0.238	IP	68Bul
	41.5% CaO+51.8% Al_2O_3 +6.7% MgO +0.52% MnO	$Mn^{2+/0}$	1450	0.436	IP	68Bul
	41.5% CaO+51.8% Al_2O_3 +6.7% MgO +0.52% MnO	$Mn^{2+/0}$	1450	0.890	IP	68Bul
	41.5% CaO+51.8% Al_2O_3 +6.7% MgO +0.58% MnO	$Mn^{2+/0}$	1450	0.415	IP	68Bul
	41.5% CaO+51.8% Al_2O_3 +6.7% MgO+0.8% MnO	$Mn^{2+/0}$	1450	0.550	IP	68Bul

Electrode [1])	Electrolyte	Educt, product, concentration [2])	T [°C]	j_0, j_{00} or k_c^{θ} [3]) [A cm^{-2} or cm s^{-1}]	Method [9])	Ref.
Mn	41.5% CaO+51.8% Al_2O_3 +6.7% MgO +1.43% MnO	$Mn^{2+/0}$	1450	0.650	IP	68Bul
	41.5% CaO+51.8% Al_2O_3 +6.7% MgO +1.72% MnO	$Mn^{2+/0}$	1450	0.695	IP	68Bul
Mercury (Hg)	27% $LiNO_3$ +18% $NaNO_3$ +55% KNO_3	$Ni^{2+/0}$	140	*0.0042*	IP	55Ran
Nickel (Ni)	eutectic of LiCl+KCl	$Ni^{2+/0}$	450	110	GD	60Lai2
	eutectic of LiCl+KCl	$Ni^{2+/0}$	450	110	GD	61Lai
(Ni)$_{(Hg)}$	27% $LiNO_3$ +18% $NaNO_3$+55% KNO_3	$Ni^{2+/0}$	140	*0.0062*	IP	55Ran
Platinum (Pt)	eutectic of LiCl+KCl	Pb^{2+}	450	172	IP	57Lai
	eutectic of LiCl+KCl	Ni^{2+}	450	172	IP	57Lai
	eutectic of LiCl+KCl	Co^{2+}	450	172	IP	57Lai
	eutectic of LiCl+KCl	$Cd^{2+/0}$	450	31.8	IP	57Lai
	eutectic of LiCl+KCl	$Zn^{2+/0}$	450	31.8	IP	57Lai
	eutectic of LiCl+KCl	$Pt^{2+/0}$	450	40	GD	60Lai2
	eutectic of LiCl+KCl	$Pt^{2+/0}$	450	40	GD	61Lai
	eutectic of LiCl+KCl	$Pt^{2+/0}$	450	*$1.32\cdot10^{-2}$*	IP [4])	72Hil1
	eutectic of LiCl+KCl	$Pt^{2+/0}$	450	*$1.76\cdot10^{-2}$*	IP [5])	72Hil1
	eutectic of LiCl+KCl	$Pt^{2+/0}$	450	*$\sim3\cdot10^{-2}$*	FR	72Hil2
	eutectic of LiCl+KCl	$Pt^{2+/0}$	450	*$\sim3\cdot10^{-2}$*	FD	72Hil2
	eutectic of LiCl+KCl	$Pt^{2+/0}$	450	*$1.5\cdot10^{-2}$*	IP	72Hil2
	eutectic of LiCl+KCl	$Pt^{2+/0}$	450	*$1.3\cdot10^{-2}$*	IP	72Hil2 [4])
	eutectic of LiCl+KCl	$Pt^{2+/0}$	450	*$1.7\cdot10^{-2}$*	IP	72Hil2 [5])
	eutectic of LiCl+KCl	$Pt^{2+/0}$	450	*$1.3\cdot10^{-2}$*	IP	72Hil2 [6])
	eutectic of LiCl+KCl	$Pt^{2+/0}$	450	*$1.7\cdot10^{-2}$*	IP	72Hil2 [7])
	eutectic of LiCl+KCl	$Pt^{2+/0}$	450	*$1.49\cdot10^{-2}$*	CS	72Hil3 [8])
	eutectic of LiCl+KCl	$Pt^{2+/0}$	450	*$1.3\cdot10^{-2}$*	IP	72Hil3
	eutectic of LiCl+KCl	$V^{3+/2+}$	450	30	PS	60Lai2 [4])
Silver (Ag)	eutectic of LiCl+KCl	Ag^{+}	450	190	GD	60Lai2
	eutectic of LiCl+KCl	Ag/Ag^{+}	410	*$7\cdot10^{-4}$*	IP	59Hil
	27wt% $LiNO_3$ +18wt% $NaNO_3$ +55wt% KNO_3	Ag/Ag^{+}	120	*$2\cdot10^{-2}$*	IP	59Hil
	eutectic of LiCl+KCl	Ag^{+}	450	*$1.35\cdot10^{-2}$*	IP	72Hil1 [4])
	eutectic of LiCl+KCl	Ag^{+}	450	*$2.04\cdot10^{-2}$*	IP	72Hil1 [5])
	eutectic of LiCl+KCl	Ag^{+}	450	*~0.1*	FR	72Hil2
	eutectic of LiCl+KCl	Ag^{+}	450	*~0.1*	FD	72Hil2
	eutectic of LiCl+KCl	Ag^{+}	450	*0.17*	IP	72Hil2
	eutectic of LiCl+KCl	Ag^{+}	450	*0.13*	IP	72Hil2 [4])
	eutectic of LiCl+KCl	Ag^{+}	450	*$0.02\cdot10^{-2}$*	IP	72Hil2 [5])

Electrode [1])	Electrolyte	Educt, product, concentration [2])	T [°C]	j_0, j_{00} or k_c^θ [3]) [A cm^{-2} or cm s^{-1}]	Method [9])	Ref.
	eutectic of LiCl+KCl	Ag^+	450	*0.19·10^{-2}*	IP	72Hil2 [6])
	eutectic of LiCl+KCl	Ag^+	450	*0.67·10^{-2}*	IP	72Hil2 [7])
	eutectic of LiCl+KCl	Ag^+	450	*~0.115*	CS	72Hil3
Titanium (Ti)	eutectic of LiCl+KCl	Ti^{2+}/Ti^{3+}	400	*0.9*	IP	59Hil
Vanadium (V)	eutectic of LiCl+KCl	$V^{3+/2+}$	450	30	PS	60Lai2
Yttrium (Y)	eutectic of LiCl+KCl	Zn^{2+}	450	*0.3*	GD	60Lai2
Zinc (Zn)	eutectic of LiCl+KCl	Zn^{2+}	450	*0.4*	PS	60Lai2
	eutectic of LiCl+KCl	Zn^{2+}	450	150	GD, PS	61Lai
	eutectic of LiCl+KCl	Zn^{2+}	450	*0.19*	IP	70Bek

[1]) The term $_{(Hg)}$ denotes an amalgam electrode.

[2]) The educt/reacting species for the charge transfer reaction to which j_0 is referring to is stated only if it is not already given by the composition of the electrolyte solution. If the concentration of the educt is given in addition the exchange current density j_0 will be stated.

[3]) If no concentration of the educt is given the standard exchange current density j_{00} is stated. Values referring to k_c^θ are marked (k).

[4]) The electrode was a flat microelectrode.

[5]) The electrode was a hemispherical microelectrode.

[6]) Flat microelectrode, complex plane analysis.

[7]) Hemispherical microelectrode, complex plane analysis.

[8]) Wire microelectrode.

[9]) CS = coulostatic method, CV = cyclic voltamogram, FD = Faradaic distortion method, FR = Faradaic rectification, GD = galvanostatic double pulse method, IP = impedance method, PS = potential step method. See also list of symbols.

References

55Ran Randles, J.E.B., White, W.: Z. Elektrochem. **59** (1955) 666.

57Lai Laitinen, H.A., Gaur, H.C.: J. Electrochem. Soc. **104** (1957) 730.

58Bar1 Barker, G.C., Faircloth, R.L., Gardner, A.W.: Nature (London) **181** (1958) 247.

59Hil Hill, D.L., Hills, G.J., Young, L., Bockris, J. O'M.: J. Electroanal. Chem. **1** (1959/60) 79.

60Lai2 Laitinen, H.A., Tischer, R.P., Roe, D.K.: J. Electrochem. Soc. **107** (1960) 546.

61Lai Laitinen, H.A., Tischer, R., Roe, D.K.: Trans. Symp. Electrode Proc. Philadelphia 1959, New York: Wiley, 1961, p. 185.

68Bul Buler, P.I., Nikitin, Yu.P.: Elektrokhimiya **4** (1968) 330.

70Bek Bek, R.Yu., Lifshits, A.S.: Izv. Sib. Otd. Akad. Nauk SSSR, Ser. Khim. Nauk (1970) 53.

71Hof Hoff, H.: Electrochim. Acta **16** (1971) 1357.

72Hil1 Hill, D.L., Fung, K.W., Hsieh, S., Lee, H.L.: J. Electroanal. Chem. **38** (1972) 57.

72Hil2 Hill, D.L., Fung, K.W., Lee, H.L.: J. Electroanal. Chem. **38** (1972) 67.

72Hil3 Hill, D.L., Fung, K.W., Lee, H.L.: J. Electroanal. Chem. **38** (1972) 75.

84Gro Gross, M., Jordan, J.: Pure Appl. Chem. **56** (1984) 1095.